


Article

Chalcogen Bonding due to the Exo-Substitution of Icosahedral Dicarborane

Jindřich Fanfrlík ^{1,*} , Drahomír Hnyk ^{2,*} and Pavel Hobza ¹

¹ Institute of Organic Chemistry and Biochemistry of the Czech Academy of Sciences, Flemingovo nam. 2, 166 10 Prague 6, Czech Republic

² Institute of Inorganic Chemistry of the Czech Academy of Sciences, 250 68 Husinec-Řež, Czech Republic

* Correspondence: fanfrlik@uochb.cas.cz (J.F.); hnyk@iic.cas.cz (D.H.)

Academic Editor: Antonio Caballero

Received: 4 June 2019; Accepted: 11 July 2019; Published: 23 July 2019



Abstract: Chalcogen atoms are a class of substituents capable of generating inner and outer derivatives of boron clusters. It is well known that chalcogenated boron clusters can form strong σ -hole interactions when a chalcogen atom is a part of an icosahedron. This paper studies σ -hole interactions of dicarboranes with two exopolyhedral chalcogen atoms bonded to carbon vertices. Specifically, a computational investigation has been carried out on the co-crystal of $(1,2\text{-C}_2\text{B}_{10}\text{H}_{10})_2\text{Se}_4$ •toluene and a single crystal of $(1,2\text{-C}_2\text{B}_{10}\text{H}_{10})_2\text{Te}_4$.

Keywords: sigma hole; heteroborane; co-crystal

1. Introduction

Polyhedral boron hydrides (boranes) are known for the presence of delocalized electron-deficient bonding. It is characterized by the aggregation of atoms to form 3-center-2-electron (3c-2e) bonds [1], which result in the formation of trigonal faces and hypercoordination. The three-dimensional deltahedral shapes typical of boron clusters are classified according to their formal electron counts, namely by the terms *closo*, *nido*, *arachno*, and *hypho* [2]. The *closo* clusters are of particular current interest as they exhibit especially high thermal and chemical stability. The number of vertices, n , can range from 5 to 12 with the formula $\text{B}_n\text{H}_n^{2-}$, with the 12-vertex icosahedral cluster (an example in Figure 1), $\text{B}_{12}\text{H}_{12}^{2-}$ (point-group symmetry I_h), being the most common and most stable [3]. The replacement of one or more boron atoms at a vertex by atoms of other elements results in the formation of *closo* heteroboranes. For example, the replacement of two formally neutral BH groups by two CH^+ moieties yields the neutral dicarboranes $\text{C}_2\text{B}_{10}\text{H}_{12}$ in a variety of three isomers differing in the relative positions of the hypercarbon atoms within the icosahedral cage. The 1,2-isomer, the so-called *o*-carbaborane with C_{2v} symmetry, is the least stable isomer and has the largest dipole moment [4] among them. This dipole moment directs from the middle point of the C-C vector toward the rest of the cluster. The fact that C vertices are the center of partial positive charge might be surprising because it contradicts the electronegativity concept. Such an electron distribution is, however, common in multicenter bonding [1]. Carboranes are considered 3D aromatic systems [5] and have rich substitution chemistry. Outer functionalization can be achieved by replacing terminal hydrogen atoms by various substituents to maintain the overall neutral charge. Chalcogen atoms belong to a class of substituents capable of generating derivatives of the icosahedral cage both as part of the cage [1] and as a peripheral group; the latter has so far been exemplified by thiolated carbaboranes. Very recently, a selenolated carbaborane with two SeH groups on both carbons has been prepared and structurally characterized [6]. However, the preparation of 1,2-(SeH)₂-*closo*-1,2- $\text{C}_2\text{B}_{10}\text{H}_{10}$ has been very difficult since the necessary intermediate for obtaining it, the lithium 1,2-diselenolato-1,2-dicarba-*closo*-dodecaboranate salt $(\text{Li}^+)_2[1,2\text{-(Se)}_2\text{-closo-1,2-C}_2\text{B}_{10}\text{H}_{10}]^{2-}$, tends to dimerize in an oxidative manner [7] via two C-Se-Se-C

diselenide bridges (see Figure 1). Analogous disulfide bridges are well known among cysteine amino acids. It has been shown that the replacement of an interchain disulfide bridge with a diselenide bridge increases the lifetime of insulin without impairing its hormonal function [8]. The same dimerization as in $(\text{Li}^+)_2[1,2-(\text{Se})_2\text{-}closo\text{-}1,2\text{-C}_2\text{B}_{10}\text{H}_{10}]^{2-}$ also applies to $(\text{Li}^+)_2[1,2-(\text{Te})_2\text{-}closo\text{-}1,2\text{-C}_2\text{B}_{10}\text{H}_{10}]^{2-}$ [9]. Consequently, $(1,2\text{-C}_2\text{B}_{10}\text{H}_{10})_2\text{E}_4$ (E = Se, Te) have been obtained and structurally characterized as reported in references [7] and [9], respectively (see Figure 1). Note that such a molecular shape with the two *o*-carborane moieties and four chalcogens that form an eight-membered ring in a chair conformation also exist for E = S, but this structure has not been structurally described [10]. Cyclic molecules containing divalent chalcogen atoms have been receiving a great deal of attention as they can form columnar structures and even nanotubes via intermolecular diE-bonds [11–13].

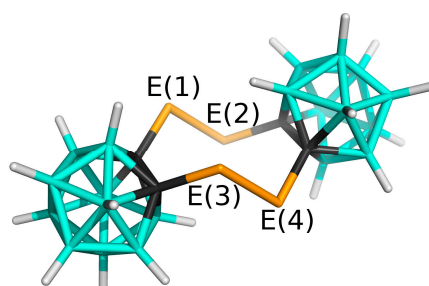


Figure 1. Molecular diagram of the Se_4C_4 and Te_4C_4 molecules that contain two icosahedral moieties connected by chalcogen bridges. The atom color coding is as follows: cyan—B; black—C; orange—E (E = Se, Te); white—H.

