


 Cite this: *RSC Adv.*, 2021, 11, 3539

 Received 22nd December 2020  
 Accepted 11th January 2021

DOI: 10.1039/d0ra10739e

[rsc.li/rsc-advances](http://rsc.li/rsc-advances)

## Visible light-driven Giese reaction with alkyl tosylates catalysed by nucleophilic cobalt†

 Kimihiro Komeyama,<sup>ID</sup>\* Takuya Michiyuki, Yoshikazu Teshima and Itaru Osaka<sup>ID</sup>

The scope of the Giese reaction is expanded using readily available alkyl tosylates as substrates and nucleophilic cobalt(i) catalysts under visible-light irradiation. The reaction proceeds preferentially with less bulky primary alkyl tosylates. This unique reactivity enables the regio-selective Giese reaction of polyol derivatives.

Organic transformations involving alkyl radicals are powerful tools for the construction of C(sp<sup>3</sup>)-containing carbon–carbon bonds.<sup>1</sup> Particularly, the conjugate radical addition to electron-deficient olefins (Giese reaction) is a versatile protocol for C(sp<sup>3</sup>)-C(sp<sup>3</sup>) bond formation in the synthesis of natural products and pharmaceuticals<sup>2</sup> because it enables the introduction of alkyl moieties in a site-selective manner.

Typically, the Giese reaction involves the generation of alkyl radicals from organohalides using stannyl radicals,<sup>3</sup> silyl radicals,<sup>4</sup> and transition-metal catalysts.<sup>5</sup> More recently, oxidative protocols have been extensively developed to generate alkyl radicals from alkyl metallic reagents.<sup>6</sup> However, these approaches are disadvantageous; halogenated waste is generated, and the preparation of most of these alkyl radical sources requires several steps from commercially available chemicals, resulting in a multi-step process to construct the desired carbon–carbon bonds. Therefore, the development of Giese reactions with naturally abundant carbon sources such as alkyl carboxylic acids<sup>7</sup> and alkyl amines<sup>8</sup> is receiving increased attention.

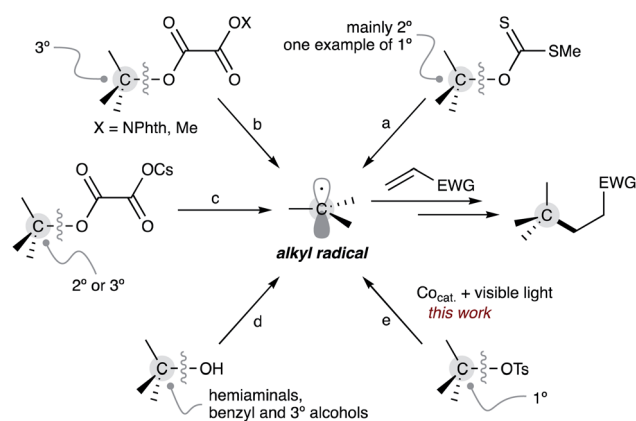
Alkanols are also among the most critical carbon sources because they are abundant in bioactive molecules and natural products in the form of sugars, steroids and among others. Furthermore, the hydroxyl substituent plays a vital role as a directing group for C–H bond functionalisations.<sup>9</sup> Therefore, the development of the Giese reaction from alkanols is of great importance in modern synthetic organic chemistry (Scheme 1).

In 2000, Togo reported the first example of a Giese reaction using alkyl xanthates as alkyl radical equivalents (Scheme 1a).<sup>10</sup> Overman applied a reductive approach for alkyl radical generation from alkyl *N*-phthalimidoyl oxalates by a photocatalyst to the Giese reaction (Scheme 1b).<sup>11</sup> Similarly, Gong demonstrated

the reductive strategy of alkyl methyl oxalates as alternative alkyl radical sources according to a unique reaction design.<sup>12</sup> Moreover, Overman and MacMillan found a radical addition using alkyl cesium oxalates as the source of alkyl radicals (Scheme 1c).<sup>13</sup> Recently, a direct approach from alkanols to alkyl radicals using a low-valent titanocene catalyst was also demonstrated (Scheme 1d).<sup>14</sup>

These protocols are pioneering works that allowed using abundant alkanols and their derivatives as alkyl radical sources for the Giese reaction. However, since these approaches involve the homolytic cleavage of the robust alcoholic C–O bond, the efficiency of the alkyl radical formation greatly depends on the thermodynamic stability of the generated alkyl radicals. Therefore, these Giese reactions are typically limited to tertiary, secondary, and benzylic alkyl radicals as well as heteroatom-linked carbon radicals generated from hemiaminals. Access to non-stabilised primary alkyl radicals from primary alcohols and their derivatives is more challenging.

We recently demonstrated various nickel and nucleophilic cobalt(i)-catalysed transformations of alkyl tosylates, including C(sp<sup>3</sup>)-C(sp<sup>2</sup>)<sup>15</sup> and C(sp<sup>3</sup>)-C(sp<sup>3</sup>) couplings<sup>16</sup> and amidation.<sup>17</sup> These reactions start from an S<sub>N</sub>2-type oxidative addition of,



Scheme 1 Examples of alcohol-based Giese reaction.

Department of Applied Chemistry, Graduate School of Advanced Science and Engineering, Hiroshima University, 1-4-1 Kagamiyama, Higashi-Hiroshima City, Hiroshima 739-8527, Japan. E-mail: [kkome@hiroshima-u.ac.jp](mailto:kkome@hiroshima-u.ac.jp)

† Electronic supplementary information (ESI) available: Additional experimental results and <sup>1</sup>H and <sup>13</sup>C NMR spectra data of all products. See DOI: 10.1039/d0ra10739e



especially, primary alkyl tosylates to a nucleophilic cobalt(II) centre, producing an alkyl-cobalt(III) species that enables a transalkylation with a nickel to form a reactive alkyl nickel intermediate. On the other hand, the resulting alkyl-cobalt(III) generates an alkyl radical through homolytic Co–C bond cleavage by visible-light irradiation. Numerous radical reactions that involve the catalytic photo-cleavage of a Co–C bond has been developed;<sup>18</sup> in contrast, the Giese reaction using unactivated alkyl tosylates as alkyl radical sources has never been reported so far. Herein, we report a Giese reaction that uses alkyl tosylates as alkyl radical sources, benefiting from the generation of non-stabilised alkyl radicals by photo-cleavage of the Co–C bond on the corresponding alkyl-cobalt(III) species (Scheme 1e).

Our investigation started by optimising the reaction of alkyl tosylate **1a** with activated olefin **2a** (Table 1). When **1a** was treated with **2a** (1.0 equiv.) in the presence of vitamin B<sub>12</sub> (VB<sub>12</sub>, 5 mol%) as a catalyst, Mn (3.0 equiv.) as a reductant, and Et<sub>3</sub>N·HCl (1.5 equiv.) as a proton donor under blue-light irradiation ( $\lambda = 454$  nm, 40 W), the expected radical adduct **3a** was obtained in 61% yield (Table 1, entry 1). In the reaction, the presence and type of proton donors greatly affected the reaction efficiency (entries 2–4). The choice of the planar-cobalt catalysts was also crucial. Thus, CoCl(dmgH)<sub>2</sub>Py, Co(salen), and Co(Pc) catalysts gave the product in negligible yields (entries 5–7),

whereas Co(TMPP) afforded the product in 58% yield (entry 8). The present standard conditions gave **3a** in 15% yield, even under ambient conditions without the light (entry 9). In contrast, the reaction rarely occurred in complete darkness (entry 10). Furthermore, a periodic light-ON/OFF switching experiment in the cobalt-catalysed reaction of **1a** with **2a** (Fig. 1) and time-dependent UV-Vis absorption spectra of methylcobalamin under blue-light irradiation (Fig. 2) indicated that the photo-induced homolytic cleavage of the Co–C bond on the alkyl-Co(III) intermediate produced a Co(II) complex and an alkyl radical.<sup>19</sup> A slight reduction of the yield was obtained using Zn instead of Mn (entry 11). Control experiments confirmed that no reaction occurred in the absence of VB<sub>12</sub> catalyst (entry 12) and Mn reductant (entry 13). Finally, increasing the amount of olefin **2a** (1.5–2.0 equiv.) improved the yield (70–82%, entries 14 and 15).

We explored the substrate scope for alkyl tosylates and activated olefins with these optimised reaction conditions in hand (Table 2). Various activated olefins possessing amide, ester, ketone, sulfonyl and pinacolboronyl functions were well tolerated in the reaction with some primary alkyl tosylates, producing the target products **3b–3h** in high yields. Notably, the radical addition was performed on a gram-scale to give the desired adduct **3h** in 78% yield.

