

Dibismuthates as Linking Units for Bis-Zwitterions and Coordination Polymers

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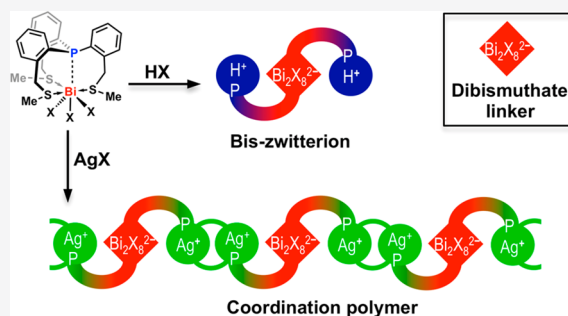
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ABSTRACT: Adducts of bismuth trihalides BiX_3 ($X = \text{Cl}, \text{Br}, \text{I}$) and the PS_3 ligand ($\text{PS}_3 = \text{P}(\text{C}_6\text{H}_4\text{-}o\text{-CH}_2\text{SCH}_3)_3$) react with HCl to form inorganic/organic hybrids with the general formula $[\text{HPS}_3\text{BiX}_4]_2$. On the basis of their solid-state structures determined by single-crystal X-ray diffraction, these compounds exhibit discrete bis-zwitterionic assemblies consisting of two phosphonium units $[\text{HPS}_3]^+$ linked to a central dibismuthate core $[\text{Bi}_2\text{X}_6]^{2-}$ via $\text{S} \rightarrow \text{Bi}$ dative interactions. Remarkably, the phosphorus center of the PS_3 ligand undergoes protonation with hydrochloric acid. This is in stark contrast to the protonation of phosphines commonly observed with hydrogen halides resulting in equilibrium. To understand the important factors in this protonation reaction, ^{31}P NMR experiments and DFT computations have been performed. Furthermore, the dibismuthate linker was utilized to obtain the coordination polymer $\{[\text{AgPS}_3\text{BiCl}_3(\text{OTf})_2](\text{CH}_3\text{CN})_2\}_\infty$, in which dicationic $[\text{Ag}(\text{PS}_3)]_2^{2+}$ macrocycles containing five-coordinate silver centers connect the dianionic $[\text{Bi}_2\text{Cl}_6(\text{OTf})_2]^{2-}$ dibismuthate fragments. The bonding situation in these dibismuthates has been investigated by single-crystal X-ray diffraction and DFT calculations (NBO analysis, AIM analysis, charge distribution).



INTRODUCTION

Halobismuthate(III) anions have been known for over 100 years. For example, dissolving bismuth trichloride in hydrochloric acid leads to the formation of the $[\text{BiCl}_4]^-$ monoanion and the $[\text{BiCl}_5]^{2-}$ dianion.^{1,2} Following these simple anions, a plethora of more complex oligomeric and polymeric halobismuthate anions of the general formula $[\text{Bi}_n\text{X}_m]^{3n-m}$ have been discovered.^{3–5} Due to the weakness of the bismuth–halogen bond, these complexes can undergo various association and dissociation processes in solution; however, in the solid state, the versatility of coordination modes around the bismuth center results in a high structural diversity. For example, halobismuthates can be discrete anions with varying nuclearity (zero dimension, 0D) or they can form one-dimensional (1D) polymeric chains, two-dimensional (2D) networks, or even three-dimensional (3D) architectures.^{5–18} Since the solid-state structure of inorganic/organic hybrids is crucial for applications and depends on several factors, such as the counterion, solvent, temperature, etc., the controlled design of these materials with bespoke properties is difficult, stimulating intense ongoing research in this field.

A common prototype of discrete anions are binuclear dibismuthates of the general formula $[\text{Bi}_2\text{X}_m]^{6-m}$, where m can range from 8 to 11.^{3,4} Among these, the rarest species known are for $[\text{Bi}_2\text{X}_8]^{2-}$, whose centrosymmetric geometry exhibits two edge-sharing square pyramids and a stereochemically active lone pair at each of the bismuth centers (see Figure 1).^{19–22} In contrast, in the other three anions ($m = 9, 10,$ or

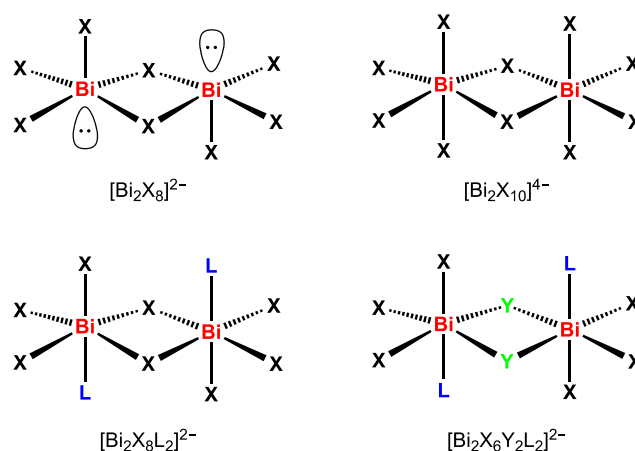


Figure 1. Structural diversity of binuclear bismuthate anions ($X = \text{Cl}, \text{Br}, \text{I}$; $L = \text{Lewis bases with O, N, or S donor atoms}$; $Y = \text{triflate etc.}$; for details see the text). Except for the $[\text{Bi}_2\text{X}_8]^{2-}$ anion the lone pairs at the bismuth centers are not shown, as they typically appear to be stereochemically inactive.

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11) the bismuth resides in an octahedral coordination environment. Similar to the $[\text{Bi}_2\text{X}_8]^{2-}$ anion, in the dibismuthate $[\text{Bi}_2\text{X}_{10}]^{4-}$ the two hexacoordinated bismuth centers are connected by two μ_2 -bridging halide ions, and the bridging Bi–X distances are longer than the terminal distances. Formally, the $[\text{Bi}_2\text{X}_{10}]^{4-}$ unit can be derived from the $[\text{Bi}_2\text{X}_8]^{2-}$ anion by coordinating two further X^- anions to the axial positions (occupied by lone pairs in $[\text{Bi}_2\text{X}_8]^{2-}$). Alternatively, other Lewis bases may play the role of electron-pair donors instead of these two halide ions; however, these $[\text{Bi}_2\text{X}_8\text{L}_2]^{2-}$ species are less common than the all-halide dibismuthates $[\text{Bi}_2\text{X}_{10}]^{4-}$. The lion's share of the $[\text{Bi}_2\text{X}_8\text{L}_2]^{2-}$ structures has been reported with solvent molecules (e.g., L = THF,^{20,23–26} H_2O ,²⁷ DMSO,^{26,28–32} acetone^{28,33,34}) coordinating to the bismuth centers (see Figure 1), but in some cases more sophisticated donor molecules such as bipyridines^{35–42} can act as Lewis bases. On the basis of our CCDC search, the most common donor atom is oxygen,^{20,23–34,38,42–44} followed by nitrogen,^{35–37,39–41} and to our knowledge there is only one example with sulfur (dimethyl sulfide⁴⁵).

