

## Article

# Processing of Waste from Enrichment with the Production of Cement Clinker and the Extraction of Zinc

Alexandr Kolesnikov <sup>1,2,\*</sup>, Roman Fediuk <sup>3,4</sup> , Olga Kolesnikova <sup>2,\*</sup>, Nurgali Zhanikulov <sup>2,5</sup>, Bibol Zhakipbayev <sup>2</sup>, Rasim Kuraev <sup>2</sup>, Elmira Akhmetova <sup>2</sup> and Aizhan Shal <sup>1,2</sup>

<sup>1</sup> Department of “Life Safety and Environmental Protection” M. Auezov South Kazakhstan University, Building B, Av. Tauke Khan, 5, Shymkent 160012, Kazakhstan; shala96@rambler.ru

<sup>2</sup> Department of Science of Production and Innovation, M. Auezov South Kazakhstan University, Shymkent 160012, Kazakhstan; zhn94@ro.ru (N.Z.); bezh84@lenta.ru (B.Z.); rmkuraev@yandex.kz (R.K.); emia86@rambler.ru (E.A.)

<sup>3</sup> Polytechnic Institute, Far Eastern Federal University, 690922 Vladivostok, Russia; fedyuk.rs@dvfu.ru

<sup>4</sup> Institute of Civil Engineering, Peter the Great St. Petersburg Polytechnic University, 195251 St. Petersburg, Russia

<sup>5</sup> Department of Inorganic and Technical Chemistry, Academician E.A. Buketov Karaganda University, Karaganda 100024, Kazakhstan

\* Correspondence: kas164@yandex.kz (A.K.); ogkolesnikova@yandex.kz (O.K.); Tel.: +7-7052566897 (A.K.)

† These authors contributed equally to this work.

**Abstract:** This paper presents studies on the processing of enrichment tailings as a component of a raw mixture in order to obtain cement clinker, with simultaneous distillation of zinc. Thermodynamic studies were carried out in the temperature range of 600–1600 °C using the software application “HSC Chemistry 6” developed by the metallurgical company Outokumpu (Finland). As a result of the conducted studies, we found that zinc contributes to the intensification of mineral formation of cement clinker. In particular, it was found that the formation of belite is possible in the temperature range from 990.7 to 1500 °C with Gibbs energy values of  $-0.01$  and  $-323.8$  kJ (which is better than the standard process by  $-11.4$  kJ), respectively; the formation of alite is possible in the temperature range from 982.9 to 1500 °C with Gibbs energy values of  $-0.05$  and  $-402.1$  kJ (better than the standard process by  $-11.4$  kJ), respectively; the formation of tricalcium aluminate is thermodynamically possible in the temperature range from 600 °C at  $\Delta G_T^\circ = -893.8$  kJ to 1500 °C at  $\Delta G_T^\circ = -1899.3$  kJ (better than the standard process by  $-1570.1$  kJ), respectively; and the formation of four calcium aluminoferrite is possible in the temperature range from 600 °C at  $\Delta G_T^\circ = -898.9$  kJ to 1500 °C at  $\Delta G_T^\circ = -1959.3$  kJ (better than the standard process by  $-1570.2$  kJ), respectively, with simultaneous distillation of zinc into a gaseous state for its further capture.

**Keywords:** thermodynamics; technogenic material; enrichment tailings; cement clinker; zinc metasilicate; zinc sublimates



**Citation:** Kolesnikov, A.; Fediuk, R.; Kolesnikova, O.; Zhanikulov, N.; Zhakipbayev, B.; Kuraev, R.; Akhmetova, E.; Shal, A. Processing of Waste from Enrichment with the Production of Cement Clinker and the Extraction of Zinc. *Materials* **2022**, *15*, 324. <https://doi.org/10.3390/ma15010324>

Academic Editors: Edward Bormashenko and Kevin M. Van Geem

Received: 28 October 2021

Accepted: 28 December 2021

Published: 3 January 2022

**Publisher’s Note:** MDPI stays neutral with regard to jurisdictional claims in published maps and institutional affiliations.



**Copyright:** © 2022 by the authors. Licensee MDPI, Basel, Switzerland. This article is an open access article distributed under the terms and conditions of the Creative Commons Attribution (CC BY) license (<https://creativecommons.org/licenses/by/4.0/>).

## 1. Introduction

The mining and metallurgical industry is the leading sector of the Kazakh economy, accounting for 15.2% of the total industrial production. Many mining and metallurgical industries are city-forming, and therefore the development of the mining and metallurgical complex is the development of cities, jobs for more than 200 thousand people, and the accumulation of a huge amount of waste from enrichment [1,2]. In the CIS countries, the total volume of mined mining, solid minerals are about 3.5 billion m<sup>3</sup> per year, and taking into account mining and preparation and processing works—about 5 billion m<sup>3</sup>, that is, 1.5 billion m<sup>3</sup> of rocks is mined incidentally (the bulk of which, after enrichment, is stored in dumps and tailing dumps) in order to ensure the extraction of the main minerals from the ground.

At the moment, it is obvious that the modern world, as before, continues to need resources, which in the future will have a separate place in the development of the world

global economy and the economy of Kazakhstan [3,4]. However, at the same time, it is necessary to critically rethink the organization of the raw material industries, namely, to revise the approaches to the management of natural resources. It is necessary to actively introduce integrated information technology platforms for natural and artificial raw materials, as well as to significantly increase the requirements for energy efficiency and resource conservation of industrial enterprises of the chemical, energy, mining, and metallurgical industries, alongside their environmental friendliness and efficiency [1,2,5–14].