Next, the substrate scope for alkyl tosylates was evaluated. Branched alkyl chain (for **3i**), aryl ethers (for **3j** and **3k**), internal (for **3l**) and terminal alkynyl groups (for **3m**), a *cis*-configured alkenyl group (for **3n**) and an ester (for **3o**) afforded the corresponding radical adducts in good yields (60–85%). An alkyl

Table 1 Optimisation of the reaction conditions<sup>a</sup>

Entry	Changes from standard conditions	Yield <sup>b</sup> (%)
1	None	61
2	Without Et <sub>3</sub> N·HCl	10
3	H <sub>2</sub> O instead of Et <sub>3</sub> N·HCl	24
4	NH <sub>4</sub> Cl instead of Et <sub>3</sub> N·HCl	54
5	CoCl(dmgH) <sub>2</sub> Py <sup>c</sup> instead of VB <sub>12</sub>	Trace
6	Co(salen) <sup>c</sup> instead of VB <sub>12</sub>	Trace
7	Cp(Pc) <sup>c</sup> instead of VB <sub>12</sub>	Trace
8	Co(TMPP) <sup>c</sup> instead of VB <sub>12</sub>	58
9	In ambient conditions <sup>d</sup>	15
10	In dark	5
11	Zn instead of Mn	51
12	Without VB <sub>12</sub>	0
13	Without Mn reductant	0
14	Olefin (1.5 equiv.)	82 (80)
15	Olefin (2.0 equiv.)	70

<sup>a</sup> Standard reaction conditions: **1a** (0.25 mmol), **2a** (0.25 mmol), VB<sub>12</sub> (12.5  $\mu$ mol, 5 mol%), Et<sub>3</sub>N·HCl (0.38 mmol), Mn powder (0.75 mmol), DMF (1.5 mL); room temperature for 8 h; the blue-light irradiation; argon atmosphere. <sup>b</sup> Yields were determined by GC using mesitylene as an internal standard. The parenthesis value indicates the isolated yield. <sup>c</sup> See ESI. <sup>d</sup> The reaction was performed without blue-light irradiation.

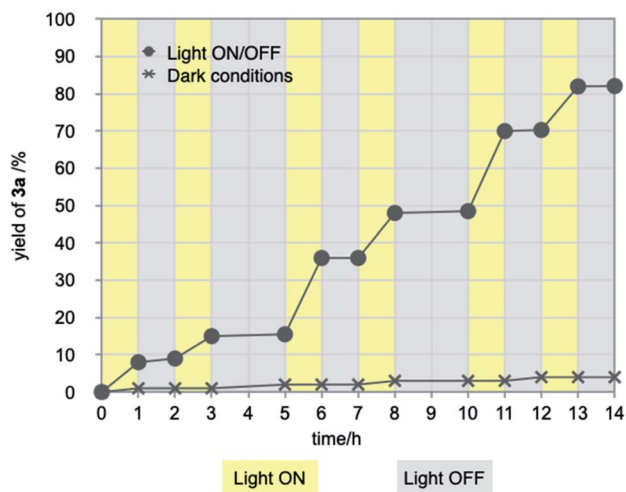
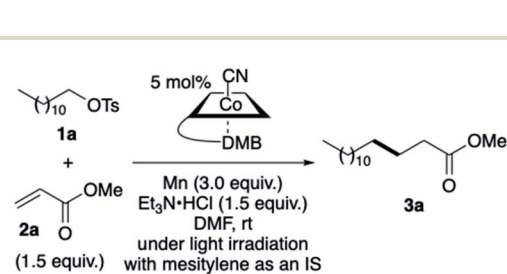


Fig. 1 Plots of product yield (%) against reaction time (h) in the cobalt-catalysed Giese reaction of **1a** and **2a**, ●: under light ON/OFF conditions, ×: in dark.

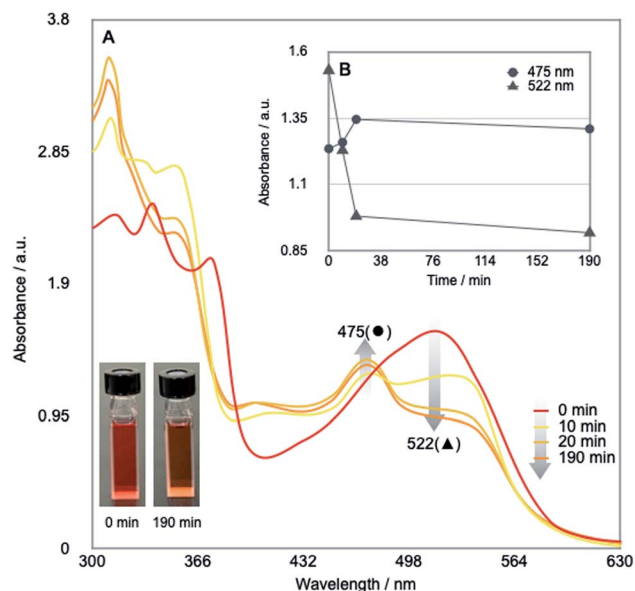


Fig. 2 (A) Time-dependent UV-Vis absorption spectra of Me-Cbl in DMF ( $10^{-4}$  mol L $^{-1}$ ) under the blue-light irradiation under argon atmosphere. (B) Variation over time of absorbances at 475 and 522 nm.

chloride was also suitable for this reaction, being converted to radical adduct **3p** in 79% yield. Interestingly, a chloro-substituent located at the sterically congested position did not participate in the transformation; chloro-substituted radical adduct **3q** was selectively obtained in 75% yield. In the alkyl radical generation from alkyl tosylates, steric factors can be expected to affect the  $S_N2$  displacement of nucleophilic cobalt(i). Indeed, the conversion of bulky alkyl tosylates such as **3r–3u** was slow under the optimised conditions. The reactivity was improved by increasing the catalyst loading up to 10 mol% (for **3r** and **3s**) or using a smaller planar-cobalt catalyst like Co(TMPP) (for **3t** and **3u**), providing the corresponding adducts in 70–85% yields, even in the reaction with secondary alkyl tosylate like 3-pyrrolidinyl tosylate. Unfortunately, secondary alkyl tosylates such as cyclohexyl and 2-hexyl tosylates were not compatible with the reaction. The low reactivity associated with the steric repulsion between the cobalt catalyst and the tosylate is a drawback of this reaction; however, it could be also considered a peculiarity. Thus, since in the tosylation of alkyl polyols the tosylation selectively occurs at the primary alcoholic position,<sup>20</sup> multi-functionalised primary alkyl tosylates could be readily synthesised by the tosylation of polyols and subsequent protection of the remaining hydroxy groups. We believed that a regioselective Giese reaction starting from polyols could be envisaged by combining the present radical addition with the classical tosylation. This strategy enabled the regioselective transformation of various diols bearing primary and secondary alcoholic moieties for the synthesis of functionalised alcohol derivatives **3v–3y**. Although the present reaction was strongly inhibited by a naked hydroxyl group, pre-treatment of 3-hydroxybutyl tosylate with trimethylsilyl chloride (1.2 equiv.) in the presence of Mn (4.2 equiv.) at 25 °C for 3 h enabled the direct conversion of the hydroxylated alkyl tosylate to the corresponding

radical adduct **3z** in 75% yield without any deprotection process. A selective Giese reaction of a triol derivative having one TBSO and two TsO groups in different steric environments was also demonstrated, in which the reaction proceeded at the less bulky TsO site without lacking the other TsO group (for **3A**).

The reaction of 5-hexenyl tosylate (**4**) with phenyl acrylate under the standard conditions afforded a mixture of cyclised **5** and linear adducts **6** in 19% and 34% yields, respectively (Scheme 2, eqn (1)). Additionally, the VB<sub>12</sub>-catalysed reaction of dodecyl tosylate (**1a**) in the presence of  $\gamma$ -terpinene (3.0 equiv.) as a hydrogen atom donor<sup>21</sup> afforded dodecane (**7**) in 80% GC yield. These results support the generation of alkyl radicals during the reaction. Additionally, in the absence of alkyl tosylate, phenyl acrylate produced diphenyl adipate in 30% yield and 60% conversion (Scheme 2, eqn (2)). This result indicates that the nucleophilic cobalt generated *in situ* could react faster with alkyl tosylates than with activated olefins.<sup>22</sup>

