

Thiazolothiazoles

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Highly Colored Boron-Doped Thiazolothiazoles from the Reductive Dimerization of Boron Isothiocyanates

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Dedicated to Professor Todd B. Marder on the occasion of his 65th birthday

Abstract: Reduction of (CAAC)BBr₂(NCS) (CAAC = cyclic alkyl(amino)carbene) in the presence of a Lewis base L yields tricoordinate (CAAC)LB(NCS) borylenes which undergo reversible E/Z-isomerization. The same reduction in the absence of L yields deep blue, bis(CAAC)-stabilized, boron-doped, aromatic thiazolothiazoles resulting from the dimerization of dicoordinate (CAAC)B(NCS) borylene intermediates.

Conjugated molecules and materials based on fused thiazolo[5,4-*d*]thiazole (TzTz) building blocks have attracted much interest in the past decade because of their interesting optoelectronic properties and high thermal stability.^[1] Theoretical studies have shown that the introduction of aromatic TzTz spacers into conjugated oligo- and polymers often enhances their π -stacking ability and light-harvesting efficiency, improves charge injection and transport, hinders charge recombination and increases dye regeneration.^[2] These properties make TzTz-based materials particularly suitable for applications in organic photovoltaic cells (OPVs), organic light-emitting diodes (OLEDs), and organic field-effect transistors (OFETs).^[3] Furthermore, TzTz-based luminescent materials and small molecules have been developed

for chemical sensing,^[4] photocatalysis,^[5] live cell imaging,^[6] and NIR-II photothermal conversion and therapy.^[7]

Studies have shown that the replacement of an endocyclic carbon unit [CR] (R = anionic substituent) by an isolobal boron unit [BR] in heteroaromatic molecules results in unique electronic properties owing to the strong electrophilic character of boron.^[8] This enables new applications of boron-heteroarenes in coordination chemistry, pharmacology, and materials science. While B,N-heteroarenes have been intensively studied,^[8] and there is an increasing interest in B,S-heteroarenes owing to their potential use in electronic devices,^[9] examples of B,N,S-heteroarenes remain essentially limited to the stable 1,3,2-benzothiazaborole motif, which is easily accessed through the reaction of borane precursors with *o*-aminothiophenols^[10] or benzothiazole derivatives.^[11] In contrast, monocyclic thiazaboroles (Tzb) have only been studied theoretically,^[12] in particular with regard to their potential as superhalogens, that is, aromatic molecules with an electron affinity higher than that of chlorine.^[13] Consequently, studies on the applications of Tzb derivatives remain rare. Benzothiazaborolyl ligands have been used in transition metal chemistry^[14] while benzothiazaborolates have been tested for their biocidal properties.^[15] The successful use of oligomers containing aromatic benzo-bis([1,3,2]Tzb) units in OFETs^[16] also highlights the potential of fused Tzb building blocks in optoelectronic devices.

Herein, we report a straightforward synthetic route towards hitherto unknown fused [1,3,2]thiazaborolo[5,4-*d*]-[1,3,2]thiazaboroles (TzbTzb) through the reductive dimerization and C–C coupling of transient isothiocyanatoborylenes of the form [LB(NCS)] (L = neutral donor ligand). Spectroscopic and computational studies highlight the differences and similarities with the parent TzTz unit.

Inspired by our previous study on the reduction of a cyclic alkyl(amino)carbene (CAAC)-stabilized cyanoborane to a self-stabilizing, tetrameric cyanoborylene, [(CAAC)B(CN)]₄, which features a 12-membered (BCN)₄ ring,^[17] we wondered if the reduction of a related (CAAC)-stabilized boron isothiocyanate would yield a similar oligomeric, self-stabilizing [(CAAC)B(NCS)]_n species. To verify the viability of (CAAC^R)B(NCS) borylenes, their tricoordinate adducts with the N-heterocyclic carbene (NHC) IPr (1,3-diisopropylimidazol-2-ylidene) were first synthesized by room temperature reduction of the (CAAC^R)BBr₂(NCS) precursors (CAAC^{Me} = 1-(2,6-diisopropylphenyl)-3,3,5,5-tetramethylpyrrolidin-2-ylidene; CAAC^{Cy} = 2-(2,6-diisopropylphenyl)-