Chalcogen atoms belong to a class of substituents capable of also generating icosahedral derivatives in which chalcogen atoms are part of the icosahedron, *closo*-1-EB₁₁H₁₁ (E = S, Se, Te) [2]. The substitution chemistry of *closo*-1-SB₁₁H₁₁ has also been researched, but to a lesser extent than in the case of *o*-carborane. Therefore, 12-ph-*closo*-1-SB₁₁H₁₀ was prepared and a very strong chalcogen bonding between a partially positively charged sulfur [14] atom and the π -electron cloud of the phenyl from the neighboring molecule in the crystal packing was detected [15]. Pnictogen bonding between 3D and 2D aromatics was also spotted in the cocrystal of *closo*-1,7-P₂B₁₀Cl₁₀ and toluene [16]. Since toluene cocrystallizes with $(1,2\text{-C}_2\text{B}_{10}\text{H}_{10})_2\text{Se}_4$ (abbreviated as Se_4C_4) and $(1,2\text{-C}_2\text{B}_{10}\text{H}_{10})_2\text{Te}_4$ (abbreviated as Te_4C_4) crystallizes itself, we have carried out a computational investigation on these two crystals to examine the intensity of chalcogen (E-) and dichalcogen (diE-) bonding in the corresponding crystal packing. We have opted for these two crystals because the chalcogen atoms are located outside the icosahedron here in contrast to the previously studied heteroboranes with chalcogens directly involved in the icosahedron [15].

2. Methods

The molecular electrostatic potential (ESP) surfaces of isolated molecules were computed on the 0.001 a.u. molecular surfaces at the HF/def2-TZVP level using the Gaussian09 [17] and Molekel4.3 [18,19] programs.

The interactions in the crystal structures considered (CCDC codes 613,551 and 608,807) [7,9] were studied by employing a cluster model. The clusters were created around each unique molecule. The first layer consisted of molecules within 5 Å of the central molecule. Similarly, the second layer was formed by molecules within 5 Å of the first layer. Hydrogen atoms of the central molecule and the first layer were optimized by the DFT-D3/BLYP/DZVP method [20]. We selected the DFT-D3/BLYP/DZVP method because it makes it possible to obtain geometries close to more expensive calculations [20] for our extended clusters consisting of about 300–800 atoms. The resolution-of-identity (RI) approximation to the DFT method was also used. Due to the size of the second layer (comprising more than 1000 atoms), hydrogen atoms of the second layer were optimized by the semiempirical quantum mechanical

PM6-D3H4X method [21]. Heavy atoms were kept in crystallographic positions. The clusters obtained were used for energy calculations. Interaction energy (ΔE) was computed as the energy difference between the energy of the dimer and the sum of monomer energies. The energies were determined using the DFT-D3/TPSS/TZVPP level. ΔE values were decomposed by symmetry-adapted perturbation-theory (SAPT) methodology. The simplest truncation of SAPT (SAPT0) decomposition [22] was performed with the recommended jun-cc-pVDZ basis set [23]. The noncovalent interaction index (NCI) [24,25] was calculated for selected motifs in order to elucidate all possible interactions in these motifs. It was carried out at the recommended B3LYP/6-31G* level [24]. The stability of the selected binding motifs was also examined by optimizations of all atoms. Here, we combined the DFT-D3/BLYP/DZVP method with the LBFGS algorithm and the strict optimization criteria (energy change < 0.0006 kcal mol⁻¹, the largest gradient component < 0.12 kcal mol⁻¹ Å⁻¹ and the root-mean-square gradient < 0.06 kcal mol⁻¹ Å⁻¹). Vibrational frequencies were calculated numerically at the same level. Turbomole (7.0) [26], PSI4 [27], MOPAC2016 [28] and Cuby4 [29] programs were used.

3. Results and Discussion

3.1. Charge Distribution of Isolated Molecules

The charge distribution in isolated molecules has been studied by computing molecular electrostatic potential (ESP) surfaces (see Table 1 and Figure 2). The σ -holes [30] of the E atoms are the most positive areas of the molecular surfaces of the **Se₄C₄** and **Te₄C₄** molecules with large $V_{S,max}$ values of 28.6 and 35.6 kcal mol⁻¹. The **Se₄C₄** and **Te₄C₄** molecules are thus supposed to form strong σ -hole interactions. The higher value of **Te₄C₄** corresponds to the larger atomic number, smaller electronegativity, and larger polarizability of the Te atoms. The large $V_{S,max}$ values of **Se₄C₄** and **Te₄C₄** molecules might be caused by the fact that the E atoms are bound to the C vertex of the carborane moiety. It has been shown that the S atom bound to the B vertex in 9,12-(SH)₂-*closo*-1,2-C₂B₁₀H₁₀ has a negative ESP molecular surface without a positive σ -hole in contrast to the carbon-bound S atom of 1,2-(SH)₂-*closo*-1,2-C₂B₁₀H₁₀, which has a σ -hole with the $V_{S,max}$ value of 16.0 kcal mol⁻¹ [31]. Analogically, halogenated dicarbaboranes could form halogen bonds only when they were bound to the C vertex [32]. It should be noted that the **Se₄C₄** and **Te₄C₄** molecules have zero dipole moments due to their symmetry, while the parental carborane molecule has a very large dipole moment. BH vertices have a negative ESP surface, as a result of which they could form attractive contacts with the σ -holes of the Se and Te atoms. The π ring of the toluene molecule has a more negative molecular ESP surface than the BH vertices (see Table 1 and Figure 2), and its interaction with the σ -holes should thus be more favorable.

