

Received 3 June 2015

Accepted 30 June 2015

Edited by M. Weil, Vienna University of
Technology, Austria

Keywords: crystal structure; picloram; Tordon; herbicide; sodium salt; coordination polymer; sodium–water cationic chain; hydrogen bonding

CCDC reference: 1409779

Supporting information: this article has supporting information at journals.iucr.org/e

Crystal structure of poly[[di- μ_2 -aqua-aquasodium] 4-amino-3,5,6-trichloropyridine-2-carboxylate trihydrate], the sodium salt of the herbicide picloram

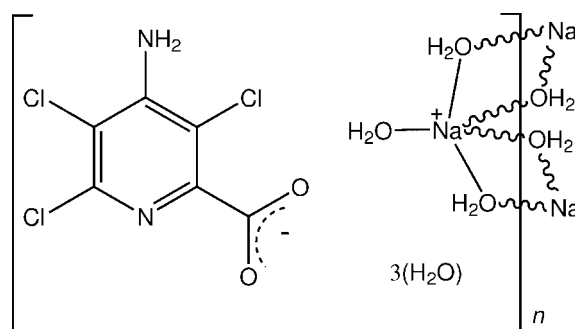
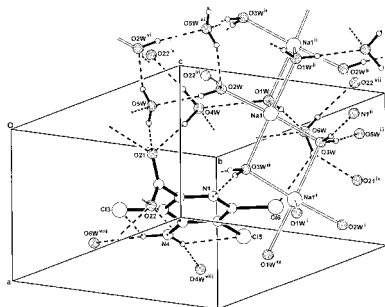
Graham Smith

Science and Engineering Faculty, Queensland University of Technology, GPO Box 2434, Brisbane, Queensland 4001, Australia. *Correspondence e-mail: g.smith@qut.edu.au

In the structure of the title complex, $\{[\text{Na}(\text{H}_2\text{O})_3](\text{C}_6\text{H}_2\text{Cl}_3\text{N}_2\text{O}_2)\cdot 3\text{H}_2\text{O}\}_n$, the sodium salt of the herbicide picloram, the cation adopts a polymeric chain structure, based on μ_2 -aqua-bridged NaO_5 trigonal-bipyramidal complex units which have, in addition, a singly bonded water molecule. Each of the bridges within the chain, which extends parallel to the *a* axis, is centrosymmetric, with $\text{Na}\cdots\text{Na}$ separations of 3.4807 (16) and 3.5109 (16) Å. In the crystal, there are three water molecules of solvation and these, as well as the coordinating water molecules and the amino group of the 4-amino-3,5,6-trichloropicolinate anion, are involved in extensive inter-species hydrogen-bonding interactions with carboxyl and water O atoms, as well as the pyridine N atom. Among these associations is a centrosymmetric cyclic tetrawater $R_4^4(8)$ motif, resulting in an overall three-dimensional structure.

1. Chemical context

4-Amino-3,5,6-trichloropyridine-2-carboxylic acid (picloram) is a commercial herbicide (Mullinson, 1985) introduced by Dow Chemicals as Tordon (O'Neil, 2001). Although it has potential as a metal-chelating ligand similar to picolinic acid, there are only five metal complexes with picloramate anions in the crystallographic literature. Examples include picloram as a bidentate *N,O* chelating ligand with Mn^{II} (Smith *et al.*, 1981a) and Cu^{II} (two structures, one a mixed-ligand complex with 2-aminopyrimidine; O'Reilly *et al.*, 1983) and caesium (Smith, 2013). In the Mg complex (Smith *et al.*, 1981b), the picloramate anions act as counter-ions to the $[\text{Mg}(\text{H}_2\text{O})_6]^{2+}$ cation. Although the structure of picloram has not been reported, that of the guanidinium salt is known (Parthasarathi *et al.*, 1982). The reaction of picloram with sodium bicarbonate in aqueous ethanol gave crystals of the title complex salt $\{[\text{Na}(\text{H}_2\text{O})_3]^+\cdot\text{C}_6\text{H}_2\text{Cl}_3\text{N}_2\text{O}_2^-\cdot 3\text{H}_2\text{O}\}_n$, and the structure is reported herein.



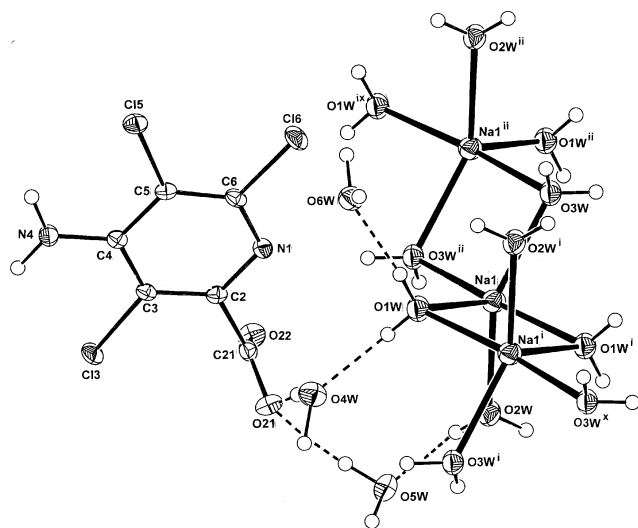


Figure 1
The atom-numbering scheme for the hydrated title salt, with non-H atoms drawn as 40% probability ellipsoids. Inter-species hydrogen bonds are shown as dashed lines. Symmetry codes: (ix) $x + 1, y, z$; (x) $x - 1, y, z$. For other symmetry codes, see Table 1.

