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A highly selective and sensitive H₂S sensor at low temperatures based on Cr-doped α -Fe₂O₃ nanoparticles

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Cr-doped α -Fe₂O₃ nanoparticles were synthesized by a low-cost and environmentally friendly hydrothermal route. Their gas sensing properties were investigated and the sensor showed high sensitivity and selectivity to H₂S gas. Different Cr doping levels from 0 to 8.0 wt% were studied, and the sensor of 4.0 wt% Cr-doped α -Fe₂O₃ showed the largest response, with a response of 213 to 50 ppm H₂S at 100 °C. The incorporation of Cr ions within α -Fe₂O₃ nanocrystals increases the specific surface area, and promotes the oxidation of H₂S and oxygen adsorption in the air. Thus, the doping of Cr into α -Fe₂O₃ nanostructures would be a promising method for designing and fabricating high performance H₂S gas sensors.

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1. Introduction

Recently, atmospheric environmental pollution has become one of the most serious problems that all the countries of the earth are facing.¹ Many steps have been taken to resolve it, such as enhancing control over pollutant discharge, researching high-performance catalysts for the catalytic conversion of auto exhausts, developing new technologies for the monitoring of toxic gases, enhancing the environmental awareness of people *etc.* Among these, controlling and monitoring poisonous gases from sources of pollution are quite important so far. Accordingly, the development of gas sensors with enhanced sensitivity and selectivity is very urgent and important.

Hydrogen sulfide (H₂S) is a colorless, highly flammable and extremely toxic gas with a characteristic foul odour of rotten eggs.^{2–5} Furthermore, it is also produced in large quantity from industrial activities and places such as petroleum natural gas drilling and refining, wastewater treatment, coke ovens, tanneries, kraft paper mills, and landfills.^{6,7} Hence, it is necessary to detect concentrations of H₂S in work environments such as oil and natural gas industries. To date, much effort has been made to prepare a number of gas-sensing nanomaterials, such as Fe₂O₃,^{8–13} WO₃,¹⁴ SnO₂,¹⁵ CuO,¹⁶ ZnO,³ and some compounds.^{17–19} Among them, α -Fe₂O₃ is an excellent potential H₂S-sensitive material.

In recent years, α -Fe₂O₃ has attracted great attention for its distinctive sensing properties toward H₂S gas. Despite some exciting results obtained, the development of more highly

sensitive and markedly selective gas sensors based on α -Fe₂O₃ nanostructures remains a challenge. Up to now, many efforts have been taken such as heterostructure constructing, element doping, and decorating in order to improve its performance.^{10,20,21} Among these methods, element doping has been turned out to be a very simple and feasible method. Chromium, as a member of transition metals, exhibits a wide range of possible oxidation states. And it is a promising candidate for efficient doping materials in sensor applications. There have been some reports showing the performance of Cr dopant in ammonia sensors,^{22–24} NO₂ sensors,^{25,26} acetone sensors,²⁷ oxygen sensor,²⁸ and so on. Herein we report a facile one-step hydrothermal method to synthesize Cr-doped α -Fe₂O₃ nanoparticles, for optimizing the gas sensing properties. At 100 °C, the sensor using 4 wt% Cr-doped α -Fe₂O₃ displayed the highest response toward hydrogen sulfide gases with a low cross-selectivity and short response time. The enhanced sensing response can be attributed to the increases in surface area and catalysis effect by Cr doping.

2. Experimental details

2.1 Synthesis of Cr-doped α -Fe₂O₃

All of the reagents used in the experiments were of analytical grade without further purification. Fe(NO₃)₃·9H₂O and Cr(NO₃)₃·9H₂O were used as iron and chromium sources, respectively. A series of Cr-doped α -Fe₂O₃ (0.0, 2.0, 4.0, 6.0, and 8.0 wt%) nanoparticles were synthesized by one-step hydrothermal reaction. In a typical experiment, Fe(NO₃)₃·9H₂O (0.8077 g), Urea (0.5 g) and Cr(NO₃)₃·9H₂O (0.0168 g) were put into 35 ml of distilled water to form homogeneous solution under vigorous stirring. Subsequently 300 μ l ammonia was

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introduced into above solution by magnetic stirring. The precursor solution was continuously stirred at room temperature to obtain a homogeneous mixture. Afterwards, the solution was transferred to a Teflon-lined stainless steel autoclave with a capacity of 50 ml and maintained at 180 °C for 12 h. After the hydrothermal procedure, the autoclave naturally cooled to room temperature. The solid products were subjected to centrifugation, rinsed alternately with distilled water and ethanol several times, and dried at 70 °C in air for 24 h. Finally, the dry precipitates were annealed in a muffle furnace at 500 °C for 2 h in air.