Table 1. The Dipole moment (μ) and the maximum and minimum values of the electrostatic potential of the molecular surfaces ($V_{S,max}$ and $V_{S,min}$, respectively) computed at the HF/def2-TZVP level. Energies are in kcal mol⁻¹ and μ in D.

	μ	$V_{S,max}$	$V_{S,min}$
Toluene	0.4	14.6	-19.8
Se₄C₄	0.0	28.6	-8.2
Te₄C₄	0.0	35.6	-10.1

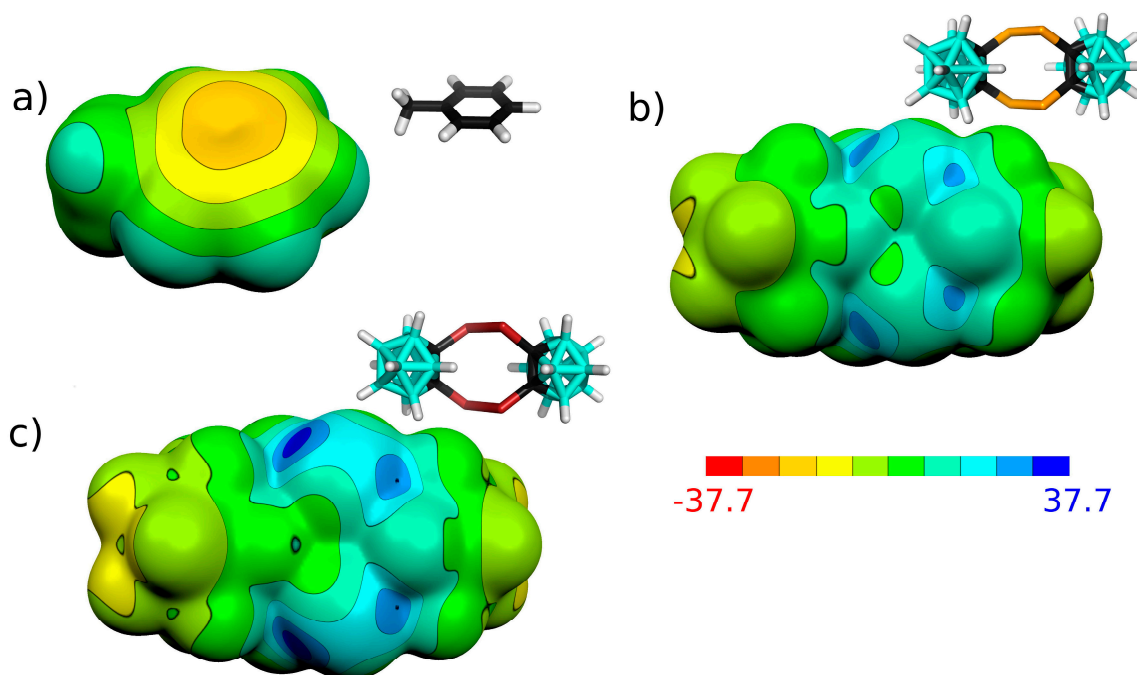


Figure 2. Molecular diagrams and computed electrostatic potential (ESP) molecular surfaces of toluene (a), Se₄C₄ (b) and Te₄C₄ (c). The ESP has been computed at the HF/def2TZVP level. The ESP color range is in kcal mol⁻¹. The atom color coding is as follows: cyan—B; black—C; orange—Se; white—H; red—Te.

3.2. Interactions in the Crystal of Se₄C₄•toluene and in the Single Crystal of Te₄C₄

The crystal packing of the studied crystals is shown in Figure 3. First, we focus on the crystal structure of Se₄C₄•toluene. The ΔE values of all crystal motifs are computed at the DFT-D3 level. Selected motifs are also examined by the SAPT0 method. The comparison of DFT-D3 and SAPT0 results verifies the accuracy of the selected DFT-D3 method (RMSE of 0.31 kcal mol⁻¹). In the case of toluene, the sums of the ΔE values for the first and second layers are -28.50 and -1.90 kcal mol⁻¹, respectively. The most negative ΔE value of the toluene molecules had the motif stabilized mainly by the C-Se $\cdots\pi$ E-bond with the separation between the Se atom and the center of the aromatic ring of the toluene being 3.30 Å, which corresponds to 91.7% of the sum of van der Waals radii (Σr_{vdW}) [33] of the C and Se atoms (Figure 4A). This motif had a highly negative ΔE value, exceeding -9 kcal mol⁻¹. It thus forms about 30% of the sum of ΔE values for the toluene molecule. The most important term of this motif in the SAPT decomposition is dispersion, followed by the electrostatic term. They form 64 and 27% of the attractive energy terms, respectively (Table 2). Similar results were reported for the EDA decomposition of the bifurcated type of Se $\cdots\pi$ interactions in diphenyl selenide [34]. The induction term which also includes charge transfer (orbital mixing) forms only about 9% of the attractive energy terms in the SATP decomposition, which agrees with findings that this contribution is smaller and does not determine the character of σ -hole interactions [35,36]. Additionally, this E-bond can be compared with the B-P $\cdots\pi$ pnictogen (Pn) bond found in the crystal of *closo*-1,7-P₂B₁₀Cl₁₀•toluene, which had the ΔE value of -9.9 kcal mol⁻¹ computed at the same DFT-D3 level, the length of 3.08 Å (88% of Σr_{vdW}), and a larger contribution of the electrostatic term in the SAPT0 decomposition (35% of the attractive energy terms) [16]. It should also be mentioned that the free optimization followed by the frequency analysis confirmed that this motif represents true minimum on potential energy surface. The second most stable interaction of the toluene molecule of the studied crystal structure is a stacking motif, which has no contact shorter than Σr_{vdW} (Figure 4B) and the ΔE value of -4.07 kcal mol⁻¹. In this motif, the dispersion is even more pronounced. It forms about 75% of the attractive energy terms in

the SAPT decomposition. This motif was not stable in gas phase optimization and the comparison of the X-ray motif with the optimized geometry is shown in Figure 4.