2. Structural commentary

In the structure of the title salt, (Fig. 1), polymeric cationic chains based on μ_2 -water-bridged NaO_5 trigonal-bipyramidal complex units are formed, comprising centrosymmetric four-membered water-bridged Na_2O_2 rings with both O1W and O3W [$\text{Na}\cdots\text{Na}^i$ and $\text{Na}\cdots\text{Na}^{ii}$ = 3.4807 (16) and 3.5109 (16) Å, respectively; for symmetry codes, see Table 1]. In the fifth Na coordination site is the third water molecule (O2W) in a non-bridging mode [overall Na–O range, 2.3183 (17)–2.4185 (16) Å: Table 1]. Although the μ_2 -water-bridged cationic chains are relatively common, the NaO_5 coordination with one non-bridging water is rare, compared to

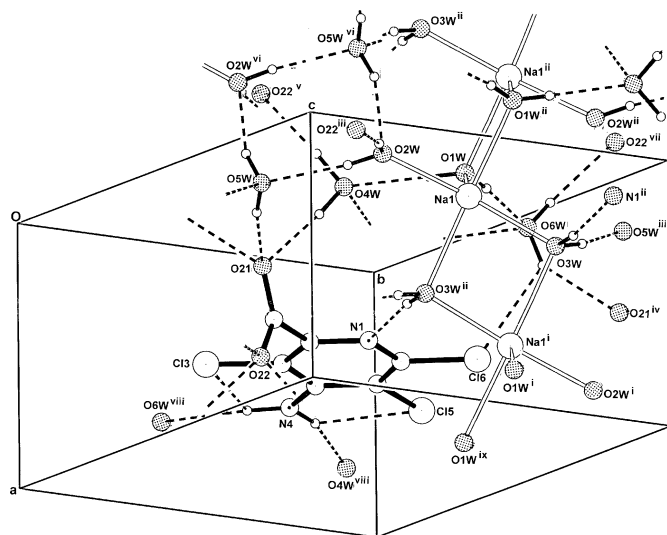


Figure 2
The three-dimensional hydrogen-bonded structure, with inter-species hydrogen bonds and intramolecular N–H \cdots Cl associations shown as dashed lines. For symmetry codes, see Fig. 1 and Table 2.

Table 1
Selected bond lengths (Å).

Na1–O1W ⁱ	2.4185 (16)	Na1–O2W	2.3183 (17)
Na1–O3W ⁱⁱ	2.3803 (16)	Na1–O3W	2.3530 (16)
Na1–O1W	2.3529 (16)		

Symmetry codes: (i) $-x, -y + 1, -z + 2$; (ii) $-x + 1, -y + 1, -z + 2$.

Table 2
Hydrogen-bond geometry (Å, °).

$D\cdots H\cdots A$	$D\cdots H$	$H\cdots A$	$D\cdots A$	$D\cdots H\cdots A$
O1W–H11W \cdots O4W	0.82	1.97	2.786 (2)	178
O1W–H12W \cdots O6W	0.89	1.99	2.877 (2)	174
O2W–H21W \cdots O22 ⁱⁱⁱ	0.94	1.79	2.721 (2)	171
O2W–H22W \cdots O5W	0.84	1.98	2.816 (2)	172
O3W–H31W \cdots O5W ^{iv}	0.88	1.93	2.781 (2)	163
O3W–H32W \cdots N1 ⁱⁱ	0.90	2.02	2.910 (2)	170
O4W–H41W \cdots O22 ^v	0.99	1.93	2.916 (2)	173
O4W–H42W \cdots O21	0.93	1.99	2.918 (2)	174
O5W–H51W \cdots O21	0.90	1.87	2.748 (2)	166
O5W–H52W \cdots O2W ^{vi}	0.90	2.01	2.8481 (19)	155
O6W–H61W \cdots O22 ^{vii}	0.93	2.10	2.972 (2)	156
O6W–H62W \cdots O21 ^{iv}	0.93	2.19	2.996 (2)	144
N4–H41 \cdots O6W ^{viii}	0.96	2.18	3.080 (2)	156
N4–H42 \cdots O4W ^{viii}	0.97	2.22	3.146 (2)	160

Symmetry codes: (ii) $-x + 1, -y + 1, -z + 2$; (iii) $-x + 1, -y, -z + 2$; (iv) $x, y + 1, z$; (v) $x - 1, y, z$; (vi) $-x, -y, -z + 2$; (vii) $x - 1, y + 1, z$; (viii) $-x + 1, -y + 1, -z + 1$.

the more usual octahedral NaO_6 coordination involving two non-bridging water molecules in other examples, *e.g.* in the biphenyl-4,4'-diphosphonate salt (Kinnibrugh *et al.*, 2012).

