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Nano-sized mesoporous phosphated tin oxide as an efficient solid acid catalyst†

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Herein, we prepared a mesoporous tin oxide catalyst (mSnO₂) activated with phosphate species by the adsorption of phosphate ions from a phosphoric acid solution onto tin oxyhydroxide (Sn(OH)₄) surface. The phosphate content ranged from 3 to 45 wt%. The nonaqueous titration of *n*-butylamine in acetonitrile was used to determine the total surface acidity level. FTIR of chemically adsorbed pyridine was used to differentiate between the Lewis and Brønsted acid sites. Thermal and X-ray diffraction analysis indicated that the addition of phosphate groups stabilized the mesostructure of mSnO₂ and enabled it to keep its crystalline size at the nanoscale. FTIR analysis indicated the polymerization of the HPO₄²⁻ groups into P₂O₇⁴⁻, which in turn reacts with SnO₂ to form a SnP₂O₇ layer, which stabilizes the mesoporous structure of SnO₂. The acidity measurements showed that the phosphate species are distributed homogeneously over the mSnO₂ surface until surface saturation coverage at 25 wt% PO₄³⁻, at which point the acid strength and surface acidity level are maximized. The catalytic activity was tested for the synthesis of hydroquinone diacetate, where it was found that the % yield of hydroquinone diacetate compound increased gradually with the increase in PO₄³⁻ loading on mSnO₂ until it reached a maximum value of 93.2% for the 25% PO₄³⁻/mSnO₂ catalyst with 100% selectivity and excellent reusability for three consecutive runs with no loss in activity.

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Introduction

The replacement of liquid acid or acid halide catalysts, such as H₂SO₄, H₃PO₄, HF and AlCl₃, is a critical factor for the design of cleaner processes for green catalytic action.¹ The problems of homogeneous liquid acid catalysis include toxicity, corrosion, catalyst waste and difficulty in separation and recovery.² The economics of the separation cost majorly determine the viability of such a process.³ Efforts have been made to design environmentally safe processes and technologies to achieve these goals. The best catalytic technology for solving the problems of the liquid acid catalyst is to replace it by a heterogeneous solid acid catalyst.⁴ The reuse and recycling of the catalyst in heterogeneous processes are the main advantages of the increased catalyst efficiency over homogeneous ones.⁵ The presence of an interconnected network of a large number of pores, acid sites of different strength and the hydrophobic properties of solid catalyst surfaces support inhibiting the processes of corrosion and separation and for limiting pollution

problems.² Oil refining and pharmaceutical industries utilize solid catalysts with acidic properties.^{6–8} SO₄²⁻, PO₄³⁻, WO₄²⁻, Cl⁻ and F⁻ supported on metal oxides have been used to increase their thermal stability, mesoporosity, and more critically, the acidity level of such catalysts.^{9–12} Most of the published works to date have been focused on SO₄²⁻- and WO₄²⁻-supported metal oxide systems,^{13–18} which have been found to be suitable for specific acid-catalyzed reactions. The acidity of such systems strongly depend on the preparation method.^{11,14,16}

A template pathway for preparing mesoporous materials is a novel process to produce materials with high activity for the synthesis of industrially relevant chemicals.^{19–22} Mesoporous materials are characterized by a high surface area, ordered pore channels, and large pore volume value. These advantages can facilitate active components achieving a highly dispersed state on the support, which raises their activity for reacting molecules.^{23,24} For better performance, an efficient heterogeneous catalyst needs a higher acid strength level and density as well as thermal stability to minimize catalyst poisoning and leaching. Therefore, mesoporous materials, primarily metal oxides, have attracted growing interest as a support in the field of catalysis.²⁵ Tin oxide is one of the most attractive functional materials because of its potential applications, mainly as a catalyst and as a carrier in supported catalysts.²⁶ Besides its intrinsic non-stoichiometry and crystal defects, tin oxide nanoparticles can be used more efficiently as a catalyst due to the faster migration of the oxide ions within the SnO₂ nanoparticles to increase the

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surface to volume ratio. Many authors have used high valence cations of metal oxide to promote the thermal stability and catalytic property of pure tin oxide.^{27,28} Therefore, it is of interest to study the novelties of phosphate-supported SnO₂ as a solid acid catalyst. Various chemical routes have been used to synthesize nanoparticles of mesoporous metal oxides, *e.g.*, precipitation, sol-gel, hydrothermal, microwave-assisted syntheses and ultrasonic spray pyrolysis.^{29,30} The addition of anion groups to a mesostructure metal oxide during the preparation methods prevent the sintering process and thus lead to improvements in their thermal stability and textural properties.^{31,32}

Further improvement is necessary, however, for the preparation of thermally stable mesoporous tin oxide. This process is still a challenge to researchers.^{9,33} Unlike sulfated metal oxides, phosphated metal oxides have received inadequate attention. Albeit, phosphate-based solid acids have proven quite efficacious in several industrially important acid-demanding reactions.

The present study aimed to adopt a phosphate-modified mesoporous tin oxide (mSnO₂) system as a catalyst. The modified catalytic systems (H₃PO₄/m-SnO₂) were then applied to the preparation of hydroquinone diacetate, which is widely used as an ingredient for skin bleaching cream. DTA was used to systematically elucidate the physico-chemical properties of the H₃PO₄/mSnO₂ catalysts, while TGA, XRD, TEM and FTIR techniques were applied for characterization, and the texture properties were determined from N₂ adsorption at -196 °C. The surface acidity was also examined.

Results and discussion

The thermograms of DTA and the TGA of dried PO₄³⁻/Sn(OH)₄ samples are shown in Fig. 1. The sample showed endothermic peaks at low temperature, which are assigned to the evolution of physically adsorbed water. The sample also showed two exothermic peaks. The first exothermic effect with a weight loss of 8.9 wt% at around 312 °C was attributed to the decomposition of the surfactant template and partial removal of the OH groups (dehydroxylation),^{34,35} while the second exothermic peak

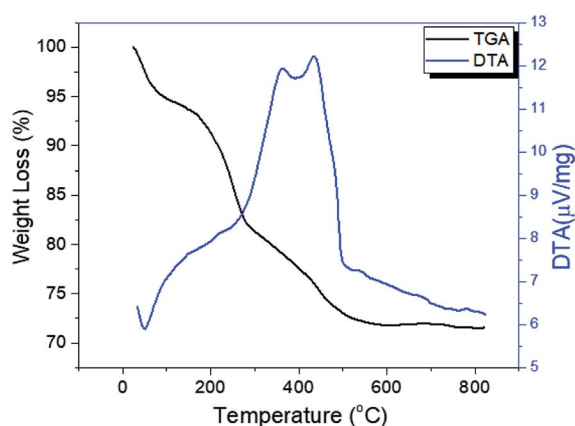


Fig. 1 DTA and TGA curves of dried Sn(OH)₄.