In the case of the Se_4C_4 molecule, the sum of the ΔE values for the first and second layers are -85.96 and -3.61 kcal mol $^{-1}$, respectively. Here, two motifs have highly negative ΔE values. One of them (i.e., the motif stabilized by the C-Se $\cdots\pi$ E-bond) was already described above. Interestingly, the Se_4C_4 dimer has even more negative ΔE value exceeding -10 kcal mol $^{-1}$, although it has no contact below Σr_{vdW} (see Figure 4c). It has two dichalcogen contacts of 3.87 Å (102% of Σr_{vdW}) and two Se \cdots H-B chalcogen contacts of 3.24 Å (108% of Σr_{vdW}). Additionally, it has numerous B-H \cdots H-B contacts with the H \cdots H separation of about 3 Å (136% of Σr_{vdW}). According to NCI analyses, these homopolar contacts also contribute to the binding. This is in agreement with the fact that the BH vector of dicarbaborane is another type of amphiphilic moiety [37], besides more the known CF_3 moiety of trifluorotoluene [38]. Interestingly, the crystal structure also contain the Se_4C_4 dimer motif stabilized exclusively via two homopolar B-H \cdots H-B contacts (see Figure S2a) with the H \cdots H separation of 2.60 Å (118% of Σr_{vdW}). This contact has the ΔE value of -2.67 kcal mol $^{-1}$ at the DFT-D3 level. Additionally, we have computed the ΔE value of a hypothetical dimer that has two B-H \cdots H-B contacts (see Figure S2b) with the elongated H \cdots H separation of 2.89 Å. This hypothetical dimer has the ΔE value of -2.22 kcal mol $^{-1}$. We can thus speculate that the B-H \cdots H-B contacts could form about 40% of the ΔE value for the interaction motif shown in Figure 4c.

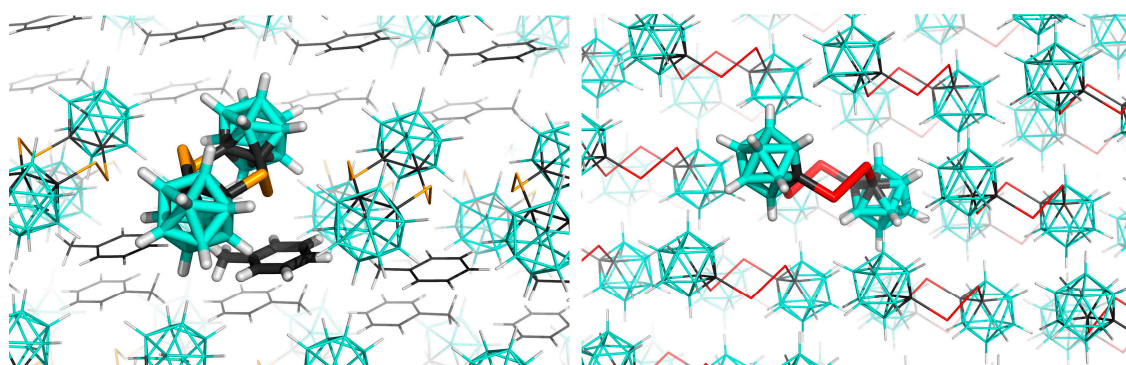


Figure 3. The crystal packing in the crystal of $\text{Se}_4\text{C}_4\bullet\text{toluene}$ (left) and the single crystal of Te_4C_4 (right). The atom color coding is as follows: cyan—B; black—C; orange—Se; white—H; red—Te.

Table 2. Computed DFT-D3/TPSS/TZVPP interaction energy (ΔE) values for the most stable motifs in the crystal of $\text{Se}_4\text{C}_4\bullet\text{toluene}$. Interaction energies were decomposed into electrostatic (E_{elec}), induction (E_{ind}), dispersion (E_{disp}), and exchange (E_{exch}) contributions using the SAPT0/jun-cc-pVDZ methodology. Energies are in kcal mol $^{-1}$. The relative values in parentheses show the contribution to the sum of all the attractive energy terms of SAPT0.

Interaction Motif	ΔE					
	DFT-D3	SAPT0				
		Total	E_{elec}	E_{ind}	E_{disp}	E_{exch}
$\text{Se}_4\text{C}_4\cdots\text{toluene}$ (E-bonding) ¹	-9.22	-9.28	-5.47 (27%)	-1.81 (9%)	-12.66 (64%)	10.66
$\text{Se}_4\text{C}_4\cdots\text{toluene}$ (stacking) ²	-4.07	-3.62	-1.32 (19%)	-0.39 (5%)	-5.13 (75%)	3.22
$\text{Se}_4\text{C}_4\cdots\text{Se}_4\text{C}_4$ (stacking_1) ³	-10.23	-10.27	-5.06 (22%)	-1.53 (7%)	-16.04 (71%)	12.36
$\text{Se}_4\text{C}_4\cdots\text{Se}_4\text{C}_4$ (stacking_2) ⁴	-5.02	-5.43	-1.86 (19%)	-0.37 (4%)	-7.61 (77%)	4.41

¹ The motif is shown in Figure 4a. ² The motif is shown in Figure 4b. ³ The motif is shown in Figure 4c. ⁴ The motif is shown in Figure 4d.