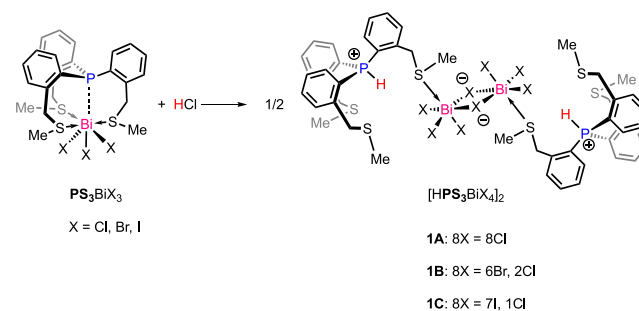
Furthermore, binuclear bismuthate anions offer the possibility of further functionalization. To our knowledge, there is only one example in which the bridging halides of $[\text{Bi}_2\text{X}_8\text{L}_2]^{2-}$ species have formally been replaced by another monoanion, resulting in a $[\text{Bi}_2\text{X}_6\text{Y}_2\text{L}_2]^{2-}$ type dibismuthate.⁴⁶ More commonly, the terminal halides can also be substituted by other anions (e.g., X = carboxylate, Y = nitrate;⁴⁷ X = Y = thiolates^{48,49}). Furthermore, the bridged $[\text{Bi}_2\text{X}_2]$ four-membered ring motif is also known in neutral dimers.^{50–55}

Here we report the reactivity of bismuth trihalide complexes with the formula PS_3BiX_3 ($\text{PS}_3 = \text{P}(\text{C}_6\text{H}_4\text{-}o\text{-CH}_2\text{SCH}_3)_3$; X = Cl, Br, I) and hydrochloric acid or silver triflate, delivering dimeric or polymeric dibismuthates of the type $[\text{Bi}_2\text{X}_8\text{L}_2]^{2-}$ or $[\text{Bi}_2\text{X}_6\text{Y}_2\text{L}_2]^{2-}$, respectively. In the starting materials PS_3BiX_3 of these reactions the bismuth center is stabilized by the cooperation of three dative S→Bi bonds and a further P··Bi pnictogen interaction, which altogether endow these complexes with significant stabilization energies of 25.9–28.3 kcal/mol.⁵⁶ In this paper we also aim to shed light on the energy requirements for the breakdown of these stabilizing interactions, leading to the formation of dibismuthate units.

RESULTS AND DISCUSSION

Synthesis and Characterization. When the adduct of the $\text{P}(\text{C}_6\text{H}_4\text{-}o\text{-CH}_2\text{SCH}_3)_3$ (PS_3) ligand and bismuth trichloride (BiCl_3) (in the following abbreviated as PS_3BiCl_3) dissolved in dichloromethane was reacted with an excess of HCl (4 M, 1,4-dioxane solution), an immediate color change of the solution from yellow to colorless was observed and a colorless, microcrystalline solid of the composition $[\text{HPS}_3\text{BiCl}_4]_2$ (**1A**, Scheme 1) was isolated in a good yield of 91% (for a single-crystal X-ray diffraction study see below). This inorganic/organic hybrid material was analyzed by multinuclear NMR spectroscopy, and the $^{31}\text{P}\{^1\text{H}\}$ NMR spectrum in CD_2Cl_2 exclusively shows a singlet resonance at -20.8 ppm, which is significantly different from that of PS_3BiCl_3 (-37.0 ppm in C_6D_6 ⁵⁶) or the free ligand PS_3 (-36.9 ppm in C_6D_6 ,⁵⁶ -36.2 ppm in CD_2Cl_2), indicating a change in the chemical environment of the phosphorus center. In the proton-coupled ^{31}P NMR spectrum this resonance appears as a doublet with a remarkably large $^1J(^{31}\text{P}\text{-}^1\text{H})$ coupling constant of 535 Hz. The downfield shift of the resonance in comparison to the free ligand, together with the magnitude of the coupling constant, is

Scheme 1. Reaction of PS_3BiX_3 (X = Cl, Br, I) with Hydrochloric Acid, Delivering the Inorganic/Organic Bis-Zwitterions **1A–C**

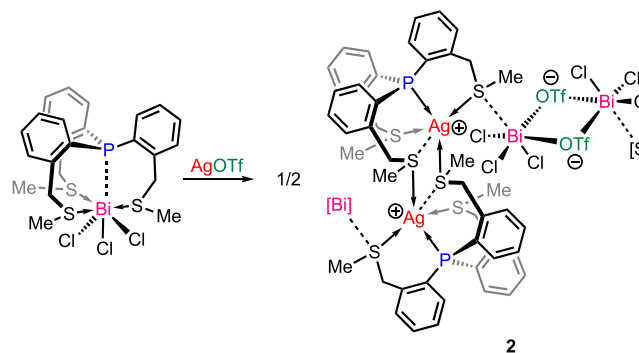


in line with the formation of a phosphonium cation (see $^1J(^{31}\text{P}\text{-}^1\text{H}) = 531$ Hz for the triphenylphosphonium cation⁵⁷ or $^1J(^{31}\text{P}\text{-}^1\text{H}) = 535$ Hz for $[\text{HPS}_3][\text{OTf}]$ ⁵⁸).

Analogously, the reaction was performed with the heavier analogues PS_3BiBr_3 and PS_3BiI_3 . Again, the ^{31}P NMR spectrum shows the formation of the phosphonium salts **1B** and **1C** ($\delta = -20.7$, $^1J(^{31}\text{P}\text{-}^1\text{H}) = 535$ Hz and $\delta = -21.8$ ppm, $^1J(^{31}\text{P}\text{-}^1\text{H}) = 535$ Hz, respectively). While the composition of the isolated **1B** can be described with six bromide and two chloride anions, the iodine content of **1C** is most likely somewhat larger than expected from the stoichiometry (for details see below and the Supporting Information). Presumably an exchange reaction of the halide substituents around the bismuth center takes place in solution, resulting in a significantly lower yield of **1C** (44%) in comparison to that of **1A** or **1B**.

To utilize the sulfur donor functionalities of the PS_3 ligand, we decided to test the reactivity of PS_3BiCl_3 toward a soft transition-metal center such as Ag^+ . Therefore, the adduct PS_3BiCl_3 was reacted with 1 equiv of silver trifluoromethanesulfonate ($\text{Ag}(\text{OTf})$) in dry acetonitrile (Scheme 2). The ^{31}P

Scheme 2. Reaction of PS_3BiCl_3 with Silver Triflate, Delivering Compound **2**^a



^a[Bi] and [S] denote the continuation of the polymeric chain.

NMR spectrum of the reaction mixture shows a broad singlet resonance at -33.1 ppm (without observable phosphorus–silver coupling) shifted downfield in comparison to the free ligand PS_3 ($\delta(^{31}\text{P}) = -38.0$ ppm in acetonitrile), indicating again a change in the chemical environment of the phosphorus center. Even though the coordination chemical shift change of $\Delta\delta(^{31}\text{P}) = \delta(\text{complex}) - \delta(\text{ligand}) = +3.8$ ppm is smaller than that for triphenylphosphine complexes $[(\text{PPh}_3)_n\text{Ag}(\text{OTf})]$ ($\Delta\delta(^{31}\text{P}) = 13.6\text{--}22.1$ ppm for $n = 1\text{--}4$ ^{59,60}), together with the broadening of the resonance this is consistent with an

interaction of the phosphorus and the silver centers. Similar behavior is widely known for silver complexes with neutral ligands and is attributed to the fluxionality of the species in solution. According to a single-crystal X-ray diffraction analysis, in the solid state this material can be described with the formula $\{[\text{AgPS}_3\text{BiCl}_3(\text{OTf})_2](\text{CH}_3\text{CN})_2\}_\infty$ (see below).