Currently, due to the decrease in valuable metals in ores and the increasing amount of refractory raw materials [15–17], it is economically feasible and necessary to comprehensively process both poor, substandard, and hard-to-reach mineral raw materials and technogenic, as secondary raw materials [18,19], found in particular in dumps and tailings in countries such as Kazakhstan, Uzbekistan, Russia, Finland, Poland, Ukraine, Canada, and Argentina.

One such form of waste is waste from the mining and processing and metallurgical industries—tailings from the enrichment of the Balkhash Concentration Plant (BCP) in Kazakhstan [20,21], which contain in their chemical composition a number of useful compounds, in particular, silicon oxides, aluminum, and iron, which are necessary in the production of cement clinker, as well as zinc compounds, which are some of the valuable non-ferrous metals in the metallurgical industry [22,23].

It is known that the main resources in the production of binders were previously provided by traditional mineral raw materials. In the new economic realities, this approach is qualitatively changing, and industrial wastes act as secondary raw materials [15,18–20,24–26]. Moreover, the cost of such raw materials is much lower, and the processing conditions are often simple. These features of economic development insistently require a high-quality study of all types of accumulated and unused waste.

Portland cement clinker consists of a close connection of four crystalline phases of alite, belite, tricalcium aluminate (aluminate phase), and four-calcium aluminoferrite (ferrite phase). The names alite, belite, aluminate and ferrite are used to distinguish them from pure phases and to distinguish them from other ions. In addition, Portland cement clinker contains a small portion of free calcium oxide and a small amount of periclase (MgO).

Portland cement clinker is produced by the synthesis of a precisely composed raw material mixture (raw flour, wet raw mass, or raw sludge). It contains elements, usually oxides, namely, CaO, SiO<sub>2</sub>, Al<sub>2</sub>O<sub>3</sub>, and Fe<sub>2</sub>O<sub>3</sub>, as well as a small amount of other materials. Raw flour, wet raw mass, or raw sludge are finely ground and thoroughly mixed. Portland cement clinker is a hydraulic material that must consist of at least two-thirds of calcium silicates (3CaO·SiO<sub>2</sub> and 2CaO·SiO<sub>2</sub>) after firing. The rest consists of aluminum and iron-containing clinker phases and other compounds. The mass ratio of CaO/SiO<sub>2</sub> should be at least 2.0. The mass of magnesium oxide (MgO) cannot exceed 5.0% [20–25].

In studies, using the calculation of the Gibbs energy ( $\Delta G_T^0$ ), we simulated the process of simultaneous synthesis of the formation of the main minerals of cement clinker (Ca<sub>2</sub>SiO<sub>4</sub>, Ca<sub>3</sub>SiO<sub>5</sub>, 3CaO·Al<sub>2</sub>O<sub>3</sub>, 4CaO·Al<sub>2</sub>O<sub>3</sub>·Fe<sub>2</sub>O<sub>3</sub>) and zinc sublimates depending on the temperature from limestone and artificial tailings from the beneficiation of non-ferrous metals at the Balkhash dressing plant. The studies were carried out in order to study various dependencies, patterns, and mechanisms during the formation of a number of mineral compounds of cement clinker and zinc stripping.

Thus, research and scientific work aimed at reducing energy costs and unit costs of raw materials, involving technogenic raw materials in the production cycle as a secondary raw material, while reducing the harmful impact on the environment through waste processing, are relevant, new, and require a comprehensive qualitative study for their further development and implementation in production.

## 2. Materials and Methods

Thermodynamic studies were carried out using the “Thermodynamics” (Moscow, Russia) software [26] and the “HSC Chemistry 6” software complex, developed by the

metallurgical company Outokumpu (Helsinki, Finland). The software package used in this work is based on the ideology of the European consortium SGTE (Scientific Group Thermodata Europe, Stockholm, Sweden), which develops, maintains, and distributes high-quality databases. The SGTE structure is represented by specialized research centers in Germany, Canada, France, Sweden, Great Britain, and the USA. The database of the software package contains information on 22,000 individual substances [27].

To calculate the thermodynamic functions characterizing an individual substance, we used the standard values of enthalpy  $H_{298}$ ; entropy  $S_{298}$ ; and polynomial coefficients  $A$ ,  $B$ ,  $C$ , and  $D$  stored in the database, from which the molar heat capacity was calculated at an arbitrarily specified temperature  $T$  in accordance with Equation (1).

The enthalpy of an individual substance at a temperature  $T$ , which differs from the standard one, equal to 298 K, was calculated by the formula:

$$H_T = H_{298} + \int_{298}^T C_p dT + \sum H_F, \quad (1)$$

where  $H_{298}$  is the enthalpy value of a given substance under standard conditions;  $C_p$ —molar heat capacity;  $\sum H_F$ —enthalpy of phase transitions (polymorphic transformations, melting, evaporation).

The entropy is defined as

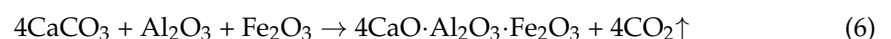
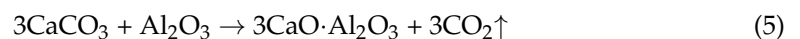
$$S_T = S_{298} + \int_{298}^T \frac{C_p}{T} dT + \frac{\sum H_F}{T} \quad (2)$$

where  $S_{298}$  is the value of the entropy of a given substance under standard conditions;  $C_p$ —molar heat capacity;  $\frac{\sum H_F}{T}$ —entropy of phase transitions (polymorphic transformations, melting, evaporation).

