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Photocrystallography of $[\text{Ru}(\text{bpy})_2(\text{dmsO})_2]^{2+}$ reveals an O-bonded metastable state†Zoe Y. Marr,^a Rajani Thapa Magar,^b Bertrand Fournier,^{cd} Jason B. Benedict^{id}*^a and Jeffrey J. Rack^{id}*^b

We report the first instance of observing the phototriggered isomerization of dmsO ligands on a bis sulfoxide complex, $[\text{Ru}(\text{bpy})_2(\text{dmsO})_2]$, in the crystalline solid state. The solid-state UV-vis spectrum of the crystal demonstrates an increase in optical density around 550 nm after irradiation, which is consistent with the solution isomerization results. Digital images of the crystal before and after irradiation display a notable color change (pale orange to red) and cleavage occurs along planes, $(\bar{1}01)$ and (100) , during irradiation. Single crystal X-ray diffraction data also confirms that isomerization is occurring throughout the lattice and a structure that contains a mix of the *S,S* and *O,O/S,O* isomer was attained from a crystal irradiated *ex situ*. *In situ* irradiation XRD studies reveal that the percentage of the O-bonded isomer increases as a function of 405 nm exposure time.

Introduction

Photocrystallography has emerged as an indispensable tool for the structural characterization of metastable states of photochromic compounds.¹ Irradiation of single crystals of these compounds or complexes yields a mixture of ground state and metastable state isomers. While complete conversion of a single crystal can be achieved, it is much more common to observe partial conversion, and thus the isomerization is modelled as a disorder in the crystal. Success of this experiment requires that the solid-state photochemistry mimic the solution photochemistry. Moreover, crystal engineering to manage density, moduli, molecular structural changes, and heat dissipation is often required for photocrystallography to be successful. Indeed, irradiation of many crystalline materials can yield explosions (termed photosalient),^{2–4} dramatic cracking or cleaving of planes,^{5,6} and even disintegration of the crystal.⁷ In stark contrast, other crystalline samples simply change color upon irradiation without any significant or substantial photochemistry beyond the surface of the crystal, revealing a complicated relationship between crystal structure and molecular

photochemistry. Nevertheless, photocrystallography has yielded the molecular structures of metastable states of a wide range of transition metal complexes with NO^+ ,^{8–11} N_2 ,¹² SO_2 ,^{13–15} NO_2^- (ref. 16–18) ligands as well as diarylethenes,¹⁹ fulgides, spiro-compounds, and has recently been employed to observe reaction intermediates in C–H bond activation.^{20–22} We have created a large family of photochromic ruthenium sulfoxides by joining the electrochemistry of $[\text{Ru}(\text{NH}_3)_5(\text{dmsO})]^{2+}$ (dmsO is dimethylsulfoxide) with the photochemistry of $[\text{Ru}(\text{bpy})_3]^{2+}$ (bpy is 2,2'-bipyridine).^{23–25} In 1982, Henry Taube first proposed an ECEC (Electrochemical Chemical–Electrochemical Chemical) mechanism to explain the peculiar cyclic voltammogram of $[\text{Ru}(\text{NH}_3)_5(\text{dmsO})]^{2+}$, where oxidation of Ru^{2+} to produce Ru^{3+} triggers intramolecular S → O isomerization of the sulfoxide, and reduction of Ru^{3+} to yield Ru^{2+} prompts intramolecular O → S isomerization.²⁶ While the ground state S-bonded isomer is well characterized by infrared spectroscopy and single crystal X-ray diffraction,^{27,28} evidence for the formation of the isomerization product was provided by comparison of the O-bonded $\text{Ru}^{3+/2+}$ coupled with other O-bonded coordination complexes and from its reactivity with exogenous ligands. Despite claims of X-ray data of the O-bonded isomer in $[\text{Ru}(\text{NH}_3)_5(\text{dmsO})]^{3+}$, the O-bonded structure produced from isomerization to our knowledge has not been reported.^{29,30}

The MLCT (Metal-to-Ligand Charge-Transfer) transition in ruthenium polypyridine complexes yields an excited state best formulated as an oxidized ruthenium ion and a reduced bipyridine ligand (e.g., $[\text{Ru}(\text{bpy})_2(\text{bpy}^-)]^{2+}$).³¹ The phototriggered S → O isomerization in ruthenium polypyridine sulfoxides is activated by the formation of Ru^{3+} in the MLCT state of these heteroleptic complexes (Scheme 1). Characterization of the metastable state of these photochromic complexes has included

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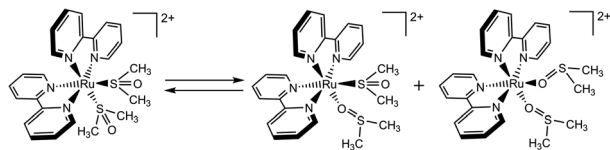
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Scheme 1 Proposed isomerization products of S,S -[Ru(bpy) $_2$ (dmsO) $_2$] $^{2+}$ to yield a mixture of O,O bonded and S,O bonded isomers upon visible light irradiation.