2.2 Sensor fabrication and sensing measurements

In a typical experiment, a small amount of the sensing material powder was mixed with terpineol followed by 2 h grinding to form a paste, then coated uniformly onto a ceramic tube. After the coating was air dried, the α -Fe₂O₃-coated Al₂O₃ microtube was fixed to a special pedestal with six poles by welding four Pt electrodes to four poles of the pedestal. A Ni–Cr heating coil was then inserted through the Al₂O₃ microtube, and its two ends were welded to the other two poles of the pedestal. A schematic representation and a micrograph image of gas sensor were shown in Fig. 1. The measurements of hydrogen sulfide gases sensing were conducted on a ZhongKe NS-4003 Smart Sensor Analyser (Beijing ZhongKe Micro-Nano Networking Technology Co., Ltd.). The relative humidity at room temperature was about 52% during the whole experiment. The sensor response is defined as $\text{response} = R_a/R_g$, where R_a is the resistance in air and R_g is the measured resistance in the presence of the test gas. The response time is defined as the period in which the sensor output change reached 90% of the stabilized value, similarly, the recovery time is defined as the time required for the sensor to reach 10% of the initial steady-state value after the sensor was taken out of the test chamber.

2.3 Characterization

X-ray powder diffraction (XRD) data of the as synthesized materials were collected on a Philips X'pert pro diffractometer using Cu-K α radiation at 40 kV and 30 mA. Scanning electron microscopy (SEM) images and EDX data were recorded on a scanning electron microscope (SEM, Zeiss SUPRA 55) equipped with an energy dispersive X-ray Spectroscopy (EDS, Oxford).

The transmission electron microscopy images (TEM) were obtained using JEOL JEM-2100. The Raman spectra were performed by using the IDSpec ARCTIC Raman system. X-ray photoelectron spectroscopy (XPS) measurements were performed with a PHI Quantum 200 spectrophotometer with Mg K radiation.

3. Results and discussion

3.1 Characterization of sensor material

The Raman spectrum of the α -Fe₂O₃ is presented in Fig. 2. In the range of 200–1500 cm⁻¹, many peaks could be observed, which were located at 225, 290, 405, 495, 611, 658, 825, 1081, and 1322 cm⁻¹, respectively. All these peaks correspond to α -Fe₂O₃ phase.^{29,30} The XRD patterns of the pure and Cr-doped α -Fe₂O₃ samples are shown in Fig. 3a. Clearly, all the diffraction peaks of all the samples could be indexed to the rhombohedral structure of α -Fe₂O₃ phase (JCPDS file no. 33-0664). No other peaks corresponding to Cr-related compounds were observed in all the XRD patterns of 0–8 wt% Cr-doped α -Fe₂O₃ samples. In addition, as shown in Fig. 3b, a high angle shift was detected from the (110) peak *via* comparing the different Cr doping α -Fe₂O₃. This is due to larger ionic radius of Fe³⁺ (0.67 Å) as compared to Cr³⁺ (0.64 Å), which according to Bragg's law shifts the peak position to high diffraction angles. The incorporation

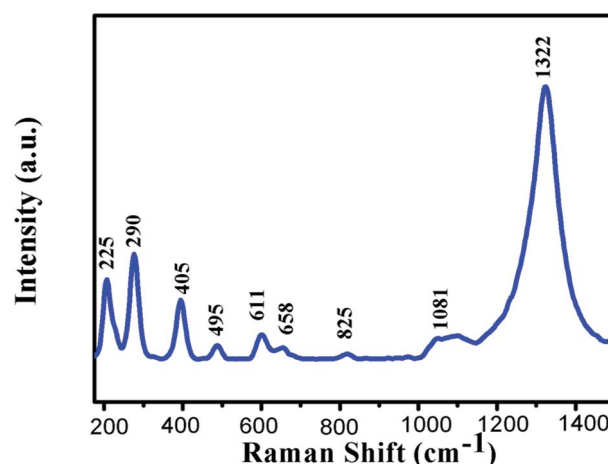


Fig. 2 Raman scattering spectra of pristine α -Fe₂O₃.