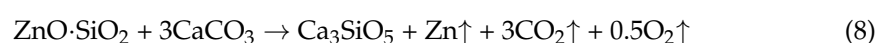
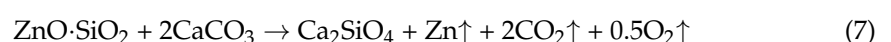
As a mineral raw material, limestone of the Mynaral deposit was considered, containing 98.3% Ca, Si, C, and O in its elemental and chemical composition, and as technogenic raw materials, tailings of copper ores from the Balkhash concentrating plant, containing 90.43% consisting of Si, Fe, Ca, Al, and O in its composition; 7.76% of Na, Mg, S, K, and Ti; 1.37% Zn; and 0.44% Pb.

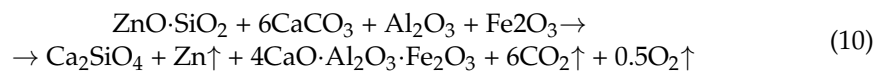
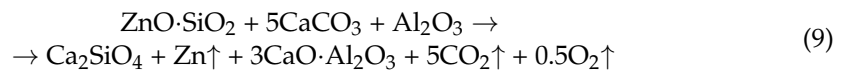
### 3. Results

As a result of the conducted thermodynamic studies, the Gibbs energy ( $\Delta G_T^\circ$ ) was calculated under conditions of modeling the synthesis of the formation of the main minerals of cement clinker ( $\text{Ca}_2\text{SiO}_4$ ,  $\text{Ca}_3\text{SiO}_5$ ,  $3\text{CaO}\cdot\text{Al}_2\text{O}_3$ ,  $4\text{CaO}\cdot\text{Al}_2\text{O}_3\cdot\text{Fe}_2\text{O}_3$ ) and zinc sublimates (Zn) in the temperature range 600–1600 °C. Limestone (containing mainly calcium carbonate) and technogenic (tailings from the enrichment of non-ferrous metals containing such compounds as  $\text{ZnO}\cdot\text{SiO}_2$  [21]) were the raw materials, considered groups of chemical reactions. In particular, group I—standard reactions (3)–(6) for the formation of clinker minerals [28–30]:

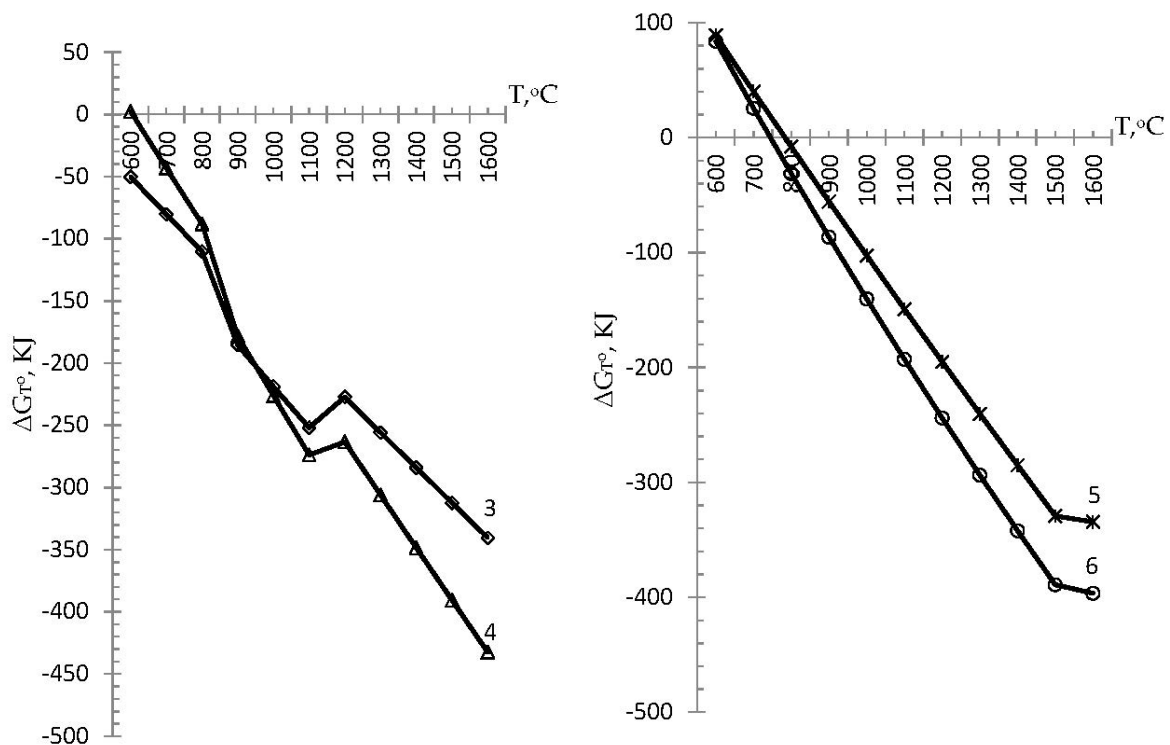


Group II is represented by non-standard reactions (7)–(10) of formation of clinker minerals in the presence of zinc metasilicate contained in the tailings:





On the basis of the results of the calculations of standard reactions of group I, we plotted the dependence of the Gibbs energy on temperature (the possibility of reactions (3)–(6) with the formation of cement clinker minerals) in Figure 1. Figure 1 shows that reaction (3) with the formation of  $\text{Ca}_2\text{SiO}_4$  and (4) with the formation of  $\text{Ca}_3\text{SiO}_5$  proceeded practically in the entire investigated temperature range. In this case, the Gibbs energy of reaction (3) was in the negative range, which indicated the possibility of a complete reaction with the formation of belite ( $\text{Ca}_2\text{SiO}_4$ ) within  $\Delta G_T^\circ$  from  $-50.2$  to  $-340.5$  kJ at 600 and 1600 °C, respectively. In contrast to reaction (3), the Gibbs energy of reaction (4) at the beginning of the investigated temperature interval had a positive value of  $\Delta G_T^\circ$  equal to 2.3 kJ at  $T = 600$  °C, which indicates that under these conditions, the present reaction does not proceed. The onset temperature of reaction (4) was 605.2 °C when a negative value of Gibbs energy appeared, in particular, the occurrence of reaction (4) and the formation of alite ( $\text{Ca}_3\text{SiO}_5$ ) was found to be possible in the temperature range 605.2–1600 °C with  $\Delta G_T^\circ$  from  $-0.006$  to  $-432.5$  kJ, respectively. At the same time, for reactions (5) and (6) shown in Figure 1 had almost the same dependence of the Gibbs energy on temperature, having almost the same positive values at 600 °C 89 and 83.9 kJ, respectively, which corresponded to their limited occurrence at the initial temperature of the research.