at 423 °C corresponded to a weight loss of 3.1 wt%, which may be due to the crystallization of tin oxyhydroxide into tetragonal cassiterite SnO₂.³⁶ In the case of PO₄³⁻/Sn(OH)₄ dried samples, the DTA in Fig. 2 shows exothermic peaks at 321 °C, 334 °C and 342 °C as the phosphate content increases for the 8%, 25% and 45% PO₄³⁻/Sn(OH)₄ samples, respectively. Also, the exothermic peaks of the crystallization of tin oxyhydroxide into mSnO₂ were also shifted to higher temperatures as the phosphate content increased, reaching 427 °C, 484 °C and 504 °C for the 8%, 25% and 45% PO₄³⁻/mSnO₂ samples, respectively, indicating that the addition of phosphate to mSnO₂ may enhance the surfactant-tin oxyhydroxide surface interaction,³⁷ thus increasing the thermal stability of the tin oxyhydroxide phase and shifting its crystallization to higher temperature. These results show no effect that could be assigned to the decomposition of PO₄³⁻ species, which was attributed to the thermal stability when treated with phosphoric acid.³⁸

Fig. 2S(A)† shows the low-angle XRD patterns of 3%, 25%, and 35% PO₄³⁻/mSnO₂ samples heated at 400 °C. The XRD curves exhibit one diffraction peak at 2θ = 2.4°, corresponding to the (100) reflection, which was assigned to the presence of hexagonal mesostructured SnO₂.^{39,40} The same peak was observed for the calcined samples at 400 °C, 550 °C and 650 °C (Fig. 2S†), which indicates that the presence of phosphate groups maintain the mesoporous structure of mSnO₂.

Fig. 3A illustrates the wide-angle XRD patterns of mSnO₂, and for the 3%, 25% and 35% PO₄³⁻/mSnO₂ calcined at 400 °C. The samples were found to display four well-distinguished peaks at 2θ = 26.64°, 33.91°, 38.10° and 51.74°, corresponding to the (110), (100), (200) and (211) reflection planes of a tetragonal cassiterite phase of mSnO₂, respectively. The addition of PO₄³⁻ to mSnO₂ until 25 wt% PO₄³⁻ shows an XRD pattern similar to that of the pure mSnO₂, indicating that the PO₄³⁻ species exists as amorphous phosphate groups on the mSnO₂ surface.⁴¹ When the loading of PO₄³⁻ increased to 35 wt%, a new reflection peak appears at 2θ = 22.61°, which can be assigned to a layer-structured SnP₂O₇.⁴² On the other hand, it is clear from Fig. 3B that the intensity of the diffraction peaks of the 3% PO₄³⁻/mSnO₂ sample increases with the increase in the calcination temperature of the sample from 400 °C to 650 °C,

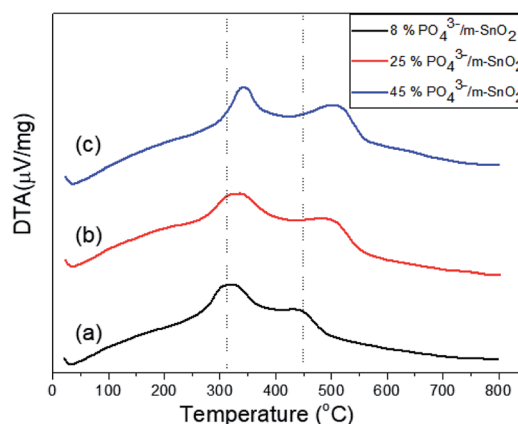


Fig. 2 DTA curves of (a) 8%, (b) 25%, (c) 45% PO₄³⁻/m-SnO₂ samples.

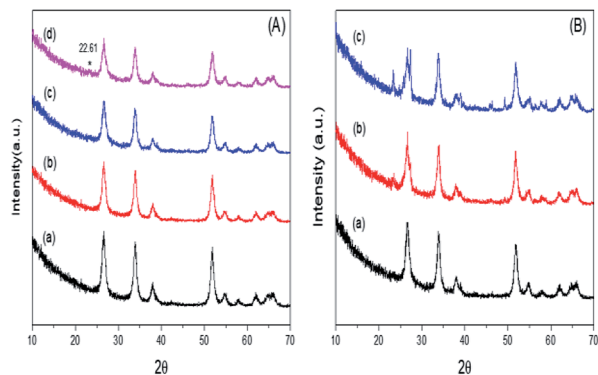


Fig. 3 (A) Wide-angle XRD pattern of (a) m-SnO₂ (400 °C) and the samples of PO₄³⁻/m-SnO₂ (400 °C) at (b) 3%, (c) 25%, (d) 35%, (B) wide-angle XRD patterns of the sample 3% PO₄³⁻/m-SnO₂ at (a) 400 °C, (b) 550 °C, (c) 650 °C.

which may be due to the increase in the dispersion of PO₄³⁻ on the mSnO₂ surface, which in turn increase the thermal stability of mSnO₂, arising from the interaction of PO₄³⁻ and mSnO₂. The degree of crystallization of cassiterite mSnO₂ gradually decreases with the increase in the phosphate content, while the width of the reflections is considerably broadened, indicating a small crystalline domain size (Table 1). The addition of phosphate anions to the template synthesis reaction medium of mSn(OH)₄, substantially slows down the crystallization process *via* retarding the crystal size growth by occupying the defects on the surface of mSn(OH)₄,⁴³ thus increasing its thermal stability and thus protecting the mesoporous structure from collapsing.⁴¹ The crystallite size data were determined from X-ray diffraction and are listed in Table 1, showing that the crystallite size of PO₄³⁻/mSnO₂ is smaller in comparison to the pure mSnO₂ nanocrystalline (13.70–11.30 nm). Furthermore, the crystallite size of 3% PO₄³⁻/mSnO₂ increases from 12.17 to 16.66 nm with the increase in calcination temperature (400–650 °C), which indicates that pure mSnO₂ and PO₄³⁻/mSnO₂ are present in the nanoscale. In the phosphate sample annealed between 400 °C and 650 °C, the crystallite size slightly changes from 13.70 to 16.66 nm, suggesting that the PO₄ species remain bonded to the mSnO₂ surface, stabilizing it against sintering. It was reported that the order of the mesostructured materials strongly depends on the texture properties of the pores.⁴⁴ An analogous phenomenon was observed with TiO₂, ZrO₂ and Fe₂O₃.^{44,45}

TEM images of mSnO₂ show an ordered structure with a uniform mesoporous arrangement into hexagons (Fig. 4a). It

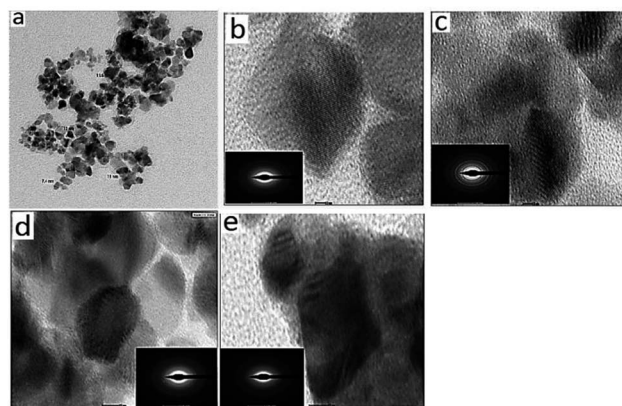


Fig. 4 TEM images of (a) mSnO₂ calcined at 550 °C, (b) 3% PO₄³⁻/m-SnO₂ calcined at 400 °C, (c) 3% PO₄³⁻/m-SnO₂ calcined at 550 °C, (d) 8% PO₄³⁻/m-SnO₂ calcined at 400 °C and (e) 25% PO₄³⁻/m-SnO₂ calcined at 400 °C.

