


 Cite this: *RSC Adv.*, 2022, 12, 2564

# Singular value decomposition analysis of the electron density changes occurring upon electrostatic polarization of water

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In-depth elucidation of how molecules are electrically polarized would be one key factor for understanding the properties of those molecules under various thermodynamic and/or spatial conditions. Here this problem is tackled for the case of hydrogen-bonded water by conducting singular value decomposition of the electron density changes that occur upon electrostatic polarization. It is shown that all those electron density changes are approximately described as linear combinations of ten orthonormal basis "vectors". One main component is the interatomic charge transfer through each OH bond, while some others are characterized as the atomic dipolar polarizations, meaning that both of these components are important for the electrostatic polarization of water. The interaction parameters that reasonably well reproduce the induced dipole moments are derived, which indicate the extent of mixing of the two components in electrostatic polarization.

Received 4th September 2021

Accepted 12th January 2022

DOI: 10.1039/d1ra06649h

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## 1. Introduction

Molecules in condensed phases are electrically polarized more or less by intermolecular interactions according to the polarity and hydrogen-bonding conditions of the surrounding environment. Such polarization often profoundly affects the system's properties. In the case of water, it is well-known that the dielectric constant changes significantly with temperature and pressure in the thermodynamic region near the critical point,<sup>1</sup> resulting in continuously varying solubility of a wide range of solutes<sup>2</sup> and thus extensive utility as a medium for various chemical reactions.<sup>3</sup> It has been recognized that, to reproduce this situation computationally in molecular dynamics (MD) simulations, it is required to treat the molecules' electric polarization explicitly;<sup>4–7</sup> fixed-charge models that include the effects of polarization in an average way are not sufficient for this purpose. Similar situations arise for water molecules on material interfaces or confined in small spaces.<sup>8,9</sup>

There are many ways that have been proposed up to now to explicitly treat electric polarization of water. They are roughly categorized as (1) atomic dipolar polarization (ADP) model,<sup>10–20</sup> (2) Drude oscillator (or shell) model,<sup>21–24</sup> and (3) fluctuating

charge (FC) model.<sup>25–28</sup> These are different in the representation of molecular electric polarization; the polarization is completely or essentially confined within each atom with the atomic partial charges being invariant in the ADP and Drude oscillator models, while it is represented only by the interatomic charge transfer (CT) within a molecule without any atomic polarization incorporated in the FC model. In other words, the electronic responses supposed in these models are totally different. It is thus expected that these models have different effects on short- and medium-ranged intermolecular interactions, even though they may all similarly be effective in reproducing the molecular induced dipole. Recently a few combined models that take into account both the atomic dipolar polarizabilities and interatomic CT have been proposed,<sup>29–32</sup> but the relative importance of these two factors, which is essential for correct modeling, is not sufficiently well clarified. Although it somewhat depends on the way of defining the boundaries of atoms in a molecule, analysis in this relation will be of great help in deepening our understanding on the polarization-induced behavior of electrons based on our chemical intuition. With respect to the interaction energies, inclusion of the polarization effects with many-body potential functions<sup>19,33,34</sup> is also a powerful way, but explicit analysis of the behavior of electrons (as charged particles) is needed for better understanding on observable electric properties such as a dipole moment and infrared intensities.

In the present study, this problem is tackled by singular value decomposition (SVD) analysis<sup>35</sup> of the electron density changes occurring upon electrostatic polarization. The electron density changes are digitally expressed as arrays of the values at sufficiently fine three-dimensional grid points, and those obtained in a number of different electrostatic situations may be

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analyzed statistically. Aside from a few applications to classical or quantum dynamics simulations and to spectral analysis for the purpose of decomposing a complex high-dimensional object into some low-dimensional components,<sup>36–41</sup> SVD has been extensively used for extracting important factors in image processing,<sup>42–44</sup> so that it is considered to be a suitable method of analysis in the present context when each array of electron density changes (calculated for each specified configuration) is regarded as an image. However, unlike usual image processing, each array of electron density changes has three ( $x$ ,  $y$ , and  $z$ ) spatial dimensions, so that the preferable number of grid points to represent them tends to be large. The present study shows how SVD is practically effective in analyzing the polarization-induced electron density changes.

## 2. Computational procedure

### 2.1. Calculations of electron density changes upon electrostatic polarization

Calculations were done mainly for the hydrogen-bonded pairs of water molecules in the (water)<sub>90</sub> clusters (this cluster size is not a magic number in a scientific sense but is a rough upper limit of the size that can be treated with the available computational resources). Firstly, 13 different initial structures of the (water)<sub>90</sub> clusters were extracted<sup>45</sup> from the snapshots of MD simulations carried out for the liquid system of 1024 molecules using the TIP4P<sup>46</sup> potential model at the temperature of mp +25 K (257 K),<sup>47</sup> and then the structures were optimized at the B3LYP/6-31+G(2df,p) level of density functional theory (DFT). Hydrogen-bonded pairs of molecules were taken with the criterion of  $r(\text{O}\cdots\text{H}) \leq 2.5 \text{ \AA}$  and  $\theta(\text{O}\cdots\text{H}-\text{O}) \geq 90^\circ$ . There were 2035 pairs of molecules in total taken in this way. Then, after transformation of coordinates using the molecular axes of each molecule in the pairs, and symmetry operation so that the hydrogen-bond counterpart is located in the specified  $xy$  quadrant ( $x > 0$  and  $y > 0$  for the hydrogen-bond donating molecules,  $x < 0$  and  $y < 0$  for the hydrogen-bond accepting molecules), with the molecular  $C_2$  axis being taken along the  $z$  direction, 500 mutually different (to the greatest extent) configurations each of hydrogen-bond donor and acceptor were extracted and were used in the subsequent analyses. The spatial distributions of such configurations are shown in Fig. 1.