Optical Properties of Bis-Zwitterions 1A–C. As isolated solids, compounds **1A** and **1B** are white and yellowish white, respectively; however, **1C** is intense bright red and shows no photoluminescence in the solid state or in solution. In dichloromethane solution, all of these compounds exhibit broad absorption bands (see Figure 2), similarly to other

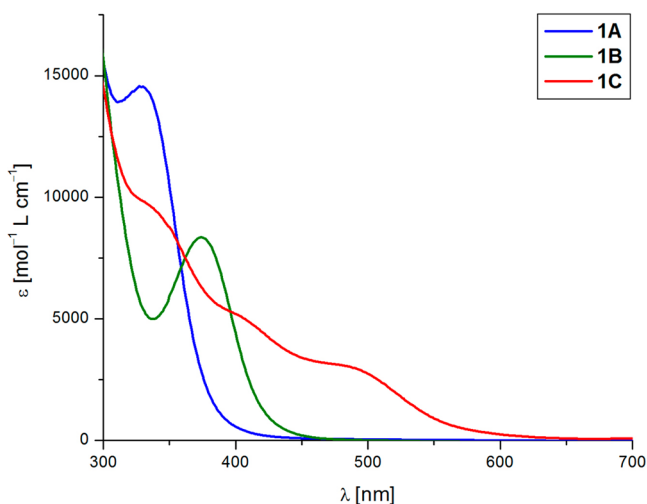


Figure 2. UV–vis spectra of compounds **1A–C** in dichloromethane ($c = 10^{-6}$ mol/L).

bismuth complexes,⁴² and a clear red shift can be observed in the direction **1A** → **1B** → **1C** ($\lambda_{\text{max}} = 328, 374,$ and 485 nm, respectively, for the lowest energy excitations). This tendency can be reproduced by time-dependent density functional theory calculations at the TD-B3LYP/def2-SVP level, which predict several closely lying transitions, matching the experimentally observed broad absorption bands. The calculated absorption bands are somewhat shifted toward lower wavelengths (297–330 nm (**1A**), 303–340 nm (**1B**), and 376–439 nm (**1C**)) in comparison to the experimental bands; nevertheless, the trend is properly described.

These low-energy excitations typically involve transitions from the dibismuthate core to the π^* system localized at the two phosphonium units. As an example, the HOMO and LUMO Kohn–Sham orbitals of **1A** are presented in Figure 3 (the LUMO+1, LUMO, HOMO, and HOMO-1 orbitals with their energies are collected in the Table S3 in the Supporting Information). The HOMO of **1A** is formed from the p-type lone pairs located at the halogen centers and the sulfur donor atoms of PS_3 with an antibonding combination of the bismuth s orbital, similarly to those reported for $[\text{BiX}_6]^{3-}$ ($X = \text{Cl}, \text{I}$).⁶¹ In contrast, the LUMO is localized at the two phosphonium units and shows π symmetry. The red shift observed in the direction of **1A**, **1B**, and **1C** can be explained on the basis of the HOMO–LUMO energies as follows. The LUMO energy of the π^* system located at the phosphonium units stays practically unchanged in these three compounds; however, the energy of the HOMO orbitals depends on the halogen substituent. As the orbital energies of the lone pairs at the

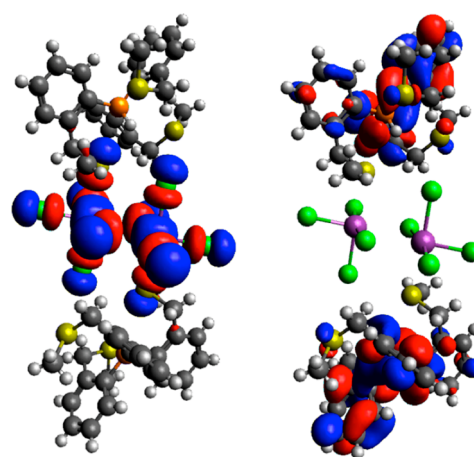


Figure 3. Kohn–Sham HOMO (left) and LUMO (right) of **1A** at the B3LYP/def2-SVP level with a contour value of 0.02.

halide centers increase in the order Cl, Br, I, the HOMO energies increase in the sequence **1A**, **1B**, **1C**, resulting in a decrease in the HOMO–LUMO gap (7.84, 7.67, and 6.61 eV, respectively) and therefore in a shifting of the absorption band to longer wavelengths.

Altogether, these results show that both inorganic and organic components play an important role in the optical properties of these bis-zwitterions. Furthermore, the UV–vis absorptions can be tuned by changing the halogen substituents of the dibismuthate cores. Alternatively, the modifications of the π^* type acceptor orbitals of the phosphonium system (e.g., via heterosubstitution in the π system or introduction of more extended aromatic units) could be employed to further tune the optical properties of such bis-zwitterionic assemblies.

Structural Investigations: X-ray Diffraction and DFT Calculations. The inorganic/organic hybrid bis-zwitterion **1A** crystallizes in two different polymorphic forms (**1A'** and **1A''**) with the same space group $P2_1/n$ and broadly similar crystal packing and lattice parameters (Table S1) but different molecular conformations. Form **1A'** is enantiotropic, undergoing a reversible, single-crystal to single-crystal phase transition between 220 and 200 K to the triclinic (space group $P\bar{1}$) form **1A'''**. Since the structure of **1A'''** shows close similarity to that of **1A'**, in the following only **1A'** and **1A'''** are discussed in detail, and the polymorphic form **1A'''** is presented in the Figure S13 in the Supporting Information. All of these polymorphs show discrete dimeric molecules without short contacts or significant van der Waals interactions between them.

Both dimeric **1A'** and **1A''** (Figure 4A,B, respectively) adopt a bis-zwitterionic structure consisting of two $[\text{HPS}_3]^+$ phosphonium units each coordinating with one thioether arm to a central dianionic $[\text{Bi}_2\text{Cl}_8]^{2-}$ motif, confirming that the chloride anions from the hydrogen chloride starting material are captured in the coordination spheres of the bismuth centers. The centrosymmetric core shows an edge-shared bioctahedral structure of the type $[\text{Bi}_2\text{X}_8\text{L}_2]^{2-}$ (cf. Figure 1). The two bismuth centers and six of the chloro ligands are nearly coplanar, while two further chlorines and the sulfur donors of two different PS_3 ligands occupy the axial positions. Again, in both polymorphic forms these sulfur atoms are mutually *trans* oriented and are *cis* relative to the bridging chlorides.

