

Surface Plasmon Resonance Sensing Detection of Mercury and Lead Ions Based on Conducting Polymer Composite

Mahnaz M. Abdi^{1*}, Luqman Chuah Abdullah¹, Amir R. Sadrolhosseini², Wan Mahmood Mat Yunus², Mohd Maarof Moxsin², Paridah Md. Tahir¹

¹ Laboratory of Biopolymer and Derivatives, Institute of Tropical Forestry and Forest Products (INTROP), Universiti Putra Malaysia, Serdang, Selangor, Malaysia,

² Department of Physics, Faculty of Science, Universiti Putra Malaysia, Serdang, Selangor, Malaysia

Abstract

A new sensing area for a sensor based on surface plasmon resonance (SPR) was fabricated to detect trace amounts of mercury and lead ions. The gold surface used for SPR measurements were modified with polypyrrole-chitosan (PPy-CHI) conducting polymer composite. The polymer layer was deposited on the gold surface by electrodeposition. This optical sensor was used for monitoring toxic metal ions with and without sensitivity enhancement by chitosan in water samples. The higher amounts of resonance angle unit (Δ RU) were obtained for PPy-CHI film due to a specific binding of chitosan with Pb^{2+} and Hg^{2+} ions. The Pb^{2+} ion bind to the polymer films most strongly, and the sensor was more sensitive to Pb^{2+} compared to Hg^{2+} . The concentrations of ions in the parts per million range produced the changes in the SPR angle minimum in the region of 0.03 to 0.07. Data analysis was done by Matlab software using Fresnel formula for multilayer system.

Citation: Abdi MM, Abdullah LC, Sadrolhosseini AR, Mat Yunus WM, Moxsin MM, et al. (2011) Surface Plasmon Resonance Sensing Detection of Mercury and Lead Ions Based on Conducting Polymer Composite. PLoS ONE 6(9): e24578. doi:10.1371/journal.pone.0024578

Editor: Matteo Rini, Joint Research Centre - European Commission, Germany

Received: May 5, 2011; **Accepted:** August 14, 2011; **Published:** September 9, 2011

Copyright: © 2011 Abdi et al. This is an open-access article distributed under the terms of the Creative Commons Attribution License, which permits unrestricted use, distribution, and reproduction in any medium, provided the original author and source are credited.

Funding: This work was supported by Ministry of Higher Education of Malaysia (GRF) and University Putra Malaysia. The funders had no role in study design, data collection and analysis, decision to publish, or preparation of the manuscript.

Competing Interests: The authors have declared that no competing interests exist.

* E-mail: mahnaz@introp.upm.edu.my

Introduction

In recent years, surface plasmon resonance (SPR) has been widely demonstrated to be an effective optical technique for in situ investigation of the optical and electrical properties of conducting polymers, such as polypyrrole films [1]. Surface plasmon resonance (SPR) is a quantum optical-electrical phenomenon arising from the interaction of light with a metal surface [2]. The charge density oscillations that exist at a metal-dielectric (or metal/vacuum) interface are called surface plasmons that propagate in a direction parallel to the metal-dielectric interface and are on the boundary of the metal and the external medium (air or water for example) [3]. So that, these oscillations are very sensitive to any change of this boundary, such as adsorption of molecules to the metal surface.

Heavy metals such as mercury and lead ions have long been recognized as a harmful environmental pollutant. Lower levels of lead can affect on the central nervous system, kidney, and blood cells [4] and in severe cases can cause convulsions, coma, and even death. Lead interferes with the development of the nervous system and is therefore particularly toxic to children and unborn babies, causing potentially permanent learning and behaviour disorders [5].

There are many techniques for analysis trace metal e.g., atomic absorption, atomic emission, and fluorescence spectrometry, inductively coupled plasma-mass spectrometry (ICP-MS), and electrochemical techniques (such as ion-selective potentiometry and anodic stripping voltammetry) [6]. In contrast to their attractive analytical “figures of merit”, each of these techniques suffers from some certain disadvantages. Applications of all of

these methods require a knowledge of chemistry and instrumentation and need exactitude apparatuses [7].