The electron densities  $\rho^{(\text{el})}(\mathbf{r})$  of electrostatically polarized water molecules were calculated by replacing the hydrogen-bond donor or acceptor molecule with a set of atomic partial charges. Here, a fifth of the CHelpG<sup>48</sup> charges ( $-0.1462764e$  for O and  $0.0731382e$  for H) was employed to limit the systems within a weak perturbation regime. Then, the electron density without placing the atomic partial charges was subtracted to derive the changes upon electrostatic polarization. In calculating the electron densities, an augmented basis set of aug-cc-pVDZ<sup>49</sup> was used for better representation of the polarization of hydrogen atoms. The evaluation points  $\mathbf{r}$  of  $\rho^{(\text{el})}(\mathbf{r})$  were taken in a rectangular box of  $10.0 \times 11.6 \times 10.6 \text{ \AA}^3$ , so that (similarly to our previous studies of electron density analysis<sup>45,50–56</sup>) each boundary of the box is at least  $5 \text{ \AA}$  from any atom in the molecule, with the interval of  $0.02 \text{ \AA}$ . As a result,

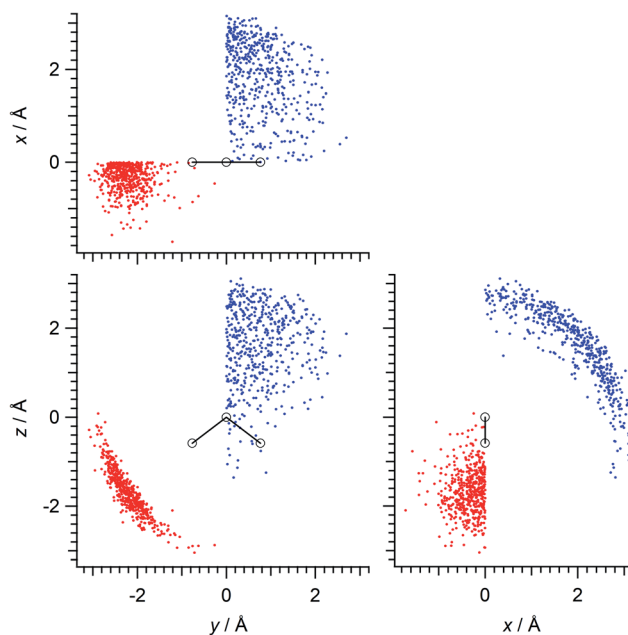


Fig. 1 Spatial distributions of the hydrogen-bond donating (blue) and accepting (red) molecules (represented by their oxygen atoms) that are included in the SVD analysis as the source of electrostatic polarization of the water molecules (shown superimposed at the center) in the (water)<sub>90</sub> clusters.

the number of the evaluation points  $\mathbf{r}$  of each  $\rho^{(\text{el})}(\mathbf{r})$  is 154 564 011 ( $=501 \times 581 \times 531$ ).

In addition, a supplementary analysis was carried out for a water molecule under dipolar electrostatic perturbation from the H atom side of the OH bond. Locating this H atom at the origin and taking the OH bond along the  $y$  axis (the O atom on the  $+y$  side), electric charges of  $-0.15e$  and  $+0.15e$  (which were made similar to those adopted above) were placed at  $r = \xi$  and  $\xi + 1.0 \text{ \AA}$ , respectively, of  $(0, -r \cos \theta, r \sin \theta)$  or  $(r \sin \theta, -r \cos \theta, 0)$ , with  $\xi = 2.0, 2.4$  or  $2.8 \text{ \AA}$  and  $\theta = 0^\circ$  or  $\pm 30^\circ$  (as a result, there are 15 configurations in total). In this case, the evaluation points  $\mathbf{r}$  of  $\rho^{(\text{el})}(\mathbf{r})$  were taken in a rectangular box of  $10.0 \times 15.6 \times 15.5 \text{ \AA}^3$  with the interval of  $0.02 \text{ \AA}$ .

The DFT calculations described above were carried out by using the Gaussian 09 program,<sup>57</sup> and the analyses before and after those DFT calculations were done with our own original programs.

### 2.2. Singular value decomposition analysis

Each of the polarization-induced electron density changes calculated in this way for the water molecules in the (water)<sub>90</sub> clusters is regarded as a 154 564 011-dimensional vector (denoted as  $\mathbf{b}_j$ ), and there are 1000 such vectors ( $j = 1, 2, \dots, 1000$ ). The matrix  $\mathbf{B}$  consisting of all these vectors (of the size of  $154\,564\,011 \times 1000$ ) was subject to the SVD analysis described below.

In the SVD analysis,<sup>35</sup>  $\mathbf{B}$  is decomposed as

$$\mathbf{B} = \mathbf{U}\mathbf{\Lambda}\mathbf{V} \quad (1)$$

where  $\mathbf{U}$  and  $\mathbf{V}$  are  $m \times m$  and  $n \times n$  orthogonal matrices ( $m = 154\ 564\ 011$  and  $n = 1000$ ),  $\mathbf{\Lambda}$  is an  $m \times n$  rectangular diagonal matrix [whose (non-negative) diagonal elements are denoted as  $\lambda_k$  ( $k = 1, 2, \dots, n$ )], and the upper left superscript  $\text{t}$  stands for transposing of the matrix. Since we have

$${}^t\mathbf{B}\mathbf{B} = (\mathbf{V}^t\mathbf{\Lambda}^t\mathbf{U})(\mathbf{U}\mathbf{\Lambda}^t\mathbf{V}) = \mathbf{V}({}^t\mathbf{\Lambda}\mathbf{\Lambda})^t\mathbf{V} \quad (2)$$

where  ${}^t\mathbf{\Lambda}\mathbf{\Lambda}$  is an  $n \times n$  diagonal matrix with the diagonal elements of  $\lambda_k^2$  ( $k = 1, 2, \dots, n$ ), it is possible to obtain matrices  $\mathbf{V}$  and  ${}^t\mathbf{\Lambda}\mathbf{\Lambda}$  by diagonalization of  ${}^t\mathbf{B}\mathbf{B}$ , which is the matrix of correlation among vectors  $\mathbf{b}_j$  ( $j = 1, 2, \dots, n$ ), as

$$({}^t\mathbf{B}\mathbf{B})\mathbf{v}_k = \lambda_k^2\mathbf{v}_k \quad (3)$$

where  $\mathbf{v}_k$  ( $k = 1, 2, \dots, n$ ) is the  $k$ th column vector constituting matrix  $\mathbf{V}$ . From eqn (1) we have

$$\mathbf{B}\mathbf{V} = (\mathbf{U}\mathbf{\Lambda}^t\mathbf{V})\mathbf{V} = \mathbf{U}\mathbf{\Lambda} \quad (4)$$

