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Silicate silver/flower-like magnalium hydroxide composites for enhanced visible light photodegradation activities

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Flower-like magnalium layered composites (MgAl-LDH) were first fabricated by a hydrothermal method, and a series of AgSiO$_x$/MgAl-LDH composites with different mole ratios was successfully prepared by loading AgSiO$_x$ on the MgAl-LDH surface. The photocatalytic activities of the composites for the degradation of methylene blue (MB) were investigated under visible light irradiation (xenon lamp). The prepared AgSiO$_x$/MgAl-LDH composites were characterized by X-ray diffraction (XRD), scanning electron microscopy (SEM), transmission electron microscopy (TEM), UV-vis diffuse reflectance spectroscopy, thermogravimetric analysis (TG), and X-ray photoelectron spectroscopy (XPS). The results revealed that AgSiO$_x$/MgAl-LDH (5 : 1) displayed a photocatalytic degradation efficiency of 99.7% in 30 min for MB. The photocatalytic degradation rate was higher than that of pure AgSiO$_x$.

Introduction

The increase in demand for textiles, leather, and other industries has resulted in the generation of a large number of organic dye pollutants, which are difficult to treat because of their large volumes and high concentrations. Therefore, the disposal of dye wastewater in an environmentally friendly manner with high efficiency has become a matter of pressing concern.[1–4] Among the multiple approaches for handling dye wastewater, photocatalytic technology can decompose organic matter into inorganic ions and water at normal temperature and pressure without secondary pollution.[5,6] Hence, semiconductor photocatalysts have received significant attention since Fujishima *et al.*[7] reported in 1972 that TiO$_2$ electrodes can achieve photocatalytic splitting of water. However, the practical application of traditional semiconductor photocatalysts has been limited due to their wide band gaps, quick combination of photogenerated holes, and narrow absorption spectra.[8,9] Thus, it is desirable to develop a highly efficient photocatalyst that is active in the visible light to improve the utilization of sunlight.

Silver silicate (AgSiO$_x$) is a novel type of photocatalyst developed in recent years. There are three kinds of ligands in silver silicate, namely, AgO$_2$, AgO$_3$, and AgO$_4$, which are favorable for the separation of photoelectrons and holes; the narrow band gap (1.58 eV) results in very strong photocatalytic activity in the entire visible light region (λ < 740 nm), and the

photocatalytic degradation efficiency of methylene blue (MB) is much higher than those of Ag$_2$O and Ag$_3$PO$_4$ under visible light.[10] In addition, silver silicate can be combined with other materials such as Ag$_6$Si$_2$O$_7$/WO$_3$,[11] Ag$_6$Si$_2$O$_7$/α-Fe$_2$O$_3$,[12] and Ag$_6$Si$_2$O$_7$@RGO to improve the photocatalytic performance.[13] However, the high cost of silver salts limits its application. Therefore, herein, to reduce the cost and enhance photocatalytic activity, we choose layered double hydroxides as a catalyst carrier.

Layered double hydroxides (LDHs) are a class of anionic clays or layered compounds. The general formula of LDHs can be represented by $[M^{2+}_{1-x}M^{3+}_x(OH)_2]^{x+}(A^{n-})_{n/x} \cdot y H_2O$, where M^{2+} and M^{3+} are the divalent and trivalent metal cations, respectively, and A^{n-} is an interlayer anion.[14–16] LDHs have received much interest because of their special layered structures, stability, exchangeable ions in the interlayer, and large surface area, due to which they can be used as ideal candidates for the preparation of composites with catalysts or catalyst supports.[17–21] In addition, LDHs can display unique optical properties through the introduction of functional anions or upon organo-modification, which can be potentially applied in the fields of fluorescence,[22–24] detection,[25] UV-blocking, and UV absorption materials.[26,27]

In particular, MgAl-LDH is the most typical representative of different kinds of hydrotalcites. However, MgAl-LDH alone shows very low photocatalytic activity under visible light irradiation. Fortunately, the layered structure of LDHs has been proven to be beneficial for inducing electron migration and inhibition of electron–hole pair recombination; additionally, the large number of hydroxyl groups on LDH laminates can generate more highly active ·OH species, which are beneficial

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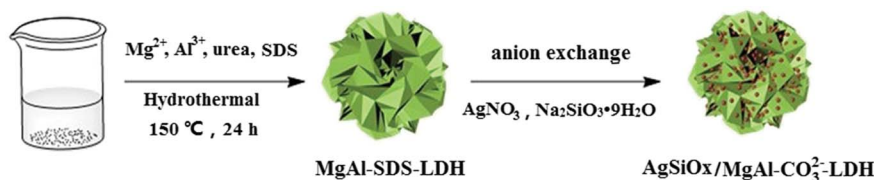
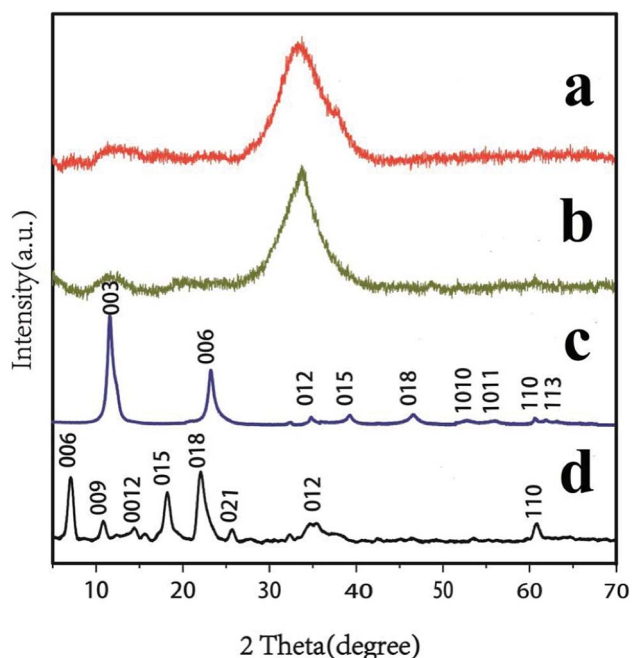**Fig. 1** Schematic diagram of the process of the synthesis of AgSiO$_x$/MgAl-LDH.

