



# Synthesis of amorphadiene from dihydroartemisinic acid

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## ARTICLE INFO

### Article history:

Received 9 November 2018

Accepted 22 December 2018

Available online 26 December 2018

### Keywords:

Amorphadiene

Artemisinin

Malaria

Elimination

Microwaves

## ABSTRACT

Amorphadiene is a natural product involved in the biosynthesis of the antimalarial drug artemisinin. A convenient four-step synthesis of amorphadiene, starting from commercially available dihydroartemisinic acid, is reported. The targeted molecule is isolated with an overall yield of 85% on a multi-gram scale in four steps with only one chromatography.

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## 1. Introduction

Amorphadiene (**AD**) is produced in plants by cyclization of farnesyl-pyrophosphate by the enzyme amorphadiene synthase (ADS) (Scheme 1) [1]. **AD** is a key intermediate in the biosynthesis of the antimalarial drug artemisinin [1,2]. In this context, the synthesis of **AD** has been described using a fermentation route (Amyris process) [2–4]. However, **AD** is not yet commercially available. Herein, we report a short and high-yielding gram-scale synthesis of **AD** starting from the commercially available dihydroartemisinic acid, **1**, which is an intermediate in the Sanofi process to prepare artemisinin [5].

With the goal of providing a direct and scalable access to **AD** from the commercially available natural product dihydroartemisinic acid, **1**, a three-step synthetic procedure, relying on the carboxylic acid reduction to the corresponding alcohol **2**, followed by an activation/elimination sequence to produce **AD**, was envisioned (Scheme 2). Bouwmeester et al. described a similar synthetic approach as ours, starting from artemisinic acid, affording **AD** with an overall yield of 25% [6]. However, only a generic route

was reported without detailed procedures, scale, and yield for each individual step.

## 2. Results and discussion

Our attempts for the direct reduction of **1** to **2** focused on the use of lithium aluminium hydride (LiAlH<sub>4</sub>) as a reducing agent. The reaction was first tested using 2.0 equivalents of LiAlH<sub>4</sub> in anhydrous THF at 0 °C (Table 1, entry 1) [7]. However, only moderate conversion of **1** was obtained (60%) and alcohol **2** was isolated in only 25% yield after purification. By increasing the amount of LiAlH<sub>4</sub> in freshly distilled Et<sub>2</sub>O and, after stirring at 23 °C for 2 h, alcohol **2** was obtained with a good yield of 85% without purification (Table 1, entry 2). This yield was further improved to 95% by stirring the reaction mixture for 24 h (Table 1, entry 3). Ultimately, we were able to optimize these conditions using 3.0 equivalents of LiAlH<sub>4</sub> in non-distilled Et<sub>2</sub>O at a concentration of 0.29 M (Table 1, entry 4).

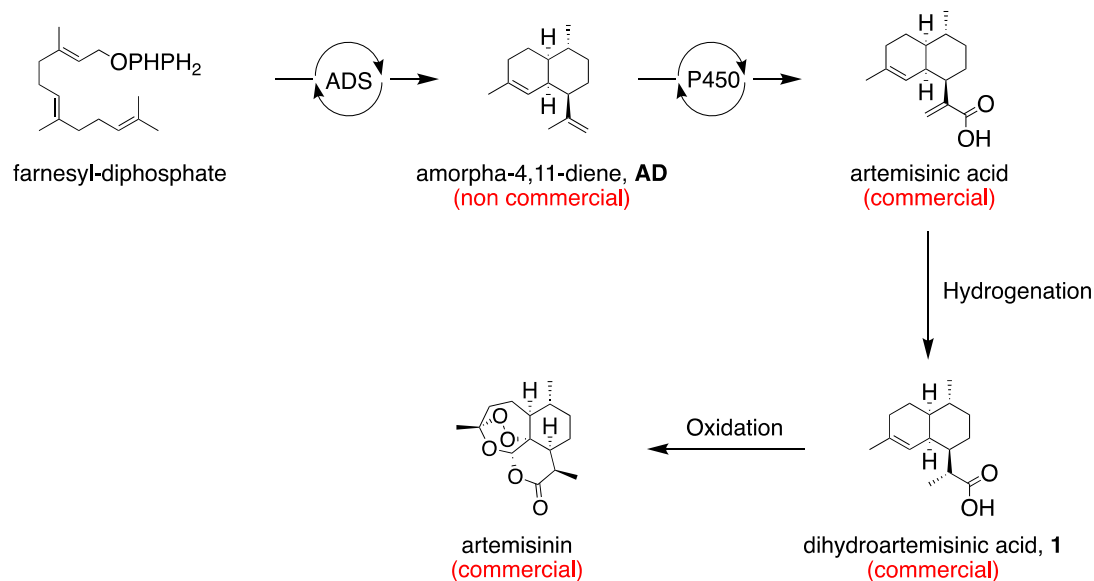
Different conditions were then evaluated to attempt direct conversion of **2** into **AD**. Unfortunately, the direct elimination of the hydroxyl group, using either the Burgess reagent [8] or a one-pot selenide strategy, inspired in the Grieco method [9], failed to give satisfying results even on purified **2** (Scheme 3) [10].

