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## 6,7-Dimethoxy-3-methoxycarbonyl-1-(2-methoxyphenyl)-3,4-dihydroisoquinoline 2-oxide

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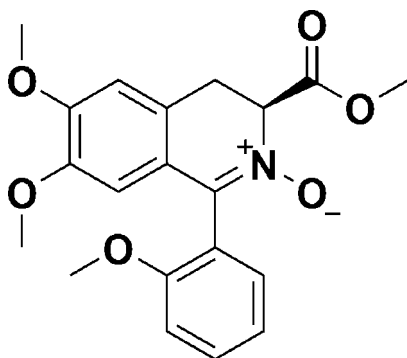
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Key indicators: single-crystal X-ray study;  $T = 173$  K; mean  $\sigma(\text{C}-\text{C}) = 0.002$  Å;  $R$  factor = 0.053;  $wR$  factor = 0.157; data-to-parameter ratio = 16.2.

In the title compound,  $\text{C}_{20}\text{H}_{21}\text{NO}_6$ , an  $N$ -oxide-based organocatalyst, the  $N$ -containing six-membered ring adopts a twisted half-chair conformation. No hydrogen bonding or  $\pi$ - $\pi$  stacking was found within the crystal structure.

### Related literature

For related structures, see: Naicker *et al.* (2010, 2011).



### Experimental

#### Crystal data

$\text{C}_{20}\text{H}_{21}\text{NO}_6$   
 $M_r = 371.38$   
 Monoclinic,  $P2_1/n$   
 $a = 5.4765$  (1) Å  
 $b = 21.9984$  (6) Å  
 $c = 15.0007$  (4) Å  
 $\beta = 92.774$  (2)°  
 $V = 1805.08$  (8) Å<sup>3</sup>  
 $Z = 4$   
 Mo  $K\alpha$  radiation  
 $\mu = 0.10$  mm<sup>-1</sup>  
 $T = 173$  K  
 $0.25 \times 0.18 \times 0.15$  mm

#### Data collection

Bruker APEXII diffractometer  
 7815 measured reflections  
 3959 independent reflections  
 3092 reflections with  $I > 2\sigma(I)$   
 $R_{\text{int}} = 0.016$

#### Refinement

$R[F^2 > 2\sigma(F^2)] = 0.053$   
 $wR(F^2) = 0.157$   
 $S = 1.06$   
 3959 reflections  
 245 parameters  
 H-atom parameters constrained  
 $\Delta\rho_{\text{max}} = 0.92$  e Å<sup>-3</sup>  
 $\Delta\rho_{\text{min}} = -0.33$  e Å<sup>-3</sup>

Data collection: *APEX2* (Bruker, 2006); cell refinement: *SAINT* (Bruker, 2006); data reduction: *SAINT*; program(s) used to solve structure: *SHELXS97* (Sheldrick, 2008); program(s) used to refine structure: *SHELXL97* (Sheldrick, 2008); molecular graphics: *OLEX2* (Dolomanov *et al.*, 2009); software used to prepare material for publication: *SHELXL97*.

The authors wish to thank Dr Hong Su of the Chemistry Department, University of Cape Town, for her assistance with the crystallographic data collection.

Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: HG5031).

### References

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**supplementary materials**

*Acta Cryst.* (2011). E67, o1352 [ doi:10.1107/S1600536811016539 ]

## 6,7-Dimethoxy-3-methoxycarbonyl-1-(2-methoxyphenyl)-3,4-dihydroisoquinoline 2-oxide

T. Naicker, T. Govender, H. G. Kruger and G. E. M. Maguire

### Comment

The title compound is a novel *N*-oxide based catalyst containing a tetrahydroisoquinoline (TIQ) backbone. This is the first X-ray crystal structure report of this type of organocatalyst within this TIQ class of molecules. This compound and its derivatives are currently being tested in our laboratory as novel organocatalysts for asymmetric allylation reactions (Naicker *et al.* 2010).