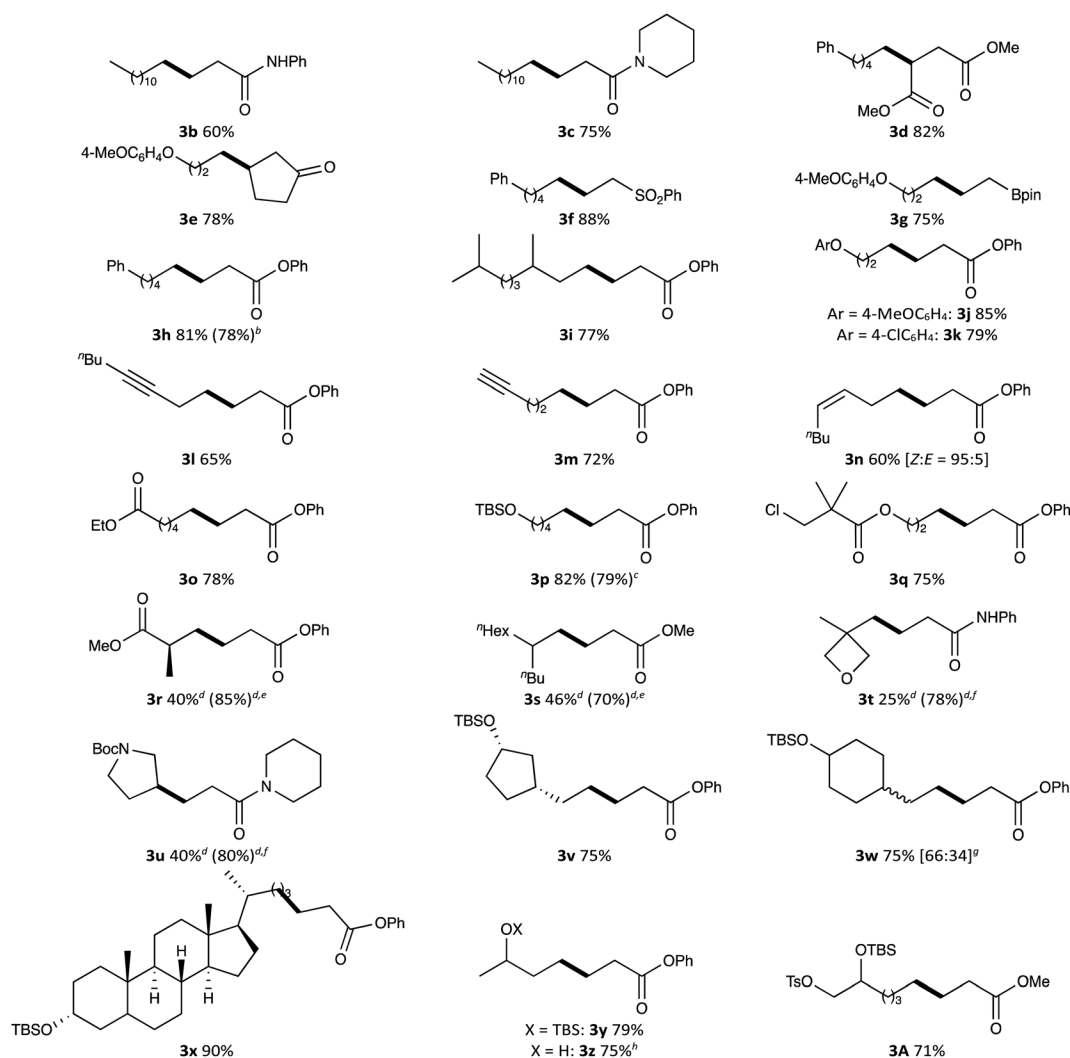
According to these results, we tentatively proposed a catalytic cycle depicted in Scheme 3. The  $S_N2$ -type oxidative addition of alkyl tosylates to nucleophilic cobalt(i) **A** would afford alkyl-cobalt(iii) **B**, which would undergo the Co–C bond cleavage under the visible-light irradiation to generate an alkyl radical and planar-cobalt(ii) **C**. The generated alkyl radical would add to the  $\beta$ -position of an activated olefin, followed by reduction with Mn and protonation to give adduct **3**. The catalytic cycle would be closed by Mn reduction of planar-cobalt(ii) **C** to regenerate the nucleophilic cobalt(i) **A**.

In conclusion, we have demonstrated a Giese reaction utilising readily available alkyl tosylates as alkyl radical sources for the first time, which proceeded efficiently in the presence of vitamin B<sub>12</sub> or Co(TTMP) catalysts, Mn reductant, and a proton donor under the visible-light irradiation. The critical step in this reaction is the generation of alkyl-cobalt(iii) intermediates by  $S_N2$ -type oxidative addition of alkyl tosylates to nucleophilic cobalt(i). Hence, the bulkiness around the TsO group on alkyl tosylates affected the reaction efficiency. Nevertheless, classical tosylation and subsequent Giese reaction enabled the regioselective functionalisation of polyol compounds. Further studies on the alkyl radical generation protocol to extend the scope of this strategy to novel organic transformations are now ongoing in our laboratory.

## Experimental

### General procedure for cobalt-catalysed Giese reaction of alkyl tosylates

In a Schlenk tube equipped with a stirring bar, Mn powder (41.2 mg, 0.75 mmol) was added and heated at 400 °C for 3 min under vacuum. After cooling and filling with argon, the tube was added triethylamine hydrochloride Et<sub>3</sub>N·HCl (51.6 mg, 0.38 mmol) and vitamin B<sub>12</sub> (VB<sub>12</sub>; 16.9 mg, 12.5  $\mu$ mol), evacuated and filled with argon three times. After adding dry DMF (1.5 mL) and stirring for 5 min, the color instantly changed to green by adding trimethylsilyl chloride (*ca.* 6  $\mu$ L). Then, alkyl tosylate (0.25 mmol) and activated olefin (0.35 mmol) were added to the solution. The reaction mixture was stirred at room temperature for the appropriate time (18–24 h) under blue-light irradiation using a PhotoRedOx Box (HeptoChem Inc.) equipped with

Table 2 Substrate scope of the cobalt-catalysed Giese reaction of alkyl tosylates with activated olefins<sup>e</sup>

<sup>a</sup> Reaction conditions: alkyl tosylate (0.25 mmol), activated olefin (0.375 mmol), vitamin B<sub>12</sub> (12.5 μmol, 5 mol%), Et<sub>3</sub>N·HCl (0.38 mmol), Mn powder (0.75 mmol), DMF (1.5 mL); room temperature for 16 h; the blue-light irradiation; argon atmosphere. Isolated yield. <sup>b</sup> The reaction was conducted with 5-phenylpentyl tosylate (5 mmol) and phenyl acrylate (7.5 mmol). <sup>c</sup> 5-(*tert*-Butyldimethylsilyl)-1-chloropentane was used instead of the corresponding tosylate. <sup>d</sup> Reaction time: 24 h. <sup>e</sup> 10 mol% of vitamin B<sub>12</sub> was used. <sup>f</sup> Co(TMPP) was used instead of vitamin B<sub>12</sub>. <sup>g</sup> Diastereomeric alkyl tosylate [65 : 35] was used. The values in bracket indicate a diastereomeric ratio of the radical adduct. <sup>h</sup> See main text.

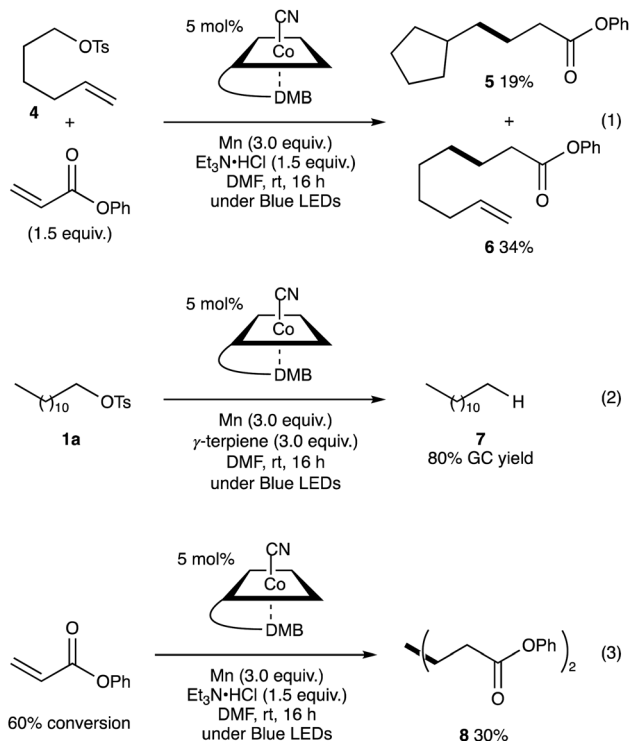
a Kessil A160W TUNE BLUE ( $\lambda = 454$  nm, 40 W) as shown in Fig. S1 (ESI).<sup>†</sup> The color of the reaction mixture drastically changed in each step, as depicted in Fig. S2 (ESI).<sup>†</sup> The obtained mixture was diluted with ethyl acetate and quenched by saturated aqueous NH<sub>4</sub>Cl. The aqueous phase was extracted with ethyl acetate, and the combined organic phase was dried over anhydrous MgSO<sub>4</sub>. After filtration and removal of the solvent, the residue was purified by silica-gel flash column chromatography to give the radical adduct.

