

# Crystal structure of 1-(2,4-dimethylphenyl)urea

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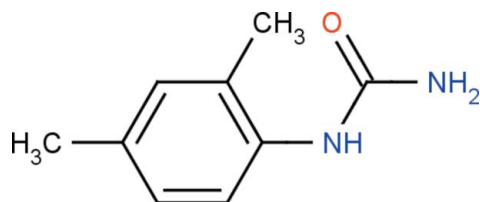
In the title urea derivative, C<sub>9</sub>H<sub>12</sub>N<sub>2</sub>O, the dihedral angle between the benzene ring and the mean plane of the urea group, N—C(=O)—N, is 86.6 (1)°. In the crystal, the urea O atom is involved in three N—H···O hydrogen bonds. Molecules are linked *via* pairs of N—H···O hydrogen bonds, forming inversion dimers with an R<sub>2</sub><sup>2</sup>(8) ring motif. The dimers are linked by further N—H···O hydrogen bonds, forming two-dimensional networks lying parallel to (100).

**Keywords:** crystal structure; urea; urea derivatives; hydrogen bonding.

**CCDC reference:** 1039538

## 1. Related literature

For general background to urea derivatives and their biological applications and properties, see: Ramalingan & Kwak (2008); Ramalingan *et al.* (2010); Yang *et al.* (2013); Safari & Gandomi-Ravandi (2014); Suzuki *et al.* (2013); Boulahjar *et al.* (2012); Zhang *et al.* (2014)



## 2. Experimental

### 2.1. Crystal data

C<sub>9</sub>H<sub>12</sub>N<sub>2</sub>O  
M<sub>r</sub> = 164.21

Monoclinic, P2<sub>1</sub>/c  
a = 14.631 (4) Å

b = 7.0633 (19) Å  
c = 8.786 (2) Å  
β = 93.530 (4)°  
V = 906.2 (4) Å<sup>3</sup>  
Z = 4

Mo Kα radiation  
μ = 0.08 mm<sup>-1</sup>  
T = 292 K  
0.20 × 0.18 × 0.16 mm

### 2.2. Data collection

Bruker SMART APEX CCD area-detector diffractometer  
8026 measured reflections

1556 independent reflections  
1284 reflections with I > 2σ(I)  
R<sub>int</sub> = 0.028

### 2.3. Refinement

R[F<sup>2</sup> > 2σ(F<sup>2</sup>)] = 0.100  
wR(F<sup>2</sup>) = 0.349  
S = 1.59  
1556 reflections  
119 parameters  
2 restraints

H atoms treated by a mixture of independent and constrained refinement  
Δρ<sub>max</sub> = 0.87 e Å<sup>-3</sup>  
Δρ<sub>min</sub> = -0.32 e Å<sup>-3</sup>

**Table 1**

Hydrogen-bond geometry (Å, °).

D—H···A	D—H	H···A	D···A	D—H···A
N1—H1···O1 <sup>i</sup>	0.86	2.23	2.941 (3)	140
N2—H2A···O1 <sup>i</sup>	0.86 (1)	2.24 (2)	2.985 (3)	145 (3)
N2—H2B···O1 <sup>ii</sup>	0.86 (1)	2.12 (1)	2.977 (3)	173 (4)

Symmetry codes: (i) x, -y + ½, z - ½; (ii) -x + 1, -y, -z + 1.

Data collection: SMART (Bruker, 2001); cell refinement: SAINT (Bruker, 2001); data reduction: SAINT; program(s) used to solve structure: SHELXS97 (Sheldrick, 2008); program(s) used to refine structure: SHELXL2013 (Sheldrick, 2008); molecular graphics: ORTEP-3 for Windows (Farrugia, 2012) and PLATON (Spek, 2009); software used to prepare material for publication: SHELXL2013 and PLATON (Spek, 2009).

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Supporting information for this paper is available from the IUCr electronic archives (Reference: SU5043).

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## supporting information

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## Crystal structure of 1-(2,4-dimethylphenyl)urea

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### S1. Comment

Urea and its derivatives are important key starting materials for the construction of biologically important heterocycles (Ramalingan & Kwak, 2008; Ramalingan *et al.*, 2010; Yang *et al.*, 2013; Safari & Gandomi-Ravandi, 2014). They display various biological activities viz. antibacterial (Suzuki *et al.*, 2013), antiproliferative and antitumor (Boulahjar *et al.*, 2012), and HIV-1 integrase (Zhang *et al.*, 2014). As a vital reactant and intermediate for the construction of heterocyclic chemical entities of biological importance, the title compound has been synthesized and single crystals were grown by slow evaporation in ethanol.