The structure of the title salt also contains non-coordinating picloramate anions and three water molecules of solvation (O4W–O6W). In this anion, the carboxyl group lies close to perpendicular to the pyridine ring [torsion angle N1–C2–C21–O21 = 89.1 (2)°], which is similar to that in the anhydrous guanidinium picloramate salt (73.3°) (Parthasarathi *et al.*, 1982), while the amine group gives lateral intramolecular N4–H41 \cdots Cl3 and N4–H41 \cdots O6W interactions [2.9956 (17), 3.080 (2) Å].

3. Supramolecular features

In the crystal there are numerous inter-species water O–H \cdots O_{carboxyl/water}, O–H \cdots N_{pyridine} and O–H \cdots Cl hydrogen-bonding interactions (Table 2), including a centrosymmetric tetra-water cyclic ring involving O2W–H \cdots O5W and O5W–H \cdots O2W^{vi} [graph set $R_4^4(8)$], giving a three-dimensional structure (Fig. 2). Cyclic tetra-water moieties such as found in the present structure are being identified in an increasing number in labile water-stabilized salt hydrates, *e.g.* in the brucinium L-glycerate 4.75-hydrate salt (Białońska *et al.*, 2005). Also found in the structure of the title salt is a short intermolecular Cl3 \cdots Cl5^{xi} contact [3.2108 (7) Å; symmetry code (xi): $x, y - 1, z$].

4. Database survey

The (μ_2 -aqua)-bridged $\text{Na}_2(\text{H}_2\text{O})_2$ units in the coordination polymeric cationic chains in the title structure have precedents

in a large number of reported crystal structures. However, with few exceptions, these are based on NaO_6 polyhedra, with octahedral or distorted octahedral stereochemistry, having two non-bridging water molecules $[\text{Na}_2(\text{H}_2\text{O})_8]^{2+}$, compared to one non-bridging water molecule in the NaO_5 coordination $[\text{Na}_2(\text{H}_2\text{O})_6]^{2+}$ of the title complex. The $[\text{Na}_2(\text{H}_2\text{O})_8]^{2+}$ dications may be discrete, such as found in the anionic aryltelluronic anhydride salt (Beckmann *et al.*, 2012) and the anionic dimethylarsenate (cacodylate) salt (Lennartson & Håkansson, 2008), or they may be found as $[\text{Na}_4(\text{H}_2\text{O})_{16}]^{4+}$ tetra-cations as found in the dianionic biphenyl-4,4'-diphosphonate salt (Kinnibrugh *et al.*, 2012) and the monoanionic salt of luminol (5-amino-2,3-dihydro-1,4-phthalazinedione; Guzei *et al.*, 2013). However, more commonly, they are polymeric $[\text{Na}_2(\text{H}_2\text{O})_8]_n$, *e.g.* in the monoanionic salt of the anti-allergic drug tranilast ($\{2\text{-}[3\text{-}(3,4\text{-dimethoxyphenyl)acryloyl]amino\}benzoic\}$ acid; Geng *et al.*, 2013), but often associated with metal complex anions, *e.g.* the Cu^{II} complex with pyrophosphate, $[\text{Cu}(\text{H}_2\text{O})(\text{phen})(\text{P}_2\text{O}_7)]^{2-}$ (phen = 1,10-phenanthroline; Marino *et al.*, 2010), the mixed-valent di- $\text{Ru}^{\text{II,III}}$ complex with 1-hydroxyethane 1,1-diphosphonate (hedp) $[\text{Ru}_2(\text{hedp})_2\text{X}]^{4-}$ ($\text{X} = \text{Cl}, \text{Br}$; Yi *et al.*, 2005) and the dioxo-Np complex anion salt with dipicolinate (dipic), $[\text{NpO}_2(\text{dipic})(\text{H}_2\text{O})_2]^-$ (Tian *et al.*, 2009).

5. Synthesis and crystallization

The title compound was synthesized by briefly heating together 0.5 mmol of 4-amino-3,5,6-trichloropicolinic acid (picloram) with excess NaHCO_3 in 10 ml of 10% (v/v) ethanol–water. Room temperature evaporation of the solution to dryness gave minor colourless crystal blocks of the title complex from which a specimen was cleaved for the X-ray analysis.

6. Refinement details

Crystal data, data collection and structure refinement details are summarized in Table 3. Hydrogen atoms of the water molecules and the amine group were located in a difference-Fourier synthesis but were subsequently constrained in the refinement with the isotropic displacement parameters allowed to ride, with $U_{\text{iso}}(\text{H}) = 1.2U_{\text{eq}}(\text{N})$ or $1.5U_{\text{eq}}(\text{O})$.