Then, the first  $n$  column vectors constituting matrix  $\mathbf{U}$  [denoted as  $\mathbf{u}_k$  ( $k = 1, 2, \dots, n$ )] are obtained as

$$\mathbf{u}_k = \frac{1}{\lambda_k}\mathbf{B}\mathbf{v}_k \quad (5)$$

By using  $\lambda_k$ ,  $\mathbf{u}_k$ , and  $\mathbf{v}_k$ , eqn (1) may also be expressed as

$$\mathbf{B} = \sum_{k=1}^n \lambda_k \mathbf{u}_k {}^t\mathbf{v}_k \quad (6)$$

Eqn (5) may be regarded as representing the transformation of the sets of vectors from  $\{\mathbf{b}_j\}$  ( $j = 1, 2, \dots, n$ , representing the original set of polarization-induced electron density changes) to  $\{\mathbf{u}_k\}$  ( $k = 1, 2, \dots, n$ , representing the important elements of polarization-induced electron density changes). Then, representing the electrostatic situations (electrostatic potentials and electric fields on the atoms) corresponding to  $\mathbf{b}_j$  as  $\mathbf{d}_j$  ( $j = 1, 2, \dots, n$ ), and the matrix consisting of them as  $\mathbf{D}$ , the electrostatic situations corresponding to  $\mathbf{u}_k$  are obtained as

$$\mathbf{s}_k = \frac{1}{\lambda_k}\mathbf{D}\mathbf{v}_k \quad (7)$$

In fact, the above procedure describes the first-round analysis of the electron density changes. The theoretical treatments after that, which are motivated by the results, are described below in Section 3.1.

The polarization-induced electron density changes calculated in the supplementary analysis were treated in a similar way. In this case, SVD was conducted for fifteen 303 634 056-dimensional vectors  $\mathbf{b}_j$  ( $j = 1, 2, \dots, 15$ ).

## 3. Results and discussion

### 3.1. SVD analysis of electron density changes

The eigenvalues corresponding to the first 50 eigenvectors obtained from the SVD analysis are shown in Fig. 2. It is seen that the eigenvalue decreases rather rapidly, by a factor of 100 in the first 6 eigenvectors, by another factor of 100 in the next 11

eigenvectors, and so on. This result suggests that it is possible to extract a set of a few important electron density changes that occur upon electrostatic polarization. The one- and two-dimensional plots of the electron density changes for the first and second eigenvectors are shown in Fig. 3. The electron density change corresponding to the first eigenvector shown in Fig. 3a mainly consists of the interatomic CT through the OH bond on the left-hand side, as clearly recognized from the large negative feature throughout the molecule of the one-dimensional running integral (light blue curve) along the  $y$  axis [the molecular in-plane axis taken perpendicular to the  $z$  ( $C_2$  molecular) axis] shown on the lower-left side. Inspecting the  $yz$  two-dimensional plot, it may also be recognized that the electron density change in the spatial region of the oxygen atom is not of a simple spherical character but rather of a quadrupolar nature. A similar quadrupolar nature of electron density modulation is also seen for the carbon atom of the C–Br bond involved in halogen bonding upon electrostatic polarization.<sup>55</sup> In contrast, the electron density change corresponding to the second eigenvector shown in Fig. 3b mainly consists of the out-of-plane dipolar polarization of the oxygen atom, with a small anti-phase inner core polarization, but the interatomic CT through the OH bond on the right-hand side is also mixed to a certain extent.

In fact, considering the structural symmetry of the water molecule, if there is an interatomic CT component through the OH bond on the left-hand side, there should also be the counterpart component through the OH bond on the right-hand side. Asymmetric distribution of the hydrogen-bond donors and acceptors adopted in the present study as shown in Fig. 1 is most suitable for obtaining localized vectors of electron density changes, but they should be partially symmetrized, handled additionally with symmetry operation, and made mutually orthogonal to obtain electron density changes that may be regarded as truly elementary components, which will be hereafter called “basis vectors”.

For example, the first basis vector (called basis vector **1a**) was derived with the following procedure: (1) symmetrize the primary component obtained from the SVD analysis (shown in Fig. 3a) with respect to the  $yz$  plane; (2) conduct a symmetry operation on it with respect to the  $xz$  plane to obtain another vector; (3) calculate the normalized mutually orthogonal symmetry-related linear combinations of these two. The electron density changes thus obtained (the interatomic CT component for the OH bond on the left-hand side) is shown in Fig. 4a. The symmetry-related counterpart (not shown) is called basis vector **1b**, which is the mirror image with respect to the  $xz$  plane.