^1H NMR spectroscopy,³² infrared spectroscopy,^{33,34} DFT calculations,^{35–41} and X-ray absorption spectroscopy.⁴² While these techniques reflect molecular structure, none of these data provide the molecular structure of the metastable state. While there are many structures of S-bonded or O-bonded sulfoxides, we do not know of any X-ray diffraction study that characterizes the conversion of an S-bonded sulfoxide to an O-bonded sulfoxide.^{43,44} Herein, we report the structure of the metastable state of a photochromic ruthenium sulfoxide complex and show that the structure is unequivocally an O-bonded sulfoxide (Scheme 1).

Results and discussion

We chose to investigate [Ru(bpy) $_2$ (dmsO) $_2$][B(Ph) $_4$] $_2$ as the large size, flexibility, weak basicity, and lack of functional groups would increase the likelihood of observing isomerization of the dmsO ligands in the single crystalline phase. As expected, the tetraphenylborate salt exhibits a limited number of short contacts to the dmsO. For dmsO1 (containing S1), and dmsO2 (containing S2; Fig. 1, left), three short contacts to the anion were observed (contacts not shown); however, for dmsO2 two of these contacts were to a symmetry related dmsO2.

Our previous investigations of bis sulfoxides have revealed that solvent and the nature of the chelate (if any) can be important in the quantum yield of isomerization and product distribution (one sulfoxide isomerization vs. two sulfoxide

isomerizations) following visible light irradiation.²² These studies have demonstrated the presence of separate one isomerization (S,O -product isomer) and two isomerization (O,O -product isomer) reaction pathways. We are interested in capturing the mechanistic details of isomerization within a crystalline environment and correlating these structural changes with electronic spectroscopic results within photochromic crystals.

Previous reports establish that the S,S -, S,O - and O,O -isomers each feature unique absorption maxima in solution. Thus, irradiation of solutions of this complex at much lower concentrations than those present in a crystal show a distinct darkening of the solution from colorless or pale yellow to a deep burgundy or purple color.

The solid-state UV-visible spectra of the crystal during irradiation demonstrates similar features to that seen in the solution data (Fig. S1†). Prior to irradiation, the crystal of the S,S -isomer absorbs at ~ 460 nm. As the crystal is irradiated with 405 nm light, a new broad absorption feature appears at ~ 550 nm, which is indicative of the formation of the O,O -isomer. These changes in the optical density of the crystal are consistent with the solution isomerization results and are suggestive of isomerization. Upon careful examination of the crystals during irradiation, the crystal not only demonstrates dramatic changes in color (Fig. 2 and S2;† consistent with solution observations), but fracture lines along $(\bar{1}01)$ are also

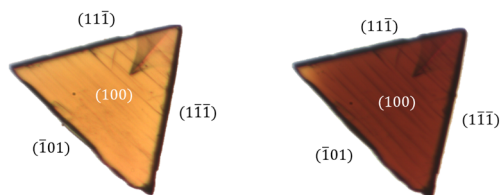


Fig. 2 Digital images of the irradiated crystal before (left) and after 600 seconds of irradiation with 405 nm light (right).