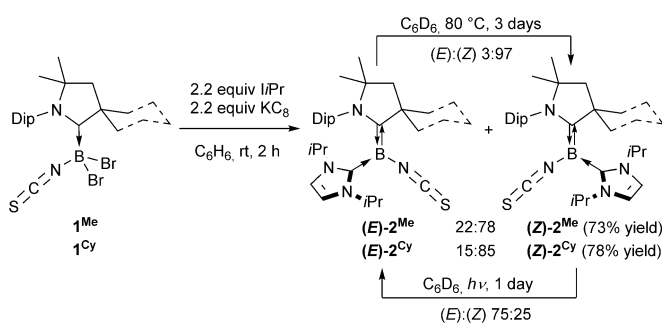
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Scheme 1. Synthesis and isomerization of CAAC–NHC-stabilized (*E*- and (*Z*-isothiocyanatoborylenes. Dip = 2,6-diisopropylphenyl.

3,3-dimethyl-2-azaspiro[4.5]decan-1-ylidene), $\mathbf{1}^{\text{Me}}$ and $\mathbf{1}^{\text{Cy}}$,^[18] with 2.2 equiv KC_8 in benzene (Scheme 1).

After workup, fractional crystallization of each borylene from benzene afforded two consecutive crops of orange crystals, each showing a different ^{11}B NMR resonance, the first at -2.6 ppm, the second at 3.8 ppm. In each case X-ray crystallographic analyses identified the first crop (62–68 % yield of isolated product) as the (*Z*) isomer of the unsymmetrical isothiocyanatoborylene, (*Z*)- $\mathbf{2}^{\text{R}}$, in which N1 and N4 are positioned on the same side of the $\text{C}2=\text{B}3$ double bond, and the second crop (ca. 10 % yield of isolated product) as the corresponding (*E*)- $\mathbf{2}^{\text{R}}$ isomer (see Figure 1 for (*Z*)- $\mathbf{2}^{\text{Me}}$ and (*E*)- $\mathbf{2}^{\text{Cy}}$, and Figures S55 and S56 in the SI for molecular structures of (*E*)- $\mathbf{2}^{\text{Me}}$ and (*Z*)- $\mathbf{2}^{\text{Cy}}$, respectively). These compounds display structural parameters typical for CAAC–NHC-stabilized tricoordinate borylenes,^[17,19] with trigonal planar boron centers and short $\text{B}-\text{C}_{\text{CAAC}}$ bonds (ca. 1.45 Å) with double bond character, indicating strong π backbonding from the borylene center to the π -acidic CAAC ligand. The NHC rings are rotated almost perpendicularly to the borylene plane (torsion angle $\text{N}4\text{-B}3\text{-C}7\text{-N}8$ $73\text{--}85^\circ$) and coordinating as pure σ donors ($\text{B}2\text{-C}7$ ca. 1.59 Å). The $\text{B}3\text{-N}4$ distance (ca. 1.49 Å) suggests partial π donation from the NCS ligand to boron.^[20]

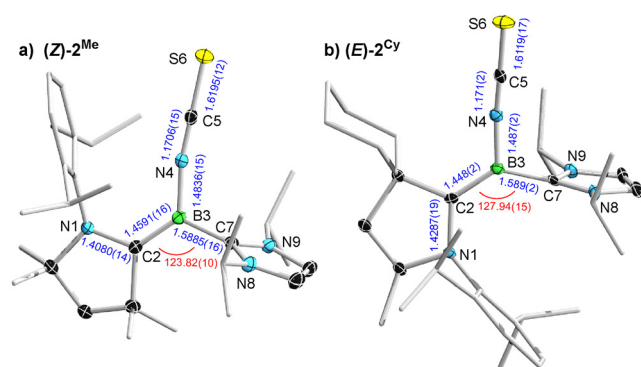


Figure 1. Crystallographically derived molecular structures of a) (*Z*)- $\mathbf{2}^{\text{Me}}$ (only one of the two molecules present in the asymmetric unit is represented) and b) (*E*)- $\mathbf{2}^{\text{Cy}}$.^[38] Atomic displacement ellipsoids drawn at 50% probability level. Ellipsoids on the CAAC ligand periphery and hydrogen atoms omitted for clarity. Selected bond lengths (Å) in blue, selected angles (°) in red.

