

CrossMark  
click for updatesCite this: *Chem. Sci.*, 2015, 6, 4784Received 25th April 2015  
Accepted 22nd May 2015

DOI: 10.1039/c5sc01515d

www.rsc.org/chemicalscience

# Double-walled pyr topology networks from a novel fluoride-bridged heptanuclear metal cluster†

Kai-Jie Chen, John J. Perry IV, Hayley S. Scott, Qing-Yuan Yang  
and Michael J. Zaworotko\*

Two isostructural metal–organic materials, **Tripp-1-M** (**Tripp** = 2,4,6-tris(4-pyridyl)pyridine; M = Co, Ni), that exhibit binodal 3,6-connected **pyr** network topology have been prepared and characterized. **Tripp-1-M** are based upon a novel  $M_7F_{12}^{2+}$  cluster that possesses 12 connection points but, because of double cross-linking by 3-connected **Tripp** ligands, it functions as a 6-connected supermolecular building block (SBB).

## Introduction

Metal–organic materials (MOMs)<sup>1</sup> have attracted rapidly increasing attention from the scientific community in the last two decades thanks to their inherent modularity. This modularity can promote diversity of composition, amenability to systematic design<sup>2</sup> and control over certain bulk properties.<sup>3</sup> However, not all MOMs are well suited to serve as prototypal platforms for the generation of families of materials with the same topology. Such platforms are important because they enable systematic fine-tuning of both pore size (*e.g.* organic linkers with different lengths) and pore chemistry (*e.g.* functional group substitution at the linker or metal substitution at the node). Most platforms are built from single metal ion or small cluster (molecular building block, MBB) nodes and are exemplified by platforms sustained by carboxylate clusters such as 4-connected (4-c)  $M_2(RCOO)_4$  (*e.g.* **HKUST-1**,<sup>4</sup> **PCN-6**,<sup>5</sup> **MOF-2**,<sup>6</sup> **DMOF-1**<sup>7</sup> and **NU-100**<sup>8</sup>), 6-c  $M_3O(RCOO)_6$  (*e.g.* **MIL-88**,<sup>9</sup> **MIL-101**<sup>10</sup> and **PCN-600**<sup>11</sup>) and 6-c  $M_4O(RCOO)_6$  (*e.g.* **MOF-5**<sup>12</sup> and **MOF-177**<sup>13</sup>). The exploitation of larger, high symmetry clusters offers the possibility of much higher levels of connectivity and even greater control over topology. Such “supermolecular building blocks”, SBBs, are exemplified by 12-connected (12-c)  $Zr_6O_4(OH)_4(RCOO)_{12}$ <sup>14</sup> and 24-c “nanoball”  $Cu_{24}(1,3-bdc)_{24}$  clusters.<sup>15</sup> High connectivity mixed carboxylate/N-donor clusters have also been utilised in this context.<sup>16</sup> In addition to affording greater control over topology because of fewer possible topological outcomes, higher connectivity nodes can result in greater robustness.<sup>17</sup>

In this contribution, we introduce a new inorganic SBB of formula  $M_7F_{12}^{2+}$  (M = Co, Ni, Fig. 1) and demonstrate that it can serve as a 6-c SBB through double cross-linking of its 12 connection points by a facile to prepare 3-c ligand, 2,4,6-tris(4-pyridyl)pyridine, **Tripp** (Scheme 1). A new type of double-walled **pyr** topology (Fig. 2) material which exhibits permanent porosity is thereby generated.

In previous work, double<sup>18,19</sup> or quadruple<sup>20</sup> cross-linking of carboxylate<sup>21</sup> or oxide<sup>22</sup> based SBBs has been shown to represent a suitable approach to build MOMs with well-known<sup>23</sup> or hitherto novel<sup>24</sup> topology. However, the use of fluoride-bridged MBBs and SBBs as nodes for the construction of three-dimensional MOMs remains rare.<sup>25</sup> This is despite the fact that discrete fluoride-bridged metal clusters are known<sup>26,27</sup> and that such structures can exhibit interesting magnetic properties.<sup>28</sup>