Fig. 2 XRD patterns of (a) AgSiO$_x$/MgAl-LDH (5 : 1); (b) AgSiO$_x$; (c) MgAl–CO$_3^{2-}$-LDH; and (d) MgAl–DS-LDH.

for improving the quantum efficiency of the photocatalytic reaction.[28,29] A wide variety of semiconductor and MgAl-LDH composites such as TiO$_2$/MgAl-LDH,[30] SnO$_2$/MgAl-LDH,[31] and
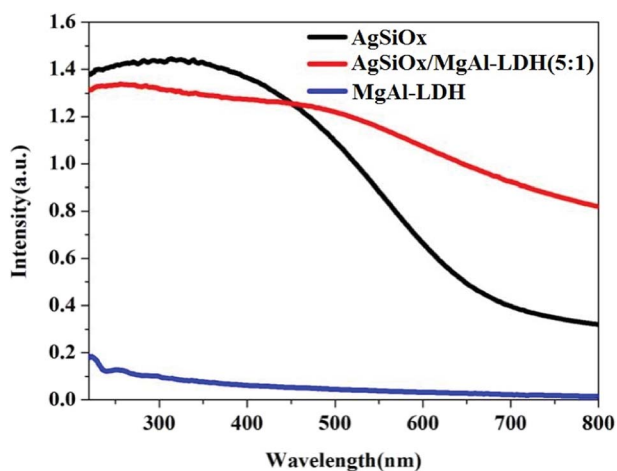
CeO$_2$/MgAl-LDH[32] has been reported for various photocatalytic applications.

Herein, a flower-like MgAl-LDH sample was prepared by a simple hydrothermal method. A series of AgSiO$_x$/MgAl-LDH photocatalysts was synthesized *in situ* using a one-step approach. The composite structures were characterized by XRD, SEM, EDS, TEM, TG, and XPS. Finally, photodegradation tests and the mechanism of AgSiO$_x$/MgAl-LDH photocatalysts were discussed.

Experimental section

Materials

Analytical grade Mg(NO$_3$)$_2$·6H$_2$O, Al(NO$_3$)$_3$·9H$_2$O, urea, Na$_2$-SiO$_3$·9H$_2$O, AgNO$_3$, methylene blue (MB), hexadecyl-trimethylammonium bromide (CTAB), trichloromethane, ethylenediaminetetraacetic acid disodium salt (EDTA-2Na), and *tert*-butanol were purchased and used without further purification.

Synthesis of MgAl–DS-LDH

Mg(NO$_3$)$_2$·6H$_2$O (0.0016 mol), 0.0008 mol Al(NO$_3$)$_3$·9H$_2$O, 0.0016 mol SDS, and 0.008 mol urea were dissolved in 960 mL deionized water. After magnetic stirring for 30 min, the solution was transferred to a 150 mL hydrothermal reactor and subsequently heated at 150 °C for 24 h. After cooling down to room temperature, the precipitate was filtered and washed thrice with
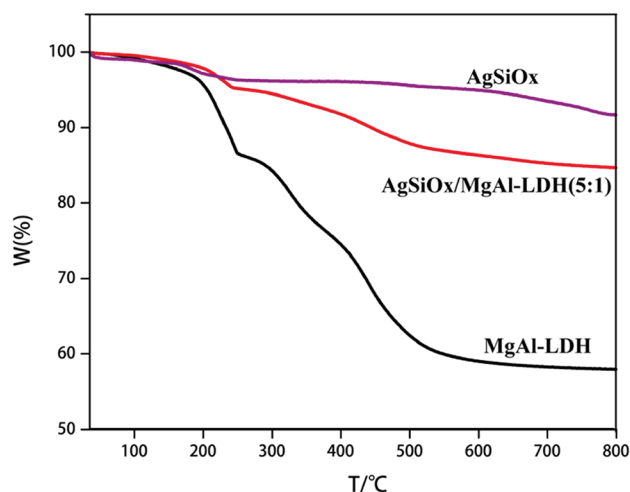

Fig. 3 UV-vis spectra of AgSiO$_x$, MgAl-LDH, and AgSiO$_x$/MgAl-LDH (5 : 1).

