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Removal of per- and polyfluoroalkyl substances (PFAS) from water by ceric(IV) ammonium nitrate†

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Ceric(IV) ammonium nitrate (CAN) in aqueous medium acts as an excellent precipitating agent for perfluorooctanesulfonic acid (PFOS). The Ce(IV) center plays a crucial role. Interestingly, Ce(III) chloride showed much less effectiveness under similar conditions. The efficacy of CAN was reduced upon changing the substrate to perfluorooctanoic acid (PFOA).

Per- and polyfluoroalkyl substances (PFAS) are toxic and xenobiotic compounds that have been widely used in fire extinguishers, carpet guards, paper, and non-stick cookware.¹ However, these groups of organofluorines are highly resistant to degradation and are not easily separable by the usual water purification techniques.^{2–4} The environmental and health hazards caused by PFAS contamination in drinking water and ground water have stimulated research for innovative strategies to remove PFASs.⁵ An additional risk while developing these removal techniques is the toxicity of PFASs even at very low concentrations.^{6,7} Easily available chemicals that can efficiently degrade/precipitate out low concentrations of PFAS from water can provide a futuristic design to develop valuable water purification materials.

The large C–F bond dissociation energy ($\sim 485 \text{ kJ mol}^{-1}$) makes it resistant to oxidation and reduction, and together with simultaneous hydro- and oleophobicity of PFAS, make the overall degradation/separation process challenging.^{8,9} Methods to degrade/remove PFAS from water include *in situ* generated radicals, heat treatment, photocatalysis, or strong reducing or oxidizing agents.^{10–12} There are also important reports on PFAS remediation techniques using carbonaceous nanomaterials that are high surface area sorbents or In_2O_3 type, which possess oxygen vacancies in the monocrystalline structure.^{13–15}

There are also some reports on the removal of PFASs directly by powdered-activated carbon (PAC) and granular-activated carbon (GAC) adsorption processes.^{16,17} The conventional coagulation processes (e.g., by iron chloride, alum *etc.*) are usually inefficient at low concentrations of PFAS. They are slow, produce a lot of sludge,

and are usually expensive.¹⁸ A natural coagulant *Moringa oleifera* seeds has been reported to be a better alternative and produces less sludge.¹⁹ Nevertheless, the effectiveness at low concentration of PFAS remains as a challenge.

Herein, we report a new approach for precipitation of the commercially available PFOS (as potassium salt) from water ($< 20 \mu\text{M}$) using ceric(IV) ammonium nitrate (CAN). This study was initiated from the perspective of generating *in situ* high valent/mixed valent oxide species/nanoparticles from FeCl_2 and CoCl_2 that can oxidatively degrade PFAS. In preliminary studies, CAN was used as an oxidant and resulted in significant reduction in the concentration of PFAS of the supernatant aqueous solution whenever CAN was used. Thereafter a range of concentrations of CAN (from 0.094 mM to 0.75 mM) were tested in consecutive experiments keeping PFOS concentration unchanged at $15 \mu\text{M}$ (Fig. S3†). After the treatment with 0.38 mM concentration CAN, the supernatant shows $\sim 80\%$ removal of PFOS (Fig. 1, column 1 and 2). Very slow formation of an off-white precipitate was seen from the solution after seven days of undisturbed standing.

Two 1st row transition metal salts [CoCl_2 and FeCl_2] and also a combination of these salts with CAN [CAN : CoCl_2 (1 : 1 equivalent); CAN : FeCl_2 (1 : 1 equivalent)] were tested under identical conditions. In all these cases, very little ($\sim 15\text{--}20\%$) diminution of PFOS concentration (in LC-MS) was observed in the absence of CAN (Fig. 1, column 4–7). Since CAN itself was very reactive and either pure Fe and Co salts or a 1 : 1 mixture appeared to be much less effective, higher ratios of CAN : CoCl_2 or CAN : FeCl_2 were not tested further.¶ The lower efficacy of these 1 : 1 combinations can be explained by possible partial use of CAN to oxidize Fe(II) and Co(II).²⁰

To check whether Ce^{IV} is playing a crucial role, another Ce^{IV} containing commercially available and strong oxidant,

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§ ICP-MS measurement was not successful due to the presence of the excess nitrate (from ammonium cerium(IV) nitrate) in the medium.

¶ The higher oxidative activity of Ce(IV) is well known at high concentration and in acidic medium.