A major disadvantage of ICP-MS is the high capital cost of the instrumentation. In addition, this instrument is bulky and not selective to different charge states of an element. The non linearity of the calibration curves is a disadvantage of the atomic absorption spectroscopy (AAS) technique (in the absorbance range higher than 0.5 to 1). Running and investment costs of inductively coupled plasma atomic emission spectroscopy (ICP-AES) are high. Another disadvantage of ICP-AES is spectral interferences (due to many emission lines). Voltammetric methods are simple, inexpensive, and portable but they suffer of interferences inherent in complex sample matrix. Anodic stripping voltammetry can only measure amalgam-forming metal species [8]. Thereby, the complementary methods have developed to overcome some of these shortcomings.

SPR is a method that provides an effective optical technique to detect monolayer thicknesses of the material on the conducting surface. This technique exhibits a good sensitivity, stability, and reproducibility; sometimes even changes in refractive index of approximately 10^{-5} can be detected by SPR technique [9]. Conducting sensors based on PPy film with different dopant were developed for detecting volatile aromatic hydrocarbons [10] and volatile organic solvent sensors have also been fabricated using PPy films on conducting glass substrates [11]. Selective SPR sensors based on PPy have been reported for the detection of Cu^{2+} [12], and also for biosensor such as *Pseudomonas aeruginosa* cells [13].

In previous work [14], we obtained the refractive indexes of PPy and PPy-CHI thin films (in contact with air) by SPR technique.

Since they produced a sharp peak of resonance angle, it was shown these polymers are capable for using in sensitive optical sensors. The aim of the current study is to prepare optical sensors based on PPy and PPy-CHI composite films to detect trace amount of Hg^{2+} and Pb^{2+} . The optical sensors were used for monitoring of toxic metal ions with and without sensitivity enhancement by chitosan. To our knowledge, this is the first report of a SPR sensor that uses PPy-CHI biocomposite for heavy metal detection.

Materials and Methods

Pyrrrol monomer (Acros Organic) was distilled and stored below 4°C . All the other reagents including chitosan (a local company with 88% degree of deacetylation) and *p*-toluene sulfonate (Fluka) were of analytical grade, and used without further purification. All of the electropolymerization of pyrrole was done via a typical three-electrode electrochemical cell arrangement. A saturated calomel electrode (SCE) and a carbon rod were used as reference and counter electrode, respectively. The thin layer of gold covered onto the glass slide was used as working electrode.

The glass microscope slides (ZF52, $n = 1.8395$) were cleaned using $\text{HCl}/\text{H}_2\text{O}$ at 60°C for 15 min following by rinsing the surface with deionized water and drying at 50°C . Each glass slide was coated with a thin layer of gold, using a sputtering coating (Model SC7940, Polaron). The thin layer of gold coated onto glass slide was used as working electrode in electrochemical deposition of polymer.

The sample of PPy-CHI was prepared in a solution containing 0.015 M Pyrrole, 0.005 M *P*-TS and 0.07% (w/v) chitosan in a 1% acetic acid solution. The PPy film was prepared, using the same composition of pyrrole and *P*-TS without chitosan. Electrochemical deposition of PPy-CHI and PPy films on the gold substrate was performed using a potentiostat (Model: PS 605, USA), imposing a constant potential of 0.85 V (vs SCE) for 150 sec. The results from FT-IR, electrical conductivity, and XRD study confirmed the incorporation of chitosan in PPy structure [15]. The electrochemical technique which can be used to electropolymerization can be chosen from all known techniques such as; potentiostatic, galvanostatic and potential/ current sweep, etc.

The surface of the polymer films was washed with distilled deionized water (DDW). The thin slides coated with polymer were attached to the prism with high index glass (ZF52, $n = 1.8395$) by using index-matching fluid ($n = 1.52$ new port company). SPR measurements were conducted with a set up assembled according to the Kretschmann's prism configuration (Figure 1). The precision of rotator stage was 0.003. Data analysis was done by Matlab software using Fresnel formula for multilayer system [16]. In the Fresnel equation for *p*-polarized light the reflectivity of the light ($R(\theta)$) is a function of angle and wavelength where θ is the angle of incidence light at the interface between the prism and surface [17]. DDW was then run on the surface of the films to establish a baseline. When a stable baseline of the SPR angle was obtained, the PPy and PPy-CHI film exhibited a SPR angular profile deep enough for precise definition of the resonance angle. The refractive indexes of the polymer films were successfully measured by SPR technique.