is apparent that mSnO₂ has a significant void fraction, which may be due to the mesoporous structure of SnO₂, and concomitantly a rather low density. Fig. 4c–e show the TEM images of the 3%, 8% and 25% PO₄³⁻/mSnO₂ catalysts, which indicate that when PO₄³⁻ species are loaded on mSnO₂ calcined at 400 °C, the lattice retains the ordered mesoporous channels and the PO₄³⁻ groups are embedded uniformly into the mSnO₂ framework without disordering the mesostructured framework. The selected area electron diffraction (SAED) pattern images (insets) of the samples exhibit diffraction rings attributed to the polycrystalline behaviour of the mSnO₂, while the crystallite degrees of the 3%, 8% and 25% PO₄³⁻/mSnO₂ catalysts decrease with the increase in PO₄³⁻ content.^{46,47} The TEM images of 3% PO₄³⁻/mSnO₂ calcined at 550 °C (Fig. 4b) illustrate that the mesoporous structure remains even after calcination at 550 °C, which reflects the role of the phosphate groups in improving the thermal stability of the tin oxide.⁴⁷ It is clear from the TEM images that the dark colour and the area occupied by the phosphate group both increase with both the phosphate content and the calcination temperature, indicating the great dispersion of phosphate groups on the mesoporous SnO₂ surfaces. Also, the side images (inset) show an increase in the crystalline size with increasing calcination temperature, which is in a good agreement with the results from the XRD analysis.

The FTIR spectra of the mSnO₂ 550 °C, xPO₄³⁻/mSnO₂ and 3PO₄³⁻/mSnO₂ samples calcined at different temperatures are shown in Fig. 5A and B. The band at 1625 cm⁻¹ appearing in the

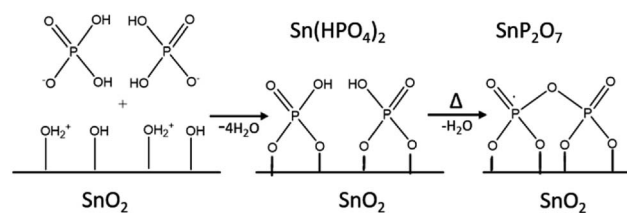
Table 1 Comparison of the structural parameters of mSnO₂ and various PO₄³⁻/mSnO₂ samples

Samples	<i>D</i> (nm)	<i>d</i> ₁₀₀ (nm)	<i>d</i> _{pore,BJH} (nm)	<i>a</i> ⁰ (nm)	<i>t</i> _{wall} (nm)
mSnO ₂ -400	13.70	3.62	3.22	4.18	0.96
3%PO ₄ ³⁻ /mSnO ₂ -400	12.17	3.63	3.15	4.19	1.04
25%PO ₄ ³⁻ /mSnO ₂ -400	11.87	3.66	2.66	4.23	1.57
35%PO ₄ ³⁻ /mSnO ₂ -400	11.30	3.69	2.52	4.26	1.74
3%PO ₄ ³⁻ /mSnO ₂ -550	14.50	3.60	3.41	4.16	0.75
3%PO ₄ ³⁻ /mSnO ₂ -650	16.66	3.57	3.82	4.13	0.31

spectra of the samples even after calcination at different temperatures corresponds to the OH bending vibrations. The FTIR spectra of phosphate mSnO₂ generally show absorption bands at 1035 and 1150, 1360, 1410 and 1520 cm⁻¹. The absorption bands at 1360, 1410 and 1520 cm⁻¹ could be assigned to P–O stretching frequency bonds in which their order is close to two P=O, phosphoryl groups.⁴⁸ By increasing the temperature to 650 °C, the broadband at 1550–1300 cm⁻¹ is reduced to one sharp band and one shoulder, due to the formation of bidentate-bound phosphate ions (C₂V point group). The band at 1035 and the shoulder at 1150 cm⁻¹ are frequencies for PO₄³⁻ and the position of such a fundamental band was observed in phosphated zirconia.⁴⁹ The fading of the shoulder and the presence of a broad and sharp band located at 1055 cm⁻¹ when the PO₄³⁻ content was increased to 45 wt% may be due to the dropping of the symmetry in the free PO₄³⁻ (Td point group) to either C₃V or C₂V.⁴⁸ This may be due to the PO₄³⁻ groups being bound to the tin oxide surface and being distributed homogeneously.^{50,51} Here, the 1055 cm⁻¹ band is divided into two bands at 1035 and 1150 cm⁻¹ at 650 °C, which could be assigned to the bidentate bound phosphate ions (C₂V point group), which motivate the polymerization of the HPO₄²⁻ groups into P₂O₇⁴⁻, which is in a good agreement with the appearance of a new peak due to SnP₂O₇ in the X-ray pattern (Fig. 3A).

These results indicate that phosphate groups are chemically adsorbed on the surface of mSnO₂ and still exist on the surface even after calcination of the samples at high temperatures. The mechanism, as explained in Scheme 1, follows the dehydroxylation reaction of P/mSnO₂ samples, with amorphous Sn(HPO₄)₂ formed on the surface, which is further heated to form tin pyrophosphate, SnP₂O₇, which is in a good agreement with the XRD pattern at 650 °C.

Fig. 6 shows the nitrogen adsorption–desorption isotherms of some representative samples of PO₄³⁻/mSnO₂, in which the isotherms exhibit a type IV feature, which is a characteristic feature of mesoporous materials.⁵² Barrett–Joyner–Halenda (BJH) analysis demonstrated that the average pore diameter (*d*_{pore, BJH}) of mSnO₂ calcined at 550 °C was 13.21 nm (Table 1). The *d*_{pore, BJH} of PO₄³⁻/mSnO₂ samples was found to decrease



Scheme 1 Formation of PO₄³⁻/mSnO₂.

with the increase in phosphate content from 3.15 to 2.52 nm (Table 1). The approximate thickness of the porous wall (*t*_{wall}) calculated from the difference between the hexagonal unit cell parameter (*a*⁰ = 2*d*₁₀₀/√3) and *d*_{pore, BJH} (Table 1), after calcination at 400 °C *t*_{wall} increased with increasing the phosphate content from 0.96 nm for mSnO₂ to 1.74 nm for the 35% PO₄³⁻/mSnO₂ samples. The *d*_{pore, BJH} and *t*_{wall} values decreased with the calcination temperature. On the other hand, the values of *d*₁₁₀-spacing and *a*⁰ were found to decrease with the increase in calcination temperature, which may indicate that some changes occur on the mesostructure as the calcination temperature of PO₄³⁻/mSnO₂ increases. The values of *S*_{BET}, as shown in Table 2, increase gradually with the rise of PO₄³⁻ content to reach a maximum for 25% PO₄³⁻/mSnO₂. This amount was optimal to form a monolayer, and the values of *S*_{BET} decreased with further increases of the PO₄³⁻ content.