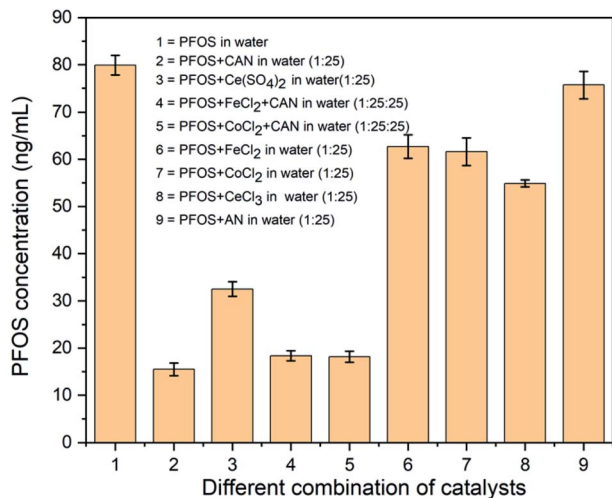


Fig. 1 (1) The LC-MS detected concentrations of 15 μM of PFOS in water before treatment, (2) after treatment with 0.38 mM CAN ($[(\text{NH}_4)_2\text{Ce}^{\text{IV}}(\text{NO}_3)_6]$), (3) 0.38 mM $\text{Ce}^{\text{IV}}(\text{SO}_4)_2$, (4) 0.38 mM CAN + 0.38 mM FeCl_2 , (5) 0.38 mM CAN + 0.38 mM CoCl_2 , (6) 0.38 mM FeCl_2 , (7) 0.38 mM CoCl_2 , (8) 0.38 mM CeCl_3 , and (9) 0.38 mM AN (NH_4NO_3) (all experiments were repeated in triplicate and the dilution details are provided in the ESI†).

$\text{Ce}^{\text{IV}}(\text{SO}_4)_2$ (0.38 mM) was tested and showed $\sim 60\%$ disappearance of PFOS under identical conditions (Fig. 1, column 1 and 3). To find if this is a general trend for any oxidation state of cerium, the impact of $\text{Ce}^{\text{III}}\text{Cl}_3$ (commercially available) was also investigated. It clearly showed significantly lower efficiency in removing PFOS (at best $\sim 33\%$) (Fig. 1 and column 8) from the aqueous medium. No precipitate was found even after two weeks of undisturbed standing. Observing the noteworthy reactivity difference between CAN and CeCl_3 , another set of experiments was performed using 0.38 mM of ammonium nitrate (AN) (Fig. 1, column 9) and 15 μM of PFOS. Almost no change ($<5\%$) in PFOS concentration indicates strongly towards the significance of the Ce(IV) center.

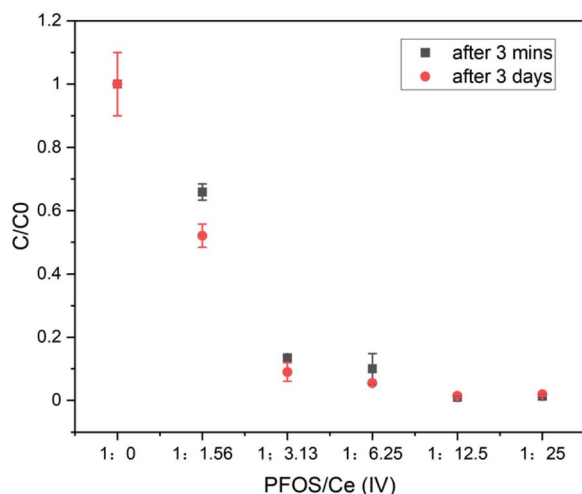


Fig. 2 The LC-MS detected 1.5 mM of PFOS in water before ($[\text{PFOS} : \text{Ce}^{\text{IV}}] = 1 : 0$) and after the treatment (3 minute) with 2.34 mM, 4.69 mM, 9.38 mM, 18.75 mM and 37.5 mM of CAN.

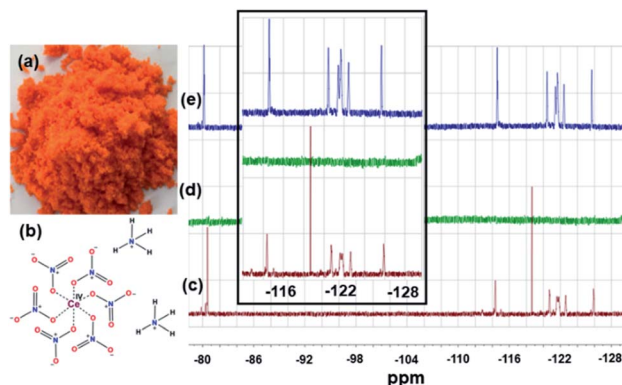


Fig. 3 (a) The colour and texture of CAN that was used for this study, (b) pictorial representation of CAN's chemical structure, (c) and (d) are the 376 MHz ^{19}F NMR of aqueous solutions (1.5 mM) of PFOS in $\text{DMSO}-d_6$, before and after the treatment with 18.75 mM CAN [Ce^{IV}] respectively showing complete disappearance of PFOS from water, and (e) 376 MHz ^{19}F NMR of the precipitate in $\text{DMSO}-d_6$. Inset shows magnified picture of c, d and e from -129 ppm to -114 ppm, indicating complexation or strong interaction between CAN [Ce^{IV}] and PFOS.