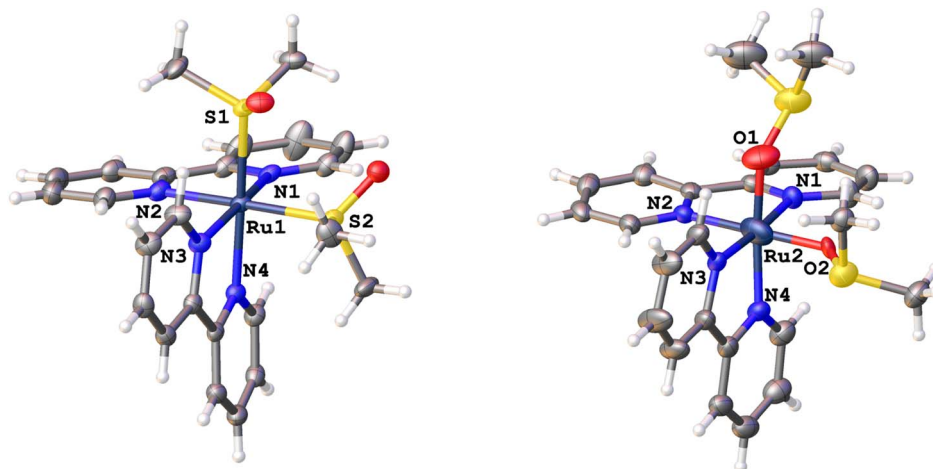


Fig. 1 Ground state structure of bis-[Ru(bpy) $_2$ (dmsO) $_2$] $^{2+}$ (left). Isolated molecular structure of bis-[O,O]-[Ru(bpy) $_2$ (dmsO) $_2$] $^{2+}$ after irradiation with 405 nm light (right). Ruthenium, navy; nitrogen, blue; oxygen, red; sulfur, yellow; carbon, grey; hydrogen, white.

evident. The emergence of fracture planes in the crystal indicate strain in the lattice presumably caused by the stress associated with the proposed sulfoxide isomerization (deep coloration) in the crystal. This is supported by the fact that the dmsoligands align along (101) (see Fig. S3†). If the crystal is irradiated with 405 nm light for a prolonged time (>30 minutes), delamination occurs at the surface seemingly parallel to (100) (Fig. S4†).

Single-crystal X-ray diffraction data of $[\text{Ru}(\text{bpy})_2(\text{dmsol})_2][\text{B}(\text{Ph})_4]_2$ study corroborates that the dmsoligands isomerize within the lattice. Photodifference maps are a useful tool for the visualization of light-induced electron density distribution changes in photocrystallography experiments.⁴⁵ A photodifference map calculated using the experimental intensities from the initial and final data collections (details in ESI†) reveals regions of positive and negative density changes near the dmsoligands consistent with $\text{S} \rightarrow \text{O}$ isomerization (Fig. 3). Red surfaces, corresponding to negative changes in electron density, are observed near the initial ground state positions of the sulfur, oxygen and carbon atoms of both dmsoligands indicating a reduction of electron density at these positions. As expected, the approximate volume of these surfaces tracks with atom size ($\text{S} > \text{O} > \text{C}$). The green surfaces near the dmsoligands, corresponding to positive changes, are observed in close proximity to the ground-state Ru–S and S–C bonds. These regions are consistent with the approximate new positions of the sulphur and oxygen atoms following light-induced S–O isomerization and served as the basis for the subsequent refinement of the position and orientation of the dmsol groups.

Shown in Fig. 1 is the ground state structure of $[\text{S,S}]-[\text{Ru}(\text{bpy})_2(\text{dmsol})_2][\text{B}(\text{Ph})_4]_2$ (tetraphenylborate anion not shown) prior to laser irradiation. After *ex situ* irradiation with 405 nm light, the crystal structure revealed a mixture of S,S - and O,O -isomers. When modelling the disorder for partly isomerized structure (S,O -isomer), the dmsol groups and the Ru

were each refined independently. The dmsol *trans* to N4 contains 10.0% in the O bonded conformation whereas the dmsol *cis* to N4 contains 6.0%. The observation of disorder in both dmsoligands clearly indicates that both ligands isomerize, though given the differences in isomerization for both sulfoxides both S,O and O,O products are likely formed in the crystal lattice. The structure of the O,O -isomer is displayed in Fig. 1, right. This structure clearly features two O-bonded dmsoligands with a tetrahedral geometry about the sulfur, indicating that the sulfur has not changed oxidation state. We note that the unit cell volume of the crystal is essentially unchanged between the two structures ($\Delta V \sim 0.13\%$; $V_{\text{S,S}} = 6139.9(4) \text{ \AA}^3$; $V_{\text{O,O}} = 6148.3(4) \text{ \AA}^3$; Table S1†). In addition to the new coordination mode of dmsol, there are other changes in the new O,O -bonded structure relative to the initial S,S -bonded complex.

Table 1 summarizes the structural parameters of each isomer. Notably, there is a distinctive shortening of the bond between the Ru and the dmsol from $\sim 2.3 \text{ \AA}$ to $\sim 1.8 \text{ \AA}$, which is consistent with isomerization from Ru–S to Ru–O coordination modes. We find Ru–S bond distances of 2.291 and 2.285 \AA , which are consistent with that found in other salts of $[\text{Ru}(\text{bpy})_2(\text{dmsol})_2]^{2+}$ as well as related bis(sulfoxides).^{46–48} Moreover, Calligaris has categorized geometric and infrared data for a wide variety of metal sulfoxide complexes in both S- and O-bonded coordination modes. He reports that the average Ru(II)–S bond distance is 2.260(2) \AA from 474 independent ruthenium sulfoxide structures, which is in accord with our present findings.⁴³ Whereas an average Ru(II)–O bond distance of 2.10(1) \AA is obtained from statistical analysis of 291 independent structures. The bond distances of the Ru(II)–O bond reported here differ slightly from literature, which is likely a consequence of modelling low occupancy disorder. For example, the Ru–O bond distances in *cis*- $[\text{Ru}(\text{bpy})_2(\text{OH}_2)_2]^{2+}$ are 2.147(4) and 2.152(4) \AA .⁴⁹ In $[\text{Ru}(\text{bpy})_2(\text{ox})]^{2+}$ (ox is oxalate), the