After extracting these basis vectors, projections of vectors  $\mathbf{u}_k$  onto the basis vectors were calculated and were subtracted from  $\mathbf{u}_k$  to represent the “remaining” components of the polarization-induced electron density changes (the resulting vectors, which are orthogonal to the basis vectors, are denoted as  $\mathbf{u}'_k$ ), and the matrix of correlation among those vectors (denoted as  ${}^t\mathbf{B}'\mathbf{B}'$ ) was reconstructed as

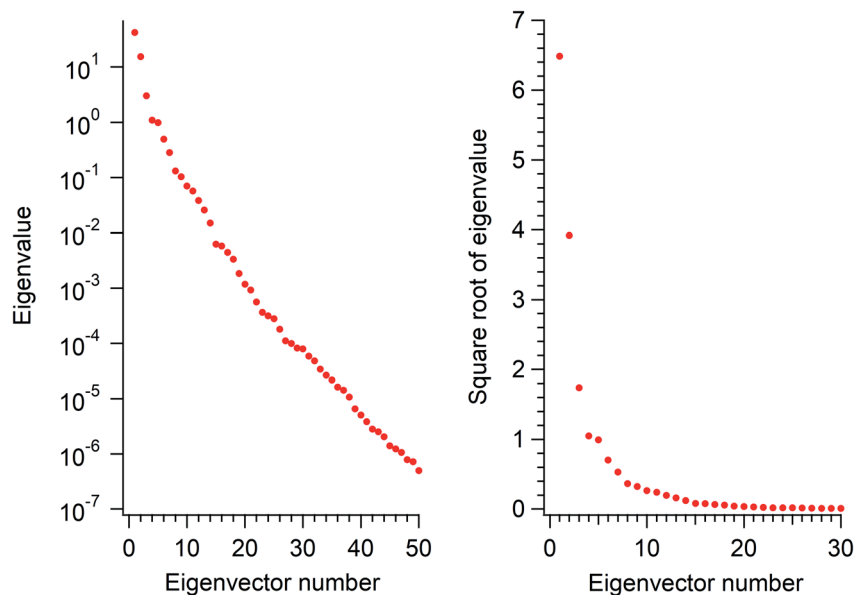


Fig. 2 Eigenvalues (logarithmic scale, left panel) and their square roots (linear scale, right panel) obtained from the SVD analysis of the polarization-induced electron density changes of the water molecules in the (water)<sub>90</sub> clusters.

$$\mathbf{B}' = \sum_{k=1}^{n'} \lambda_k \mathbf{u}'_k \mathbf{v}'_k \quad (8)$$

$$\mathbf{B}'\mathbf{B}' = \sum_{j,k=1}^{n'} \lambda_j \lambda_k \mathbf{v}_j (\mathbf{u}'_j \mathbf{u}'_k)^t \mathbf{v}_k \quad (9)$$

Here the summation was truncated to  $n' = 32$  (numbered from larger eigenvalues) as an approximation. In eqn (9), the scalar products  $\mathbf{u}'_j \mathbf{u}'_k$  were calculated first for computational efficiency. Then, this  $\mathbf{B}'\mathbf{B}'$  was diagonalized to derive a new set of eigenvalues and eigenvectors. The vector of the largest eigenvalue this time was mainly related to the polarization of the O atom along the  $x$  axis, so that only symmetrization and

normalization were conducted. The resultant basis vector (called basis vector 2) represents the truly out-of-plane (along the  $x$  axis) dipolar polarization component of the oxygen atom as shown in Fig. 4e.

The remaining basis vectors (3, 4a, 5, 6a, and 7) shown in Fig. 4 and the symmetry-related ones (4b, and 6b) were obtained by repeating these procedures, which are summarized as a diagram in Chart 1. Owing to the asymmetric distribution of the hydrogen-bond donors and acceptors included in the present analysis as shown in Fig. 1, these localized and partially symmetrized basis vectors have successfully been obtained. The main characters of these may be described as rather skewed dipolar polarizations of the oxygen atom along the  $z$  and  $y$  axes (vectors 3 and 5), in-plane and out-of-plane vertical (to the bond)

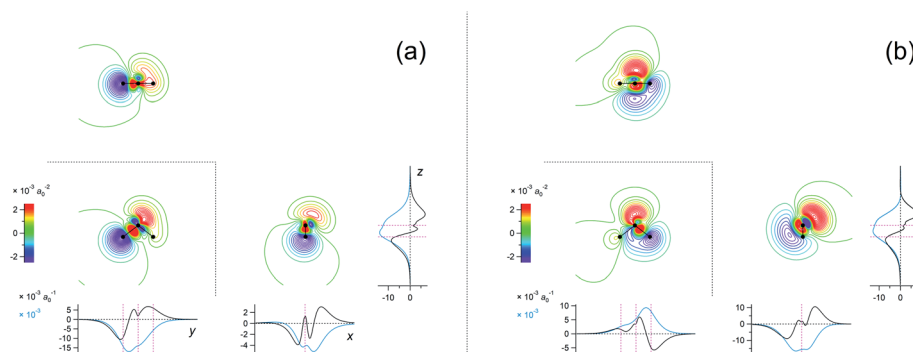


Fig. 3 Two-dimensional ( $xy$ ,  $xz$ , and  $yz$ ) contour plots of  $\int dz\rho^{(el)}(\mathbf{r})$ ,  $\int dy\rho^{(el)}(\mathbf{r})$ , and  $\int dx\rho^{(el)}(\mathbf{r})$ , one-dimensional ( $x$ ,  $y$ , and  $z$ ) plots of  $\iint dydz\rho^{(el)}(\mathbf{r})$ ,  $\iint dx dz\rho^{(el)}(\mathbf{r})$ , and  $\iint dx dy\rho^{(el)}(\mathbf{r})$  (black curves), and their running integrals along the respective axis (light blue curves) of the electron density changes corresponding to the (a) first and (b) second eigenvectors obtained from the SVD analysis. The contours in the two-dimensional plot are drawn with the interval of  $0.36 \times 10^{-3} a_0^{-2}$  in the range from  $-9 \times 10^{-3} a_0^{-2}$  to  $9 \times 10^{-3} a_0^{-2}$ , with the color code shown on the left-hand side. The atomic positions of the central water molecule are indicated with black filled circles in the two-dimensional plots and with pink dotted vertical lines in the one-dimensional plots. The coordinate system is shown in Fig. 1.