Acknowledgements

The author acknowledges financial support from the Science and Engineering Faculty, Queensland University of Technology, Brisbane.

References

- Agilent (2014). *CrysAlis PRO*. Agilent Technologies Ltd., Yarnton, England.
 Altomare, A., Cascarano, G., Giacovazzo, C. & Guagliardi, A. (1993). *J. Appl. Cryst.* **26**, 343–350.
 Beckmann, J., Duthie, A., Gesing, T. M., Koehne, T. & Lork, E. (2012). *Organometallics*, **31**, 3451–3454.

Table 3
Experimental details.

Crystal data	
Chemical formula	$[\text{Na}(\text{H}_2\text{O})_3](\text{C}_6\text{H}_2\text{Cl}_3\text{N}_2\text{O}_2)\cdot 3\text{H}_2\text{O}$
M_r	371.53
Crystal system, space group	Triclinic, $P\bar{1}$
Temperature (K)	200
a, b, c (Å)	6.5625 (5), 8.4574 (6), 13.8553 (10)
α, β, γ (°)	78.747 (6), 79.374 (6), 88.864 (6)
V (Å ³)	741.17 (9)
Z	2
Radiation type	Mo $K\alpha$
μ (mm ⁻¹)	0.68
Crystal size (mm)	0.35 × 0.35 × 0.22
Data collection	
Diffractometer	Oxford Diffraction Gemini-S CCD detector
Absorption correction	Multi-scan (<i>CrysAlis PRO</i> ; Agilent, 2014)
$T_{\text{min}}, T_{\text{max}}$	0.947, 0.980
No. of measured, independent and observed [$I > 2\sigma(I)$] reflections	6040, 2905, 2487
R_{int}	0.025
$(\sin \theta/\lambda)_{\text{max}}$ (Å ⁻¹)	0.617
Refinement	
$R[F^2 > 2\sigma(F^2)], wR(F^2), S$	0.031, 0.088, 1.00
No. of reflections	2905
No. of parameters	182
H-atom treatment	H-atom parameters constrained
$\Delta\rho_{\text{max}}, \Delta\rho_{\text{min}}$ (e Å ⁻³)	0.30, -0.28

Computer programs: *CrysAlis PRO* (Agilent, 2014), *SIR92* (Altomare *et al.*, 1993), *SHELXL97* (Sheldrick, 2008) within *WinGX* (Farrugia, 2012), *PLATON* (Spek, 2009).

- Białońska, A., Ciunik, Z., Popek, T. & Lis, T. (2005). *Acta Cryst.* **C61**, o88–o91.
 Farrugia, L. J. (2012). *J. Appl. Cryst.* **45**, 849–854.
 Geng, N., Chen, J. M., Li, Z.-J., Jiang, L. & Lu, T.-B. (2013). *Cryst. Growth Des.* **13**, 3546–3553.
 Guzei, I. A., Kim, M.-H. & West, R. (2013). *J. Coord. Chem.* **66**, 3722–3739.
 Kinnibrugh, T., Garcia, N. & Clearfield, A. (2012). *J. Solid State Chem.* **187**, 149–158.
 Lennartson, A. & Håkansson, M. (2008). *Acta Cryst.* **C64**, m13–m16.
 Marino, N., Vortherms, A. R., Hoffman, R. L. & Doyle, R. P. (2010). *Inorg. Chem.* **49**, 6790–6792.
 Mullinson, W. R. (1985). *Proc. Western Soc. Weed Sci.* **38**, 21–92.
 O'Neil, M. J. (2001). Editor. *The Merck Index*, 13th ed., pp. 1325–1326. Whitehouse Station, NJ, USA: Merck & Co. Inc.
 O'Reilly, E. J., Smith, G., Kennard, C. H. L. & White, A. H. (1983). *Aust. J. Chem.* **36**, 183–190.
 Parthasarathi, V., Wolfrum, S., Noordik, J. H., Beurskens, P. T., Smith, G., O'Reilly, E. J. & Kennard, C. H. L. (1982). *Cryst. Struct. Commun.* **11**, 1519–1524.
 Sheldrick, G. M. (2008). *Acta Cryst.* **A64**, 112–122.
 Smith, G. (2013). *Acta Cryst.* **E69**, m22–m23.
 Smith, G., O'Reilly, E. J. & Kennard, C. H. L. (1981a). *Aust. J. Chem.* **34**, 891–896.
 Smith, G., O'Reilly, E. J. & Kennard, C. H. L. (1981b). *Cryst. Struct. Commun.* **10**, 1277–1282.
 Spek, A. L. (2009). *Acta Cryst.* **D65**, 148–155.
 Tian, G., Rao, L. & Teat, S. J. (2009). *Inorg. Chem.* **48**, 10158–10164.
 Yi, X.-Y., Liu, B., Jiménez-Aparicio, R., Urbanos, F. A., Gao, S., Xu, W., Chen, J.-S., Song, Y. & Zheng, L.-M. (2005). *Inorg. Chem.* **44**, 4309–4314.

