

Acta Crystallographica Section E **Structure Reports** Online

ISSN 1600-5368

1-[4-(4-Fluorophenyl)-6-methyl-2sulfanylidene-1,2,3,4-tetrahydropyrimidin-5-yl]ethanone

N. Anuradha,^a A. Thiruvalluvar,^a* S. Chitra,^b D. Devanathan^c and R. J. Butcher^d

^aPG Research Department of Physics, Rajah Serfoji Government College (Autonomous), Thanjavur 613 005, Tamilnadu, India, ^bDepartment of Chemistry, K.S.R. College of Engineering, K.S.R. Kalvi Nagar, Tiruchengode 637 215, Tamilnadu, India, ^cDepartment of Chemistry, Government Arts College, C. Mutlur 608 102, Chidambaram, Tamilnadu, India, and d Department of Chemistry, Howard University, 525 College Street NW, Washington, DC 20059, USA. Correspondence e-mail: thiruvalluvar.a@gmail.com

Received 24 July 2012; accepted 26 July 2012

Key indicators: single-crystal X-ray study; T = 123 K; mean σ (C–C) = 0.007 Å; R factor = 0.080; wR factor = 0.233; data-to-parameter ratio = 15.2.

In the title molecule, $C_{13}H_{13}FN_2OS$, the heterocyclic ring adopts a slightly distorted flattened boat conformation, and the plane through the four coplanar atoms makes a dihedral angle of 87.45 $(14)^{\circ}$ with the benzene ring. The thione, acetyl and methyl groups lie on the opposite side of the heterocyclic mean plane to the fluorophenyl group, which has an axial orientation. N-H···O, N-H···S, C-H···F and C-H···O intermolecular hydrogen bonds and a weak $C-H \cdots \pi$ interaction involving the benzene ring are found in the crystal structure.

Related literature

For chemical and biological applications of dihydropyrimidine derivatives and for the closely related crystal structure of the chloro derivative, see: Anuradha et al. (2009).



Experimental

Crystal data C13H13FN2OS

 $M_{\rm w} = 264.32$

Triclinic, $P\overline{1}$ a = 7.1775 (11) Åb = 8.1099 (13) Å c = 12.490 (2) Å $\alpha = 103.529 \ (15)^{\circ}$ $\beta = 91.933 (14)^{\circ}$ $\gamma = 106.971 (14)^{\circ}$

Data collection

Oxford Diffraction Xcalibur Ruby Gemini diffractometer Absorption correction: analytical (CrysAlis PRO; Agilent, 2012) $T_{\rm min}=0.515,\;T_{\rm max}=0.891$

Refinement

$R[F^2 > 2\sigma(F^2)] = 0.080$	H atoms treated by a mixture of
$wR(F^2) = 0.233$	independent and constrained
S = 1.12	refinement
2629 reflections	$\Delta \rho_{\rm max} = 0.82 \text{ e } \text{\AA}^{-3}$
173 parameters	$\Delta \rho_{\rm min} = -0.37 \ {\rm e} \ {\rm \AA}^{-3}$

Table 1

Cg2 is the centroid of the C41-C46 benzene ring.

Hydrogen-bond geometry (Å, °).

$D - H \cdots A$	D-H	$H \cdot \cdot \cdot A$	$D \cdots A$	$D - \mathbf{H} \cdots \mathbf{A}$
$N1-H1\cdotsO15^{i}$	0.82 (5)	2.10 (5)	2.867 (5)	156 (5)
N3−H3···S2 ⁱⁱ	0.85 (6)	2.50 (6)	3.350 (4)	174 (6)
C16−H16A····F4 ⁱⁱⁱ	0.98	2.50	3.358 (6)	146
$C61 - H61A \cdots F4^{iii}$	0.98	2.47	3.319 (6)	145
$C61 - H61B \cdots O15^{i}$	0.98	2.54	3.383 (5)	144
$C16-H16C\cdots Cg2^{iv}$	0.98	2.92	3.633 (5)	130