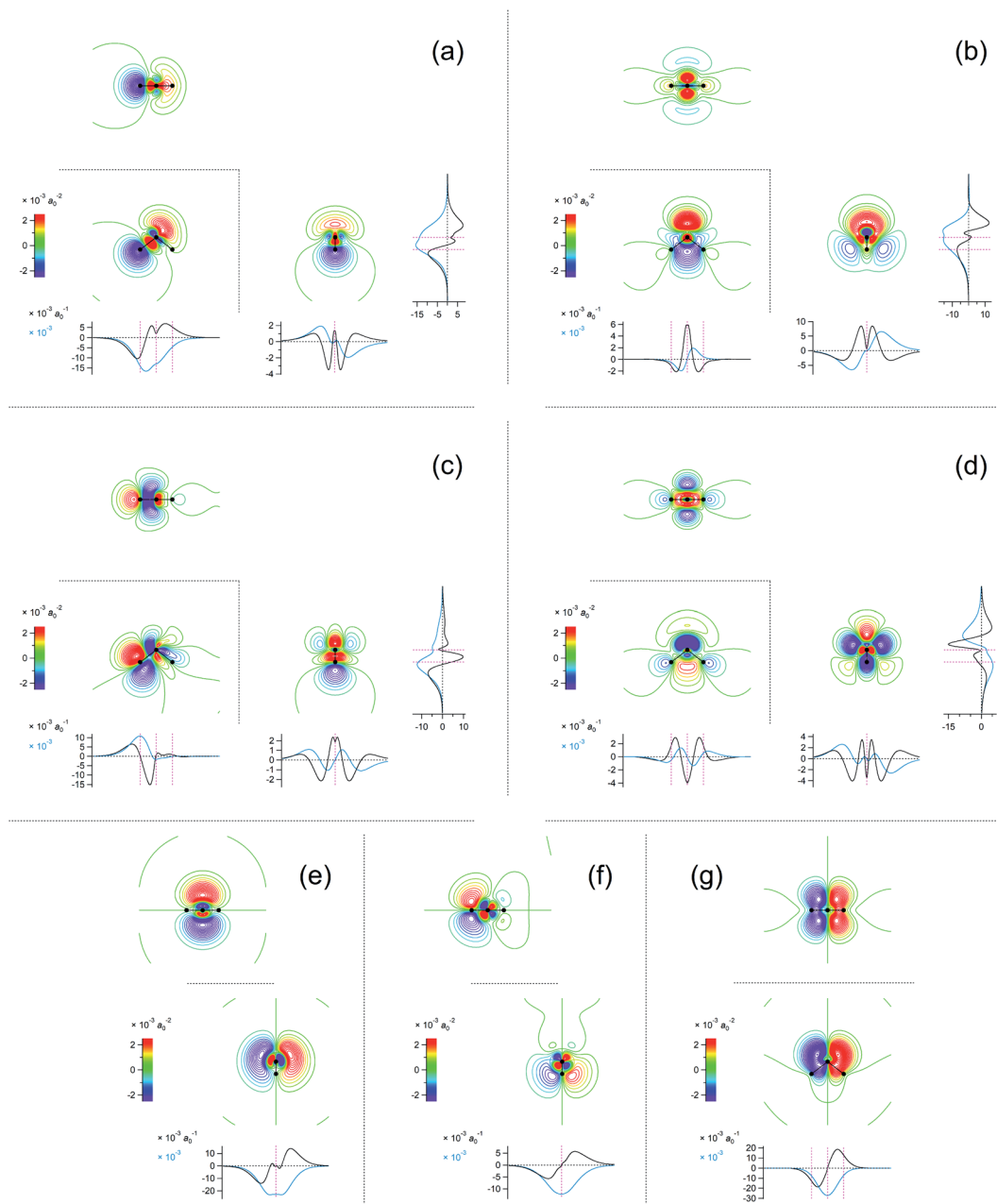


Fig. 4 Two-dimensional ( $xy$ ,  $xz$ , and  $yz$ ) contour plots of  $\int dz\rho^{(el)}(\mathbf{r})$ ,  $\int dy\rho^{(el)}(\mathbf{r})$ , and  $\int dx\rho^{(el)}(\mathbf{r})$ , one-dimensional ( $x$ ,  $y$ , and  $z$ ) plots of  $\iint dydz\rho^{(el)}(\mathbf{r})$ ,  $\iint dxz\rho^{(el)}(\mathbf{r})$ , and  $\iint dxdy\rho^{(el)}(\mathbf{r})$  (black curves), and their running integrals along the respective axis (light blue curves) of the basis vectors of polarization-induced electron density changes of the water molecules in the  $(\text{water})_{90}$  clusters: (a) vector **1a**, (b) vector **3**, (c) vector **6a**, (d) vector **7**, (e) vector **2**, (f) vector **4a**, and (g) vector **5** (some of the plots are omitted in the last three panels because of the symmetry). The contours in the two-dimensional plots are drawn with the interval of  $0.36 \times 10^{-3} a_0^{-2}$  in the range from  $-14.4 \times 10^{-3} a_0^{-2}$  to  $14.4 \times 10^{-3} a_0^{-2}$ , with the color code shown on the left-hand side. The atomic positions of the central water molecule are indicated with black filled circles in the two-dimensional plots and with pink dotted vertical lines in the one-dimensional plots. The coordinate system is shown in Fig. 1.

dipolar polarizations of one of the hydrogen atoms (vectors **6a** and **4a**) as well as their symmetry-related counterparts (vectors **6b** and **4b**), and a higher-order component with a relatively small induced dipole (vector **7**). Judging from the eigenvalues of the SVD analysis, the set of these ten basis vectors covers nearly 99% of the electron density changes occurring upon electrostatic polarizations between hydrogen-bonded water molecules. In other words, the electron density changes occurring upon

electrostatic polarization of water are nearly ten-dimensional, and the ten basis vectors described above (**1a**, **1b**, **2**, **3**, **4a**, **4b**, **5**, **6a**, **6b**, and **7**) constitute an orthonormal basis of the ten-dimensional polarization “space”.