Alternatively, we considered the activation of the alcohol followed by an elimination. Alcohol **2** was quantitatively converted to its corresponding mesylate **3** (MsCl, 1.1 equiv; Et<sub>3</sub>N, 1.5 equiv; in

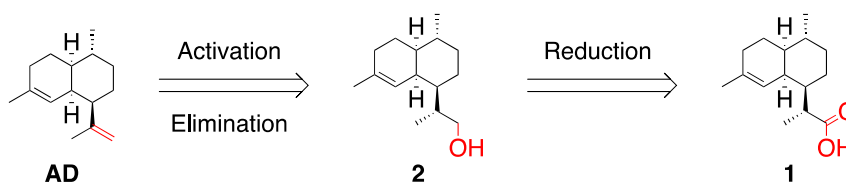
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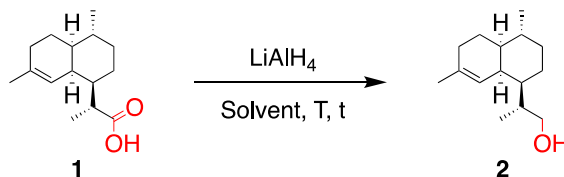


**Scheme 1.** Semi-synthetic approach for the production of artemisinin (Amyris/Sanofi processes).



**Scheme 2.** Formation of amorpha-4,11-diene, **AD**, from dihydroartemisinic acid, **1**.

**Table 1**  
Reduction of **1** to alcohol **2**.



Entry	LiAlH <sub>4</sub> (equiv)	Solvent	T	t [h]	Conversion	Yield <sup>a</sup>
1	2.0	THF (0.18 M)	0 °C	1	ca. 60%	25%
2	5.0	Et <sub>2</sub> O (0.07 M)	0 °C → 23 °C	2	100%	85%
3	5.0	Et <sub>2</sub> O (0.07 M) <sup>b</sup>	0 °C → 23 °C	24	100%	95%
4	3.0	Et <sub>2</sub> O (0.29 M) <sup>b</sup>	0 °C → 23 °C	16	100%	95%

<sup>a</sup> Isolated yields.

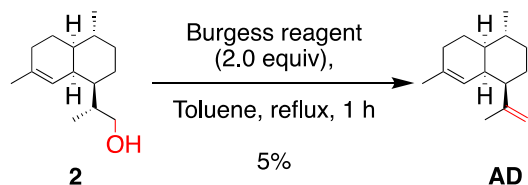
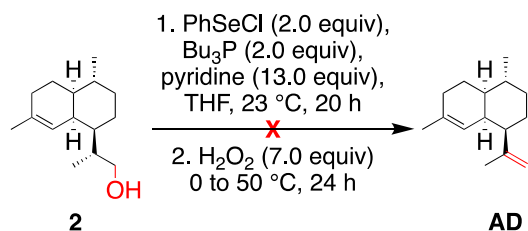
<sup>b</sup> The reaction was conducted with non-distilled solvent.

anhydrous CH<sub>2</sub>Cl<sub>2</sub>, 1 M) (Scheme 4) [11]. Bouwmeester et al. reported the same transformation of **2** to **3** using pyridine as solvent and base, which however, required purification by column chromatography to isolate pure mesylate **3** [6]. We replaced pyridine with dichloromethane as solvent and used only 1.5 equivalents of base, which afforded pure **3** without further purification.

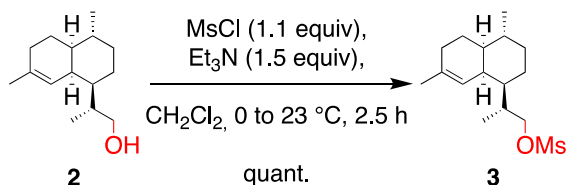
The elimination of the leaving group and formation of the C=C double bond turned out to be more challenging than anticipated. Several conditions were attempted based on related literature procedures (Table 2) [12–14]. As the direct elimination of the mesylate group failed to give satisfying results (Table 2, entry 1), **3** was treated with sodium iodide (NaI, 5.0 equiv) and 1,8-

diazabicyclo[5.4.0]undec-7-ene (DBU, 3.0 equiv) in a one-pot reaction, upon which **AD** was isolated, after column chromatography, in 35% yield (Table 2, entry 2). Alternatively, DBU was added only after completion of the Finkelstein reaction using NaI (2.0–10.0 equiv) but yields in **AD** remained moderate (30–40%; Table 2, entries 3–5).

As the one-pot procedure was not high yielding, another method used by Baran et al. on a similar scaffold as **3**, [15] consisting of isolating the iodo intermediate **4**, after reacting **3** with NaI prior to the addition of base was realized (Scheme 5). The substitution of the mesylate moiety by an iodine proceeded well, affording **4** in yields to 97%.

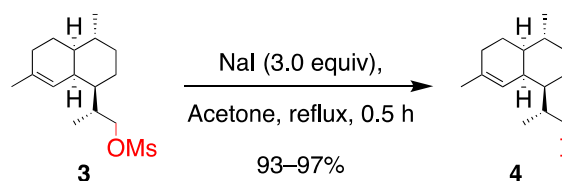


Scheme 3. Attempts for the direct synthesis of AD from alcohol 2.



Scheme 4. Transformation of alcohol 2 to the corresponding mesylate 3.