The single crystal X-ray analysis confirmed the molecular structure, as illustrated in Fig. 1. Methyl carbon atoms, C7 and C8, deviate by  $-0.000(1)$  and  $-0.040(1)$  Å, respectively, from the attached benzene ring. The dihedral angle between benzene ring and the mean plane through the urea atoms (N1/C9/O1/N2) is  $86.6(1)^\circ$ .

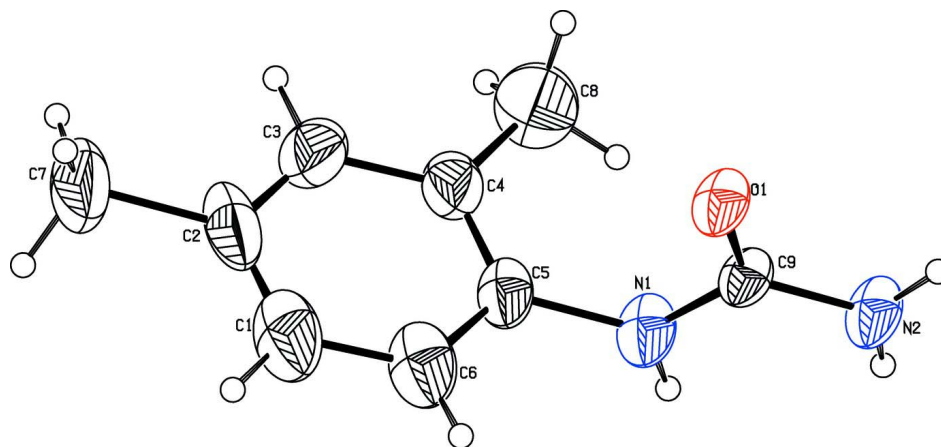
In the crystal, three strong N—H $\cdots$ O hydrogen bonds stabilize the molecular packing (Fig. 2 and Table 1). Molecules are linked via pairs of N—H $\cdots$ O hydrogen bonds forming inversion dimers with an  $R^2_2(8)$  ring motif. The dimers are linked by further N—H $\cdots$ O hydrogen bonds forming two-dimensional networks lying parallel to (100); see Table 1 and Fig. 2.

### S2. Experimental

To a solution of 2,6-dimethylaniline (0.1 mol) in glacial acetic acid (30 ml), was added distilled water (70 ml). Sodium cyanate (0.1 mol) in medium-hot water (50 ml) was then added in a slow manner with constant stirring. The resulted solution was allowed to stand for 60 min. and then cooled in ice. It was then filtered using a Buchner funnel and the solid obtained was dried using high-vacuum. Single crystals of the title compound were obtained by slow evaporation of a solution in ethanol at room temperature.

