

catena-Poly[[diaquamanganese(II)]- μ -7-oxabicyclo[2.2.1]heptane-2,3-dicarboxylato]

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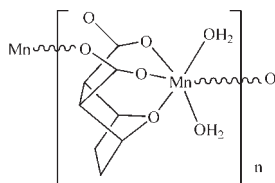
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Key indicators: single-crystal X-ray study; $T = 296$ K; mean $\sigma(\text{C}-\text{C}) = 0.003$ Å; R factor = 0.024; wR factor = 0.057; data-to-parameter ratio = 15.3.

In the title polymer, $[\text{Mn}(\text{C}_8\text{H}_8\text{O}_5)(\text{H}_2\text{O})_2]_n$, the Mn^{II} atom is in a distorted octahedral coordination mode, binding to the bridging O atom of the bicycloheptane unit, two O atoms from corresponding carboxylate groups, one carboxylate O atom from a symmetry-related bridging ligand and two water O atoms. This arrangement leads to the formation of polymeric chains propagating parallel to [001]. The crystal structure is stabilized by several O—H...O hydrogen-bonding interactions involving the coordinated water molecules as donors and the carboxylate O atoms as acceptors.

Related literature

7-Oxabicyclo[2.2.1]heptane-2,3-dicarboxylic anhydride (norcantharidin) is a lower toxicity anticancer drug, see: Shimi *et al.* (1982). For the preparation of disodium demethylcantharate, see: Yin *et al.* (2003).



Experimental

Crystal data

$[\text{Mn}(\text{C}_8\text{H}_8\text{O}_5)(\text{H}_2\text{O})_2]$	$V = 2062.04$ (11) Å ³
$M_r = 275.12$	$Z = 8$
Orthorhombic, <i>Iba</i> 2	Mo $K\alpha$ radiation
$a = 10.3513$ (2) Å	$\mu = 1.30$ mm ⁻¹
$b = 18.9903$ (4) Å	$T = 296$ K
$c = 10.4899$ (5) Å	$0.37 \times 0.22 \times 0.14$ mm

Data collection

Bruker APEXII area-detector diffractometer	14198 measured reflections
Absorption correction: multi-scan (<i>SADABS</i> ; Sheldrick, 1996)	2401 independent reflections
$T_{\text{min}} = 0.722$, $T_{\text{max}} = 0.838$	2196 reflections with $I > 2\sigma(I)$
	$R_{\text{int}} = 0.030$

Refinement

$R[F^2 > 2\sigma(F^2)] = 0.024$	H atoms treated by a mixture of independent and constrained refinement
$wR(F^2) = 0.057$	$\Delta\rho_{\text{max}} = 0.21$ e Å ⁻³
$S = 1.02$	$\Delta\rho_{\text{min}} = -0.28$ e Å ⁻³
2401 reflections	Absolute structure: Flack (1983), 1171 Friedel pairs
157 parameters	Flack parameter: 0.019 (15)
7 restraints	

Table 1

Selected bond lengths (Å).

Mn1—O3 ⁱ	2.0910 (13)	Mn1—O1W	2.1892 (13)
Mn1—O4	2.1527 (13)	Mn1—O2	2.1929 (15)
Mn1—O2W	2.1722 (16)	Mn1—O1	2.2660 (11)

Symmetry code: (i) $x, -y + 1, z - \frac{1}{2}$

Table 2

Hydrogen-bond geometry (Å, °).

$D-H\cdots A$	$D-H$	$H\cdots A$	$D\cdots A$	$D-H\cdots A$
O1W—H1WA...O4 ⁱⁱ	0.87 (2)	1.83 (2)	2.6965 (17)	175 (3)
O2W—H2WA...O5 ⁱⁱⁱ	0.88 (2)	1.95 (2)	2.796 (2)	161 (3)
O1W—H1WB...O5 ⁱ	0.82 (2)	2.01 (2)	2.809 (2)	166 (3)
O2W—H2WB...O2 ⁱ	0.80 (2)	2.13 (2)	2.822 (2)	145 (3)

Symmetry codes: (i) $x, -y + 1, z - \frac{1}{2}$; (ii) $-x, -y + 1, z$; (iii) $-x, y, z - \frac{1}{2}$.