After isolation, the iodo-compound 4 was reacted with DBU (5.0 equiv) in boiling acetone and AD was isolated in 45% yield (Table 3, entry 1). It is worth noting that 4-hydroxy-4-methylpentan-2-one was also isolated from this reaction, which presumably indicates that a self-condensation of acetone occurred and thus, acetone was not considered as a relevant solvent to carry out this transformation. In THF, a longer reaction time (24 h instead of 5 h) was required to achieve full conversion of 4 (Table 3, entry 2) but the yield in AD remained modest (45%). With these results in hand, we postulated that DBU could promote undesirable side reactions and we decided to use alternative bases. No conversion was observed with Et<sub>3</sub>N (5.0 equiv) (Table 3, entry 3), whereas a moderate conversion was recorded with 5.0 equivalents of *t*-BuOK after 40 h at 65 °C (55%) (Table 3, entry 4). This result was improved by using microwave irradiation and gratifyingly, a complete conversion of 4 was obtained after 1 h producing AD in 89% yield (Table 1, entry 5). Reducing the amount of *t*-BuOK to 3.0 equiv did not give full



Scheme 5. Conversion of mesylate 3 into the iodinated intermediate 4.

conversion of 4 after 1 h in THF (Table 3, entry 6). By using a freshly prepared solution of *t*-BuOK in THF (1 M), the excess of base was reduced to 1.3 equivalents and the reaction time to 0.5 h, affording 1 in an excellent yield of 96% (Table 3, entry 8). More importantly, it allowed us to increase the concentration in 4 to 0.75 M, which also enabled to scale-up the reaction to 4 g of 4 per batch using a standard 20 mL sealed tube.

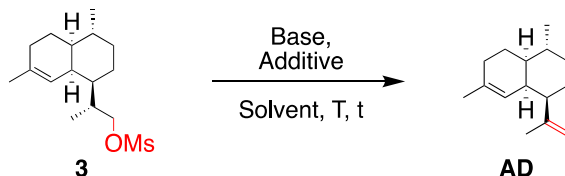
Because thermal heating is usually preferred over microwave-assisted reaction for large scale synthesis, we also investigated the preparation of AD using conventional heating. A moderate conversion (50%) of 4 was achieved when the reaction was performed in *tert*-butanol (*t*-BuOH) at 65 °C (Table 3, entry 9), whereas the iodo intermediate 4 was fully converted to AD when the reaction was carried out in boiling *t*-BuOH (Table 3, entry 10). Ultimately, we were able to perform the elimination with only 2.0 equivalents of *t*-BuOK (on a 6.5 g scale) and to isolate AD in 63% yield (Table 3, entry 11). We found that, in this case, 10% of the alcohol 2 was formed along with other unidentified side products (estimated yield based on <sup>1</sup>H NMR).

### 3. Conclusion

In summary, a screening of conditions allowed us to secure a straightforward access to the important terpene, amorphadiene (AD), from commercially available dihydroartemisinic acid 1 (Scheme 6). The most challenging step was the elimination step, which required considerable optimization. Eventually, 1.3 equiv of *t*-BuOK (1 M in THF) were sufficient to achieve full conversion of 4 in 0.5 h and to produce AD in 96% yield. Alternatively, we also developed a thermal approach in boiling *t*-BuOH affording AD in 63% yield. It is noteworthy to mention that alcohol 2 could be transformed in one step to its corresponding bromo derivative 5 by employing an Appel bromination (see Experimental Section) [16], however, a lower yield (35%) was obtained for the elimination step.

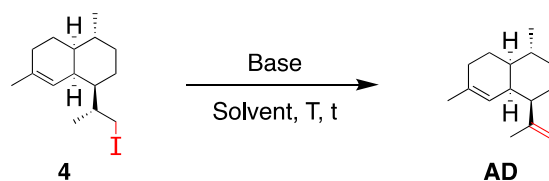
A reliable and straightforward multi-gram scale synthesis of amorphadiene from dihydroartemisinic acid has been established

**Table 2**  
Primary screening experiments for the elimination step.



Entry	Base (equiv)	Additive (equiv)	"steps"	Solvent	T	t [h]	Conversion	Yield
1	DBU (3.0)	—	1	THF	reflux	20	0%	—
2	DBU (3.0)	NaI (5.0)	1	DMF	80 °C → 23 °C	48	100%	35%
3	DBU (3.0)	NaI (2.0)	2	DME	60 °C → reflux	5 then 3	100%	30%
4	DBU (5.0)	NaI (5.0)	2	Acetone	Reflux	2.5 then 13	100%	40%
5	DBU (5.0)	NaI (10.0)	2	Acetone THF	Reflux 65 °C	13 3.5	100%	30%