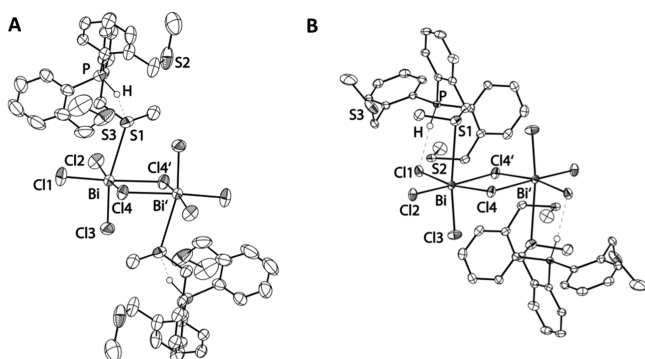


Figure 4. ORTEP representations and selected atomic distances (Å) and bond angles (deg): (A) **1A'**, Bi–S 3.1576(8), Bi–Cl1 2.5846(8), Bi–Cl2 2.5457(8), Bi–Cl3 2.5471(9), Bi–Cl4 2.9055(8), Bi–Cl4' 2.8192(7), C–P–C bond angles 108.63(15), 110.03(14), 112.17(13); (B) **1A''**, Bi–S1 3.2416(6), Bi–Cl1 2.6411(5), Bi–Cl2 2.5542(5), Bi–Cl3 2.5008(6), Bi–Cl4 2.7611(5), Bi–Cl4' 2.8957(5), C–P–C bond angles 107.23(8), 110.36(9), 111.90(9). Thermal ellipsoids are drawn at the 50% probability level. Hydrogen atoms (except for that at the phosphorus center) have been omitted for clarity. Bond valences around the bismuth centers (s , vu; for details see text): **1A'**, Bi–S 0.194, Bi–Cl1 0.754, Bi–Cl2 0.837, Bi–Cl3 0.834, Bi–Cl4 0.217, Bi–Cl4' 0.400, $\sum s = 3.335$; **1A''**, Bi–S 0.154, Bi–Cl1 0.647, Bi–Cl2 0.818, Bi–Cl3 0.945, Bi–Cl4 0.468, Bi–Cl4' 0.325, $\sum s = 3.358$.

The question of stereochemical activity or inactivity of the lone electron pair may arise in connection with bismuth(III) cations. In contrast to the lighter congeners P or Sb, the inertness (stereochemical activity) of the lone pair at Bi(III) centers is less pronounced and it is often difficult to recognize and verify, especially in hexacoordinated species.^{61,62} As the coordination environment around the bismuth centers in both **1A'** and **1A''** is distorted from the ideal octahedral environment, we inspected the bond angles more closely in order to find evidence of stereochemical activity. A significant deviation from the octahedral bond angles may be indicative of the existence of an inert lone pair, and the largest distortion from the ideal geometry was found for the Cl1–Bi–Cl3 angle in **1A'**: namely, 98.8°. On the basis of a comparison of our system to previously reported structures with stereochemically active^{63,64} or inactive^{65,66} bismuth lone pairs, we may consider that the bismuth lone pairs in this study have no appreciable stereochemical activity. The situation is further disturbed by the fact that two different kinds of atoms (S/Cl) coordinate to the bismuth centers, which clearly affects the bonding parameters (bond lengths and bond angles). According to Wheeler et al. the stereochemical activity of the lone pair at Bi is consistent with an unsymmetrical distortion of the HOMO;⁶¹ however, on the basis of Figure 3 the Bi center in **1A** is stereochemically inactive.

A difference between the two polymorphs shown in Figure 4 is that **1A''** exhibits a PH⋯Cl1 hydrogen bond between the phosphonium proton and one of the terminal chloro ligands of the dibismuthate core, while in **1A'** only a PH⋯S1 interaction can be found. The presence of the PH⋯Cl1 intramolecular hydrogen bond in **1A''** also affects the orientation of the [HPS₃]⁺ moiety with respect to the halobismuthate core.

For the metric parameters, in structure **1A'** the bridging Bi–Cl bond distances are, as expected, significantly longer than either the terminal or the axial distances, and the values match those observed for example in [Li(THF)₄]₂[Bi₂Cl₈].²³ The

Bi–S1 bond length of 3.1576(8) Å is slightly longer than that in the only reported octaiododibismuthate complex with SME₂ (3.054(8) Å).⁴⁵ In general, the structural parameters of **1A''** are similar to those of **1A'**. The only remarkable differences are related to the previously mentioned PH⋯Cl1 hydrogen bond in the former, which leads to a slight elongation of the Bi–Cl1 bond (2.6411(5) Å) in comparison to the other terminal Bi–Cl2 bond (2.5542(5) Å).

The bond valences (s) were calculated on the basis of the method by Brown⁶⁷ employing the data set reported by Brese and O'Keeffe.⁶⁸ In both structures **1A'** and **1A''**, the bond valences for the bridging Bi–Cl bonds ($s = 0.317$ – 0.468 valence units (vu)) are significantly smaller than those of the terminal bonds ($s = 0.647$ – 0.945 vu), outlining weaker covalent interactions for the former. The bond valences of the Bi–S interaction are $s = 0.194$ and 0.154 vu for **1A'** and **1A''**, respectively, and indicate a dative bond with non-negligible covalent character. The sums of bond valences around the bismuth center in **1A'** and **1A''** ($\sum s = 3.335$ and 3.358 vu, respectively) are similar to that of the starting material PS₃BiCl₃ ($\sum s = 3.233$ vu).⁵⁸

The gas-phase optimized structure of **1A** (at the ω B97XD/def2-SVP(PCM) level; see the Supporting Information) resembles that of **1A'** in the solid state. The Wiberg bond indices (WBI) follow the tendencies of the bond valences. The sum of NPA (natural population analysis) charges for the [Bi₂Cl₈] moiety is -2.13 e (instead of -2 e), which reveals electron donation from the sulfur atoms to the bismuth centers. The bis-zwitterionic charge distribution of **1A** is clearly visible on the molecular electrostatic potential map plotted on the van der Waals surface (Figure 5).

The mixed halide analogues **1B** and **1C** exhibit molecular structures similar to that of the polymorph **1A'**, and the ORTEP representations are shown in Figures S14 and S15, respectively, in the Supporting Information. The crystal structure of **1C** (monoclinic space group $P2_1/n$) is also similar

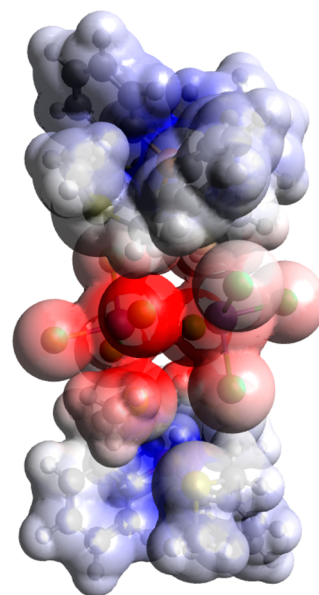


Figure 5. Molecular electrostatic potential map of **1A** at the ω B97XD/def2-SVP(PCM) level. The red and blue areas indicate attractive and repulsive electrostatic potentials toward a positive point charge, respectively.

to that of **1A'**, while the triclinic structures of **1B** (space group $P\bar{1}$) are analogous to that of **1A'''**: i.e., distorted **1A'**. The occupancies were refined independently for each halogen site, and the different halides show no remarkable preference at each of the positions. While for **1B** the bromine to chlorine ratio is near to that expected (5.94:2.06), in **1C** the I:Cl ratio (7.19:0.81) deviates significantly from 6:2. In both **1B** and **1C** no evidence of stereochemically active lone pairs at the Bi(III) centers is visible.