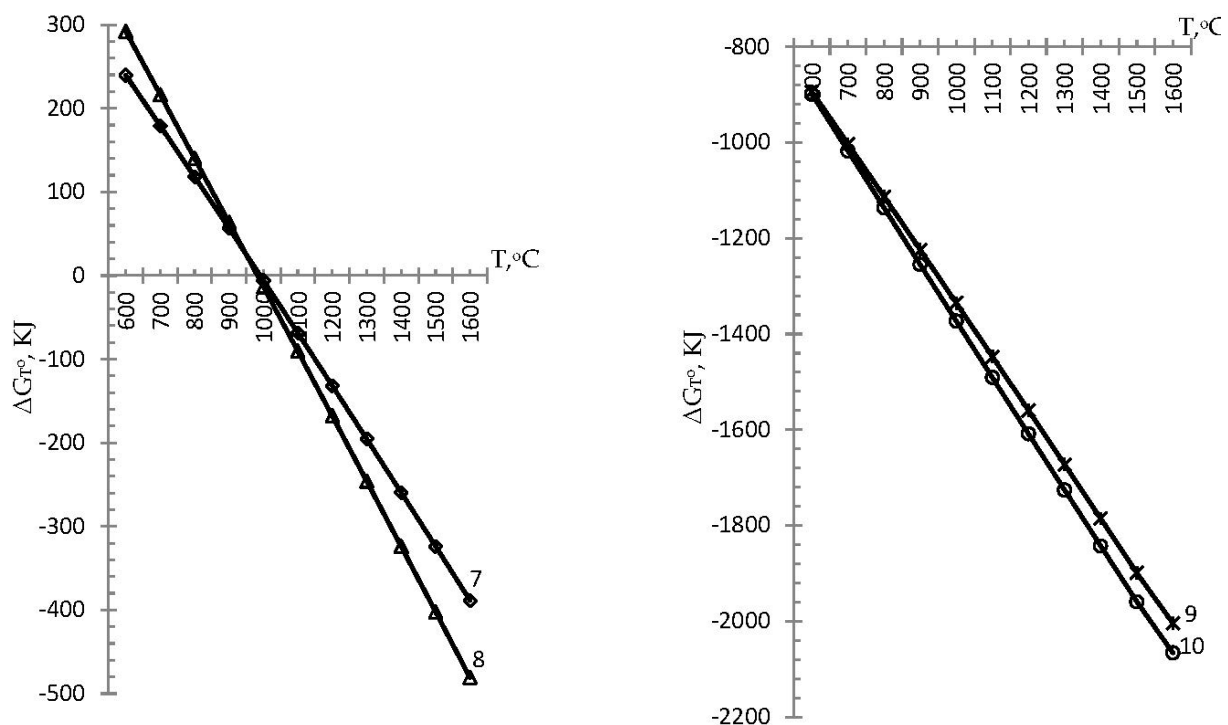
- (3)  $2\text{CaCO}_3 + \text{SiO}_2 \rightarrow \text{Ca}_2\text{SiO}_4 + 2\text{CO}_2$ ; (4)  $3\text{CaCO}_3 + \text{SiO}_2 \rightarrow \text{Ca}_3\text{SiO}_5 + 3\text{CO}_2$ ; (5)  $3\text{CaCO}_3 + \text{Al}_2\text{O}_3 \rightarrow 3\text{CaO} \cdot \text{Al}_2\text{O}_3 + 3\text{CO}_2$ ;  
 (6)  $4\text{CaCO}_3 + \text{Al}_2\text{O}_3 + \text{Fe}_2\text{O}_3 \rightarrow 4\text{CaO} \cdot \text{Al}_2\text{O}_3 \cdot \text{Fe}_2\text{O}_3 + 4\text{CO}_2$ .

**Figure 1.** Influence of temperature on the Gibbs energy of the possibility of standard reactions with the formation of clinker minerals.

The Gibbs energy of reaction (5) with the possible formation of the clinker mineral  $3\text{CaO} \cdot \text{Al}_2\text{O}_3$  became negative at a temperature of 783.58 °C and was  $-0.001$  kJ, reaching

−334.4 kJ at  $T = 1600\text{ }^{\circ}\text{C}$ . The Gibbs energy of reaction (6) with the possible formation of the clinker mineral  $4\text{CaO}\cdot\text{Al}_2\text{O}_3\cdot\text{Fe}_2\text{O}_3$  became negative at a temperature of  $746\text{ }^{\circ}\text{C}$  and amounted to  $-0.528\text{ kJ}$ , reaching a value of  $-396.5\text{ kJ}$  at a maximum research temperature of  $1600\text{ }^{\circ}\text{C}$ .