Fig. 4 TG curves of AgSiO$_x$, AgSiO$_x$/MgAl-LDH (5 : 1), and MgAl-LDH.
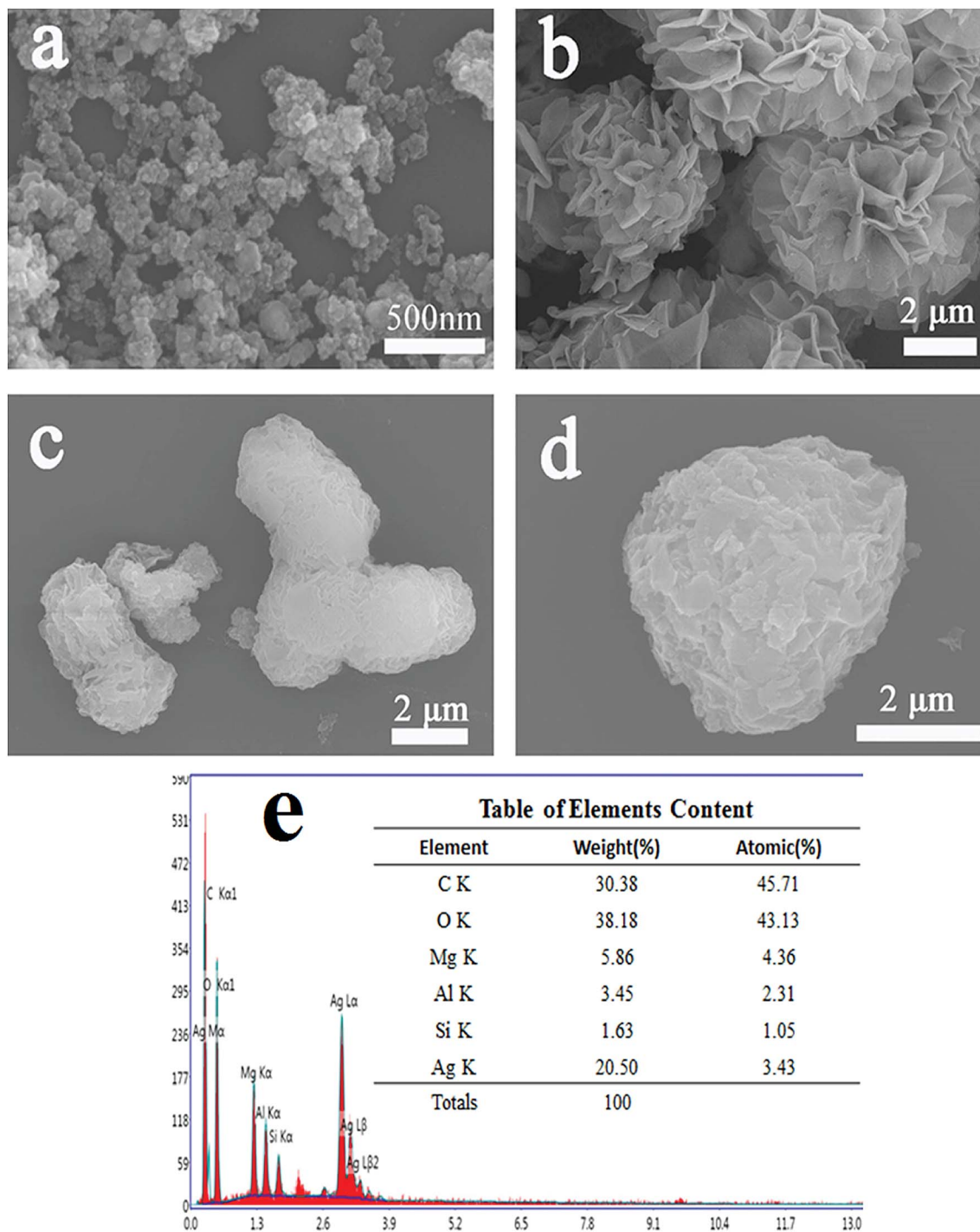
Fig. 5 SEM images of (a) AgSiO$_x$; (b) MgAl-LDH; (c and d) AgSiO$_x$/MgAl-LDH (5 : 1); and (e) EDS of AgSiO$_x$/MgAl-LDH (5 : 1).

ethanol and deionized water. Finally, the white products were dried in a vacuum oven at 80 °C for 8 h.

Synthesis of MgAl–CO$_3^{2-}$-LDH

CTAB (0.15 mol L^{-1}, 50 mL), 0.2 mol L^{-1} Na$_2$CO$_3$ (100 mL), 60 mL chloroform, and 2 g MgAl–DS-LDHs were mixed by magnetic stirring for 1 h. Then, the products were washed thrice with deionized water and ethanol and dried at 80 °C.

Synthesis of AgSiO$_x$

A total of 150 mL of deionized AgNO$_3$ solution (solution A) and 50 mL of deionized Na$_2$SiO$_3$·9H$_2$O solution (solution B) was prepared with a concentration of 0.1 mol L^{-1}. Under magnetic stirring, the solution B was added dropwise to the solution A using a constant pressure funnel. The mixture was further stirred for 2 h. The final reddish brown products were obtained by centrifuging and washing thrice with deionized water and

ethanol; the products were then dried at room temperature (Fig. 1).

Synthesis of AgSiO$_x$/MgAl-LDH

The AgSiO$_x$/MgAl-LDH composites were prepared by *in situ* deposition. In a typical procedure, 0.5 g of MgAl–CO$_3^{2-}$-LDH powder and a certain amount of AgNO$_3$ solution (0.1 mol L^{-1}) were added to 150 mL water. The solution mixture was magnetically stirred for 30 min and then, a certain amount of 0.1 mol L^{-1} Na$_2$SiO$_3$·9H$_2$O solution was added dropwise. After two hours of reaction, the resultant product was separated by filtration, washed thrice with deionized water and ethanol, and dried at room temperature. Changing the amount of AgNO$_3$ and Na$_2$SiO$_3$·9H$_2$O solutions yielded composite materials with different mass ratios of AgSiO$_x$ and MgAl-LDH (3 : 1, 4 : 1, 5 : 1, 6 : 1, 10 : 1).