Moreover, the *S*_{BET} values were found to decrease when the calcination temperature increased up to 650 °C.^{50,51} The increase in the surface area values of PO₄³⁻/mSnO₂ over that of mSnO₂ may be due to the formation of two-dimensional surface layers from Sn-PO₄³⁻, which stabilize the defects and inhibit the collapsing of the mSnO₂ surface,⁵³ which also explains the role of PO₄³⁻ in enhancing the porous structure of PO₄³⁻/mSnO₂. The increase in PO₄³⁻ content beyond 25 wt% led to the production of more phosphate layers and/or the phosphate groups attacking the pore mouths blocking them, and consequently the surface area values decreased.^{54,55} The above results suggest that the interaction between PO₄³⁻ and mSnO₂ crystallites not only stabilizes the crystallite structure in lower dimensions, but also it keeps phosphate groups at the SnO₂ surfaces, inhibiting the crystallite aggregation and acting as a structure progeny director mediating nanoparticle growth and assembly and stabilizing the mesostructure walls and

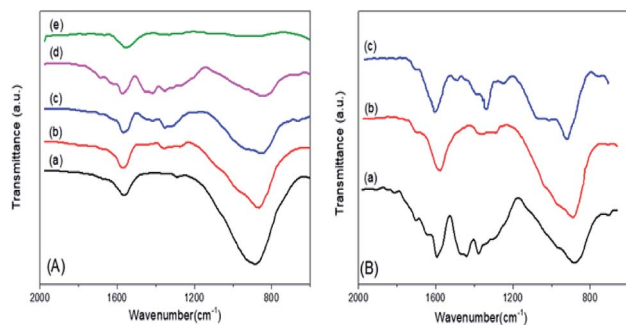


Fig. 5 (A) FTIR spectra of (a) m-SnO₂ (550 °C) and (b) 3%, (c) 25%, (d) 35%, (e) 45% PO₄³⁻/m-SnO₂ catalysts calcined at 550 °C. (B) FTIR spectra of 3% PO₄³⁻/m-SnO₂ calcined at: (a) 450 °C, (b) 550 °C, (c) 650 °C.

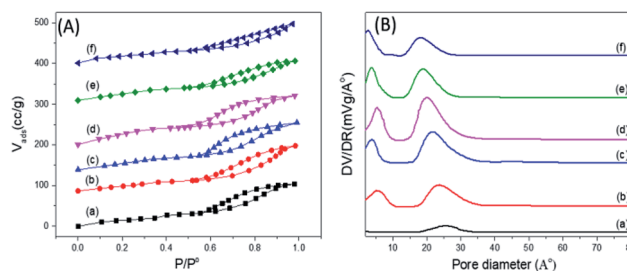


Fig. 6 (A) Adsorption–desorption isotherms of nitrogen at –196 °C (B) pore size distributions for PO₄³⁻/m-SnO₂ at (a) 3%, (b) 8%, (c) 18%, (d) 25%, (e) 35%, (f) 45% calcined at 400 °C.

Table 2 Surface characteristics of PO₄³⁻/m-SnO₂ catalysts

Catalyst	<i>T</i> (°C)	<i>S</i> _{BET} m ² g ⁻¹	<i>d</i> _{pore,BJH} (nm)	VT ml g ⁻¹
m-SnO ₂	550	29	3.22	0.13
3% PO ₄ ³⁻ /m-SnO ₂	400	65	3.15	0.16
8% PO ₄ ³⁻ /m-SnO ₂	400	95	2.96	0.17
18% PO ₄ ³⁻ /m-SnO ₂	400	130	2.81	0.18
25% PO ₄ ³⁻ /m-SnO ₂	400	155	2.66	0.18
35% PO ₄ ³⁻ /m-SnO ₂	400	113	2.52	0.14
45% PO ₄ ³⁻ /m-SnO ₂	400	75	2.46	0.14
3% PO ₄ ³⁻ /m-SnO ₂	450	63	3.35	0.15
3% PO ₄ ³⁻ /m-SnO ₂	550	56	3.41	0.15
3% PO ₄ ³⁻ /m-SnO ₂	650	41	3.82	0.14

increasing the thermal stability.⁵⁶ Again, the increase in surface area and pore volume and the decrease in the average pore diameter of the 3% to 25% PO₄³⁻/mSnO₂ samples over pure mSnO₂ with increasing the phosphate content indicate a relatively high thermal stability of the prepared samples. The decrease in the pore volume and *d*_{pore,BJH} upon increasing the phosphate content over 25 wt% may be due to the formation of multi-phosphate layers, as proposed in Scheme 1, whereby the surface hydroxyl groups of Sn(OH)₄ are replaced by phosphate groups and therefore there is an increased energy barrier and less crystal growth sites, consequently inhibiting the crystal growth.

The samples showed a bimodal pore size distribution consisting of smaller and larger interparticle pores, as illustrated in Fig. 6B. The addition of PO₄³⁻ to mSnO₂ was associated with an increase in the number of mesopores and micropores, as evident from the increase in the height of the distribution maxima and slight shift to lower rH values until a maximum at 25 wt% phosphate was reached, where it begins to decrease. Upon increasing the calcination temperature, (Fig. 3S†), both the smaller and larger interparticle pores decreased, which is in line with the decrease in the surface (Table 2). The width of the mesopore type increased with phosphate content, which confirmed the role of PO₄³⁻ in increasing the porosity of mSnO₂.⁵⁶

The nonaqueous potentiometric titration curves of *n*-butylamine of Fig. 4S† were used to calculate the total number of acid sites and the acid strength of the PO₄³⁻/mSnO₂ catalysts.^{57,58} As evident from Fig. 4S† and the data in Table 3, the acidity and the acid strength values of PO₄³⁻/mSnO₂ changed with the phosphate loading and reached their maximum value for 25% PO₄³⁻/mSnO₂ (Ei = +480.6 mV) calcined at 400 °C, which may be due to the completion of a bidentate phosphate monolayer on the tin oxide surface, as evident from the FTIR data. The decrease in the total acidity and the acid strength upon increasing the PO₄³⁻ loading may be due to the formation of more phosphate layers on the mSnO₂ surface. The calcination temperature 400 °C was the most suitable calcination temperature for the generation of strong acidic sites.

The loss of surface acidity at higher calcination temperatures up to 650 °C may be due to the loss of H⁺ by dehydration with the surface OH or OH of the next phosphate groups, forming

a bridging phosphate group on the surface, as shown in Scheme 1, which causes a reduction in acidic hydrogen and a decrease in the total acidity.