From the crystal structure it is evident that the *N*-containing six membered ring assumes a twisted half chair conformation (Fig. 1). This differs from a similar structure that we recently reported which displays a twisted half boat conformation (Naicker *et al.* 2011). A possible reason for the change is the introduction of the oxygen atom O2 onto the sp<sup>2</sup> hybridized nitrogen atom. All our previous examples have either a hydrogen or methyl group at that position.

Interestingly there is no classic hydrogen bonding within the crystal packing however, there are various intermolecular and intramolecular short contact interactions that link the molecules together within the crystal lattice. The *N*-oxide oxygen O2 displays two potential hydrogen bond interactions to C16—H16 and C20A—H20A which are 3.36 Å and 3.34 Å respectively. These interactions result in a layered packing within the crystal structure as shown along the (100) axis in Fig. 2. A centroid distance of 7.156 Å indicated that there is no  $\pi$ - $\pi$  stacking within the crystal matrix.

### Experimental

(*S*)-methyl 6,7-dimethoxy-1-(2-methoxyphenyl)-3,4-dihydroisoquinoline-3-carboxylate (1.30 g, 3.7 mmol) was dissolved in dry methylene chloride (50 ml). Potassium carbonate (1.0 g, 7.5 mmol) was added and the reaction cooled to  $-78$  °C. Meta-chloroperbenzoic acid (0.86 g of 75% pure, net 0.65 g, 3.7 mmol) was then added, and the reaction was allowed to stir at  $-78$  °C for 3 h. At this time, the reaction was allowed to warm to room temperature. After stirring for a further 2 h at room temperature, methylene chloride (50 ml) was added to dilute the reaction and celite (500 mg) was added to aid filtration. The reaction was filtered, and the methylene chloride concentrated to dryness affording the crude product which was purified by column chromatography (methylene chloride:methanol, 99:1,  $R_f = 1/5$ ) (1.20 g, 87% yield).

Melting point 423 K.  $[\alpha]_D^{20}$  5.128 (c 0.13 in CHCl<sub>3</sub>).

IR (neat): 2923, 1742, 1508, 1285, 729 cm<sup>-1</sup>.

<sup>1</sup>H NMR (400 MHz, CDCl<sub>3</sub>)  $\delta$  7.55 (dd,  $J = 7.5, 1.7$  Hz, 1H), 7.51 – 7.42 (m, 1H), 7.20 – 6.97 (m, 3H), 6.74 (d,  $J = 6.5$  Hz, 1H), 6.27 (d,  $J = 13.6$  Hz, 1H), 4.94 (dt,  $J = 5.8, 2.9$  Hz, 1H), 3.90 (d,  $J = 1.8$  Hz, 4H), 3.77 (d,  $J = 4.3$  Hz, 3H), 3.73 – 3.54 (m, 6H), 3.49 – 3.33 (m, 1H).

## supplementary materials

$^{13}\text{C}$  NMR (101 MHz,  $\text{CDCl}_3$ )  $\delta$  168.74, 157.23, 149.52, 148.17, 131.24, 131.12, 130.43, 122.89, 121.93, 121.49, 121.01, 120.80, 119.71, 112.05, 111.50, 110.64, 110.60, 110.08, 109.50, 77.34, 77.03, 76.71, 70.84, 70.72, 56.17, 56.11, 56.05, 55.93, 55.62, 53.24, 53.10, 30.80, 30.73.

Recrystallization from ethyl acetate at room temperature afforded colourless crystals suitable for X-ray analysis.

### Refinement

All non-hydrogen atoms were refined anisotropically. The hydrogen atoms were placed in idealized positions in a riding model with  $U_{\text{iso}}$  set at 1.2 or 1.5 times those of their parent atoms (1.2 for tertiary C—H, secondary C—H<sub>2</sub> and aromatic C—H or N—H groups and 1.5 for methyl C—H<sub>3</sub>) and fixed C—H bond lengths (e.g. 0.88 Å for N—H and others ranging from 0.95 Å to 1.00 Å).