Characterization

Microstructures of the nanocomposites were observed with a Zeiss Supra 55 field-emission scanning electron microscope (SEM) and a JEOL JEM-3010 transmission electron microscope (TEM). X-ray diffraction (XRD) measurements were carried out on a Rigaku D/Max 2500 diffractometer with Cu Kα radiation (λ = 1.54 Å) at a generator voltage of 40 kV and a generator current of 40 mA. AgSiO$_x$, MgAl-LDH, and AgSiO$_x$/MgAl-LDH were characterized using a Thermo VG RSCAKAB 250X high-resolution X-ray photoelectron spectroscope (XPS). UV-vis diffuse reflectance spectra were recorded on a UV-vis spectrometer (UV-2450 Shimadzu). Thermogravimetric analysis (TG) was conducted using a STA449-type thermogravimetric analyzer under nitrogen flow; the heating rate was 20 °C min^{-1}, and the gas flow velocity was 27 mL min^{-1}.

Photocatalytic measurements

The photocatalytic experiments were carried out in a photocatalytic reaction chamber under a CEL-HXUV300 xenon lamp with a cutoff filter (λ > 420 nm), and the photocatalytic activities were evaluated by the degradation of MB. A total of 100 mg of photocatalyst was dispersed into 50 mL of MB (40 mg L^{-1}) solution, and the mixture was stirred in the dark for 30 min to achieve adsorption equilibrium. The reaction temperature was maintained at 20 °C, and 5 mL of the solution mixture was

taken out every 5 min and centrifuged (8000 rpm). The supernatant was tested using a UV-6100s Double BEAM spectrophotometer.

Results and discussion

XRD analysis

The XRD patterns of MgAl–DS-LDH, MgAl–CO$_3^{2-}$-LDH, AgSiO$_x$, and AgSiO$_x$/MgAl-LDH (5 : 1) are shown in Fig. 2. As shown in Fig. 2b, AgSiO$_x$ retained the unique characteristic diffraction peak in accordance with the standard card (JCPDS no. 85-0281), and this was consistent with previously reported results,[10] indicating the successful synthesis of AgSiO$_x$. As shown in Fig. 2c, the pattern of MgAl–CO$_3^{2-}$-LDH exhibited sharp and clear peaks at 11.71, 23.58, 39.67, 47.10, 47.10 and 47.10°, which corresponded to the basal reflections of (003), (006), (015), (018), (110), and (113). The structure of our compound matched with the standard structure of MgAl–CO$_3^{2-}$-LDH (PDF#35-0964), suggesting the synthesis of MgAl–CO$_3^{2-}$-LDHs.[33] The XRD patterns (Fig. 2d) illustrate that C$_{12}$H$_{25}$SO$_3^-$ anions were intercalated in MgAl-LDH, which was consistent with previously reported results.[34] When the prepared AgSiO$_x$ was uniformly loaded on the MgAl-LDH surface, the XRD pattern (Fig. 2a) of the composites mainly exhibited the features of silver silicate; the diffraction peak signals of MgAl-LDH were obscured due to its small particles, high signals, and low crystallinity. However, MgAl-LDH as a carrier did not affect the structure of AgSiO$_x$.

UV-vis analysis

The UV-visible spectra of AgSiO$_x$/MgAl-LDH (5 : 1), AgSiO$_x$, and MgAl-LDH were recorded, and the results are shown in Fig. 3. MgAl-LDH exhibited low absorption in the ultraviolet region and almost no absorption in the visible region. For silver silicate, apart from the ultraviolet region, it also exhibited strong absorption in the visible light region (λ > 420 nm), thus showing good photocatalytic efficiency in the visible light region. The absorptions of the composites and pure AgSiO$_x$ in the ultraviolet light region were similar. However, the absorption of the composite in the visible region of λ > 420 nm was much higher than those of pure AgSiO$_x$ and MgAl-LDH, indicating that MgAl-LDH could enhance the absorption in the visible light region as
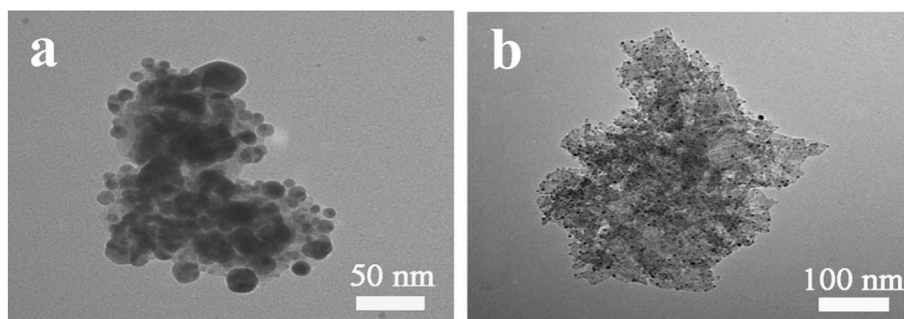