Symmetry codes: (i) x - 1, y, z; (ii) -x + 1, -y + 1, -z + 1; (iii) -x + 1, -y, -z; (iv) x, y - 1, z

Data collection: CrysAlis PRO (Agilent, 2012); cell refinement: CrysAlis PRO; data reduction: CrysAlis PRO; program(s) used to solve structure: DIRDIF2008 (Beurskens et al., 2008); program(s) used to refine structure: SHELXL97 (Sheldrick, 2008); molecular graphics: ORTEP-3 (Farrugia, 1997) and PLATON (Spek, 2009); software used to prepare material for publication: PLATON.

RJB acknowledges the NSF MRI program (grant No. CHE-0619278) for funds to purchase an X-ray diffractometer.

Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: WN2486).

References

- Agilent (2012). CrysAlis PRO. Agilent Technologies, Yarnton, Oxfordshire, England.
- Anuradha, N., Thiruvalluvar, A., Pandiarajan, K., Chitra, S. & Butcher, R. J. (2009). Acta Cryst. E65, 0564-0565.
- Beurskens, P. T., Beurskens, G., de Gelder, R., García-Granda, S., Gould, R. O. & Smits, J. M. M. (2008). The DIRDIF2008 Program System. Crystallography Laboratory, University of Nijmegen, The Netherlands.
- Farrugia, L. J. (1997). J. Appl. Cryst. 30, 565.
- Sheldrick, G. M. (2008). Acta Cryst. A64, 112-122.
- Spek, A. L. (2009). Acta Cryst. D65, 148-155.

V = 672.0 (2) Å³

Cu Ka radiation

 $0.59 \times 0.36 \times 0.06 \text{ mm}$

3949 measured reflections 2629 independent reflections

2109 reflections with $I > 2\sigma(I)$

 $\mu = 2.17 \text{ mm}^{-1}$

T = 123 K

 $R_{\rm int} = 0.055$

7 - 2

supplementary materials

Acta Cryst. (2012). E68, o2625 [doi:10.1107/S1600536812033727]

1-[4-(4-Fluorophenyl)-6-methyl-2-sulfanylidene-1,2,3,4-tetrahydropyrimidin-5yl]ethanone

N. Anuradha, A. Thiruvalluvar, S. Chitra, D. Devanathan and R. J. Butcher

Comment

As part of our investigations of dihydropyrimidine derivatives (Anuradha *et al.*, 2009) to compare their chemical and biological activities, we have undertaken the X-ray cryatal structure analysis of the title compound.

In the title molecule, $C_{13}H_{13}FN_2OS$ (Fig.1), the heterocyclic ring adopts a slightly distorted flattened boat conformation, and the plane through the four coplanar atoms (C2,N3,C5 and C6) makes a dihedral angle of 87.45 (14)° with the benzene ring. The thione, acetyl and methyl groups have equatorial orientations and the fluorophenyl group has an axial orientation. N1—H1…O15, N3—H3…S2, C16—H16A…F4, and C61—H61B…O15 intermolecular hydrogen bonds and a weak C16—H16C… π interaction involving the benzene (C41—C46) ring are found in the crystal structure (Fig.2, Table 1).

Experimental

A solution of acetylacetone (1.0012 g, 0.01 mol), 4-fluorobenzaldehyde (1.25 g, 0.01 mol) and thiourea (1.14 g, 0.015 mol) was heated under reflux in the presence of calcium fluoride (0.07 g, 0.001 mol) for 2 h (monitored by TLC). After completion of the reaction, the reaction mixture was cooled to room temperature and poured into crushed ice. The solid product was filtered under suction and purified by recrystallization from hot methanol to give the product in the pure form. Yield 1.92 g (96%).