### Figures

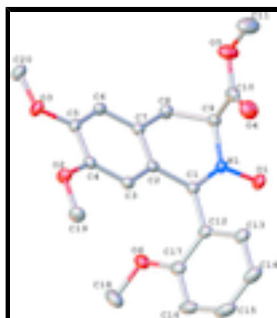


Fig. 1. The molecular structure of the title compound with the atom numbering scheme. Displacement ellipsoids are drawn at the 40% probability level. Hydrogen atoms have been omitted for clarity.

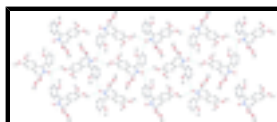


Fig. 2. A partial projection of the title compound, viewed along [100] plane.

### 6,7-dimethoxy-3-(methoxycarbonyl)-1-(2-methoxyphenyl)-3,4-dihydroisoquinolin-2-ium-2-olate

#### Crystal data

$\text{C}_{20}\text{H}_{21}\text{NO}_6$

$M_r = 371.38$

Monoclinic,  $P2_1/n$

Hall symbol: -P 2yn

$a = 5.4765$  (1) Å

$b = 21.9984$  (6) Å

$c = 15.0007$  (4) Å

$\beta = 92.774$  (2)°

$V = 1805.08$  (8) Å<sup>3</sup>

$Z = 4$

$F(000) = 784$

$D_x = 1.367$  Mg m<sup>-3</sup>

Melting point: 423 K

Mo  $K\alpha$  radiation,  $\lambda = 0.71073$  Å

Cell parameters from 7815 reflections

$\theta = 2.3$ – $27.1$ °

$\mu = 0.10$  mm<sup>-1</sup>

$T = 173$  K

Block, colourless

$0.25 \times 0.18 \times 0.15$  mm

*Data collection*

Bruker APEXII diffractometer	3092 reflections with $I > 2\sigma(I)$
Radiation source: fine-focus sealed tube	$R_{\text{int}} = 0.016$
graphite	$\theta_{\text{max}} = 27.1^\circ$ , $\theta_{\text{min}} = 2.3^\circ$
1.2° $\varphi$ and $\omega$ scans	$h = -7 \rightarrow 7$
7815 measured reflections	$k = -28 \rightarrow 28$
3959 independent reflections	$l = -19 \rightarrow 19$

*Refinement*

Refinement on $F^2$	Secondary atom site location: difference Fourier map
Least-squares matrix: full	Hydrogen site location: inferred from neighbouring sites
$R[F^2 > 2\sigma(F^2)] = 0.053$	H-atom parameters constrained
$wR(F^2) = 0.157$	$w = 1/[\sigma^2(F_o^2) + (0.080P)^2 + 1.0977P]$
$S = 1.06$	where $P = (F_o^2 + 2F_c^2)/3$
3959 reflections	$(\Delta/\sigma)_{\text{max}} < 0.001$
245 parameters	$\Delta\rho_{\text{max}} = 0.92 \text{ e } \text{Å}^{-3}$
0 restraints	$\Delta\rho_{\text{min}} = -0.33 \text{ e } \text{Å}^{-3}$
Primary atom site location: structure-invariant direct methods	Extinction correction: <i>SHELXL97</i> (Sheldrick, 2008), $F_c^* = kFc[1+0.001xFc^2\lambda^3/\sin(2\theta)]^{-1/4}$
	Extinction coefficient: 0.0055 (19)

*Special details*

**Experimental.** Half sphere of data collected using *COLLECT* strategy (Nonius, 2000). Crystal to detector distance = 30 mm; combination of  $\varphi$  and  $\omega$  scans of 1.0°, 40 s per °, 2 iterations.

**Geometry.** All e.s.d.'s (except the e.s.d. in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell e.s.d.'s are taken into account individually in the estimation of e.s.d.'s in distances, angles and torsion angles; correlations between e.s.d.'s in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell e.s.d.'s is used for estimating e.s.d.'s involving l.s. planes.