On the basis of the results of the calculations of the reactions of group II, we plotted the graphs of the Gibbs energy dependence on temperature (the possibility of reactions (7)–(10) in the presence of zinc metasilicate with the formation of cement clinker minerals and zinc stripping) in Figure 2. Figure 2 shows that the reaction (7) with the distillation of zinc, the formation of  $\text{Ca}_2\text{SiO}_4$ , and (8) with the formation of  $\text{Ca}_3\text{SiO}_5$  with the simultaneous distillation of zinc did not occur in the entire investigated temperature range. Thus, the Gibbs energy of reaction (7) was in the positive range at  $T = 600\text{--}990.6\text{ }^{\circ}\text{C}$ , which corresponded to its impossible occurrence in this temperature range. With a further increase in temperature, the Gibbs energy of reaction (7) became negative, reaching a value of  $-0.015\text{ kJ}$  at  $T = 990.7\text{ }^{\circ}\text{C}$  and becoming more negative at a maximum study temperature of  $1600\text{ }^{\circ}\text{C}$ , reaching a value of  $-388.6\text{ kJ}$ , which indicates that the reaction proceeded in the range of  $990.7\text{--}1600\text{ }^{\circ}\text{C}$  with simultaneous stripping of zinc into the gas phase and the formation of the clinker mineral belite— $\text{Ca}_2\text{SiO}_4$ . Similarly to reaction (7), the Gibbs energy of reaction (8) at the beginning of the investigated temperature interval had a positive value of  $\Delta G_{\text{T}}^{\circ}$  equal to  $292\text{ kJ}$  at  $T = 600\text{ }^{\circ}\text{C}$ , which indicates that the present reaction did not proceed under these conditions.



(7)  $\text{ZnO}\cdot\text{SiO}_2 + 2\text{CaCO}_3 \rightarrow \text{Ca}_2\text{SiO}_4 + \text{Zn} + 2\text{CO}_2 + 0.5\text{O}_2$ ; (8)  $\text{ZnO}\cdot\text{SiO}_2 + 3\text{CaCO}_3 \rightarrow \text{Ca}_3\text{SiO}_5 + \text{Zn} + 3\text{CO}_2 + 0.5\text{O}_2$ ;  
 (9)  $\text{ZnO}\cdot\text{SiO}_2 + 5\text{CaCO}_3 + \text{Al}_2\text{O}_3 \rightarrow \text{Ca}_2\text{SiO}_4 + \text{Zn} + 3\text{CaO}\cdot\text{Al}_2\text{O}_3 + 5\text{CO}_2 + 0.5\text{O}_2$ ; (10)  $\text{ZnO}\cdot\text{SiO}_2 + 6\text{CaCO}_3 + \text{Al}_2\text{O}_3 + \text{Fe}_2\text{O}_3 \rightarrow \text{Ca}_2\text{SiO}_4 + \text{Zn} + 4\text{CaO}\cdot\text{Al}_2\text{O}_3\cdot\text{Fe}_2\text{O}_3 + 6\text{CO}_2 + 0.5\text{O}_2$ .

**Figure 2.** Influence of temperature on the Gibbs energy of the possibility of reactions in the presence of zinc metasilicate with the formation of clinker minerals and the simultaneous stripping of zinc.

The temperature of the onset of reaction (8) was  $982.9\text{ }^{\circ}\text{C}$  when a negative value of Gibbs energy appeared; in particular, the occurrence of reaction (8) with the transformation of zinc into a gaseous state and with the formation of the clinker mineral alite ( $\text{Ca}_3\text{SiO}_5$ ) was possible in the temperature range of  $982.9\text{--}1600\text{ }^{\circ}\text{C}$  with  $\Delta G_{\text{T}}^{\circ}$  from  $-0.05$  to  $-480.7\text{ kJ}$ , respectively. In this case, the Gibbs energy of reactions (7) and (8) at a temperature of



1000 °C became almost the same as the curves began to intersect, wherein the Gibbs energy of reaction (8) became more negative in contrast to reaction (7).