To gain information on the relative energy of geometrical isomers of the mixed halides with the formula $[\text{HPS}_3\text{BiX}_3\text{Cl}]_2$ (with both $\text{X} = \text{Br}$ and $\text{X} = \text{I}$), the structures of 10 possible geometrical (*cis/trans*) isomers were generated by placing one chlorine atom at each bismuth center at different positions (bridging, axial, or any equatorial) and the optimized results at the $\omega\text{B97XD}/\text{def2-SVP}(\text{PCM})$ level are collected in the [Supporting Information](#). In the case of the bromine analogue all 10 isomers have rather similar relative energies and the maximum difference is only 3.1 kcal/mol (see [Table S4](#)), which agrees with the observed equal distribution of the different halides in the solid-state structures. In contrast, the relative energy difference among the 10 geometrical isomers is somewhat larger (6.9 kcal/mol) in the case of the iodine analogue. On the basis of the relative energies, the chlorine atoms disfavor the bridging positions and prefer the axial or equatorial positions (for details see [Table S5](#)).

The adduct of PS_3BiCl_3 and $\text{Ag}(\text{OTf})$ forms a one-dimensional polymeric chain with the formula $\{[\text{AgPS}_3\text{BiCl}_3(\text{OTf})]_2(\text{CH}_3\text{CN})_2\}_\infty$ ([Figure 6A](#)). Between

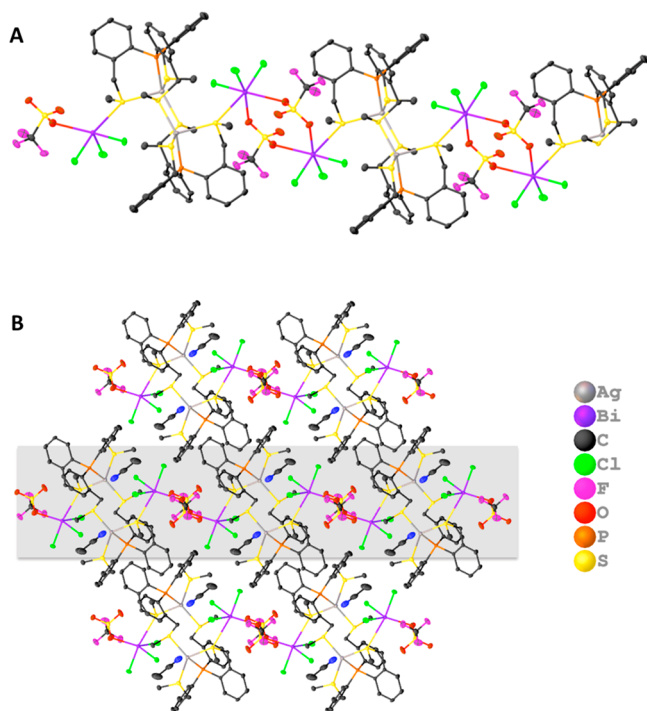


Figure 6. (A) ORTEP representation showing the ribbonlike polymeric chain of $\{[\text{AgPS}_3\text{BiCl}_3(\text{OTf})]_2(\text{CH}_3\text{CN})_2\}_\infty$ composed of dibismuthate units $[\text{Bi}_2\text{Cl}_6]^{2-}$ and cationic $[\text{Ag}(\text{PS}_3)]_2^{2+}$. Hydrogen atoms and solvent molecules have been omitted for clarity. (B) Packing representation of $\{[\text{AgPS}_3\text{BiCl}_3(\text{OTf})]_2(\text{CH}_3\text{CN})_2\}_\infty$ along the a axis. Hydrogen atoms are not shown. The gray area represents the polymeric chain shown in A. Thermal ellipsoids are drawn at the 50% probability level in both images.

the polymeric chains no significant van der Waals interactions can be found. Along the a axis a void contains acetonitrile solvent molecules ([Figure 6B](#)). In this polymer, dibismuthate building blocks with the formula $[\text{Bi}_2\text{X}_6\text{Y}_2\text{L}_2]^{2-}$ ([Figure 1](#)) can be found. While **1A–C** contain discrete bis-zwitterions, the silver/bismuth heterometallic polymer is built up by an infinite chain of dianionic and dicationic repeating units (triflate-bridged dibismuthates and $[\text{Ag}(\text{PS}_3)]_2^{2+}$ macrocycles, respectively). In contrast to the more common polymeric bismuthate ions outlined in the [Introduction](#),^{3,5} in the present case discrete dibismuthate anions are the repeating units connecting dicationic macrocycles via organic linkers.

The question may arise as to how strong is the interaction which holds together the assembly of the polymeric chain. The heterolytic bond dissociation energy of an $\text{S} \rightarrow \text{Bi}$ dative bond amounts to only 14.3 kcal/mol (as calculated for $\text{Me}_2\text{S} \rightarrow \text{BiCl}_3$ at the $\omega\text{B97XD}/\text{def2-SVP}$ level), which implies that the formation of such bonds can be reversible. We have estimated the (BSSE corrected) total interaction energy associated with the assembly of two neutral $[(\text{AgPS}_3)_2\text{Bi}_2\text{Cl}_6(\text{OTf})_2]$ units (as depicted in [Scheme 2](#)) via an $\text{S} \rightarrow \text{Bi}$ bond (to form $[(\text{AgPS}_3)_2\text{Bi}_2\text{Cl}_6(\text{OTf})_2]_2$), which is as large as 46.6 kcal/mol ($\omega\text{B97XD}/\text{def2-SVP}$, employing solid-state structures as references). This shows that, in addition to the dative bond, significant electrostatic attraction arises between the partial charges of the fragments. Nevertheless, due to the heterolytic dissociation of the dative $\text{S} \rightarrow \text{Bi}$ bond, the solution of compound **2** presumably contains $[\text{AgPS}_3\text{BiCl}_3(\text{OTf})]$ or its dimeric form rather than a polymer, as suggested by DOSY experiments (for details see the [Supporting Information](#)).

In the centrosymmetric $[\text{Bi}_2\text{Cl}_6(\text{OTf})_2]^{2-}$ core ([Figure 7A](#)), the triflate anions occupy the bridging positions and, similarly to **1A–C**, the sulfur donor atoms in the octahedral coordination sphere of the bismuth are mutually *trans* oriented. The Bi–Cl bond lengths in the dibismuthate $[\text{Bi}_2\text{Cl}_6(\text{OTf})_2]^{2-}$

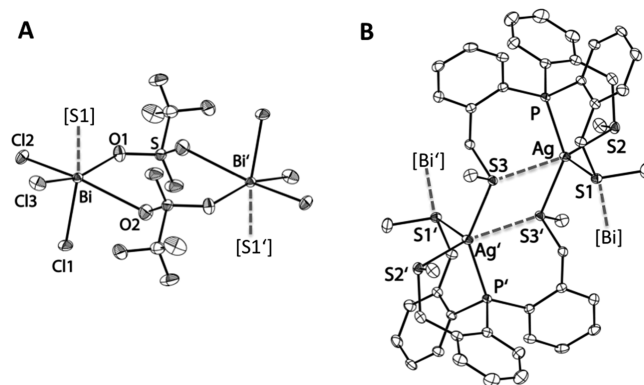


Figure 7. ORTEP representations of $\{[\text{AgPS}_3\text{BiCl}_3(\text{OTf})]_2(\text{CH}_3\text{CN})_2\}_\infty$ and selected atomic distances (Å) and bond angles (deg): (A) $[\text{Bi}_2\text{Cl}_6(\text{OTf})_2]^{2-}$, Bi–O1 2.7564(17), Bi–O2 2.776, Bi–S1 3.2304(5), Bi–Cl1 2.4918(6), Bi–Cl2 2.4808(6), Bi–Cl3 2.4949(5); (B) $[\text{Ag}(\text{PS}_3)]_2^{2+}$, Ag–P 2.4658(5), Ag–S1 2.6823(5), Ag–S2 2.5854(5), Ag–S3' 2.5050(5), Ag–S3 3.418. Thermal ellipsoids are drawn at the 50% probability level. Hydrogen atoms and solvent molecules have been omitted for clarity. Bond valences (s , vu ; for details see the text): around the bismuth centers, Bi–S1 0.159, Bi–Cl1 0.969, Bi–Cl2 0.998, Bi–Cl3 0.961, Bi–O1 0.165, Bi–O2 0.157, $\sum s = 3.408$; around the silver centers, Ag–P 0.515, Ag–S1 0.237, Ag–S2 0.308, Ag–S3' 0.383, Ag–S3 0.032, $\sum s = 1.476$.