#### Procedure for direct Giese reaction of 3-hydroxybutyl tosylate: synthesis of **3z**

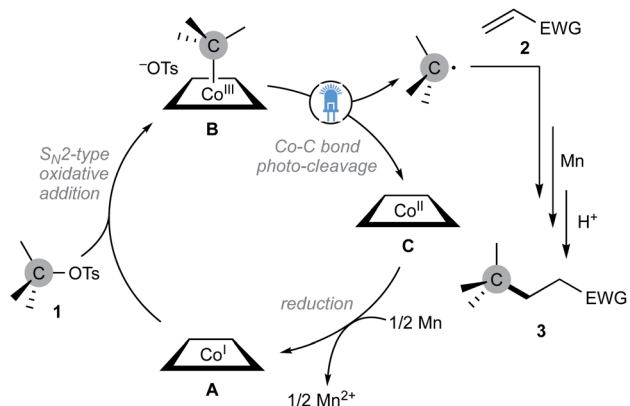
In a Schlenk tube equipped with a stirring bar, Mn powder (57.7 mg, 1.05 mmol) was added and heated at 400 °C for 3 min

under vacuum. After cooling and filling with argon, the tube was added Et<sub>3</sub>N·HCl (51.6 mg, 0.38 mmol), DMF (1.5 mL) and 3-hydroxybutyl tosylate (61.1 mg, 0.25 mmol). Trimethylsilyl chloride (38 μL, 0.30 mmol) was added into the mixture. The solution was stirred for 3 h at 25 °C. Then, vitamin B<sub>12</sub> (VB<sub>12</sub>; 16.9 mg, 12.5 μmol) and phenyl acrylate (55.6 mg, 0.38 mmol) were successively added. The reaction mixture was stirred at room temperature for 16 h under blue-light irradiation. After reaction, the obtained mixture was diluted with ethyl acetate and quenched by saturated aqueous NH<sub>4</sub>Cl. The aqueous phase was extracted with ethyl acetate, and the combined organic phase was dried over anhydrous MgSO<sub>4</sub>. After filtration and removal of the solvent, the residue was purified by silica-gel flash column chromatography using as 10% ethyl acetate in hexane an eluent to give the radical adduct **3z** (41.7 mg, colorless oil) in 75% yield.





**Scheme 2** Radical clock experiment using 5-hexenyl tosylate **6** (eqn (1)). Radical trapping reaction with  $\gamma$ -terpinene (eqn (2)). Reductive dimerization of phenyl acrylate (eqn (3)).



**Scheme 3** Plausible reaction mechanism of nucleophilic cobalt-catalysed Giese reaction with alkyl tosylates under the light irradiation.

### Characterisation of products

**3a.** Isolated as a colorless oil in 80% yield;  $^1\text{H NMR}$  (500 MHz,  $\text{CDCl}_3$ )  $\delta$  3.66 (s, 3H), 2.30 (t,  $J = 7.6$  Hz, 2H), 1.66–1.57 (m, 2H), 1.35–1.16 (m, 22H), 0.87 (t,  $J = 7.0$  Hz, 3H);  $^{13}\text{C NMR}$  (126 MHz,  $\text{CDCl}_3$ )  $\delta$  174.4, 51.4, 34.1, 31.9, 29.7, 29.7, 29.6, 29.6, 29.4, 29.3, 29.2, 29.1, 24.9, 22.7, 14.1; HRMS calcd for  $\text{C}_{16}\text{H}_{33}\text{O}_2$  [ $\text{M} + \text{H}$ ] $^+$ : 257.2481, found 257.2479.

**3b.** Isolated as a colorless oil in 60% yield;  $^1\text{H NMR}$  (500 MHz,  $\text{CDCl}_3$ )  $\delta$  7.51 (d,  $J = 8.0$  Hz, 2H), 7.35–7.28 (m, 2H), 7.13–7.06 (m, 1H), 2.35 (t,  $J = 7.6$  Hz, 2H), 1.73 (p,  $J = 7.5$  Hz, 2H),

1.42–1.27 (m, 4H), 1.25 (s, 19H), 0.88 (t,  $J = 6.9$  Hz, 3H);  $^{13}\text{C NMR}$  (126 MHz,  $\text{CDCl}_3$ )  $\delta$  171.3, 138.0, 129.0, 124.1, 119.7, 37.9, 31.9, 29.7, 29.7, 29.6, 29.6, 29.5, 29.4, 29.3, 29.3, 25.6, 22.7, 14.1; HRMS calcd for  $\text{C}_{21}\text{H}_{36}\text{NO}$  [ $\text{M} + \text{H}$ ] $^+$ : 318.2797, found 304.2642.

**3c.** Isolated as a colorless oil in 75% yield;  $^1\text{H NMR}$  (500 MHz,  $\text{CDCl}_3$ )  $\delta$  3.54 (t,  $J = 5.6$  Hz, 2H), 3.38 (t,  $J = 5.4$  Hz, 2H), 2.34–2.26 (m, 2H), 1.71–1.44 (m, 10H), 1.24 (s, 24H), 0.87 (t,  $J = 6.8$  Hz, 3H);  $^{13}\text{C NMR}$  (126 MHz,  $\text{CDCl}_3$ )  $\delta$  171.5, 46.7, 42.6, 33.5, 31.9, 29.7, 29.7, 29.6, 29.6, 29.5, 29.5, 29.4, 29.3, 26.6, 25.6, 25.5, 24.6, 22.7, 14.1; HRMS calcd for  $\text{C}_{20}\text{H}_{40}\text{NO}$  [ $\text{M} + \text{H}$ ] $^+$ : 310.3110, found 310.3108.

**3d.** Isolated as a colorless oil in 82% yield;  $^1\text{H NMR}$  (500 MHz,  $\text{CDCl}_3$ )  $\delta$  7.27 (t,  $J = 7.5$  Hz, 2H), 7.17 (t,  $J = 8.8$  Hz, 3H), 3.69 (s, 3H), 3.67 (s, 3H), 2.83 (dddd,  $J = 9.3, 7.4, 6.3, 5.1$  Hz, 1H), 2.71 (dd,  $J = 16.5, 9.3$  Hz, 1H), 2.62–2.55 (m, 2H), 2.42 (dd,  $J = 16.5, 5.2$  Hz, 1H), 1.69–1.56 (m, 3H), 1.55–1.45 (m, 1H), 1.32 (qt,  $J = 6.6, 2.6$  Hz, 4H);  $^{13}\text{C NMR}$  (126 MHz,  $\text{CDCl}_3$ )  $\delta$  175.4, 172.4, 142.5, 128.4, 128.2, 125.6, 51.8, 51.7, 41.1, 35.8, 31.8, 31.2, 29.0, 26.8; HRMS calcd for  $\text{C}_{17}\text{H}_{25}\text{O}_4$  [ $\text{M} + \text{H}$ ] $^+$ : 293.1753, found 293.1754.

**3e.** Isolated as a colorless oil in 78% yield;  $^1\text{H NMR}$  (500 MHz,  $\text{CDCl}_3$ )  $\delta$  6.83 (s, 4H), 3.92 (t,  $J = 6.3$  Hz, 2H), 3.76 (s, 3H), 2.46–2.37 (m, 1H), 2.37–2.26 (m, 1H), 2.28–2.10 (m, 3H), 1.89–1.73 (m, 3H), 1.69–1.46 (m, 3H);  $^{13}\text{C NMR}$  (126 MHz,  $\text{CDCl}_3$ )  $\delta$  219.5, 153.8, 153.1, 115.4, 114.6, 68.4, 55.7, 45.2, 38.5, 37.0, 32.1, 29.5, 27.7; HRMS calcd for  $\text{C}_{15}\text{H}_{21}\text{O}_3$  [ $\text{M} + \text{H}$ ] $^+$ : 249.1491, found 249.1493.

**3f.** Isolated as a colorless oil in 88% yield;  $^1\text{H NMR}$  (500 MHz,  $\text{CDCl}_3$ )  $\delta$  7.91 (dd,  $J = 8.4, 1.4$  Hz, 2H), 7.69–7.62 (m, 1H), 7.60–7.53 (m, 2H), 7.30–7.23 (m, 2H), 7.20–7.12 (m, 3H), 3.11–3.04 (m, 2H), 2.61–2.54 (m, 2H), 1.70 (ddt,  $J = 13.4, 10.6, 6.4$  Hz, 2H), 1.61–1.53 (m, 2H), 1.41–1.19 (m, 6H);  $^{13}\text{C NMR}$  (126 MHz,  $\text{CDCl}_3$ )  $\delta$  142.5, 139.1, 133.6, 129.2, 128.3, 128.2, 128.0, 125.6, 56.2, 35.8, 31.2, 28.8, 28.8, 28.1, 22.5; HRMS calcd for  $\text{C}_{19}\text{H}_{25}\text{O}_2\text{S}$  [ $\text{M} + \text{H}$ ] $^+$ : 317.1575, found 317.1574.