#### 4. Discussion

From the given results of the Gibbs energy of formation according to standard reactions (3)–(6) of cement clinker minerals and according to non-standard reactions (7)–(10) with simultaneous distillation of zinc, it can be seen that both groups of reactions of formation of clinker minerals are capable of proceeding within the studied temperature. Thus, in particular, the Gibbs energy of the standard reaction (3) with the formation of belite ( $\text{Ca}_2\text{SiO}_4$ ) had negative values in the entire investigated temperature range and ranged from  $-50.2$  to  $-340.5$  kJ, which indicates the possibility of the complete course of the reaction. In contrast to the standard reaction (3), the non-standard reaction (7) with the formation of  $\text{Ca}_2\text{SiO}_4$  in the presence of zinc metasilicate did not occur in the entire investigated temperature range, and it was limited at the initial stage of modeling. The possibility of reaction (7) was observed at  $T = 990.7$  °C with a negative value of the Gibbs energy ( $-0.015$  kJ) and became more negative, reaching a value of  $-388.6$  kJ at a maximum research temperature of  $1600$  °C. From this comparison, it can be seen that the standard reaction (3) had an advantage over the non-standard reaction (7) in the temperature range of  $600$ – $1400$  °C; then, they changed places, and the advantage of the reaction (7) was observed at a temperature of  $1500$ – $1600$  °C (with the advantage of energy values Gibbs at  $-11.1$  and  $-48.1$  kJ). The values of the Gibbs energy of the standard reaction (4) for the formation of alite ( $\text{Ca}_3\text{SiO}_5$ ) at  $T = 600$  °C had a positive value of  $\Delta G_T^\circ$  equal to  $2.3$  kJ and was limited. In particular, the course of reaction (4) and the formation of alite ( $\text{Ca}_3\text{SiO}_5$ ) was possible in the temperature range  $605.2$ – $1600$  °C with Gibbs energy from  $-0.006$  to  $-432.5$  kJ, respectively. The Gibbs energy of the non-standard reaction (8) with the formation of alite and the distillation of zinc, similarly to the standard reaction (4), was limited at the initial stage and was capable of proceeding in the temperature range  $982.9$ – $1600$  °C, with a Gibbs energy from  $-0.05$  to  $-480.7$  kJ, respectively. Comparing the standard reaction (4) with the non-standard reaction (8), one can see the advantage of reaction (4) at a temperature of  $600$ – $1400$  °C; then, at a temperature of  $1500$ – $1600$  °C, the advantage of proceeding goes to a non-standard reaction (8) (with the advantage of Gibbs energy values by  $-11$ ,  $4$ , and  $-48.2$  kJ). The Gibbs energy of the standard reaction (5) with the formation of the clinker mineral  $3\text{CaO}\cdot\text{Al}_2\text{O}_3$  was limited and was capable of proceeding in the temperature range  $783.58$ – $1600$  °C, amounting to  $-0.001$  kJ and  $-334.4$  kJ, respectively. In contrast to the standard reaction (5), the non-standard reaction (9) proceeded over the entire temperature range under study, with negative Gibbs energies from  $-893$  kJ to  $-20,041.1$  kJ, respectively. When comparing the values of reactions (5) and (9), a significant advantage of the non-standard reaction (9) was obvious. The Gibbs energy of the standard reaction (6) was positive, which made the course of the reaction of formation of the clinker mineral  $4\text{CaO}\cdot\text{Al}_2\text{O}_3\cdot\text{Fe}_2\text{O}_3$  limited. Reaction (6) was capable of proceeding in the temperature range  $746$ – $1600$  °C with Gibbs energies of  $-0.528$  kJ and  $-396.5$  kJ, respectively. Non-standard reaction (10) proceeded in the entire investigated temperature range, having negative Gibbs energies from  $-898.9$  kJ to  $-20,066.2$  kJ, respectively, having a significant advantage over the standard reaction (6). The data obtained were consistent with the research carried out by a number of scientists [25,26,31–40] and complement them. Regarding the distillation of non-ferrous metals, in particular, zinc, in the simulated reactions, it was shown that non-ferrous metals in the form of zinc are able to pass into the gas phase and be captured in bag filters. There are known studies by a group of authors [41,42], where experimental studies of the formation of solid solutions in zinc and lead oxides with partial substitution of non-ferrous metals of the main elements in cement clinker minerals are given. However, the manuscript considers a model with zinc metasilicate, and experimental studies of zinc behavior are envisaged in upcoming laboratory experiments.

This research will contribute to the development of a technology for the integrated processing of waste (tailings) from enrichment, as a secondary raw material with the production of cement clinker and with the simultaneous distillation and capture of non-ferrous metals in order to reduce the anthropogenic load on the natural environment of the region.

## 5. Conclusions

In group II, non-standard reactions (7)–(10) of the formation of the main minerals of cement clinker and the stripping of zinc, in particular, the occurrence of reaction (7) with the formation of belite, is possible in the temperature range from 990.7 to 1500 °C with Gibbs energies of  $-0.01$  and  $-323.8$  kJ, respectively, better than the standard process (reaction (3)) at  $-11.4$  kJ:

- the formation of alite during the course of reaction (8) was found to be possible in the temperature range from 982.9 to 1500 °C with Gibbs energies of  $-0.05$  and  $-402.1$  kJ, respectively, better than the standard process (reaction (4)) at  $-11.4$  kJ;
- the formation of tricalcium aluminate during reaction (9) was found to be thermodynamically possible in the temperature range from 600 °C at  $\Delta G_T^\circ = -893.8$  kJ to 1500 °C at  $\Delta G_T^\circ = -1899.3$  kJ, better than the standard process (reaction (5)) at  $-1570.1$  kJ;
- the formation of four calcium alumoferrite ( $4\text{CaO}\cdot\text{Al}_2\text{O}_3\cdot\text{Fe}_2\text{O}_3$ ) during reaction (10) was found to be possible in the temperature range from 600 °C at  $\Delta G_T^\circ = -898.9$  kJ to 1500 °C at  $\Delta G_T^\circ = -1959.3$  kJ, better than the standard process (reaction (5)) at  $-1570.2$  kJ;
- the formation of the main minerals of cement clinker in reaction group II at a temperature of 1500 °C, depending on the Gibbs energy, was found to be represented by the following series  $4\text{CaO}\cdot\text{Al}_2\text{O}_3\cdot\text{Fe}_2\text{O}_3 > 3\text{CaO}\cdot\text{Al}_2\text{O}_3 > \text{Ca}_3\text{SiO}_5 > \text{Ca}_2\text{SiO}_4$ .

**Author Contributions:** Conceptualization, A.K. and O.K.; methodology, A.K.; investigation, A.K., R.F., O.K., N.Z., B.Z., E.A., R.K. and A.S.; data curation, A.K.; writing—original draft preparation, A.K.; writing—review and editing, O.K.; supervision, N.Z. and R.F.; project administration, A.K. and R.F.; funding acquisition, A.K. All authors have read and agreed to the published version of the manuscript.

**Funding:** This research was funded research by the Science Committee of the ministry of Education and Science of the Republic of Kazakhstan (grant no. AP08053015).

