

## Supramolecular Chemistry

## Pathway Dependence in Redox-Driven Metal–Organic Gels

Santanu Panja and Dave J. Adams\*<sup>[a]</sup>

**Abstract:** Pathway dependence is common in self-assembly. Herein, the importance of pathway dependence for redox-driven gels is shown by constructing a Fe<sup>II</sup>/Fe<sup>III</sup> redox-based metal–organic gel system is shown. In situ oxidation of the Fe<sup>II</sup> ions at different rates results in conversion of a Fe<sup>II</sup> gel into a Fe<sup>III</sup> organic gel, which controls the material properties, such as gel stiffness, gel strength, and an unusual swelling behaviour, is described. The rate of formation of Fe<sup>III</sup> ions determines the extent of intermolecular interactions and so whether gelation or precipitation occurs.

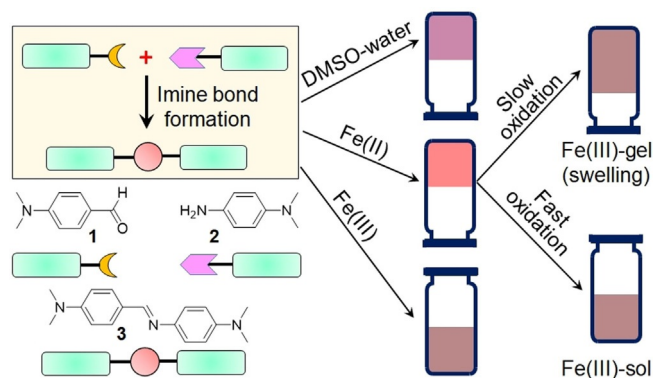
Supramolecular low molecular weight hydrogels (LMWGs) formed by the self-assembly of small organic molecules induced by non-covalent interactions are fascinating smart materials, which have multifunctional applications.<sup>[1]</sup> Of the different kinds of supramolecular gels, metal–organic gels have received significant interest in recent years because of their widespread applications, particularly in optoelectronics, pharmaceuticals and catalysis.<sup>[2]</sup> Metal–organic gels are a special class of supramolecular gels that incorporate a metallic element into the ligand during self-assembly. Conceptually, metal–organic gels are synthesized based on strong metal–ligand interactions, in which the organic ligand may compose of a single component or can be derived from the reaction between multiple functional groups. Incorporation of metal ions into the organic frameworks often dramatically changes the optical and chemical properties of the ligand and therefore can be used as a powerful strategy to modify the material properties.<sup>[2a,b,3]</sup>

One interesting property of supramolecular gels is their responsiveness towards various stimuli including heat, pH, irradiation, chemical entities, and redox reactions.<sup>[3b,4]</sup> Redox responses make the gel systems desirable for biomimicry, as well as for numerous possible applications.<sup>[4f,5]</sup> However, most of the redox-fuelled gels found in the literature are polymeric in nature and are usually developed from the intermolecular di-

sulfide exchange reaction-based molecular systems.<sup>[4f,5c–e,g]</sup> Therefore, designing and construction of new supramolecular low molecular weight redox-based gels is highly desirable; these are expected to have very different underlying properties.<sup>[6]</sup>

A key issue for many supramolecular gels is that the properties significantly depend on the preparative pathway.<sup>[7]</sup> Because of such effect, even though the composition of the final materials remains same, the material properties can vary depending on the self-assembly kinetics.<sup>[7a]</sup> Gels formed at a high rate are often kinetically trapped, which means that they can be hard to reproduce and control. To avoid this kinetic trapping during the gelation process, the environmental conditions need to be well controlled to achieve homogeneous and reproducible gels.

Herein, we designed a new redox-responsive metal–organic hydrogel system and discuss the pathway dependence of these redox-based gels (Figure 1). Unlike other redox systems, instead of using sulfide/disulfide-based ligands,<sup>[4f,5c,d,g]</sup> herein, we utilize dynamic imine bond formation between an aldehyde (1) and an amine (2) as the key chemical reaction to synthesize the ligand (3). To make the organic framework redox responsive, we incorporated Fe<sup>II</sup> ions into the gel medium. In situ oxidation of the Fe<sup>II</sup> ions by an oxidising agent results in formation of a Fe<sup>III</sup>–organic gel. The final properties of the Fe<sup>III</sup> gel significantly depend upon the rate of oxidation of Fe<sup>II</sup>. Although a slow rate of oxidation gives Fe<sup>III</sup> gels with high stiffness, a very fast oxidation drives the system towards precipitation. Precipitation also occurred on direct treatment of the mixture of the aldehyde (1) and amine (2) with Fe<sup>III</sup>. Hence, we showed that we can prepare Fe<sup>III</sup> metallogels, which cannot be prepared directly by controlling the reaction pathway. In some



**Figure 1.** Cartoon representing the phase transformations of the mixture of 1 and 2 under different conditions (the pictures of the inverted vials represent gel states).

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Supporting information and the ORCID identification number(s) for the author(s) of this article can be found under:  
<https://doi.org/10.1002/chem.202001051>; it contains experimental details.

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cases, we also find that the materials exhibit a highly unusual swelling, which is very uncommon for such supramolecular gels.