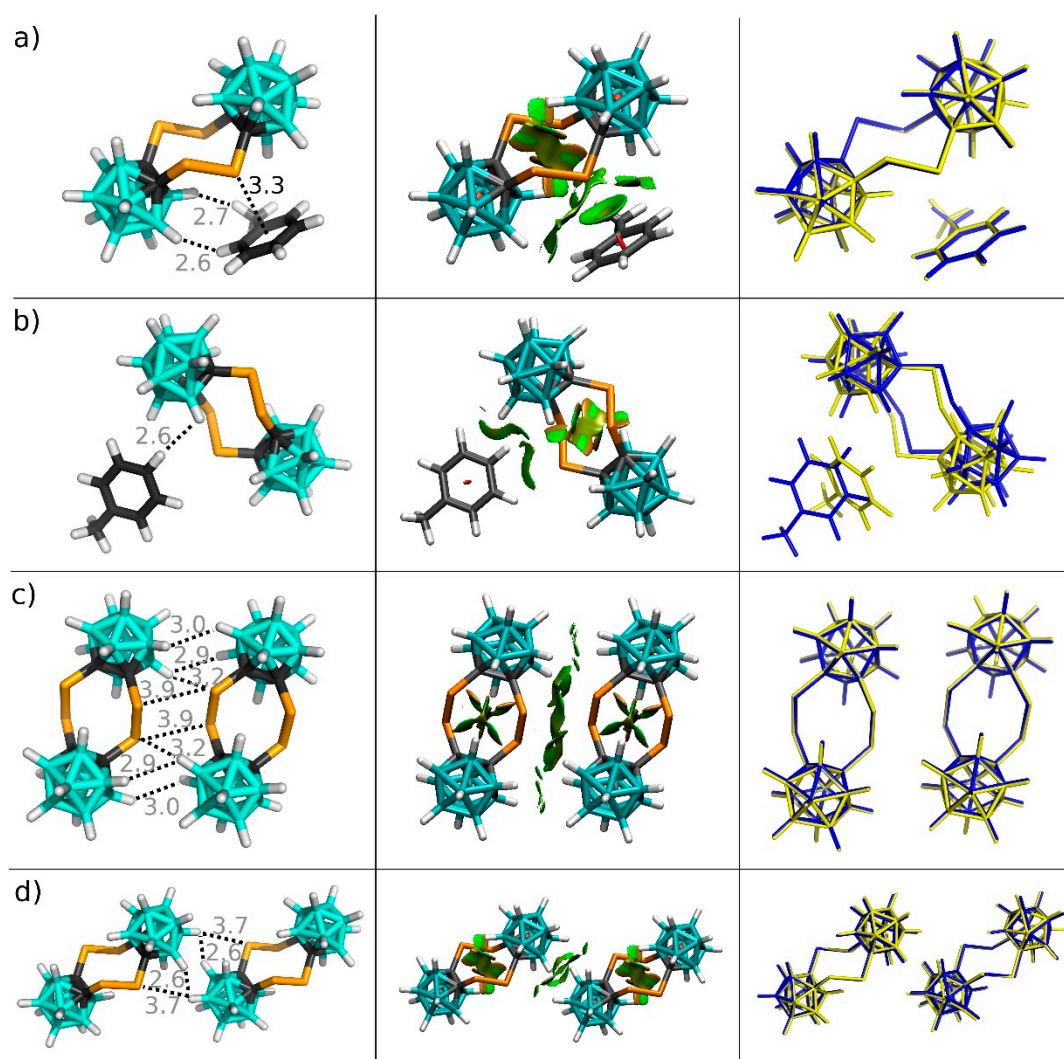


Figure 4. Four most stable binding motifs of the crystal structure of $\text{Se}_4\text{C}_4\cdot\text{toluene}$ (a), (b), (c) and (d). The left column: X-ray geometries (only the positions of H atoms optimized) and distances in Å. Distances larger than Σr_{vdW} are in gray. The middle column: NCI bonding isosurfaces. The NCI color scale is from -0.04 (blue) to 0.02 (red) a.u. The right column: A comparison of X-ray geometries (blue) with fully optimized structures in the gas phase (yellow).

In the second step, we focus on the single crystal of Te_4C_4 . Here, the sums of the ΔE values of the first and second layers are -91.86 and -4.58 kcal mol $^{-1}$, respectively. The most stable motif is stabilized by the C-Te \cdots Te-C diE-bond with the Te \cdots Te separation of 4.05 Å (98.3% of Σr_{vdW}), two C-Te \cdots H-B E-bonds with the Te \cdots H separation of 3.13 Å (99.1% of Σr_{vdW}), and numerous homopolar B-H \cdots H-B contacts with the H \cdots H separation of about 2.8 Å (127% of Σr_{vdW})—see Figure 5A. It has the ΔE value of -11.77 kcal mol $^{-1}$, which is approximately 24.4% of the sum of the ΔE values of Te_4C_4 considering that each molecule has this motif twice. The free optimization and frequency analysis have again confirmed the local minimum character of this motif. The second most stable motif has similar ΔE , specifically of -11.14 kcal mol $^{-1}$. This motif has two C-Te \cdots H-B E-bonds with the Te \cdots H separation of 3.10 Å (98.1% of Σr_{vdW})—see Figure 5B. This motif also has several C-Te \cdots H-B contacts with the Te \cdots H separation of about 3.4 Å. Together, these diE- and E- bonds thus represent about 48% of the sum of the ΔE values of Te_4C_4 . In contrast to unusual C-Te \cdots H-B E-bonds, the Te \cdots Te diE-bonds are considerably less rare; as an example, one can mention a very short Te \cdots Te contact with the separation of 3.78 Å (91.7% of Σr_{vdW}) in the crystal structure of 2,7-ditelluraocta-3,5-diyne [12].

