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## Hydrogen separation through tailored dual phase membranes with nominal composition $\text{BaCe}_{0.8}\text{Eu}_{0.2}\text{O}_{3-\delta}:\text{Ce}_{0.8}\text{Y}_{0.2}\text{O}_{2-\delta}$ at intermediate temperatures

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Hydrogen permeation membranes are a key element in improving the energy conversion efficiency and decreasing the greenhouse gas emissions from energy generation. The scientific community faces the challenge of identifying and optimizing stable and effective ceramic materials for  $\text{H}_2$  separation membranes at elevated temperature (400–800 °C) for industrial separations and intensified catalytic reactors. As such, composite materials with nominal composition  $\text{BaCe}_{0.8}\text{Eu}_{0.2}\text{O}_{3-\delta}:\text{Ce}_{0.8}\text{Y}_{0.2}\text{O}_{2-\delta}$  revealed unprecedented  $\text{H}_2$  permeation levels of 0.4 to 0.61 mL·min<sup>-1</sup>·cm<sup>-2</sup> at 700 °C measured on 500 μm-thick-specimen. A detailed structural and phase study revealed single phase perovskite and fluorite starting materials synthesized via the conventional ceramic route. Strong tendency of Eu to migrate from the perovskite to the fluorite phase was observed at sintering temperature, leading to significant Eu depletion of the proton conducting  $\text{BaCe}_{0.8}\text{Eu}_{0.2}\text{O}_{3-\delta}$  phase. Composite microstructure was examined prior and after a variety of functional tests, including electrical conductivity,  $\text{H}_2$ -permeation and stability in  $\text{CO}_2$  containing atmospheres at elevated temperatures, revealing stable material without morphological and structural changes, with segregation-free interfaces and no further diffusive effects between the constituting phases. In this context, dual phase material based on  $\text{BaCe}_{0.8}\text{Eu}_{0.2}\text{O}_{3-\delta}:\text{Ce}_{0.8}\text{Y}_{0.2}\text{O}_{2-\delta}$  represents a very promising candidate for  $\text{H}_2$  separating membrane in energy- and environmentally-related applications.

Hydrogen permeation membranes are a key element to reach high energy conversion efficiency and decreasing the greenhouse gas emissions from power generation and energy-intensive industries, i.e. by capturing and utilizing  $\text{CO}_2$  or moving towards hydrogen-based systems by extracting highly pure  $\text{H}_2$  from gas mixtures<sup>1–3</sup>. In this context, the integration of high performance  $\text{H}_2$ -permeation ceramic membranes with competitive manufacturing cost and life time (>3 years) in industrial processes can significantly intensify them by (i) maximizing products yield and energy efficiency and (ii) promoting the reaction rates<sup>4</sup>. Such integrative membrane approach of *in situ*  $\text{H}_2$  extraction (used as fuel) or consumption (as raw material for chemical production) in a proton conductor-based membrane reactor would shift the thermodynamic equilibrium towards the product side, hence boosting process efficiency, saving energy and reducing final product cost- all factors of high environmental and economic impact. Proton conducting ceramic membranes may therefore be of interest for integrated gasification process<sup>5,6</sup>, and more specifically in water gas shift reactors at elevated temperatures (600–900 °C)<sup>7</sup>, in catalytic membrane reactors (CMRs) for accomplishing chemical and petro-chemical reactions (e.g. ammonia synthesis<sup>8,9</sup>, non-oxidative de-hydrogenation reaction<sup>10</sup>).

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Although every specific application field sets particular requirements in terms of membrane stability, as a general rule, a good ceramic candidate has to remain phase and chemically stable at elevated temperatures in atmospheres possibly containing  $H_2$ ,  $H_2O$ ,  $CO$ ,  $CO_2$ ,  $CH_4$ ,  $SO_x$ ,  $H_2S$ , and different levels of ash. Such atmospheres represent extreme operating environments for the majority of  $H_2$  selective material classes, leading to material decomposition, membrane disintegration and performance degradation. Specifically, formation of undesired phases (carbonates, hydrates, sulfates) is critical for ceramic based membranes, while hydrogen-induced embrittlement and sulfur poisoning are particularly critical for Pd-based membranes<sup>7,11</sup>.

Properties like high selectivity for  $H_2$ , significant durability, mechanical and hydrothermal stability in reducing environments, as well as operating temperature ranges above 500 °C, make dense ceramic mixed proton-electronic conductors a promising alternative to precious metals and related alloys (Ag-Pd)<sup>12</sup>, polymers<sup>13</sup>, metal-organic frameworks (MOFs)<sup>14,15</sup>, zeolite<sup>16</sup> or other ceramic microporous materials<sup>17,18</sup>.

Apart from the stability issues which emerge to be solved, performance target for  $H_2$ -flux of 1–2 mL<sub>n</sub>·min<sup>-1</sup>·cm<sup>-2</sup> at 600–700 °C has been set for  $H_2$ -separation dense ceramic materials (based on the technical targets for dense metallic membranes for  $H_2$  separation and purification). However, these values have not been yet achieved.

In the context of above, scientific community still faces the challenge of identifying highly performing and at the same time stable mixed proton-electronic conducting ceramic materials and to explore and utilize them as membranes in a number of chemical reactions and  $H_2$  separation tasks at elevated temperatures.

In search for promising candidates for  $H_2$  separation applications, electrical and proton transport properties of large number of oxide ceramics from several structural classes have been explored with particular intensity in the last years. Hence, a wide range of state-of-the art perovskite-based proton conductors<sup>19,20</sup> to materials with less trivial structures and relatively novel to  $H_2$  extraction application, as defective fluorites<sup>21–26</sup>, pyrochlores<sup>27</sup>, fergusonites<sup>28–32</sup> has been covered. A comprehensive review on the electrical properties of a number of proton conductors could be found in ref. 33, while Table 1 presents a comparison of thickness-normalized  $H_2$  permeation flux values from literature. As it could be inferred from the table,  $H_2$ -permeation rates still remain lower than the milestone value at the target temperature.

The largest number of studies is however dedicated to the perovskite-based oxide ceramic materials, in particular with large lattice constant of the type  $SrCeO_3$ ,  $BaCeO_3$  and  $BaZrO_3$  (Table 1). Substituted  $BaCeO_3$  exhibits the highest levels of proton thermodynamic stability in the crystal structure remaining fully hydrated up to temperatures of about 600 °C but it is unstable in acidic gas environments due to its alkaline character<sup>19,66–68</sup>.

Apart from their good proton conduction, perovskite-based materials e.g. of the type  $BaCe_{1-x}M_xO_{3-\delta}$  exhibit also electronic conduction upon doping/substitution with M-cations with mixed valence<sup>40,42,69</sup>. However, the levels of electronic conduction are significantly low relative to the ionic conduction. Indeed, a raise in the level of electronic conductivity can be pursued by adding a second electronically conducting crystalline phase to the pristine proton conductor. As reported by Elangovan *et al.*<sup>61</sup>, a successful approach to obtain  $H_2$ -permeating material with high performance, is to combine (i) a protonic conducting perovskite phase  $BaCe_{1-x}M_xO_{3-\delta}$  and (ii) an electronic conducting fluorite phase  $Ce_{1-y}M_yO_{2-\delta}$  (M is metal dopant) into the so called *dual-phase ceramic material*, which at the end possesses increased ambipolar conductivity. The two crystalline phases form the *cer-cer composite* with certain degree of percolation, which provides efficient pathways for proton and electronic transport across the membrane. They are furthermore in a close contact at high temperature and reducing environments, therefore sufficient chemical and thermal compatibility between them is required. Elangovan *et al.*<sup>61</sup>, Ricote *et al.*<sup>70</sup>, Medvedev *et al.*<sup>71</sup> and Rebollo *et al.*<sup>63</sup> have shown the chemical compatibility and stability of systems consisting of  $BaCe_{1-x}Ln_xO_{3-\delta}$ : $Ce_{1-x}Ln_xO_{2-\delta}$  ( $x = 0.1–0.2$ ) and  $BaCe_{0.65}Zr_{0.2}Y_{0.15}O_{3-\delta}$ : $Ce_{0.85}Ln_{0.15}O_{2-\delta}$  ( $Ln = Y, Gd$ ), while Huang *et al.*<sup>72</sup> reported the appearance of minor additional phase in the grain boundaries when Zr is doped in the perovskite phase. Furthermore, the addition of the doped ceria phase would enhance the stability of otherwise easily attacked by  $CO_2$  and  $H_2O$ -containing environments  $BaCeO_3$  phase due to a shift in thermodynamic equilibrium towards the reactant side<sup>73</sup>. This leads to overall suppression of  $BaCeO_3$  decomposition to  $BaCO_3$ ,  $Ba(OH)_2$  and  $CeO_2$ .

In the present work composite ceramic membrane material with nominal composition  $BaCe_{0.8}Eu_{0.2}O_{3-\delta}$ : $Ce_{0.8}Y_{0.2}O_{2-\delta}$  (labeled BCEO:CYO) at 50:50 vol.% ratio was explored as  $H_2$  permeation membrane in the temperature range 600–700 °C. This ratio was selected to ensure the largest degree of interaction between the two phases and sufficient transport pathways both for protons and electrons. In the context of above, acceptor substitution with 20 mol. %  $Eu^{3+}$  in  $BaCeO_3$  and the same amount of  $Y^{3+}$  in  $CeO_2$  was undertaken as a strategy to pursue an improvement and fine-tuning of transport properties of the composite ceramic material.