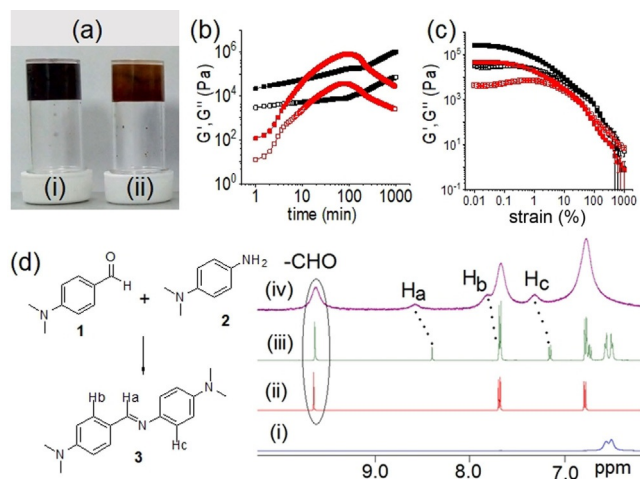
To prepare the gel, we employed dynamic imine bond formation<sup>[8]</sup> reaction between 4-(dimethylamino)benzaldehyde (**1**) and *N,N*-dimethyl-*p*-phenylenediamine (**2**) in DMSO/H<sub>2</sub>O (25:75 v/v). When a mixture of equimolar amount of **1** and **2** (0.134 M) in DMSO is diluted with water, a brown self-supporting gel was rapidly formed (Figure 2a). The gelation process was followed by rheology. By time sweep rheology, initially the storage modulus (*G'*) was considerably higher than the loss modulus (*G''*), indicating that gelation was very quick and occurred before the measurement could be begun (Figure 2a). The gel continued to evolve with time and *G'* and *G''* reach a plateau after approximately 14 hours (Figures 2b and S1 in the Supporting Information). The gel exhibits a high stiffness ( $\approx 2 \times 10^5$  Pa), but starts to collapse at a low strain of approximately 0.2% (the critical strain; Figures 2c and S1).

To characterize the chemical component responsible for gelation, <sup>1</sup>H NMR spectroscopy and high-resolution mass spectroscopy (HRMS) of the gel state were performed (Figure 2d). By <sup>1</sup>H NMR spectroscopy, the appearance of a new peak at 8.39 ppm clearly demonstrated the imine bond formation between the aldehyde and amine. Integration of the <sup>1</sup>H NMR spectra showed around 26% conversion to the imine **3** after 16 hours. This presumably represents the position of the equilibrium of the reaction under these conditions, because imines are susceptible to hydrolysed in water.<sup>[8,9]</sup> The imine bond formation was further confirmed by recording the HRMS of the gel (Figure S2 in the Supporting Information). The appearance of the mass at 268.1792 shows the formation of compound **3**

[expected mass = 268.1814 for the formula (*M*+H)<sup>+</sup>] in the gel state. Moreover, by FTIR spectroscopy, the stretching signal for the aldehyde carbonyl of **1** appeared at 1656 cm<sup>-1</sup>, whilst in the gel state it remained almost unaffected (Figure S3 in the Supporting Information). However, a broad peak appeared at 1680 cm<sup>-1</sup> for the C=N bond, again confirming the formation of the imine **3** in the gel state.

We incorporated Fe<sup>II</sup> ions (as sulfate salt) into the gel medium to convert this supramolecular gel into a redox responsive metal-organic gel.<sup>[2c,5a,f,10]</sup> We used 0.134 M of Fe<sup>II</sup> to prepare the Fe<sup>II</sup> gel (1 molar equivalent with respect to the aldehyde). Addition of aqueous solution of Fe<sup>II</sup> to the mixture of **1** and **2** not only modified the gelation kinetics, but also changed the final mechanical properties of the gels. Time sweep rheology indicated that the initial values of both *G'* and *G''* were significantly lower compared to the case when Fe<sup>II</sup> was absent (Figures 2b and S1). With time, both *G'* and *G''* started to increase rapidly. After approximately 2 hours, *G'* and *G''* started to decrease and became almost constant after approximately 13 hours (Figure S1 in the Supporting Information). The presence of Fe<sup>II</sup> resulted in around approximately six times decrease in both *G'* and *G''* of the gel (Figures 2c and S1). However, no significant change in gel strength (critical strain) was observed (Figure 2c). Surprisingly, when we tried to make a control gel with preformed imine **3**, no gel formation was noticed either in absence or presence of Fe<sup>II</sup> (Figure S4 in the Supporting Information). Compound **3** is poorly soluble in DMSO. Upon addition of H<sub>2</sub>O to a suspension of **3** in DMSO, a yellow and orange yellow precipitate was formed in absence and presence of Fe<sup>II</sup> respectively. Hence, in situ formation of **3** is necessary for gelation to occur.