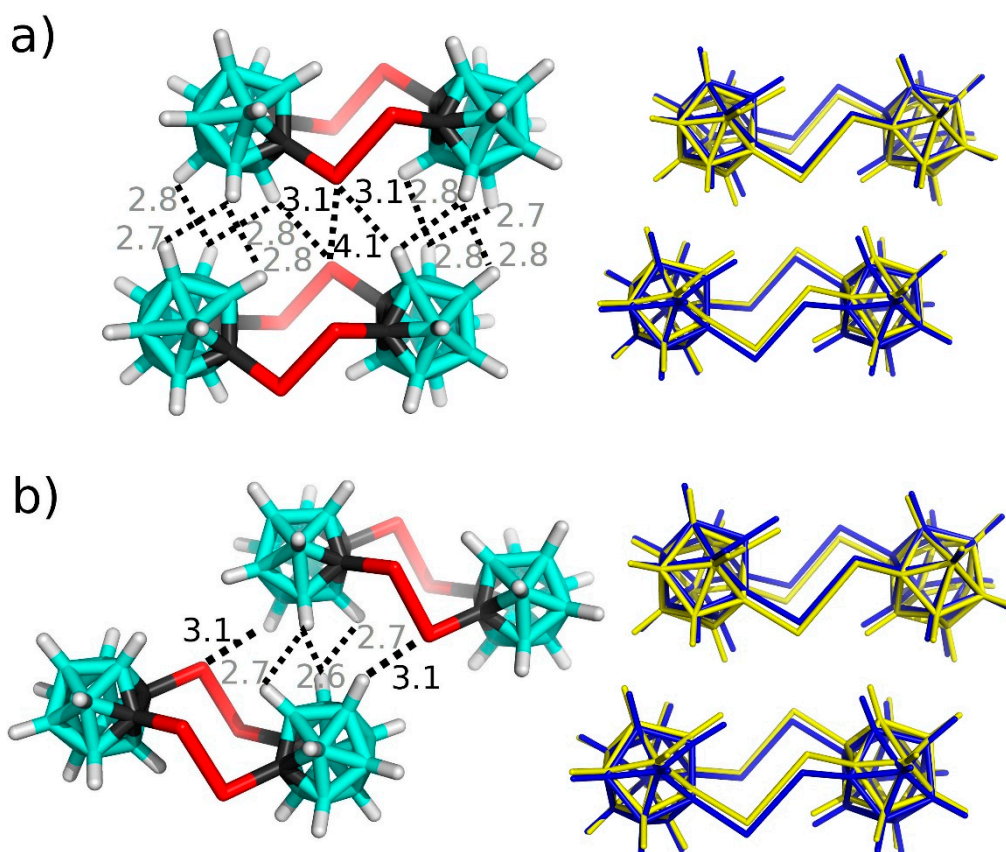


Figure 5. Two most stable binding motifs of the single crystal of Te_4C_4 (a) and (b). The left column: X-ray geometries (only the positions of H atoms optimized) and distances in Å. Distances larger than Σr_{vdW} are in gray. The right column: Comparison of X-ray geometries (blue) with fully optimized structures in the gas phase (yellow). Distances are in Å. The positions of H atoms were optimized at the DFT-D3/BLYP/DZVP level.

In summary, we have shown that the dicarbaboranes with two exopolyhedral chalcogen atoms bonded to the carbon vertices have highly positive σ -holes and thus a great ability to form chalcogen bonds. A computational investigation carried out on the crystal of $\text{Se}_4\text{C}_4 \bullet \text{toluene}$ and the single crystal of Te_4C_4 has highlighted the importance of $\text{E} \cdots \pi$, $\text{E} \cdots \text{HB}$ and $\text{E} \cdots \text{E}$ types of E- and diE-bonding.

Supplementary Materials: The following are available online, Figure S1: 2D-NCI plot of reduced density gradient (RDG) vs. $\text{sign}(\lambda_2)\rho$, Figure S2: Binding motifs of Se_4C_4 dimer stabilized via homopolar B-H \cdots H-B contacts.

Author Contributions: Methodology: J.F., D.H. and P.H.; writing—original draft preparation: J.F.; writing—review and editing: J.F., D.H. and P.H.; visualization: J.F.; project administration: J.F., D.H. and P.H.

Funding: This work was supported by the research project RVO 61388963 of the Czech Academy of Sciences. This research was funded by the Czech Science Foundation grant number 17-08045S, and by the European Regional Development Fund; OP RDE; Project: “Chemical biology for drugging undruggable targets (ChemBioDrug)” (No. CZ.02.1.01/0.0/0.0/16_019/0000729).

Conflicts of Interest: The authors declare no conflict of interest.