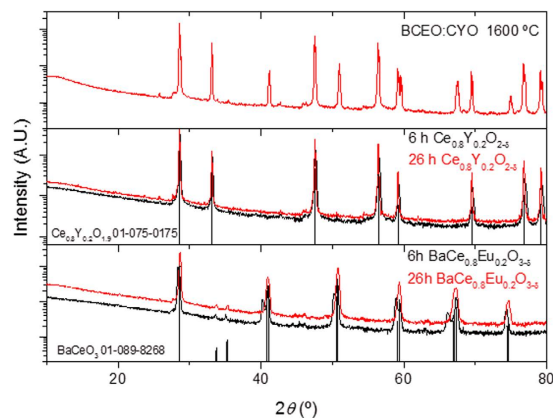
For the selection of the two substituents several factors were accounted.  $BaCeO_3$  doped by Gd from rare earths (RE) in addition to Y, exhibits the highest ionic (protonic) conductivity among solid solutions  $BaCe_{1-x}RE_xO_{3-\delta}$ <sup>74–76</sup>.  $Eu^{3+}$  ( $4f^6$ ) and  $Gd^{3+}$  ( $4f^7$ ) have the same valence state, nearly the same ionic radii and good geometrical compatibility with Ce as B-site ion in Ba-cerate. The ionic radius increases in the row  $Gd \rightarrow Eu \rightarrow Sm$  and decreased protonic conductivity has been found for Sm due to the increased lattice distortion. However, similar lattice distortions and conductivities are expected to be produced by  $Eu^{3+}$  and  $Gd^{3+}$  in  $BaCeO_3$ . As demonstrated elsewhere<sup>77</sup>, slightly higher values of the total conductivity in  $H_2$  environments were determined for  $BaCe_{0.85}Gd_{0.15}O_{3-\delta}$  in comparison with  $BaCe_{0.85}Eu_{0.15}O_{3-\delta}$ . Furthermore, there was practically no difference in conductivities under  $H_2/N_2$  and air atmospheres measured for  $BaCe_{0.85}Eu_{0.15}O_{3-\delta}$ , while for  $BaCe_{0.85}Tb_{0.15}O_{3-\delta}$  largely higher values were recorded under reducing environments. According to Radojkovic<sup>78</sup>,  $Eu^{3+}$  substitution in  $BaCeO_3$  leads to clear advantages over that with  $Y^{3+}$  or  $Gd^{3+}$  reported as the best proton conductors. One may highlight (i) the lower sintering temperature (reported to be below 1500 °C); (ii) the larger conductivity under wet  $H_2$  conditions ( $1.21 \cdot 10^{-2}$  S·cm<sup>-1</sup> at 600 °C) due to the larger unit cell volume and lattice distortion favoring intra-octahedral proton migration of lower activation energy; (iii) improved grain boundary conductivity ascribed to a decrease

Material	Ref.	Gas atmosphere Feed – Sweep	T (°C)	$j_{H_2, norm}$ ( $mL \cdot min^{-1} \cdot cm^{-1}$ )
<i>Defective fluorite structured materials</i>				
La <sub>5.5</sub> WO <sub>11.25-δ</sub>	25	wet 50% H <sub>2</sub> in He – wet Ar	750/900	1.5·10 <sup>-3</sup> /4.7·10 <sup>-3</sup>
La <sub>5.5</sub> W <sub>0.8</sub> Mo <sub>0.2</sub> O <sub>11.25-δ</sub>	25	wet 50% H <sub>2</sub> in He – wet Ar	700	2.7·10 <sup>-3</sup>
La <sub>5.5</sub> W <sub>0.8</sub> Re <sub>0.2</sub> O <sub>11.25-δ</sub>	25	wet 50% H <sub>2</sub> in He – wet Ar	700	5.9·10 <sup>-3</sup>
Nd <sub>5.5</sub> WO <sub>11.25-δ</sub>	34	wet 20% H <sub>2</sub> in He – wet Ar	1000	1.2·10 <sup>-3</sup>
(Nd <sub>5/6</sub> La <sub>1/6</sub> ) <sub>5.5</sub> WO <sub>12-δ</sub>	35	wet 50% H <sub>2</sub> in He – wet Ar	900	1.2·10 <sup>-3</sup>
(La <sub>5/6</sub> Nd <sub>1/6</sub> ) <sub>5.5</sub> WO <sub>12-δ</sub>	36	wet 50% H <sub>2</sub> in He – wet Ar	900	1.4·10 <sup>-3</sup>
Nd <sub>5.5</sub> W <sub>0.5</sub> Mo <sub>0.5</sub> O <sub>11.25-δ</sub>	37	wet 50% H <sub>2</sub> in He – wet Ar	900	7.5·10 <sup>-3</sup>
Nd <sub>5.5</sub> W <sub>0.5</sub> Re <sub>0.5</sub> O <sub>11.25-δ</sub>	38	wet 50% H <sub>2</sub> in He – wet Ar	900	4.1·10 <sup>-3</sup>
<i>Perovskite structured materials</i>				
BaCe <sub>0.80</sub> Y <sub>0.10</sub> Ru <sub>0.10</sub> O <sub>3-δ</sub>	39	wet H <sub>2</sub> in Ar – Ar	800	4.3·10 <sup>-3</sup>
BaCe <sub>0.95</sub> Nd <sub>0.05</sub> O <sub>3-δ</sub>	40	wet 80% H <sub>2</sub> in He – dry Ar + Ne	900	1.3·10 <sup>-3</sup>
BaZr <sub>0.80</sub> Y <sub>0.15</sub> Mn <sub>0.05</sub> O <sub>3-δ</sub>	41	wet 50% H <sub>2</sub> in He – wet Ar	900	1.4·10 <sup>-3</sup>
SrCe <sub>0.95</sub> Tm <sub>0.05</sub> O <sub>3-δ</sub>	42	10% H <sub>2</sub> in He – air	700/900	1.2·10 <sup>-3</sup> /1.9·10 <sup>-3</sup>
SrCe <sub>0.95</sub> Tm <sub>0.05</sub> O <sub>3-δ</sub>	43	10% H <sub>2</sub> in He – air	700/900	2.6·10 <sup>-3</sup> /6.4·10 <sup>-3</sup>
SrCe <sub>0.95</sub> Tm <sub>0.05</sub> O <sub>3-δ</sub>	44	10% H <sub>2</sub> dry in He – 20% O <sub>2</sub> in Ar	750/900	4.3·10 <sup>-3</sup> /6.8·10 <sup>-3</sup>
SrCe <sub>0.95</sub> Tb <sub>0.05</sub> O <sub>3-δ</sub>	45	20% H <sub>2</sub> in He – CO in Ar	750/900	4.0·10 <sup>-4</sup> /1.6·10 <sup>-3</sup>
SrCe <sub>0.95</sub> Y <sub>0.05</sub> O <sub>3-δ</sub>	46	80% H <sub>2</sub> in He – Ar	900	3.2·10 <sup>-4</sup>
SrCe <sub>0.95</sub> Yb <sub>0.05</sub> O <sub>3-δ</sub>	47	10% H <sub>2</sub> in He – air	677	2.7·10 <sup>-3</sup>
SrCe <sub>0.95</sub> Eu <sub>0.05</sub> O <sub>3-δ</sub>	48	100% H <sub>2</sub> – He	700/850	2.3·10 <sup>-4</sup> /7.4·10 <sup>-4</sup>
SrCe <sub>0.95</sub> Sm <sub>0.05</sub> O <sub>3-δ</sub>	48	100% H <sub>2</sub> – He	850	5.4·10 <sup>-4</sup>
SrCe <sub>0.75</sub> Zr <sub>0.20</sub> Tm <sub>0.05</sub> O <sub>3-δ</sub>	49	H <sub>2</sub> in He – wet Ar	900	5.0·10 <sup>-3</sup>
SrCe <sub>0.75</sub> Zr <sub>0.20</sub> Tm <sub>0.05</sub> O <sub>3-δ</sub>	44	10% H <sub>2</sub> dry in He – 20% O <sub>2</sub> in Ar	750/900	8.0·10 <sup>-4</sup> /2.4·10 <sup>-3</sup>
SrCe <sub>0.70</sub> Zr <sub>0.25</sub> Ln <sub>0.05</sub> O <sub>3-δ</sub> (Ln = Tm, Yb)	50	wet 20% H <sub>2</sub> – wet sweep (not specified)	900	2.3·10 <sup>-4</sup>
SrCe <sub>0.65</sub> Zr <sub>0.20</sub> Eu <sub>0.15</sub> O <sub>3-δ</sub>	51	100% H <sub>2</sub> – He	900	8.5·10 <sup>-4</sup>
SrZr <sub>0.95</sub> Y <sub>0.05</sub> O <sub>3-δ</sub>	47	20% H <sub>2</sub> – air in He	700	<2.3·10 <sup>-5</sup>
<i>Cer-Met dual phase materials</i>				
BaCe <sub>0.95</sub> Tb <sub>0.05</sub> O <sub>3-δ</sub> :Ni (50:50 wt.%)	52	50% H <sub>2</sub> in N <sub>2</sub> – He	850	8.3·10 <sup>-3</sup>
BaCe <sub>0.90</sub> Y <sub>0.10</sub> O <sub>3-δ</sub> :Ni (60:40 vol.%)	53	4% H <sub>2</sub> in He – 100 ppm H <sub>2</sub> in N <sub>2</sub>	800	1.7·10 <sup>-2</sup>
BaCe <sub>0.80</sub> Y <sub>0.20</sub> O <sub>3-δ</sub> :Ni (60:40 vol.%)	54	3.8% H <sub>2</sub> in N <sub>2</sub> – 100 ppm H <sub>2</sub> in N <sub>2</sub>	900	2.0·10 <sup>-3</sup>
BaCe <sub>0.85</sub> Zr <sub>0.10</sub> Tb <sub>0.05</sub> O <sub>3-δ</sub> :Ni (50:50 wt.%)	55	50% H <sub>2</sub> in 50% He – Ar	800	8.5·10 <sup>-3</sup>
BaCe <sub>0.70</sub> Zr <sub>0.10</sub> Y <sub>0.20</sub> O <sub>3-δ</sub> :Ni (60:40 vol.%)	56	4% H <sub>2</sub> in He – 100 ppm H <sub>2</sub> in N <sub>2</sub>	900	5.6·10 <sup>-3</sup>
BaCe <sub>0.70</sub> Zr <sub>0.10</sub> Y <sub>0.10</sub> Yb <sub>0.10</sub> O <sub>3-δ</sub> :Ni (60:40 vol.%)	57	20% H <sub>2</sub> wet, 60% CO <sub>2</sub> , 20% He – N <sub>2</sub>	900	3.5·10 <sup>-3</sup>
Ce <sub>0.50</sub> La <sub>0.4875</sub> Ca <sub>0.0125</sub> O <sub>2-δ</sub> :Ni (60:40 vol.%)	58	wet 20% H <sub>2</sub> , 77% N <sub>2</sub> – Ar	900	1.5·10 <sup>-3</sup>
YSZ:Pd (40:60 vol.%)	59	90% H <sub>2</sub> in He – N <sub>2</sub>	400/900	4.7·10 <sup>-2</sup> /9.4·10 <sup>-2</sup>
<i>Cer-Cer dual phase materials</i>				
La <sub>5.5</sub> WO <sub>11.25-δ</sub> :La <sub>0.87</sub> Sr <sub>0.13</sub> CrO <sub>3-δ</sub> (50:50 vol.%)	60	wet 50% H <sub>2</sub> in He – wet Ar	700	5.5·10 <sup>-3</sup>
BaCe <sub>0.80</sub> Eu <sub>0.20</sub> O <sub>3-δ</sub> :Ce <sub>0.80</sub> Eu <sub>0.20</sub> O <sub>2-δ</sub> (50:50 vol.%)	61	H <sub>2</sub> , CH <sub>4</sub> , H <sub>2</sub> O, CO, CO <sub>2</sub> – He	900	<7.0·10 <sup>-2</sup>
BaCe <sub>0.80</sub> Y <sub>0.20</sub> O <sub>3-δ</sub> :Ce <sub>0.80</sub> Y <sub>0.20</sub> O <sub>2-δ</sub> (50:50 wt.%)	62	wet 50% H <sub>2</sub> in He – wet Ar	900	1.1·10 <sup>-2</sup>
BaCe <sub>0.65</sub> Zr <sub>0.20</sub> Y <sub>0.15</sub> O <sub>3-δ</sub> :Ce <sub>0.85</sub> Gd <sub>0.15</sub> O <sub>2-δ</sub> (50:50 vol.%)	63	wet 50% H <sub>2</sub> in He – wet Ar	755	1.76·10 <sup>-2</sup>
BaCe <sub>0.20</sub> Zr <sub>0.70</sub> Y <sub>0.10</sub> O <sub>3-δ</sub> :Sr <sub>0.95</sub> Ti <sub>0.90</sub> Ni <sub>0.10</sub> O <sub>3-δ</sub> (50:50 vol.%)	64	9% H <sub>2</sub> in He – dry Ar	800	1.1·10 <sup>-3</sup>
SrZrO <sub>3</sub> :SrFeO <sub>3</sub> (80:20 vol.%)	65	H <sub>2</sub> in He – wet Ar	900	4.8·10 <sup>-3</sup>