The presence of Fe<sup>II</sup> also changed the visual appearance of the gel (Figure 2a). The colour of the gel changed from brown to reddish brown in presence of Fe<sup>II</sup>. UV/Vis and emission spectroscopy measurements of **1** and **2** were conducted under different conditions to highlight the aggregation properties (Figures S5 and S6 in the Supporting Information). By UV/Vis analysis, **1** and **2** exhibited absorption at 353 and 305 nm, respectively. In comparison, the gel state obtained from the mixture of **1** and **2** showed a strong absorption at 338 nm with two shoulder peaks at 310 and 434 nm. Time-dependent emission experiments showed that as the reaction proceeds, the strong emission of the aldehyde at 410 nm started to decrease, and a new band appeared at 468 nm. In presence of Fe<sup>II</sup>, the absorption peak at 310 nm became more intense, whereas the shoulder at 434 nm remained unaffected. By fluorescence, the emission of the gel at 468 nm blueshifted by 8 nm in presence of Fe<sup>II</sup> along with the generation of a new band at 550 nm. These data suggest existence of different intermolecular interactions in the gel matrices formed in absence and presence of Fe<sup>II</sup>. To confirm this, the <sup>1</sup>H NMR spectrum of the Fe<sup>II</sup> gel was superimposed with that obtained in absence of Fe<sup>II</sup> (Figure 2d). Comparison of the data shows that the signal for the imine proton H<sub>a</sub> of **3** moved to the downfield region by 0.2 ppm due to the interaction with Fe<sup>II</sup>. Moreover, due to metal coordination, the aromatic protons H<sub>b-c</sub> also showed approximately 0.2 ppm downfield shift. Interestingly, while the



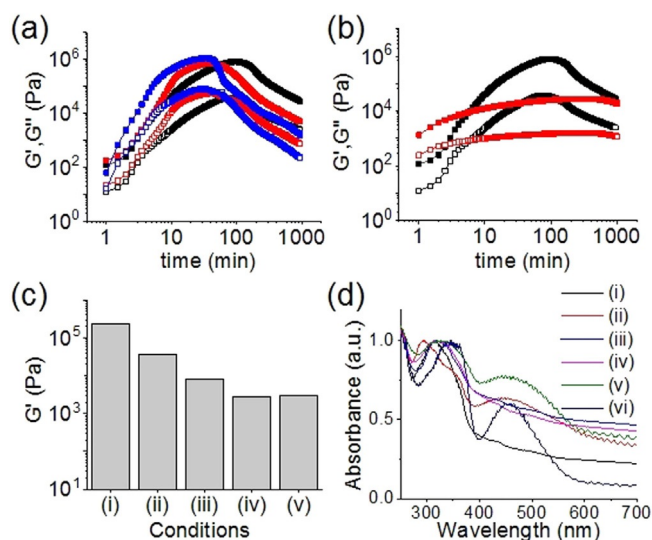
**Figure 2.** (a) Photograph of the hydrogels obtained from the mixture of **1** and **2** in absence (i) and presence (ii) of Fe<sup>II</sup>. (b) Variation of *G'* (closed symbol) and *G''* (open symbol) with time for the mixture of **1** and **2** in absence (black) and presence (red) of Fe<sup>II</sup>. (c) Strain sweeps of the gels prepared in absence (black) and presence (red) of Fe<sup>II</sup> after 16 hours. The closed symbols represent *G'*, open symbols *G''*. (d) Partial <sup>1</sup>H NMR (in [D<sub>6</sub>]DMSO) spectra of (i) **2**; (ii) **1**; (iii) gels obtained from the mixture of **1** and **2** in absence of Fe<sup>II</sup>; and (iv) gels obtained from the mixture of **1** and **2** in presence of Fe<sup>II</sup>. In all cases (a–d), the initial concentrations of **1**, **2**, and Fe<sup>II</sup> are 0.134 M.

signals for aromatic protons, as well as the carbonyl  $-\text{CH}$  of **1**, exhibited no shift in  $^1\text{H}$  NMR upon interaction with  $\text{Fe}^{\text{II}}$ , the aromatic protons of **2** became broad and shifted downfield by 0.2 ppm. These results indicate that  $\text{Fe}^{\text{II}}$  not only binds with the imine bond of **3**, but also interacts with the amine functionality of **2**. By HRMS, no evidence of formation of **3**- $\text{Fe}^{\text{II}}$  complex was found (Figure S7 in the Supporting Information). This indicates that although the interaction of the imine bond with  $\text{Fe}^{\text{II}}$  ion is subtle, it causes significant change at the macroscopic level.

The presence of  $\text{Fe}^{\text{II}}$  ions makes the gel medium redox responsive.<sup>[2c,5a,f,10]</sup> Practical uses of  $\text{Fe}^{\text{II}}$ / $\text{Fe}^{\text{III}}$  redox systems involving LMWGs are limited in the literature. For example, recently, Das et al. reported a reusable transient hydrogel system based on  $\text{Fe}^{\text{II}}$ / $\text{Fe}^{\text{III}}$  redox conversion and explored those transient aggregates in mimicking peroxidase activity.<sup>[5a]</sup> Panja and Ghosh utilized a  $\text{Fe}^{\text{II}}$  metallogel for visual recognition of  $\text{H}_2\text{O}_2$  from other reactive oxygen species (ROS) by performing Fenton reaction inside the gel medium.<sup>[11]</sup> Inspired by their work, we attempted to convert our  $\text{Fe}^{\text{II}}$  gel into a  $\text{Fe}^{\text{III}}$  gel through an in situ oxidation of the  $\text{Fe}^{\text{II}}$  ions by different oxidizing agents.