In order to obtain clarity on the possible species that are precipitating out from the CAN–PFOS interaction, higher concentrations were used. LC-MS analysis indicated that CAN is more effective at higher concentration and complete removal of PFOS (1.5 mM) is possible at around 18.75 mM concentration of CAN (by LC-MS) (Fig. 2) within 3 minutes of continuous stirring after mixing. An off-white precipitate formed over couple of days after mixing CAN and PFOS.

The solutions before and after the treatment with CAN and the precipitate formed during the experiment were also investigated by 400 MHz ^1H and 376 MHz ^{19}F NMRs (Fig. 3, S6–S9).† These NMR studies (Fig. 3) of the supernatant liquid showed complete removal (in the NMR detection limit) of PFOS from water. The slight change

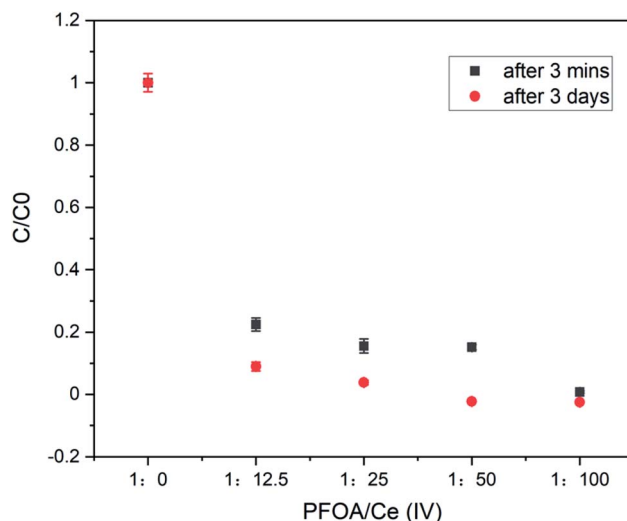


Fig. 4 The LC-MS detected concentrations of 1.5 mM of PFOA in water before ($[\text{PFOA} : \text{Ce}^{\text{IV}}] = 1 : 0$) and after the treatment (3 minute) with 18.75 mM, 37.5 mM, 75 mM and 150 mM of CAN.

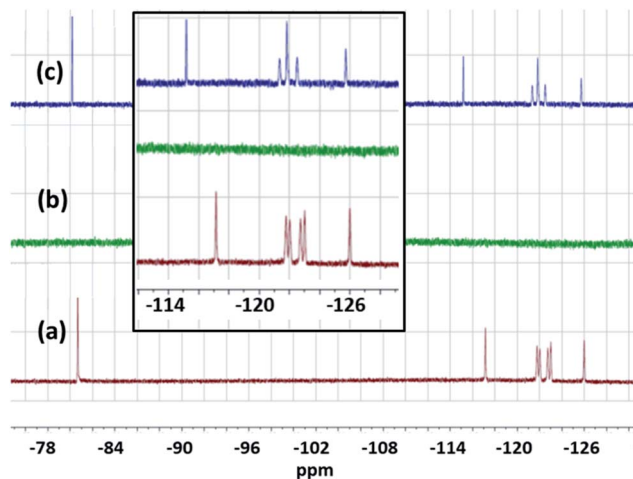


Fig. 5 (a) and (b) are the 376 MHz ^{19}F NMR of an aqueous solution (1.5 mM) of PFOA in $\text{DMSO}-d_6$, before and after the treatment with 0.15 M CAN $[\text{Ce}(\text{iv})]$ respectively showing complete disappearance of PFOA from water. (c) Represents the 376 MHz ^{19}F NMR of the precipitate in $\text{DMSO}-d_6$. Inset shows magnified picture of a, b and c from -128 ppm to -113 ppm, indicating complexation between $\text{Ce}(\text{iv})$ and PFOA.

in the chemical shift values and the splitting pattern of the NMR signals indicate CAN–PFOS interaction or possible complex formation.

To generalize our findings, we also tested another commercially available and highly regulated PFAS, perfluorooctanoic acid (PFOA, Fig. 6) as a substrate.²¹ At 15 μM starting concentration, at best 65% of the PFOA could be removed with 0.75 mM of CAN (Fig. S4†). At 1.5 mM PFOA, 78% removal was possible, with 18.75 mM CAN. However, complete removal of PFOA (1.5 mM) was detected (by LC-MS) at a much higher CAN concentration (0.15 M), after 3 minutes of continuous stirring (Fig. 4).