Results and Discussion

Figure 2 shows the surface plasmon resonance (SPR) reflectivity curves for the glass/Au/DDW, (Au), and glass/Au/PPy/DDW, (Au-PPy), and glass/Au/PPy-CHI/DDW, (Au-PPy-CHI), systems. The values of the real and imaginary parts of the refractive indexes of films were successfully obtained by nonlinear least square fitting using Fresnel equations for multilayer system. The resonance angle (θ_{spr}) and the optical constant for the PPy film were

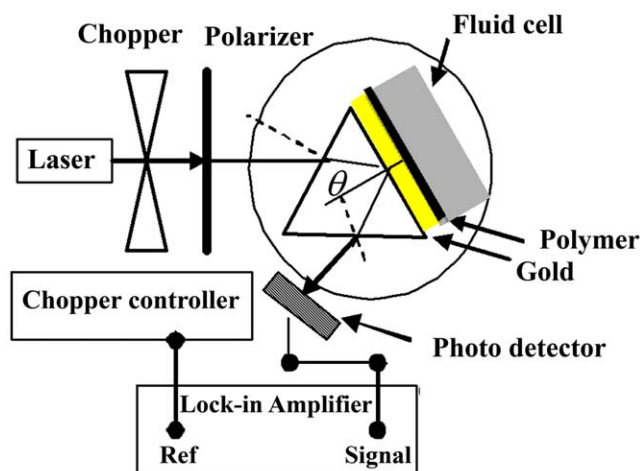


Figure 1. Set up configuration for SPR measurements.
doi:10.1371/journal.pone.0024578.g001

$(57.67 \pm 0.01)^{\circ}$ and $1.598 + 0.159i$, respectively. The formation of a thin film of polypyrrole with chitosan on the gold surface was monitored as a decreasing SPR angle for the PPy-CHI film.

The resonance angle (θ_{spr}) and the optical constant of the PPy-CHI film were $(56.75 \pm 0.01)^{\circ}$ and $1.601 + 0.165i$, respectively. The dip widths of the curves for the PPy-CHI and PPy films were larger than Au/DDW, which was proportional to the attenuation of the surface plasmon. The thickness of the thin films was calculated using the Fresnel equation and the values of 26 nm and 28 nm were obtained for the PPy-CHI and PPy thin films, respectively. When a stable baseline of SPR angle was obtained for PPy and PPy-CHI films, the surface of the films were ready for binding tests for Hg^{2+} and Pb^{2+} samples.

Binding of Hg^{2+} and Pb^{2+} on to PPy and PPy-CHI films

Different concentrations of Hg^{2+} and Pb^{2+} in aqueous solution were injected in to the fluid cell attached to the films, and the SPR angle (resonance units = RU) was monitored to detect any binding interactions. The increase in the SPR angle (ΔRU) for each Hg^{2+} and Pb^{2+} concentration was determined based on the resonance angle of the ions (after the ion binding event) and the DDW baselines before running ions samples over a time duration of Δt .

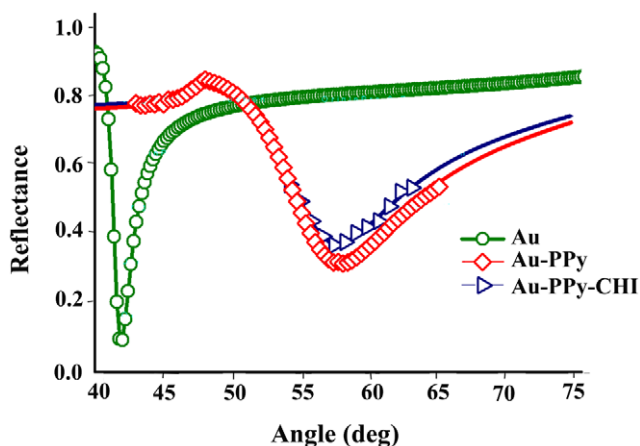


Figure 2. SPR angular profiles for DDW on the Au, Au-PPy and Au-PPy-CHI surface.
doi:10.1371/journal.pone.0024578.g002

These ΔRU results indicated that Hg^{2+} and Pb^{2+} were binding successfully with the PPy and PPy-CHI. Each sample was run for 20 min. The concentrations of ions in the parts per million range produced the changes in the SPR angle minimum in the region of 0.03 to 0.07. Due to the high sensitivity of SPR, it is also possible to detect concentrations in the parts per billion ranges. A control experiment for the Hg^{2+} and Pb^{2+} ions with bare Au surface was done for 30 min, which produced no changes in the resonance angle indicating that the polymer was necessary to react with the ions. After exposure to each sample concentration, all of the PPy and PPy-CHI films were washed with DDW to establish a baseline, as the history of the films was irrelevant.