Prior to this, we investigated the role of dissolved oxygen on our  $\text{Fe}^{\text{II}}$  gel. For this purpose, instead of deionized water, we used degassed, deionized water to prepare the gel. Rheological studies showed that the rheological moduli, as well as viscosity, follow similar trends as in the case with normal water. Interestingly, final values of both  $G'$  and  $G''$  of the gel formed with normal water are considerably lower than the gel formed with degassed water (Figure S8 in the Supporting Information). However, no significant change in the gel strength (the critical strain) was observed. These results point out that the dissolved oxygen has a subtle effect on the stability of  $\text{Fe}^{\text{II}}$  ions and presumably oxidise some  $\text{Fe}^{\text{II}}$  ions into  $\text{Fe}^{\text{III}}$  ions inside the gel medium, resulting in the decrease in stiffness of the material (by  $\approx 3$  times). However, no significant change in the absorption and emission spectra of the gels were noticed (Figure S9).

Next, we used  $\text{NaNO}_2$  (0.067 M) as an in situ oxidizing agent and monitored the self-assembly kinetics by time sweep rheology. Because the  $\text{NaNO}_2$  is a mild oxidant, it causes slow conversion of  $\text{Fe}^{\text{II}}$  ions into  $\text{Fe}^{\text{III}}$ . By time sweep rheology, at the early stages, a slight increase in the rate of increase of both  $G'$  and  $G''$  was noticed (Figure 3a). Interestingly, after reaching the maxima, the rheological moduli started to decrease earlier than the case with no oxidizing agent before the values become almost constant after approximately 12 hours. Viscosity data recorded over time follows a similar trend as that of rheology (Figure S10 in the Supporting Information). Instead of  $\text{NaNO}_2$ , when same concentration of  $\text{H}_2\text{O}_2$  (0.067 M) was used, which is a stronger oxidant, the self-assembly kinetics behave differently (Figures 3b and S10 in the Supporting Information). In this case, the variation of the rheological moduli was straightforward, in which  $G'$  and  $G''$  increase as the aggregation proceeds and finally reached the plateau after approximately three hours. However, in both cases, formation of  $\text{Fe}^{\text{III}}$  resulted in a significant decrease in the stiffness of the final gels, whereas the extent of reduction in the values of  $G'$  depends on the rate of oxidation of  $\text{Fe}^{\text{II}}$  (Figures 3c, S11, S12, and



**Figure 3.** Variation of  $G'$  (closed symbol) and  $G''$  (open symbol) with time for the mixture of **1**, **2** and  $\text{Fe}^{\text{II}}$  in presence of redox reaction involving (a)  $\text{NaNO}_2$  and (b)  $\text{H}_2\text{O}_2$ . The black data is for no oxidizing agent, the red data for 0.067 M and the blue data for 0.134 M oxidising agent. (c) Bar graph representing the stiffness ( $G'$ ) of the final gels obtained from: (i) the mixture of **1** and **2**; (ii) the mixture of **1**, **2** and  $\text{Fe}^{\text{II}}$ ; (iii)–(v) the mixture of **1**, **2** and  $\text{Fe}^{\text{II}}$  in presence of redox reaction involving 0.067 M of  $\text{NaNO}_2$  (iii); 0.067 M of  $\text{H}_2\text{O}_2$  (iv) and 0.134 M of  $\text{NaNO}_2$  (v). (d) Normalized UV/Vis spectra of the gel (i) and sol (ii) obtained from the mixture of **1** and **2** in presence of  $\text{Fe}^{\text{II}}$  and  $\text{Fe}^{\text{III}}$  respectively. (iii)–(v) Normalized UV/Vis spectra of the gels obtained from the mixture of **1**, **2** and  $\text{Fe}^{\text{II}}$  in presence of redox reaction involving 0.067 M of  $\text{NaNO}_2$  (iii), 0.134 M of  $\text{NaNO}_2$  (iv) and 0.067 M of  $\text{H}_2\text{O}_2$  (v). (vi) Normalized UV/Vis spectra of the sol obtained from the mixture of **1**, **2** and  $\text{Fe}^{\text{II}}$  in presence of redox reaction involving 0.134 M of  $\text{H}_2\text{O}_2$ . For (a)–(d), initial concentrations of **1**, **2** and  $\text{Fe}^{\text{II}}$  are 0.134 M.

Table S1 in the Supporting Information). While a slow oxidation of  $\text{Fe}^{\text{II}}$  by  $\text{NaNO}_2$  causes approximately four times reduction in  $G'$ , fast oxidation involving  $\text{H}_2\text{O}_2$  resulted in a tenfold decrease in the stiffness of the gel compared to the pristine  $\text{Fe}^{\text{II}}$  gel. However, irrespective of rate of oxidation of  $\text{Fe}^{\text{II}}$ , the final  $\text{Fe}^{\text{III}}$  gels showed approximately four times increase in strength of the materials.