**Table 1.** Based on literature summary of thickness ( $L_{mem}$ )-normalized values of H<sub>2</sub> permeation rates  $j_{H_2, norm}$  through ceramic membranes.

in band gap energy related to the transition between O 2p valence band to Ce 4f conduction band; and (iv) dominating proton conductivity under OCV conditions at 650 °C and 700 °C.

On the other hand, in reducing atmospheres, Y-substituted CeO<sub>2</sub> exhibits predominantly electronic conduction via small polaron hopping mechanism due to Ce<sup>4+/3+</sup> variable valence. Besides, detailed structural and phase characterization via Rietveld refinement on XRD will help to evaluate the phase formation and phase stability of the cer-cer composite supported by microstructural study via SEM and HR-TEM. Electrical conductivity properties of the composite material and its constituting phases were studied to elucidate the relative contributions of each phase in the composite. In addition, hydrogen permeation of the composite membrane with thickness of 500 μm was characterized as a function of the temperature, the hydration conditions and the hydrogen concentration in the feed. Membrane stability in CO<sub>2</sub>-containing atmospheres was studied *in situ* with H<sub>2</sub> permeation measurements to conclude on eventual phase, microstructural and performance degradation effects.



**Figure 1.** XRD diffraction patterns of BCEO (bottom); CYO (middle) synthesized by solid-state reaction at 1400 °C for 6 and 26 h; and the BCEO:CYO (top) composite pattern sintered for 10 h at 1600 °C.

## Results and Discussion

**Structural and phase composition study.** Figure 1 shows the XRD patterns of BCEO and CYO synthesized via the conventional solid-state route at 1400 °C, and the composite pattern resulting from mixing them in a 50:50 vol.% ratio and sinter for 10 h at 1600 °C. Rietveld refinement on the XRD performed on the BCEO and CYO separately showed that the single phases were not formed after 6 h of calcination. XRD patterns revealed a mixture of tetragonal and orthorhombic perovskite phases for BCEO and two fluorite phases with different Y content (details are summarized in the Table 2). An extra thermal treatment of 20 h (in total 26 h) allows single phase samples with sharp defined peaks to be obtained. Besides, once the pressed pellet is sintered at 1600 °C, the observed splitting of peaks disappear and XRD reveals that a single BCEO perovskite phase and a single CYO fluorite were formed. The presence of other oxides was not observed.

As it can be observed from Table 2, the lattice parameters of the two phases constituting the dual phase ceramics disclosed certain deviation from the nominal stoichiometry. The ceria phase shows a cubic fluorite structure with a lattice parameter of 5.411 (1) Å, clearly shifted from the theoretical value of 5.405 Å estimated by Kim's empirical formula<sup>79</sup> for 20 mol.% yttria substitution in CeO<sub>2</sub> (and from the 5.405 Å obtained experimentally for the single doped ceria phase). The calculation took into account the ionic radii in 8-fold coordinated Ce<sup>4+</sup> and Y<sup>3+</sup> cations, which are 0.97 and 1.019 Å, respectively. Since the ionic radii of Eu<sup>3+</sup> is 1.066 Å, the substitution of Ce<sup>4+</sup> cations by Eu<sup>3+</sup> is suggested. By using the experimental cell parameter obtained for the fluorite phase, Kim's formula can be reversely used to estimate to which extent Eu migrated to the ceria phase. As there are not yttrium related impurities, the yttrium content in the fluorite was fixed to 20 mol.%, which gives a value of  $x = 10.5$  mol.% for Eu<sup>3+</sup> and corresponds to stoichiometry Ce<sub>0.695</sub>Y<sub>0.2</sub>Eu<sub>0.105</sub>O<sub>2.6</sub>. This value indicates that practically all the Eu introduced by the perovskite phase has migrated to the fluorite phase. The mass balance by ICP-OES (Supplementary information, Table S1) indicates that the barium was not deficient after the sintering (see Experimental Section, Sample preparation paragraph). The unit cell volumes of pure BaCeO<sub>3</sub>, BaCe<sub>0.9</sub>Eu<sub>0.1</sub>O<sub>2.95</sub> according to<sup>78</sup>, BaCe<sub>0.8</sub>Eu<sub>0.2</sub>O<sub>2.9</sub> after 26 h at 1400 °C and for Ba-cerate phase in the composite material BaCe<sub>0.8</sub>Eu<sub>0.2</sub>O<sub>2.9</sub>:Ce<sub>0.8</sub>Y<sub>0.2</sub>O<sub>1.9</sub> after 10 h at 1600 °C are respectively: 340 Å<sup>3</sup>, 341.79 Å<sup>3</sup>, 342.4 Å<sup>3</sup> and 340.73 Å<sup>3</sup>, showing that (i) with increasing the Eu amount the volume expands as expected; and (ii) Ba-cerate in the composite is practically Eu-depleted.

**Microstructural characterization.** FE-SEM analysis of the polished cross-section of the BCEO:CYO membrane revealed a high relative density (>98%), although some occluded porosity is still present (Fig. 2a–f). The thermal expansion coefficient (TEC) of the composite material in air ( $12.4 \cdot 10^{-6} \text{ K}^{-1}$ ) shows identical behavior during heating and cooling cycles in air (Supplementary information, Figure S2). No cracks related to the TEC thermochemical cycles were detected due to the good thermal compatibility of the two ceramic components. BSE-SEM (Fig. 2c–e) and EDS analysis (Fig. 2f) show dark grains corresponding to the CYO phase, as well as a bright phase corresponding to BCEO. In addition, Fig. 2c,f shows the EDS linescan analysis of a fracture cross-section of the membrane and the corresponding Ba and Y concentration profiles. Neither Ba (in Fig. 2f the blue line) nor Y (in Fig. 2f the red line) interdiffusion between both phases occurs. The topography and distinct textures observed in the polished cross-section for each phase originates from the difference in hardness, i.e., the dark phase (CeO<sub>2</sub>) is harder, and this effect was previously reported for BaCeO<sub>3</sub>:CeO<sub>2</sub> composites<sup>70</sup>.