Sensitive optical sensor for Hg^{2+} and Pb^{2+} detection

The increase in the SPR angle for each Hg^{2+} and Pb^{2+} concentration indicated that these ions were binding successfully with PPy. The typical sensograms for 0.5–12 ppm Hg^{2+} and Pb^{2+} binding on PPy are presented in Figure 3. It can be observed that, first, ΔRU increased over the time for each sample, but after almost 700 sec the graphs showed a constant value. This result can be explained by saturation of the binding sites available on the PPy surface.

It seems that there is a chemical binding of ions in water with the PPy thin film that was immobilized on the gold surface. The interaction between Hg^{2+} and PPy was also reported by Yu *et al.* [9], as the natural preference for a combination of soft acid (Hg^{2+}) and soft base (PPy).

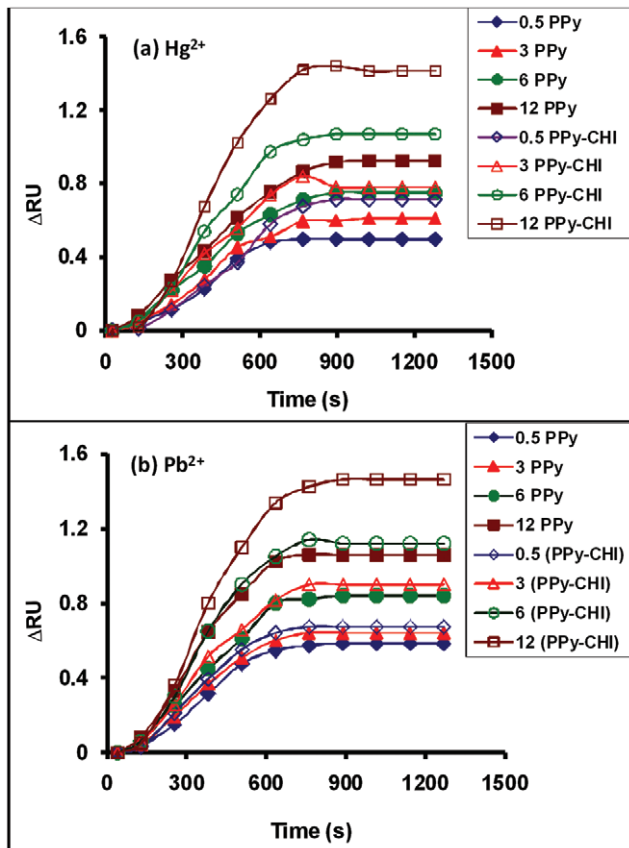


Figure 3. The sensograms for 0.5–12 ppm ion binding on PPy and PPy-CHI surface. Filled symbols (PPy surface), Open symbols (PPy-CHI surface). SPR sensogram for (a) Hg^{2+} and (b) Pb^{2+} . doi:10.1371/journal.pone.0024578.g003

On the other hand, there are some studies on the treatment of solutions of metal ions with redox polymers, such as polypyrroles, polyanilines and polythiophenes, to reduce the ions to a lower valence [18]. These literatures show that PPy can interact with certain heavy metals through an acid and base interaction or an oxidation reaction.

Sensitivity Enhancement Using CHI

Chitosan was used to prepare the PPy-CHI composite film to function as a sensitivity enhancement agent for detecting for heavy metals ions. DDW was run to establish a baseline. The decreasing SPR angle for the PPy-CHI film when compared to the PPy film indicated that chitosan incorporated inside the PPy structure. The sensograms for 0.5–12 ppm Hg^{2+} and Pb^{2+} binding on the PPy-CHI are also presented in Figure 3. The composite films of PPy-CHI showed the higher ΔRU compared to the PPy film for each time and concentration, indicating a higher sensitivity of the PPy in the presence of chitosan. In fact, the natural biopolymers have adsorption properties toward certain heavy metals. Of the many absorbents identified, chitosan has the highest sorption capacity for several metal ions [19]. Following this interaction of CHI and metal ions to enhance the original ΔRU , lower concentrations of ions could be readily detected with less interference.