### 3.2. Electrostatic properties and induced dipole moments

The electrostatic properties and the induced dipole moments related to these basis vectors are summarized in Table 1. Then,

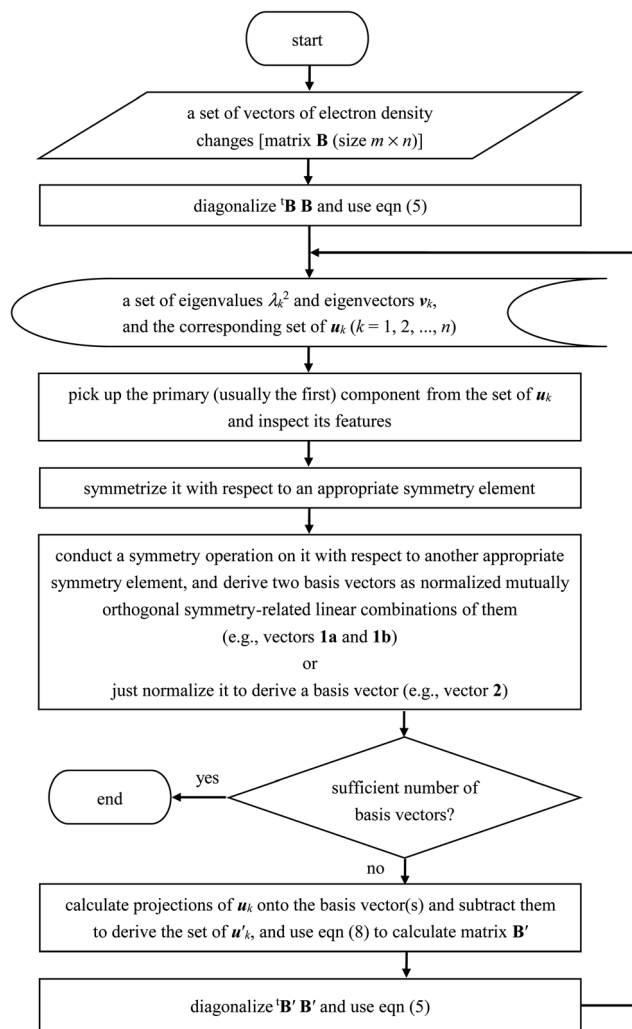


Chart 1 Diagram of the procedures of the SVD analysis of electron density changes adopted in the present study.

how can we set up and derive the electrostatic interaction parameters that are usable in MD simulations? In the present case, the out-of-plane components are rather easy to handle; since there are three basis vectors (**2**, **4a**, and **4b**), with the main characters being described as the atomic dipolar polarizations, it would be most reasonable to place dipolar polarizabilities ( $\alpha_{xx}$ ) on the three atomic sites and derive their values by linear algebra. With regard to the in-plane components, it is needed to identify the electrostatic properties related to the interatomic CT components.

For this purpose, a separate supplementary set of calculations was conducted for the situations of dipolar electrostatic perturbations mimicking the effect of a hydrogen-bond acceptor as described in the last part of Section 2.1. The result is shown in Fig. 5. It is clearly seen that, integrating along the two directions perpendicular to the OH bond on the left-hand side (to derive the one-dimensional plot along the  $y$  axis in this figure), the electron density changes obtained for the electrostatic perturbations with different angular settings are nearly invariant after scaling with the electrostatic potential

difference between the H and O atoms. As discussed above in relation to the results shown in Fig. 3a and 4a, the component shown in Fig. 5 is characterized as the interatomic CT through the OH bond. Thus, one may conclude that it is the electrostatic potential difference between the H and O atoms rather than the electric field on the H atom along the OH bond that is most effective for the interatomic CT components, in a way similar to the through-bond polarization discussed for halogen-bonding systems<sup>55</sup> and for hydrogen fluoride,<sup>56</sup> and in accord with the idea of the charge response kernel.<sup>58,59</sup>

As a result, in the present case, four interaction parameters [dipolar polarizabilities  $\alpha_{yy}$  and  $\alpha_{zz}$  on O,  $\alpha_{\perp}$  (vertical to the bond) on H, and the charge response parameter  $\kappa_{\text{OH}}$ ] are selected, and their values are obtained by least-squares fitting. In this least-squares fitting, the square root of the eigenvalue related to each basis vector in the respective cycle of the repeated SVD analyses was used as the relative weight. The values of the parameters thus obtained are shown in Table 2. By using these values, the induced dipole moments corresponding to the basis vectors are reasonably well reproduced as shown in Table 3, although there is a noticeable deviation with regard to the angle of the induced dipole corresponding to basis vector **6a** (by  $36.3^\circ$ ).

The values of the parameters shown in Table 2 indicate the extent of mixing of the atomic dipolar polarization component and the interatomic CT component in the molecular electric polarization of water. For example, when the uniform electric field of  $0.04 E_{\text{h}} e^{-1} a_0^{-1}$  is operating along the out-of-plane ( $x$ ) axis, the induced dipole moment arises totally from the atomic dipolar polarizations, which amounts to  $0.31 ea_0$  in total. In contrast, when the uniform electric field of  $0.04 E_{\text{h}} e^{-1} a_0^{-1}$  is operating along one of the OH bond, the interatomic CT along this bond (related to  $\kappa_{\text{OH}}$ ) gives rise to  $0.19 ea_0$  of induced dipole moment, while the atomic dipolar polarizations of the oxygen (related to  $\alpha_{yy}$  and  $\alpha_{zz}$ ) and the other hydrogen atom ( $\alpha_{\perp}$ ) gives rise to  $0.16 ea_0$  in almost the same direction, with a small almost perpendicular induced dipole of  $0.05 ea_0$  due to the interatomic CT related to this other H atom. When the electric field is non-uniform, the extent of mixing of these components would vary. Since the values of the parameters shown in Table 2 are derived for the hydrogen-bonding pairs of molecules, it is expected that they are more pertinent to those non-uniform situations.