**Refinement.** Refinement of  $F^2$  against ALL reflections. The weighted  $R$ -factor  $wR$  and goodness of fit  $S$  are based on  $F^2$ , conventional  $R$ -factors  $R$  are based on  $F$ , with  $F$  set to zero for negative  $F^2$ . The threshold expression of  $F^2 > \sigma(F^2)$  is used only for calculating  $R$ -factors(gt) *etc.* and is not relevant to the choice of reflections for refinement.  $R$ -factors based on  $F^2$  are statistically about twice as large as those based on  $F$ , and  $R$ -factors based on ALL data will be even larger.

*Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters ( $\text{Å}^2$ )*

	$x$	$y$	$z$	$U_{\text{iso}}^*/U_{\text{eq}}$
O1	0.2176 (3)	0.22907 (6)	0.35593 (9)	0.0354 (3)
O2	0.6006 (2)	-0.03624 (6)	0.11617 (9)	0.0322 (3)
O3	0.2824 (3)	-0.10179 (6)	0.19301 (9)	0.0344 (3)
O4	0.5317 (3)	0.12825 (8)	0.45676 (10)	0.0441 (4)

## supplementary materials

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O5	0.2474 (3)	0.07631 (8)	0.52628 (10)	0.0485 (4)
O6	0.2093 (3)	0.16305 (7)	0.07924 (9)	0.0391 (4)
N1	0.2146 (3)	0.17212 (7)	0.33291 (10)	0.0251 (3)
C1	0.3125 (3)	0.14938 (8)	0.26149 (11)	0.0246 (4)
C2	0.3023 (3)	0.08366 (8)	0.24573 (11)	0.0228 (4)
C3	0.4608 (3)	0.05630 (8)	0.18643 (11)	0.0247 (4)
H3	0.5749	0.0806	0.1567	0.030*
C4	0.4517 (3)	-0.00558 (8)	0.17115 (11)	0.0246 (4)
C5	0.2793 (3)	-0.04178 (8)	0.21365 (12)	0.0260 (4)
C6	0.1259 (3)	-0.01473 (8)	0.27287 (12)	0.0262 (4)
H6	0.0115	-0.0390	0.3025	0.031*
C7	0.1370 (3)	0.04758 (8)	0.28959 (11)	0.0238 (4)
C8	-0.0255 (3)	0.07830 (8)	0.35355 (12)	0.0283 (4)
H8A	-0.0748	0.0488	0.3991	0.034*
H8B	-0.1752	0.0932	0.3210	0.034*
C9	0.1083 (3)	0.13129 (8)	0.39883 (12)	0.0259 (4)
H9	-0.0107	0.1550	0.4335	0.031*
C10	0.3260 (4)	0.11235 (9)	0.46255 (13)	0.0344 (5)
C11	0.4434 (5)	0.06056 (13)	0.59251 (17)	0.0572 (7)
H11A	0.3780	0.0341	0.6382	0.086*
H11B	0.5086	0.0978	0.6206	0.086*
H11C	0.5744	0.0393	0.5629	0.086*
C12	0.4564 (3)	0.19020 (8)	0.20530 (12)	0.0278 (4)
C13	0.6553 (3)	0.22127 (8)	0.24385 (14)	0.0314 (4)
H13	0.6881	0.2188	0.3065	0.038*
C14	0.8070 (4)	0.25591 (9)	0.19251 (17)	0.0410 (5)
H14	0.9427	0.2770	0.2195	0.049*
C15	0.7571 (4)	0.25908 (10)	0.10160 (17)	0.0455 (6)
H15	0.8616	0.2823	0.0660	0.055*
C16	0.5593 (4)	0.22943 (10)	0.06094 (15)	0.0418 (5)
H16	0.5270	0.2328	-0.0017	0.050*
C17	0.4075 (4)	0.19452 (8)	0.11258 (13)	0.0323 (4)
C18	0.1512 (5)	0.16733 (12)	-0.01499 (14)	0.0520 (6)
H18A	0.0059	0.1428	-0.0303	0.078*
H18B	0.2892	0.1523	-0.0479	0.078*
H18C	0.1187	0.2098	-0.0311	0.078*
C19	0.7699 (4)	-0.00092 (10)	0.06834 (13)	0.0351 (5)
H19A	0.8650	-0.0278	0.0311	0.053*
H19B	0.8806	0.0207	0.1108	0.053*
H19C	0.6803	0.0286	0.0303	0.053*
C20	0.0983 (4)	-0.13944 (9)	0.22980 (15)	0.0364 (5)
H20A	0.1191	-0.1815	0.2100	0.055*
H20B	-0.0640	-0.1248	0.2094	0.055*
H20C	0.1140	-0.1378	0.2951	0.055*