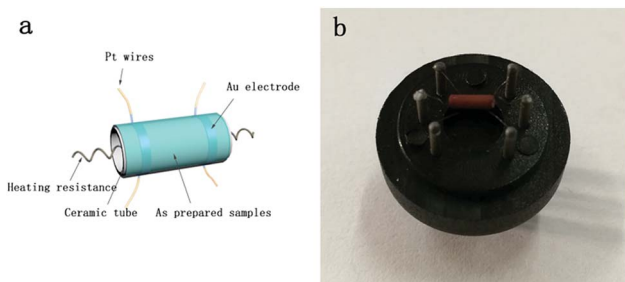


Fig. 1 (a) A schematic representation and (b) a micrograph image of gas sensor.

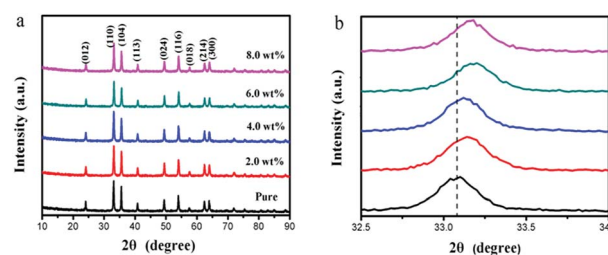


Fig. 3 (a) XRD patterns and (b) high-resolution of (110) peak of the pure, 2.0 wt%, 4.0 wt%, 6.0 wt% and 8.0 wt% Cr-doped α -Fe₂O₃ nanoparticles.

of atom with low ionic radius atom results in shrink of unit cell that decreases the lattice parameter and decreases d -spacing thus shifting peak positions to higher angles. Therefore, the substitution of Fe^{3+} by Cr^{3+} induced the high angle shift of diffraction peaks, suggesting that Cr element is incorporated into the $\alpha\text{-Fe}_2\text{O}_3$ lattice.

Morphology of the pristine and Cr-doped $\alpha\text{-Fe}_2\text{O}_3$ structures were also discerned by SEM observations. As shown in Fig. 4a, spherical particles coexisted with irregular particle aggregates. As the amount of Cr^{3+} was increased to 4 wt%, the obtained product consisted of small nanospheres with good monodispersity, as shown in Fig. 4c. However, particle aggregates without an orderly shape were collected when the amount of Cr^{3+} increased to 8 wt%, as shown in Fig. 4e. $\alpha\text{-Fe}_2\text{O}_3$ nanoparticles can be obtained by using a one-step hydrothermal process at an appropriate concentration of ammonia. And Cr^{3+} may control the morphology of $\alpha\text{-Fe}_2\text{O}_3$ because of its acid nature.^{23,31} The detailed mechanism of Cr on Fe_2O_3 nanoparticle is not clear now and needs further study. The EDS patterns of 4 wt% Cr-doped $\alpha\text{-Fe}_2\text{O}_3$ nanoparticles shown in Fig. 3f indicate that the as-prepared 4 wt% Cr-doped $\alpha\text{-Fe}_2\text{O}_3$ nanoparticles are composed of Fe, O and Cr, where the x -axis represents X-ray energy (in keV) and the y -axis represents the intensity. And it is evident that the peaks associated with chromium in the doped samples confirm the presence of chromium in $\alpha\text{-Fe}_2\text{O}_3$ nanoparticles.

Microstructures of the as-prepared products were studied by selected area electron diffraction (SAED) and high resolution transmission electron microscopy (HRTEM) characterizations. As shown in Fig. 5a, 4 wt% Cr-doped $\alpha\text{-Fe}_2\text{O}_3$ nanoparticles were observed by low magnification transmission electron microscopy. In the HRTEM images shown in Fig. 5b and c, we found