The ca. 5 ppm separation of the ^{11}B NMR shifts of the two isomers, with (*Z*)- $\mathbf{2}^{\text{R}}$ at -2.6 ppm and (*E*)- $\mathbf{2}^{\text{R}}$ at 3.8 ppm, can be explained by Hirshfeld charge analysis,^[21] which indicates slightly positive (ca. $+0.003$) and negative (ca. -0.012) boron atoms for (*E*)- $\mathbf{2}^{\text{R}}$ and (*Z*)- $\mathbf{2}^{\text{R}}$, respectively. The different steric constraints were also apparent in the ^1H NMR spectra, which showed symmetrical ligand resonances indicative of free $\text{B}-\text{C}_{\text{NHC}}$ bond rotation for (*Z*)- $\mathbf{2}^{\text{R}}$ and very broad ligand resonances split into two magnetically inequivalent sets, typical of hindered rotation, for (*E*)- $\mathbf{2}^{\text{R}}$.^[22] NMR spectroscopic analyses of the crude borylene solutions prior to recrystallization showed (*E*)/(*Z*) ratios of 22:78 and 15:85 for $\mathbf{2}^{\text{Me}}$ and $\mathbf{2}^{\text{Cy}}$. In both cases the preference for the (*Z*) isomer is explained by the strong steric repulsion between the bulky CAAC-Dip substituent and $i\text{Pr}$ ligand. Heating of a C_6D_6 solution of isolated (*E*)- $\mathbf{2}^{\text{Me}}$ at 80°C for 3 days resulted in 97 % conversion to the thermodynamically preferred (*Z*)- $\mathbf{2}^{\text{Me}}$ isomer.^[23] Conversely, irradiation of isolated (*Z*)- $\mathbf{2}^{\text{Me}}$ under a UV lamp for 1 day afforded a 75:25 mixture of the (*E*) and (*Z*) isomers, respectively. The greater thermodynamic stability of the (*Z*) isomers is also corroborated by DFT calculations at the OLYP^[24]/TZP level, with computed $\Delta G[(\text{Z})\text{-}(\text{E})\text{-}\mathbf{2}^{\text{R}}]$ values of -6.86 kcal mol $^{-1}$ for $\text{R} = \text{Me}$ and -5.01 kcal mol $^{-1}$ for $\text{R} = \text{Cy}$. While the vast majority of unsymmetrical borylenes of the type $(\text{CAAC})(\text{L})\text{BR}$ are formed exclusively as a single (*Z*) or (*E*) isomer, depending on the relative sterics of L and R,^[19,25] a similar, thermally induced, reversible isomerization has been observed only recently for the first time in a $(\text{CAAC}, \text{PMe}_3)$ -stabilized borylborylene, where it was likely induced by the thermal lability of PMe_3 .^[26]

In the absence of a stabilizing Lewis base the twofold reduction of $\mathbf{1}^{\text{R}}$ with KC_8 in benzene resulted in a color change to intense blue and a new downfield-shifted ^{11}B NMR resonance around 32 ppm. After workup, dark blue crystals of the bis(CAAC)-stabilized thiazaborolo[5,4-*d*]thiazaboroles $\mathbf{3}^{\text{Me}}$ and $\mathbf{3}^{\text{Cy}}$ were isolated in 38% and 67% yield, respectively.^[27]