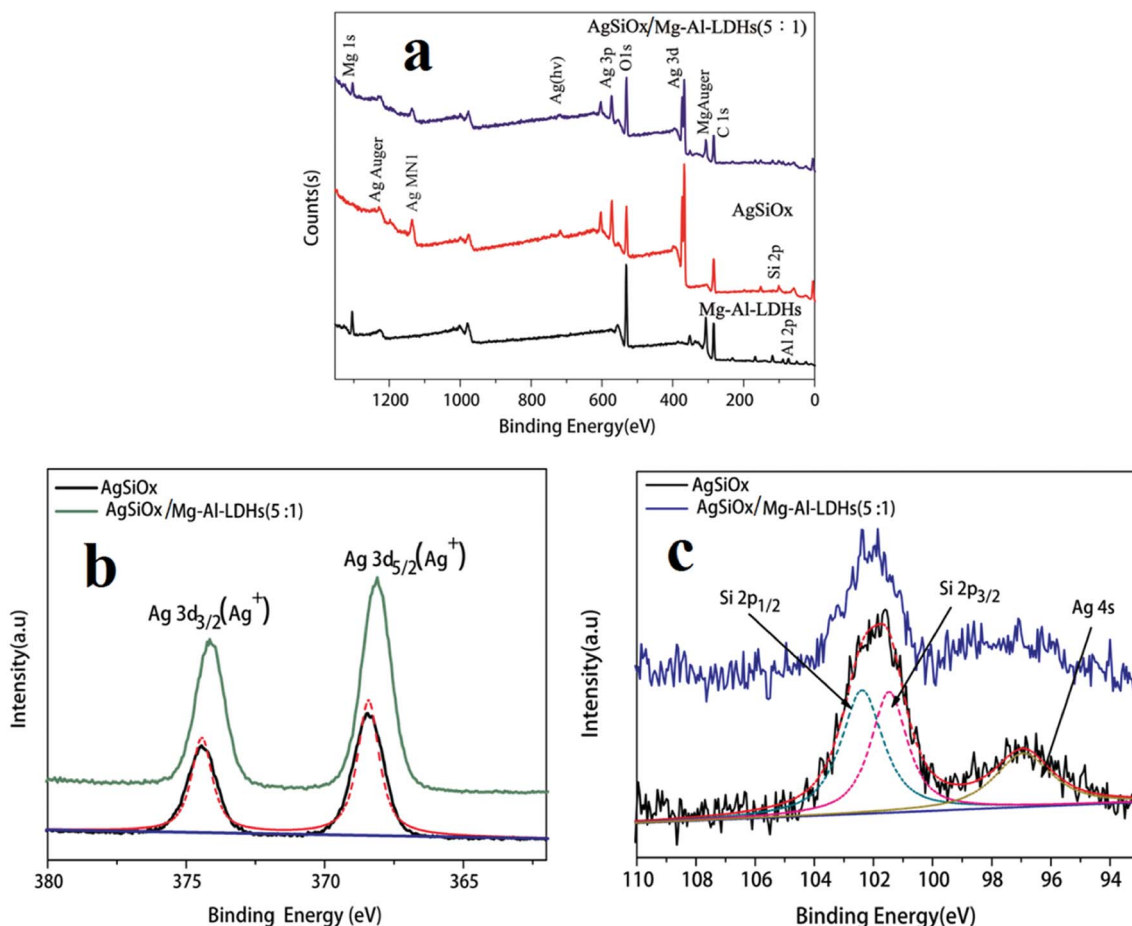Fig. 6   TEM images of (a) AgSiO$_x$; (b) AgSiO$_x$/MgAl-LDH (5 : 1).

Fig. 7 (a) The complete XPS spectra of pure AgSiO$_x$, MgAl-LDH, AgSiO$_x$/MgAl-LDH (5 : 1); (b) main peaks of Ag 3d$_{5/2}$ and Ag 3d$_{3/2}$ for pure AgSiO$_x$ and AgSiO$_x$/MgAl-LDH (5 : 1); (c) main peaks of Si 2p$_{3/2}$ and Si 2p$_{5/2}$ for pure AgSiO$_x$ and AgSiO$_x$/MgAl-LDH.

a carrier, which was beneficial for the enhancement of photocatalytic efficiency.

TG analysis

As shown in Fig. 4, AgSiO$_x$ started exhibiting weight loss at 200 °C, with a mass loss of about 8.32%, which was mainly caused by the removal of physically adsorbed water. The thermal decomposition of MgAl-LDH mainly consisted of two stages: the first weight loss at temperatures below 250 °C was mainly due to the loss of physically adsorbed water and interlaminar water, corresponding to a mass loss of ~13.65%. The

second stage was seen in the temperature range of 250–550 °C. In this process, interlayer water, interlayer hydroxyl removal and interlayer anion decomposition, such as CO_3^{2-} decomposition to release CO_2, were mainly carried out, and the mass loss at this stage was about 25.79%.[20,21,28,29] With the addition of AgSiO$_x$, the weight loss ratio of the composite photocatalyst decreased, which also illustrated that the above-mentioned two materials have been effectively combined.