Data collection: *APEX2* (Bruker, 2006); cell refinement: *SAINT* (Bruker, 2006); data reduction: *SAINT*; program(s) used to solve structure: *SHELXS97* (Sheldrick, 2008); program(s) used to refine structure: *SHELXL97* (Sheldrick, 2008); molecular graphics: *SHELXTL* (Sheldrick, 2008); software used to prepare material for publication: *SHELXL97*.

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Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: WM2346).

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supplementary materials

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***catena*-Poly[[diaquamanganese(II)]- μ -7-oxabicyclo[2.2.1]heptane-2,3-dicarboxylato]**

N. Wang, J. Feng, D.-H. Wang and Q.-Y. Lin

Comment

7-oxabicyclo[2,2,1]heptane-2,3-dicarboxylic anhydride (norcantharidin) derived from cantharidin is a lower toxicity anti-cancer drug (Shimi *et al.*, 1982) which makes this compound and its derived metal complexes interesting for structural research.

In the title polymer, each Mn^{II} atom is six-coordinated in a distorted octahedral coordination, defined by the bridging oxygen atom (O1) of the bicycloheptane unit, two oxygen atoms (O2 and O4) from carboxylate groups, one carboxylate oxygen atom (O3A) from a symmetry-related bridging ligand, and two oxygen atoms (O1W and O2W) from two water molecules. Each demethylcantharate anion adopts simultaneously a bridging coordination mode (O2 and O3 towards neighbouring Mn^{II} atoms) and a monodentate coordination mode (through O4 towards Mn1). Owing to the binding of the bridging oxygen atom (O1) with Mn^{II}, two six-membered rings (Mn1/O2/C8/C5/C4/O1 and Mn1/O4/C7/C6/C1/O1) are created. In addition, a seven-membered ring (Mn1/O2/C8/C5/C6/C7/O4) is formed because of the coordination of the carboxylate oxygen atoms (O2 and O4).

The crystal structure is stabilized by several hydrogen bonding interactions of the type O—H \cdots O, involving the coordinated water molecules as donors and carboxylate O atoms as acceptors (Table 2).

Experimental

Disodium demethylcantharate was prepared in accordance with the literature technique (Yin *et al.*, 2003).

A solution of manganese(II) acetate (1 mmol) and disodium demethylcantharate (1 mmol) was stirred at the room temperature. After 1 h, a solution of 2-amino-1,3,4-thiadiazole (2 mmol) was added dropwise to the mixed solution. Crystals suitable for single-crystal X-ray diffraction were obtained as colourless blocks over a period of several weeks.

Refinement

H atoms bonded to C atoms were positioned geometrically and were refined using a riding model [$d(\text{C—H}) = 0.97\text{--}0.98 \text{ \AA}$, $U_{\text{iso}}(\text{H}) = 1.2U_{\text{eq}}(\text{C})$]. H atoms bonded to water O atoms were located in a difference Fourier maps and were refined with an O—H distance restraint of $0.85(2) \text{ \AA}$ and $U_{\text{iso}}(\text{H}) = 1.5U_{\text{eq}}(\text{O})$.

Figures

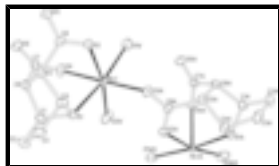


Fig. 1. A view of the molecule of (I) showing the atom-labelling scheme with displacement ellipsoids drawn at the 30% probability.