Further detailed insight in composite microstructure was achieved by the TEM and high resolution scanning transmission electron microscope with high-angle annular dark-field imaging (HRSTEM-HAADF) (Fig. 3). A low-magnification TEM image acquired from the composite sample as-sintered at 1600 °C is shown in Fig. 3a (Composite sample, nominally labeled BCEO:CYO, was actually named BaCeO:CeYEuO in the TEM image after its actual composition with Eu predominantly located in the CYO phase). The sample is composed of randomly oriented, well-connected and packed grains with sizes ranging from 0.5 μm to 2 μm. The grains have a polygonal shape, which minimizes their surface energy. Two types of grains were observed: some grain's interiors were decorated with dotty/fuzzy contrast, while other grains surface was even. Furthermore, dotty contrast inside grains is sensitive to the grains relative orientation with respect to the electron beam. The slight tilt of the grain

Nominal starting compounds	Phase composition	Space group	Lattice parameter(s), (Å)	Phase wt. %	Phase $D_{th}$ (g/cm <sup>3</sup> )	Rwp, %
BaCe <sub>0.8</sub> Eu <sub>0.2</sub> O <sub>2.9</sub> 1400 °C/6h	BaCe <sub>0.8</sub> Eu <sub>0.2</sub> O <sub>2.9</sub> tetragonal	P-42m	a = 4.388 (1) c = 4.465 (1)	86	6.30	6.7
	BaCe <sub>0.8</sub> Eu <sub>0.2</sub> O <sub>2.9</sub> orthorhombic	Pmcn	a = 6.245 (1) b = 8.770 (1) c = 6.272 (1)	14	6.31	
BaCe <sub>0.8</sub> Eu <sub>0.2</sub> O <sub>2.9</sub> 1400 °C/26h	BaCe <sub>0.8</sub> Eu <sub>0.2</sub> O <sub>2.9</sub> orthorhombic	Pmcn	a = 6.229 b = 8.782 c = 6.257 (1)	100	6.35	6.6
Ce <sub>0.8</sub> Y <sub>0.2</sub> O <sub>1.9</sub> 1400 °C/6h	Ce <sub>0.9</sub> Y <sub>0.1</sub> O <sub>1.95</sub> cubic	Fm-3m	a = 5.411(1)	45	6.93	5.3
	Ce <sub>0.8</sub> Y <sub>0.2</sub> O <sub>1.9</sub> cubic	Fm-3m	a = 5.401(1)	55	6.75	
Ce <sub>0.8</sub> Y <sub>0.2</sub> O <sub>1.9</sub> 1400 °C/26h	Ce <sub>0.8</sub> Y <sub>0.2</sub> O <sub>1.9</sub> cubic	Fm-3m	a = 5.405(1)	100	6.8	5.9
BaCe <sub>0.8</sub> Eu <sub>0.2</sub> O <sub>2.9</sub> ; Ce <sub>0.8</sub> Y <sub>0.2</sub> O <sub>1.9</sub> 1600 °C/10h	BaCe <sub>1-x</sub> Eu <sub>x</sub> O <sub>3-δ</sub> orthorhombic	Pmcn	a = 8.786(1) b = 6.245(1) c = 6.210(1)	46	6.34	5.4
	Ce <sub>0.695</sub> Y <sub>0.2</sub> Eu <sub>0.105</sub> O <sub>2.8</sub> cubic	Fm-3m	a = 5.411(1)	54	6.77	

**Table 2. Rietveld refinement results from the XRD patterns of BCEO and CYO synthesized via the conventional solid-state route at 1400 °C during 6 h and 20 h (in total 26 h), and of BCEO:CYO dual-phase pellet (50:50 vol.%) sintered at 1600 °C for 10 h. For reference, the lattice parameters of pure BaCeO<sub>3</sub> are a = 8.776(1) Å, b = 6.234(1) Å, c = 6.214(1) Å.**

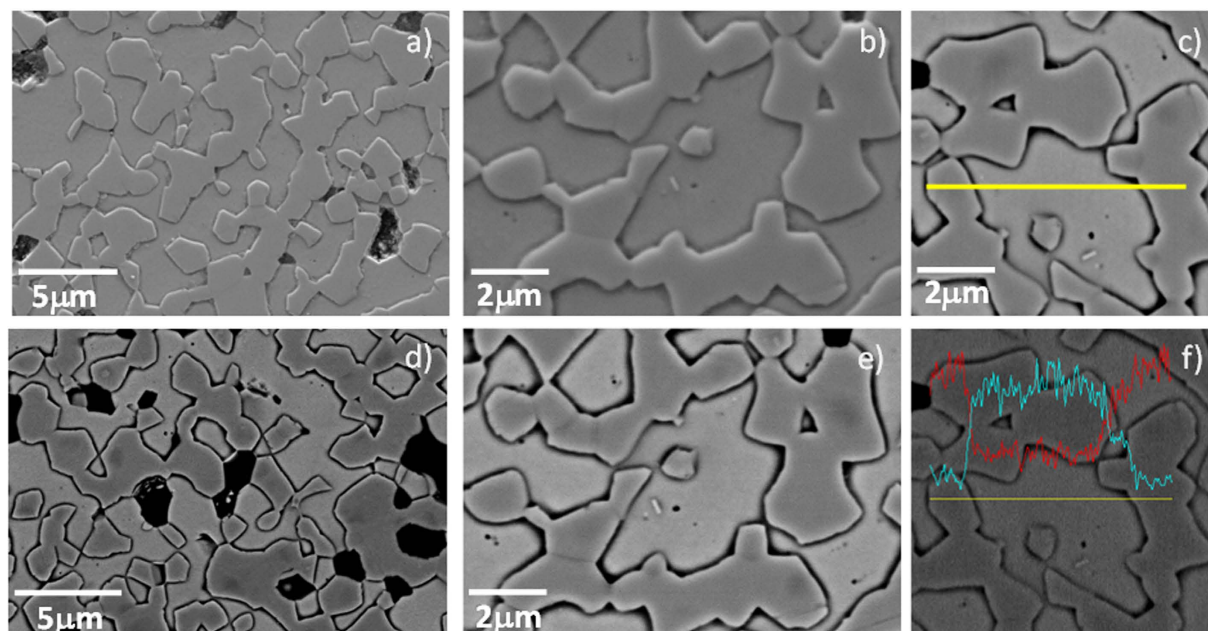
could enhance or diminish the dotted contrast significantly (Fig. 3b,c). HR-TEM investigation does not reveal any interface phases at the grain boundaries (not shown). STEM-HAADF images acquired using strong elemental contrast (Z-contrast) conditions are shown in Fig. 3d. The dotted/fuzzy contrast is evidenced under such imaging conditions as well. EDX analysis (Fig. 3d,e) revealed that dotted/fuzzy contrast rich grains are composed of Ce, Y, Eu and O elements, while other grains consist of Ba, Ce and O. The dotted pattern contrast might be attributed to Eu presence in CYO grains, which is associated with high point-defect density caused by incorporation of Eu into this phase. Furthermore, no migration of Y from the fluorite to the perovskite could be evidenced (Fig. 3(e)). HR-STEM-HAADF images recorded from the two grain types in range of crystallographic projections exhibit single crystal structure as shown in Fig. 3f. Dotted contrast is absent in HRSTEM-HAADF imaged for CYO (Ce-Y-Eu-O) grain due to channeling contrast mechanisms dominant at those imaging conditions. By correlating selected area electron diffraction (SAED) and HRSTEM-HAADF images with crystal structure data, the CYO (actually CYEuO) grains showed Fm-3m space group, while BCEO (actually BCO) grains exhibit Pmcn space group crystal structure, both in agreement with the Rietveld refinement made on the XRD patterns. As a summary, the microstructural analysis confirms that Eu diffused from the BCEO perovskite to the CYO fluorite when exposed to high temperature, as previously shown in the XRD section with no evidence for Y diffusion from the fluorite to the perovskite.

Since the composition and microstructure of both phases is pure, without grain boundary segregations electrical and electrochemical study was performed in order to evaluate the transport properties of this dual-phase material.

**Electrochemical characterization.** Figure 4 displays the total conductivity corresponding to the single phase materials BaCe<sub>0.8</sub>Eu<sub>0.2</sub>O<sub>3-δ</sub> (BCEO) and Ce<sub>0.8</sub>Y<sub>0.2</sub>O<sub>2-δ</sub> (CYO), and the composite material BCEO:CYO with nominal composition BaCe<sub>0.8</sub>Eu<sub>0.2</sub>O<sub>3-δ</sub>:Ce<sub>0.8</sub>Y<sub>0.2</sub>O<sub>2-δ</sub>. The samples have been measured as a function of inverse temperature under H<sub>2</sub>, H<sub>2</sub>+H<sub>2</sub>O, D<sub>2</sub> and D<sub>2</sub>+D<sub>2</sub>O atmospheres. Within the studied range of temperatures (from 800 °C to 400 °C), BCEO exhibits mainly protonic conductivity, ascertained from the H/D isotopic and hydration effect, i.e. lower conductivity in D<sub>2</sub> than in H<sub>2</sub> atmospheres and higher conductivity in wet than in dry atmospheres. Above 750 °C, the different conductivity curves converge due to the progressive oxide dehydration with temperature, leading to significant drop in proton concentration. On the other hand, CYO possesses higher conductivity in dry atmospheres than in wet conditions (wet atmospheres signify less reducing conditions than dry atmospheres) suggesting that electronic transport prevails under these conditions. Finally, BCEO:CYO composite shows prevailing *n*-type conduction behaviour and conductivity values similar to CYO, indicating proper percolation of this electron conducting component, as previously suggested by the SEM analysis.

In order to disclose the predominant transport of these compounds, *p*O<sub>2</sub> effect on the conductivity was also studied under wet reducing conditions (Fig. 5). The BCEO behavior as a prevailing proton conductor deduced from the observed H/D isotopic effect is confirmed by the relationship  $\sigma \propto pO_2^0$  in all studied temperatures i.e., the conductivity is independent on the *p*O<sub>2</sub> values. On the other hand, CYO exhibits predominant *n*-type electronic conduction following a power law of  $\sigma \propto pO_2^{-1/4}$  as it has been reported previously for CYO material<sup>80,81</sup>. Finally, from the relationship  $\sigma \propto pO_2^{-1/6}$  observed for the composite BCEO:CYO, it can be assumed an intermediate behaviour, i.e. it is mainly *n*-type electronic conductor but presents a significant ionic contribution, arising from the physical mixture of the two different conductors.