Refinement

The two N-bound H atoms were located in a difference Fourier map and refined freely; N1—H1 = 0.82 (5) Å and N3— H3 = 0.85 (6) Å. The remaining H atoms were positioned geometrically and allowed to ride on their parent atoms, with Csp^2 —H = 0.95, *C*(methyl)—H = 0.98, and C(methine)—H = 1.00 Å; $U_{iso}(H) = kU_{eq}(C)$, where k = 1.5 for methyl and 1.2 for all other H atoms.

Computing details

Data collection: *CrysAlis PRO* (Agilent, 2012); cell refinement: *CrysAlis PRO* (Agilent, 2012); data reduction: *CrysAlis PRO* (Agilent, 2012); program(s) used to solve structure: *DIRDIF2008* (Beurskens *et al.*, 2008); program(s) used to refine structure: *SHELXL97* (Sheldrick, 2008); molecular graphics: *ORTEP-3* (Farrugia, 1997) and *PLATON* (Spek, 2009); software used to prepare material for publication: *PLATON* (Spek, 2009).



Figure 1

The molecular structure of the title compound, with displacement ellipsoids drawn at the 30% probability level. H atoms are shown as small spheres of arbitrary radius.



Figure 2

The packing of the title compound, viewed along the b axis. Dashed lines indicate hydrogen bonds. H atoms not involved in hydrogen bonding have been omitted.

1-[4-(4-Fluorophenyl)-6-methyl-2-sulfanylidene-1,2,3,4-tetrahydropyrimidin- 5-yl]ethanone

Z = 2
F(000) = 276
$D_{\rm x} = 1.306 {\rm ~Mg} {\rm ~m}^{-3}$
Melting point: 509 K
Cu <i>K</i> α radiation, $\lambda = 1.54184$ Å
Cell parameters from 1186 reflections
$\theta = 3.7 - 75.9^{\circ}$
$\mu = 2.17 \text{ mm}^{-1}$
T = 123 K
Plate, colourless
$0.59 \times 0.36 \times 0.06 \text{ mm}$

Data collection

Oxford Diffraction Xcalibur Ruby Gemini diffractometer	3949 measured reflections
Radiation source: Enhance (Cu) X-ray Source	2109 reflections with $L > 2\sigma(I)$
Craphite managhromator	P = 0.055
	$R_{\rm int} = 0.055$
Detector resolution: 10.5081 pixels mm ⁻¹	$\theta_{\text{max}} = /6.1^\circ, \ \theta_{\text{min}} = 3.7^\circ$
ω scans	$h = -7 \rightarrow 8$
Absorption correction: analytical	$k = -8 \rightarrow 10$
(CrysAlis PRO; Agilent, 2012)	$l = -15 \rightarrow 11$
$T_{\min} = 0.515, \ T_{\max} = 0.891$	
Refinement	
Refinement on F^2	Secondary atom site location: difference Fourier
Least-squares matrix: full	map
$R[F^2 > 2\sigma(F^2)] = 0.080$	Hydrogen site location: inferred from
$wR(F^2) = 0.233$	neighbouring sites
S = 1.12	H atoms treated by a mixture of independent
2629 reflections	and constrained refinement
173 parameters	$w = 1/[\sigma^2(F_o^2) + (0.1003P)^2 + 1.2879P]$
0 restraints	where $P = (F_o^2 + 2F_c^2)/3$
Primary atom site location: structure-invariant	$(\Delta/\sigma)_{\rm max} = 0.001$
direct methods	$\Delta \rho_{\rm max} = 0.82 \text{ e } \text{\AA}^{-3}$
	$\Delta \rho_{\rm min} = -0.37 \ {\rm e} \ {\rm \AA}^{-3}$

Special details

Geometry. Bond distances, angles *etc*. have been calculated using the rounded fractional coordinates. All su's are estimated from the variances of the (full) variance-covariance matrix. The cell e.s.d.'s are taken into account in the estimation of distances, angles and torsion angles

Refinement. Refinement of F^2 against ALL reflections. The weighted *R*-factor *wR* and goodness of fit *S* are based on F^2 , conventional *R*-factors *R* are based on *F*, with *F* set to zero for negative F^2 . The threshold expression of $F^2 > 2\sigma(F^2)$ is used only for calculating *R*-factors(gt) *etc.* and is not relevant to the choice of reflections for refinement. *R*-factors based on F^2 are statistically about twice as large as those based on *F*, and *R*-factors based on ALL data will be even larger.

Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters $(Å^2)$

	x	у	Ζ	$U_{ m iso}$ */ $U_{ m eq}$	
S2	0.18125 (14)	0.34721 (13)	0.46715 (10)	0.0405 (3)	
F4	0.7221 (6)	0.3280 (5)	-0.0587 (3)	0.0829 (16)	
015	0.7912 (4)	-0.1016 (4)	0.3039 (3)	0.0520 (12)	
N1	0.2057 (5)	0.0276 (4)	0.3682 (3)	0.0387 (10)	
N3	0.4983 (5)	0.2512 (4)	0.4147 (3)	0.0349 (10)	
C2	0.3046 (5)	0.2035 (5)	0.4140 (3)	0.0351 (11)	
C4	0.6032 (5)	0.1414 (5)	0.3474 (3)	0.0344 (11)	
C5	0.4912 (5)	-0.0550 (5)	0.3312 (3)	0.0346 (11)	
C6	0.2961 (6)	-0.1043 (5)	0.3369 (3)	0.0352 (11)	
C15	0.6152 (6)	-0.1729 (5)	0.3037 (4)	0.0394 (14)	
C16	0.5357 (6)	-0.3717 (6)	0.2740 (4)	0.0455 (14)	
C41	0.6337 (6)	0.1890 (5)	0.2361 (4)	0.0380 (11)	
C42	0.4755 (7)	0.1672 (6)	0.1611 (4)	0.0450 (14)	
C43	0.5034 (8)	0.2138 (7)	0.0611 (4)	0.0521 (17)	
C44	0.6923 (9)	0.2814 (7)	0.0386 (4)	0.0581 (17)	
C45	0.8513 (8)	0.3036 (7)	0.1095 (5)	0.0604 (17)	

C46	0.8225 (7)	0.2573 (6)	0.2092 (4)	0.0473 (14)
C61	0.1533 (5)	-0.2885 (5)	0.3148 (4)	0.0403 (13)
H1	0.087 (7)	0.005 (6)	0.370 (4)	0.036 (12)*
H3	0.572 (8)	0.356 (8)	0.446 (4)	0.051 (14)*
H4	0.73478	0.16638	0.38822	0.0414*
H16A	0.45343	-0.41343	0.20285	0.0679*
H16B	0.45700	-0.40981	0.33158	0.0679*
H16C	0.64443	-0.42238	0.26819	0.0679*
H42	0.34597	0.11967	0.17837	0.0543*
H43	0.39509	0.19929	0.00998	0.0629*
H45	0.98007	0.35005	0.09089	0.0727*
H46	0.93223	0.27237	0.25940	0.0573*
H61A	0.15509	-0.35405	0.23847	0.0604*
H61B	0.02124	-0.28040	0.32542	0.0604*
H61C	0.19019	-0.35119	0.36620	0.0604*

Atomic displacement parameters (A^2)