SEM and TEM analysis

The SEM image of the prepared pure AgSiO$_x$ is displayed in Fig. 5a. The microscopic morphology of pure AgSiO$_x$ was a solid pellet of agglomerate. The prepared MgAl-LDH with spherical and flower-like structure is shown in Fig. 5b. During the formation of the flower-like MgAl-LDH sample, excessive surfactants formed micelles in the solution and imposed a bending force on the growth of LDH; thus, the LDH sheets grew along the curved micellar interface. Thus, a unique flower-like structure with a size of about 3–5 microns was formed,[34] which was much larger than that of pure AgSiO$_x$; it possessed high specific surface area and adsorption capacity, which are

Table 1 Decolorization rate constants of MB obtained for different photocatalysts

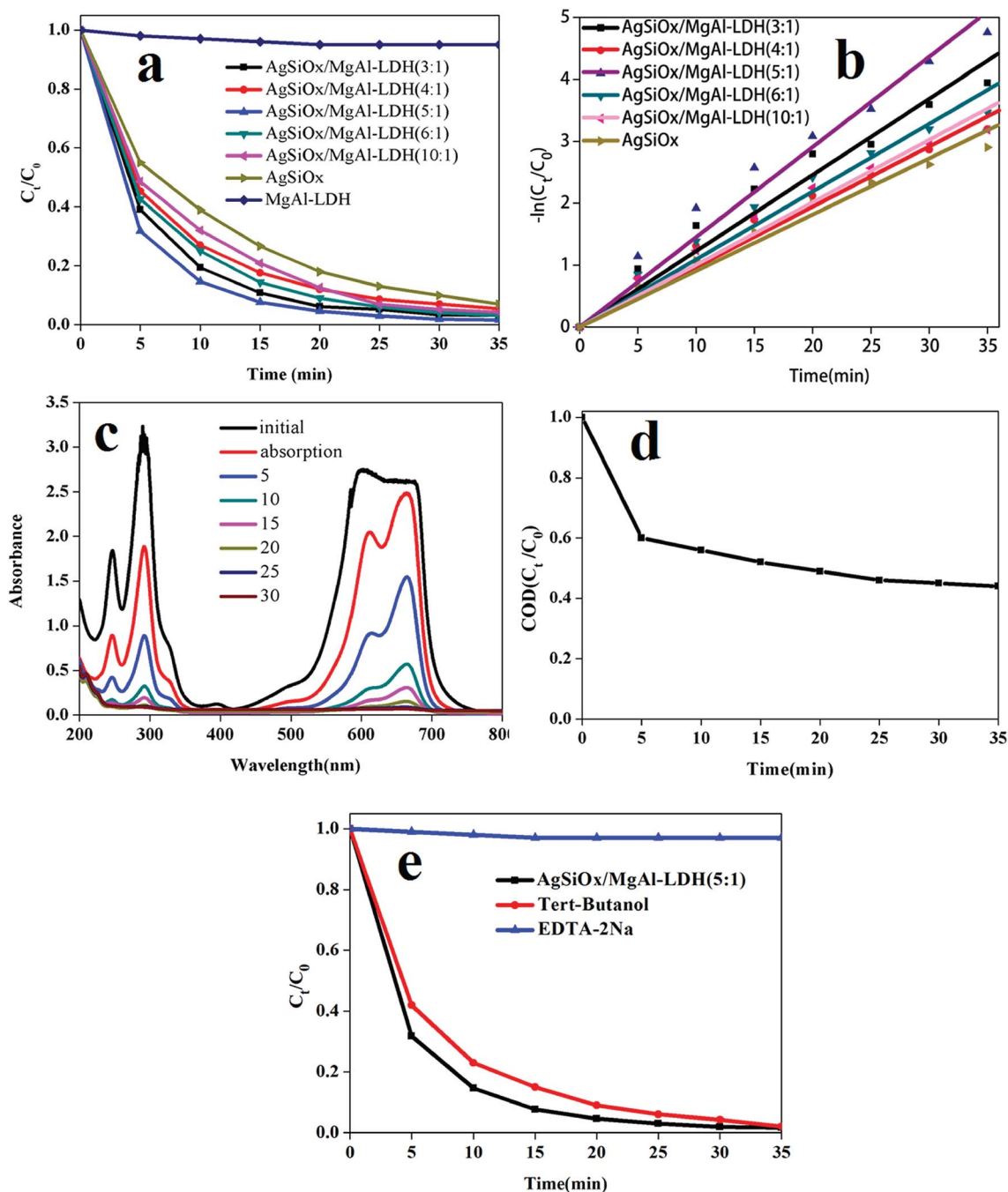| Samples	k (min^{-1})
AgSiO$_x$/Mg–Al-LDHs (3 : 1)	0.1159
AgSiO$_x$/Mg–Al-LDHs (4 : 1)	0.1007
AgSiO$_x$/Mg–Al-LDHs (5 : 1)	0.1367
AgSiO$_x$/Mg–Al-LDHs (6 : 1)	0.1094
AgSiO$_x$/Mg–Al-LDHs (10 : 1)	0.1138
AgSiO$_x$	0.0935
Mg–Al-LDHs	—

Fig. 8 (a) Photocatalytic degradation curves; (b) curve of the fitting equation of photocatalytic degradation rate; (c) the UV-visible absorption spectra of MB during photodegradation process with AgSiO$_x$/MgAl-LDH (5 : 1); (d) the curve of COD (C_t/C_0) in the process of photodegradation of MB with AgSiO$_x$/MgAl-LDH (5 : 1); (e) photodegradation of MB by AgSiO$_x$/MgAl-LDH (5 : 1) with the addition of EDTA-2Na or *tert*-butanol.

beneficial for the dispersal of AgSiO$_x$ and enriching of the pollutants.