**Table 3**  
Optimization of the conditions for the elimination step.

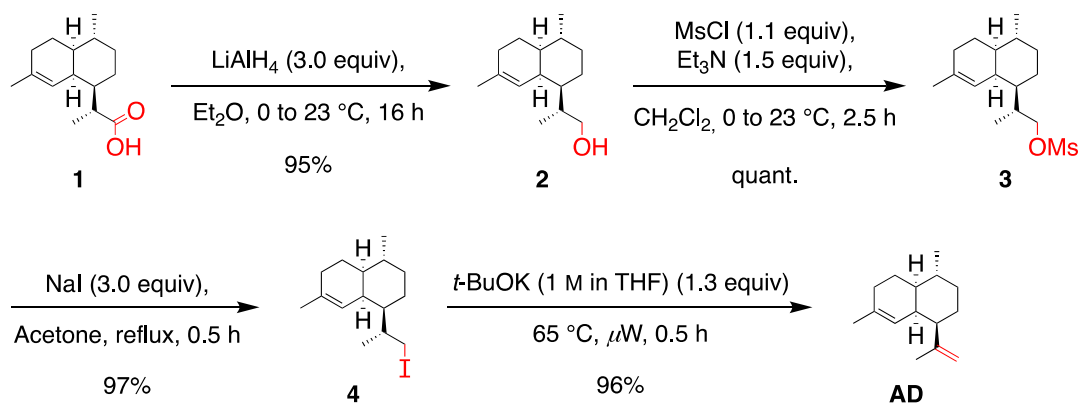


Entry	Base (equiv)	Solvent	[C]	T [°C]	Time [h]	Conversion	Yield
1	DBU (5.0)	Acetone	0.15 M	Reflux	5	100%	45% <sup>a</sup>
2	DBU (5.0)	THF	0.15 M	65	24	100%	45% <sup>a</sup>
3	Et <sub>3</sub> N (5.0)	THF	0.15 M	65	40	0%	—
4	<i>t</i> -BuOK (5.0)	THF	0.15 M	65	40	55%	47% <sup>b</sup>
5	<i>t</i> -BuOK (5.0)	THF	0.15 M	65 ( $\mu$ W) <sup>c</sup>	1	100%	89% <sup>a</sup>
6	<i>t</i> -BuOK (3.0)	THF	0.15 M	65 ( $\mu$ W)	1	55%	47% <sup>b</sup>
7	<i>t</i> -BuOK (6.7) (1 M in THF)	THF	0.15 M	65 ( $\mu$ W)	0.5	100%	90% <sup>b</sup>
8	<i>t</i> -BuOK (1.3) (1 M in THF)	THF	0.75 M	65 ( $\mu$ W)	0.5	100%	96% <sup>a</sup>
9	<i>t</i> -BuOK (5.0)	<i>t</i> -BuOH	0.15 M	65	5	50%	48% <sup>b</sup>
10	<i>t</i> -BuOK (5.0)	<i>t</i> -BuOH	0.15 M	90	5	100%	86% <sup>b</sup>
11	<i>t</i> -BuOK (2.0)	<i>t</i> -BuOH	1.50 M	90	3.5	100%	63% <sup>a</sup>

<sup>a</sup> Isolated yield.

<sup>b</sup> Estimated yield based on the GC/MS chromatogram of crude product.

<sup>c</sup>  $\mu$ W = microwave irradiation.



**Scheme 6.** Synthesis of amorphadiene, **AD**, from dihydroartemisinic acid, **1**.

with an overall yield of 85%. This will enable future scale-up and access to this key molecule for studying its transformation into artemisinic acid and artemisinin derivatives.

## 4. Experimental section

### 4.1. General information

Reagents (Aldrich) were purchased as reagent grade and used without further purification. Reactions in the absence of air and moisture were performed in oven-dried glassware under Ar atmosphere. Flash column chromatography was performed using SiO<sub>2</sub> (60 Å, 230–400 mesh, particle size 0.040–0.063 mm, Merck) at 23 °C with a head pressure of 0.0–0.5 bar. The solvent compositions are reported individually in parentheses. Analytical thin layer chromatography (TLC) was performed on aluminium sheets coated with silica gel 60 F254 (Merck, Macherey-Nagel) or with silica gel 60 RP-18 F254s (Merck, Macherey-Nagel). Visualization was achieved using an alkaline aqueous solution of potassium permanganate (KMnO<sub>4</sub>). Evaporation *in vacuo* was performed at 25–35 °C and 900–10 mbar. Reported yields refer to spectroscopically and chromatographically pure compounds that were dried

under high vacuum (0.1–0.05 mbar) before analytical characterization. <sup>1</sup>H and <sup>13</sup>C nuclear magnetic resonance (NMR) spectra were recorded on a Bruker AV 400 spectrometer at 400 MHz (<sup>1</sup>H) and 101 MHz (<sup>13</sup>C). Chemical shifts  $\delta$  are reported in ppm upfield using the residual deuterated solvent signals as an internal reference (CDCl<sub>3</sub>:  $\delta_{\text{H}} = 7.26$  ppm,  $\delta_{\text{C}} = 77.16$  ppm). For <sup>1</sup>H NMR, coupling constants *J* are given in Hz and the resonance multiplicity is described as s (singlet), d (doublet), t (triplet), q (quartet), m (multiplet), and br (broad). All spectra were recorded at 298 K. Infrared (IR) spectra were recorded on a Bruker Tensor 27 FT-IR spectrometer and are reported as wavenumbers  $\tilde{\nu}$ (cm<sup>-1</sup>). High-resolution mass spectrometry (HRMS) was performed by the Laboratoire de Spectrométrie de Masse from Sorbonne Université, Paris. Gas Chromatography coupled to Mass Spectrometry (GC/MS) analysis was performed on a Shimadzu GCMS-QP2010S using an electronic impact (EI) spectrometer. Low-resolution mass spectra (LRMS) result from ionization by electronic impact (EI-LRMS). The abundance indicated for each mass number (*m/z* values) is given in percentage relative to the strongest peak of 100% abundance (base peak). Melting points were determined using a Büchi melting point apparatus in open capillaries. Nomenclature follows the suggestions proposed by the software ChemDraw Professional 2016.