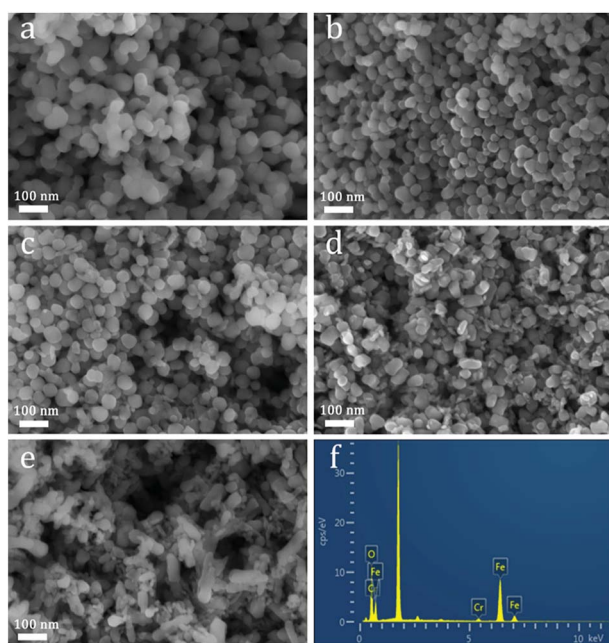


Fig. 4 (a–e) Typical SEM images of the pure, 2.0 wt%, 4.0 wt%, 6.0 wt% and 8.0 wt% $\alpha\text{-Fe}_2\text{O}_3$ nanoparticles, (f) the EDX spectrum of the 4.0 wt% $\alpha\text{-Fe}_2\text{O}_3$ nanoparticles.

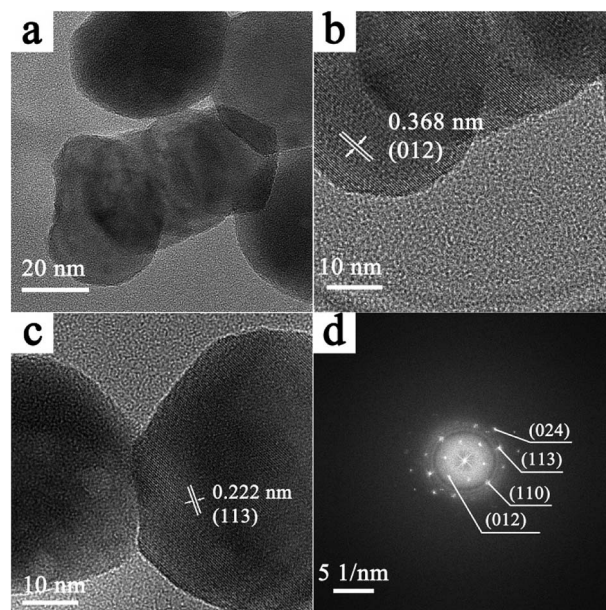


Fig. 5 (a) Low magnification TEM image of the 4 wt% Cr-doped $\alpha\text{-Fe}_2\text{O}_3$. (b and c) HRTEM images and (d) SAED pattern of the 4 wt% Cr-doped $\alpha\text{-Fe}_2\text{O}_3$.

the lattice distances were about 0.368 nm and 0.222 nm, which can be corresponded to the (012) and (113) lattice plane of $\alpha\text{-Fe}_2\text{O}_3$. The diffraction spots in a SAED pattern (Fig. 5d) matched well with (012), (110), (113) and (024) planes of $\alpha\text{-Fe}_2\text{O}_3$.

X-ray photoelectron spectroscopy (XPS) was performed to illuminate the surface composition of the studied 4 wt% Cr-doped $\alpha\text{-Fe}_2\text{O}_3$ nanoparticles. The spectra of Fe 2p and Cr 2p are shown in Fig. 6. In Fig. 6a, the binding energies of Fe $2p_{3/2}$ and Fe $2p_{1/2}$ are 710.7 and 724.3 eV, respectively, which are in good agreement with the literature values of Fe^{3+} in $\alpha\text{-Fe}_2\text{O}_3$.^{32–34} The binding energy of Cr $2p_{3/2}$ is 577.8 eV and Cr $2p_{3/2}$ –Cr $2p_{1/2}$ binding energy separation is 9.6 eV, that is typical for Cr^{3+} centers dispersed in an oxide environment (Fig. 6b).^{35–37}