The SEM images of the prepared AgSiO$_x$/MgAl-LDH (5 : 1) composite are shown in Fig. 5c and d. We can clearly see that AgSiO$_x$ was uniformly and densely loaded on the surface of MgAl-LDH. The EDS spectrum of AgSiO$_x$/MgAl-LDH (Fig. 5e) shows the presence of Mg, Al, O, Si, and Ag, with atomic percentages of 4.36%, 2.31%, 43.13%, 1.05% and 3.43%, respectively; these values indicated that the estimated atomic

ratio of Mg to Al was ∼2 : 1 and Ag to Si was ∼3 : 1, which confirmed that the prepared samples are AgSiO$_x$/MgAl-LDHs.

The TEM images of the prepared pure AgSiO$_x$ and AgSiO$_x$/MgAl-LDH (5 : 1) composite are shown in Fig. 6. We can see that AgSiO$_x$ was composed of solid spherical particles with a size of 30–50 nm, and it could easily be aggregated. The TEM image of AgSiO$_x$/MgAl-LDH (5 : 1) is shown in Fig. 6b, and it reveals that the AgSiO$_x$ particles were highly distributed on the MgAl-LDH surface or in the fractures, and the size of AgSiO$_x$ was smaller

than that of pure $AgSiO_x$; this indicated that MgAl-LDH can improve the dispersibility of $AgSiO_x$.

XPS analysis

The full-scan XPS spectra of pure $AgSiO_x$, MgAl-LDH, and $AgSiO_x$/MgAl-LDH (5 : 1) materials are shown in Fig. 7a. The Al 2p and Mg 1s peaks were observed for MgAl-LDH; Ag 3d, Ag 3p, O 1s, and Si 2p peaks were seen for pure $AgSiO_x$. All the above-mentioned characteristic peaks could be found for the $AgSiO_x$/MgAl-LDH (5 : 1) composite. The high resolution XPS spectrum for the Ag 3d peak is shown in Fig. 7b. The peaks located at 374.4 and 368.3 eV correspond to Ag $3d_{3/2}$ and Ag $3d_{5/2}$, which were detected for $AgSiO_x$ and $AgSiO_x$/MgAl-LDH. The XPS diffraction peaks of Si 2p for pure $AgSiO_x$ and $AgSiO_x$/MgAl-LDH (5 : 1) are shown in Fig. 7c; the diffraction peaks resulted from Si $2p_{1/2}$, Si $2p_{3/2}$, and Ag 4s. The divided peaks were located at 102.3, 101.2, and 97.1 eV for Si $2p_{1/2}$, Si $2p_{3/2}$, and Ag 4s, respectively, and these observations were consistent with previously reported results.[12] The XPS results further illustrated the successful preparation of the composite materials.

Photocatalytic performance

The performance of the samples for photodegradation of MB was studied in a photocatalytic reaction chamber under a CEL-HXUV300 xenon lamp with a cutoff filter (λ > 420 nm). As shown in Fig. 7a, all the samples except pure MgAl-LDH exhibited good performance. All the composites showed better degradation efficiency than pure $AgSiO_x$, and $AgSiO_x$/MgAl-LDH (5 : 1) showed the best performance under the experimental conditions. As shown in Fig. 7b, upon fitting the correlation data, the performance of the samples in the photodegradation process was found to be consistent with the first order reaction kinetics:

$$\ln(C_0/C_t) = kt$$

here, C_0 is the concentration of MB at time $t = 0$ (as the light is turned on, after the period in the dark), C_t is the concentration of MB at photocatalytic reaction time t, and k is the decolorization rate constant. The MB decolorization rate constants under different conditions are shown in Table 1. The results demonstrated that the photocatalytic rate of $AgSiO_x$/MgAl-LDH (5 : 1) was higher than that of pure silver, and the full UV-visible absorption spectra for MB photodegradation are shown in Fig. 8c.

To further explore the degradation of methylene blue, the COD_{Gr} (O_2, mg L^{-1}) values at different reaction times ($t = 0$–35 min, at 5 minutes intervals) were measured and shown in Fig. 8d. The experimental results showed that the $AgSiO_x$/MgAl-LDH (5 : 1) catalyst exhibits good COD removal performance for methylene blue (56%) after 35 minutes.

To investigate the active species in the photocatalytic process, a photoinduced hole removal agent EDTA-2Na and hydroxyl scavenger tert-butanol were added to the reaction mixture. The addition of EDTA-2Na almost stopped the photo-catalytic degradation process (Fig. 8e), whereas the addition of tert-butanol hardly changed the photocatalytic degradation

process. The results illustrated that photogenerated holes were the main active species in the photocatalytic degradation process, which was consistent with previously reported results.[13]

To further evaluate the cyclic performance of $AgSiO_x$/MgAl-LDH (5 : 1), it was repeatedly used for three cycles; the specific performance (Fig. 9) revealed good stability and reus-ability. It is easy to find a conclusion from the diagram of cycling runs that if we want to get the same photocatalytic degradation rate, we should extend the reaction time in the photocatalytic cycle test; this could be because some of the Ag^+ ions are reduced to Ag^0, which is frequently observed in most Ag-based compounds. Additionally, after each cycle test, a large amount of water is used to wash the samples, which can result in the loss of catalyst, leading to the decrease in photocatalytic efficiency (Fig. 10).