Temperature programmed reduction (TPR) measurements (Figure S3 in Supplementary Information) were performed in order to study the redox behavior of the single compounds and the composite, and to correlate with the conductivity data. The CYO sample shows the maximum of its reduction peak at 725 °C, ascribed to the reducible bulk Ce<sup>4+</sup>. This reduction temperature agrees with previous studies on doped CeO<sub>2</sub><sup>82</sup>. Regarding BCEO, a broad reduction peak is observed between 300–700 °C, which can be attributed to the Ce<sup>4+</sup> reduction and some oxygen release. The reduction of Eu<sup>3+</sup> to Eu<sup>2+</sup> is not expected in the considered temperature range as it was observed previously for Eu doped NdWO compound<sup>83</sup>. Finally, the Ce<sup>4+</sup> reduction in the sintered BCEO:CYO samples is shifted to higher temperatures presenting its maximum at 800 °C. The lower reducibility of Ce<sup>4+</sup> in this



**Figure 2.** (a,b) SEM, (c–e) BSE-SEM and (f) EDS linescan analysis of the fractured cross-section of the BCEO:CYO membrane sintered at 1600 °C.

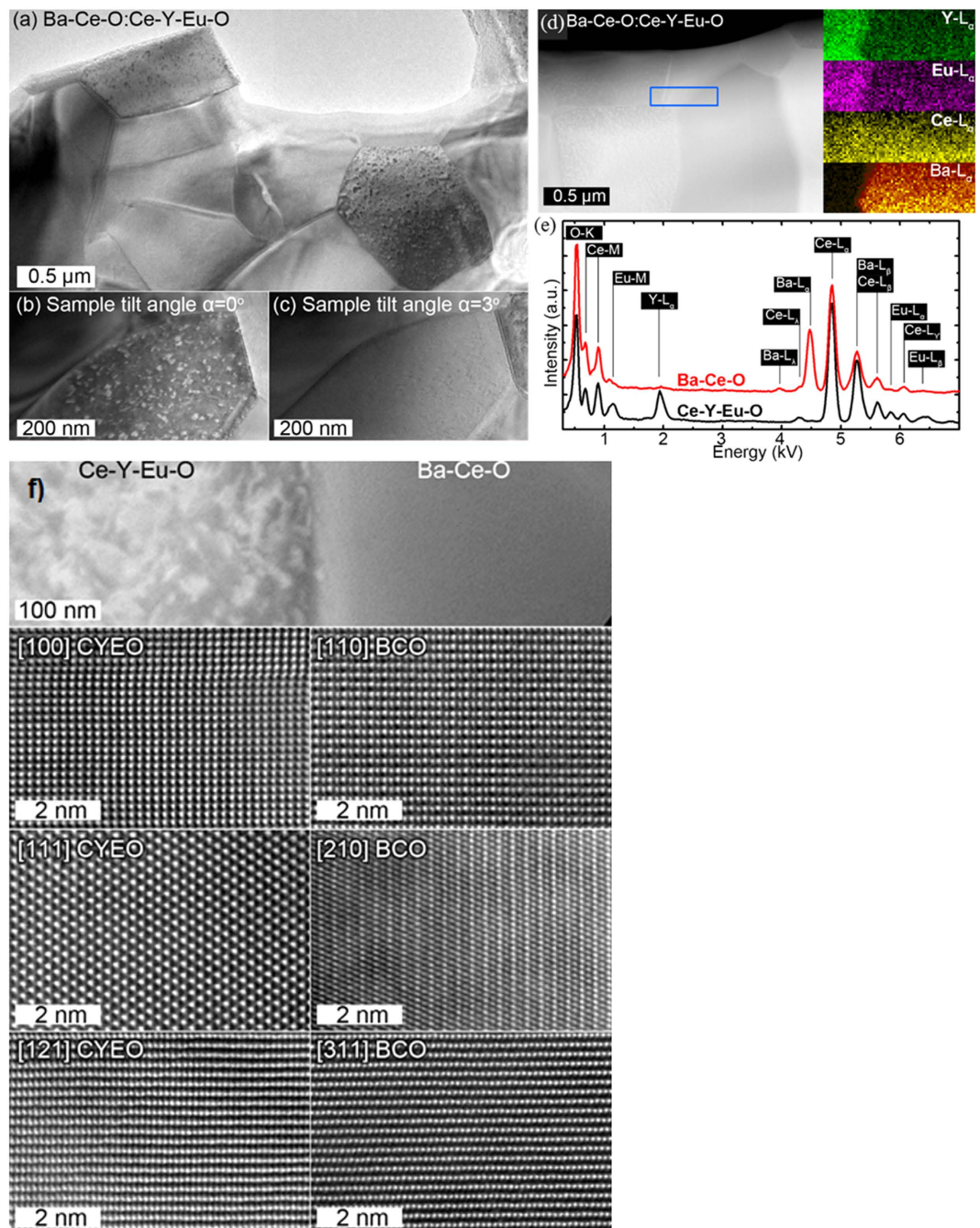
sample could be ascribed to the different stoichiometry of the BCEO and doped ceria grains as compared with the separate single material, as this is previously stated for instance in heavily doped cerias<sup>84</sup>.

**Hydrogen permeation measurements.** Permeation properties were thoroughly studied by analyzing the influence of the environment humidification and temperature on the H<sub>2</sub> permeation. Three different configurations, depicted in Fig. 6a, were selected in order to study the effect of humidification: (A) only feed side humidified, (B) both membrane sides humidified and (C) only sweep side humidified. 50 vol.% H<sub>2</sub> in He was employed as feed gas and Ar was used as sweep gas. Figure 6b plots the H<sub>2</sub> permeation behavior through the composite membrane as a function of the temperature in the three configurations. When only the feed side is humidified (configuration A), H<sub>2</sub> flow is below 0.05 mL·min<sup>-1</sup>·cm<sup>-2</sup> due to the low concentration of protons as a result of the large pH<sub>2</sub>O gradient across the membrane. When both sides are humidified (configuration B), the obtained H<sub>2</sub> flow is one order of magnitude higher than in configuration A (0.4 mL·min<sup>-1</sup>·cm<sup>-2</sup> at 700 °C). This important increase is attributed to the faster proton transport through the membrane, as a consequence of the higher hydration of the material and hence the higher concentration of protonic defects<sup>63</sup>. In addition to proton transport, H<sub>2</sub> is also generated at the sweep side by water splitting reaction mediated by the oxygen ion diffusion from the high pO<sub>2</sub> side (Ar sweep) to the lower pO<sub>2</sub> side (H<sub>2</sub> feed). The oxygen transport takes place preferentially through the CYO phase, since this phase presents significant oxygen ion conductivity at elevated temperatures<sup>80</sup>, although it is still one order of magnitude lower than the electronic conductivity in reducing conditions. Finally, for humidified sweep gas (configuration C), the H<sub>2</sub> flow increases further, reaching values of 0.61 mL·min<sup>-1</sup>·cm<sup>-2</sup> at 700 °C. This increase in H<sub>2</sub> production stems from the higher magnitude of water splitting process since a higher pO<sub>2</sub> gradient is imposed across the membrane as compared to configuration B.

H<sub>2</sub> permeation was also studied as a function of the hydrogen concentration (pH<sub>2</sub>) in the feed side (Fig. 6c,d). Irrespective of the humidification conditions, H<sub>2</sub> flow rises with increasing pH<sub>2</sub> as it is postulated by the Wagner equation, which describes the transport of both protons and oxygen ions. Furthermore, H<sub>2</sub> permeation was investigated by using lower pH<sub>2</sub>O at the sweep side, pH<sub>2</sub>O = 0.0094 atm instead of 0.042 atm. Results are shown in Fig. 7a–c. H<sub>2</sub> flows obtained at 700 °C sweeping pH<sub>2</sub>O = 0.042 atm are 19% and 8% higher than the values obtained for pH<sub>2</sub>O = 0.0094 atm, in configurations B and C, respectively (Fig. 7a). Assuming that the partial conductivities are constant in the studied pO<sub>2</sub> and pH<sub>2</sub>O ranges, and that H<sub>2</sub> transport can be described by Wagner's, the H<sub>2</sub> flow could be expressed as it is indicated in eq. 1, where the first term corresponds to the H<sub>2</sub> permeating through the membrane and the second one responds to the H<sub>2</sub> produced by water splitting.

$$J_{H_2} \approx \frac{RT}{4F^2L} \frac{(\sigma_{H^+} \cdot \sigma_{e^-})}{\sigma_T} \ln \frac{p(H_2, feed)}{p(H_2, perm)} + \frac{RT}{8F^2L} \frac{(\sigma_{O^{2-}} \cdot \sigma_{e^-})}{\sigma_T} \ln \frac{p(O_2, perm)}{p(O_2, feed)} \quad (1)$$

With pH<sub>2</sub>O = 0.042 atm,  $\ln \frac{p(O_2, perm)}{p(O_2, feed)}$  is 1.285 and 1.121 times higher than the corresponding to pH<sub>2</sub>O = 0.0094 atm in configuration B and C, respectively. On the other hand, the variation in the term  $\ln \frac{p(H_2, feed)}{p(H_2, perm)}$  by changing the pH<sub>2</sub>O is almost negligible. From these results, it could be concluded that an