3.2 Gas sensing properties

3.2.1 Gas-sensing measurement. Gas sensing experiments to 50 ppm hydrogen sulfide gases at different temperatures are performed as shown in Fig. 6a. The responses of pure, 2.0 wt%, 4.0 wt%, 6.0 wt%, and 8.0 wt% Cr-doped $\alpha\text{-Fe}_2\text{O}_3$ samples are tested at the temperatures from 60 to 160 °C which are found to increase with the increasing temperature, reach the maximum at a certain temperature, and then decrease with a higher

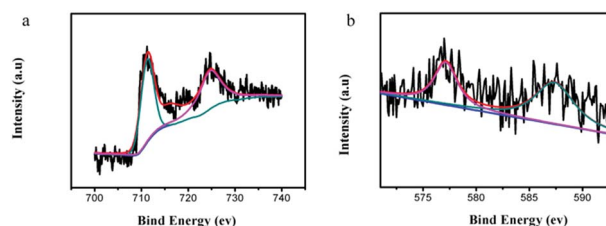


Fig. 6 XPS spectra of 4.0 wt% Cr-doped $\alpha\text{-Fe}_2\text{O}_3$ nanoparticles (a) Fe 2p, (b) Cr 2p.

operating temperature. It is a common behavior on the surface of α -Fe₂O₃ or similar semiconducting metal oxides which can be explained by the kinetics and mechanics of gas adsorption and desorption. When the heating temperature is low, the chemical activation of nanoparticles is accordingly low, and it causes a low response. The increasing temperature leads to positive chemical activation, and the sensor gets the maximum response at a certain heating temperature. The further increasing operating temperature results in some adsorbed gas molecules escaping from surface before gas reaction, and then the response decreases correspondingly.

Meanwhile, Fig. 7a shows that the Cr-doped α -Fe₂O₃ gas sensors achieve a maximum sensitivity at 100 °C, while the pure α -Fe₂O₃ sensor obtain the maximum sensitivity at 120 °C. The gas responses of the sensors based on the pure, 2.0 wt% 4.0 wt% 6.0 wt% and 8.0 wt% Cr-doped α -Fe₂O₃ to 50 ppm H₂S at 100 °C is 112, 155, 213, 102 and 65, respectively. The result revealed that sensor based on 4.0 wt% Cr-doped α -Fe₂O₃ showed the highest response to 50 ppm H₂S and the value was about two times higher than that of pure α -Fe₂O₃. According to the above test result, we could surmise that the Cr on sensor would play a crucial role in the governing mechanism. Therefore, further measurements were carried out using this specific composition (4 wt%) and operating temperature (100 °C) to explore the effect of other factors on the sensing properties of Cr-doped α -Fe₂O₃.

Subsequently, variations in the electrical conductance were measured on α -Fe₂O₃ nanoparticles. Butane, ammonia, ethanol and methane common gases (50 ppm) used as probe molecules to investigate the cross-sensitivity of the sensor at an operating temperature of 100 °C (Fig. 7b). Compared with the conductance change in response to 50 ppm H₂S at 100 °C, the responses of the 4.0 wt% Cr-doped α -Fe₂O₃ sensor to other gases are negligible. The 2.0 wt% Cr-doped α -Fe₂O₃ sensor shows a low response to ammonia and ethanol. Consequently, it was concluded that the sensor using as-synthesized 4.0 wt%

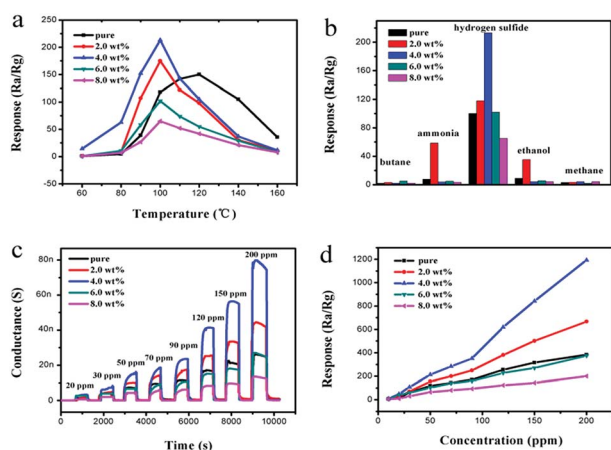


Fig. 7 (a) Gas responses of the sensors based on the pure, 2.0 wt%, 4.0 wt%, 6.0 wt%, and 8.0 wt% Cr-doped α -Fe₂O₃ nanoparticles vs. operating temperatures to 50 ppm H₂S. (b) Gas responses of five sensors to 50 ppm various target gases at 100 °C. (c) Real-time response curves of five sensors to different H₂S concentration at 100 °C. (d) Gas responses of five sensors as a function of the H₂S concentration at 100 °C.