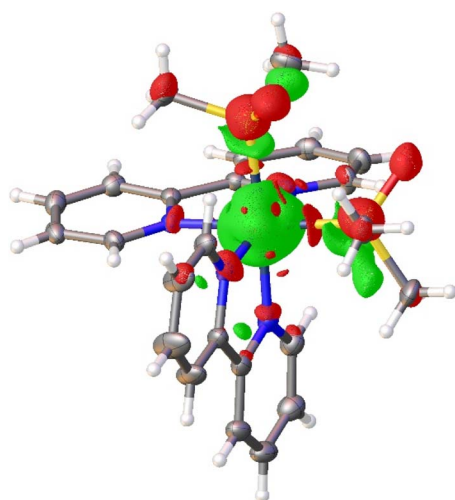


Fig. 3 Photodifference map of the Ru–dmsol complex with iso-meshes of 0.766 e \AA^{-3} showing regions of positive and negative density changes near the dmsoligands. The positive and negative electron density iso-meshes are represented in green and red respectively.

Table 1 Crystallographically determined bond lengths and angles for the ground state and metastable state isomers

<u>S,S-Bonded isomer</u>		<u>O,O-Bonded isomer</u>	
Bond	Distance (\AA)	Bond	Distance (\AA)
Ru1–S1	2.291	Ru2–O1	1.813
Ru1–S2	2.285	Ru2–O2	1.803
Ru1–N1	2.073	Ru2–N1	2.028
Ru1–N2	2.099	Ru2–N2	1.983
Ru1–N3	2.080	Ru2–N3	2.099
Ru1–N4	2.083	Ru2–N4	1.883
S1–O1	1.481	S1–O1	1.561
S2–O2	1.482	S2–O2	1.533
Bond angle	Angle ($^\circ$)	Bond angle	Angle ($^\circ$)
N2–Ru1–N4	86.763 $^\circ$	N2–Ru2–N4	96.23 $^\circ$
S1–Ru1–N2	91.186 $^\circ$	O1–Ru2–N2	90.569 $^\circ$
S2–R1–N4	94.207 $^\circ$	O2–Ru2–N4	81.314 $^\circ$
S1–Ru1–S2	88.067 $^\circ$	O1–Ru2–O2	92.514 $^\circ$
Ru1–S1–O1	116.518 $^\circ$	Ru2–O1–S1	136.729 $^\circ$
Ru1–S2–O2	115.338 $^\circ$	Ru2–O2–S2	126.280 $^\circ$

Ru–O bond distances are 2.087(7) and 2.093(6) Å.⁵⁰ Lastly, Ru–O bond distances of 2.050(8) and 2.030(8) Å were found in [Ru(bpy)₂(DBSQ)]²⁺, where DBSQ is 3,5-di-*tert*-butyl-1,2-semiquinato.⁵¹

The S–O bond distances for the ground state *S,S*-isomer (1.481 Å and 1.482 Å) are in accord with many existing structures (1.489(1) Å) and the S–O bond distances for the *O,O*-isomer are similar (1.561 Å and 1.533 Å) to what is typically observed (1.545(3) Å).^{43,44} The S–C bond distances for a Ru(II) S-bonded and O-bonded sulfoxide are 1.784(1) and 1.783(5) Å, respectively. The ground state structure reveals commensurate distances (1.768–1.786 Å) while the S–C bond distances for the O-bonded metastable state are close to the expected value (ranging from 1.659 to 1.804 Å). Likewise, the Ru–S–O angle (116.5° and 115.3°) is in good agreement with the literature data (Ru–S–O: 117.6°), while the Ru–O–S angles (136.7° and 126.2°) in the metastable state are just slightly wider than that outlined by Calligaris (122.0°).

There is also a distortion of the geometry about Ru between the ground state and metastable state structures. Specifically, the N2–Ru1–N4 angle is 86.763° for the *S,S*-isomer, while this angle is 96.23° for the *O,O*-isomer (Fig. 1). Similarly, the S1–Ru1–S2 angle for the *S,S*-isomer is 88.067°, whereas the corresponding angle in the *O,O*-isomer (O1–Ru2–O2) is 92.514°. We are in the process of preparing new complexes to define these distances and angles more precisely of the metastable state. Nevertheless, the diffraction data unequivocally show changes in the coordination mode of dmsol following irradiation of the crystal.