Atomic displacement parameters ( $\text{\AA}^2$ )

$U^{11}$

$U^{22}$

$U^{33}$

$U^{12}$

$U^{13}$

$U^{23}$

O1	0.0496 (8)	0.0207 (7)	0.0370 (7)	-0.0003 (6)	0.0121 (6)	-0.0065 (6)
O2	0.0374 (7)	0.0271 (7)	0.0335 (7)	0.0029 (6)	0.0143 (6)	-0.0035 (5)
O3	0.0425 (8)	0.0194 (6)	0.0422 (8)	-0.0021 (6)	0.0122 (6)	-0.0042 (6)
O4	0.0335 (8)	0.0543 (10)	0.0441 (9)	0.0016 (7)	-0.0020 (6)	-0.0001 (7)
O5	0.0505 (10)	0.0559 (10)	0.0386 (8)	-0.0023 (8)	-0.0030 (7)	0.0157 (7)
O6	0.0542 (9)	0.0366 (8)	0.0261 (7)	0.0019 (7)	-0.0020 (6)	0.0010 (6)
N1	0.0295 (8)	0.0213 (7)	0.0247 (7)	-0.0006 (6)	0.0040 (6)	-0.0014 (6)
C1	0.0292 (9)	0.0223 (8)	0.0223 (8)	0.0019 (7)	0.0024 (7)	-0.0001 (7)
C2	0.0256 (9)	0.0225 (8)	0.0202 (8)	0.0007 (7)	0.0006 (6)	0.0006 (6)
C3	0.0279 (9)	0.0240 (9)	0.0225 (8)	-0.0005 (7)	0.0038 (7)	0.0002 (7)
C4	0.0275 (9)	0.0256 (9)	0.0211 (8)	0.0031 (7)	0.0034 (7)	-0.0019 (7)
C5	0.0315 (9)	0.0203 (8)	0.0261 (9)	0.0002 (7)	-0.0003 (7)	-0.0012 (7)
C6	0.0277 (9)	0.0250 (9)	0.0260 (9)	-0.0032 (7)	0.0033 (7)	0.0002 (7)
C7	0.0251 (8)	0.0240 (9)	0.0225 (8)	0.0010 (7)	0.0016 (6)	-0.0010 (7)
C8	0.0286 (9)	0.0269 (9)	0.0299 (9)	-0.0014 (7)	0.0071 (7)	-0.0023 (7)
C9	0.0304 (9)	0.0247 (9)	0.0230 (9)	0.0025 (7)	0.0074 (7)	0.0003 (7)
C10	0.0522 (13)	0.0269 (10)	0.0245 (9)	0.0061 (9)	0.0068 (8)	-0.0033 (7)
C11	0.0544 (15)	0.0673 (17)	0.0485 (14)	-0.0081 (13)	-0.0119 (11)	0.0216 (13)
C12	0.0350 (10)	0.0195 (8)	0.0294 (9)	0.0038 (7)	0.0081 (7)	0.0010 (7)
C13	0.0341 (10)	0.0227 (9)	0.0379 (10)	0.0024 (7)	0.0063 (8)	0.0016 (8)
C14	0.0370 (11)	0.0267 (10)	0.0603 (14)	-0.0005 (8)	0.0124 (10)	0.0053 (9)
C15	0.0492 (13)	0.0313 (11)	0.0583 (14)	0.0040 (9)	0.0258 (11)	0.0149 (10)
C16	0.0611 (14)	0.0319 (11)	0.0339 (11)	0.0092 (10)	0.0179 (10)	0.0106 (9)
C17	0.0436 (11)	0.0242 (9)	0.0296 (10)	0.0070 (8)	0.0071 (8)	0.0017 (7)
C18	0.0755 (17)	0.0523 (14)	0.0275 (11)	0.0137 (13)	-0.0059 (11)	-0.0040 (10)
C19	0.0360 (11)	0.0378 (11)	0.0325 (10)	0.0011 (8)	0.0132 (8)	-0.0016 (8)
C20	0.0423 (11)	0.0236 (9)	0.0438 (12)	-0.0061 (8)	0.0059 (9)	-0.0007 (8)