supporting information

Acta Cryst. (2015). E71, 931-933 [doi:10.1107/S2056989015012633]

Crystal structure of poly[[di- μ_2 -aqua-aquasodium] 4-amino-3,5,6-trichloropyridine-2-carboxylate trihydrate], the sodium salt of the herbicide picloram

Graham Smith

Computing details

Data collection: *CrysAlis PRO* (Agilent, 2014); cell refinement: *CrysAlis PRO* (Agilent, 2014); data reduction: *CrysAlis PRO* (Agilent, 2014); program(s) used to solve structure: *SIR92* (Altomare *et al.*, 1993); program(s) used to refine structure: *SHELXL97* (Sheldrick, 2008) within *WinGX* (Farrugia, 2012); molecular graphics: *PLATON* (Spek, 2009); software used to prepare material for publication: *PLATON* (Spek, 2009).

Poly[[di- μ_2 -aqua-aquasodium] 4-amino-3,5,6-trichloropyridine-2-carboxylate trihydrate]

Crystal data

$[\text{Na}(\text{H}_2\text{O})_3](\text{C}_6\text{H}_2\text{Cl}_3\text{N}_2\text{O}_2)\cdot 3\text{H}_2\text{O}$

$M_r = 371.53$

Triclinic, $P1$

Hall symbol: -P 1

$a = 6.5625$ (5) Å

$b = 8.4574$ (6) Å

$c = 13.8553$ (10) Å

$\alpha = 78.747$ (6)°

$\beta = 79.374$ (6)°

$\gamma = 88.864$ (6)°

$V = 741.17$ (9) Å³

$Z = 2$

$F(000) = 380$

$D_x = 1.665$ Mg m⁻³

Mo $K\alpha$ radiation, $\lambda = 0.71073$ Å

Cell parameters from 2548 reflections

$\theta = 3.8$ – 28.8 °

$\mu = 0.68$ mm⁻¹

$T = 200$ K

Block, colourless

$0.35 \times 0.35 \times 0.22$ mm

Data collection

Oxford Diffraction Gemini-S CCD-detector diffractometer

Radiation source: Enhance (Mo) X-ray source

Graphite monochromator

Detector resolution: 16.077 pixels mm⁻¹

ω scans

Absorption correction: multi-scan

(*CrysAlis PRO*; Agilent, 2014)

$T_{\min} = 0.947$, $T_{\max} = 0.980$

6040 measured reflections

2905 independent reflections

2487 reflections with $I > 2\sigma(I)$

$R_{\text{int}} = 0.025$

$\theta_{\max} = 26.0$ °, $\theta_{\min} = 3.3$ °

$h = -8$ → 7

$k = -10$ → 10

$l = -17$ → 13

Refinement

Refinement on F^2

Least-squares matrix: full

$R[F^2 > 2\sigma(F^2)] = 0.031$

$wR(F^2) = 0.088$

$S = 1.00$

2905 reflections

182 parameters

0 restraints

Primary atom site location: structure-invariant direct methods

Secondary atom site location: difference Fourier map

Hydrogen site location: inferred from neighbouring sites

H-atom parameters constrained

$$w = 1/[\sigma^2(F_o^2) + (0.0478P)^2 + 0.2264P]$$

$$\text{where } P = (F_o^2 + 2F_c^2)/3$$

$$(\Delta/\sigma)_{\max} = 0.001$$

$$\Delta\rho_{\max} = 0.30 \text{ e } \text{\AA}^{-3}$$

$$\Delta\rho_{\min} = -0.28 \text{ e } \text{\AA}^{-3}$$

Extinction correction: *SHELXL97* (Sheldrick, 2008) within *WinGX* (Farrugia, 2012)

Extinction coefficient: 0.050 (3)

Special details

Geometry. Bond distances, angles *etc.* have been calculated using the rounded fractional coordinates. All su's are estimated from the variances of the (full) variance-covariance matrix. The cell e.s.d.'s are taken into account in the estimation of distances, angles and torsion angles

Refinement. Refinement of F^2 against ALL reflections. The weighted R -factor wR and goodness of fit S are based on F^2 , conventional R -factors R are based on F , with F set to zero for negative F^2 . The threshold expression of $F^2 > \sigma(F^2)$ is used only for calculating R -factors(gt) *etc.* and is not relevant to the choice of reflections for refinement. R -factors based on F^2 are statistically about twice as large as those based on F , and R -factors based on ALL data will be even larger.

Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters (\AA^2)

	<i>x</i>	<i>y</i>	<i>z</i>	$U_{\text{iso}}^*/U_{\text{eq}}$
Cl3	0.73580 (8)	0.09272 (5)	0.52593 (4)	0.0257 (2)
Cl5	0.76678 (8)	0.74102 (5)	0.46993 (4)	0.0240 (2)
Cl6	0.67863 (9)	0.70381 (6)	0.70472 (4)	0.0295 (2)
O21	0.4767 (2)	0.04778 (17)	0.77535 (11)	0.0287 (5)
O22	0.8188 (2)	0.04407 (17)	0.77067 (11)	0.0292 (5)
N1	0.6684 (2)	0.39495 (18)	0.71307 (12)	0.0196 (5)
N4	0.7849 (3)	0.42747 (19)	0.40294 (12)	0.0213 (5)
C2	0.6814 (3)	0.2631 (2)	0.67196 (14)	0.0172 (5)
C3	0.7193 (3)	0.2696 (2)	0.57064 (14)	0.0170 (5)
C4	0.7473 (3)	0.4178 (2)	0.50250 (14)	0.0168 (6)
C5	0.7335 (3)	0.5539 (2)	0.54712 (14)	0.0173 (5)
C6	0.6946 (3)	0.5361 (2)	0.64975 (15)	0.0186 (6)
C21	0.6561 (3)	0.1034 (2)	0.74534 (14)	0.0204 (6)
Na1	0.24455 (12)	0.42819 (9)	1.01558 (6)	0.0268 (3)
O1W	0.0817 (2)	0.53070 (17)	0.87914 (10)	0.0299 (5)
O2W	0.1503 (2)	0.15728 (17)	1.05390 (11)	0.0300 (5)
O3W	0.4202 (2)	0.62279 (17)	1.07140 (10)	0.0272 (5)
O4W	0.1371 (2)	0.27585 (17)	0.77676 (12)	0.0317 (5)
O5W	0.2700 (2)	-0.10534 (18)	0.95986 (11)	0.0325 (5)
O6W	0.1709 (2)	0.82429 (17)	0.73270 (11)	0.0298 (5)
H41	0.80460	0.32820	0.37840	0.0260*
H42	0.80680	0.53330	0.36040	0.0260*
H11W	0.09740	0.45450	0.85000	0.0450*
H12W	0.10510	0.62480	0.83710	0.0450*
H21W	0.15180	0.09520	1.11820	0.0450*
H22W	0.19310	0.08550	1.02150	0.0450*
H31W	0.38250	0.71930	1.04480	0.0410*
H32W	0.40980	0.61800	1.13770	0.0410*
H41W	0.02360	0.20490	0.77150	0.0480*
H42W	0.25100	0.20920	0.77350	0.0480*
H51W	0.34890	-0.07030	0.89970	0.0490*
H52W	0.15590	-0.13130	0.93860	0.18 (2)*

H61W	0.06160	0.89510	0.72480	0.0450*
H62W	0.30130	0.87260	0.72460	0.0450*

Atomic displacement parameters (Å²)

	U^{11}	U^{22}	U^{33}	U^{12}	U^{13}	U^{23}
Cl3	0.0359 (3)	0.0159 (2)	0.0269 (3)	0.0007 (2)	-0.0044 (2)	-0.0091 (2)
Cl5	0.0280 (3)	0.0146 (2)	0.0282 (3)	-0.0001 (2)	-0.0049 (2)	-0.0012 (2)
Cl6	0.0412 (3)	0.0196 (3)	0.0309 (3)	0.0023 (2)	-0.0063 (2)	-0.0131 (2)
O21	0.0247 (8)	0.0260 (8)	0.0314 (9)	-0.0065 (6)	-0.0013 (6)	0.0013 (6)
O22	0.0290 (8)	0.0250 (8)	0.0318 (9)	0.0043 (6)	-0.0089 (7)	0.0013 (6)
N1	0.0202 (9)	0.0195 (8)	0.0195 (9)	0.0004 (6)	-0.0025 (7)	-0.0060 (7)
N4	0.0263 (9)	0.0197 (8)	0.0178 (8)	0.0009 (7)	-0.0031 (7)	-0.0042 (6)
C2	0.0128 (9)	0.0164 (9)	0.0220 (10)	0.0008 (7)	-0.0024 (7)	-0.0037 (7)
C3	0.0152 (9)	0.0147 (9)	0.0224 (10)	0.0005 (7)	-0.0038 (7)	-0.0063 (7)
C4	0.0106 (9)	0.0193 (10)	0.0216 (10)	0.0008 (7)	-0.0044 (7)	-0.0053 (8)
C5	0.0141 (9)	0.0144 (9)	0.0228 (10)	0.0002 (7)	-0.0041 (7)	-0.0019 (7)
C6	0.0160 (10)	0.0167 (9)	0.0252 (10)	0.0015 (7)	-0.0040 (8)	-0.0090 (8)
C21	0.0251 (11)	0.0192 (9)	0.0177 (10)	0.0014 (8)	-0.0028 (8)	-0.0068 (7)
Na1	0.0249 (5)	0.0264 (4)	0.0291 (5)	-0.0002 (3)	-0.0031 (3)	-0.0068 (3)
O1W	0.0360 (9)	0.0275 (8)	0.0252 (8)	0.0022 (6)	-0.0021 (7)	-0.0065 (6)
O2W	0.0390 (9)	0.0239 (8)	0.0246 (8)	0.0020 (6)	-0.0008 (7)	-0.0039 (6)
O3W	0.0322 (9)	0.0264 (8)	0.0223 (8)	0.0022 (6)	-0.0020 (6)	-0.0064 (6)
O4W	0.0282 (8)	0.0283 (8)	0.0399 (9)	0.0023 (6)	-0.0067 (7)	-0.0097 (7)
O5W	0.0342 (9)	0.0353 (9)	0.0247 (8)	0.0009 (7)	-0.0010 (7)	-0.0020 (7)
O6W	0.0257 (8)	0.0311 (8)	0.0330 (9)	-0.0003 (6)	-0.0043 (7)	-0.0079 (6)