We performed an *in situ* SC-XRD data collection study to determine if the extent of dmsol isomerization was related to the total exposure to visible light (405 nm). Five complete data sets were collected sequentially on the same crystal during continuous irradiation (405 nm), and then the crystal structure was solved after each run to determine the percent of the O-bonded isomer as a function of total time irradiated. It was difficult to reproducibly observe the isomerization at cryogenic temperatures, so these data sets were collected at room temperature unlike the *ex situ* studies described above. This led to lower quality data sets; however, relevant conclusions may be drawn from this experiment. Chiefly, we confirmed that the concentration of the bis O-bonded isomer increased as a function of total light exposure (Fig. S5†). From the first data set to the final data set, the percent of the O-bonded dmsol gradually increased from ~10.1% to ~12.2% and ~7.7% to ~10.4% for dmsol1 and dmsol2, respectively. Although XRD is unable to determine the precise distribution of S and O bonded isomers in the lattice (*e.g.*, whether it is a mix of *S,S* and *O,O* or a mix of *S,O/O,S*), the dmsol ligands refined independently and only differ in occupancies of the S- and O-bonded isomer by ~3% throughout the entire data collection. This suggests that the isomerization of the dmsol ligands is correlated with light exposure. The plateau observed is consistent with achieving a photostationary state within the crystalline environment.

Consistent with solution measurements, the isomerization is thermally reversible. When a dark red crystal containing a mixture of isomers (~10% and 6%) was isolated in the dark

for a week, the crystal reverted to a pale orange color, characteristic of the S-bonded isomer (Fig. S6†). SC-XRD measurements on this crystal (in ESI†) revealed conversion back to the ground state structure. Subsequent refinement showed only 2.6% of dmsol1 remained O-bonded, whereas dmsol2 displayed no evidence of the O-bonded isomer.

Conclusions

In aggregate, the data presented here establish that the metastable product from irradiation of photochromic [Ru(bpy)₂(-dmsol)₂][BPh₄]₂ is an O-bonded sulfoxide. This result is consistent with the proposal first outlined by Taube in 1982. This study is enabled by the incorporation of large, flexible, weakly coordinating anion within the salt. Future studies will focus on revealing the mechanistic details of the isomerization by photocrystallography and transient absorption spectroscopy in this and related complexes.

Data availability

Crystallographic data for the structures reported herein have been deposited as CCDC 2226309–2226311. Additional experimental details and data are provided in the ESI,† including a full description of synthetic procedures, crystallographic modelling techniques, UV-visible spectra of single crystals, and images showing changes to a single crystal upon irradiation.

Author contributions

ZYM, RTM, JBB, and JJR were responsible for conceptualization. ZYM, RTM, and BF were responsible for the investigation. ZYM and BF for formal analysis. JBB and JJR for funding acquisition. All authors participated in writing – review and editing.

Conflicts of interest

There are no conflicts to declare.

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Notes and references

- 1 P. Coppens, I. Novozhilova and A. Kovalevsky, Photoinduced linkage isomers of transition-metal nitrosyl compounds and related complexes, *Chem. Rev.*, 2002, **102**, 861–883.
- 2 R. Medishetty, S. C. Sahoo, C. E. Mulijanto, P. Naumov and J. J. Vittal, Photosensitive Behavior of Photoreactive Crystals, *Chem. Mater.*, 2015, **27**, 1821–1829.