*Geometric parameters (Å, °)*

O1—N1	1.2995 (19)	C8—H8B	0.9900
O2—C4	1.366 (2)	C9—C10	1.549 (3)
O2—C19	1.429 (2)	C9—H9	1.0000
O3—C5	1.356 (2)	C11—H11A	0.9800
O3—C20	1.436 (2)	C11—H11B	0.9800
O4—C10	1.187 (3)	C11—H11C	0.9800
O5—C10	1.329 (3)	C12—C13	1.388 (3)
O5—C11	1.468 (3)	C12—C17	1.407 (3)
O6—C17	1.362 (3)	C13—C14	1.388 (3)
O6—C18	1.437 (2)	C13—H13	0.9500
N1—C1	1.320 (2)	C14—C15	1.380 (3)
N1—C9	1.476 (2)	C14—H14	0.9500
C1—C2	1.466 (2)	C15—C16	1.381 (4)
C1—C12	1.484 (2)	C15—H15	0.9500
C2—C7	1.393 (2)	C16—C17	1.394 (3)
C2—C3	1.408 (2)	C16—H16	0.9500
C3—C4	1.381 (2)	C18—H18A	0.9800
C3—H3	0.9500	C18—H18B	0.9800
C4—C5	1.411 (3)	C18—H18C	0.9800

## supplementary materials

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C5—C6	1.386 (3)	C19—H19A	0.9800
C6—C7	1.394 (2)	C19—H19B	0.9800
C6—H6	0.9500	C19—H19C	0.9800
C7—C8	1.501 (2)	C20—H20A	0.9800
C8—C9	1.520 (2)	C20—H20B	0.9800
C8—H8A	0.9900	C20—H20C	0.9800
C4—O2—C19	117.10 (15)	O5—C11—H11A	109.5
C5—O3—C20	117.19 (15)	O5—C11—H11B	109.5
C10—O5—C11	112.08 (18)	H11A—C11—H11B	109.5
C17—O6—C18	117.57 (18)	O5—C11—H11C	109.5
O1—N1—C1	125.52 (15)	H11A—C11—H11C	109.5
O1—N1—C9	114.13 (14)	H11B—C11—H11C	109.5
C1—N1—C9	120.15 (15)	C13—C12—C17	119.13 (18)
N1—C1—C2	119.40 (16)	C13—C12—C1	119.27 (16)
N1—C1—C12	118.66 (15)	C17—C12—C1	121.42 (17)
C2—C1—C12	121.51 (15)	C12—C13—C14	121.2 (2)
C7—C2—C3	119.32 (16)	C12—C13—H13	119.4
C7—C2—C1	120.45 (16)	C14—C13—H13	119.4
C3—C2—C1	120.23 (16)	C15—C14—C13	118.7 (2)
C4—C3—C2	120.54 (16)	C15—C14—H14	120.6
C4—C3—H3	119.7	C13—C14—H14	120.6
C2—C3—H3	119.7	C14—C15—C16	121.8 (2)
O2—C4—C3	124.72 (16)	C14—C15—H15	119.1
O2—C4—C5	115.23 (15)	C16—C15—H15	119.1
C3—C4—C5	120.05 (16)	C15—C16—C17	119.4 (2)
O3—C5—C6	125.37 (16)	C15—C16—H16	120.3
O3—C5—C4	115.48 (16)	C17—C16—H16	120.3
C6—C5—C4	119.13 (16)	O6—C17—C16	124.20 (18)
C5—C6—C7	121.00 (16)	O6—C17—C12	116.05 (17)
C5—C6—H6	119.5	C16—C17—C12	119.8 (2)
C7—C6—H6	119.5	O6—C18—H18A	109.5
C2—C7—C6	119.92 (16)	O6—C18—H18B	109.5
C2—C7—C8	117.68 (15)	H18A—C18—H18B	109.5
C6—C7—C8	122.40 (16)	O6—C18—H18C	109.5
C7—C8—C9	110.07 (15)	H18A—C18—H18C	109.5
C7—C8—H8A	109.6	H18B—C18—H18C	109.5
C9—C8—H8A	109.6	O2—C19—H19A	109.5
C7—C8—H8B	109.6	O2—C19—H19B	109.5
C9—C8—H8B	109.6	H19A—C19—H19B	109.5
H8A—C8—H8B	108.2	O2—C19—H19C	109.5
N1—C9—C8	111.39 (14)	H19A—C19—H19C	109.5
N1—C9—C10	105.07 (14)	H19B—C19—H19C	109.5
C8—C9—C10	114.19 (15)	O3—C20—H20A	109.5
N1—C9—H9	108.7	O3—C20—H20B	109.5
C8—C9—H9	108.7	H20A—C20—H20B	109.5
C10—C9—H9	108.7	O3—C20—H20C	109.5
O4—C10—O5	124.7 (2)	H20A—C20—H20C	109.5
O4—C10—C9	125.50 (18)	H20B—C20—H20C	109.5
O5—C10—C9	109.75 (18)		