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Crystal data

[Mn(C₈H₈O₅)(H₂O)₂]

$M_r = 275.12$

Orthorhombic, *Iba*2

Hall symbol: I 2 -2c

$a = 10.3513$ (2) Å

$b = 18.9903$ (4) Å

$c = 10.4899$ (5) Å

$V = 2062.04$ (11) Å³

$Z = 8$

$F(000) = 1128$

$D_x = 1.772$ Mg m⁻³

Mo $K\alpha$ radiation, $\lambda = 0.71073$ Å

Cell parameters from 6907 reflections

$\theta = 2.1$ – 27.8°

$\mu = 1.30$ mm⁻¹

$T = 296$ K

Block, colourless

$0.37 \times 0.22 \times 0.14$ mm

Data collection

Bruker APEXII area-detector diffractometer

Radiation source: fine-focus sealed tube graphite

ω scans

Absorption correction: multi-scan (*SADABS*; Sheldrick, 1996)

$T_{\min} = 0.722$, $T_{\max} = 0.838$

14198 measured reflections

2401 independent reflections

2196 reflections with $I > 2\sigma(I)$

$R_{\text{int}} = 0.030$

$\theta_{\max} = 27.8^\circ$, $\theta_{\min} = 2.1^\circ$

$h = -13 \rightarrow 12$

$k = -23 \rightarrow 24$

$l = -13 \rightarrow 13$

Refinement

Refinement on F^2

Least-squares matrix: full

$R[F^2 > 2\sigma(F^2)] = 0.024$

$wR(F^2) = 0.057$

$S = 1.02$

2401 reflections

157 parameters

7 restraints

Secondary atom site location: difference Fourier map

Hydrogen site location: inferred from neighbouring sites

H atoms treated by a mixture of independent and constrained refinement

$w = 1/[\sigma^2(F_o^2) + (0.0349P)^2]$

where $P = (F_o^2 + 2F_c^2)/3$

$(\Delta/\sigma)_{\max} = 0.001$

$\Delta\rho_{\max} = 0.21$ e Å⁻³

$\Delta\rho_{\min} = -0.28$ e Å⁻³

Absolute structure: Flack (1983), 1171 Friedel pairs

Primary atom site location: structure-invariant direct methods Flack parameter: 0.019 (15)

Special details

Geometry. All esds (except the esd in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell esds are taken into account individually in the estimation of esds in distances, angles and torsion angles; correlations between esds in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell esds is used for estimating esds involving l.s. planes.

Refinement. Refinement of F^2 against ALL reflections. The weighted R -factor wR and goodness of fit S are based on F^2 , conventional R -factors R are based on F , with F set to zero for negative F^2 . The threshold expression of $F^2 > \sigma(F^2)$ is used only for calculating R -factors(gt) etc. and is not relevant to the choice of reflections for refinement. R -factors based on F^2 are statistically about twice as large as those based on F , and R -factors based on ALL data will be even larger.

Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters (\AA^2)

	<i>x</i>	<i>y</i>	<i>z</i>	$U_{\text{iso}}^*/U_{\text{eq}}$
Mn1	0.23964 (2)	0.466637 (12)	0.21088 (5)	0.02642 (8)
O1	0.37728 (10)	0.37412 (6)	0.22354 (12)	0.0264 (3)
O3	0.35327 (13)	0.46639 (7)	0.59926 (14)	0.0309 (3)
C8	0.35590 (15)	0.45297 (10)	0.48293 (17)	0.0235 (4)
O1W	0.12217 (13)	0.56135 (8)	0.23773 (15)	0.0428 (4)
C6	0.28402 (16)	0.32972 (9)	0.4091 (2)	0.0266 (4)
H6A	0.2856	0.2879	0.4636	0.032*
O5	0.08680 (12)	0.35748 (7)	0.51658 (13)	0.0365 (3)
O2	0.32170 (11)	0.49510 (7)	0.39659 (14)	0.0300 (3)
O4	0.11371 (12)	0.40095 (7)	0.32282 (15)	0.0359 (3)
C5	0.40058 (15)	0.37989 (9)	0.44426 (17)	0.0250 (4)
H5A	0.4519	0.3589	0.5130	0.030*
C7	0.15121 (15)	0.36557 (9)	0.41784 (18)	0.0255 (4)
O2W	0.15899 (16)	0.41688 (10)	0.04225 (15)	0.0490 (4)
C4	0.47817 (16)	0.37949 (10)	0.31993 (18)	0.0289 (4)
H4A	0.5344	0.4206	0.3094	0.035*
C1	0.32042 (18)	0.30945 (10)	0.27210 (19)	0.0309 (4)
H1A	0.2471	0.2926	0.2215	0.037*
C3	0.54794 (18)	0.30928 (12)	0.3051 (2)	0.0394 (5)
H3A	0.5904	0.2955	0.3837	0.047*
H3B	0.6112	0.3111	0.2370	0.047*
C2	0.4359 (2)	0.25889 (11)	0.2721 (2)	0.0439 (5)
H2A	0.4481	0.2373	0.1892	0.053*
H2B	0.4262	0.2224	0.3361	0.053*
H2WA	0.0832 (18)	0.4022 (14)	0.017 (3)	0.066*
H1WA	0.0462 (18)	0.5712 (14)	0.268 (2)	0.066*
H2WB	0.193 (2)	0.4298 (16)	-0.022 (2)	0.066*
H1WB	0.125 (2)	0.5871 (13)	0.176 (2)	0.066*