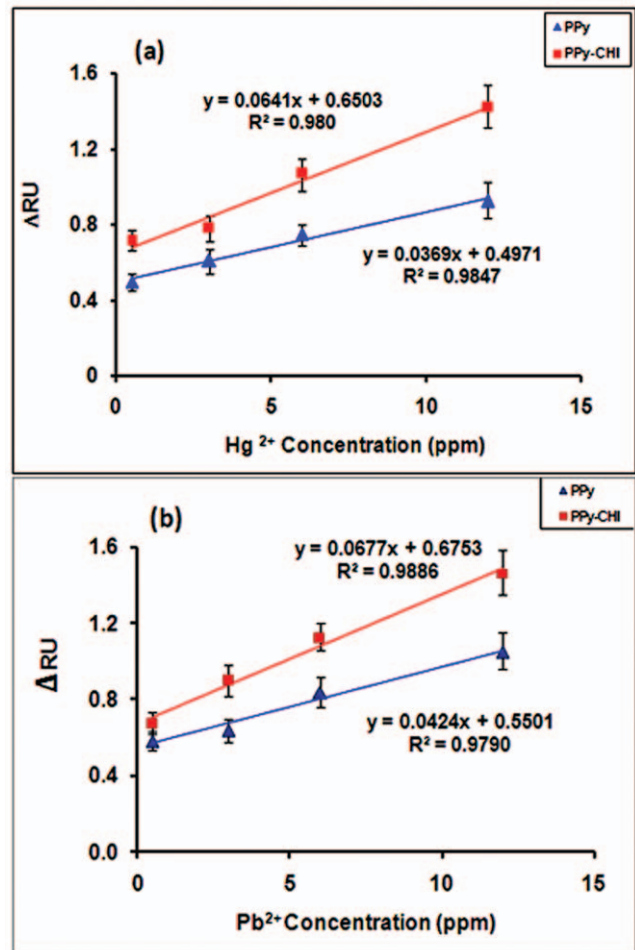


Figure 4. Standard calibration curves of ΔRU versus ion concentration for PPy and PPy-CHI films. (a) Hg^{2+} and (b) Pb^{2+} . doi:10.1371/journal.pone.0024578.g004

In order to show the sensitivity of the PPy and PPy-CHI films for metal ions, standard calibration curves were plotted for sensors and are shown in Figure 4. Different concentrations of ions were run for 20 min and all ΔRU were taken from the first 13 min of each binding curves. Each concentration was repeated 3 times and the standard deviation data were in the region of 0.06 to 0.18 and it is shown as error bar on the diagram of sensitivity.

These figures clearly show the enhancement of sensitivity for sample determination in the presence of chitosan. In the both sensograms the SPR dip shift is directly correlated with the ions concentration ($[Hg^{2+}]$ or $[Pb^{2+}]$). Linear regression analysis of the sensors yielded the following equations with all correlation coefficients R^2 greater than 0.97, revealing good linearity for the relationship between the SPR dip shift and ions concentration:

$$\Delta\theta = 6.50 \times 10^{-1} + 6.41 \times 10^{-2} [Hg^{2+}] \quad (\text{PPy-CHI})$$

$$\Delta\theta = 4.97 \times 10^{-1} + 3.69 \times 10^{-2} [Hg^{2+}] \quad (\text{PPy})$$

$$\Delta\theta = 6.75 \times 10^{-1} + 6.77 \times 10^{-2} [Pb^{2+}] \quad (\text{PPy-CHI})$$

$$\Delta\theta = 5.50 \times 10^{-1} + 4.24 \times 10^{-2} [Pb^{2+}] \quad (\text{PPy})$$

The slop of the standard calibration curve in the PPy-CHI sensor was almost 2 times higher than that for the PPy sensor