O1—N1—C1—C2	177.89 (16)	O1—N1—C9—C8	146.52 (15)
C9—N1—C1—C2	3.4 (2)	C1—N1—C9—C8	-38.4 (2)
O1—N1—C1—C12	5.3 (3)	O1—N1—C9—C10	-89.34 (17)
C9—N1—C1—C12	-169.18 (15)	C1—N1—C9—C10	85.72 (18)
N1—C1—C2—C7	17.8 (2)	C7—C8—C9—N1	51.46 (19)
C12—C1—C2—C7	-169.84 (16)	C7—C8—C9—C10	-67.4 (2)
N1—C1—C2—C3	-161.82 (16)	C11—O5—C10—O4	3.0 (3)
C12—C1—C2—C3	10.6 (3)	C11—O5—C10—C9	-175.79 (18)
C7—C2—C3—C4	0.5 (3)	N1—C9—C10—O4	-0.3 (3)
C1—C2—C3—C4	-179.88 (16)	C8—C9—C10—O4	122.1 (2)
C19—O2—C4—C3	-2.9 (3)	N1—C9—C10—O5	178.51 (15)
C19—O2—C4—C5	176.99 (16)	C8—C9—C10—O5	-59.1 (2)
C2—C3—C4—O2	-178.80 (16)	N1—C1—C12—C13	56.9 (2)
C2—C3—C4—C5	1.3 (3)	C2—C1—C12—C13	-115.55 (19)
C20—O3—C5—C6	5.5 (3)	N1—C1—C12—C17	-128.06 (19)
C20—O3—C5—C4	-175.62 (16)	C2—C1—C12—C17	59.5 (2)
O2—C4—C5—O3	-1.0 (2)	C17—C12—C13—C14	-0.6 (3)
C3—C4—C5—O3	178.87 (16)	C1—C12—C13—C14	174.57 (17)
O2—C4—C5—C6	177.95 (15)	C12—C13—C14—C15	0.0 (3)
C3—C4—C5—C6	-2.2 (3)	C13—C14—C15—C16	0.8 (3)
O3—C5—C6—C7	-179.98 (17)	C14—C15—C16—C17	-1.0 (3)
C4—C5—C6—C7	1.2 (3)	C18—O6—C17—C16	-1.3 (3)
C3—C2—C7—C6	-1.5 (2)	C18—O6—C17—C12	178.88 (18)
C1—C2—C7—C6	178.87 (16)	C15—C16—C17—O6	-179.39 (19)
C3—C2—C7—C8	178.87 (15)	C15—C16—C17—C12	0.4 (3)
C1—C2—C7—C8	-0.7 (2)	C13—C12—C17—O6	-179.81 (16)
C5—C6—C7—C2	0.7 (3)	C1—C12—C17—O6	5.1 (3)
C5—C6—C7—C8	-179.74 (16)	C13—C12—C17—C16	0.4 (3)
C2—C7—C8—C9	-33.2 (2)	C1—C12—C17—C16	-174.66 (17)
C6—C7—C8—C9	147.19 (17)		