**3g.** Isolated as a colorless oil in 75% yield;  $^1\text{H NMR}$  (500 MHz,  $\text{CDCl}_3$ )  $\delta$  6.82 (s, 4H), 3.89 (t,  $J = 6.6$  Hz, 2H), 3.76 (s, 3H), 1.88 (ddt,  $J = 16.2, 9.6, 6.8$  Hz, 2H), 1.75 (p,  $J = 6.8$  Hz, 1H), 1.46 (dtd,  $J = 7.8, 5.0, 2.4$  Hz, 2H), 1.28–1.20 (m, 13H), 0.80 (t,  $J = 7.4$  Hz, 2H);  $^{13}\text{C NMR}$  (126 MHz,  $\text{CDCl}_3$ )  $\delta$  153.6, 153.3, 115.4, 114.6, 82.9, 68.6, 55.7, 29.7, 29.1, 24.8, 23.8, one signal was obscured; HRMS calcd for  $\text{C}_{18}\text{H}_{30}\text{BO}_4$  [ $\text{M} + \text{H}$ ] $^+$ : 321.2237, found 321.2235.

**3h.** Isolated as a colorless oil in 81% yield;  $^1\text{H NMR}$  (500 MHz,  $\text{CDCl}_3$ )  $\delta$  7.44–7.35 (m, 2H), 7.35–7.30 (m, 1H), 7.30–7.15 (m, 5H), 7.14–7.05 (m, 2H), 2.67–2.60 (m, 2H), 2.57 (t,  $J = 7.5$  Hz, 2H), 1.77 (t,  $J = 7.3$  Hz, 2H), 1.70–1.60 (m, 2H), 1.48–1.34 (m, 6H);  $^{13}\text{C NMR}$  (126 MHz,  $\text{CDCl}_3$ )  $\delta$  172.3, 150.7, 142.7, 129.4, 128.4, 128.2, 125.7, 125.6, 121.5, 35.9, 34.3, 31.4, 29.1, 29.1, 29.0, 24.9; HRMS calcd for  $\text{C}_{20}\text{H}_{24}\text{O}_2$  [ $\text{M} + \text{H}$ ] $^+$ : 297.1855, found 297.1850.

**3i.** Isolated as a colorless oil in 77% yield;  $^1\text{H NMR}$  (500 MHz,  $\text{CDCl}_3$ )  $\delta$  7.41–7.34 (m, 2H), 7.25–7.19 (m, 1H), 7.11–7.05 (m, 2H), 2.56 (t,  $J = 7.5$  Hz, 2H), 1.82–1.66 (m, 2H), 1.53 (dt,  $J = 13.3, 6.6$  Hz, 1H), 1.49–1.03 (m, 11H), 0.87 (d,  $J = 6.6$  Hz, 9H);  $^{13}\text{C NMR}$  (126 MHz,  $\text{CDCl}_3$ )  $\delta$  172.3, 150.7, 129.4, 125.7, 121.6, 39.3,

37.2, 36.6, 34.4, 32.6, 28.0, 26.6, 25.3, 24.8, 22.7, 22.6, 19.6; HRMS calcd for  $C_{19}H_{30}O_2$   $[M + H]^+$ : 291.2324, found 291.2326.

**3j.** Isolated as a colorless oil in 85% yield;  $^1H$  NMR (500 MHz,  $CDCl_3$ )  $\delta$  7.42–7.35 (m, 2H), 7.28–7.20 (m, 1H), 7.13–7.05 (m, 2H), 6.85 (s, 3H), 3.95 (t,  $J = 6.4$  Hz, 2H), 3.77 (s, 3H), 2.61 (t,  $J = 7.5$  Hz, 2H), 1.84 (ddd,  $J = 13.7, 8.8, 6.7$  Hz, 5H), 1.61 (ddt,  $J = 9.5, 7.0, 2.0$  Hz, 2H);  $^{13}C$  NMR (126 MHz,  $CDCl_3$ )  $\delta$  172.1, 153.7, 153.1, 150.7, 129.4, 125.7, 121.5, 115.4, 114.6, 68.2, 55.7, 34.2, 29.0, 25.6, 24.6; HRMS calcd for  $C_{19}H_{22}O_4$   $[M + H]^+$ : 315.1596, found 315.1597.

**3k.** Isolated as a colorless oil in 79% yield;  $^1H$  NMR (500 MHz,  $CDCl_3$ )  $\delta$  7.42–7.34 (m, 2H), 7.25–7.20 (m, 3H), 7.11–7.05 (m, 2H), 6.86–6.79 (m, 2H), 3.95 (t,  $J = 6.4$  Hz, 2H), 2.61 (t,  $J = 7.4$  Hz, 2H), 1.89–1.79 (m, 4H), 1.66–1.56 (m, 2H);  $^{13}C$  NMR (126 MHz,  $CDCl_3$ )  $\delta$  172.0, 157.6, 150.6, 129.4, 129.2, 125.7, 125.3, 121.5, 115.7, 67.8, 34.2, 28.8, 25.5, 24.6; HRMS calcd for  $C_{18}H_{20}ClO_3$   $[M + H]^+$ : 319.1101, found 319.1099.

**3l.** Isolated as a colorless oil in 65% yield;  $^1H$  NMR (500 MHz,  $CDCl_3$ )  $\delta$  7.41–7.34 (m, 2H), 7.22 (ddt,  $J = 7.7, 7.0, 1.1$  Hz, 1H), 7.11–7.04 (m, 2H), 2.58 (dd,  $J = 7.9, 7.1$  Hz, 2H), 2.23 (tt,  $J = 7.0, 2.4$  Hz, 2H), 2.15 (tt,  $J = 7.1, 2.4$  Hz, 2H), 1.91–1.82 (m, 2H), 1.66–1.57 (m, 2H), 1.51–1.34 (m, 4H), 0.90 (t,  $J = 7.2$  Hz, 3H);  $^{13}C$  NMR (126 MHz,  $CDCl_3$ )  $\delta$  172.0, 150.7, 129.4, 125.7, 121.5, 80.8, 79.3, 33.9, 31.2, 28.4, 24.1, 21.9, 18.5, 18.4, 13.6; HRMS calcd for  $C_{17}H_{23}O_2$   $[M + H]^+$ : 259.1698, found 259.1696.

**3m.** Isolated as a colorless oil in 72% yield;  $^1H$  NMR (500 MHz,  $CDCl_3$ )  $\delta$  7.38 (t,  $J = 8.0$  Hz, 2H), 7.22 (t,  $J = 7.4$  Hz, 1H), 7.07 (d,  $J = 7.6$  Hz, 2H), 2.58 (t,  $J = 7.5$  Hz, 2H), 2.23 (td,  $J = 6.8, 2.6$  Hz, 2H), 1.96 (t,  $J = 2.6$  Hz, 1H), 1.78 (p,  $J = 7.4$  Hz, 2H), 1.65–1.49 (m, 4H);  $^{13}C$  NMR (126 MHz,  $CDCl_3$ )  $\delta$  172.1, 150.7, 129.4, 125.7, 121.6, 84.3, 68.4, 34.2, 28.1, 28.1, 24.4, 18.3; HRMS calcd for  $C_{14}H_{17}O_2$   $[M + H]^+$ : 217.1229, found 217.1227.

**3n.** Isolated as *cis*-rich stereoisomers (*cis/trans* = 95 : 5, colorless oil) in 60% yield;  $^1H$  NMR (500 MHz,  $CDCl_3$ )  $\delta$  7.41–7.33 (m, 2H), 7.28–7.15 (m, 1H), 7.11–7.04 (m, 2H), 5.40 (ddt,  $J = 10.8, 6.8, 1.4$  Hz, 1H), 5.36 (ddt,  $J = 10.8, 6.9, 1.4$  Hz, 1H), 2.57 (t,  $J = 7.4$  Hz, 2H), 2.16–1.99 (m, 3H), 1.83–1.73 (m, 2H), 1.57 (s, 1H), 1.53–1.43 (m, 2H), 1.38–1.24 (m, 4H), 0.90 (t,  $J = 7.1$  Hz, 3H);  $^{13}C$  NMR (126 MHz,  $CDCl_3$ )  $\delta$  172.2, 150.7, 130.6, 129.4, 129.0, 125.7, 121.6, 34.3, 31.9, 29.1, 26.9, 26.8, 24.6, 22.3, 14.0; HRMS calcd for  $C_{17}H_{25}O_2$   $[M + H]^+$ : 261.1855, found 261.1859.

**3o.** Isolated as a colorless oil in 78% yield;  $^1H$  NMR (500 MHz,  $CDCl_3$ )  $\delta$  7.41–7.33 (m, 2H), 7.22 (ddt,  $J = 7.1, 6.5, 1.1$  Hz, 1H), 7.10–7.03 (m, 2H), 4.12 (q,  $J = 7.1$  Hz, 2H), 2.55 (t,  $J = 7.5$  Hz, 2H), 2.30 (t,  $J = 7.5$  Hz, 2H), 1.79–1.71 (m, 2H), 1.68–1.57 (m, 2H), 1.47–1.30 (m, 6H), 1.25 (t,  $J = 7.1$  Hz, 3H);  $^{13}C$  NMR (126 MHz,  $CDCl_3$ )  $\delta$  173.8, 172.2, 150.7, 129.4, 125.7, 121.5, 60.2, 34.3, 34.3, 28.9, 28.9, 24.9, 24.8, 14.2; HRMS calcd for  $C_{17}H_{24}O_4$   $[M + H]^+$ : 293.1753, found 293.1755.