### 3.3. Discussion toward application to larger systems

To apply the present method to larger systems, some technical issues need to be considered. First, the number of the evaluation points  $r$  of each  $\rho^{(\text{el})}(r)$  increases with the system size, but not proportionally. This is because the evaluation points are taken in a rectangular box, each boundary of which is set as being at least  $5 \text{ \AA}$  away from any atom in the system. For example, if we suppose a system with the frame of the atomic centers fitting in a cube of  $10 \times 10 \times 10 \text{ \AA}^3$ , the rectangular box of the evaluation points will be of the size of  $20 \times 20 \times 20 \text{ \AA}^3$ . If we retain  $0.02 \text{ \AA}$  for the interval, the number of the evaluation points will be  $m = 10^9$  (more exactly,  $1001^3$ ), which is only about 6.5 times of that for the case of water examined in the present

Table 1 The electrostatic properties and the induced dipole moments related to the electron density basis vectors of water

Atom <sup>a</sup> /molecule	Property <sup>b</sup>	Electron density basis vector						
		1a	2	3	4a	5	6a	7
<b>Electrostatic properties</b>								
O	$E_x$	0	-1.357	0	-0.242	0	0	0
	$E_y$	-0.394	0	0	0	-1.297	0.269	0
	$E_z$	-0.508	0	-1.152	0	0	-0.116	0.028
H <sub>1</sub>	$\Delta\phi_{\text{OH}}$	-1.812	0	-0.445	0	-0.234	0.502	0.037
	$E_x$	0	-0.553	0	-2.275	0	0	0
	$E_y$	-1.268	0	0.679	0	0.302	2.159	-0.978
H <sub>2</sub>	$E_z$	-0.966	0	0.085	0	0.783	-2.676	0.063
	$\Delta\phi_{\text{OH}}$	-0.037	0	-0.445	0	0.234	-0.115	0.037
	$E_x$	0	-0.553	0	0.121	0	0	0
	$E_y$	-0.127	0	-0.679	0	0.302	0.119	0.978
	$E_z$	-0.273	0	0.085	0	-0.783	0.134	0.063
	<b>Induced dipole moment</b>							
Molecule	$\mu_x$	0	-8.201	0	-4.312	0	0	0
	$\mu_y$	-5.111	0	0	0	-6.134	1.931	0
	$\mu_z$	-4.950	0	-5.074	0	0	-3.576	-0.158

<sup>a</sup> H<sub>1</sub> and H<sub>2</sub> are located on the  $-y$  and  $+y$  side. The coordinate system is shown in Fig. 1. <sup>b</sup> In units of  $10^{-2} E_{\text{h}} e^{-1}$  for the electrostatic potential (the difference from the value at the oxygen atom,  $\Delta\phi_{\text{OH}}$ ),  $10^{-2} E_{\text{h}} e^{-1} a_0^{-1}$  for the electric field ( $E_x, E_y, E_z$ ), and  $10^{-2} e a_0$  for the dipole moment ( $\mu_x, \mu_y, \mu_z$ ).