References

- Melichar, P.; Hnyk, D.; Fanfrlík, J. Systematic Examination of Classical and Multi-Center Bonding in Heteroborane Clusters. *Phys. Chem. Chem. Phys.* **2018**, *20*, 4666–4675. [[CrossRef](#)] [[PubMed](#)]
- Hnyk, D.; McKee, M. Molecular Structures of Free Boron Clusters. In *Challenges and Advances in Computational Chemistry and Physics, (Boron: The Fifth Element)*; Hnyk, D.; McKee, M. Springer: Dordrecht, The Netherlands, 2016; Volume 20, pp. 17–48.
- Grimmes, R. Supercarboranes. *Angew. Chem. Int. Ed.* **2003**, *42*, 1198–1200. [[CrossRef](#)] [[PubMed](#)]

4. Hnyk, D.; Vřetečka, V.; Drož, L.; Exner, O. Charge Distribution within 1,2-Dicarba-closo-dodecaborane: Dipole Moments of its Phenyl Derivatives. *Collect. Czech. Chem. Commun.* **2001**, *66*, 1375–1379. [[CrossRef](#)]
5. Schleyer, P.v.R.; Najafian, K. Stability and Three-Dimensional Aromaticity of *closo*-Monocarbaborane Anions, *CBn-1Hn-*, and *closo*-Dicarbaboranes, *C2Bn-2Hn*. *Inorg. Chem.* **1998**, *37*, 3454–3470. [[CrossRef](#)] [[PubMed](#)]
6. Baše, T.; Holub, J.; Fanfrlík, J.; Hnyk, D.; Lane, P.D.; Wann, D.A.; Vishnevskiy, Y.V.; Tikhonov, D.; Reuter, C.G.; Mittel, N.W. Icosahedral Carbaboranes with Peripheral Hydrogen-Chalcogenide Groups: Structures from Gas Electron Diffraction and Chemical Shielding in Solution. *Chem. Eur. J.* **2019**, *25*, 2313–2321. [[CrossRef](#)] [[PubMed](#)]
7. Wrackmeyer, B.; García Hernández, Z.; Kempe, R.; Herbehold, M. Novel 1,2-Dicarba-closo-dodecaborane(12) Derivatives of Selenium. *Eur. J. Inorg. Chem.* **2007**, 239–246. [[CrossRef](#)]
8. Arai, K.; Takei, T.; Okumura, M.; Watanabe, S.; Amagai, Y.; Asahina, Y.; Moroder, L.; Hojo, H.; Inaba, K.; Iwaoka, M. Preparation of Selenoinsulin as a Long-Lasting Insulin Analogue. *Angew. Chem. Int. Ed. Engl.* **2017**, *56*, 5522–5526. [[CrossRef](#)] [[PubMed](#)]
9. Herbehold, M.; Milius, W.; Guo-Xin, J.; Kremnitz, W.; Wrackmeyer, B. Molecular Structures of Some Tellurium Derivatives of 1,2-Dicarba-closo-dodecaborane(12). *Z. Anorg. Allg. Chem.* **2006**, *632*, 2031–2036. [[CrossRef](#)]
10. Teixidor, F.; Rudolph, R.W. Synthesis of Cyclic and Polymeric Derivatives of 1,2-Dithiol-*o*-carborane. *J. Organomet. Chem.* **1983**, *241*, 301–312. [[CrossRef](#)]
11. Gleiter, R.; Haberhauer, G.; Werz, D.B.; Roming, F.; Bleiholder, C. From noncovalent Chalcogen-chalcogen Interactions to Supramolecular Aggregates: Experiments and Calculations. *Chem. Rev.* **2018**, *118*, 2010–2014. [[CrossRef](#)]
12. Werz, D.B.; Gleiter, R.; Roming, F. Nanotube Formation Favored by Chalcogen-Chalcogen Interactions. *J. Am. Chem. Soc.* **2002**, *124*, 10638–10639. [[CrossRef](#)] [[PubMed](#)]
13. Werz, D.B.; Staeb, T.H.; Benisch, C.; Rausch, B.J.; Rominger, F.; Gleiter, R. Self-Organization of Chalcogen-Containing Cyclic Alkynes and Alkenes to Yield Columnar Structures. *Org. Lett.* **2002**, *4*, 339–342. [[CrossRef](#)] [[PubMed](#)]
14. Macháček, J.; Plešek, J.; Holub, J.; Hnyk, D.; Vřetečka, V.; Císařová, I.; Kaupp, M.; Štíbr, B. New Route to 1-Thia-closo-dodecaborane(11), *closo*-1-SB11H11, and its Halogenation Reactions. The Effect of the Halogen on the Dipole Moments and the NMR Spectra and the Importance of Spin–Orbit Coupling for the 11B Chemical Shifts. *Dalton Trans.* **2006**, 1024–1029. [[CrossRef](#)] [[PubMed](#)]
15. Fanfrlík, J.; Přáda, A.; Padělková, Z.; Pecina, A.; Macháček, J.; Lepšík, M.; Holub, J.; Růžička, A.; Hnyk, D.; Hobza, P. The Dominant Role of Chalcogen Bonding in the Crystal Packing of 2D/3D Aromatics. *Angew. Chem. Int. Ed.* **2014**, *53*, 10139–10142. [[CrossRef](#)] [[PubMed](#)]
16. Fanfrlík, J.; Hnyk, D. Dihalogen and Pnictogen Bonding in Crystalline Icosahedral Phosphaboranes. *Crystals* **2018**, *8*, 390–399. [[CrossRef](#)]
17. Frisch, M.J.; Trucks, G.W.; Schlegel, H.B.; Scuseria, G.E.; Robb, M.A.; Cheeseman, J.R.; Scalmani, G.; Barone, V.; Mennucci, B.; Petersson, G.A.; et al. *Gaussian 09, Revision, D.01*; Gaussian, Inc.: Wallingford, CT, USA, 2009.
18. Flükiger, P.; Lüthi, H.P.; Portmann, S.; Weber, J. *MOLEKEL 4.3*; Swiss Center for Scientific Computing: Manno, Switzerland, 2000.
19. Portmann, S.; Lüthi, H.P. *MOLEKEL: An Interactive Molecular Graphic Tool*. *CHIMIA Int. J. Chem.* **2000**, *54*, 766–770.
20. Hostaš, J.; Řezáč, J. Accurate DFT-D3 Calculations in a Small Basis Set. *J. Chem. Theory Comput.* **2017**, *13*, 3575–3585. [[CrossRef](#)]
21. Řezáč, J.; Hobza, P. Advanced Corrections of Hydrogen Bonding and Dispersion for Semiempirical Quantum Mechanical Methods. *J. Chem. Theory Comput.* **2012**, *8*, 141–151. [[CrossRef](#)]
22. Jeziorski, B.; Moszynski, R.; Szalewicz, K. Perturbation Theory Approach to Intermolecular Potential Energy Surfaces of van der Waals Complexes. *Chem. Rev.* **1994**, *94*, 1887–1930. [[CrossRef](#)]
23. Parker, T.M.; Burns, L.A.; Parrish, R.M.; Ryno, A.G.; Sherrill, C.D. Levels of Symmetry Adapted Perturbation Theory (SAPT). I. Efficiency and Performance for Interaction Energies. *J. Chem. Phys.* **2014**, *140*, 094106. [[CrossRef](#)]
24. Johnson, E.R.; Keinan, S.; Mori-Sánchez, P.; Contreras-García, J.; Cohen, A.J.; Yang, W. Revealing Noncovalent Interactions. *J. Am. Chem. Soc.* **2010**, *132*, 6498–6506. [[CrossRef](#)] [[PubMed](#)]