X-ray crystallographic analyses of $\mathbf{3}^{\text{R}}$ (see Figure 2a for $\mathbf{3}^{\text{Me}}$ and Figure S57 in the SI for $\mathbf{3}^{\text{Cy}}$) show fully planar, centrosymmetric structures, with a central TzbTzb unit, formally resulting from dimerization of two dicoordinate $(\text{CAAC}^{\text{R}})\text{B}(\text{NCS})$ borylenes through $\text{S}\rightarrow\text{B}$ adduct formation and $\text{C}-\text{C}$ coupling at the $\text{C}5$ positions. The reaction is somewhat reminiscent of the coupling of four CO molecules at diborynes, which afforded fused bicyclic bis(borolactones).^[28] We investigated plausible mechanistic pathways for the reductive cyclization of $\mathbf{1}^{\text{Me}}$ using DFT calculations, of which the main intermediates are shown in Scheme 2 (see Figures S62–S63 for a full description). In path 1, the eight-membered ring $\mathbf{A1}^{\text{Me}}$ is obtained after a sequence of stepwise $\text{S}\rightarrow\text{B}$ adduct formation, with barriers of 14.9 (TS1) and 9.2 kcal mol $^{-1}$ (TS2). A $\text{C}-\text{C}$ coupling step from $\mathbf{A1}^{\text{Me}}$ leads to $\mathbf{3}^{\text{Me}}$, the formation of which is exergonic by a total of -81.5 kcal mol $^{-1}$, and with a barrier (TS3) of 10.3 kcal mol $^{-1}$ from $\mathbf{A1}^{\text{Me}}$. Our calculations reveal a second viable pathway (path 2) where an early $\text{C}-\text{C}$ coupling takes place after the first $\text{S}\rightarrow\text{B}$ adduct formation and leads to the stable intermediate $\mathbf{A2}^{\text{Me}}$. The corresponding energy barriers (TS2 and TS3) are lower than those of path 1. Our results thus indicate

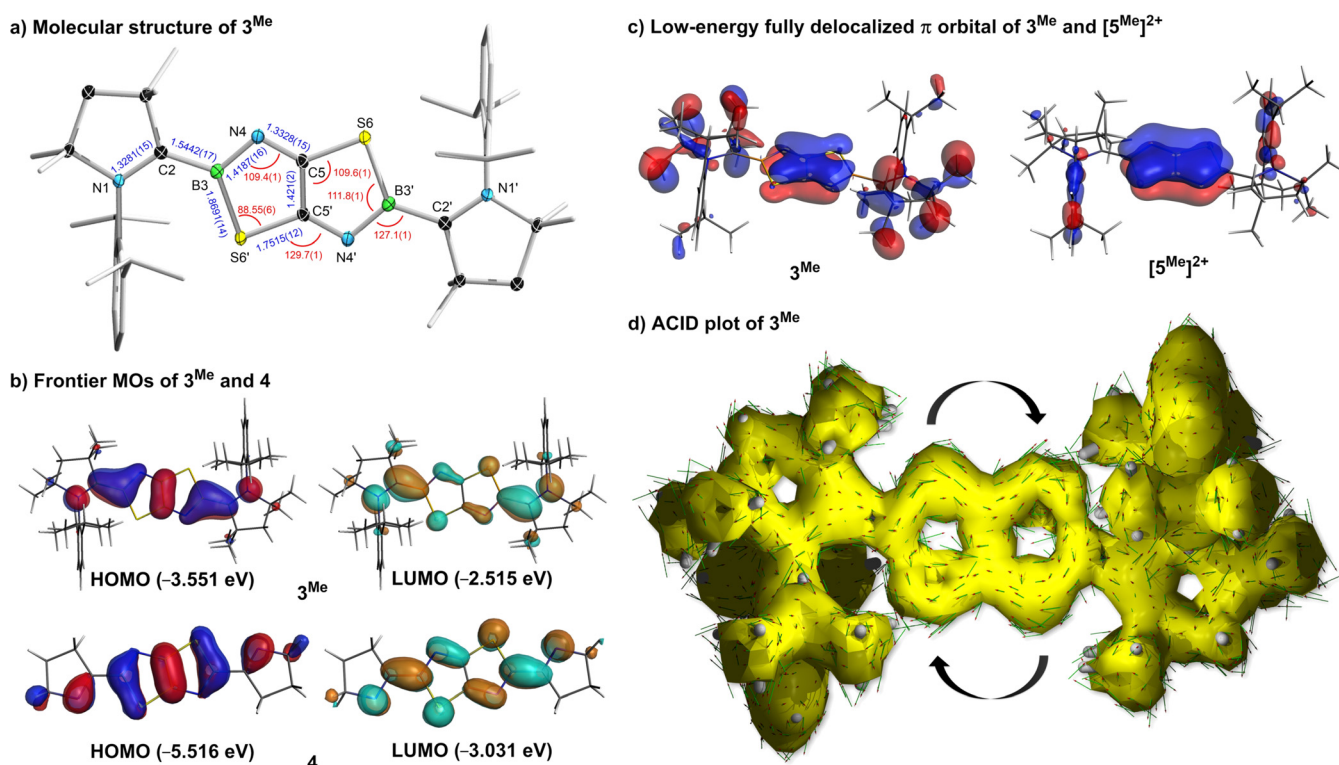
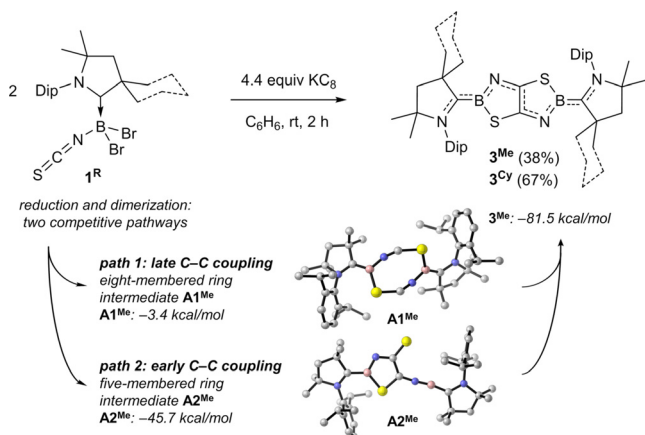