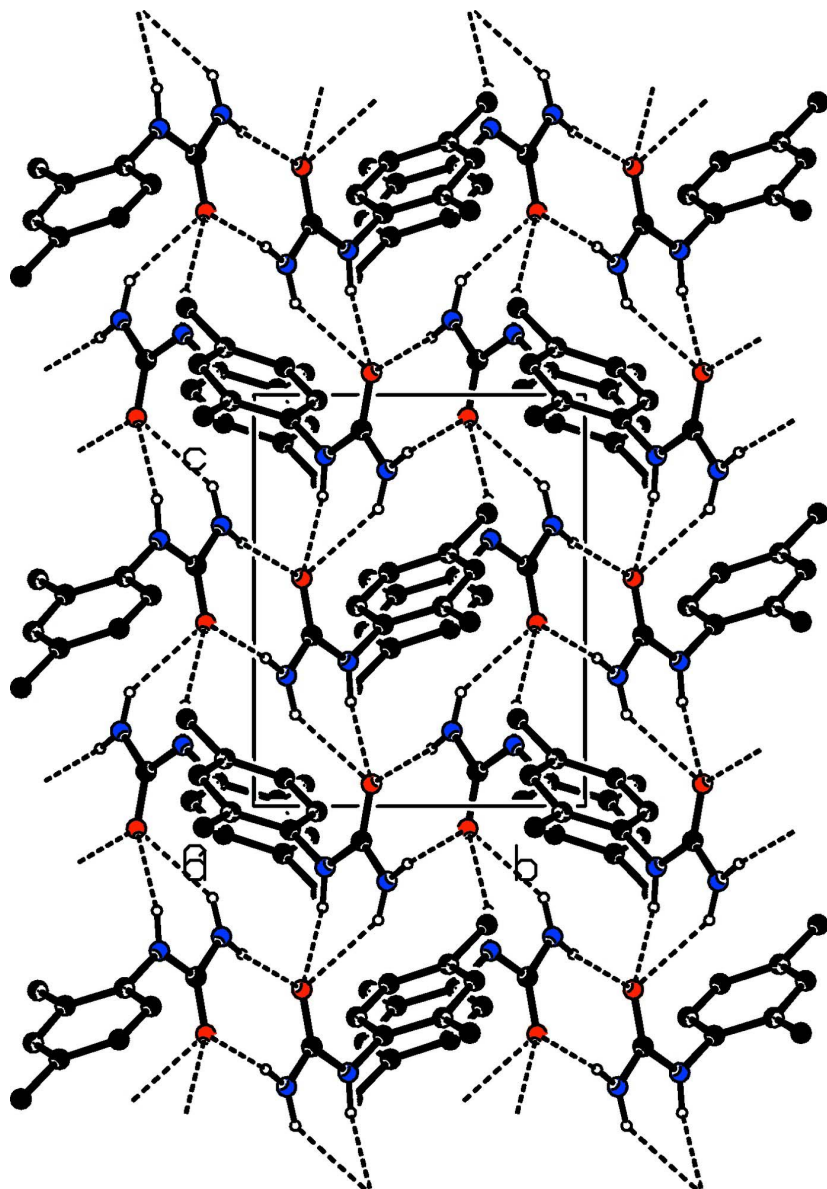
### S3. Refinement

Atoms H2A and H2B were located from a difference Fourier map and freely refined. The remaining H atoms were positioned geometrically and allowed to ride on their parent atoms, with N—H = 0.86 Å and C—H = 0.93–0.96 Å and with  $U_{\text{iso}}(\text{H}) = 1.5U_{\text{eq}}(\text{C})$  for methyl H atoms and =  $1.2U_{\text{eq}}(\text{N,C})$  for other H atoms.



**Figure 1**

The molecular structure of the title compound, with atom labelling. Displacement ellipsoids are drawn at the 30% probability level.



**Figure 2**

A projection of the crystal packing of the title compound, along the *a* axis. Hydrogen bonds are shown as dashed lines (see Table 1 for details; H atoms not involved in hydrogen bonding have been omitted for clarity).