25. Contreras-García, J.; Yang, W.; Johnson, E.R. Analysis of Hydrogen-Bond Interaction Potentials from Electron Density: Integration of Noncovalent Interaction Region. *J. Phys. Chem. A* **2011**, *115*, 12983–12990. [[CrossRef](#)] [[PubMed](#)]
26. Ahlrichs, R.; Bar, M.; Haser, M.; Horn, H.; Kolmel, C. Electronic Structure Calculations on Workstation Computers: The Program System Turbomole. *Chem. Phys. Lett.* **1989**, *162*, 165–169. [[CrossRef](#)]
27. Turney, J.M.; Simmonett, A.C.; Parrish, R.M.; Hohenstein, E.G.; Evangelista, F.; Fermann, J.T.; Mintz, B.J.; Burns, L.A.; Wilke, J.J.; Abrams, M.L.; et al. Psi4: An Open-Source Ab Initio Electronic Structure Program. *WIREs Comput. Mol. Sci.* **2012**, *2*, 556–565. [[CrossRef](#)]
28. Stewart, J.P. Optimization of parameters for semiempirical methods IV: Extension of MNDO, AM1, and PM3 to more main group elements. *J. Mol. Model.* **2004**, *10*, 155–164. [[CrossRef](#)]
29. Řezáč, J. Cuby: An Integrative Framework for Computational Chemistry. *J. Comput. Chem.* **2016**, *37*, 1230–1237. [[CrossRef](#)]
30. Clark, T.; Hennemann, M.; Murray, J.S.; Politzer, P. Halogen Bonding: The σ -Hole. *J. Mol. Model.* **2007**, *13*, 291–296. [[CrossRef](#)]
31. Fanfrlík, J.; Hnyk, D. Chalcogens Act as Inner and Outer Heteroatoms in Borane Cages with Possible Consequences for σ -Hole Interactions. *CrystEngComm* **2016**, *47*, 8973–9162. [[CrossRef](#)]
32. Fanfrlík, J.; Holub, J.; Růžicková, Z.; Řezáč, J.; Lane, P.D.; Wann, D.A.; Hnyk, D.; Růžička, A.; Hobza, P. Competition between Halogen, Hydrogen and Dihydrogen Bonding in Brominated Carboranes. *ChemPhysChem* **2016**, *17*, 3373–3376. [[CrossRef](#)]
33. Mantina, M.; Chamberlin, A.C.; Valero, R.; Cramer, C.J.; Truhlar, D.G. Consistent van der Waals Radii for the Whole Main Group. *J. Phys. Chem. A* **2009**, *113*, 5806–5812. [[CrossRef](#)]
34. Bhandary, S.; Sirohiwal, A.; Kadu, R.; Kumar, S.; Chopra, D. Dipsresion Stabilized Se/Te $\cdots\pi$ Double Chalcogen Bonding Synthons in in Situ Cryocrystallized Divalent Organochalcogen Liquids. *Cryst. Growth Des.* **2018**, *18*, 3734–3739. [[CrossRef](#)]
35. Tsuzuki, S.; Sato, N. Origin of Attraction in Chalcogen-nitrogen Interaction in 1,2,5-Chalcogenadiazole Dimers. *J. Phys. Chem. A* **2013**, *117*, 6849–6855. [[CrossRef](#)] [[PubMed](#)]
36. Řezáč, J.; De la Lande, A. On the Role of Charge Transfer in Halogen Bonding. *Phys. Chem. Chem. Phys.* **2017**, *19*, 791–803. [[CrossRef](#)] [[PubMed](#)]
37. Fanfrlík, J.; Pecina, A.; Řezáč, J.; Sedlak, R.; Hnyk, D.; Lepšík, M.; Hobza, P. B-H $\cdots\pi$: A Nonclassical Hydrogen Bond or Dispersion Contact? *Phys. Chem. Chem. Phys.* **2017**, *19*, 18194–18200. [[CrossRef](#)] [[PubMed](#)]
38. Esterhuysen, C.; Heßelman, A.; Clark, T. Trifluoromethyl: An Amphiphilic Noncovalent Bonding Partner. *ChemPhysChem* **2017**, *18*, 772–784. [[CrossRef](#)] [[PubMed](#)]



© 2019 by the authors. Licensee MDPI, Basel, Switzerland. This article is an open access article distributed under the terms and conditions of the Creative Commons Attribution (CC BY) license (<http://creativecommons.org/licenses/by/4.0/>).