Geometric parameters (Å, °)

Na1—O1W ⁱ	2.4185 (16)	O4W—H42W	0.9300
Na1—O3W ⁱⁱ	2.3803 (16)	O4W—H41W	0.9900
Na1—O1W	2.3529 (16)	O5W—H52W	0.9000
Na1—O2W	2.3183 (17)	O5W—H51W	0.9000
Na1—O3W	2.3530 (16)	O6W—H61W	0.9300
Cl3—C3	1.7216 (18)	O6W—H62W	0.9300
Cl5—C5	1.7213 (18)	N1—C2	1.342 (2)
Cl6—C6	1.7289 (19)	N1—C6	1.330 (2)
O21—C21	1.243 (2)	N4—C4	1.342 (2)
O22—C21	1.250 (2)	N4—H41	0.9600
O1W—H12W	0.8900	N4—H42	0.9700
O1W—H11W	0.8200	C2—C3	1.370 (3)
O2W—H21W	0.9400	C2—C21	1.515 (3)
O2W—H22W	0.8400	C3—C4	1.407 (3)
O3W—H31W	0.8800	C4—C5	1.403 (2)
O3W—H32W	0.9000	C5—C6	1.376 (3)
O1W ⁱ —Na1—O3W ⁱⁱ	173.66 (6)	H41W—O4W—H42W	103.00
O1W—Na1—O3W	114.69 (6)	H51W—O5W—H52W	98.00

O1W—Na1—O1W ⁱ	86.32 (5)	H61W—O6W—H62W	116.00
O1W—Na1—O3W ⁱⁱ	100.02 (6)	C2—N1—C6	116.35 (16)
O2W—Na1—O3W	141.97 (6)	C4—N4—H41	118.00
O1W ⁱ —Na1—O2W	85.88 (6)	C4—N4—H42	118.00
O2W—Na1—O3W ⁱⁱ	92.71 (6)	H41—N4—H42	124.00
O1W—Na1—O2W	103.20 (6)	N1—C2—C3	123.14 (17)
O3W—Na1—O3W ⁱⁱ	84.24 (5)	N1—C2—C21	115.53 (16)
O1W ⁱ —Na1—O3W	93.06 (5)	C3—C2—C21	121.32 (16)
Na1—O1W—Na1 ⁱ	93.68 (5)	Cl3—C3—C2	119.28 (14)
Na1—O3W—Na1 ⁱⁱ	95.76 (6)	Cl3—C3—C4	119.39 (14)
Na1—O1W—H11W	100.00	C2—C3—C4	121.32 (16)
Na1 ⁱ —O1W—H11W	121.00	N4—C4—C3	122.49 (16)
Na1—O1W—H12W	128.00	N4—C4—C5	122.95 (17)
H11W—O1W—H12W	112.00	C3—C4—C5	114.55 (17)
Na1 ⁱ —O1W—H12W	103.00	Cl5—C5—C4	118.07 (14)
Na1—O2W—H22W	128.00	Cl5—C5—C6	121.71 (14)
H21W—O2W—H22W	97.00	C4—C5—C6	120.22 (16)
Na1—O2W—H21W	121.00	Cl6—C6—N1	115.37 (15)
Na1—O3W—H32W	119.00	Cl6—C6—C5	120.21 (14)
H31W—O3W—H32W	108.00	N1—C6—C5	124.42 (17)
Na1—O3W—H31W	109.00	O21—C21—O22	127.24 (18)
Na1 ⁱⁱ —O3W—H32W	115.00	O21—C21—C2	116.83 (17)
Na1 ⁱⁱ —O3W—H31W	109.00	O22—C21—C2	115.91 (17)
O3W ⁱⁱ —Na1—O3W—Na1 ⁱⁱ	0.00 (6)	N1—C2—C21—O21	89.1 (2)
O1W—Na1—O1W ⁱ —Na1 ⁱ	0.00 (6)	N1—C2—C3—Cl3	179.26 (15)
O2W—Na1—O1W ⁱ —Na1 ⁱ	-103.53 (6)	N1—C2—C3—C4	0.0 (3)
O3W—Na1—O1W ⁱ —Na1 ⁱ	114.57 (6)	C21—C2—C3—Cl3	0.7 (3)
O1W—Na1—O3W ⁱⁱ —Na1 ⁱⁱ	114.11 (6)	C3—C2—C21—O22	89.4 (2)
O2W—Na1—O3W ⁱⁱ —Na1 ⁱⁱ	-141.98 (6)	N1—C2—C21—O22	-89.3 (2)
O3W—Na1—O3W ⁱⁱ —Na1 ⁱⁱ	-0.02 (8)	C3—C2—C21—O21	-92.2 (2)
O2W—Na1—O3W—Na1 ⁱⁱ	87.03 (10)	C2—C3—C4—C5	0.1 (3)
O1W ⁱ —Na1—O3W—Na1 ⁱⁱ	174.25 (5)	Cl3—C3—C4—N4	0.6 (3)
O2W—Na1—O1W—Na1 ⁱ	84.89 (6)	Cl3—C3—C4—C5	-179.13 (15)
O3W—Na1—O1W—Na1 ⁱ	-91.68 (6)	C2—C3—C4—N4	179.9 (2)
O1W ⁱ —Na1—O1W—Na1 ⁱ	0.00 (5)	C3—C4—C5—Cl5	179.71 (15)
O3W ⁱⁱ —Na1—O1W—Na1 ⁱ	-179.91 (5)	N4—C4—C5—Cl5	-0.1 (3)
O1W—Na1—O3W—Na1 ⁱⁱ	-98.40 (6)	N4—C4—C5—C6	-180.0 (2)
C6—N1—C2—C21	178.63 (17)	C3—C4—C5—C6	-0.2 (3)
C2—N1—C6—Cl6	-179.62 (14)	C4—C5—C6—N1	0.2 (3)
C6—N1—C2—C3	0.0 (3)	Cl5—C5—C6—Cl6	-0.2 (3)
C2—N1—C6—C5	-0.1 (3)	Cl5—C5—C6—N1	-179.72 (15)
C21—C2—C3—C4	-178.59 (18)	C4—C5—C6—Cl6	179.74 (16)