### 1-(2,4-Dimethylphenyl)urea

#### Crystal data

$C_9H_{12}N_2O$

$M_r = 164.21$

Monoclinic,  $P2_1/c$

$a = 14.631(4) \text{ \AA}$

$b = 7.0633(19) \text{ \AA}$

$c = 8.786(2) \text{ \AA}$

$\beta = 93.530(4)^\circ$

$V = 906.2(4) \text{ \AA}^3$

$Z = 4$

$F(000) = 352$

$D_x = 1.204 \text{ Mg m}^{-3}$

Mo  $K\alpha$  radiation,  $\lambda = 0.71073 \text{ \AA}$

Cell parameters from 6568 reflections

$\theta = 2.8\text{--}24.6^\circ$

$\mu = 0.08 \text{ mm}^{-1}$

$T = 292 \text{ K}$

Block, colourless

$0.20 \times 0.18 \times 0.16 \text{ mm}$

*Data collection*

Bruker SMART APEX CCD area-detector  
diffractometer

Radiation source: fine-focus sealed tube

$\omega$  scans

8026 measured reflections

1556 independent reflections

1284 reflections with  $I > 2\sigma(I)$

$R_{\text{int}} = 0.028$

$\theta_{\text{max}} = 25.0^\circ$ ,  $\theta_{\text{min}} = 2.8^\circ$

$h = -17 \rightarrow 17$

$k = -8 \rightarrow 8$

$l = -10 \rightarrow 10$

*Refinement*

Refinement on  $F^2$

Least-squares matrix: full

$R[F^2 > 2\sigma(F^2)] = 0.100$

$wR(F^2) = 0.349$

$S = 1.59$

1556 reflections

119 parameters

2 restraints

Hydrogen site location: mixed

H atoms treated by a mixture of independent  
and constrained refinement

$w = 1/[\sigma^2(F_o^2) + (0.2P)^2]$

where  $P = (F_o^2 + 2F_c^2)/3$

$(\Delta/\sigma)_{\text{max}} = 0.001$

$\Delta\rho_{\text{max}} = 0.87 \text{ e } \text{\AA}^{-3}$

$\Delta\rho_{\text{min}} = -0.32 \text{ e } \text{\AA}^{-3}$

*Special details*

**Geometry.** All esds (except the esd in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell esds are taken into account individually in the estimation of esds in distances, angles and torsion angles; correlations between esds in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell esds is used for estimating esds involving l.s. planes.

*Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters ( $\text{\AA}^2$ )*

	<i>x</i>	<i>y</i>	<i>z</i>	$U_{\text{iso}}^*/U_{\text{eq}}$
O1	0.40043 (14)	0.1453 (3)	0.55423 (18)	0.0634 (8)
N1	0.32526 (19)	0.2851 (4)	0.3520 (3)	0.0709 (10)
H1	0.3186	0.2927	0.2543	0.085*
N2	0.4476 (2)	0.0982 (4)	0.3174 (3)	0.0682 (9)
H2A	0.437 (2)	0.125 (4)	0.2228 (15)	0.060 (8)*
H2B	0.4934 (19)	0.036 (5)	0.360 (4)	0.089 (11)*
C1	0.1253 (4)	0.4077 (12)	0.5679 (5)	0.1144 (17)
H1A	0.0700	0.3530	0.5912	0.137*
C2	0.1467 (3)	0.5877 (12)	0.6154 (4)	0.122 (2)
C3	0.2259 (4)	0.6656 (8)	0.5753 (5)	0.1049 (17)
H3	0.2404	0.7876	0.6085	0.126*
C4	0.2873 (2)	0.5700 (6)	0.4854 (4)	0.0794 (11)
C5	0.2652 (2)	0.3904 (5)	0.4421 (3)	0.0666 (10)
C6	0.1852 (3)	0.3106 (8)	0.4870 (4)	0.0921 (13)
H6	0.1721	0.1855	0.4606	0.111*
C7	0.0822 (4)	0.7033 (13)	0.7118 (7)	0.178 (4)
H7A	0.0806	0.6477	0.8114	0.267*
H7B	0.0216	0.7030	0.6630	0.267*
H7C	0.1040	0.8312	0.7212	0.267*
C8	0.3708 (4)	0.6517 (7)	0.4485 (7)	0.1135 (16)
H8A	0.3942	0.5853	0.3638	0.170*
H8B	0.4142	0.6433	0.5348	0.170*
H8C	0.3612	0.7822	0.4217	0.170*

C9	0.39175 (19)	0.1749 (4)	0.4154 (3)	0.0519 (8)
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*Atomic displacement parameters (Å<sup>2</sup>)*

	$U^{11}$	$U^{22}$	$U^{33}$	$U^{12}$	$U^{13}$	$U^{23}$
O1	0.0806 (14)	0.0787 (15)	0.0316 (11)	0.0198 (9)	0.0094 (9)	0.0014 (7)
N1	0.0824 (18)	0.0971 (19)	0.0338 (12)	0.0270 (14)	0.0098 (11)	0.0074 (11)
N2	0.0857 (18)	0.0849 (18)	0.0353 (14)	0.0251 (13)	0.0137 (11)	0.0013 (10)
C1	0.090 (3)	0.193 (5)	0.063 (2)	0.031 (3)	0.025 (2)	0.007 (3)
C2	0.068 (2)	0.245 (7)	0.053 (2)	0.063 (3)	0.0003 (17)	-0.015 (3)
C3	0.104 (3)	0.123 (3)	0.084 (3)	0.036 (3)	-0.019 (3)	-0.036 (2)
C4	0.0685 (19)	0.105 (3)	0.064 (2)	0.0184 (17)	-0.0010 (15)	-0.0088 (16)
C5	0.0682 (19)	0.090 (2)	0.0417 (16)	0.0245 (15)	0.0069 (13)	0.0057 (13)
C6	0.086 (2)	0.128 (3)	0.064 (2)	0.018 (2)	0.0218 (17)	0.0214 (19)
C7	0.106 (4)	0.316 (9)	0.112 (4)	0.091 (5)	-0.002 (3)	-0.091 (5)
C8	0.112 (3)	0.101 (3)	0.129 (4)	-0.017 (3)	0.024 (3)	0.004 (3)
C9	0.0652 (16)	0.0572 (15)	0.0342 (14)	0.0069 (11)	0.0094 (11)	-0.0013 (9)