**Institutional Review Board Statement:** Not applicable.

**Informed Consent Statement:** Not applicable.

**Data Availability Statement:** Data sharing is not applicable to this article.

**Conflicts of Interest:** The authors declare no conflict of interest.

## References

1. Petrenko, E.S.; Vechkinzova, E.A.; Urazbekov, A.K. Context analysis and prospects of development of the mining and metallurgical industry of Kazakhstan. *Ekon. Otnosh.* **2019**, *9*, 2661–2676. [CrossRef]
2. Satbaev, B.N.; Koketaev, A.I.; Aimbetova, E.O.; Shalabaev, N.T.; Satbaev, A.B. Environmental technology for the integrated disposal of man-made wastes of the metallurgical industry: Self-curing, chemically resistant refractory mass. *Refract. Ind. Ceram.* **2019**, *60*, 318–322. [CrossRef]
3. Kazakhstan in a New Reality: Time for Action. President of Kazakhstan Kassym-Jomart Tokayev's State of the Nation Address. Available online: <https://www.akorda.kz/en/president-kassym-jomart-tokayev-delivers-his-state-of-the-nation-address-to-the-people-of-kazakhstan-28624> (accessed on 1 September 2020).
4. The Strategic Plan for Development of the Republic of Kazakhstan until the Year 2025, Has Been Approved by the Decree of the President of the Republic of Kazakhstan No. 636. Available online: <https://adilet.zan.kz/rus/docs/U1800000636> (accessed on 15 February 2018).
5. Khoroshavin, L.B.; Perpelitsyn, V.A.; Kochkin, D.K. Problems of technogenic resources. *Refract. Ind. Ceram.* **1998**, *39*, 366–368. [CrossRef]
6. Kolesnikov, A.S.; Kenzhibaeva, G.S.; Botabaev, N.E.; Kutzhanova, A.N.; Iztleuov, G.M.; Suigenbaeva, A.Z.; Ashirbaev, K.A.; Kolesnikova, O.G. Thermodynamic Modeling of Chemical and Phase Transformations in a Waelz Process-Slag—Carbon System. *Refract. Ind. Ceram.* **2020**, *61*, 289–292. [CrossRef]
7. Vasilieva, N.V.; Fedorova, E.R. Process control quality analysis. *Tsvetnye Met.* **2020**, *10*, 70–76. [CrossRef]
8. Ferreira, W.L.; Reis, É.L.; Lima, R.M. Incorporation of residues from the minero-metallurgical industry in the production of clay–lime brick. *J. Clean. Prod.* **2015**, *87*, 505–510. [CrossRef]
9. Kolesnikov, A.S.; Naraev, V.N.; Natorkhin, M.I.; Saipov, A.A.; Kolesnikova, O.G. Review of the processing of minerals and technogenic sulfide raw material with the extraction of metals and recovering elemental sulfur by electrochemical methods. *Rasayan J. Chem.* **2020**, *13*, 2420–2428. [CrossRef]
10. Peng, Z.; Gregurek, D.; Wenzl, C.; White, J.F. White Slag Metallurgy and Metallurgical Waste Recycling. *J. Metall.* **2016**, *68*, 2313–2315. [CrossRef]