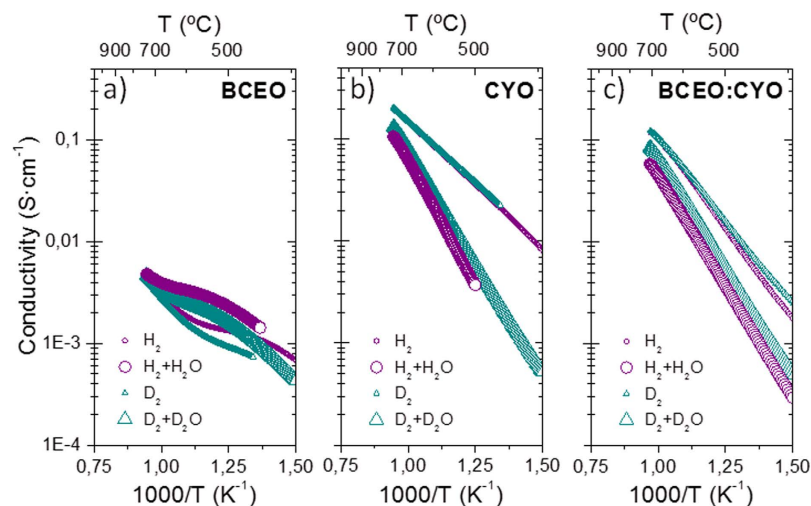


**Figure 3.** (a) Overview TEM image acquired from BCEO:CYO sample sintered at 1600 °C for 10 h; (b)-(c) shows the grain contrast dependence with distinct sample orientation (tilt); (d) Overview STEM-HAADF acquired from BCEO:CYO samples together with corresponding elemental EDX areal maps; (e) integrated EDX spectrum from the CYO and BCEO grains; and (f) High-resolution STEM-HAADF images acquired from CYO (Ce-Y-Eu-O) grain (left column) and BCEO (Ba-Ce-O) grain (right column) in the BCEO:CYO dual phase sample utilizing different crystallographic projections.

important contribution to the  $H_2$  flow observed comes from the water splitting and the changes observed when  $p_{H_2O}$  decreases are mainly ascribed to the reduction of the oxygen transport through the membrane.

Step-changes from configurations C to A were monitored for both conditions (Fig. 7b,c). Hydrogen fluxes decrease changing from configuration C to B and A for  $p_{H_2O} = 0.042$  atm (Fig. 7b) and they increase changing from configuration A to B and C for  $p_{H_2O} = 0.0094$  atm (Fig. 7c). The quick response for both conditions is mainly related to the important changes in  $H_2$  production by water splitting<sup>11</sup>.

In the light of the previous reports on similar cercer composites<sup>63,64,85</sup>, the effect of Pt layer applied on the surface of the tested membrane on water splitting reaction must be also taken into consideration. According to<sup>62</sup>, no contribution of water splitting to the  $H_2$  permeation could be evidenced at lower water vapor pressure at the



**Figure 4.** Total conductivity as a function of inverse temperature measured under four different reducing atmospheres: 5 vol.% H<sub>2</sub> in He and 5 vol.% D<sub>2</sub> in He both dry and moist (H<sub>2</sub> and D<sub>2</sub> humidified with H<sub>2</sub>O and D<sub>2</sub>O at room temperature, respectively) for (a) BCEO, (b) CYO and (c) BCEO:CYO composite.

sweep side due to omitting Pt catalyst layer. However, cerce composites were just recently identified as promising candidates for H<sub>2</sub> permeation membranes, therefore detailed studies have not been yet done to assign different reaction contributions to the overall H<sub>2</sub> flux.

**Stability tests.** The assessment of the H<sub>2</sub> permeation stability under CO<sub>2</sub> atmospheres was carried out at 700 °C during 6 days by using 15 vol.% CO<sub>2</sub> in Ar as sweep gas, 50 vol.% H<sub>2</sub> in He as feed gas and humidifying both sides of the membrane (configuration B). Figure 8a shows that the H<sub>2</sub> flow obtained under CO<sub>2</sub> containing atmosphere is significantly lower than that obtained by using pure Ar, around 0.15 mL·min<sup>-1</sup>·cm<sup>-2</sup>. This drop in the H<sub>2</sub> flux by using CO<sub>2</sub> could be ascribed to the CO<sub>2</sub>/H<sub>2</sub> and/or CO<sub>2</sub>/O<sub>2</sub> competitive adsorption on the membrane surface<sup>11,63</sup> that slows down the gas exchange. Besides, the flux equilibration in these conditions takes more than 2 days.

TG analysis was performed under 5 vol.% CO<sub>2</sub> balanced with Ar by using crushed samples sintered at 1600 °C. Figure 8b indicates that CYO sample does not suffer any CO<sub>2</sub> uptake under CO<sub>2</sub> containing atmospheres within the studied temperature range. On the other hand, a mass gain is observed for BCEO sample above 500 °C, which is ascribed to the carbonation of the sample<sup>37,66,73</sup>. However, the mass of the composite BCEO:CYO does not increase indicating that no carbonation process is taking place. It can be concluded that the addition of ceria to the cerate allows shifting the carbonation equilibrium and consequently improves the stability of the composite.

The structural stability of BCEO:CYO membrane after 13 days on stream (total duration of the permeation measurements, including CO<sub>2</sub> stability measurement) was verified by XRD (Fig. 9a) and FE-SEM (Fig. 9b–e) analysis. Figure 9a depicts the XRD patterns recorded on dual-phase BCEO:CYO sintered samples before permeation (labeled BP-1600/10 h) and after the H<sub>2</sub>-permeation tests (labeled AP). Both the feed and sweep sides of the membrane were investigated to detect any phase changes or formation of carbonates (denoted with AP-F and AP-S for feed and sweep side of the membrane, respectively). As reference, the peak positions and related intensities for BaCeO<sub>3</sub> and Ce<sub>0.8</sub>Y<sub>0.2</sub>O<sub>1.9</sub> are presented at the bottom of the figure. Up to the analytical limits of the XRD device, no changes in the structure were found and both fluorite and perovskite phases are well distinguished. The peaks marked by a diamond correspond to Pt traces from the catalytic layer. The XRD pattern of the sweep side of the membrane after permeation shows broader peaks, which could indicate some modification in the surface microstructure.

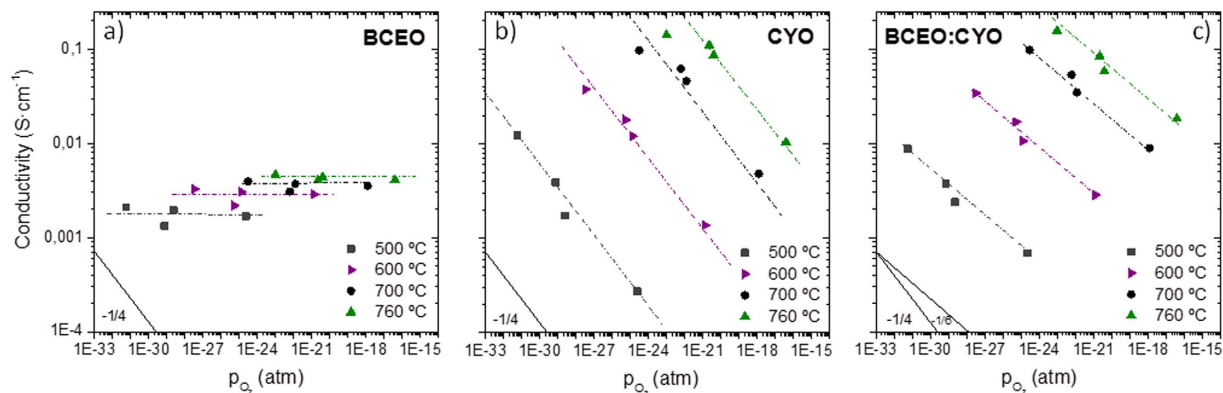
Figure 9b–e presents the FE-SEM (b), and the BSE-SEM (c,d) micrographs and EDS linescan (e) analysis of the membrane after the permeation measurements (note that the membrane is the same specimen for all the H<sub>2</sub> permeation measurements described in the work). Neither morphological nor structural changes were detected.

Finally, stability of dual phase composite material was checked after running the sintered membranes in continuous electrical tests for 670 h under 4% H<sub>2</sub>-containing dry reducing conditions, including several cycling steps from room temperature up to 900 °C and 1000 °C with holding times at these temperatures of 48 h and 20 h, respectively. TEM investigation of a reference sample and of post-treatment specimens is shown in Figure S4 in Supplementary Information. No morphological differences between the reference and post-treatment membranes and no grain boundary segregations were detected resulting from the 670 h continuous operation under reducing conditions.