Cr-doped α -Fe₂O₃ nanoparticles showed excellent selectivity to H₂S as opposed to other test gases.

The concentrations of hydrogen sulfide were varied from 20 to 200 ppm. Transient response curves of the sensor (composed of pure, 2.0 wt%, 4.0 wt%, 6.0 wt%, and 8.0 wt% Cr-doped α -Fe₂O₃) at the optimal temperature of 100 °C are shown in Fig. 7c. The nearly rectangle-shaped response and recovery curves indicate the excellent hydrogen sulfide response and rapid reaction rate of the sensor. Exposing the sensor to H₂S significantly increases conductance until it reaches saturation. After the removal of H₂S, the conductance approaches its initial value.

Fig. 7d shows the relationship between the response of the sensor and H₂S concentration from 20 to 200 ppm calculated from Fig. 7c. Increasing the concentration accelerates the response. Among them, the response of the sensor based on the 4.0 wt% Cr-doped α -Fe₂O₃ was apparently higher than the other sensors. The response of Cr-doped Fe₂O₃ sensors showed almost no saturation with the variation of H₂S concentrations between 20 and 200 ppm, indicating a wide response range of the Cr-doped sensor and its possible application in detecting high H₂S concentration environment.

From these, the Cr doping amount of 4.0 wt% was considered as the optimum doping concentration. Then the response and recovery characteristics of the sensor based on 4.0 wt% Cr-doped α -Fe₂O₃ nanoparticles were investigated to 50 ppm hydrogen sulfide gases at 100 °C. As shown in Fig. 8a, the conductance increased upon exposure to hydrogen sulfide gases, which was consistent with the gas sensing behavior of n-type oxide semiconductor. The response and recovery times of the sensors based on the 4.0 wt% Cr-doped α -Fe₂O₃ were 29 and 80 s, respectively. Stability of the sensors is also investigated. The 4.0 wt% Cr-doped α -Fe₂O₃ sensor was stored in air and kept working at 100 °C for subsequent sensing property tests at a level of 50 ppm H₂S after aging for 30 days. It can be seen that

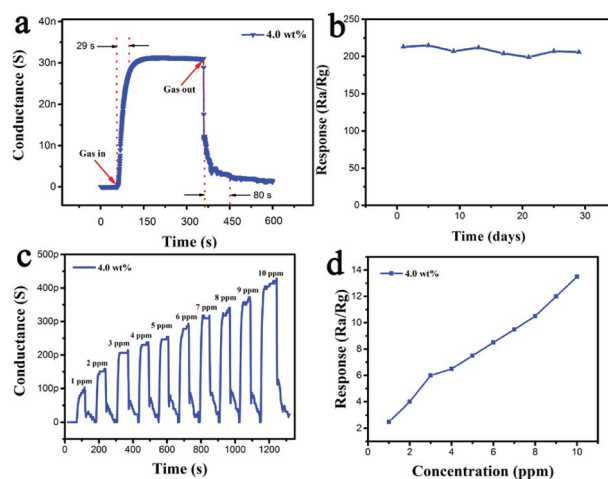


Fig. 8 (a) Dynamic response of the sensor based on 4 wt% Cr-doped α -Fe₂O₃ to 50 ppm H₂S at 100 °C. (b) Stability of the sensor to 50 ppm H₂S. (c) Dynamic response of the sensors to different concentrations H₂S from 1 ppm to 10 ppm. (d) Responses of the sensor to different concentrations of H₂S.