*Geometric parameters (Å, °)*

O1—C9	1.236 (3)	C3—C4	1.406 (6)
N1—C9	1.340 (4)	C3—H3	0.9300
N1—C5	1.428 (4)	C4—C5	1.358 (6)
N1—H1	0.8600	C4—C8	1.407 (7)
N2—C9	1.337 (4)	C5—C6	1.377 (6)
N2—H2A	0.857 (10)	C6—H6	0.9300
N2—H2B	0.863 (10)	C7—H7A	0.9600
C1—C6	1.351 (7)	C7—H7B	0.9600
C1—C2	1.368 (10)	C7—H7C	0.9600
C1—H1A	0.9300	C8—H8A	0.9600
C2—C3	1.348 (9)	C8—H8B	0.9600
C2—C7	1.541 (6)	C8—H8C	0.9600
C9—N1—C5	121.9 (2)	C6—C5—N1	120.5 (4)
C9—N1—H1	119.1	C1—C6—C5	122.2 (6)
C5—N1—H1	119.1	C1—C6—H6	118.9
C9—N2—H2A	117 (2)	C5—C6—H6	118.9
C9—N2—H2B	115 (3)	C2—C7—H7A	109.5
H2A—N2—H2B	128 (4)	C2—C7—H7B	109.5
C6—C1—C2	119.2 (6)	H7A—C7—H7B	109.5
C6—C1—H1A	120.4	C2—C7—H7C	109.5
C2—C1—H1A	120.4	H7A—C7—H7C	109.5
C3—C2—C1	119.0 (4)	H7B—C7—H7C	109.5
C3—C2—C7	119.5 (7)	C4—C8—H8A	109.5
C1—C2—C7	121.5 (6)	C4—C8—H8B	109.5
C2—C3—C4	122.7 (5)	H8A—C8—H8B	109.5
C2—C3—H3	118.7	C4—C8—H8C	109.5
C4—C3—H3	118.7	H8A—C8—H8C	109.5

C5—C4—C3	117.1 (4)	H8B—C8—H8C	109.5
C5—C4—C8	121.0 (4)	O1—C9—N2	122.5 (2)
C3—C4—C8	121.8 (4)	O1—C9—N1	122.4 (2)
C4—C5—C6	119.7 (3)	N2—C9—N1	115.1 (2)
C4—C5—N1	119.8 (3)		
C6—C1—C2—C3	2.3 (7)	C8—C4—C5—N1	-2.8 (5)
C6—C1—C2—C7	-178.3 (4)	C9—N1—C5—C4	90.7 (4)
C1—C2—C3—C4	0.5 (7)	C9—N1—C5—C6	-88.6 (4)
C7—C2—C3—C4	-178.9 (4)	C2—C1—C6—C5	-4.1 (7)
C2—C3—C4—C5	-1.6 (6)	C4—C5—C6—C1	2.9 (6)
C2—C3—C4—C8	-178.1 (5)	N1—C5—C6—C1	-177.8 (3)
C3—C4—C5—C6	-0.1 (5)	C5—N1—C9—O1	6.1 (5)
C8—C4—C5—C6	176.5 (4)	C5—N1—C9—N2	-174.4 (3)
C3—C4—C5—N1	-179.4 (3)		

*Hydrogen-bond geometry (Å, °)*

<i>D</i> —H $\cdots$ <i>A</i>	<i>D</i> —H	H $\cdots$ <i>A</i>	<i>D</i> $\cdots$ <i>A</i>	<i>D</i> —H $\cdots$ <i>A</i>
N1—H1 $\cdots$ O1 <sup>i</sup>	0.86	2.23	2.941 (3)	140
N2—H2A $\cdots$ O1 <sup>i</sup>	0.86 (1)	2.24 (2)	2.985 (3)	145 (3)
N2—H2B $\cdots$ O1 <sup>ii</sup>	0.86 (1)	2.12 (1)	2.977 (3)	173 (4)

Symmetry codes: (i)  $x, -y+1/2, z-1/2$ ; (ii)  $-x+1, -y, -z+1$ .