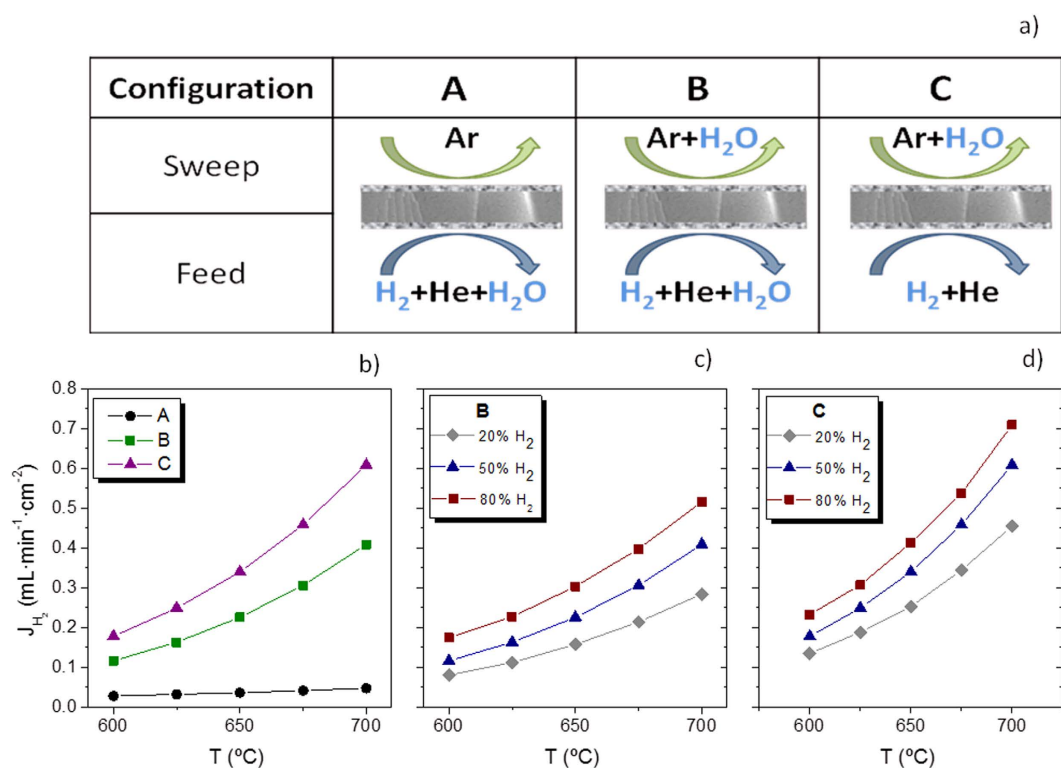
## Conclusions

We have presented a dual phase material based on perovskite and fluorite exhibiting both high electronic and H<sup>+</sup>/O<sup>2-</sup> co-ionic conductivity. Nominal BaCe<sub>0.8</sub>Eu<sub>0.2</sub>O<sub>3-δ</sub> (BCEO) and Ce<sub>0.8</sub>Y<sub>0.2</sub>O<sub>2-δ</sub> (CYO) were mixed in a 50:50 vol.% ratio. Eu<sup>3+</sup> cations originally present in BCEO diffused to CYO during the sintering step (Ce<sub>0.695</sub>Y<sub>0.2</sub>Eu<sub>0.105</sub>O<sub>2-δ</sub>;





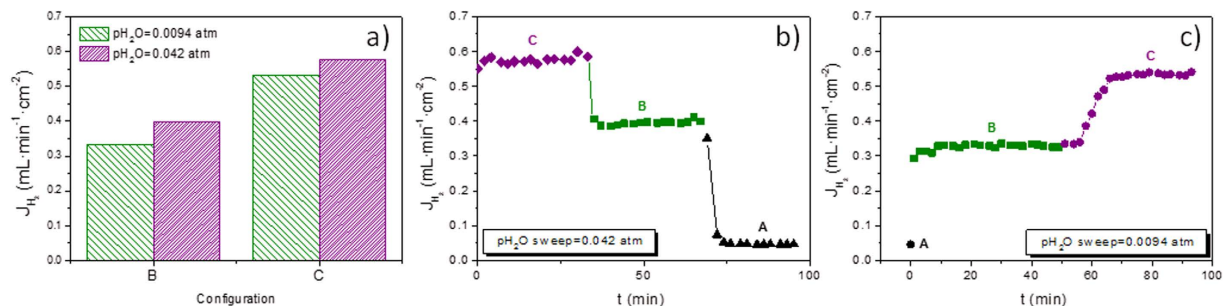
**Figure 5.** Total conductivity as a function of  $p_{O_2}$  under wet reducing atmospheres at four different temperatures for (a) BCEO, (b) CYO and (c) BCEO:CYO.



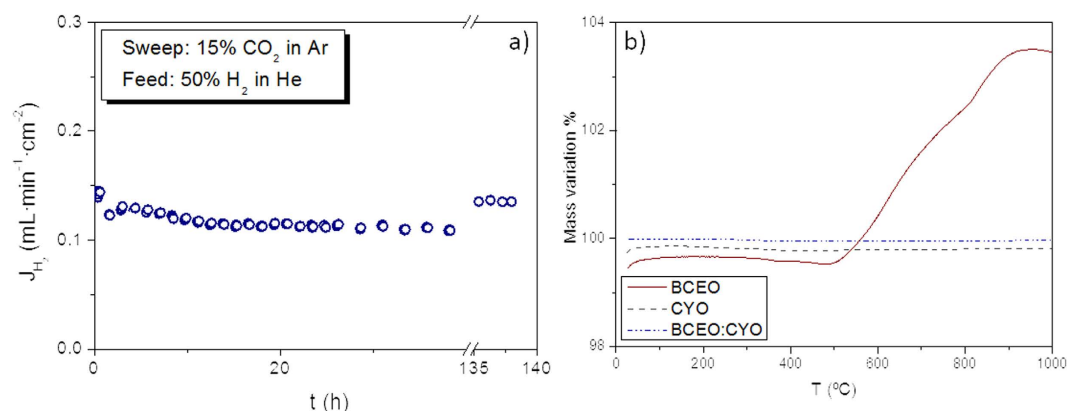
**Figure 6.** (a) Representation of configurations for membrane testing; (b)  $H_2$  flow as a function of temperature in the three configurations A, B and C with different hydration degree, feed stream 50%  $H_2$  in He; and (c,d)  $H_2$  flow as a function of temperature for different  $p_{H_2}$  feeding in configuration B and C.

CYEuO) as revealed by TEM analysis. DC-conductivity using a deuterium tracer showed that the composite is mixed protonic-electronic conductor (MPEC). Indeed, BCEO behaves as a prevailing proton conductor and CYO (CYEuO) as a mixed oxygen ionic and electronic conductor.

$H_2$  permeation through a thick planar BCEO:CYO membrane was measured in different gas hydration configuration. The reached  $H_2$  fluxes are among the highest reported for ceramic membranes up to date 0.4 and 0.61 mL·min<sup>-1</sup>·cm<sup>-2</sup> at 700 °C under configuration B and C, respectively (or thickness-normalized values of 0.02 and 0.0305 mL·min<sup>-1</sup>·cm<sup>-1</sup> at 700 °C under configuration B and C, respectively). This remarkably high permeability originates from the transport of protons as well as from the hydrogen generated by water splitting on the permeate side (the high oxygen partial pressure side) of the membrane due to the oxygen ion transport in ceria. The membrane material was stable after 670 h continuous operation under 4%  $H_2$ -containing dry reducing conditions, as well as in 15%  $CO_2$  containing atmosphere, although the resulting flux was lower than in pure Ar due to the competitive adsorption in the surface between  $H_2/O_2$  and  $CO_2$ . The stability in  $CO_2$  is tentatively attributed to the protective effect of the ceria phase over the cerate phase.



**Figure 7.** (a) H<sub>2</sub> flows obtained at 700 °C in configurations B and C for  $p_{H_2O}(\text{sweep}) = 0.042$  atm and  $p_{H_2O}(\text{sweep}) = 0.0094$  atm. (b,c) H<sub>2</sub> flow variation produced by the step-change from C to B to A with (b)  $p_{H_2O}(\text{sweep}) = 0.042$  atm and (c)  $p_{H_2O}(\text{sweep}) = 0.0094$  atm at 700 °C feeding 50 vol.% H<sub>2</sub>.



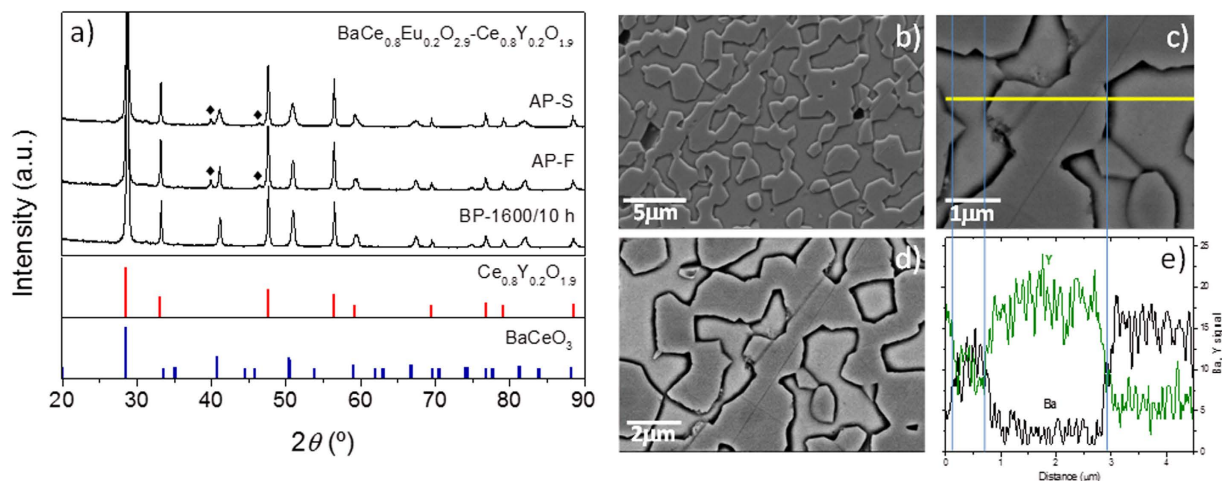
**Figure 8.** (a) H<sub>2</sub> flow as a function of time by using 15 vol.% CO<sub>2</sub> in Ar as sweep gas and 50 vol.% H<sub>2</sub> in He as feed gas at 700 °C. Both sides of the membrane were humidified (configuration B). (b) TG measurements of BCEO, CYO and BCEO:CYO under 5 vol.% CO<sub>2</sub> in Ar.