Pyridine adsorption on the PO₄³⁻/mSnO₂ samples was used to specify the Brönsted (B) and Lewis (L) acid sites on the sample surfaces. The FTIR spectra of pyridine adsorbed over the Brönsted (B) and Lewis (L) acid sites of the samples are shown in Fig. 7A and B. Bands at around 1444 and 1452 cm⁻¹ were characteristic of pyridine adsorbed on Lewis acid sites.⁵⁹ The bands at 1536 and 1553 cm⁻¹ were due to the pyridinium ion adsorbed on Brönsted acid sites.⁶⁰ Moreover, other FTIR bands appear at 1484 and 1503 cm⁻¹, attributed to the overlap of Lewis and Brönsted acid sites.⁶¹ The integrated areas of the bands at 1536 and 1452 cm⁻¹ were used to calculate the number of Brönsted and Lewis acid sites. It was evident from Fig. 7C that the samples have Brönsted and Lewis acid sites and these increase with the increase in phosphate loading up to 25% PO₄³⁻/mSnO₂, and then decrease with further increases in PO₄³⁻ loading above 25 wt%. Moreover, the number of Brönsted and Lewis acid sites decreases gradually with the increase in calcination temperature over 400 °C (Table 3). The surface distribution and interaction of phosphate species with the SnO₂ surface increase the number of Brönsted and Lewis acid sites. In analogy with sulfated tin oxide,⁶² the increase in Lewis acidity by phosphate loading may be due to the formation of a tin oxide surface phosphorus complex containing P=O covalent bonds, and so the inductive effect of these covalent bonds of P=O strengthen the Lewis acid character of the sample, as suggested in Scheme 2.^{63,64} Fig. 5S† shows the NH₃-TPD results for the pure m-SnO₂ calcined at 400 °C, which shows a broad peak of NH₃ desorption at 240 °C, which indicates the presence of weak acid sites. After modifying the mSnO₂ with PO₄³⁻, this peak starts to disappear and two peaks start to appear at 334 °C and 440 °C, corresponding to moderate and strong acid sites, respectively. For 45% PO₄³⁻, only one peak at 440 °C is seen, corresponding to the strong acid sites due to the saturation of the surface with the PO₄³⁻ groups.²⁶

Fig. 8A and B show the effect of the phosphate content and consequently, the number of Brönsted and Lewis acid sites on the % formation of hydroquinone diacetate compound. As can be seen, the % yield of hydroquinone diacetate compound increases gradually with the increase in PO₄³⁻ loading on mSnO₂ until it reaches a maximum value of 93.2% for the 25% PO₄³⁻/mSnO₂ catalyst with a selectivity of 100%. The surface acidity and acid strength of PO₄³⁻/mSnO₂ are enhanced and attain a maximum at 25 wt% PO₄³⁻. Moreover, the Brönsted and Lewis acid sites change with the same trend as the formation of hydroquinone diacetate compound, which may indicate that Lewis and/or Brönsted acid sites can catalyse the formation of hydroquinone diacetate compound. No activity was observed for pure mSnO₂, which may be due to its low acidity and site strength (+74 mV). As evident from Table 3 and Scheme 2, increasing the calcination temperature from 400 °C to 650 °C lead to a decrease in the catalytic activity, which is in line with the decrease in surface area and the decrease of the number of Brönsted and Lewis acid sites. The calcination temperature is an essential aspect in the synthesis of phosphated metal oxide

Table 3 Acidity characteristics of $\text{PO}_4^{3-}/\text{m-SnO}_2$ catalysts

Catalyst	T ($^\circ\text{C}$)	Ei (mV)	No. of acid sites/ $\text{g} \times 10^{-19}$	B/L	Conversion%	TOF h^{-1}
m-SnO ₂	400	74	0.84	0.54	0	0
3% $\text{PO}_4^{3-}/\text{m-SnO}_2$	400	320.4	0.96	1.04	8.6	222.88
8% $\text{PO}_4^{3-}/\text{m-SnO}_2$	400	372.6	1.08	0.82	50.3	1158.76
18% $\text{PO}_4^{3-}/\text{m-SnO}_2$	400	450.7	1.20	1.12	87.7	1818.31
25% $\text{PO}_4^{3-}/\text{m-SnO}_2$	400	480.6	1.33	1.68	93.2	1743.47
35% $\text{PO}_4^{3-}/\text{m-SnO}_2$	400	421.8	0.96	1.14	80.4	2083.70
45% $\text{PO}_4^{3-}/\text{m-SnO}_2$	400	382.3	0.72	1.48	63.1	2180.45
25% $\text{PO}_4^{3-}/\text{m-SnO}_2$	450	295.6	0.90	1.07	85.4	2360.83
25% $\text{PO}_4^{3-}/\text{m-SnO}_2$	550	274.6	0.84	1.52	61.6	1824.53
25% $\text{PO}_4^{3-}/\text{m-SnO}_2$	650	203.5	0.78	1.36	30.5	

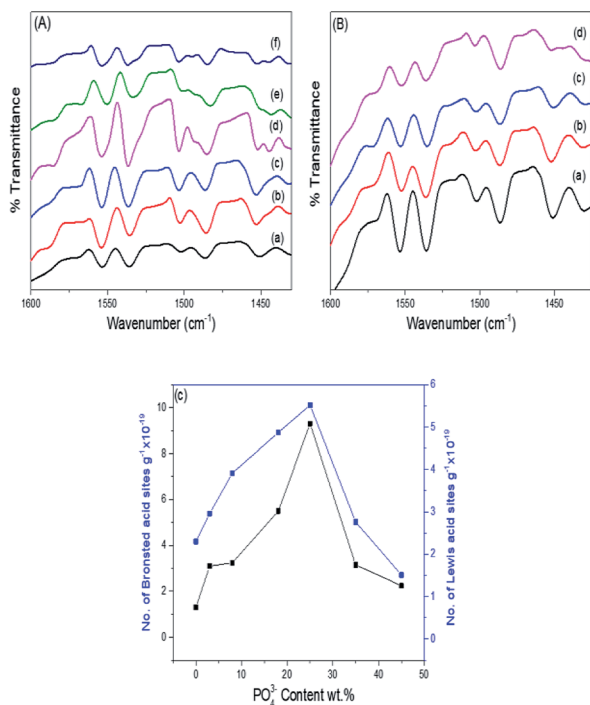
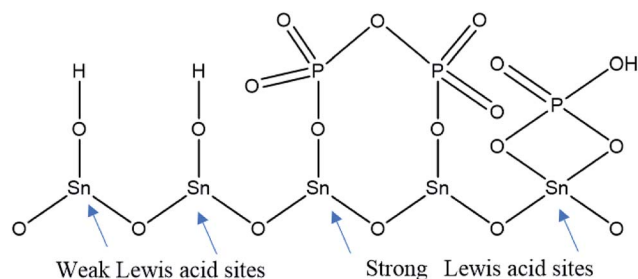


Fig. 7 (A) FTIR spectra of pyridine adsorbed on (a) 3%, (b) 8%, (c) 18%, (d) 25%, (e) 35%, (f) 45% $\text{PO}_4^{3-}/\text{m-SnO}_2$ calcined at 400°C . (B) FTIR spectra of pyridine adsorbed on the 25% $\text{PO}_4^{3-}/\text{m-SnO}_2$ sample calcined at (a) 400°C , (b) 450°C , (c) 550°C and (d) 650°C . (C) Effect of PO_4^{3-} content on the number of Brønsted and Lewis acid sites.



Scheme 2 Lewis acid on $\text{PO}_4^{3-}/\text{m-SnO}_2$.