## 4.2. Synthesis of amorpho-4,11-diene (**AD**)

### 4.2.1. (*R*)-2-((1*R*,4*R*,4*aS*,8*aS*)-4,7-Dimethyl-1,2,3,4,4*a*,5,6,8*a*-octahydronaphthalen-1-yl)propan-1-ol (**2**)

A suspension of  $\text{LiAlH}_4$  (12.2 g, 321.5 mmol) in  $\text{Et}_2\text{O}$  (120 mL) was cooled below  $0^\circ\text{C}$  with an ice/sodium chloride/water bath and treated over 1.5 h with a solution of dihydroartemisinic acid (**1**) (25.0 g, 105.78 mmol) in  $\text{Et}_2\text{O}$  (250 mL) while maintaining the internal temperature below  $7^\circ\text{C}$ . The mixture was maintained in the cold bath, allowed to warm up to  $23^\circ\text{C}$ , stirred for 16 h, cooled below  $0^\circ\text{C}$ , and treated dropwise with  $\text{H}_2\text{O}$  (12.5 mL; 1.0 mL/g of  $\text{LiAlH}_4$ ) while maintaining the internal temperature below  $5^\circ\text{C}$ . The mixture was stirred at  $5^\circ\text{C}$  for 20 min, treated carefully with a 15%  $\text{NaOH}$  solution in  $\text{H}_2\text{O}$  (12.5 mL; 1.0 mL/g of  $\text{LiAlH}_4$ ), stirred for 10 min at  $5^\circ\text{C}$ , and treated with  $\text{H}_2\text{O}$  (37.0 mL; 3.0 mL/g of  $\text{LiAlH}_4$ ) while maintaining the internal temperature below  $5^\circ\text{C}$ . Anhydrous  $\text{MgSO}_4$  was added to absorb the remaining water and the mixture was filtered. The filtrate was evaporated to afford **2** (22.40 g, 95%) as a white solid.

m.p.  $83.9\text{--}84.3^\circ\text{C}$ ;  $R_f = 0.22$  ( $\text{SiO}_2$ ; petroleum ether/ $\text{EtOAc}$  95:05;  $\text{KMnO}_4$ );  $[\alpha]_D^{20} = -8.8^\circ$  (c 1.04,  $\text{CHCl}_3$ );  $^1\text{H NMR}$  (400 MHz,  $\text{CDCl}_3$ ):  $\delta = 5.21$  (br s, 1H), 3.75 (m, 1H), 3.52 (m, 1H), 2.47 (br s, 1H), 2.08–1.72 (m, 3H), 1.69–1.57 (m, 3H), 1.61 (s, 3H), 1.53 (m, 1H), 1.41 (m, 1H), 1.30–1.14 (m, 2H), 1.00 (d,  $J = 6.8$  Hz, 3H), 1.08–0.88 (m, 2H), 0.86 ppm (d,  $J = 6.4$  Hz, 3H), OH signal too weak to be observed;  $^{13}\text{C NMR}$  (101 MHz,  $\text{CDCl}_3$ ):  $\delta = 135.3$ , 120.8, 67.0, 42.8, 42.2, 37.7, 36.8, 35.8, 27.8, 26.8, 26.5, 26.0, 24.0, 20.0, 15.1 ppm; IR (ATR):  $\tilde{\nu} = 3335$ , 3328, 2920, 2865, 1467, 1451, 1432, 1370, 1025, 991, 955  $\text{cm}^{-1}$ ; HR-ESI-MS:  $m/z$ : 245.1875 ( $[\text{M} + \text{Na}]^+$ , calcd for  $\text{C}_{15}\text{H}_{26}\text{O}\text{Na}^+$ : 245.1876); GC/MS:  $m/z$  (%): 222 (11,  $[\text{M}]^+$ ), 191 (14,  $[\text{M} - \text{CH}_3\text{O}]$ ), 163 (100,  $[\text{M} - \text{C}_3\text{H}_7\text{O}]^+$ ), 149 (10), 135 (15), 121 (25), 107 (36), 105 (14). Analytical data correspond to the literature [17].

### 4.2.2. (*R*)-2-((1*R*,4*R*,4*aS*,8*aS*)-4,7-Dimethyl-1,2,3,4,4*a*,5,6,8*a*-octahydronaphthalen-1-yl)propyl methanesulfonate (**3**)

A solution of **2** (18.89 g, 85.0 mmol) in anhydrous  $\text{CH}_2\text{Cl}_2$  (85 mL) was cooled to  $0^\circ\text{C}$ , treated with  $\text{Et}_3\text{N}$  (18.4 mL, 127.5 mmol), stirred for 5 min at  $0^\circ\text{C}$ , and treated dropwise with  $\text{MsCl}$  (7.6 mL, 93.5 mmol) leading to a white precipitate. The mixture was stirred at  $23^\circ\text{C}$  for 2.5 h and diluted with  $\text{CH}_2\text{Cl}_2$  (40 mL). The organic layer was washed with 1 N  $\text{HCl}$  ( $2 \times 20$  mL), dried over  $\text{MgSO}_4$ , filtered, and evaporated to afford mesylate **3** (25.64 g, quant.) as a white crystalline solid.