**3p.** Isolated as a colorless oil in 82% yield;  $^1H$  NMR (500 MHz,  $CDCl_3$ )  $\delta$  7.37 (dd,  $J = 8.5, 7.4$  Hz, 2H), 7.27–7.18 (m, 1H), 7.08 (dd,  $J = 8.7, 1.2$  Hz, 2H), 3.61 (t,  $J = 6.6$  Hz, 2H), 2.56 (t,  $J = 7.5$  Hz, 2H), 1.76 (p,  $J = 7.5$  Hz, 2H), 1.58–1.49 (m, 2H), 1.41–1.31 (m, 6H), 0.90 (s, 9H), 0.06 (s, 6H);  $^{13}C$  NMR (126 MHz,  $CDCl_3$ )  $\delta$  172.3, 150.7, 129.4, 125.7, 121.5, 63.2, 34.4, 32.8, 29.1, 29.0, 26.0, 25.6, 24.9, 18.4, –5.3; HRMS calcd for  $C_{20}H_{35}O_3Si$   $[M + H]^+$ : 351.2355, found 351.2359.

**3q.** Isolated as a yellow-viscous oil in 75% yield;  $^1H$  NMR (500 MHz,  $CDCl_3$ )  $\delta$  7.41–7.34 (m, 2H), 7.22 (td,  $J = 7.3, 1.1$  Hz, 1H), 7.11–7.04 (m, 2H), 4.15 (t,  $J = 6.5$  Hz, 2H), 3.61 (s, 2H), 2.58 (t,  $J = 7.5$  Hz, 2H), 1.80 (p,  $J = 7.5$  Hz, 1H), 1.72 (dq,  $J = 8.2, 6.6$  Hz, 0H), 1.56–1.46 (m, 2H), 1.29 (s, 6H);  $^{13}C$  NMR (126 MHz,  $CDCl_3$ )  $\delta$  175.0, 172.0, 150.7, 129.4, 125.8, 121.5, 64.7, 52.1, 44.6, 34.2, 28.3, 25.5, 24.5, 23.2; HRMS calcd for  $C_{17}H_{24}ClO_4$   $[M + H]^+$ : 327.1363, found 327.1365.

**3r.** Isolated as a colorless oil in 85% yield;  $^1H$  NMR (500 MHz,  $CDCl_3$ )  $\delta$  7.38 (t,  $J = 7.8$  Hz, 2H), 7.22 (t,  $J = 7.4$  Hz, 1H), 7.08 (t,  $J = 6.6$  Hz, 2H), 3.69 (s, 3H), 2.64 (ddt,  $J = 7.1, 4.4, 2.3$  Hz, 1H), 2.60–2.54 (m, 1H), 2.50 (p,  $J = 6.9$  Hz, 1H), 1.93–1.86 (m, 1H), 1.84–1.68 (m, 2H), 1.61–1.49 (m, 1H), 1.20 (d,  $J = 7.0$  Hz, 3H);  $^{13}C$  NMR (126 MHz,  $CDCl_3$ )  $\delta$  176.8, 171.8, 150.6, 129.4, 125.7, 121.5, 51.6, 39.2, 34.1, 33.0, 24.3, 22.6, 17.1; HRMS calcd for  $C_{14}H_{19}O_4$   $[M + H]^+$ : 251.1283, found 251.1286.

**3s.** Isolated as a yellow-viscous oil in 70% yield;  $^1H$  NMR (500 MHz,  $CDCl_3$ )  $\delta$  3.67 (s, 3H), 2.28 (t,  $J = 7.6$  Hz, 2H), 1.64–1.54 (m, 3H), 1.34–1.24 (m, 12H), 1.28–1.19 (m, 6H), 0.88 (td,  $J = 7.0, 2.3$  Hz, 6H);  $^{13}C$  NMR (126 MHz,  $CDCl_3$ )  $\delta$  174.3, 51.4, 37.2, 34.5, 33.5, 33.2, 33.1, 31.9, 29.8, 28.9, 26.6, 23.1, 22.7, 22.2, 14.2, 14.1; HRMS calcd for  $C_{16}H_{33}O_2$   $[M + H]^+$ : 257.2481, found 257.2492.

**3t.** Isolated as a red oil in 78% yield;  $^1H$  NMR (400 MHz,  $CDCl_3$ )  $\delta$  7.51 (d,  $J = 7.9$  Hz, 2H), 7.36 (s, 1H), 7.31 (t,  $J = 7.9$  Hz, 2H), 7.10 (t,  $J = 7.4$  Hz, 1H), 4.43 (d,  $J = 5.6$  Hz, 2H), 4.35 (d,  $J = 5.6$  Hz, 2H), 2.41–2.33 (m, 2H), 1.76–1.64 (m, 4H), 1.30 (s, 3H);  $^{13}C$  NMR (126 MHz,  $CDCl_3$ )  $\delta$  170.8, 137.8, 129.0, 124.2, 119.8, 82.7, 39.1, 38.5, 37.7, 23.1, 20.5; HRMS calcd for  $C_{14}H_{20}NO_2$   $[M + H]^+$ : 234.1494, found 234.1491.

**3u.** Isolated as a colorless oil in 80% yield;  $^1H$  NMR (500 MHz,  $CDCl_3$ )  $\delta$  3.60–3.45 (m, 3H), 3.38 (dq,  $J = 10.9, 5.0$  Hz, 3H), 3.24 (dtd,  $J = 16.5, 10.0, 7.0$  Hz, 1H), 2.96–2.81 (m, 1H), 2.41–2.26 (m, 2H), 2.23–2.08 (m, 1H), 2.06–1.93 (m, 1H), 1.84–1.60 (m, 5H), 1.60–1.49 (m, 4H), 1.45 (s, 9H);  $^{13}C$  NMR (126 MHz,  $CDCl_3$ )  $\delta$  170.6, 154.6, 79.0, 51.5, 50.9, 46.6, 45.7, 45.4, 42.7, 38.8, 37.9, 31.8, 31.8, 31.0, 29.7, 28.6, 28.5, 26.5, 25.5, 24.5; HRMS calcd for  $C_{17}H_{31}N_2O_3$   $[M + H]^+$ : 311.2335, found 311.2339.

**3v.** Isolated as a colorless oil in 75% yield;  $^1H$  NMR (500 MHz,  $CDCl_3$ )  $\delta$  7.41–7.35 (m, 2H), 7.22 (ddt,  $J = 7.8, 7.0, 1.1$  Hz, 1H), 7.10–7.05 (m, 2H), 4.20 (q,  $J = 5.9$  Hz, 1H), 2.56 (t,  $J = 7.5$  Hz, 2H), 2.06–1.98 (m, 1H), 1.84–1.65 (m, 5H), 1.62–1.52 (m, 1H), 1.47–1.31 (m, 5H), 1.17 (dddd,  $J = 12.8, 8.9, 6.2, 0.8$  Hz, 1H), 0.89 (s, 9H), 0.04 (s, 6H);  $^{13}C$  NMR (126 MHz,  $CDCl_3$ )  $\delta$  172.3, 150.7, 129.4, 125.7, 121.6, 74.1, 42.6, 37.5, 36.4, 35.3, 34.4, 29.9, 28.0, 25.9, 25.1, 18.1, –4.7, –4.7; HRMS calcd for  $C_{22}H_{37}O_3Si$   $[M + H]^+$ : 377.2512, found 377.2510.

**3w.** Isolated as a *cis/trans*-mixture (colorless oil) in 75% [66 : 34] yield;  $^1H$  NMR (500 MHz,  $CDCl_3$ )  $\delta$  7.43–7.34 (m, 2H), 7.22 (ddd,  $J = 8.0, 7.0, 1.1$  Hz, 1H), 7.08 (ddt,  $J = 7.6, 2.0, 1.0$  Hz, 2H), 3.92 (q,  $J = 3.5, 3.1$  Hz, 0.66H), 3.56–3.46 (m, 0.34H), 2.56 (tdd,  $J = 7.6, 3.1, 0.9$  Hz, 2H), 1.88–1.81 (m, 1H), 1.79–1.68 (m, 3H), 1.66–1.56 (m, 2H), 1.40 (p,  $J = 4.0, 3.5$  Hz, 2H), 1.34–1.19 (m, 4H), 1.22 (s, 1H), 0.90 (d,  $J = 2.7$  Hz, 9H), 0.06 (s, 1H);  $^{13}C$  NMR (126 MHz,  $CDCl_3$ )  $\delta$  172.3, 172.2, 150.8, 129.5, 129.2, 125.8, 125.5, 121.7, 121.4, 72.0, 71.9, 67.4, 67.2, 36.4, 36.0, 34.6, 34.4, 34.4, 34.2, 33.3, 33.2, 31.5, 31.3, 27.1, 26.7, 26.5, 26.0, 26.0, 25.9,

25.9, 25.9, 25.8, 25.7, 25.2, 25.0, 18.2, 18.1, -4.5, -4.7, -4.7, -4.9; HRMS calcd for  $C_{23}H_{39}O_3Si [M + H]^+$ : 391.2668, found 391.2668.