Figure 2. a) Crystallographically derived molecular structure of 3^{Me} .^[38] Ellipsoids on the CAAC ligand periphery and hydrogen atoms omitted for clarity. Selected bond lengths (Å) in blue, selected angles (°) in red. b) Plots of HOMO and LUMO of 3^{Me} and **4** at the OLYP/TZP level of theory. c) Low-energy fully delocalized π orbital of 3^{Me} and $[5^{\text{Me}}]^{2+}$. d) ACID plot of 3^{Me} (isosurface: 0.025).



Scheme 2. Reductive cyclization of 1^{R} in the absence of a Lewis base.

that both mechanisms are viable, with a preference for path 2, and suggest that the C–C coupling can also occur immediately after the first S→B adduct formation. Conversely, direct C–C coupling from two isolated borylenes is inaccessible, as this occurs via a barrier of 31.9 kcal mol $^{-1}$ (**TS4**). Finally, the formation of $A1^{\text{Me}}$ from isolated borylenes through a concerted transition state also seems unlikely, and all attempts to optimize such a transition state converged directly to **TS1**.

In comparison to the similarly planar structure of PhTzTzPh,^[29] the presence of the boron atoms in the TzBTzB unit results in shortening of the C–N bonds and lengthening of the central C–C bond by less than 3%, while