We further increased the rate of oxidation of  $\text{Fe}^{\text{II}}$  by increasing the concentration of the oxidizing agents. Similar trends in  $G'$ ,  $G''$ , and viscosity were monitored when we increased the initial concentration of  $\text{NaNO}_2$  from 0.067 M to 0.134 M (Figures 3a and S10 in the Supporting Information). The final values of  $G'$  and  $G''$  of the gels concomitantly decreases with an increase in the initial concentration of  $\text{NaNO}_2$  (Figures 3c, S11, and Table S1 in the Supporting Information). Notably, when we increased the concentration of  $\text{H}_2\text{O}_2$  (0.134 M), instead of a gel, precipitation occurred (Figure S12). These results suggest that the formation of  $\text{Fe}^{\text{III}}$  gel depends significantly on the rate of oxidation of  $\text{Fe}^{\text{II}}$ . Interestingly, direct treatment of the mixture of **1** and **2** with  $\text{Fe}^{\text{III}}$  (0.134 M) produced precipitation (Figure S12). Correlation of these results indicates a complex mechanism for the formation of  $\text{Fe}^{\text{III}}$  gels via oxidation processes, in which two phenomena are occurring simultaneously by the formation of the imine **3** and the conversion of  $\text{Fe}^{\text{II}}$  to  $\text{Fe}^{\text{III}}$ . A slow conversion to  $\text{Fe}^{\text{III}}$  allows formation of continuous net-

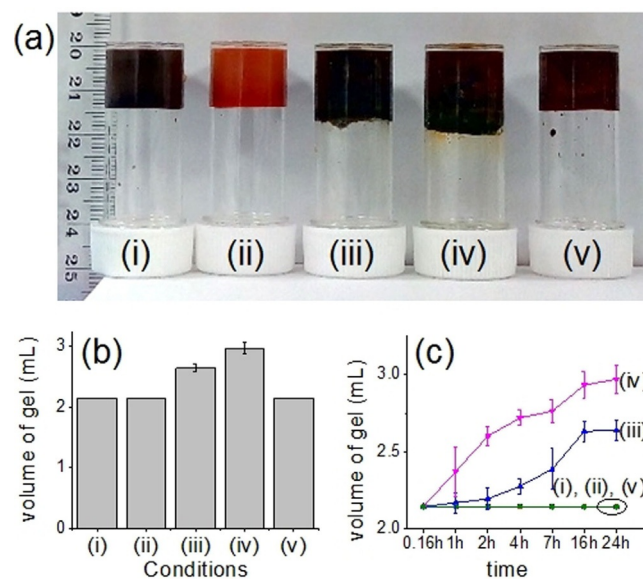
work structures involving **3**, whereas fast oxidation drives the system towards kinetically trapped states, in which the intermolecular interactions involving  $\text{Fe}^{\text{III}}$  were strong enough to produce precipitation.<sup>[12]</sup>

The visual appearance of the gels also depends on the initial reaction conditions. Oxidation of  $\text{Fe}^{\text{II}}$  either by  $\text{NaNO}_2$  or  $\text{H}_2\text{O}_2$  turned the reddish brown  $\text{Fe}^{\text{II}}$  gels into deep brown  $\text{Fe}^{\text{III}}$  gels (Figure S12 in the Supporting Information). However, these gels behaved differently by spectroscopy. In fluorescence, the emission of all  $\text{Fe}^{\text{III}}$  gels was quenched at 550 nm (Figure S13). In UV/Vis data, all the  $\text{Fe}^{\text{III}}$  gels showed absorption in the region 330–360 nm (Figure 3d). Interestingly, the absorption intensity in this region increased as the formation of  $\text{Fe}^{\text{III}}$  was faster. When the rate of oxidation of  $\text{Fe}^{\text{II}}$  was significantly high, particularly with  $\text{H}_2\text{O}_2$ , a distinct peak at 550 nm appeared. A similar spectral appearance was also observed for the sol obtained from direct treatment of  $\text{Fe}^{\text{III}}$  with **1** and **2**. These results suggest that when the rate of oxidation of  $\text{Fe}^{\text{II}}$  is extremely high, almost all the  $\text{Fe}^{\text{II}}$  is converted to  $\text{Fe}^{\text{III}}$  very rapidly and the binding interactions follow the similar pattern as in the case with  $\text{Fe}^{\text{III}}$  alone. To get more insight, FTIR studies were conducted with the metallogels prepared under different conditions (Figure S14 in the Supporting Information). FTIR studies showed that irrespective of rate of oxidation, for the  $\text{Fe}^{\text{III}}$  gels (as well as sols), C=N bond formation occurs as in all cases with appearance of a broad peak at  $1680\text{ cm}^{-1}$ . To understand the interactions with  $\text{Fe}^{\text{III}}$ , we attempted to collect  $^1\text{H}$  NMR spectra of the  $\text{Fe}^{\text{III}}$ -containing gels and sols obtained under different conditions. First, we recorded the  $^1\text{H}$  NMR spectrum of **1** and **2** in presence of  $\text{NaNO}_2$  and  $\text{H}_2\text{O}_2$  separately to investigate if oxidation leads to any chemical changes in the systems. From Figures S15–S17 in the Supporting Information, it is evident that no chemical changes occur to **1** and **2** in presence of the oxidizing agents. Similarly, the presence of  $\text{NaNO}_2$  does not alter the composition of the gel formed from **1** and **2** (Figure S18). We were unable to record the  $^1\text{H}$  NMR spectra of the gels obtained from mixture of  $\text{Fe}^{\text{II}}$  with **1** and **2** in the presence of  $\text{NaNO}_2$ . However, HRMS experiments showed formation of **3** both in absence and presence of  $\text{Fe}^{\text{II}}$  involving  $\text{NaNO}_2$  (Figure S19). For the systems formed from **1** and **2** involving  $\text{H}_2\text{O}_2$ , the aromatic protons of **2** became broad both in absence and presence of  $\text{Fe}^{\text{II}}$  (Figure S20 in the Supporting Information). A similar spectral appearance was observed in the sol obtained from direct treatment of the mixture of **1** and **2** with  $\text{Fe}^{\text{III}}$  (Figure S21). Correlation of the results from Figures S16 and S20 shows that  $\text{H}_2\text{O}_2$  readily reacts with  $\text{Fe}^{\text{II}}$  (as the peak at 10.48 ppm corresponds to  $\text{H}_2\text{O}_2$  disappeared in presence of  $\text{Fe}^{\text{II}}$ ) but causes no chemical changes to **3**. HRMS analysis also confirms the formation of compound **3** in all cases (Figures S22, S23). Furthermore, analysis of  $^1\text{H}$  NMR spectra showed the presence of chemical analytes (oxidizing agents, metal ions) have no significant effect on conversion of **3** (the percentage conversion of **3** varies between 20–26% in all cases). The slight variations in conversion of **3** is probably due to the fact that during recording of the NMR spectra, a small amount of hydrolysis may occur.<sup>[8b,9]</sup> Notably, in the mixture of **1** and **2**, unlike  $\text{Fe}^{\text{II}}$ , no shift of the imine proton  $\text{H}_a$  (Figures S20, S21)