**3x.** Isolated as a yellow viscous oil in 90% yield;  $^1H$  NMR (500 MHz,  $CDCl_3$ )  $\delta$  7.37 (t,  $J = 7.7$  Hz, 2H), 7.22 (t,  $J = 7.5$  Hz, 1H), 7.08 (d,  $J = 8.6$  Hz, 2H), 3.58 (tt,  $J = 10.5, 4.6$  Hz, 1H), 2.55 (t,  $J = 7.5$  Hz, 2H), 1.96 (dt,  $J = 12.4, 3.2$  Hz, 1H), 1.88–1.70 (m, 7H), 1.59–1.51 (m, 2H), 1.50–1.31 (m, 13H), 1.29–0.83 (m, 9H), 0.94–0.86 (m, 15H), 0.64 (s, 3H), 0.06 (s, 6H);  $^{13}C$  NMR (126 MHz,  $CDCl_3$ )  $\delta$  172.3, 150.7, 129.3, 125.7, 121.6, 72.8, 56.4, 56.2, 42.7, 42.3, 40.2, 40.1, 36.9, 35.8, 35.7, 35.6, 34.6, 34.4, 31.0, 29.6, 28.3, 27.3, 26.4, 26.0, 25.7, 25.0, 24.2, 23.4, 20.8, 18.6, 18.3, 12.0, -4.6; HRMS calcd for  $C_{39}H_{65}O_3Si [M + H]^+$ : 609.4703, found 609.4701.

**3y.** Isolated as a thin-yellow oil in 79% yield;  $^1H$  NMR (500 MHz,  $CDCl_3$ )  $\delta$  7.41–7.33 (m, 2H), 7.25–7.18 (m, 1H), 7.10–7.03 (m, 2H), 3.86–3.76 (m, 1H), 2.56 (t,  $J = 7.5$  Hz, 2H), 1.82–1.70 (m, 2H), 1.56–1.35 (m, 4H), 1.13 (d,  $J = 6.1$  Hz, 3H), 0.89 (s, 9H), 0.05 (s, 3H), 0.05 (s, 3H);  $^{13}C$  NMR (126 MHz,  $CDCl_3$ )  $\delta$  172.2, 150.7, 129.4, 125.7, 121.6, 68.4, 39.3, 34.4, 25.9, 25.3, 25.0, 23.8, 18.1, -4.4, -4.7; HRMS calcd for  $C_{19}H_{33}O_3Si [M + H]^+$ : 337.2199, found 337.2185.

**3z.** Isolated as a colorless oil in 75% yield;  $^1H$  NMR (500 MHz,  $CDCl_3$ )  $\delta$  7.41–7.34 (m, 2H), 7.25–7.18 (m, 1H), 7.10–7.05 (m, 2H), 3.87–3.78 (m, 1H), 2.58 (t,  $J = 7.5$  Hz, 2H), 1.86–1.70 (m, 2H), 1.59–1.41 (m, 5H), 1.20 (d,  $J = 6.1$  Hz, 3H);  $^{13}C$  NMR (126 MHz,  $CDCl_3$ )  $\delta$  172.2, 150.7, 129.4, 125.7, 121.5, 67.8, 38.8, 34.3, 25.2, 24.8, 23.5; HRMS calcd for  $C_{13}H_{19}O_3 [M + H]^+$ : 223.1334, found 223.1339.

**3A.** Isolated as a colorless oil in 71% yield;  $^1H$  NMR (500 MHz,  $CDCl_3$ )  $\delta$  7.81–7.75 (m, 2H), 7.37–7.31 (m, 2H), 3.90–3.77 (m, 3H), 3.66 (s, 3H), 2.45 (s, 3H), 2.29 (t,  $J = 7.5$  Hz, 2H), 1.64–1.54 (m, 3H), 1.45–1.33 (m, 1H), 1.32–1.18 (m, 6H), 0.83 (s, 9H), 0.01 (s, 3H), -0.00 (s, 3H);  $^{13}C$  NMR (126 MHz,  $CDCl_3$ )  $\delta$  174.2, 144.7, 133.0, 129.8, 127.9, 73.1, 69.9, 51.4, 34.0, 33.9, 29.2, 28.9, 25.7, 24.8, 24.6, 21.6, 18.0, -4.6, -4.8; HRMS calcd for  $C_{23}H_{41}O_6SSi [M + H]^+$ : 473.2393, found 473.2389.

Isolated as a colorless mixture of **5** and **6** (**5** : **6** = 65 : 35) in 53% yield;  $^1H$  NMR (500 MHz,  $CDCl_3$ )  $\delta$  7.44–7.31 (m, 2H), 7.29–7.18 (m, 1H), 7.14–7.02 (m, 2H), 5.82 (ddt,  $J = 16.9, 10.2, 6.7$  Hz, 0.53H), 5.01 (dq,  $J = 17.1, 1.7$  Hz, 0.52H), 4.95 (ddt,  $J = 10.2, 2.3, 1.2$  Hz, 0.53H), 2.56 (td,  $J = 7.5, 1.0$  Hz, 2H), 2.07 (tdd,  $J = 6.7, 5.3, 1.5$  Hz, 1H), 1.88–1.70 (m, 3H), 1.69–1.48 (m, 2H), 1.48–1.32 (m, 4H), 1.18–1.03 (m, 1H);  $^{13}C$  NMR (126 MHz,  $CDCl_3$ )  $\delta$  172.3, 172.2, 150.7, 150.7, 138.9, 129.4, 125.7, 125.7, 121.6, 121.5, 114.3, 39.8, 35.6, 34.6, 34.3, 33.7, 32.6, 28.9, 28.7, 28.7, 25.1, 24.9, 24.2; HRMS calcd for  $C_{15}H_{21}O_2 [M + H]^+$ : 233.1542, found 233.1538.

**8 (ref. 23).** Isolated as a colorless oil in 30% yield;  $^1H$  NMR (500 MHz,  $CDCl_3$ )  $\delta$  7.43–7.34 (m, 4H), 7.27–7.20 (m, 2H), 7.12–7.06 (m, 4H), 2.70–2.59 (m, 4H), 1.95–1.84 (m, 4H);  $^{13}C$  NMR (126 MHz,  $CDCl_3$ )  $\delta$  171.7, 150.6, 129.4, 125.8, 121.5, 34.0, 24.3.

#### Periodic light ON/OFF switching experiment in the cobalt-catalysed Giese reaction (Fig. 1)

In a Schlenk tube equipped with a stirring bar, Mn powder (41.2 mg, 0.75 mmol) was added and heated at 400 °C for 3 min under vacuum. After cooling and filling with argon, the tube was

charged with  $Et_3N \cdot HCl$  (51.6 mg, 0.38 mmol) and  $VB_{12}$  (16.9 mg, 12.5  $\mu$ mol), evacuated and filled with argon for three times. After the addition of dry DMF (1.5 mL), trimethylsilyl chloride (*ca.* 6  $\mu$ L), and mesitylene (12.0 mg, 0.15 mmol) as an internal standard (IS), the reaction mixture was stirred for 10 min. The tube was attached to an argon-filled balloon and moved to a light-tight room. Then, dodecyl tosylate (**1a**, 85.1 mg, 0.25 mmol) and methyl acrylate (**2a**, 32.3 mg, 0.38 mmol) were added into the solution. The light ON/OFF experiment was conducted under a periodic photo-irradiation with the photo-reaction apparatus mentioned above at the ambient temperature for 14 h. The product yield was determined by GC measurement of an aliquot of the reaction mixture in each reaction time. The Giese reaction under dark conditions was carried out in the same procedure.

#### Time-dependent UV-Vis absorption spectra of Me-Cbl under the blue-light irradiation (Fig. 2)

The UV-Vis absorption spectra methylcobalamin (Me-Cbl) in DMF ( $10^{-4}$  mol  $L^{-1}$ ) were recorded on the Shimadzu UV-3600 spectrometer under argon atmosphere in a light-tight room, as depicted in Fig. S4.† Usually, cobalamin(III) and (II) species exhibit red and orange colors in solution, respectively.<sup>6</sup> Before the blue-light irradiation (0 min), the solution showed red color and a strong absorption peak at 522 nm (red line). As the solution was exposed to the blue-light, the above-mentioned peak decreased and a new peak at 475 nm appeared. After the blue-light irradiation for 190 min, the color of the solution turned to orange. The results would indicate that the blue-light irradiation of Me-Cbl affords a Co(II) complex with methyl radical through the cobalt-carbon bond cleavage.