## Experimental Section

**Sample preparation.** Materials with nominal composition BaCe<sub>0.8</sub>Eu<sub>0.2</sub>O<sub>3- $\delta$</sub>  (BCEO) and Ce<sub>0.8</sub>Y<sub>0.2</sub>O<sub>2- $\delta$</sub>  (CYO) were synthesized via the conventional solid-state route. Stoichiometric amounts of high-purity BaCO<sub>3</sub> and CeO<sub>2</sub>, Eu<sub>2</sub>O<sub>3</sub>, Y<sub>2</sub>O<sub>3</sub> (Sigma Aldrich) were weighted, mixed and ball-milled in ethanol. The solid-state synthesis for both compounds took place at 1400 °C for 6 h in air. A second calcination step (20 h) was required for the complete formation of the perovskite single phase, and resulting products were milled in ethanol. After drying, the two powder products were mixed in 50:50 vol.% ratio and were homogenized 24 h in ethanol, dried and sieved. Uniaxially pressed samples were sintered at 1600 °C for 10 h leading to formation of dual phase BCEO:CYO material with nominal stoichiometry BaCe<sub>0.8</sub>Eu<sub>0.2</sub>O<sub>3- $\delta$</sub> :Ce<sub>0.8</sub>Y<sub>0.2</sub>O<sub>2- $\delta$</sub> . Samples were grinded to remove the defective top layer and no Ba-deficiency or secondary oxide exclusions compensating eventual Ba evaporation were detected by ICP-OES. Relative density of more than 98% was determined via the Archimedes approach on sintered dual phase composite samples.

**Characterization techniques.** Inductively coupled plasma optical emission spectrometry (ICP-OES) analysis was used to monitor chemical composition of the dual phase ceramics by quantification of the cation content. 50 mg sample powder was mixed with 2 mL HClO<sub>4</sub> (NORMATOM®, VWR Chemicals, Germany) and heated for 30 min to fuming. After cooling, 1 mL H<sub>2</sub>O<sub>2</sub> (NORMAPUR) and 1 mL HCl (Merck Suprapur, Germany) were added. Samples were heated at 70 °C for 30 min until complete dissolution of solid content and finally 50 mL volumes were prepared. Samples were measured at dilution 1:20 with a Thermo Scientific iCAP 7600 dual-view spectrometer, each sample was measured twice and the mean result of three emission lines per element was used for quantification. External calibration was performed with standards prepared by dilution of Merck Certipur® certified plasma emission standards with diluted acids. Relative standard deviation is 1–3%.

Powder X-ray diffraction (XRD) patterns of BCEO, CYO and composite BCEO:CYO as prepared and sintered were recorded in the 2 theta range from 10° to 80° using D4 ENDEAVOR diffractometer by Bruker AXS with CuK $\alpha$  radiation ( $\lambda = 1.54$  Å). XRD patterns of BCEO:CYO membrane after permeation tests were recorded in the 2 theta range from 20° to 90° using PANalytical Cubix fast diffractometer with CuK $\alpha_{1,2}$  radiation and an X'Celerator detector in Bragg-Brentano geometry. Phase identification was carried out with ICDD PDF2-Database (Release 2004) and X'Pert Highscore Plus (by PANalytical). The TOPAS V4 software (Bruker AXS) was used to determine the lattice parameters and for Rietveld refinements. The stoichiometry of each phase



**Figure 9.** (a) Diffraction patterns for BCEO:CYO dual phase ceramic samples: as-sintered sample (BP-1600/10h) and sample after H<sub>2</sub>-permeation tests (feed side AP-F and sweep sides AP-S). Peak positions corresponding to Pt traces from the catalytic layer are denoted by symbol (♦). As a reference (bottom): peak positions and their intensities for BaCeO<sub>3</sub> and Ce<sub>0.8</sub>Y<sub>0.2</sub>O<sub>1.9</sub>. (b) SEM and (c,d) BSE-SEM micrographs and (e) EDS linescan analysis of the fractured cross-section of the BCEO:CYO membrane after the permeation measurements.

is fixed for the Rietveld refinement and the starting crystal structure model was obtained from the Inorganic Crystal Structure Database (ICSD). The final weighted R-factor (R<sub>w</sub>, %) for each refinement is listed in Table 2.

Microstructural and chemical analyses of BCEO and CYO, as well as of the dual phase samples were performed by means of field emission scanning electron microscopy (FE-SEM) (Zeiss Ultra 55) equipped with energy-dispersive X-ray spectroscopy (EDS) (INCA, Oxford). Transmission electron microscopy (TEM) analysis was carried out using a Tecnai TF 20 UT equipped with EDS and a Gatan's imaging filter (GIF) operated at 200 kV. High-resolution scanning TEM high angle annular dark field (STEM-HAADF) imaging was performed by using C<sub>s</sub> probe corrected FEI-Titan 80–300 STEM instrument operating at 300 kV. Electron-transparent specimens were prepared from as-sintered pellets using the traditional approach which includes ultrasonic drilling, mechanical polishing and Ar<sup>+</sup> ion milling at 5 keV.

Thermal expansion coefficient (TEC) of the dual phase ceramics was measured on sintered bar samples with dimensions 25 × 4 × 4 mm<sup>3</sup> from 30 to 1400 °C (heating rate 3 °C/min) in air using Netzsch Dil 402C. Linear thermal expansion coefficients  $\alpha$  under heating and cooling were calculated using Eq. (2), where L<sub>0</sub> and dL are respectively the initial length and the length change of the sample, while T<sub>0</sub> and T are the starting and the final temperature, respectively:

$$\alpha = \frac{dL}{L_0(T - T_0)} \quad (2)$$

Electrical conductivity measurements were carried out by the standard four-point DC technique on sintered bars. Rectangular bar specimens were prepared from the resulting powders through uniaxial pressing at 100 MPa and subsequently sintered at 1600 °C. Silver paste and wire were used for contacting. The constant current was supplied by a programmable current source (Keithley 2601), while the voltage drop through the sample was detected by a multimeter (Keithley 3706). The voltage was measured by the current in both forward and reverse directions, in order to eliminate the thermal effect and to avoid non-ohmic responses. H/D isotopic effect and hydration effect for the separate materials and the dual phase material BCEO:CYO, were studied under reducing conditions: 5 vol.% H<sub>2</sub> in He and 5 vol.% D<sub>2</sub> in He both dry and moist (H<sub>2</sub> and D<sub>2</sub> humidified with H<sub>2</sub>O and D<sub>2</sub>O at room temperature, respectively). The influence of the pO<sub>2</sub> on total conductivity was also evaluated under wet reducing atmospheres by using: wet 0.05 vol.%, 5 vol.% and 100 vol.% H<sub>2</sub> and wet 5 vol.% D<sub>2</sub>.

Temperature-programmed reduction (TPR) was performed using 2910 Micromeritics equipment on crushed and ground samples sintered at 1600 °C. Powders were degassed under a dry Ar flow for 1 h and then were subjected to reduction under a dry H<sub>2</sub>/Ar (1/9) flow, and a heating rate of 10 °C·min<sup>-1</sup> until 900 °C. H<sub>2</sub> consumption was measured by a TCD (thermal conductivity detector).

H<sub>2</sub> permeation measurements were performed in a multistep continuous process using a BCEO:CYO dense disk with a diameter of 15 mm and thickness of 500 μm sintered at 1600 °C. Both disk surfaces were coated with a 20 μm screen printed Pt porous layer aiming to improve the surface catalytic activity and evaluate predominantly the bulk H<sub>2</sub> transport. Permeation measurements were performed on a sealed sample in double chamber quartz reactor following the methodology reported elsewhere<sup>25,86</sup>. Three different hydration conditions were evaluated: (A) feed side humidified and sweep side dry; (B) both sides humidified; (C) only sweep side humidified, where humidification means pH<sub>2</sub>O = 0.042 atm and dry corresponds to pH<sub>2</sub>O = 2 · 10<sup>-5</sup> atm (bottle dry). Ar was used as sweep gas and a 50 vol.% H<sub>2</sub> in He mixture was employed as feed (1.18 atm absolute pressure) under all the mentioned conditions. The flow rates used were 100 mL·min<sup>-1</sup> for feed and 150 mL·min<sup>-1</sup> for sweep and they

were controlled using mass flow controllers (MFCs). The hydrogen content in the permeate side was analyzed using micro-GC Varian CP-4900 equipped with Molsieve5A and PoraPlot-Q glass capillary modules. Sealing was obtained using a silver ring and applying a spring load. Sealing was confirmed by continuous monitoring the He concentration in the permeate stream and it was considered adequate when the helium concentration was lower than 5% of the H<sub>2</sub> permeated.

In order to evaluate the stability in operation of the composite under CO<sub>2</sub> containing atmospheres, H<sub>2</sub> permeation measurements were also performed at 700 °C for 6 days by using 15 vol.% CO<sub>2</sub> in Ar as sweep gas. After permeation measurements, integrity of the sample was evaluated by means of XRD and SEM analysis. Scheme of the steps of the hydrogen permeation measurements in a continuous process is given in Supplementary information, Figure S1.

Additionally, TG measurements were carried out on the BCEO, CYO and the composite BCEO:CYO by using crushed samples sintered at 1600 °C. TG was performed in a flow of dry 5% CO<sub>2</sub> in Ar from 25 to 1000 °C with a heating ramp of 10 K·min<sup>-1</sup> by using a Mettler-Toledo StarE balance.

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### Author Contributions

M.E.I., W.A.M., O.G., S.E. and J.M.S. designed the project and lead the interpretation and writing of the manuscript. M.E.I. and S.E. performed the electrochemical measurements. J.P. and J.M. carried out the microstructural analysis, Y.J.S. and M.B. contributed to the structural characterization. All authors contributed to the concept and analysis of results, and revised the manuscript.

### Additional Information

**Supplementary information** accompanies this paper at <http://www.nature.com/srep>

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