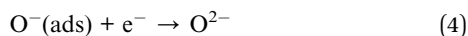
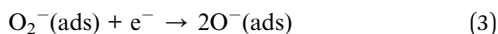
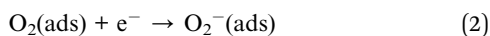
Table 1 Comparison of the sensing performance of Fe₂O₃ sensors

Material	H ₂ S concentration (ppm)	Temperature (°C)	Response	Detection limit (ppm)	Reference
Porous α -Fe ₂ O ₃	50	350	66.7	1	12
Ag loaded Fe ₂ O ₃	100	160	220	50	8
Pd decorated Fe ₂ O ₃	50	160	75	10	10
Au modified Fe ₂ O ₃	10	250	6.4	1	9
Fe ₂ O ₃ @WO ₃	10	150	190	0.5	14
Cr-doped α -Fe ₂ O ₃	50	100	213	1	This work

the sensor presented a nearly constant response to H₂S during the test, showing that the sensor exhibited long-term stability (Fig. 8b). We also investigated the detection limit of H₂S by the developed sensor. As shown in Fig. 8c, we detected dynamic response curve of sensor based 4 wt% Cr-doped α -Fe₂O₃ to H₂S ranging from 1 ppm to 10 ppm at 100 °C. Fig. 8d showed the relationship between the response of the sensor and H₂S concentration from 1 to 10 ppm. The detection of 4 wt% Cr-doped α -Fe₂O₃ to 1 ppm H₂S had the responses of about 2.5. When we tested H₂S below 1 ppm at optimal temperature, the sensor did not have response.

A comparison of the sensing performances between the sensor in this work and other H₂S sensors based on α -Fe₂O₃ reported previously is summarized in Table 1. From the table, comprehensively considering the gas response and operating temperature, it is obvious that 4.0 wt% Cr-doped α -Fe₂O₃ nanoparticles showed a high gas response at a low operating temperature. This was conducive to practical application and energy conservation. Therefore, the sensor based on 4.0 wt% Cr-doped α -Fe₂O₃ nanoparticles had more superiority than those reported in the literature.

3.2.2 Gas sensing mechanism. It is known that metal oxide gas sensing depends on the interaction between the test gas and the adsorbed oxygen species on the surface of the material.^{38,39} α -Fe₂O₃ is a n-type semiconductor. Therefore, when the α -Fe₂O₃ are exposed in an air atmosphere, oxygen molecules will adsorb on the α -Fe₂O₃ surface and capture electrons from α -Fe₂O₃ to generate the chemisorbed oxygen species at higher temperatures, leading to an increase in resistance. The main chemical equations involved in the process are shown as follows (eqn (1)–(4)):^{40,41}



When the α -Fe₂O₃ sensor is exposed to H₂S atmosphere, the surface oxygen ions react with the H₂S molecule and release a number of electrons back into the semiconductor, resulting in a decrease in resistance of the sensor. This reaction process may be as eqn (5)



However, α -Fe₂O₃ presented high catalyst for H₂S oxidation, and the dominant product was SO₂. Steijns *et al.* studied the H₂S oxidation reaction, they showed that the SO₂ formation rate was very slow at low temperatures. It was attributed to the high activation energy of SO₂ formation, which would be about 125 kJ mol⁻¹. Whereas the activation energy for sulfur formation was between 8 and 50 kJ mol⁻¹. So, Berben *et al.* prepared α -Fe₂O₃-based catalysts, with chromium as additives. It showed good activity and good selectivity, which directly and selectively oxidized H₂S with O₂ into elemental sulfur.^{42,43}

Fig. 9 shows in a double logarithmic plot the variation of the response of α -Fe₂O₃ sensor *versus* H₂S concentration. The linear trend can be related to the conductance model limited by the electron transport across the intergranular Schottky barrier^{24,44} where

$$\sigma \propto \exp(-eVs/kT) \quad (6)$$

(σ is the conductance and eVs represents the surface barrier height at the intergranular contact).

According to this model, the response of the α -Fe₂O₃ nanoparticles is explained by using the following equation:^{45,46}

$$\log \sigma/\sigma_0 = A_g + \beta \log P_g \quad (7)$$

where σ_0 denotes the conductance in the absence of the target gas, A_g is a prefactor, P_g is the gas partial pressure, and β is the response order. Mechanisms were formulated to describe the