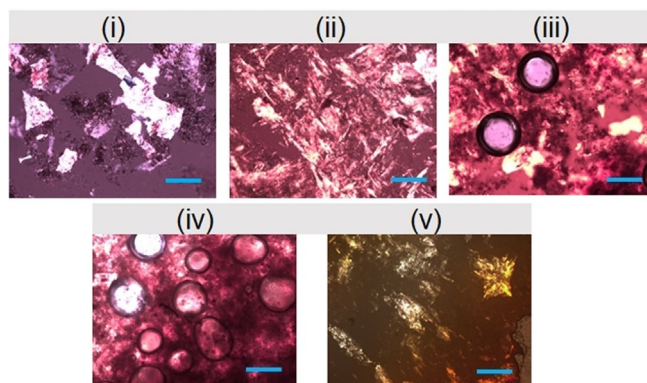
was observed whether  $\text{Fe}^{\text{III}}$  is used directly or generated in situ by oxidation of  $\text{Fe}^{\text{II}}$ . Hence, formation of **3** occurred in all cases, and the interaction of **3** with  $\text{Fe}^{\text{III}}$  is not the only the determining factor for the formation a gel or sol, but instead depends upon the assembly of the underlying structures, in which the rate of formation of  $\text{Fe}^{\text{III}}$  also determines how the Fe salt interacts with the fibres. SEM images of the gels (and sols) clearly demonstrate different aggregation depending upon the preparation pathways (Figure S24 in the Supporting Information).<sup>[12,13]</sup>

The resulting  $\text{Fe}^{\text{III}}$  gels showed unusual swelling behaviour depending on the rate of oxidation of  $\text{Fe}^{\text{II}}$  (Figure 4a). The volume of the  $\text{Fe}^{\text{II}}$  gels increases on conversion to the  $\text{Fe}^{\text{III}}$  ions by  $\text{NaNO}_2$  and the degree of swelling is proportional to the initial concentration of  $\text{NaNO}_2$  (Figure 4b and c). When 0.067 M of  $\text{NaNO}_2$  was used as oxidant, the resulting  $\text{Fe}^{\text{III}}$  gel showed approximately 23% increase in volume compared to the pristine  $\text{Fe}^{\text{II}}$  gel. An increase in  $\text{NaNO}_2$  concentration from 0.067 to 0.134 M caused a further increase in volume of the gel ( $\approx 38\%$ ). Figure 4c shows the increase in gel volume with time under different rate of oxidation of  $\text{Fe}^{\text{II}}$ . Interestingly, when  $\text{H}_2\text{O}_2$  was used as oxidant, no such swelling was noticed.