There have been many studies on the degradation mechanism of methylene blue.[35,36] Based on the above-mentioned results, the following photocatalytic degradation mechanism was proposed. Under visible light irradiation, $AgSiO_x$ generated electron–hole pairs,[11–13] and the holes reacted with abundant hydroxyl groups on the surface of LDH to generate hydroxyl radicals ($^\cdot$OH) with strong oxidability.[37–39] The excited electrons were captured by dissolved oxygen species in aqueous solution to form superoxide radicals ($^\cdot O_2^-$).[40,41] The reactive $^\cdot O_2^-$ and $^\cdot$OH species with high activities degraded organic matter into inorganic small molecules.[42,43] In addition, the particle size of $AgSiO_x$ was reduced to reveal more active sites, and the contact area with the target pollutants was increased, which was bene-ficial for the enhancement of photocatalytic efficiency. The photocatalytic decolorization reactions of MB can be written as follows:
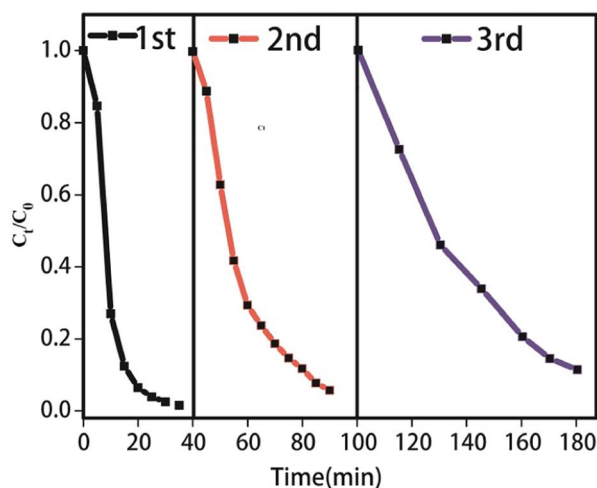
$$AgSiO_x/LDH/MB + h\nu \rightarrow AgSiO_x/LDH/MB^* \qquad (1)$$

$$AgSiO_x/LDH/MB^* \rightarrow AgSiO_x(e_{CB}^-)/LDH/MB^+ \qquad (2)$$

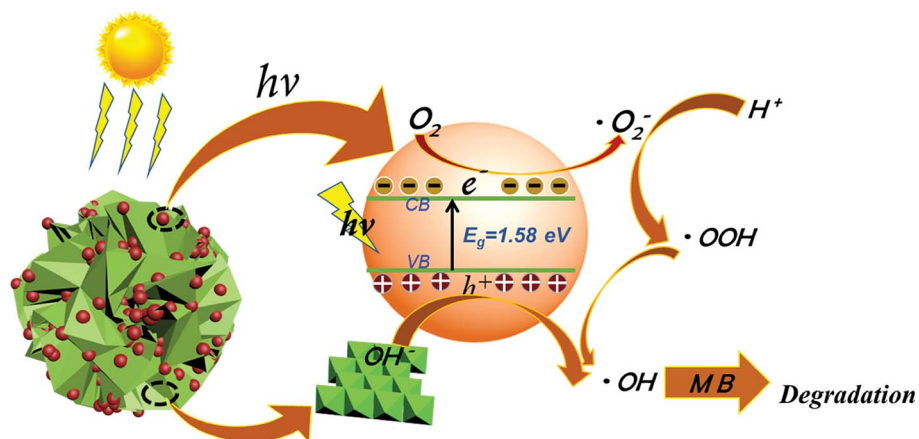Fig. 9   Cycling runs of $AgSiO_x$/MgAl-LDH (5 : 1) composite for the degradation of MB.

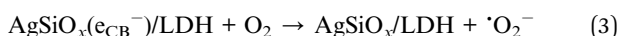**Fig. 10** Proposed photocatalytic mechanism of AgSiO$_x$/MgAl-LDH under visible light irradiation.

$$AgSiO_x(e_{CB}^-)/LDH + O_2 \rightarrow AgSiO_x/LDH + {}^\bullet O_2^- \quad (3)$$

$${}^\bullet O_2^- + H^+ \rightarrow HOO^\bullet \rightarrow {}^\bullet OH \quad (4)$$

$$h_{VB}^+ + OH^- \rightarrow {}^\bullet OH \quad (5)$$

$${}^\bullet OH + MB \rightarrow degradation\ products \quad (6)$$
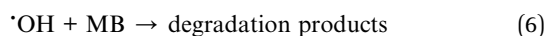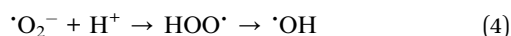
Conclusions

MgAl-LDH having a flower-like structure was prepared by a hydrothermal method, and a series of AgSiO$_x$/MgAl-LDH composites with different mole ratios was successfully obtained by loading AgSiO$_x$ on the MgAl-LDH surface. The photocatalytic activities of these composites for the photodegradation of MB were investigated under visible light irradiation. The results showed that the photodegradation efficiency of MB by AgSiO$_x$/MgAl-LDH (5 : 1) was nearly 99.7% in 30 min, and the photodegradation rate was higher than that of pure AgSiO$_x$. Meanwhile, AgSiO$_x$/MgAl-LDH (5 : 1) maintained good photocatalytic efficiency after 3 cycling experiments. The hydroxyl-rich structure and adsorption capacity of MgAl-LDH are important contributors for improving the photocatalytic ability.

Conflicts of interest

There are no conflicts to declare.

Acknowledgements

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