#### Radical trapping reaction with $\gamma$ -terpinene (eqn (2) in Scheme 2)

In a Schlenk tube equipped with a stirring bar, Mn powder (41.2 mg, 0.75 mmol) was added and heated at 400 °C for 3 min under vacuum. After cooling and filling with argon, the tube was added vitamin  $B_{12}$  ( $VB_{12}$ ; 16.9 mg, 12.5  $\mu$ mol), evacuated and filled with argon three times. After adding dry DMF (1.5 mL) and trimethylsilyl chloride (*ca.* 6  $\mu$ L), the mixture was stirred for 5 min. Then, dodecyl tosylate (**1a**, 85.1 mg, 0.25 mmol) was added to the solution. The reaction mixture was stirred at room temperature for 16 h under blue-light irradiation. The obtained mixture was diluted with ethyl acetate and quenched by saturated aqueous  $NH_4Cl$ . The product **7** was characterised by an authentic sample. The yield was evaluated by GC measurement using mesitylene as an internal standard.

## Conflicts of interest

There are no conflicts to declare.

## Acknowledgements

This work was supported by JSPS KAKENHI (grant number: JP15K05106).

## Notes and references

- (a) C. K. Prier, D. A. Rankic and D. W. C. MacMillan, *Chem. Rev.*, 2013, **113**, 5322–5363; (b) M. Yan, J. C. Lo, J. T. Edwards and P. S. Baran, *J. Am. Chem. Soc.*, 2016, **138**, 12692–12714; (c) J. Choi and G. C. Fu, *Science*, 2017, **356**, 152–160.
- (a) G. S. C. Srikanth and S. L. Castle, *Tetrahedron*, 2005, **61**, 10377–10441; (b) J. Streuff and A. Gansäuer, *Angew. Chem., Int. Ed.*, 2015, **54**, 14232–14242; (c) G. J. Lovinger and J. P. Morken, *Eur. J. Org. Chem.*, 2020, 2362–2368.
- B. Giese, J. A. González-Gómez and T. Witzel, *Angew. Chem., Int. Ed.*, 1984, **23**, 69–70.
- M. Ballestri, C. Chatgililoglu, K. B. Clark, D. Griller, B. Giese and B. Kopping, *J. Org. Chem.*, 1991, **56**, 678–683.
- (a) R. Scheffold, S. Abrecht, R. Orłinski, H. R. Ruf, P. Stamouli, O. Tinembart, L. Walder and C. Weymuth, *Pure Appl. Chem.*, 1987, **59**, 363–372; (b) H. Gong, R. S. Andrews, J. L. Zuccarello, S. J. Lee and M. R. Gagné, *Org. Lett.*, 2009, **11**, 879–882; (c) X. Wu, W. Hao, K.-Y. Ye, B. Jiang, G. Pombar, Z. Song and S. Lin, *J. Am. Chem. Soc.*, 2018, **140**, 14836–14843.
- (a) G. Sorin, R. Martinez Mallorquin, Y. Contie, A. Baralle, M. Malacria, J. Goddard and L. Fensterbank, *Angew. Chem., Int. Ed.*, 2010, **49**, 8721–8723; (b) V. Corcé, L. Chamoreau, E. Derat, J. Goddard, C. Ollivier and L. Fensterbank, *Angew. Chem., Int. Ed.*, 2015, **54**, 11414–11418.
- (a) L. Chu, C. Ohta, Z. Zuo and D. W. C. MacMillan, *J. Am. Chem. Soc.*, 2014, **136**, 10886–10889; (b) T. Qin, L. R. Malins, J. T. Edwards, R. R. Merchant, A. J. E. Novak, J. Z. Zhong, R. B. Mills, M. Yan, C. Yuan, M. D. Eastgate and P. S. Baran, *Angew. Chem., Int. Ed.*, 2017, **56**, 260–265; (c) N. P. Ramirez and J. C. Gonzalez-Gomez, *Eur. J. Org. Chem.*, 2017, 2154–2163.
- J. Wu, P. S. Grant, X. Li, A. Noble and V. K. Aggarwal, *Angew. Chem., Int. Ed.*, 2019, **58**, 5697–5701.
- (a) A. Hu, J.-J. Guo, H. Pan, H. Tang, Z. Gao and Z. Zuo, *J. Am. Chem. Soc.*, 2018, **140**, 1612–1616; (b) I. Kim, B. Park, G. Kang, J. Kim, H. Jung, H. Lee, M. Baik and S. Hong, *Angew. Chem., Int. Ed.*, 2018, **57**, 15517–15522; (c) Y. Zhu, K. Huang, J. Pan, X. Qiu, X. Luo, Q. Qin, J. Wei, X. Wen, L. Zhang and N. Jiao, *Nat. Commun.*, 2018, **9**, 2625; (d) C. Karmel, B. Li and J. F. Hartwig, *J. Am. Chem. Soc.*, 2018, **140**, 1460–1470.
- H. Togo, S. Matsubayashi, O. Yamazaki and M. Yokoyama, *J. Org. Chem.*, 2000, **65**, 2816–2819.
- G. L. Lackner, K. W. Quasdorf and L. E. Overman, *J. Am. Chem. Soc.*, 2013, **135**, 15342–15345.
- Y. Ye, H. Chen, J. L. Sessler and H. Gong, *J. Am. Chem. Soc.*, 2019, **141**, 820–824.
- C. C. Nawrat, C. R. Jamison, Y. Slutskyy, D. W. C. MacMillan and L. E. Overman, *J. Am. Chem. Soc.*, 2015, **137**, 11270–11273.
- (a) X. Zheng, X. J. Dai, H. Q. Yuan, C. X. Ye, J. Ma and P. Q. Huang, *Angew. Chem., Int. Ed.*, 2013, **52**, 3494–3498; (b) T. Suga, S. Shimazu and Y. Ukaji, *Org. Lett.*, 2018, **20**, 5389–5392; (c) H. Xie, J. Guo, Y.-Q. Wang, K. Wang, P. Guo, P.-F. Su, X. Wang and X.-Z. Shu, *J. Am. Chem. Soc.*, 2020, **142**, 16787–16794.
- (a) K. Komeyama, R. Ohata, S. Kiguchi and I. Osaka, *Chem. Commun.*, 2017, **53**, 6401–6404; (b) K. Komeyama, Y. Yamahata and I. Osaka, *Org. Lett.*, 2018, **20**, 4375–4378.
- (a) K. Komeyama, T. Michiyuki and I. Osaka, *ACS Catal.*, 2019, **9**, 9285–9291; (b) K. Komeyama, R. Tsunemitsu, T. Michiyuki, H. Yoshida and I. Osaka, *Molecules*, 2019, **24**, 1458.
- T. Michiyuki, I. Osaka and K. Komeyama, *Chem. Commun.*, 2020, **56**, 1247–1250.
- (a) M. E. Weiss, L. M. Kreis, A. Lauber and E. M. Carreira, *Angew. Chem., Int. Ed.*, 2011, **50**, 11125–11128; (b) G. P. Cerai and B. Morandi, *Chem. Commun.*, 2016, **52**, 9769–9772; (c) M. Ociepa, O. Baka, J. Narodowicz and D. Gryko, *Adv. Synth. Catal.*, 2017, **359**, 3560–3565; (d) M. J. Hossain, T. Ono, K. Wakiya and Y. Hisaeda, *Chem. Commun.*, 2017, **53**, 10878–10881; (e) L. Chen, Y. Kametani, K. Imamura, T. Abe, Y. Shiota, K. Yoshizawa, Y. Hisaeda and H. Shimakoshi, *Chem. Commun.*, 2019, **55**, 13070–13073; (f) L. Chen, Y. Hisaeda and H. Shimakoshi, *Adv. Synth. Catal.*, 2019, **361**, 2877–2884; (g) M. J. Hossain, T. Ono, Y. Yano and Y. Hisaeda, *ChemElectroChem*, 2019, **6**, 4199–4203; (h) M. Ociepa, A. J. Wierzba, J. Turkowska and D. Gryko, *J. Am. Chem. Soc.*, 2020, **142**, 5355–5361.
- S. Izumi, H. Shimakoshi, M. Abe and Y. Hisaeda, *Dalton Trans.*, 2010, **39**, 3302–3307.
- T. W. Greene and P. G. M. Wuts, *Protective Groups in Organic Synthesis*, Wiley, New York, 3rd edn, 1999.
- K. C. MacLeod, B. O. Patrick and K. M. Smith, *Organometallics*, 2010, **29**, 6639–6641.
- G. N. Schrauzer and R. J. Windgassen, *J. Am. Chem. Soc.*, 1967, **89**, 1999–2007.
- Y. Shi, P. C. J. Kamer, D. J. Cole-Hamilton, M. Harvie, E. F. Baxter, K. J. C. Lim and P. Pogorzelec, *Chem. Sci.*, 2017, **8**, 6911–6917.