	U^{11}	U^{22}	U^{33}	U^{12}	U^{13}	U^{23}
S2	0.0293 (5)	0.0314 (5)	0.0635 (7)	0.0129 (4)	0.0089 (4)	0.0121 (4)
F4	0.093 (3)	0.086 (3)	0.062 (2)	0.003 (2)	0.0100 (18)	0.0354 (18)
O15	0.0259 (14)	0.0370 (16)	0.099 (3)	0.0146 (12)	0.0106 (15)	0.0218 (16)
N1	0.0226 (16)	0.0336 (17)	0.061 (2)	0.0104 (13)	0.0064 (14)	0.0118 (15)
N3	0.0276 (16)	0.0267 (16)	0.052 (2)	0.0093 (13)	0.0063 (14)	0.0114 (14)
C2	0.0295 (18)	0.0285 (18)	0.051 (2)	0.0105 (15)	0.0064 (16)	0.0149 (16)
C4	0.0245 (17)	0.0282 (18)	0.053 (2)	0.0101 (14)	0.0051 (15)	0.0127 (16)
C5	0.0275 (18)	0.0242 (17)	0.054 (2)	0.0083 (14)	0.0037 (15)	0.0133 (15)
C6	0.0305 (18)	0.0291 (18)	0.050 (2)	0.0115 (15)	0.0048 (15)	0.0148 (16)
C15	0.0275 (19)	0.040 (2)	0.059 (3)	0.0162 (16)	0.0079 (17)	0.0208 (18)
C16	0.033 (2)	0.038 (2)	0.072 (3)	0.0183 (17)	0.0082 (19)	0.017 (2)
C41	0.037 (2)	0.0273 (18)	0.054 (2)	0.0138 (15)	0.0089 (17)	0.0129 (16)
C42	0.039 (2)	0.043 (2)	0.056 (3)	0.0142 (18)	0.0075 (19)	0.0161 (19)
C43	0.056 (3)	0.049 (3)	0.053 (3)	0.017 (2)	0.000 (2)	0.016 (2)
C44	0.073 (3)	0.051 (3)	0.047 (3)	0.010 (2)	0.011 (2)	0.017 (2)
C45	0.055 (3)	0.059 (3)	0.059 (3)	0.001 (2)	0.015 (2)	0.019 (2)
C46	0.037 (2)	0.046 (2)	0.057 (3)	0.0072 (18)	0.0087 (19)	0.016 (2)
C61	0.0263 (18)	0.0289 (19)	0.067 (3)	0.0078 (15)	0.0098 (17)	0.0148 (18)

Geometric parameters (Å, °)

S2—C2	1.692 (4)	C41—C42	1.390 (7)	
F4—C44	1.360 (6)	C42—C43	1.392 (7)	
O15—C15	1.226 (5)	C43—C44	1.372 (9)	
N1—C2	1.366 (5)	C44—C45	1.363 (8)	
N1—C6	1.395 (5)	C45—C46	1.388 (8)	
N3—C2	1.329 (5)	C4—H4	1.0000	
N3—C4	1.466 (5)	C16—H16A	0.9800	
N1—H1	0.82 (5)	C16—H16B	0.9800	
N3—H3	0.85 (6)	C16—H16C	0.9800	
C4—C41	1.534 (6)	C42—H42	0.9500	