We highlight that swelling of such a supramolecular, low molecular weight gel is very unusual. Normally, such swelling is limited to cross-linked polymer gels. To determine the reason, polarising optical microscopic (POM) images of the gels were recorded, which showed the existence of spherical gas bubbles inside the gel medium obtained from  $\text{NaNO}_2$  oxidation (Figure 5). The gas bubbles are formed because of the generation of NO and  $\text{NO}_2$  due to the redox reaction,<sup>[14]</sup> which create internal stresses resulting swelling.<sup>[15]</sup> The density of the



**Figure 4.** (a) Photograph representing the pathway-driven swelling of the gels. The gels are prepared from **1** and **2** in absence (i) and presence (ii) of  $\text{Fe}^{\text{II}}$ . In situ oxidation of the  $\text{Fe}^{\text{II}}$  gels by (iii) 0.067 M of  $\text{NaNO}_2$ , (iv) 0.134 M of  $\text{NaNO}_2$ , and (v) 0.067 M of  $\text{H}_2\text{O}_2$  gives  $\text{Fe}^{\text{III}}$  gels. (b) Bar graph representing the final volume of the respective gels obtained from (a). (c) Time-variable change in the volume of the respective gels from (a). For (a)–(d), initial concentrations of **1**, **2**, and  $\text{Fe}^{\text{II}}$  are 0.134 M.



**Figure 5.** POM images of the gels (the scale bars represent 0.2 mm). The gels are prepared from **1** and **2** in absence (i) and presence (ii) of  $\text{Fe}^{\text{II}}$ . In situ oxidation of the  $\text{Fe}^{\text{II}}$  gels by (iii) 0.067 M of  $\text{NaNO}_2$ , (iv) 0.134 M of  $\text{NaNO}_2$  and (v) 0.067 M of  $\text{H}_2\text{O}_2$  gives  $\text{Fe}^{\text{III}}$  gels. For (i)–(v), initial concentrations of **1**, **2** and  $\text{Fe}^{\text{II}}$  are 0.134 M.

gas bubbles increases with as the increase in initial  $\text{NaNO}_2$  concentration, which governs the amount of volume increase in the  $\text{Fe}^{\text{III}}$  gels. In comparison, no such gas bubbles were observed in the POM images of other metallogels. The ability of the gels towards swelling before the destruction was also verified (Figure S25 in the Supporting Information). For this purpose, we increased the initial concentration of  $\text{NaNO}_2$  further. Swelling of the gel occurred up to a concentration of 0.134 M of  $\text{NaNO}_2$ . Above this concentration of  $\text{NaNO}_2$ , the volume of the gels increases, but some amount of the gel from the upper surface was destroyed and appeared as sol upon inversion of the vials. These observations suggest that the gel network is strong enough to allow swelling until a certain point, after which the internal stresses produced by the air bubbles becomes predominant and causes deformation of the network structures at the upper surface although the rest of gels remained intact.

In conclusion, we have shown that the pathway dependence is applicable to the redox-driven gels by utilizing a  $\text{Fe}^{\text{II}}/\text{Fe}^{\text{III}}$  redox-based metal–organic gel system. To establish this, we utilize dynamic imine bond formation between an aldehyde (**1**) and an amine (**2**) as the key chemical reaction and incorporated  $\text{Fe}^{\text{II}}$  ions into the gel medium during the self-assembly process. Significantly, direct preparation of the  $\text{Fe}^{\text{III}}$ -gel from the mixture of **1**, **2** and  $\text{Fe}^{\text{III}}$  ions is not feasible in our case. However, in situ oxidation of the  $\text{Fe}^{\text{II}}$  ions by various oxidising agent results in conversion to a  $\text{Fe}^{\text{III}}$ -organic gel, where the material properties like gel stiffness, gel strength, swelling etc. can be controlled just by controlling the rate of oxidation of the  $\text{Fe}^{\text{II}}$  ions. We established that the rate of formation of  $\text{Fe}^{\text{III}}$  ions actually determines the extent of intermolecular interactions whether to produce gels or precipitations. Hence, for the  $\text{Fe}^{\text{III}}$ -metallogels, which cannot be prepared directly, we can achieve those gel states in an indirect way by employing a redox reaction. We envisage that, our approach will open up opportunities to construct new functional redox gels.

## Acknowledgements

S.P. thanks the Royal Society and SERB of India for a Newton International Fellowship. D.A. thanks the EPSRC for a Fellowship (EP/L021978/1). The authors thank Dr. Bart Dietrich and Valentina Gauci for NMR and HRMS experiments.

## Conflict of interest

The authors declare no conflict of interest.

**Keywords:** kinetic control • metal–organic gels • pathway dependence • redox responsiveness • supramolecular gels • swelling