In this case, an off-white precipitate was observed after a week of undisturbed standing. Similar NMR investigations were performed by 400 MHz ^1H and 376 MHz ^{19}F NMR (Fig. 5, and S10–S13†) spectroscopy. The complete disappearance of the fluoride signal from water (in the NMR detection limit) and the change in the chemical shift values together with the splitting pattern of the precipitate also suggest complexation between Ce and PFOA.²² It is interesting to see more prominent changes in the chemical shift values in the case of the Ce–PFOA complex than that of the Ce–PFOS complex. In the case of Ce–PFOA complex, initial deprotonation of PFOA is expected prior to complexation/interaction with $\text{Ce}(\text{iv})$.

Very similar precipitation was observed with $\text{Ce}^{\text{IV}}(\text{SO}_4)_2$ at the higher concentration study. PFOS (1.5 mM) precipitated out

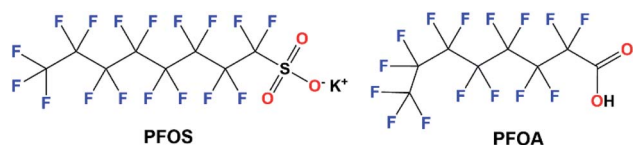


Fig. 6 Chemical structures of the two commercially available PFASs used in this study.

much faster (within a few hours) and complete removal (in the NMR detection limit) (Fig. S16†) was possible at 18.75 mM of $\text{Ce}^{\text{IV}}(\text{SO}_4)_2$. In comparison, faint white precipitation from PFOA (1.5 mM) and 18.75 mM of $\text{Ce}^{\text{IV}}(\text{SO}_4)_2$ was observed over a period of two days. NMR of the supernatant aqueous solution shows incomplete removal of PFOA under identical condition (Fig. S17†).

It is clear that PFOS reacts with CAN and $\text{Ce}^{\text{IV}}(\text{SO}_4)_2$ more efficiently (Fig. 2 and 4) and precipitates out faster than PFOA. This difference in reactivity can be due to the better donor ability (nucleophilicity) of the sulfonate group in PFOS than the carboxylic acid group in PFOA. Due to the higher acidity of the sulfonic acid, the conjugate base sulfonate (nucleophile) remains as the major species in solution in the presence of CAN in aqueous medium.²³ In the case of PFOA, to be equally effective, it needs higher concentration of CAN. Moreover, being shorter than PFOS, PFOA has been reported to show a higher tendency to stay in the aqueous phase, and therefore is difficult to precipitate out.²⁴ With AN (NH_4NO_3), no such precipitation was observed, and NMR results also confirm no detectable interaction between AN and PFOS (Fig. S14 and S15†).

All these observations suggest a possible complexation between $\text{Ce}(\text{iv})$ of CAN and these PFASs, followed by precipitation where PFOS and PFOA act as a nucleophile (ligand). No distinct new proton signal in 400 MHz ^1H NMR spectra of the aqueous solution and the precipitate after the CAN treatment, confirms the absence of any newly formed hydrocarbons.

HRMS analysis of the diluted (1 : 20 DMSO : water) solutions of the precipitates (from PFOA + CAN and PFOS + CAN) were quite complicated (Fig. S18 and S19†). It is reported that on the way to complexation with $\text{Ce}(\text{iv})/\text{Ce}(\text{iii})$, PFAS can trigger radical induced pathways to undergo C–C and C–F bond cleavages.^{25,26} In this regard, the formation of the polymeric compounds [e.g. $(\text{Ce})_m(\text{PFOA})_n(\text{NO}_3)_x(\text{NH}_4)_y(\text{H}_2\text{O})_z$, where m , n , x , y and z can be any arbitrary numbers from 0, 1, 2 to n] cannot be completely ruled out. We were unable to identify any new species with reasonable precision from the HRMS experiments. The precise identification of these species needs further detailed investigation with various other spectroscopic techniques.

In conclusion, CAN effectively precipitates out both PFOS and PFOA even at low concentration ($<20 \mu\text{M}$) and clearly reacts more efficiently with the former. A clear trend of $\text{CAN} > \text{Ce}(\text{SO}_4)_2 > \text{CeCl}_3 > \text{CoCl}_2 \geq \text{FeCl}_2 > \text{NH}_4\text{NO}_3$ was observed for precipitating out PFOS and PFOA from the aqueous medium. These PFASs are reported among the most difficult ones to degrade.²⁷ The Ce^{IV} center plays an important part in this, as reflected by the comparative studies using CAN, $\text{Ce}(\text{SO}_4)_2$, CeCl_3 and AN. The oxidation state of the cerium center, as well as the size of the PFASs play crucial roles in the successful removal/precipitation from the aqueous medium. Thus, a similar simple strategy can potentially be used in future for the design of highly effective filter beds (e.g., using a combination of GAC/PAC and CAN) for purifying PFAS contaminated water at larger scale.

Conflicts of interest

Authors declare no conflict of interests.

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