References

- Damos FS, Luz RCS, Kubota LT (2006) Investigations of ultrathin polypyrrole films: Formation and effects of doping/dedoping processes on its optical properties by electrochemical surface plasmon resonance (ESPR). *Electrochim Acta* 51: 304–1312.
- Garland JE, Assingbon KA, Pettit CM, Roy D (2003) Surface plasmon resonance transients at an electrochemical Interface: Time resolved measurements using a bicell photodiode. *Anal Chim Acta* 475: 47–58.
- Peyghambarian N, Koch SW, Mysyrowicz, A (1993) Introduction to semiconductor optics. Cliffs E, ed. Prentice Hall: Newjersey.
- Sathawara NG, Parikh DJ, Agarwal YK (2004) Essential Heavy Metals in Environmental Samples from Western India. *Bulletin of Environmental Contamination and Toxicology* 73: 756–761.
- Pearce JM (2007) Burton's line in lead poisoning. *European neurology* 57(2): 118–9.
- Skoog DA, Holler FJ, Nieman TA (1992) Principles of Instrumental Analysis. 5th ed. Philadelphia: Saunders College Publishing. pp 206–229.
- Tang Y, Zhai YF, Xiang JJ, Wang H, Liu B, et al. (2010) Colloidal gold probe-based immunochromatographic assay for the rapid detection of lead ions in water samples. *Environmental Pollution* 158: 2074–2077.
- Zhang Y, Xu M, Wang, Toledo F, Zhoua F (2007) Studies of metal ions binding by apo-metallothioneins attached onto preformed self-assembled monolayers using a highly sensitive surface plasmon resonance spectrometer Sensors and Actuators B. 123: 784–792.
- Yu JCC, Lai EPC, Sadeghi S (2004) Surface plasmon resonance sensor for Hg(II) detection by binding interactions with polypyrrole and mercaptobenzothiazole. *Sens. Actuators B* 101: 236–241.
- Barisci JN, Wallace GG, Andrews MK, Partridge AC, Harris PD (2002) Conducting polymer sensors for monitoring aromatic hydrocarbons using an electronic nose. *Sens. Actuators B: Chem.* 84: 252–257.
- Souza JEG, Santos FL, Neto BB, Santos CG, Santos MVB, et al. (2003) Free-grown polypyrrole thin films as aroma sensors. *Sens. Actuators B: Chem.* 88: 246–259.
- Kim DY, Cho HN, Kim CY (2000) Blue light emitting polymers. *Prog Polym Sci.* 25: 1089–1139.
- Park JS, Lee CM, Lee KY (2007) A surface plasmon resonance biosensor for detecting *Pseudomonas aeruginosa* cells with self-assembled chitosan-alginate multilayers. *Talanta* 72: 859–862.
- Abdi MM, Kassim A, Mahmud HNME, Mat Yunus WM, Talib ZB, et al. (2009) Physical, optical, and electrical properties of a new conducting polymer. *J Mater Sci* 44: 3682–3686.
- Abdi MM, Mahmud HNME, Kassim A, Mat Yunus WM, Talib ZBMohd, et al. (2010) Synthesis and Characterization of a New Conducting Polymer Composite. *Polymer Science Ser B* 52: Nos. 11–12, 662–669.
- Schasfoort RBM, Tudos AJ (2008) Handbook of surface Plasmon resonance, RSC publishing UK: Cambridge. pp 15–25.
- Pedrorotti LF, Pedrotti LS, Pedrotti LM (2007) Introduction to optics. 3rd ed. New York: Addison Wesley. pp 476–482.
- Sanjay B, Krishna R, Chang W (1994) Redox polymer films for metal recovery applications US Patent 5368632.
- Nomanbhay SM, Palanisamy K (2005) Removal of heavy metal from industrial wastewater using chitosan coated oil palm shell charcoal. *Electron j biotechn* 8: 44–53.
- Zhang X, Bai R (2002) Adsorption behavior of humic acid onto polypyrrole-coated nylon 6,6 granules. *J Mater Chem* 12: 2733–2739.

without chitosan. It can be seen that all the standard calibration curves showed a significant deviation from 0. This deviation was due to saturation of the binding sites available on the polymer surface, a fundamental limitation of maximum uptake capacity [20].

Conclusion

The binding interactions of heavy metals with PPy and CHI on the gold surface were monitoring using surface plasmon resonance technique. This technique was used to fabricate the PPy and PPy-CHI optical sensors for monitoring trace amount of Hg^{2+} and Pb^{2+} . The higher ΔRU results are due to specific binding of chitosan with Hg^{2+} and Pb^{2+} ions. Pb^{2+} ion binds to the polymer films most strongly and the sensors (based on PPy and PPy-CHI thin films) are more sensitive to Pb^{2+} compared to the Hg^{2+} . The refractive indexes of the polymer films were successfully measured by SPR technique.

Acknowledgments

The authors would like to thank the Department of Chemistry and Physics, University Putra Malaysia for their continued support throughout this project.

Author Contributions

Conceived and designed the experiments: MMA ARS WMMY. Performed the experiments: MMA ARS. Analyzed the data: MMA ARS WMMY. Contributed reagents/materials/analysis tools: WMMY MMHAM LCA PMT. Wrote the paper: MMA. Directed the research: MMA WMMY.