m.p.  $71.7\text{--}73.5^\circ\text{C}$ ;  $R_f = 0.24$  ( $\text{SiO}_2$ ; petroleum ether/ $\text{EtOAc}$  95:05;  $\text{KMnO}_4$ );  $[\alpha]_D^{20} = -9.8^\circ$  (c 1.02,  $\text{CHCl}_3$ );  $^1\text{H NMR}$  (400 MHz,  $\text{CDCl}_3$ ):  $\delta = 5.14$  (br s, 1H), 4.29 (dd,  $J = 9.5$ , 3.2 Hz, 1H), 4.12 (dd,  $J = 9.5$ , 6.2 Hz, 1H), 3.00 (s, 3H), 2.47 (br s, 1H), 2.00–1.72 (m, 4H), 1.69–1.48 (m, 3H), 1.63 (s, 3H), 1.40 (m, 1H), 1.31–1.16 (m, 2H), 1.04 (d,  $J = 6.8$  Hz, 3H), 1.02–0.87 (m, 2H), 0.86 ppm (d,  $J = 6.5$  Hz, 3H);  $^{13}\text{C NMR}$  (101 MHz,  $\text{CDCl}_3$ ):  $\delta = 135.9$ , 120.0, 74.1, 42.5, 42.0, 37.4, 37.3, 35.5, 34.6, 27.7, 26.8, 26.3, 25.9, 24.0, 19.8, 15.2 ppm; IR (ATR):  $\tilde{\nu} = 2911$ , 2868, 1449, 1353, 1175, 1110, 953  $\text{cm}^{-1}$ ; HR-ESI-MS:  $m/z$ : 323.1650 ( $[\text{M} + \text{Na}]^+$ , calcd for  $\text{C}_{16}\text{H}_{28}\text{O}_3\text{SNa}^+$ : 323.1651); GC/MS:  $m/z$  (%): 300 (11,  $[\text{M}]^+$ ), 163 (100,  $[\text{M} - \text{C}_4\text{H}_9\text{O}_3\text{S}]^+$ ), 147 (17), 133 (10), 121 (29), 119 (19), 107 (31), 105 (20).

### 4.2.3. (1*R*,4*R*,4*aS*,8*aS*)-1-((*R*)-1-Iodopropan-2-yl)-4,7-dimethyl-1,2,3,4,4*a*,5,6,8*a*-octahydronaphthalene (**4**)

A solution of mesylate **3** (12.03 g, 40.0 mmol) in acetone (80 mL) was treated with  $\text{NaI}$  (18.0 g, 120.1 mmol), stirred at  $70^\circ\text{C}$  for 0.5 h, cooled to  $23^\circ\text{C}$ , diluted with  $\text{Et}_2\text{O}$  (100 mL) and  $\text{H}_2\text{O}$  (50 mL). The two layers were separated and the aqueous layer was extracted with  $\text{Et}_2\text{O}$  ( $6 \times 50$  mL). The combined organic layers were washed twice with a  $\text{Na}_2\text{S}_2\text{O}_3$  sat. aqueous solution (100 + 50 mL), dried over  $\text{MgSO}_4$ , filtered, and evaporated to afford **4** (12.33 g, 93%) as a

light yellow oil that crystallizes upon cooling in a freezer.

$R_f = 0.70$  ( $\text{SiO}_2$ ; hexane;  $\text{KMnO}_4$ );  $[\alpha]_D^{20} = -32.0^\circ$  (c 1.02,  $\text{CHCl}_3$ );  $^1\text{H NMR}$  (400 MHz,  $\text{CDCl}_3$ ):  $\delta = 5.17$  (br s, 1H), 3.44 (dd,  $J = 9.7$ , 2.5 Hz, 1H), 3.30 (dd,  $J = 9.7$ , 5.4 Hz, 1H), 2.49 (br s, 1H), 2.03–1.84 (m, 2H), 1.79 (m, 1H), 1.64 (m, 1H), 1.62 (s, 3H), 1.59–1.48 (m, 2H), 1.41 (m, 1H), 1.33–1.10 (m, 3H), 0.99 (d,  $J = 6.4$  Hz, 3H), 0.97–0.89 (m, 2H), 0.87 ppm (d,  $J = 6.5$  Hz, 3H);  $^{13}\text{C NMR}$  (101 MHz,  $\text{CDCl}_3$ ):  $\delta = 135.8$ , 121.1, 46.0, 42.1, 37.7, 35.6, 34.8, 27.7, 26.8, 26.1, 25.9, 24.0, 20.2, 19.9, 19.0 ppm; IR (ATR):  $\tilde{\nu} = 2906$ , 2867, 1447, 1433, 1377, 1303, 1291, 1254, 1227, 1191, 1171, 1159, 1110, 1031, 990, 957, 942, 924  $\text{cm}^{-1}$ ; HR-ESI-MS: molecule not ionized in ESI; GC/MS:  $m/z$  (%): 332 (4,  $[\text{M}]^+$ ), 163 (100,  $[\text{M} - \text{C}_3\text{H}_6]^{+}$ ), 121 (19), 107 (25), 105 (7).

### 4.2.4. (1*R*,4*R*,4*aS*,8*aS*)-1-((*R*)-1-Bromopropan-2-yl)-4,7-dimethyl-1,2,3,4,4*a*,5,6,8*a*-octahydronaphthalene (**5**)

A solution of **2** (100 mg, 0.45 mmol) in  $\text{CH}_2\text{Cl}_2$  (2.5 mL) was treated with  $\text{PPh}_3$  (236 mg, 0.90 mmol), cooled to  $0^\circ\text{C}$ , treated with  $\text{CBr}_4$  (375 mg, 1.13 mmol), stirred at  $0^\circ\text{C}$  for 1 h, and evaporated. The crude was absorbed on  $\text{SiO}_2$  and column chromatography ( $\text{SiO}_2$  pre-treated with hexane/ $\text{Et}_3\text{N}$  99:01; hexane) gave **5** (128 mg, quant.) as a colorless oil.