are shorter than those observed in the polymorphs of **1A** (*vide supra*), and the larger bond valences also show a strengthening of these bonds (see legend of Figure 7). The Bi–O distances (on average 2.77 Å) are significantly shorter than the sum of the van der Waals radii of Bi and O (3.59 Å)⁶⁹ and are in the range of those observed in other Bi-triflate structures.^{46,58} Nevertheless, the bond valences of the Bi–O interactions (on average $s = 0.16$ vu) only indicate weak coordination of the triflate anions, enabling the three terminal chlorides to build up a stronger coordination than in dibismuthate **1A**. Hence, the total occupancy of the coordination sphere around the bismuth center in $[\text{Bi}_2\text{Cl}_6(\text{OTf})_2]^{2-}$ is similar to that in **1A**, which is also reflected in the similar sum of bond valences ($\sum s = 3.408$ vu in comparison to $\sum s = 3.335$ and 3.358 vu in **1A'** and **1A''**, respectively). In order to clarify the degree of stereochemical activity of the lone pair at the bismuth center, again the bond angles were examined around the central atom and compared with literature examples (*vide supra*). Significantly, the O1–Bi–S bond angle (108.5°) shows a substantial deviation from the ideal 90° . Therefore, in contrast to the bis-zwitterionic dimers **1A–C** discussed above, the lone pair at the bismuth center in this silver complex exhibits a considerable stereochemical activity, pointing between the O1 and S atoms.

The tetradentate PS_3 ligand coordinates in a scorpionate fashion with three dative interactions ($\text{P} \rightarrow \text{Ag}$, $\text{S1} \rightarrow \text{Ag}$, and $\text{S2} \rightarrow \text{Ag}$) and a weak $\text{S3} \cdots \text{Ag}$ interaction to the same silver center, while the last sulfur center is also involved in a bridging $\text{S3} \rightarrow \text{Ag}'$ bond to the second silver cation. The Ag–P bond length (2.4658(5) Å) is in the expected range for silver complexes with phosphine ligands.^{70,71} The $\text{S3} \cdots \text{Ag}$ atom distance (3.418 Å) is clearly shorter than the sum of van der Waals radii of these elements (3.8 Å),⁷² and these weak interactions act as cross-connections to stabilize the 12-membered $\text{Ag}_2\text{P}_2\text{S}_2\text{C}_6$ “macrocycle”. The dative $\text{S} \rightarrow \text{Ag}$ bond lengths are similar to previously reported bond lengths. Among these, the $\text{S1} \rightarrow \text{Ag}$ bond is the longest, because the S1 center binds in a bridging μ_2 fashion between the silver and the adjacent bismuth center, linking the dianionic and dicationic parts (Figure 7B). The bond valences show that the interaction of the silver with the phosphorus center ($s = 0.515$ vu) is stronger than that with the sulfur atoms ($s = 0.032$ – 0.383 vu).

The gas-phase optimized structure of the dicationic $[\text{Ag}(\text{PS}_3)_2]^{2+}$ ($\omega\text{B97XD}/\text{def2-SVP}(\text{PCM} = \text{CH}_2\text{Cl}_2)$; for details see Table S6) is very similar to that observed in the solid state. An atoms in molecules (AIM) analysis of the $[\text{Ag}(\text{PS}_3)_2]^{2+}$ dication (see Figure 8) located five bond critical points around each of the silver centers. The electron density at the bond critical points shows the same tendency as that observed for the bond valences. Importantly, the sign of the total electronic energy density^{73,74} at the bond critical point (H) reveals a difference between the four dative bonds and the fifth weaker $\text{S} \cdots \text{Ag}$ interaction (shown with blue arrows in Figure 8). In the case of the former dative bonds $H < 0$ indicates high covalent character (clearly a dative interaction), while for the latter $H > 0$ suggests a noncovalent van der Waals interaction. Therefore, the fifth, weakest $\text{S} \cdots \text{Ag}$ interaction is best described as a van der Waals interaction rather than a dative bond. Note that recently the concept of σ -hole interactions has also been extended to coinage metals (coinage-metal bond or regium bond)^{75–77} and group 12 metals (spodium bonds).⁷⁸

Altogether, the silver center in the centrosymmetric dicationic $[\text{Ag}(\text{PS}_3)_2]^{2+}$ macrocycle (Figure 7B) is stabilized

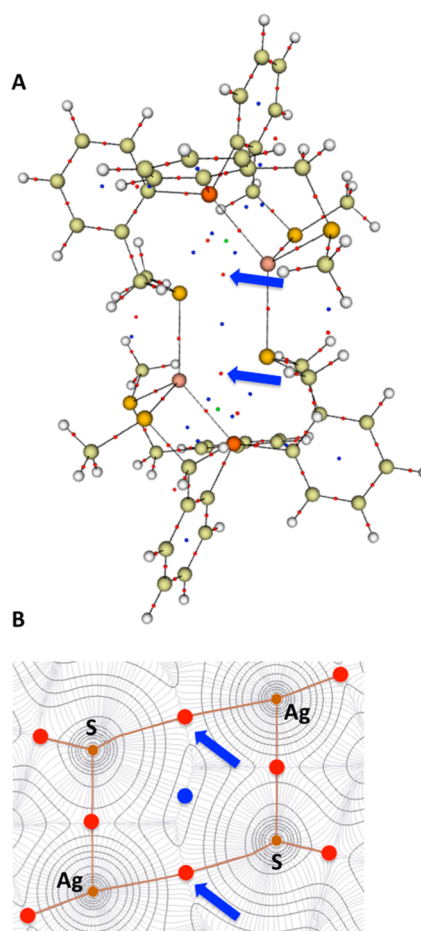


Figure 8. Plots of the results obtained by AIM analysis of the dimer $[\text{Ag}(\text{PS}_3)_2]^{2+}$ at the $\omega\text{B97XD}/\text{def2-SVP}(\text{PCM} = \text{CH}_2\text{Cl}_2)$ level of theory. Red, blue, and green dots indicate bond, ring, and cage critical points, respectively. The blue arrows show the weak $\text{S} \cdots \text{Ag}$ interactions. (A) Three-dimensional representation of bond, ring, and cage critical points. (B) Contour plot of the electron density in the Ag_2S_2 plane.