- 3 P. Naumov, S. Chizhik, M. K. Panda, N. K. Nath and E. Boldyreva, Mechanically Responsive Molecular Crystals, *Chem. Rev.*, 2015, **115**, 12440–12490.
- 4 P. Naumov, D. P. Karothu, E. Ahmed, L. Catalano, P. Commins, J. M. Halabi, M. B. Al-Handawi and L. Li, The Rise of the Dynamic Crystals, *J. Am. Chem. Soc.*, 2020, **142**, 13256–13272.
- 5 K. R. Chalek, X. N. Dong, F. Tong, R. A. Kudla, L. Y. Zhu, A. D. Gill, W. W. Xu, C. Yang, J. D. Hartman, A. Magalhaes, R. O. Al-Kaysi, R. C. Hayward, R. J. Hooley, G. J. O. Beran, C. J. Bardeen and L. J. Mueller, Bridging photochemistry and photomechanics with NMR crystallography: the molecular basis for the macroscopic expansion of an anthracene ester nanorod, *Chem. Sci.*, 2021, **12**, 453–463.
- 6 T. Kim, L. Y. Zhu, R. O. Al-Kaysi and C. J. Bardeen, Organic Photomechanical Materials, *ChemPhysChem*, 2014, **15**, 400–414.
- 7 D. J. Shields, D. P. Karothu, K. Sambath, I. Ranaweera, S. Schramm, A. Duncan, B. Duncan, J. A. Krause, A. D. Gudmundsdottir and P. Naumov, Cracking under Internal Pressure: Photodynamic Behavior of Vinyl Azide Crystals through N₂ Release, *J. Am. Chem. Soc.*, 2020, **142**, 18565–18575.
- 8 M. D. Carducci, M. R. Pressprich and P. Coppens, Diffraction studies of photoexcited crystals: Metastable nitrosyl-linkage isomers of sodium nitroprusside, *J. Am. Chem. Soc.*, 1997, **119**, 2669–2678.
- 9 A. Y. Kovalevsky, G. King, K. A. Bagley and P. Coppens, Photoinduced oxygen transfer and double-linkage isomerism in a cis-(NO)(NO₂) transition-metal complex by photocystallography, FT-IR spectroscopy and DFT calculations, *Chem.–Eur. J.*, 2005, **11**, 7254–7264.
- 10 A. A. Mikhailov, E. Wenger, G. A. Kostin and D. Schaniel, Room-Temperature Photogeneration of Nitrosyl Linkage Isomers in Ruthenium Nitrosyl Complexes, *Chem.–Eur. J.*, 2019, **25**, 7569–7574.
- 11 G. A. Kostin, A. O. Borodin, A. A. Mikhailov, N. V. Kuratieva, B. A. Kolesov, D. P. Pishchur, T. Woike and D. Schaniel, Photocystallographic, Spectroscopic, and Calorimetric Analysis of Light-Induced Linkage NO Isomers in RuNO(NO₂)(2)(pyridine)(2)OH, *Eur. J. Inorg. Chem.*, 2015, 4905–4913, DOI: [10.1002/ejic.201500702](https://doi.org/10.1002/ejic.201500702).
- 12 D. V. Fomitchev, I. Novozhilova and P. Coppens, Photo-induced linkage isomerism of transition metal nitrosyl and dinitrogen complexes studied by photocystallographic techniques, *Tetrahedron*, 2000, **56**, 6813–6820.
- 13 J. M. Cole, J. d. J. Velazquez-Garcia, D. J. Gosztola, S. G. Wang and Y.-S. Chen, Eta(2)-SO₂ Linkage Photoisomer of an Osmium Coordination Complex, *Inorg. Chem.*, 2018, **57**, 2673–2677.
- 14 J. M. Cole, J. d. J. Velazquez-Garcia, D. J. Gosztola, S. G. Wang and Y.-S. Chen, Light-Induced Macroscopic Peeling of Single Crystal Driven by Photoisomeric Nano-Optical Switching, *Chem. Mater.*, 2019, **31**, 4927–4935.
- 15 S. O. Sylvester, J. M. Cole and P. G. Waddell, Photoconversion Bonding Mechanism in Ruthenium Sulfur Dioxide Linkage Photoisomers Revealed by in Situ Diffraction, *J. Am. Chem. Soc.*, 2012, **134**, 11860–11863.