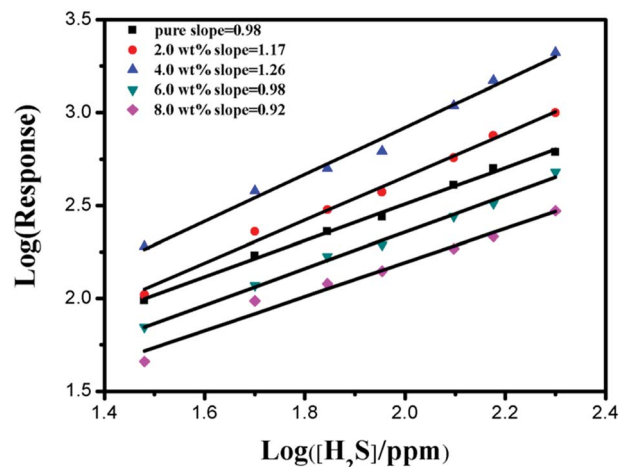
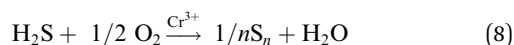


Fig. 9 Response variation for the pure, 2.0 wt%, 4.0 wt%, 6.0 wt% and 8.0 wt% Cr-doped α -Fe₂O₃ nanoparticles.

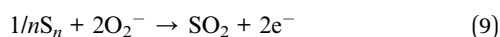
surface reaction and predict the β value in eqn (7). This is usually 1 or 1/2, depending on the charge state of the surface species, the stoichiometry of the elementary surface reactions, and the microstructure of the material. These properties contribute to the fluctuations of the response order value by different weights.

In the case of pristine α -Fe₂O₃ nanoparticles (Fig. 9), the calculated β value (*i.e.*, the slope of the curve described by eqn (7)) was about 0.98. Differently, 2.0 and 4.0 wt% Cr-doped α -Fe₂O₃ sensors show β values about 1.17 and 1.26. This can be reasonably ascribed to the presence of Cr element which may give rise to a different sensing mechanism at the microscale level.

Therefore, a mechanism was suggested to explain the excellent gas response of Cr-doped α -Fe₂O₃. First, when H₂S gas flows in, the H₂S molecules are oxidized into elemental sulfur and H₂O by the Cr³⁺ catalyst in α -Fe₂O₃. The adsorption process is as follows (eqn (8)):



In the meantime, some of the chemisorbed oxygen species reacted with elemental sulfur (eqn (9))



The ability of a reacting species to overcome a free energy barrier that is preventing a reaction is commonly modeled by the Arrhenius equation. For most of the chemical reactions, including the primitive reaction and primitive action, the Arrhenius equation gives the dependence of rates, K , of a chemical reaction on the absolute temperature T (in Kelvin) as shown in eqn (10):

$$K = A \exp(-E_a/RT) \quad (10)$$

This equation shows that the rate constant exponentially increases as activation energy (E_a) decreases. For the pristine α -Fe₂O₃ sensor, higher activation energy is required for produce SO₂. Therefore, the operating temperature of the reaction is higher. In the presence of Cr³⁺, the α -Fe₂O₃ sensor provides a reaction path with lower activation energy, providing a high response at a low temperature. For these reasons, the working temperature of the Cr-doped α -Fe₂O₃ sensors can be reduced to low temperature while the sensitivity was still kept high.

Finally, the further increasing of Cr doping content, the size of α -Fe₂O₃ gas sensing material became larger and the agglomeration of the nanoparticles became more obvious (Fig. 4d and e), which would reduce surface-to-volume ratio and efficient diffusion of the gas molecules. This may be the reason for the poor performance of 6.0 wt%, and 8.0 wt% Cr-doped α -Fe₂O₃ sensors.

4. Conclusions

In summary, we have developed a simple one-step hydrothermal method for the synthesis of Cr-doped α -Fe₂O₃

nanoparticles with different doping levels. The results of X-ray diffraction, EDS and XPS data proved that Cr ion successfully entered into the crystal lattice of α -Fe₂O₃ without deteriorating the original crystal structure. Furthermore, an appropriate amount of Cr-doped α -Fe₂O₃ (4.0 wt% Cr-doped α -Fe₂O₃ sensor in this work) display, upon exposure to H₂S, outstanding sensing properties even at relatively low operating temperature, and the sensitivity resulted as much higher than that of the pristine sensor. This behaviour could be attributed to the nanometer size effect and catalysis effect. Hence, we can confirm that the doping of Cr into α -Fe₂O₃ is a promising strategy for improving the gas sensing performance of the α -Fe₂O₃ based gas sensor.

Conflicts of interest

There are no conflicts to declare.

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