$R_f = 0.60$  ( $\text{SiO}_2$ ; hexane;  $\text{KMnO}_4$ );  $[\alpha]_D^{20} = -16.1^\circ$  (c 0.94,  $\text{CHCl}_3$ );  $^1\text{H NMR}$  (400 MHz,  $\text{CDCl}_3$ ):  $\delta = 5.17$  (br s, 1H), 3.61 (dd,  $J = 9.9$ , 2.6 Hz, 1H), 3.49 (dd,  $J = 9.9$ , 5.6 Hz, 1H), 2.49 (br s, 1H), 2.02–1.74 (m, 4H), 1.70–1.48 (m, 3H), 1.63 (s, 3H), 1.42 (m, 1H), 1.35–1.18 (m, 2H), 1.05 (d,  $J = 6.5$  Hz, 3H), 1.03–0.90 (m, 2H), 0.87 ppm (d,  $J = 6.5$  Hz, 3H);  $^{13}\text{C NMR}$  (101 MHz,  $\text{CDCl}_3$ ):  $\delta = 135.7$ , 120.7, 44.2, 42.5, 42.1, 37.6, 35.8, 35.6, 27.7, 26.8, 26.1, 25.9, 24.0, 19.9, 17.0 ppm; IR (ATR):  $\tilde{\nu} = 2907$ , 2869, 1447, 1435, 1379, 1334, 1306, 1290, 1261, 1236, 1218, 1185, 1159, 1140, 1110, 1075, 1037, 991, 958, 942, 925  $\text{cm}^{-1}$ ; HR-ESI-MS: molecule not ionized in ESI; GC/MS:  $m/z$  (%): 286 (4,  $[\text{M}]^+$ ), 284 (4,  $[\text{M}]^+$ ), 163 (100,  $[\text{M} - \text{C}_3\text{H}_6\text{Br}]^+$ ), 121 (18), 107 (22), 105 (5).

### 4.2.5. (1*R*,4*R*,4*aS*,8*aR*)-4,7-Dimethyl-1-(prop-1-en-2-yl)-1,2,3,4,4*a*,5,6,8*a*-octahydronaphthalene (amorpho-4,11-diene, **AD**)

**4.2.5.1. Microwave-assisted approach.** Iodo derivative **4** (11.85 g, 35.7 mmol) was treated with *t*-BuOK (1 M in THF) (46.4 mL, 46.4 mmol), stirred at  $65^\circ\text{C}$  in a microwave oven for 30 min, diluted with  $\text{Et}_2\text{O}$  (150 mL) and  $\text{H}_2\text{O}$  (150 mL). The two layers were separated and the aqueous layer was extracted with  $\text{Et}_2\text{O}$  ( $4 \times 75$  mL). The combined organic layers were washed brine, dried over  $\text{MgSO}_4$ , filtered, and evaporated. Column chromatography ( $\text{SiO}_2$ ; hexane) gave **AD** (7.03 g, 96%) as a colorless oil.

**4.2.5.2. Thermal approach.** A solution of iodo derivative **4** (6.46 g, 19.43 mmol) in melted *t*-BuOH (13 mL) was treated with *t*-BuOK (4.4 g, 39.2 mmol), stirred at reflux for 3.5 h, diluted with  $\text{Et}_2\text{O}$  (60 mL) and  $\text{H}_2\text{O}$  (60 mL). The two layers were separated and the aqueous layer was further extracted with  $\text{Et}_2\text{O}$  ( $4 \times 30$  mL). The combined organic layers were dried over  $\text{MgSO}_4$ , filtered, and evaporated. Column chromatography ( $\text{SiO}_2$ ; cyclohexane) gave **AD** (2.52 g, 63%) as a colorless oil.

$R_f = 0.77$  ( $\text{SiO}_2$ ; hexane;  $\text{KMnO}_4$ );  $R_f = 0.34$  (RP-18; MeOH;  $\text{KMnO}_4$ );  $[\alpha]_D^{20} = -13.9^\circ$  (c 1.00,  $\text{CHCl}_3$ ) ([18]);  $[\alpha]_D^{20} = -14.0^\circ$  (c 0.4,  $\text{CHCl}_3$ );  $^1\text{H NMR}$  (400 MHz,  $\text{CDCl}_3$ ):  $\delta = 5.06$  (br s, 1H), 4.87 (br s, 1H), 4.64 (br s, 1H), 2.55 (br s, 1H), 2.06–1.83 (m, 3H), 1.77 (m, 1H), 1.74 (s, 3H), 1.67 (dq,  $J = 12.9$ , 3.5 Hz, 1H), 1.60 (s, 3H), 1.59–1.46 (m, 2H), 1.40 (m, 1H), 1.35–1.18 (m, 2H), 0.98 (qd,  $J = 12.8$ , 3.4 Hz, 1H), 0.89 ppm (d,  $J = 6.4$  Hz, 3H);  $^{13}\text{C NMR}$  (101 MHz,  $\text{CDCl}_3$ ):  $\delta = 148.5$ , 135.1, 121.3, 110.2, 48.1, 42.3, 38.1, 35.9, 28.3, 26.9, 26.5, 26.3, 24.1, 23.1, 20.3 ppm; IR (ATR):  $\tilde{\nu} = 2918$ , 2866, 1643, 1446, 1376, 1240, 1173, 1139, 1108, 993, 968, 936  $\text{cm}^{-1}$ ; HR-ESI-MS: molecule not ionized in ESI; GC/MS:  $m/z$  (%): 204 (45,  $[\text{M}]^+$ ), 189 (73,  $[\text{M} - \text{CH}_3]^+$ ), 119 (100). Analytical data correspond to the literature [18].