- [1] a) P. Dastidar, *Gels* **2019**, *5*, 15; b) T. Christoff-Tempesta, A. J. Lew, J. H. Ortony, *Gels* **2018**, *4*, 40; c) N. Mehwish, X. Dou, Y. Zhao, C.-L. Feng, *Mater. Horiz.* **2019**, *6*, 14–44; d) J. Li, L. Geng, G. Wang, H. Chu, H. Wei, *Chem. Mater.* **2017**, *29*, 8932–8952; e) J. Li, R. Xing, S. Bai, X. Yan, *Soft Matter* **2019**, *15*, 1704–1715; f) Y. Li, D. J. Young, X. J. Loh, *Mater. Chem. Front.* **2019**, *3*, 1489–1502.
- [2] a) P. Dastidar, S. Ganguly, K. Sarkar, *Chem. Asian J.* **2016**, *11*, 2484–2498; b) H. Wu, J. Zheng, A.-L. Kjøniksen, W. Wang, Y. Zhang, J. Ma, *Adv. Mater.* **2019**, *31*, 1806204; c) S. Xiao, P. J. Paukstelis, R. D. Ash, P. Y. Zavalij, J. T. Davis, *Angew. Chem. Int. Ed.* **2019**, *58*, 18434–18437; *Angew. Chem.* **2019**, *131*, 18605–18608.
- [3] a) A. Panja, K. Ghosh, *New J. Chem.* **2019**, *43*, 934–945; b) A. J. McConnell, C. S. Wood, P. P. Neelakandan, J. R. Nitschke, *Chem. Rev.* **2015**, *115*, 7729–7793.
- [4] a) C. Echeverria, S. N. Fernandes, M. H. Godinho, J. P. Borges, P. I. P. Soares, *Gels* **2018**, *4*, 54; b) M. D. Segarra-Maset, V. J. Nebot, J. F. Miravet, B. Escuder, *Chem. Soc. Rev.* **2013**, *42*, 7086–7098; c) G. R. Deen, X. J. Loh, *Gels* **2018**, *4*, 13; d) P. K. Bolla, V. A. Rodriguez, R. S. Kalhapure, C. S. Kolli, S. Andrews, J. Renukuntla, *J. Drug Delivery Sci. Technol.* **2018**, *46*, 416–435; e) C. D. Jones, J. W. Steed, *Chem. Soc. Rev.* **2016**, *45*, 6546–6596; f) W. Cheng, Y. Liu in *Biopolymer-Based Composites* (Eds.: S. Jana, S. Maiti, S. Jana), Woodhead, Duxford, **2017**, pp. 31–60.
- [5] a) S. Ahmed, A. Chatterjee, K. Das, D. Das, *Chem. Sci.* **2019**, *10*, 7574–7578; b) S. Dhiman, K. Jalani, S. J. George, *ACS Appl. Mater. Interface* **2019**, *12*, 5259–5264; c) Y. Che, S. Zschoche, F. Obst, D. Appelhans, B. Voit, *J. Polym. Sci. Part A* **2019**, *57*, 2590–2601; d) W. A. Ogden, Z. Guan, *ChemSystemsChem* **2019**, *1*, e1900030; e) X. Sui, X. Feng, M. A. Hempenius, G. J. Vancso, *J. Mater. Chem. B* **2013**, *1*, 1658–1672; f) F. Peng, G. Li, X. Liu, S. Wu, Z. Tong, *J. Am. Chem. Soc.* **2008**, *130*, 16166–16167; g) J. P. Wojciechowski, A. D. Martin, P. Thordarson, *J. Am. Chem. Soc.* **2018**, *140*, 2869–2874.
- [6] R. Eelkema, A. Pich, *Adv. Mater.* **2020**, 1906012.
- [7] a) J. Raeburn, A. Zamith Cardoso, D. J. Adams, *Chem. Soc. Rev.* **2013**, *42*, 5143–5156; b) S. Panja, B. Dietrich, D. J. Adams, *ChemSystemsChem* **2020**, *2*, e1900038; c) E. R. Draper, D. J. Adams, *Langmuir* **2019**, *35*, 6506–6521; d) S. Panettieri, R. V. Ulijn, *Curr. Opin. Struct. Biol.* **2018**, *51*, 9–18.
- [8] a) C. B. Minkenberg, W. E. Hendriksen, F. Li, E. Mendes, R. Eelkema, J. H. van Esch, *Chem. Commun.* **2012**, *48*, 9837–9839; b) T. Jiao, G. Wu, Y. Zhang, L. Shen, Y. Lei, C.-Y. Wang, A. C. Fahrenbach, H. Li, *Angew. Chem. Int. Ed.* **2020**, <https://doi.org/10.1002/anie.201910739>; *Angew. Chem. Int. Ed.* **2020**, <https://doi.org/10.1002/ange.201910739>.
- [9] V. Saggiomo, U. Lüning, *Tetrahedron Lett.* **2009**, *50*, 4663–4665.
- [10] J. T. Auletta, G. J. LeDonne, K. C. Gronborg, C. D. Ladd, H. Liu, W. W. Clark, T. Y. Meyer, *Macromolecules* **2015**, *48*, 1736–1747.
- [11] A. Panja, K. Ghosh, *Mater. Chem. Front.* **2018**, *2*, 1866–1875.
- [12] R. K. Grötsch, C. Wanzke, M. Speckbacher, A. Angli, B. Rieger, J. Boekhoven, *J. Am. Chem. Soc.* **2019**, *141*, 9872–9878.
- [13] S. Panja, D. J. Adams, *Chem. Commun.* **2019**, *55*, 10154–10157.

[14] a) D.-J. Kim, D. Pradhan, K.-H. Park, J.-G. Ahn, S.-W. Lee, *Mater. Trans.* **2008**, *49*, 2389–2393; b) M. S. Ibrahim, A. H. Gemeay, S. E.-d. H. Etaiv, *Transition Met. Chem.* **2001**, *26*, 44–49.

[15] a) K. Krishnamurti, *Nature* **1929**, *123*, 242–243; b) J. Singh, M. E. Weber, *Chem. Eng. Sci.* **1996**, *51*, 4499–4508.

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Manuscript received: February 27, 2020

Accepted manuscript online: March 3, 2020

Version of record online: April 30, 2020

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