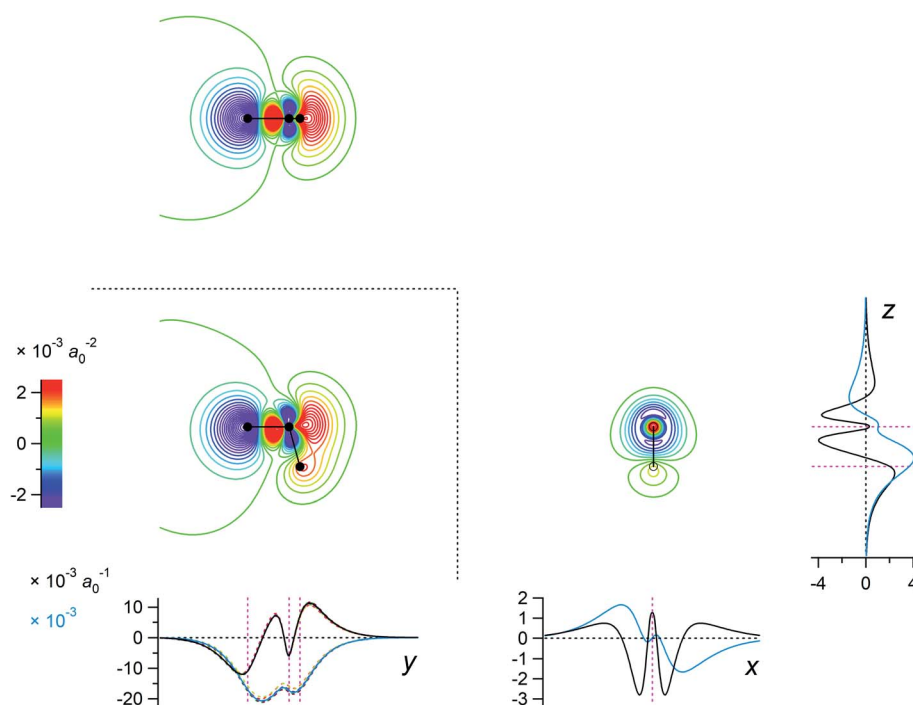


Fig. 5 Two-dimensional ( $xy$ ,  $xz$ , and  $yz$ ) contour plots of  $\int dz\rho^{(\text{el})}(\mathbf{r})$ ,  $\int dy\rho^{(\text{el})}(\mathbf{r})$ , and  $\int dx\rho^{(\text{el})}(\mathbf{r})$ , one-dimensional ( $x$ ,  $y$ , and  $z$ ) plots of  $\iint dydz\rho^{(\text{el})}(\mathbf{r})$ ,  $\iint dx dz\rho^{(\text{el})}(\mathbf{r})$ , and  $\iint dx dy\rho^{(\text{el})}(\mathbf{r})$  (black curves), and their running integrals along the respective axis (light blue curves) of the primary component obtained from the supplementary SVD analysis conducted for the electron density changes induced on a water molecule by the dipolar electrostatic perturbation mimicking the effect of a hydrogen-bond acceptor [*i.e.*, from the left-hand ( $-y$ ) side in the figure, as described in the last part of Section 2.1]. In the one-dimensional plot on the lower-left side (along the  $y$  axis), the plots of some important elements [with  $\xi = 2.0$ , 2.4 and 2.8 Å and  $\theta = 0^\circ$  (red, purple, and dark red dotted lines), and with  $\xi = 2.0$  Å and  $\theta = \pm 30^\circ$  (yellow, green, and blue dotted lines)] included in the SVD analysis are overlapped with scaling based on the electrostatic potential difference between the primarily perturbed H and O atoms [located along the  $y$  (horizontal) axis]. The contours in the two-dimensional plot are drawn with the interval of  $0.36 \times 10^{-3} a_0^{-2}$  in the range from  $-9 \times 10^{-3} a_0^{-2}$  to  $9 \times 10^{-3} a_0^{-2}$ , with the color code shown on the left-hand side. The atomic positions of the central water molecule are indicated with black filled circles in the two-dimensional plots and with pink dotted vertical lines in the one-dimensional plots.

**Table 2** Values of the parameters related to the electrostatic polarization of water

Atom/bond	Property	Value <sup>a</sup>
O	$\alpha_{xx}$	4.856
	$\alpha_{yy}$	3.339
	$\alpha_{zz}$	3.158
H	$\alpha_{xx}$	1.457
	$\alpha_{\perp}$	0.823
OH	$\kappa_{OH}$	2.605

<sup>a</sup> In units of  $e^2 a_0^2 E_h^{-1}$  for the atomic polarizability ( $\alpha_{xx}$ ,  $\alpha_{yy}$ ,  $\alpha_{zz}$ ,  $\alpha_{\perp}$ ) and  $e^2 a_0 E_h^{-1}$  for the charge response parameter ( $\kappa_{OH}$ ).

**Table 3** Estimated and reference values of the induced dipole moments of water

Electron density basis vector	Property	Value <sup>a</sup>	
		Estimated	Reference
<b>1a</b>	$\mu_y$	-5.142	-5.111
	$\mu_z$	-4.713	-4.950
<b>2</b>	$\mu_x$	-8.201	-8.201
<b>3</b>	$\mu_z$	-5.492	-5.074
<b>4a</b>	$\mu_x$	-4.312	-4.312
<b>5</b>	$\mu_y$	-5.737	-6.134
<b>6a</b>	$\mu_y$	3.980	1.931
	$\mu_z$	-1.889	-3.576
<b>7</b>	$\mu_z$	0.111	-0.158

<sup>a</sup> In units of  $10^{-2} ea_0$ .

study. Widening of the interval is not considered to be desirable. For example, in the 2D plot on the  $yz$  plane of vector **1a** shown in Fig. 4a, the plotted value changes significantly within the range of 0.1 Å around the center of the oxygen atom. The computational cost is considered to scale just linearly to the number of the evaluation points. The size of matrix  ${}^t\mathbf{BB}$  defined in eqn (2), which is subject to diagonalization in the SVD analysis, is determined by the number of vectors  $\mathbf{b}_j$  ( $n = 1000$  in the present case), not by their dimension. In this sense, the computational time needed for matrix diagonalization depends nonlinearly to the number of vectors  $\mathbf{b}_j$ , but diagonalization of a matrix of the size of  $1000 \times 1000$ , for example, is not a problem with the current computational power. More importantly, the size of matrix  $\mathbf{B}$  itself, which is equal to the dimension multiplied by the number of vectors  $\mathbf{b}_j$ , affects the disk space needed to store it. In the case of the present study ( $m = 154\ 564\ 011$  and  $n = 1000$ , double precision floating point numbers), it amounted to  $1.24 \times 10^{12}$  bytes. The number of vectors  $\mathbf{b}_j$  depends on the extent of distortion, *etc.*, of the hydrogen-bond configurations taken into account in the analysis. If we narrow the range of them, the weight of the main components derived from the analysis will increase.

Secondly, it is expected that the eigenvalues decay more slowly (as compared to the situation shown in Fig. 2) for larger systems and, hence, the number of the basis vectors required to describe the system will increase rather proportionally to the

system size. It means that the procedures described in the present study (summarized in Chart 1) will be needed to be repeated more times. To clarify the situations in this relation, however, further studies on some other typical functional groups will be needed.

## 4. Concluding remarks

In summary, it is shown by conducting SVD analysis on the electron density changes that both the interatomic CT through the OH bonds and the dipolar polarizations of the atoms are important for the molecular electrostatic polarization of hydrogen-bonded water. Taking only one of these factors is not sufficient to correctly represent it. The seven basis vectors (shown in Fig. 4) and the symmetry-related three basis vectors that are extracted in this SVD analysis constitute an orthonormal basis of the polarization “space” that covers nearly 99% of all the electron density changes. The parameters of the interatomic CT and the atomic dipolar polarizations derived from the present analysis (shown in Table 2) reasonably well reproduce the molecular induced dipole moments as shown in Table 3. As a result, these parameters indicate the extent of mixing and the relative importance of the two factors of molecular electrostatic polarization.

In modeling electric polarizations of molecules, it has sometimes been suggested that the use of point dipoles is too simple, especially for describing the interactions between closely-located molecules, so that the use of distributed charges is more appropriate.<sup>17,18</sup> In this sense, the electron density basis vectors derived in the present study demonstrate the spatial characteristics of the charge distributions related to the electrostatic polarizations, and thus are considered to form a basis for further developing sophisticated polarizable models of water.

## Conflicts of interest

There are no conflicts to declare.

## Acknowledgements

This study was supported in part by JSPS KAKENHI Grant Number JP20H05215. The calculations were carried out partly by using the computers at the Research Center for Computational Science, Okazaki, Japan.

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