Symmetry codes: (i) $-x, -y+1, -z+2$; (ii) $-x+1, -y+1, -z+2$.

Hydrogen-bond geometry (Å, °)

<i>D</i> —H··· <i>A</i>	<i>D</i> —H	H··· <i>A</i>	<i>D</i> ··· <i>A</i>	<i>D</i> —H··· <i>A</i>
O1 <i>W</i> —H11 <i>W</i> ···O4 <i>W</i>	0.82	1.97	2.786 (2)	178
O1 <i>W</i> —H12 <i>W</i> ···O6 <i>W</i>	0.89	1.99	2.877 (2)	174
O2 <i>W</i> —H21 <i>W</i> ···O22 ⁱⁱⁱ	0.94	1.79	2.721 (2)	171
O2 <i>W</i> —H22 <i>W</i> ···O5 <i>W</i>	0.84	1.98	2.816 (2)	172
O3 <i>W</i> —H31 <i>W</i> ···O5 <i>W</i> ^{iv}	0.88	1.93	2.781 (2)	163
O3 <i>W</i> —H32 <i>W</i> ···N1 ⁱⁱ	0.90	2.02	2.910 (2)	170
O4 <i>W</i> —H41 <i>W</i> ···O22 ^v	0.99	1.93	2.916 (2)	173
O4 <i>W</i> —H42 <i>W</i> ···O21	0.93	1.99	2.918 (2)	174
O5 <i>W</i> —H51 <i>W</i> ···O21	0.90	1.87	2.748 (2)	166
O5 <i>W</i> —H52 <i>W</i> ···O2 <i>W</i> ^{vi}	0.90	2.01	2.8481 (19)	155
O6 <i>W</i> —H61 <i>W</i> ···O22 ^{vii}	0.93	2.10	2.972 (2)	156
O6 <i>W</i> —H62 <i>W</i> ···C16	0.93	2.83	3.4400 (15)	124
O6 <i>W</i> —H62 <i>W</i> ···O21 ^{iv}	0.93	2.19	2.996 (2)	144
N4—H41···C13	0.96	2.54	2.9956 (17)	109
N4—H41···O6 <i>W</i> ^{viii}	0.96	2.18	3.080 (2)	156
N4—H42···C15	0.97	2.52	2.9690 (17)	108
N4—H42···O4 <i>W</i> ^{viii}	0.97	2.22	3.146 (2)	160

Symmetry codes: (ii) $-x+1, -y+1, -z+2$; (iii) $-x+1, -y, -z+2$; (iv) $x, y+1, z$; (v) $x-1, y, z$; (vi) $-x, -y, -z+2$; (vii) $x-1, y+1, z$; (viii) $-x+1, -y+1, -z+1$.