supplementary materials

Atomic displacement parameters (\AA^2)

	U^{11}	U^{22}	U^{33}	U^{12}	U^{13}	U^{23}
Mn1	0.02597 (12)	0.02875 (14)	0.02454 (14)	-0.00026 (9)	0.00046 (14)	0.00452 (15)
O1	0.0268 (5)	0.0279 (6)	0.0246 (7)	0.0008 (4)	0.0014 (5)	-0.0003 (6)
O3	0.0305 (6)	0.0387 (7)	0.0235 (7)	0.0061 (5)	-0.0003 (5)	-0.0052 (5)
C8	0.0159 (7)	0.0306 (9)	0.0240 (10)	-0.0027 (6)	0.0019 (7)	-0.0023 (8)
O1W	0.0354 (7)	0.0479 (9)	0.0452 (10)	0.0157 (6)	0.0108 (6)	0.0147 (7)
C6	0.0284 (8)	0.0222 (8)	0.0292 (10)	-0.0010 (7)	0.0010 (7)	0.0041 (8)
O5	0.0320 (7)	0.0453 (8)	0.0321 (8)	0.0008 (6)	0.0049 (6)	0.0033 (6)
O2	0.0379 (7)	0.0265 (6)	0.0257 (7)	0.0034 (6)	-0.0002 (6)	0.0008 (6)
O4	0.0233 (6)	0.0449 (8)	0.0395 (9)	0.0013 (6)	0.0017 (6)	0.0161 (7)
C5	0.0218 (8)	0.0282 (9)	0.0250 (10)	0.0026 (7)	-0.0014 (7)	-0.0011 (8)
C7	0.0228 (8)	0.0243 (9)	0.0294 (10)	-0.0053 (7)	-0.0008 (7)	0.0013 (8)
O2W	0.0458 (9)	0.0724 (11)	0.0289 (8)	-0.0249 (8)	-0.0036 (7)	0.0029 (8)
C4	0.0210 (8)	0.0357 (10)	0.0301 (11)	0.0036 (7)	0.0013 (7)	-0.0052 (8)
C1	0.0345 (9)	0.0256 (9)	0.0325 (10)	-0.0025 (7)	-0.0015 (8)	-0.0038 (8)
C3	0.0342 (10)	0.0475 (12)	0.0365 (12)	0.0185 (9)	0.0023 (9)	-0.0050 (10)
C2	0.0553 (13)	0.0305 (11)	0.0461 (13)	0.0136 (9)	0.0063 (11)	-0.0045 (10)

Geometric parameters (\AA , $^\circ$)