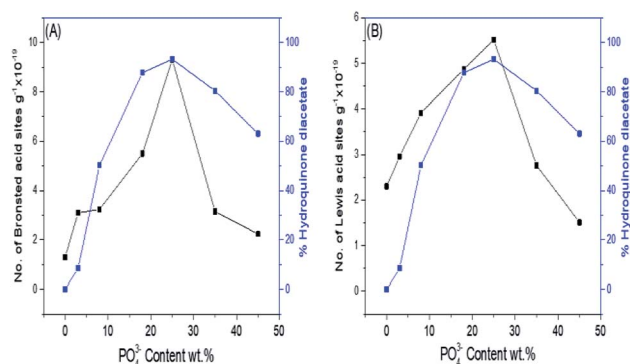


Fig. 8 (A) Effect of PO_4^{3-} content wt%/mSnO₂ calcined at 400°C on the number of Brønsted acid sites and % hydroquinone diacetate. (B) Effect of PO_4^{3-} content wt%/mSnO₂ calcined at 400°C on the number of Lewis acid sites and % hydroquinone diacetate.

catalyst due to its role in the interaction between the metal oxide surface and phosphate groups for the generation of active acid sites. The optimum operating conditions for an anion-metal oxide catalyst was found to depend on the type of metal oxides, supported anions and the calcination temperature.⁵²

It was observed that with increasing the molar ratio from 1 : 1 to 1 : 2 and 1 : 3, the % yield of hydroquinone diacetate increased gently from 60.2%, to 77.5% to 93.2%, respectively, with 100% selectivity. However, a further increase in the acetic anhydride molar content was accompanied by a decrease in the percentage yield of hydroquinone diacetate to 83.1% (molar ratio 1 : 4). As acetic anhydride acts as both a solvent and reactant, an excess of acetic anhydride is used to facilitate the reaction until the molar ratio of hydroquinone to acetic anhydride is 1 : 3. Many factors may affect the catalyst activity; the dilution effect of acetic anhydride (solvent effect) and/or the catalyst active sites may be blocked due to a higher acetic anhydride concentration, thereby decreasing the yield of hydroquinone diacetate. Because the activation of acetic anhydride is the key step in the formation of hydroquinone diacetate, the decrease in % yield of hydroquinone diacetate may be due to the presence of low (molar ratio 1 : 1 and 1 : 2) or excess (molar ratio 1 : 4) activated acetic anhydride – Brønsted acid sites and/or activated acetic anhydride – Lewis acid sites

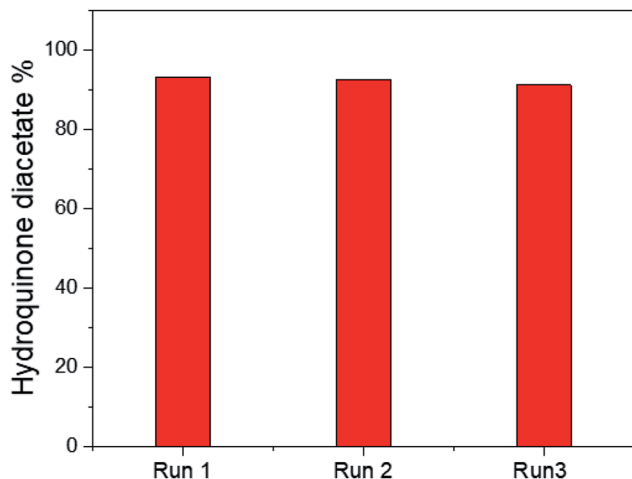


Fig. 9 Reusability test for 25% PO₄³⁻/mSnO₂.

intermediates. Therefore, 1 : 3 molar ratio of reactants was considered optimum for the further studies.

The reusability of the 25% PO₄³⁻/mSnO₂ catalyst was checked with three consecutive experiments by using the recovered catalyst. The catalyst was separated by filtration, washed with acetone several times and dried overnight at 100 °C. The reaction was reperformed on this reactivated catalyst. It was observed that there was no significant loss in activity, as shown in Fig. 9, and the catalytic activity of the catalyst was 93.2%, 92.4% and 91.1%, respectively. Moreover Fig. 7SA and B† show the TEM images of the 25% PO₄³⁻/mSnO₂ after the third run, which showed no change in the morphology of the reused catalyst; also the FTIR of the pyridine adsorbed on the reused sample showed no significant loss in the acidity. The obtained results indicated that the catalyst had excellent reusability over three consecutive runs.

Reaction mechanism

The first step in the formation of hydroquinone diacetate is chemisorption of the carbonyl group of acetic anhydride on the Brønsted acid sites or Lewis acid sites of the catalyst, then on the adsorbed acetic anhydride on the Brønsted acid sites, Scheme 3A shows a nucleophilic attack by OH of hydroquinone to form an intermediate that breaks down to lose acetic acid in

the next step. Finally, the proton of the Brønsted acid site is eliminated to give hydroquinone acetate. The hydroquinone acetate reacts with another acetic anhydride – Brønsted acid site to give the final product. The mechanism of the formation of hydroquinone diacetate on Lewis acid sites starts by the chemisorption of the carbonyl group on the acetic anhydride, which breaks down to form acylium ion. Then the OH of hydroquinone attacks acylium ion to form an intermediate, which lose a proton to give hydroquinone acetate, which undergoes the same steps to give the product hydroquinone diacetate, Scheme 3B.

Experimental

Synthesis of PO₄³⁻/mSnO₂

For the synthesis of mesoporous tin oxide (mSnO₂), 15 ml of 25 wt% NH₄OH solution was added to 100 ml solution containing 2.0 g CTAB dropwise and stirred for 15 min until the solution became homogeneous. 4.6 g SnCl₄ · 5H₂O was dissolved in 100 ml of distilled water and then added to NH₄OH CTAB solution slowly.^{65,66} The pH was fixed at 8.5. The product was filtered and washed with 2 wt% CH₃COONa solution until excess surfactant, and Cl⁻ ions were removed. The sample was then dried overnight at 800 °C.

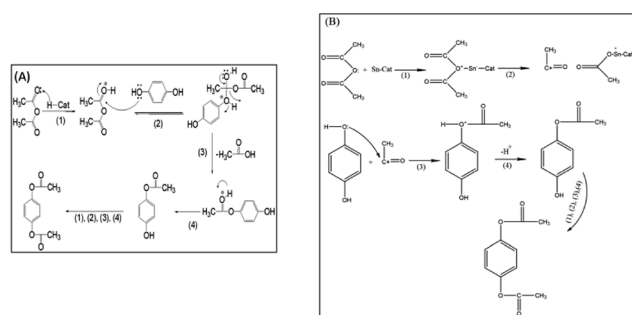
A series of 3, 8, 18, 25, 35, and 45 wt% PO₄³⁻-supported mSnO₂ samples were prepared by suspending 3.0 g of dried mSnO₂ in 25 ml of distilled water; then the calculated amount of 1 M H₃PO₄ solution was added dropwise to the suspension during stirring for 2 h. Then the samples were calcined at 400 °C, 450 °C, 550 °C and 650 °C for 4 h. 3, 8, 18, 25, 35 and 45 are denoted for the wt% of PO₄³⁻.

Characterization

DTA and TGA of the uncalcined samples were carried out using a Shimadzu thermal analyzer, type 50-H. The measurements were carried out under N₂ flow of 20 ml min⁻¹ with a heating rate of 10 °C min⁻¹. The XRD patterns were recorded using a PW 150 (Philips). The crystallite size (nm), unit cell *d*₁₁₀ and *a*⁰ were calculated from the strongest peak 110 of mSnO₂ at 2θ = 26.54° by Bragg's law.⁶⁷

$$n\lambda = 2d \sin(\theta)$$

$$a^0 = 2d/\sqrt{3}$$



Scheme 3 Reaction mechanism for the formation of hydroquinone diacetate on: (A) Brønsted acid sites of a solid acid catalyst, (B) Lewis acid sites of a solid acid catalyst.