## Acknowledgements

This work was supported by the Bill & Melinda Gates Foundation [OPP1190174]. We would like to thank Dr. Kai Rossen for his advice and for helpful discussions. We also thank Cédric Przybylski at the Laboratoire de Spectrométrie de Masse of the Sorbonne Université, Paris, for HRMS measurements.

## Appendix A. Supplementary data

Supplementary data to this article can be found online at <https://doi.org/10.1016/j.tet.2018.12.050>.

## References

- [1] G.D. Brown, *Molecules* 15 (2010) 7603–7698.
- [2] P. J. Westfall, D.J. Pitera, J.R. Lenihan, D. Eng, F.X. Woolard, R. Regentin, T. Horning, H. Tsuruta, D.J. Melis, A. Owens, S. Fickes, D. Diola, K.R. Benjamin, J.D. Keasling, M.D. Leavell, D.J. McPhee, N.S. Renninger, J.D. Newman, C.J. Paddon, *Proc. Natl. Acad. Sci.* 109 (2012) E111–E118.
- [3] C.J. Paddon, P.J. Westfall, D.J. Pitera, K. Benjamin, K. Fisher, D. McPhee, M.D. Leavell, A. Tai, A. Main, D. Eng, D.R. Polichuk, K.H. Teoh, D.W. Reed, T. Treynor, J. Lenihan, M. Fleck, S. Bajad, G. Dang, D. Dengrove, D. Diola, G. Dorin, K.W. Ellens, S. Fickes, J. Galazzo, S.P. Gaucher, T. Geistlinger, R. Henry, M. Hepp, T. Horning, T. Iqbal, H. Jiang, L. Kizer, B. Lieu, D. Melis, N. Moss, R. Regentin, S. Secrest, H. Tsuruta, R. Vazquez, L.F. Westblade, L. Xu, M. Yu, Y. Zhang, L. Zhao, J. Lievense, P.S. Covello, J.D. Keasling, K.K. Reiling, N.S. Renninger, J.D. Newman, *Nature* 496 (2013) 528–532.
- [4] S.D. Kung, S. Lund, A. Murarka, D. McPhee, C.J. Paddon, *Front. Plant Sci.* 9 (2018) 1–7.
- [5] J. Turconi, F. Griolet, R. Guevel, G. Oddon, R. Villa, A. Geatti, M. Hvala, K. Rossen, R. Göller, A. Burgard, *Org. Process Res. Dev.* 18 (2014) 417–422.
- [6] H.J. Bouwmeester, T.E. Wallaart, M.H.A. Janssen, B. van Loo, B.J.M. Jansen, M.A. Posthumus, C.O. Schmidt, J.-W.D. Kraker, W.A. König, M.C.R. Franssen, *Phytochemistry* 52 (1999) 843–854.
- [7] A.J. Pearson, H. Paramahamsan, J.D. Dudones, *Org. Lett.* 6 (2004) 2121–2124.
- [8] R.E. Dolle, K.C. Nicolaou, *J. Am. Chem. Soc.* 107 (1985) 1691–1694.
- [9] G.S. Weatherhead, G.A. Cortez, R.R. Schrock, A.H. Hoveyda, *Proc. Natl. Acad. Sci.* 101 (2004) 5805–5809.
- [10] In both cases, only traces amounts of **AD** were detected (<5%)
- [11] In parallel, **2** was also converted into the corresponding tosylate. The synthesis was easily achieved by treatment of **2** with TsCl (2 equiv) in CH<sub>2</sub>Cl<sub>2</sub> at 23 °C in the presence of 4-DMAP (2.4 equiv). The tosyl derivative was isolated in 99% yield without further purification. However, the subsequent elimination step was low yielding.
- [12] M.J. Williams, H.L. Deak, M.L. Snapper, *J. Am. Chem. Soc.* 129 (2007) 486–487.
- [13] Y.C. Jung, C.H. Yoon, E. Turos, K.S. Yoo, K.W. Jung, *J. Org. Chem.* 72 (2007) 10114–10122.
- [14] T. Tokoroyama, K. Fujimori, T. Shimizu, Y. Yamagiwa, M. Monden, H. Iio, *Tetrahedron* 44 (1988) 6607–6622.
- [15] T.J. Maimone, Y. Ishihara, P.S. Baran, *Tetrahedron* 71 (2015) 3652–3665.
- [16] S. Rana, H.S. Hong, L. Barrigan, L.W. Jin, D.H. Hua, *Bioorg. Med. Chem. Lett.* 19 (2009) 670–674.
- [17] L.-K. Sy, K.-S. Ngo, G.D. Brown, *Tetrahedron* 55 (1999) 15127–15140.
- [18] K.-S. Ngo, G.D. Brown, *Tetrahedron* 55 (1999) 15099–15108.