11. Efremova, S. Scientific and technical solutions to the problem of utilization of waste from plant- and mineral-based industries. *Russ. J. Gen. Chem.* **2012**, *82*, 963–968. [[CrossRef](#)]
12. Mamyrbekova, A.; Mamitova, A.D.; Mamyrbekova, A. Electrochemical Behavior of Sulfur in Aqueous Alkaline Solutions. *Russ. J. Phys. Chem. A* **2018**, *92*, 582–586. [[CrossRef](#)]
13. Nadirov, K.S.; Zhantasov, M.K.; Sakybayev, B.A. The study of the gossypol resin impact on adhesive properties of the intermediate layer of the pipeline three-layer rust protection coating. *Inter. J. Adhes. Adhesiv.* **2017**, *78*, 195–199. [[CrossRef](#)]
14. Myrzabekov, B.E.; Bayeshov, A.B.; Makhanbetov, A.B.; Mishra, B.; Baigenzhenov, O.S. Dissolution of Platinum in Hydrochloric Acid Under Industrial-Scale Alternating Current Polarization. *Metal. Mat. Trans. B Proc. Metal. Mat. Proc. Sci.* **2018**, *49*, 23–27. [[CrossRef](#)]
15. Reznichenko, V.A.; Lipikhina, M.S.; Morozov, A.A. *Complex Use of Ores and Concentrates*; Nauka: Moscow, Russia, 1989; pp. 1–72.
16. Lis, T.; Nowacki, K.; Żelichowska, M.; Kania, H. Innovation in metallurgical waste management. *Metal.-Sisak Zagreb.* **2015**, *54*, 283–285.
17. Borisov, D.; Stefanov, B.; Stoyanov, S.K. An algorithm for metallurgical waste minimization. *J. Chem. Technol. Metal.* **2014**, *49*, 99–105.
18. Zharmenov, A.A. *Complex Processing of Mineral Raw Materials in Kazakhstan*; Folio: Astana, Kazakhstan, 2003; pp. 1–272.
19. Alshanov, R.A. *Kazakhstan on the World Mineral Resource Market: Problems and Their Solution*; LLP «Print-S»: Almaty, Kazakhstan, 2004; pp. 1–220.
20. Kolesnikov, A.S.; Zhakipbaev, B.Y.; Zhanikulov, N.N.; Kolesnikova, O.G.; Akhmetova, E.K.; Kuraev, R.M.; Shal, A.L. Review of technogenic waste and methods of its processing for the purpose of complex utilization of tailings from the enrichment of non-ferrous metal ores as a component of the raw material mixture in the production of cement clinker. *Rasayan J. Chem.* **2021**, *14*, 997–1005. [[CrossRef](#)]
21. Fechet, R.; Zlagnean, M.; Moanta, A.; Ciobanu, L. Mining wastes—Sampling, processing and using in anufacture portland cement. *Roman. J. Miner. Depos.* **2010**, *84*, 67–70.
22. Aghazadeh-Ghomi, M.; Pourabbas, Z. Rapid synthesis of zinc oxide nanoparticles from an alkaline zinc solution via direct precipitation. *J. Mat. Sci. Mat. Electron.* **2021**, *32*, 24363–24368. [[CrossRef](#)]
23. Kolesnikov, A.S.; Serikbaev, B.E.; Zolkin, A.L.; Kenzhibaeva, G.S. Processing of Non-Ferrous Metallurgy Waste Slag for its Complex Recovery as a Secondary Mineral Raw Material. *Refract. Ind. Ceram.* **2021**, *62*, 375–380. [[CrossRef](#)]
24. Satbaev, B.; Yefremova, S.; Zharmenov, A.; Kablanbekov, A.; Yermishin, S.; Shalabaev, N.; Satbaev, A.; Khen, V. Rice Husk Research: From Environmental Pollutant to a Promising Source of Organo-Mineral Raw Materials. *Materials* **2021**, *14*, 4119. [[CrossRef](#)]
25. Khudyakova, T.M.; Kenzhibaeva, G.S.; Kutzhanova, A.N.; Iztleuov, G.M.; Mynbaeva, E. Optimization of Raw Material Mixes in Studying Mixed Cements and Their Physicomechanical Properties. *Refract. Ind. Ceram.* **2019**, *60*, 76–81. [[CrossRef](#)]
26. Babushkin, V.I.; Matveyev, G.M.; Mchedlov-Petrosyan, O.P. *Thermodynamics of Silicates*; Stroyizdat: Moscow, Russia, 1986; 408p.
27. Roine, A. *Outokumpu HSC Chemistry for Windows. Chemical Reaction and Equilibrium Software with Extensive Thermochemical Database*; Outokumpu Research OY: Pori, Finland, 2002; 268p.
28. Zeleznik, F.J.; Gordon, S. Calculation of complex chemical equilibria. *Ind. Eng. Chem.* **1968**, *60*, 27–57. [[CrossRef](#)]
29. Rog, G.; Kozłowska-Rog, A.; Zakula-Sokol, K. Thermodynamic functions of calcium aluminate. *J. Chem. Thermodyn.* **1993**, *25*, 807–810.
30. Zhanikulov, N.N.; Khudyakova, T.M.; Taimasov, B.T. Receiving portland cement from technogenic raw materials of South Kazakhstan portlandcement. *Eurasian Chem. Technol. J.* **2019**, *21*, 334–340. [[CrossRef](#)]
31. Demidov, A.I. Thermodynamic characteristics of a quasi-binary system CaO—Si1/2O in a solid state. *Sci. Technol. Ved. SPbPU* **2018**, *24*, 134–139. [[CrossRef](#)]
32. Fediuk, R.S.; Mochalov, A.V.; Bituev, A.V.; Zayakhanov, M.E. Structuring Behavior of Composite Materials Based on Cement, Limestone, and Acidic Ash. *Inorg. Mat.* **2019**, *55*, 1079–1085. [[CrossRef](#)]
33. Pratskova, S.E.; Burmistrov, V.A.; Starikova, A.A. Thermodynamic modeling of oxide melts of CaO—Al<sub>2</sub>O<sub>3</sub>—SiO<sub>2</sub>: Systems. *Russ. J. Chem. Chem. Tech.* **2020**, *63*, 45–50. (In Russian) [[CrossRef](#)]
34. Tennis, P.D.; Jennings, H.M. A model for two types of calcium silicate hydrate in the microstructure of Portland cement pastes. *Cem. Concr. Res.* **2000**, *30*, 855–863. [[CrossRef](#)]
35. Taimasov, B.T.; Sarsenbayev, B.K.; Khudyakova, T.M. Development and testing of low-energy-intensive technology of receiving sulphate-resistant and road Portland cement. *Eurasian Chem. Technol. J.* **2017**, *19*, 347–355. [[CrossRef](#)]
36. Loganina, V.I.; Fediuk, R.S. Thermodynamic Approach to Assessing the Curing of Protective and Decorative Coatings of Exterior Walls of Buildings. *Mater. Sci. Forum.* **2020**, *974*, 3–8. [[CrossRef](#)]
37. Fukuda, K. Crystal Chemistry of Cement-Clinker Minerals and Melt Differentiation Reaction of Interstitial Melt. *Nihon Kessho Gakkaishi* **2011**, *53*, 81–85. [[CrossRef](#)]
38. Hanein, T.; Glasser, F.; Bannerman, M. Thermodynamic data for cement clinkering. *Cem. Concr. Res.* **2020**, *132*, 106043. [[CrossRef](#)]
39. Zaleska, M.; Pavlik, Z.; Pavlikova, M.; Scheinherrova, L.; Pokorný, J.; Trník, A.; Svora, P.; Fort, J.; Jankovský, O.; Suchorab, Z. Biomass ash-based mineral admixture prepared from municipal sewage sludge and its application in cement composites. *Clean Technol. Environ. Policy* **2018**, *20*, 159–171. [[CrossRef](#)]
40. Jankovský, O.; Pavlikova, M.; Sedmidubský, D.; Bousa, D.; Lojka, M.; Pokorný, J.; Zaleska, M.; Pavlik, Z. Study on pozzolana activity of wheat straw ash as potential admixture for blended cements. *Ceram.-Silik.* **2017**, *61*, 327–339. [[CrossRef](#)]
41. Zhu, J.; Chen, Y.; Zhang, L.; Yang, K.; Guan, X.; Zhao, R. Insights on Substitution Preference of Pb Ions in Sulfoaluminate Cement Clinker Phases. *Materials* **2021**, *14*, 44. [[CrossRef](#)]
42. Zhu, J.; Yang, K.; Chen, Y.; Fan, G.; Zhang, L.; Guo, B.; Guan, X.; Zhao, R. Revealing the substitution preference of zinc in ordinary Portland cement clinker phases: A study from experiments and DFT calculations. *J. Hazard. Mater.* **2021**, *409*, 124504. [[CrossRef](#)]