C4—C5	1.522 (6)	C43—H43	0.9500
C5—C6	1.349 (6)	C45—H45	0.9500
C5—C15	1.479 (6)	C46—H46	0.9500
C6—C61	1.501 (6)	C61—H61A	0.9800
C15—C16	1.494 (6)	C61—H61B	0.9800
C41—C46	1.391 (7)	C61—H61C	0.9800
C2—N1—C6	124.1 (4)	C43—C44—C45	122.8 (5)
C2—N3—C4	124.0 (3)	F4—C44—C45	118.6 (6)
C2—N1—H1	113 (3)	C44—C45—C46	119.1 (5)
C6—N1—H1	123 (3)	C41—C46—C45	120.3 (5)
C2—N3—H3	123 (4)	N3—C4—H4	108.00
C4—N3—H3	113 (4)	C5—C4—H4	108.00
S2—C2—N3	123.5 (3)	C41—C4—H4	108.00
N1—C2—N3	116.3 (3)	C15—C16—H16A	110.00
S2—C2—N1	120.3 (3)	C15—C16—H16B	109.00
N3—C4—C5	109.7 (3)	C15—C16—H16C	109.00
N3—C4—C41	110.6 (3)	H16A—C16—H16B	109.00
C5—C4—C41	111.5 (3)	H16A—C16—H16C	109.00
C4—C5—C15	113.3 (3)	H16B—C16—H16C	109.00
C6—C5—C15	127.3 (4)	C41—C42—H42	119.00
C4—C5—C6	119.3 (3)	C43—C42—H42	119.00
N1—C6—C5	118.9 (4)	C42—C43—H43	121.00
N1—C6—C61	112.2 (4)	C44—C43—H43	121.00
C5—C6—C61	128.9 (4)	C44—C45—H45	120.00
C5-C15-C16	123.3 (4)	C46—C45—H45	120.00
O15—C15—C5	117.4 (4)	C41—C46—H46	120.00
O15—C15—C16	119.3 (4)	C45—C46—H46	120.00
C4—C41—C46	120.0 (4)	C6—C61—H61A	109.00
C42—C41—C46	118.8 (4)	C6—C61—H61B	109.00
C4—C41—C42	121.2 (4)	C6—C61—H61C	109.00
C41—C42—C43	121.2 (5)	H61A—C61—H61B	109.00
C42—C43—C44	117.9 (5)	H61A—C61—H61C	110.00
F4—C44—C43	118.6 (5)	H61B—C61—H61C	109.00
C6—N1—C2—S2	-170.5 (3)	C4—C5—C6—C61	174.5 (4)
C6—N1—C2—N3	8.8 (6)	C15—C5—C6—N1	177.7 (4)
C2—N1—C6—C5	-13.1 (6)	C15—C5—C6—C61	-1.6 (7)
C2—N1—C6—C61	166.4 (4)	C4—C5—C15—O15	4.4 (6)
C4—N3—C2—S2	-165.1 (3)	C4—C5—C15—C16	-174.3 (4)
C4—N3—C2—N1	15.6 (5)	C6-C5-C15-O15	-179.3 (4)
C2—N3—C4—C5	-31.5 (5)	C6-C5-C15-C16	1.9 (7)
C2—N3—C4—C41	91.8 (4)	C4—C41—C42—C43	178.7 (4)
N3—C4—C5—C6	25.6 (5)	C46—C41—C42—C43	-0.6 (7)
N3—C4—C5—C15	-157.8 (3)	C4—C41—C46—C45	-178.9 (4)
C41—C4—C5—C6	-97.2 (4)	C42—C41—C46—C45	0.4 (7)
C41—C4—C5—C15	79.4 (4)	C41—C42—C43—C44	0.3 (8)
N3—C4—C41—C42	-61.6 (5)	C42—C43—C44—F4	-179.9 (5)
N3—C4—C41—C46	117.7 (4)	C42—C43—C44—C45	0.2 (8)

supplementary materials

C5—C4—C41—C42	60.6 (5)	F4—C44—C45—C46	179.7 (5)
C5—C4—C41—C46	-120.1 (4)	C43—C44—C45—C46	-0.4 (9)
C4—C5—C6—N1	-6.2 (5)	C44—C45—C46—C41	0.1 (8)

Hydrogen-bond geometry (Å, °)

Cg2 is the centroid of the C41–C46 benzene ring.

D—H···A	<i>D</i> —Н	H···A	$D^{\dots}A$	D—H···A
N1—H1…O15 ⁱ	0.82 (5)	2.10 (5)	2.867 (5)	156 (5)
N3—H3···S2 ⁱⁱ	0.85 (6)	2.50 (6)	3.350 (4)	174 (6)
C16—H16A…F4 ⁱⁱⁱ	0.98	2.50	3.358 (6)	146
C61—H61A····F4 ⁱⁱⁱ	0.98	2.47	3.319 (6)	145
C61—H61 <i>B</i> ···O15 ⁱ	0.98	2.54	3.383 (5)	144
C16—H16 C ··· $Cg2^{iv}$	0.98	2.92	3.633 (5)	130

Symmetry codes: (i) *x*-1, *y*, *z*; (ii) -*x*+1, -*y*+1, -*z*+1; (iii) -*x*+1, -*y*, -*z*; (iv) *x*, *y*-1, *z*.