Mn1—O3 ⁱ	2.0910 (13)	C6—H6A	0.9800
Mn1—O4	2.1527 (13)	O5—C7	1.241 (2)
Mn1—O2W	2.1722 (16)	O4—C7	1.263 (2)
Mn1—O1W	2.1892 (13)	C5—C4	1.532 (2)
Mn1—O2	2.1929 (15)	C5—H5A	0.9800
Mn1—O1	2.2660 (11)	O2W—H2WA	0.875 (16)
O1—C1	1.454 (2)	O2W—H2WB	0.800 (17)
O1—C4	1.457 (2)	C4—C3	1.524 (3)
O3—C8	1.247 (2)	C4—H4A	0.9800
O3—Mn1 ⁱⁱ	2.0910 (13)	C1—C2	1.534 (3)
C8—O2	1.259 (2)	C1—H1A	0.9800
C8—C5	1.518 (2)	C3—C2	1.543 (3)
O1W—H1WA	0.867 (15)	C3—H3A	0.9700
O1W—H1WB	0.815 (16)	C3—H3B	0.9700
C6—C1	1.534 (3)	C2—H2A	0.9700
C6—C7	1.537 (2)	C2—H2B	0.9700
C6—C5	1.581 (2)		
O3 ⁱ —Mn1—O4	176.92 (5)	C4—C5—C6	101.44 (14)
O3 ⁱ —Mn1—O2W	91.42 (6)	C8—C5—H5A	109.9
O4—Mn1—O2W	87.68 (6)	C4—C5—H5A	109.9
O3 ⁱ —Mn1—O1W	83.39 (5)	C6—C5—H5A	109.9
O4—Mn1—O1W	93.99 (5)	O5—C7—O4	123.99 (16)
O2W—Mn1—O1W	104.40 (7)	O5—C7—C6	118.40 (16)
O3 ⁱ —Mn1—O2	97.45 (5)	O4—C7—C6	117.60 (16)

O4—Mn1—O2	83.84 (6)	Mn1—O2W—H2WA	137.1 (18)
O2W—Mn1—O2	168.38 (6)	Mn1—O2W—H2WB	113 (2)
O1W—Mn1—O2	84.17 (6)	H2WA—O2W—H2WB	103 (2)
O3 ⁱ —Mn1—O1	98.67 (5)	O1—C4—C3	101.98 (14)
O4—Mn1—O1	84.23 (4)	O1—C4—C5	102.44 (13)
O2W—Mn1—O1	87.25 (6)	C3—C4—C5	109.84 (16)
O1W—Mn1—O1	168.16 (6)	O1—C4—H4A	113.8
O2—Mn1—O1	84.00 (5)	C3—C4—H4A	113.8
C1—O1—C4	96.10 (13)	C5—C4—H4A	113.8
C1—O1—Mn1	114.89 (9)	O1—C1—C2	102.31 (15)
C4—O1—Mn1	115.91 (9)	O1—C1—C6	102.45 (14)
C8—O3—Mn1 ⁱⁱ	133.21 (12)	C2—C1—C6	110.39 (16)
O3—C8—O2	124.60 (17)	O1—C1—H1A	113.5
O3—C8—C5	117.07 (16)	C2—C1—H1A	113.5
O2—C8—C5	118.31 (16)	C6—C1—H1A	113.5
Mn1—O1W—H1WA	136.7 (17)	C4—C3—C2	102.08 (15)
Mn1—O1W—H1WB	112.0 (19)	C4—C3—H3A	111.4
H1WA—O1W—H1WB	100.8 (19)	C2—C3—H3A	111.4
C1—C6—C7	112.76 (16)	C4—C3—H3B	111.4
C1—C6—C5	100.51 (14)	C2—C3—H3B	111.4
C7—C6—C5	113.69 (14)	H3A—C3—H3B	109.2
C1—C6—H6A	109.8	C1—C2—C3	101.40 (15)
C7—C6—H6A	109.8	C1—C2—H2A	111.5
C5—C6—H6A	109.8	C3—C2—H2A	111.5
C8—O2—Mn1	126.25 (12)	C1—C2—H2B	111.5
C7—O4—Mn1	123.55 (11)	C3—C2—H2B	111.5
C8—C5—C4	113.07 (15)	H2A—C2—H2B	109.3
C8—C5—C6	112.38 (13)		
O3 ⁱ —Mn1—O1—C1	-167.13 (12)	C1—C6—C5—C4	1.19 (16)
O4—Mn1—O1—C1	11.83 (12)	C7—C6—C5—C4	-119.54 (17)
O2W—Mn1—O1—C1	-76.11 (12)	Mn1—O4—C7—O5	141.81 (15)
O1W—Mn1—O1—C1	93.7 (3)	Mn1—O4—C7—C6	-39.0 (2)
O2—Mn1—O1—C1	96.22 (12)	C1—C6—C7—O5	149.40 (16)
O3 ⁱ —Mn1—O1—C4	82.04 (11)	C5—C6—C7—O5	-97.02 (19)
O4—Mn1—O1—C4	-99.00 (11)	C1—C6—C7—O4	-29.8 (2)
O2W—Mn1—O1—C4	173.06 (12)	C5—C6—C7—O4	83.8 (2)
O1W—Mn1—O1—C4	-17.1 (3)	C1—O1—C4—C3	56.75 (16)
O2—Mn1—O1—C4	-14.61 (11)	Mn1—O1—C4—C3	178.24 (11)
Mn1 ⁱⁱ —O3—C8—O2	27.8 (3)	C1—O1—C4—C5	-56.96 (15)
Mn1 ⁱⁱ —O3—C8—C5	-150.61 (12)	Mn1—O1—C4—C5	64.53 (14)
O3—C8—O2—Mn1	-150.12 (13)	C8—C5—C4—O1	-86.60 (16)
C5—C8—O2—Mn1	28.30 (19)	C6—C5—C4—O1	33.96 (16)
O3 ⁱ —Mn1—O2—C8	-133.89 (13)	C8—C5—C4—C3	165.60 (15)
O4—Mn1—O2—C8	48.92 (13)	C6—C5—C4—C3	-73.84 (17)
O2W—Mn1—O2—C8	5.5 (4)	C4—O1—C1—C2	-56.40 (16)
O1W—Mn1—O2—C8	143.59 (13)	Mn1—O1—C1—C2	-178.67 (11)
O1—Mn1—O2—C8	-35.89 (13)	C4—O1—C1—C6	58.02 (14)