the C–S bonds remain unaffected. The average N1–C2 (1.33 Å), C2–B3 (1.54 Å), B3–N4 (1.42 Å), N4–C5 (1.33 Å), and C5–C5' (1.42 Å) bond lengths are all within the range of partial double bonds, indicating π electron delocalization over the entire (NCBNC) $_2$ framework. The B–S distance of ca. 1.86 Å, however, is relatively long for a single bond (compare B(SPh) $_3$: B–S 1.81 Å),^[30] suggesting that the sulfur atoms participate less effectively in the delocalized π system of 3^{R} . In order to investigate the role of π delocalization, we analyzed the electronic structure of 3^{Me} and other sulfur-containing heterocycles, particularly the TzTz analogue of 3^{Me} , 2,5-bis(2,2,4,4-tetramethyl-3,4-dihydro-2H-pyrrol-5-yl)thiazolo-[5,4-d]thiazole, herein labelled **4**, and its dicationic N-arylated derivative, $[5^{\text{Me}}]^{2+}$, the latter being isoelectronic to 3^{Me} . We show that the HOMO of 3^{Me} , unlike that of **4**, is π -delocalized over the (NCBNC) $_2$ framework with the exclusion of the sulfur atoms (Figure 2b). Furthermore, analysis of the fully delocalized π orbital of 3^{Me} and $[5^{\text{Me}}]^{2+}$ (Figure 2c) reveals that delocalization is affected by the presence of the boron atoms due to the increase in energy mismatch between the participating 2p orbitals,^[31] and becomes less prominent in 3^{Me} . As a consequence, the aromaticity of the carbene-stabilized BNC $_2$ S cycles is decreased relative to those of the corresponding thiazole rings.^[32] Accordingly, calculations using the anisotropy of the induced current density (ACID)^[33] method (Figure 2d, see Section S10 for more details) revealed the presence of a diatropic, clockwise π electron circulation, typical of aromatic compounds, along the bicyclic central structure of 3^{Me} . The attribution of a weak

but distinct aromatic character for 3^R is also supported by nucleus-independent chemical shift (NICS) calculations,^[34] with the NICS(0), NICS(1) and the zz tensor component at 1 Å above the ring, NICS_{zz}(1), values for 3^{Me} being -6.3 , -7.0 , and -12.5 ppm, respectively (see SI for comparison with other systems).

In contrast to planar ArTzTzAr compounds (Ar = phenyl, thienyl, thiazenyl, pyridyl), which are generally colorless to yellow,^[28,35] 3^{Me} and 3^{Cy} absorb around 675 nm, accounting for the intense blue color (3^{Me} : $\epsilon = 66200 \text{ M}^{-1} \text{ cm}^{-1}$; 3^{Cy} : $\epsilon = 83300 \text{ M}^{-1} \text{ cm}^{-1}$), with secondary absorption maxima around 628 and 453 nm (see Figure S41 in the SI). TDDFT calculations for 3^{Me} (see SI for more details and the relevant MOs) using the double-hybrid functional B2PLYP^[36] with the def2SVPD^[37] basis set indicate that the features around 675 and 628 nm are the $\pi \rightarrow \pi^*$ transitions HOMO \rightarrow LUMO (93%) and HOMO-1 \rightarrow LUMO (93%), respectively. In turn, two states contribute to the band around 453 nm, namely S_7 and S_9 . The former mainly results from the charge transfer excitation from HOMO-1 to LUMO + 4 (88%), while the latter is related to a $n \rightarrow \pi^*$ transition (HOMO-7 \rightarrow LUMO, 44%; HOMO-2 \rightarrow LUMO + 1, 41%). Substitution of the two carbon atoms in **4** with electron-deficient boron atoms is followed by a significant reduction of the HOMO-LUMO gap of 3^{Me} , and explains the pronounced red-shift of its lowest-energy absorption band. This highlights once more the significant electronic differences between the boron-doped thiazolothiazole and its carbon analogue.

To conclude we have shown that the reduction of CAAC-stabilized isothiocyanatoboranes in the presence of a Lewis base results in the formation of doubly base-stabilized borylenes, whereas in the absence of a Lewis base dimerization and C-C coupling of two (NCS) units leads to [1,3,2]thiazaborolo[5,4-d][1,3,2]thiazaboroles. While the latter are less aromatic than their thiazolothiazole counterparts, they display an intense blue coloration, whereas their carbon analogues are colorless. We are currently studying the reactivity and photophysics of these novel boron-doped fused heterocycles and will report our findings in due course.

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Conflict of interest

The authors declare no conflict of interest.

Keywords: aromaticity · borylenes · DFT · fused thiazaboroles · reduction

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- [18] See the Supporting Information for details of the syntheses of all precursors and their crystallographically derived structures (Figures S50–S54).
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