- 16 L. E. Hatcher, J. Christensen, M. L. Hamilton, J. Trincao, D. R. Allan, M. R. Warren, I. P. Clarke, M. Towrie, S. Fuertes, C. C. Wilson, C. H. Woodall and P. R. Raithby, Steady-State and Pseudo-Steady-State Photocystallographic Studies on Linkage Isomers of Ni(Et(4)dien)(n(2)-O,ON)(n(1)-NO₂): Identification of a New Linkage Isomer, *Chem.–Eur. J.*, 2014, **20**, 3128–3134.
- 17 L. E. Hatcher and P. R. Raithby, Dynamic single-crystal diffraction studies using synchrotron radiation, *Coord. Chem. Rev.*, 2014, **277**, 69–79.
- 18 M. R. Warren, S. K. Brayshaw, A. L. Johnson, S. Schiffers, P. R. Raithby, T. L. Easun, M. W. George, J. E. Warren and S. J. Teat, Reversible 100% Linkage Isomerization in a Single-Crystal to Single-Crystal Transformation: Photocystallographic Identification of the Metastable Ni(dppe)(eta(1)-ONO)Cl Isomer, *Angew. Chem., Int. Ed.*, 2009, **48**, 5711–5714.
- 19 J. M. Cox, I. M. Walton, D. G. Patel, M. Xu, Y.-S. Chen and J. B. Benedict, The Temperature Dependent Photoswitching of a Classic Diarylethene Monitored by in Situ X-ray Diffraction, *J. Phys. Chem. A*, 2015, **119**, 884–888.
- 20 M. I. Gonzalez, D. Gygi, Y. Qin, Q. Zhu, E. J. Johnson, Y.-S. Chen and D. G. Nocera, Taming the Chlorine Radical: Enforcing Steric Control over Chlorine-Radical-Mediated C-H Activation, *J. Am. Chem. Soc.*, 2022, **144**, 1464–1472.
- 21 D. Gygi, M. I. Gonzalez, S. J. Hwang, K. T. Xia, Y. Qin, E. J. Johnson, F. Gygi, Y.-S. Chen and D. G. Nocera, Capturing the Complete Reaction Profile of a C-H Bond Activation, *J. Am. Chem. Soc.*, 2021, **143**, 6060–6064.
- 22 D. C. Powers, S. J. Hwang, B. L. Anderson, H. Yang, S.-L. Zheng, Y.-S. Chen, T. R. Cook, F. P. Gabbai and D. G. Nocera, Stereoelectronic Effects in Cl-2 Elimination from Binuclear Pt(III) Complexes, *Inorg. Chem.*, 2016, **55**, 11815–11820.
- 23 A. W. King, L. Wang and J. J. Rack, Excited State Dynamics and Isomerization in Ruthenium Sulfoxide Complexes, *Acc. Chem. Res.*, 2015, **48**, 1115–1122.
- 24 J. J. Rack, Electron transfer triggered sulfoxide isomerization in ruthenium and osmium complexes, *Coord. Chem. Rev.*, 2009, **253**, 78–85.
- 25 S. B. Vittardi, R. T. Magar, D. J. Breen and J. J. Rack, A Future Perspective on Phototriggered Isomerizations of Transition Metal Sulfoxides and Related Complexes, *J. Am. Chem. Soc.*, 2021, **143**, 526–537.
- 26 A. Yeh, N. Scott and H. Taube, S to O and O to S Linkage Isomerization in Sulfoxide Complexes of Pentaammine-Ruthenium, *Inorg. Chem.*, 1982, **21**, 2542–2545.
- 27 F. C. March and G. Ferguson, Crystal Structure of Dimethylsulfoxidepentaammineruthenium(II) Hexafluorophosphate, *Can. J. Chem.*, 1971, **49**, 3590–3595.
- 28 C. V. Senoff, E. Maslowski and R. G. Goel, Dimethylsulfoxidepentaammineruthenium(II) Hexafluorophosphate - Spectroscopic Study, *Can. J. Chem.*, 1971, **49**, 3585–3589.