Fig. 1

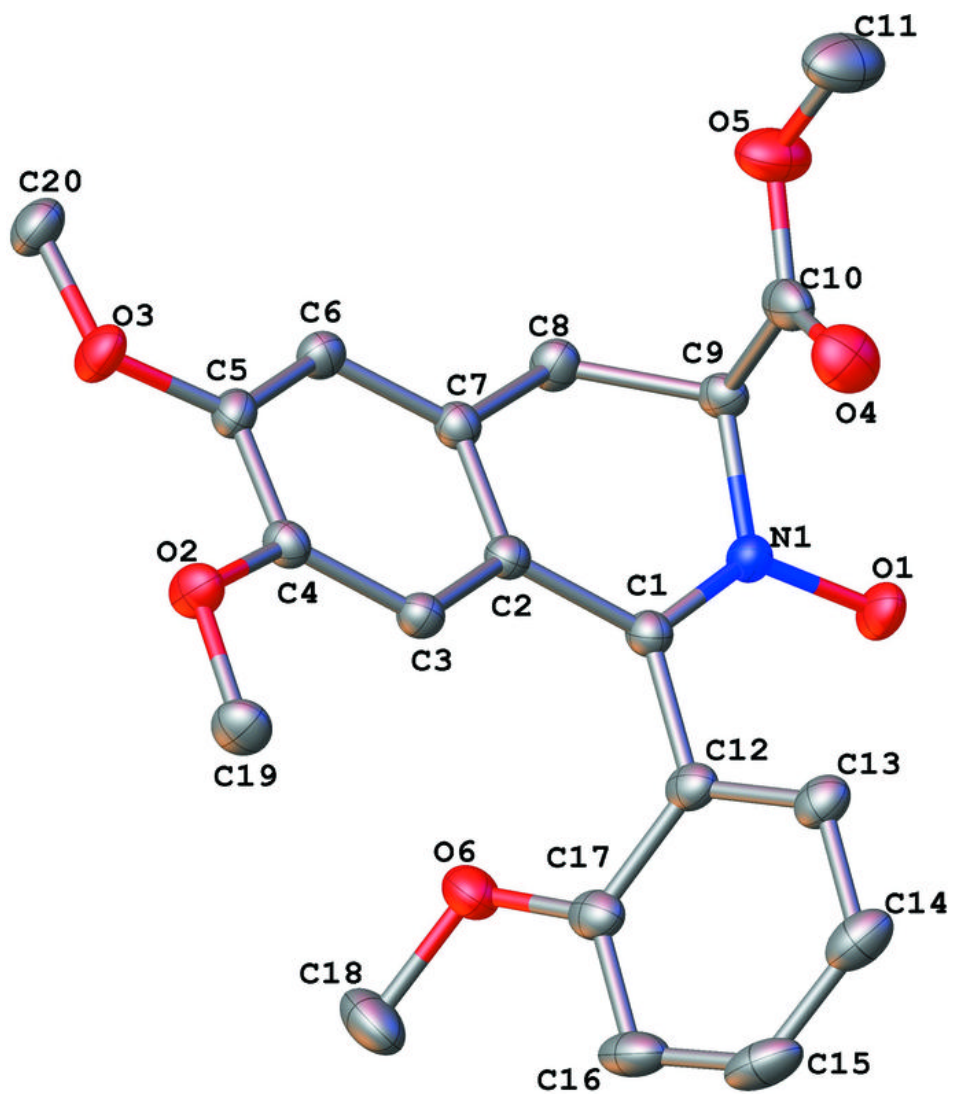


Fig. 2