by coordination of one phosphorus and three sulfur donors as well as a further van der Waals $\text{S} \cdots \text{Ag}$ interaction. On the basis of the four dative bonds the silver centers could be considered as 4-coordinated. On the other hand, the number of bond critical points representing the interactions suggests a coordination number of 5. In analogy with the limited number of reported examples with 5-coordinated silver centers,^{79–88} we therefore describe the silver centers in the $[\text{Ag}(\text{PS}_3)_2]^{2+}$ units as 5-coordinated. The structural parameter value of $\tau_5 = 0.47$ indicates that the coordination around the silver atom is intermediate between trigonal bipyramidal and square pyramidal.^{89,90}

³¹P NMR Experiments and Thermodynamic Considerations. As described above, the reaction of PS_3BiX_3 ($X = \text{Cl}, \text{Br}, \text{I}$) with HCl delivers the bis-zwitterionic compounds **1A–C**. This reaction involves the heterolytic splitting of HCl, accompanied by the coordination of the Cl^- anion to the Bi center and the protonation of the phosphorus. The latter is clearly visible in the ³¹P NMR spectrum, which contains a doublet resonance at -20.8 ppm with a $^1J(^{31}\text{P}-^1\text{H})$ coupling constant of 535 Hz. In general, phosphines are known to be weak Brønsted bases in organic solvents: e.g., for triphenyl-

phosphine $pK_a(\text{HPPH}_3^+) = 7.64$ in acetonitrile⁹¹ and has an estimated value in THF of 3.⁹²

The protonation of phosphines with hydrogen halides HX to form the triorganylphosphonium halide $[\text{HPR}_3]^+\text{X}^-$ strongly depends on the basicity of the phosphines and the strength of the acids. According to the solid-state studies, the triethylphosphonium salt $[\text{HPEt}_3]^+\text{X}^-$ exists with X = Cl, Br, I; however, the triphenyl analogue $[\text{HPPH}_3]^+\text{X}^-$ is only known with X = Br, I counterions. This is due to the lower basicity of triphenylphosphine in comparison to triethylphosphine⁹³ and the weaker acidity of HCl in comparison to HBr and HI. Sheldon and Tyree reported the reaction product of triphenylphosphine and HCl with the formula $[\text{Ph}_3\text{PHCl}]_3 \cdot \text{HCl}$.⁹⁴ Later on, however, van den Akker and Jellinek showed that this adduct arose from the partial decomposition of the labile hydrogendichloride salt $[\text{HPPH}_3]^+[\text{HCl}_2]^-$ (which readily decomposed finally to PPh_3 and 2 HCl).⁹³ Clearly, a second molecule of HCl is necessary to capture the chloride anion in the form of hydrogendichloride $[\text{HCl}_2]^-$ to facilitate the protonation. Alternatively, strong Lewis acids such as Sn(IV), Fe(III), and Mo(III) can also be employed as chloride scavengers to form ionic phosphonium salts in the solid state.^{94,95}

In contrast to the observations for the solid state, in solution, the reaction of phosphines with hydrogen halides typically results in an equilibrium: $\text{R}_3\text{P} + \text{HX} \rightleftharpoons [\text{R}_3\text{PH}]^+ + \text{X}^-$. For example, even HBr, which is a stronger acid in comparison to HCl, can only protonate triphenylphosphine reversibly, resulting in a broad resonance without resolved $^1J(^{31}\text{P}-^1\text{H})$ coupling in the ^{31}P NMR spectra.⁹⁶ Furthermore, the corresponding chemical shift represents a superposition between those of the protonated and nonprotonated phosphine. This reversibility is due to the competing affinities of the phosphine and the rather basic halide anion toward the same H^+ . The protonation of phosphines with HCl in organic solvents requires special conditions, and to our knowledge only two studies have reported on this. One possibility is to activate HCl via capturing the chloride ion: for example, by a hydrogen-bonding network (to the thiourea moiety $\text{R}-\text{NH}-\text{C}(\text{S})-\text{NH}-\text{R}'$).⁹⁷ Alternatively, the proton of HCl can be encapsulated into a rigid and sterically congested cage at the P center with further stabilization of hydrogen bonding to another phosphorus (as reported for an *in, in* diphosphine).⁹⁸ In light of these studies, the selective formation of compounds **IA**–**C** prompted us to investigate these reactions further.

To gain more insights as to why the protonation shown in Scheme 1 does not lead to an equilibrium, we performed NMR experiments in which the roles of the ligand and the bismuth center were separately considered. We also performed DFT calculations employing the ωB97XD functional with different basis sets to further bolster the experimental studies. In the following we only discuss the results obtained at the $\omega\text{B97XD}/(\text{aug-cc-pVDZ}(-\text{PP}))$ level, which has been used successfully for similar systems before.⁵⁶ The solvent effects were simulated using the polarizable continuum model (PCM) with dichloromethane as solvent.

First, we checked whether the PS_3 ligand is basic enough to be protonated with HCl in the absence of BiCl_3 . The reaction of the free ligand PS_3 with an excess of HCl yields a broad singlet peak at -32.0 ppm in the ^{31}P NMR spectrum of the CD_2Cl_2 solution. This chemical shift is situated between those of the protonated form **IA** (-20.8 ppm in CD_2Cl_2) and the (nonprotonated) free ligand PS_3 (-36.2 ppm in CD_2Cl_2), and

the unresolved broad peak indicates an equilibrium (Figure 9a–c). On the basis of the gas-phase proton affinities, the

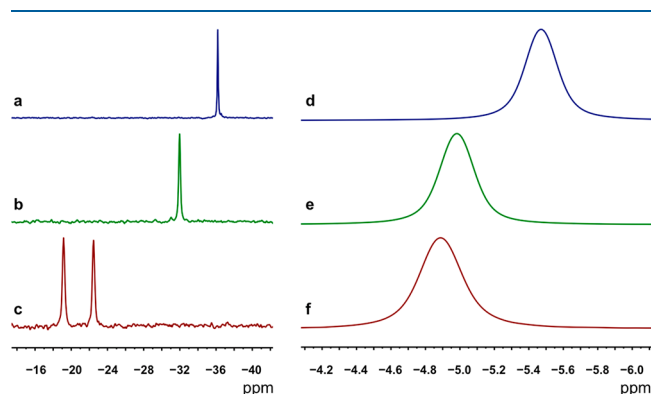


Figure 9. Solution ^{31}P NMR spectra of (a) PS_3 , (b) $\text{PS}_3 + \text{HCl}$, (c) **IA**, (d) PPh_3 , (e) $\text{PPh}_3 + \text{HCl}$, and (f) $\text{PPh}_3 + \text{HCl} + \text{BiCl}_3$: (a–c) in CD_2Cl_2 ; (d, e) in dioxane; (f) in a dioxane/ CH_2Cl_2 (1/1) mixture.

ligand PS_3 (-252.2 kcal/mol) is somewhat more basic than *ortho*-tolylphosphine (-244.5 kcal/mol) or triphenylphosphine (-241.1 kcal/mol). However, this slightly higher basicity of PS_3 is apparently not enough to compete with the chloride anion for the proton (see above).