supplementary materials

O2W—Mn1—O4—C7	131.49 (16)	Mn1—O1—C1—C6	-64.25 (15)
O1W—Mn1—O4—C7	-124.24 (15)	C7—C6—C1—O1	85.26 (17)
O2—Mn1—O4—C7	-40.55 (15)	C5—C6—C1—O1	-36.13 (16)
O1—Mn1—O4—C7	44.01 (15)	C7—C6—C1—C2	-166.38 (15)
O3—C8—C5—C4	-144.55 (16)	C5—C6—C1—C2	72.23 (17)
O2—C8—C5—C4	36.9 (2)	O1—C4—C3—C2	-35.23 (18)
O3—C8—C5—C6	101.34 (18)	C5—C4—C3—C2	72.87 (18)
O2—C8—C5—C6	-77.21 (19)	O1—C1—C2—C3	34.25 (19)
C1—C6—C5—C8	122.23 (15)	C6—C1—C2—C3	-74.20 (19)
C7—C6—C5—C8	1.5 (2)	C4—C3—C2—C1	0.64 (19)

Symmetry codes: (i) $x, -y+1, z-1/2$; (ii) $x, -y+1, z+1/2$.

Hydrogen-bond geometry (\AA , $^\circ$)

$D-H\cdots A$	$D-H$	$H\cdots A$	$D\cdots A$	$D-H\cdots A$
O1W—H1WA \cdots O4 ⁱⁱⁱ	0.87 (2)	1.83 (2)	2.6965 (17)	175 (3)
O2W—H2WA \cdots O5 ^{iv}	0.88 (2)	1.95 (2)	2.796 (2)	161 (3)
O1W—H1WB \cdots O5 ⁱ	0.82 (2)	2.01 (2)	2.809 (2)	166 (3)
O2W—H2WB \cdots O2 ⁱ	0.80 (2)	2.13 (2)	2.822 (2)	145 (3)

Symmetry codes: (iii) $-x, -y+1, z$; (iv) $-x, y, z-1/2$; (i) $x, -y+1, z-1/2$.

Fig. 1