TEM images of the calcined PO₄³⁻/mSnO₂ samples were obtained using a Joel JEM-1230 operated at 120 KV. FTIR spectra of the calcined samples were recorded by using MATTSON 5000 FTIR spectrophotometer (4 cm⁻¹ resolution and 16 scans) in dried KBr (Merck). Physically adsorbed nitrogen at -196 °C was used for elucidation of the texture properties of the calcined P/mSnO₂ samples. The total acidity of the solid samples was measured by suspending 0.05 g of solid catalyst in 10 ml of acetonitrile for 3 h and titrating with 0.1 N *n*-butylamine base in acetonitrile.^{35,57} An Orion 420 digital A pH-meter was used for measuring the potential variation. Lewis and Brønsted acid

levels were determined from the exposure of the degassed solid samples at 200 °C to dried pyridine.⁶⁸ After the removal of excess pyridine, a MATTSON 5000 FTIR spectrophotometer was used to record the spectra. For the NH₃-TPD experiment, 200 mg of the samples were kept under ammonia gas for 4 h at 100 °C. Then the samples were flushed with helium, and then the temperature was raised from 50 °C to 700 °C at a ramping rate of 5 °C min⁻¹ in the presence of helium, and the amount of NH₃ desorbed was recorded using a TCD detector.

Catalytic activity

The synthesis of hydroquinone diacetate (1,4-diacetoxybenzene) was used to test the catalytic activity of the samples. 0.1 g of the activated catalyst at 120 °C was added to 3 ml (0.032 mol) of acetic anhydride and 1.1 g (0.01 mol) of hydroquinone in a 50 ml flask, and the mixture then stirred for 15 min. After 15 min, the clear solution was filtered to separate the catalyst, and the filtrate was poured into 500 ml of crushed ice. The crystalline solids were dried to constant weight. The product was identified by FTIR, NMR spectroscopy and melting point determination. The % yield of hydroquinone diacetate was then calculated.

Conclusions

Mesoporous SnO₂ was successfully prepared through a simple method and loaded with different amounts of phosphate species. The XRD and TEM images showed that the samples were mesoporous. Phosphate species were homogeneously adsorbed on the surface of mesoporous SnO₂ up to surface saturation coverage at 25 wt% PO₄³⁻ at which point the surface, area surface acidity level and acid strength were at a maximum. The acidity of the samples was found to decrease when the phosphate loading exceeded 25 wt%, at which point polyphosphate multilayers were formed. Both Lewis and/or Brønsted acid sites were present on the catalyst surface. The synthesis of hydroquinone diacetate was used as a model reaction for testing the activity of the solid sample for the formation of bulky molecules. The maximum formation of hydroquinone diacetate and the surface acidities were found at 25 wt% phosphate loading and 400 °C calcination temperature, respectively. The catalyst could be successfully used many times without significant loss of its catalytic activity.

Conflicts of interest

There are no conflicts to declare.

References

- 1 A. I. Ahmed, S. A. El-Hakam, A. S. Khder and W. S. Abo El-Yazeed, *J. Mol. Catal. A: Chem.*, 2013, **366**, 99–108.
- 2 E. Lotero, Y. Liu, D. E. Lopez, K. Suwannakarn, D. A. Bruce and J. G. Goodwin, *Ind. Eng. Chem. Res.*, 2005, **44**, 5353–5363.
- 3 A. Corma and H. García, *Chem. Rev.*, 2003, **103**, 4307–4366.
- 4 X. Zhang, K. Wilson and A. F. Lee, *Chem. Rev.*, 2016, **116**, 12328–12368.
- 5 M. A. Rezvani, A. F. Shojaie and M. H. Loghmani, *Catal. Commun.*, 2012, **25**, 36–40.
- 6 A. Bordoloi and S. B. Halligudi, *J. Catal.*, 2008, **257**, 283–290.
- 7 D. Hua, S.-L. Chen, G. Yuan, Y. Wang, Q. Zhao, X. Wang and B. Fu, *Microporous Mesoporous Mater.*, 2011, **143**, 320–325.
- 8 M. B. Smith and J. March, *J. Med. Chem.*, 2007, **50**, 2279–2280.
- 9 S. K. Samantaray and K. Parida, *J. Mol. Catal. A: Chem.*, 2001, **176**, 151–163.
- 10 S. K. Bharadwaj, S. Hussain, M. Kar and M. K. Chaudhuri, *Appl. Catal., A*, 2008, **343**, 62–67.
- 11 S. K. Samantaray, T. Mishra and K. M. Parida, *J. Mol. Catal. A: Chem.*, 2000, **156**, 267–274.
- 12 S. K. Samantaray and K. Parida, *Appl. Catal., A*, 2001, **220**, 9–20.
- 13 S. Furuta, H. Matsushashi and K. Arata, *Catal. Commun.*, 2004, **5**, 721–723.
- 14 K. N. Rao, A. Sridhar, A. F. Lee, S. J. Tavener, N. A. Young and K. Wilson, *Green Chem.*, 2006, **8**, 790–797.
- 15 X. Chen, G. Clet, K. Thomas and M. Houalla, *J. Catal.*, 2010, **273**, 236–244.
- 16 G. Lu, X. Li, Z. Qu, Q. Zhao, H. Li, Y. Shen and G. Chen, *Chem. Eng. J.*, 2010, **159**, 242–246.
- 17 N. R. Shiju, M. AnilKumar, W. F. Hoelderich and D. R. Brown, *J. Phys. Chem. C*, 2009, **113**, 7735–7742.
- 18 R. Kourieh, S. Bennici, M. Marzo, A. Gervasini and A. Auroux, *Catal. Commun.*, 2012, **19**, 119–126.
- 19 R. M. Martín-Aranda and J. Čejka, *Top. Catal.*, 2010, **53**, 141–153.
- 20 N. Pal and A. Bhaumik, *RSC Adv.*, 2015, **5**, 24363–24391.
- 21 A. Taguchi and F. Schüth, *Microporous Mesoporous Mater.*, 2005, **77**, 1–45.
- 22 J. Fan, Y. Dai, Y. Li, N. Zheng, J. Guo, X. Yan and G. D. Stucky, *J. Am. Chem. Soc.*, 2009, **131**, 15568–15569.
- 23 Q. Yuan, A.-X. Yin, C. Luo, L.-D. Sun, Y.-W. Zhang, W.-T. Duan, H.-C. Liu and C.-H. Yan, *J. Am. Chem. Soc.*, 2008, **130**, 3465–3472.
- 24 L. Xu, H. Zhao, H. Song and L. Chou, *Int. J. Hydrogen Energy*, 2012, **37**, 7497–7511.
- 25 Z. Miao, H. Zhao, H. Song and L. Chou, *RSC Adv.*, 2014, **4**, 22509–22519.
- 26 M. Srinivas, G. Raveendra, G. Parameswaram, P. S. Sai Prasad, S. Loridant and N. Lingaiah, *J. Chem. Sci.*, 2015, **127**, 897–908.
- 27 G. M. Maksimov, G. S. Litvak, A. A. Budneva, E. A. Paukshtis, A. N. Salanov and V. A. Likhoholov, *Kinet. Catal.*, 2006, **47**, 564–571.
- 28 B. Malleshham, P. Sudarsanam, G. Raju and B. M. Reddy, *Green Chem.*, 2013, **15**, 478–489.
- 29 J. Zhu, B. Y. Tay and J. Ma, *Mater. Lett.*, 2006, **60**, 1003–1010.
- 30 M. Bhagwat, P. Shah and V. Ramaswamy, *Mater. Lett.*, 2003, **57**, 1604–1611.
- 31 Y. Leng, J. Liu, P. Jiang and J. Wang, *Chem. Eng. J.*, 2014, **239**, 1–7.