Second, we aimed to clarify whether the chloride affinity (Lewis acidity) of BiCl_3 facilitates the protonation of a phosphine with hydrogen chloride in the form of $[\text{HPR}_3]^+[\text{BiCl}_4]^-$. Therefore, solutions of triphenylphosphine and HCl were combined, and in a second step, BiCl_3 was added to this solution. In both cases, only broad singlet peaks shifted slightly downfield in comparison to triphenylphosphine were observed and no spin–spin $^1J(^{31}\text{P}-^1\text{H})$ coupling was detected (Figure 9d–f). This indicates that BiCl_3 is not effective enough to capture the Cl^- from HCl, which is in line with the lability of the $[\text{BiCl}_4]^-$ anion (the stability constant of the reaction $\text{BiCl}_3 + \text{Cl}^- \rightleftharpoons \text{BiCl}_4^-$ is only $K = 2.7 \pm 1$).² We have estimated the chloride anion affinity of bismuth trichloride by computing the following reaction $\text{BiCl}_3 + \text{Cl}^- \rightleftharpoons \text{BiCl}_4^-$, and the Gibbs free energy of this reaction (-10.4 kcal/mol) indicates that BiCl_4^- seems to be more stable according to the gas-phase calculations (including solvent effects) than in the solution experiments. Nevertheless, the energetic consequences of this complexation reaction are minor in comparison to those of the protonation of the phosphine. Furthermore, our findings demonstrate the moderate Lewis acidity of BiCl_3 in contrast to highly Lewis acidic bismuthenium cations or neutral bismuth triamides, as outlined by several recent experimental and computational investigations.^{99–104}

We may conclude from the ^{31}P NMR test experiments that the protonation with HCl cannot be achieved alone by either the PS_3 ligand or BiCl_3 , and the computations suggest that the energetic effect of the protonation at the P center is more substantial than that of the Cl^- coordination to the BiCl_3 unit. So far, however, we have not considered the possible stabilizing effect of the dimerization leading to the dibismuthate core. Therefore, we computationally investigated the energetics of the reaction between PS_3BiCl_3 and HCl resulting in the bis-zwitterionic compound **IA** ($\omega\text{B97XD}/(\text{aug-cc-pVDZ}(-\text{PP}))$ -PCM = CH_2Cl_2) level). As this reaction is associative in nature, in addition to the reaction energies we include the Gibbs free energies as well. Since the entropy in the gas-phase

calculations is remarkably different from that expected in solution, the computed reaction Gibbs free energies only give an upper limit and we primarily focus on the reaction energies.

Presumably, in the first step of the reaction the monomeric form of dimer **1A** arises from PS_3BiCl_3 and an HCl molecule, which then dimerizes to form the final product. We attempted to optimize the structure of this monomeric unit from different starting geometries, and these optimizations resulted in the contact ion pair $[\text{HPS}_3]^+[\text{BiCl}_4]^-$ without Bi–S close contacts. The reaction energy and Gibbs free energy of the reaction $\text{PS}_3\text{BiCl}_3 + \text{HCl} \rightarrow [\text{HPS}_3]^+[\text{BiCl}_4]^-$ are $\Delta E = -5.8$ kcal/mol and $\Delta G^\circ_{298\text{K}} = +2.4$ kcal/mol, respectively, indicating no substantial thermodynamic stabilization. In contrast, the dimerization reaction $2[\text{HPS}_3]^+[\text{BiCl}_4]^- \rightarrow \mathbf{1A}$ is exothermic by $\Delta E = -34.2$ kcal/mol ($\Delta G^\circ_{298\text{K}} = -11.4$ kcal/mol). This shows that the dimerization is an important stabilizing factor in the formation of **1A** and likely this compound exists in its dimeric, bis-zwitterionic form in solution. The complete reaction of $2\text{PS}_3\text{BiCl}_3 + 2\text{HCl} \rightarrow \mathbf{1A}$ is rather exothermic ($\Delta E = -45.8$ kcal/mol and $\Delta G^\circ_{298\text{K}} = -6.6$ kcal/mol), which explains why this is not an equilibrium reaction. Furthermore, the importance of the PS_3 ligand in this reaction is highlighted by the following hypothetical exchange reaction: $[\text{HPPH}_3]_2[\text{Bi}_2\text{Cl}_8] + 2\text{PS}_3 \rightarrow \mathbf{1A} + 2\text{PPH}_3$, which indicates a remarkable stabilization in **1A** ($\Delta E = -21.0$ kcal/mol, $\Delta G^\circ_{298\text{K}} = -10.0$ kcal/mol).

CONCLUSION

We have shown that dibismuthates can serve as bridging units for constructing bis-zwitterionic inorganic/organic hybrid assemblies as well as coordination polymers. While the PS_3 ligand coordinates in a monodentate fashion in the 0-dimensional bis-zwitterions, in the 1D coordination polymer $\{[\text{AgPS}_3\text{BiCl}_3(\text{OTf})_2](\text{CH}_3\text{CN})_2\}_\infty$ it binds with all four donor atoms. Among these, two sulfur donors act as μ_2 bridges: one of them cross-links the 12-membered $\text{Ag}_2\text{P}_2\text{S}_2\text{C}_6$ “macrocycle”, while the other connects the 5-coordinated silver and 6-coordinated bismuth centers. Altogether, the phosphorus and the three sulfur donor atoms endow the PS_3 ligand with versatile coordination ability, complementing the previously reported tridentate and tetradentate coordination modes.⁵⁶ The UV–vis absorption bands connected to excitations from the dibismuthate core into the π^* system of the phosphonium units can be tuned by varying the halogen substitutions at the bismuth centers in the bis-zwitterions.

Despite the remarkable differences between hydrogen chloride and silver triflate, their reactivities toward the PS_3BiCl_3 adduct show several analogous features. (1) The connections between the ligand and the bismuth center are cleaved except for one Bi–S bond, which acts as a linker between the cationic and anionic fragments in the final products. (2) The phosphorus center is attacked by the electrophiles H^+ or Ag^+ replacing the bismuth center. (3) The counteranions Cl^- and OTf^- are scavenged in the coordination sphere of the bismuth to form dibismuthate linkers.

Triarylphosphines are typically rather weak bases, and their reaction with hydrogen chloride in organic solvents leads to an equilibrium reaction. This can be explained by the competition of the phosphine and the halide anion for the proton. Remarkably, in this case the reaction does not lead to an equilibrium and the energy requirements of this process were studied by DFT calculations, which highlighted that the formation of the dimeric dibismuthate plays a major role.

To the best of our knowledge, herein we report the first coordination polymer incorporating dibismuthate units and organometallic linkers. As our study shows the potential of dibismuthates as anionic building blocks for versatile structural motifs, in the future we intend to investigate the structure and properties of further derivatives (especially with main-group and transition metals).

ASSOCIATED CONTENT

Supporting Information

The Supporting Information is available free of charge at <https://pubs.acs.org/doi/10.1021/acs.inorgchem.0c01619>.

Description of experimental procedures, characterization of compounds, X-ray crystallographic studies, and computational details (PDF)

Accession Codes

CCDC 1998214–1998218 and 1998469 contain the supplementary crystallographic data for this paper. These data can be obtained free of charge via www.ccdc.cam.ac.uk/data_request/cif, or by emailing data_request@ccdc.cam.ac.uk, or by contacting The Cambridge Crystallographic Data Centre, 12 Union Road, Cambridge CB2 1EZ, UK; fax: +44 1223 336033.

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Notes

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DEDICATION

Dedicated to Prof. Péter Huszthy on the occasion of his 70th birthday.

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