- 29 A. Tomita and M. Sano, Linkage Isomerizations of (Sulfoxide)Ammineruthenium Complexes Induced by Electrochemical Processes, *Inorg. Chem.*, 1994, **33**, 5825–5830.
- 30 A. Tomita and M. Sano, Preparations and electrochemical properties of pyrazine-bridged ruthenium-binuclear complexes exhibiting molecular hysteresis, *Inorg. Chem.*, 2000, **39**, 200–205.
- 31 E. M. Kober and T. J. Meyer, Concerning the Absorption-Spectra of the Ions $M(\text{Bpy})_3^{2+}$ ($M=\text{Fe}, \text{Ru}, \text{Os}$, $\text{Bpy}=2,2'$ -Bipyridine), *Inorg. Chem.*, 1982, **21**, 3967–3977.
- 32 D. P. Butcher Jr, A. A. Rachford, J. L. Petersen and J. J. Rack, Phototriggered S \rightarrow O isomerization of a ruthenium-bound chelating sulfoxide, *Inorg. Chem.*, 2006, **45**, 9178–9180.
- 33 C. S. Keating, B. A. McClure, J. J. Rack and I. V. Rubtsov, Sulfoxide stretching mode as a structural reporter via dual-frequency two-dimensional infrared spectroscopy, *J. Chem. Phys.*, 2010, **133**, 144513.
- 34 C. S. Keating, B. A. McClure, J. J. Rack and I. V. Rubtsov, Mode Coupling Pattern Changes Drastically Upon Photoisomerization in Ru-II Complex, *J. Phys. Chem. C*, 2010, **114**, 16740–16745.
- 35 D. A. Lutterman, A. A. Rachford, J. J. Rack and C. Turro, Theoretical Insight on the S \rightarrow O Photoisomerization of DMSO Complexes of Ru(II), *J. Phys. Chem. A*, 2009, **113**, 11002–11006.
- 36 D. A. Lutterman, A. A. Rachford, J. J. Rack and C. Turro, Electronic and Steric Effects on the Photoisomerization of Dimethylsulfoxide Complexes of Ru(II) Containing Picolinate, *J. Phys. Chem. Lett.*, 2010, **1**, 3371–3375.
- 37 A. J. Goettle, I. M. Dixon, F. Alary, J.-L. Heully and M. Boggio-Pasqua, Adiabatic Versus Nonadiabatic Photoisomerization in Photochromic Ruthenium Sulfoxide Complexes: A Mechanistic Picture from Density Functional Theory Calculations, *J. Am. Chem. Soc.*, 2011, **133**, 9172–9174.
- 38 I. Ciofini, C. A. Daul and C. Adamo, Phototriggered linkage isomerization in ruthenium -dimethylsulfoxide complexes: insights from theory, *J. Phys. Chem. A*, 2003, **107**, 11182–11190.
- 39 H. Li, L. Zhang, X. Fan and Y. Zhao, Theoretical Studies on the Redox-Stimulated Isomerization in Electrochromic Osmium Sulfoxide Complexes, *J. Phys. Chem. A*, 2015, **119**, 4244–4251.
- 40 H. Li, L. Zhang, H. Lin and X. Fan, A DFT-D study on the electronic and photophysical properties of ruthenium (II) complex with a chelating sulfoxide group, *Chem. Phys. Lett.*, 2014, **604**, 10–14.
- 41 H. Li, L. Zhang, M. Ye and X. Fan, A Comparative Study of the Redox-Induced Linkage Isomerization of Ruthenium and Osmium Complexes, *Eur. J. Inorg. Chem.*, 2015, 5074–5080, DOI: [10.1002/ejic.201500666](https://doi.org/10.1002/ejic.201500666).
- 42 A. A. Cordones, J. H. Lee, K. Hong, H. Cho, K. Garg, M. Boggio-Pasqua, J. J. Rack, N. Huse, R. W. Schoenlein and T. K. Kim, Transient metal-centered states mediate isomerization of a photochromic ruthenium sulfoxide complex, *Nat. Commun.*, 2018, **9**, 1989.
- 43 M. Calligaris, Structure and bonding in metal sulfoxide complexes: an update, *Coord. Chem. Rev.*, 2004, **248**, 351–375.
- 44 M. Calligaris and O. Carugo, Structure and bonding in metal sulfoxide complexes, *Coord. Chem. Rev.*, 1996, **153**, 83–154.
- 45 B. Fournier and P. Coppens, On the assessment of time-resolved diffraction results, *Acta Crystallogr., Sect. A: Found. Adv.*, 2014, **70**, 291–299.
- 46 M. K. Smith, J. A. Gibson, C. G. Young, J. A. Broomhead, P. C. Junk and F. R. Keene, Photoinduced Ligand Isomerization in Dimethyl Sulfoxide Complexes of Ruthenium(II), *Eur. J. Inorg. Chem.*, 2000, **2000**, 1365–1370.
- 47 B. A. McClure and J. J. Rack, Two-Color Reversible Switching in a Photochromic Ruthenium Sulfoxide Complex, *Angew. Chem., Int. Ed.*, 2009, **48**, 8556–8558.
- 48 N. V. Mockus, D. Rabinovich, J. L. Petersen and J. J. Rack, Femtosecond Isomerization in a Photochromic Molecular Switch, *Angew. Chem., Int. Ed.*, 2008, **47**, 1458–1461.
- 49 M. G. Sauer, E. Tfouni, R. H. d. A. Santos, M. T. d. P. Gambardella, M. P. F. M. Del Lama, L. F. Guimarães and R. S. da Silva, Use of HPLC in the identification of cis and trans-diaquabis(2,2'-bipyridine) ruthenium(II) complexes: crystal structure of cis-[Ru(H₂O)₂(bpy)₂](PF₆)₂, *Inorg. Chem. Commun.*, 2003, **6**, 864–868.
- 50 H.-Y. Shen, W.-M. Bu, D.-Z. Liao, Z.-H. Jiang, S.-P. Yan and G.-L. Wang, A new three-dimensional ruthenium(II) complex via hydrogen bonds: Ru(bpy)₂(ox)·4H₂O (bpy=2,2'-bipyridine, ox=oxalate ion), *Inorg. Chem. Commun.*, 1999, **2**, 615–618.
- 51 S. R. Boone and C. G. Pierpont, Charge delocalization in ruthenium-quinone complexes. Structural characterization of bis(bipyridine)(3,5-di-tert-butylsemiquinonato) ruthenium(II) perchlorate and trans-bis(4-tert-butylpyridine)bis(3,5-di-tert-butylquinone)ruthenium, *Inorg. Chem.*, 1987, **26**, 1769–1773.