- 32 J. I. Moreno, R. Jaimes, R. Gómez and M. E. Niño-Gómez, *Catal. Today*, 2011, **172**, 34–40.
- 33 M. K. Lam and K. T. Lee, *Fuel*, 2010, **89**, 3866–3870.
- 34 M. K. Lam, K. T. Lee and A. R. Mohamed, *Appl. Catal., B*, 2009, **93**, 134–139.
- 35 Y. Wang, C. Ma, X. Sun and H. Li, *Microporous Mesoporous Mater.*, 2001, **49**, 171–178.
- 36 B. Tyagi, M. K. Mishra and R. V. Jasra, *J. Mol. Catal. A: Chem.*, 2007, **276**, 47–56.
- 37 R. Kim, H. S. Park, T. Yu, J. Yi and W.-S. Kim, *Chem. Phys. Lett.*, 2013, **575**, 71–75.
- 38 L. Körösi and I. Dékány, *Colloids Surf., A*, 2006, **280**, 146–154.
- 39 S. Furuta, H. Matsuhashi and K. Arata, *Appl. Catal., A*, 2004, **269**, 187–191.
- 40 M. Hayashi, T. Hyodo, Y. Shimizu and M. Egashira, *Sens. Actuators, B*, 2009, **141**, 465–470.
- 41 X. Fan, T. Yu, Y. Wang, J. Zheng, L. Gao, Z. Li, J. Ye and Z. Zou, *Appl. Surf. Sci.*, 2008, **254**, 5191–5198.
- 42 W. Wang, Y. Pang and S. N. B. Hodgson, *J. Sol-Gel Sci. Technol.*, 2010, **54**, 19–28.
- 43 V. L. Struzhko, E. V. Senchilo and V. G. Il'in, *Theor. Exp. Chem.*, 2006, **42**, 53–59.
- 44 A. Martín, R. A. García, D. S. Karaman and J. M. Rosenholm, *J. Mater. Sci.*, 2014, **49**, 1437–1447.
- 45 R. A. García-Muñoz, V. Morales, M. Linares, P. E. González, R. Sanz and D. P. Serrano, *J. Mater. Chem. B*, 2014, **2**, 7996–8004.
- 46 L. M. Fang, X. T. Zu, Z. J. Li, S. Zhu, C. M. Liu, L. M. Wang and F. Gao, *J. Mater. Sci.: Mater. Electron.*, 2008, **19**, 868–874.
- 47 S. Takenaka, R. Takahashi, S. Sato, T. Sodesawa, F. Matsumoto and S. Yoshida, *Microporous Mesoporous Mater.*, 2003, **59**, 123–131.
- 48 V. V. Fomichev and M. E. Poloznikova, *Russ. Chem. Rev.*, 1994, **63**, 399–409.
- 49 G. Ranga Rao, T. Rajkumar and B. Varghese, *Solid State Sci.*, 2009, **11**, 36–42.
- 50 F. M. Bautista, J. M. Campelo, A. García, D. Luna, J. M. Marinas, R. A. Quirós and A. A. Romero, *Appl. Catal., A*, 2003, **243**, 93–107.
- 51 L. Korösi, S. Papp, V. Meynen, P. Cool, E. F. Vansant and I. Dékány, *Colloids Surf., A*, 2005, **268**, 147–154.
- 52 G. Srinivasa Rao, N. Pethan Rajan, M. Hari Sekhar, S. Ammaji and K. V. R. Chary, *J. Mol. Catal. A: Chem.*, 2014, **395**, 486–493.
- 53 R. Gutiérrez-Báez, J. A. Toledo-Antonio, M. A. Cortes-Jácome, P. J. Sebastian and A. Vázquez, *Langmuir*, 2004, **20**, 4265–4271.
- 54 S. Benadji, P. Eloy, A. Leonard, B. L. Su, C. Rabia and E. M. Gaigneaux, *Microporous Mesoporous Mater.*, 2012, **154**, 153–163.
- 55 A. I. Ahmed, S. A. El-Hakam, M. A. A. Elghany and W. S. A. El-Yazeed, *Appl. Catal., A*, 2011, **407**, 40–48.
- 56 E. Drzymała, G. Gruzeł, J. Depciuch, A. Budziak, A. Kowal and M. Parlinska-Wojtan, *J. Phys. Chem. Solids*, 2017, **107**, 100–107.
- 57 D. O. Bennardi, G. P. Romanelli, J. C. Autino and L. R. Pizzio, *Appl. Catal., A*, 2007, **324**, 62–68.
- 58 L. R. Pizzio, P. G. Vázquez, C. V. Cáceres and M. N. Blanco, *Appl. Catal., A*, 2003, **256**, 125–139.
- 59 L. Marosi, G. Cox, A. Tenten and H. Hibst, *J. Catal.*, 2000, **194**, 140–145.
- 60 T.-H. Kim, M. Jang and J. K. Park, *Microporous Mesoporous Mater.*, 2008, **108**, 22–28.
- 61 J. W. Harris, M. J. Cordon, J. R. Di Iorio, J. C. Vega-Vila, F. H. Ribeiro and R. Gounder, *J. Catal.*, 2016, **335**, 141–154.
- 62 Y. Shimizu, T. Hyodo and M. Egashira, *J. Eur. Ceram. Soc.*, 2004, **24**, 1389–1398.
- 63 A. Yu and R. Frech, *J. Power Sources*, 2002, **104**, 97–100.
- 64 C. Morterra, G. Cerrato, S. Ardizzone, C. L. Bianchi, M. Signoretto and F. Pinna, *Phys. Chem. Chem. Phys.*, 2002, **4**, 3136–3145.
- 65 W. Yu-de, M. Chun-lai, S. Xiao-dan and L. Heng-de, *Nanotechnology*, 2002, **13**, 565.
- 66 A. S. Khder and A. I. Ahmed, *Appl. Catal., A*, 2009, **354**, 153–160.
- 67 P. Scherrer, *Bestimmung der Größe und der inneren Struktur von Kolloidteilchen mittels Röntgenstrahlen*, Nachrichten von der Gesellschaft der Wissenschaften zu Göttingen, Mathematisch-Physikalische Klasse, 1918, vol. 2, pp. 98–100.
- 68 A. S. Khder, E. A. El-Sharkawy, S. A. El-Hakam and A. I. Ahmed, *Catal. Commun.*, 2008, **9**, 769–777.