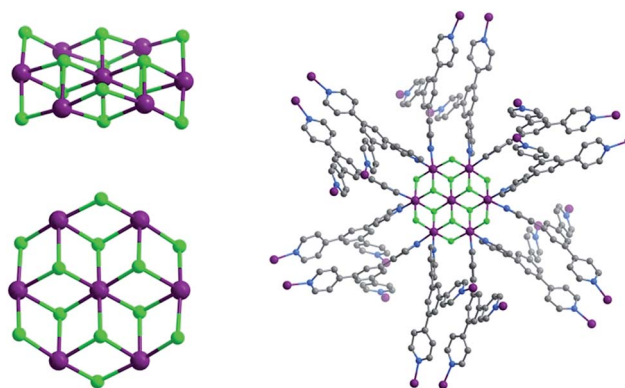
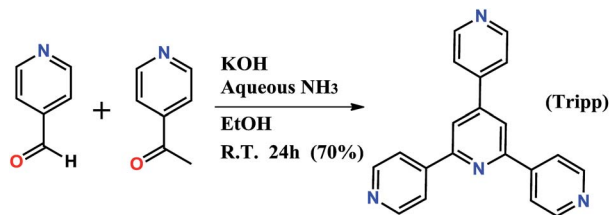
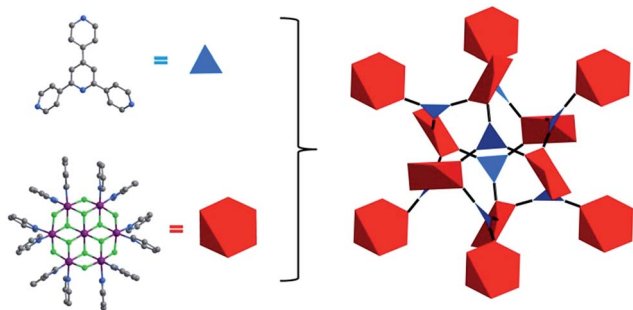


Fig. 1 (Left) Perspective and above views of the novel  $M_7F_{12}^{2+}$  cluster that sustains **Tripp-1-M**. (right) Illustration of the 12 connection points of the  $Co_7F_{12}$  SBB in **Tripp-1-Co** (Co, F, N and C atoms in purple, green, blue and grey). Solvent molecules, hydrogen atoms and counter-ions are omitted for the sake of clarity.

Department of Chemical & Environmental Sciences, University of Limerick, Limerick, Republic of Ireland. E-mail: Michael.Zaworotko@ul.ie

† Electronic supplementary information (ESI) available: Experimental details, single-crystal XRD data, PXRD patterns, TGA curves, sorption data fit and  $Q_{st}$  plot. CCDC 1060661. For ESI and crystallographic data in CIF or other electronic format see DOI: 10.1039/c5sc01515d

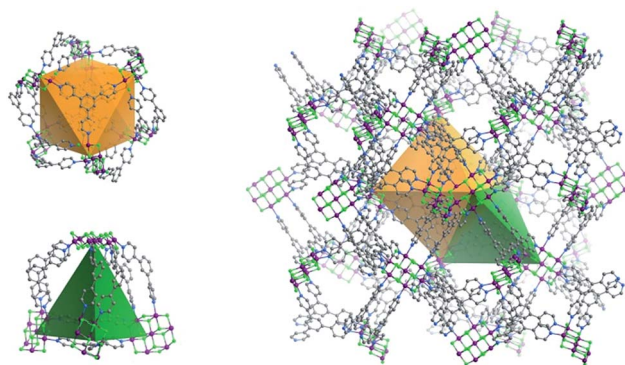


Scheme 1 Synthesis of 2,4,6-tris(4-pyridyl)pyridine (**Tripp**).Fig. 2 Illustration of the topology of **Tripp-1-M**. The **Tripp** ligand is represented by a 3-c triangle (blue) whereas the  $M_7F_{12}^{2+}$  SBB is reduced to a 6-c octahedron (red).

## Results and discussion

**Tripp** was prepared by the cyclization reaction of 4-acetylpyridine and 4-pyridinecarbaldehyde (Scheme 1).<sup>29</sup> Single crystals of **Tripp-1-Co** were initially obtained by solvothermal reaction between  $Co(NO_3)_2 \cdot 6H_2O$ , **Tripp** and  $(NH_4)_2SiF_6$  in DMF (for full details of synthetic procedures see ESI†). **Tripp-1-Co** crystallizes in the cubic space group  $Pa\bar{3}$ . A crystallographic 3-fold axis runs through the centre of the **Tripp** node and the disordered atoms of the central pyridine ring were therefore refined as 2/3 carbon and 1/3 nitrogen. All atoms of central pyridine ring are presented as carbon atoms for clarity in Fig. 1 and 3. In **Tripp-1-Co**, every **Tripp** ligand links three  $M_7F_{12}^{2+}$  SBBs and every SBB is connected by 12 **Tripp** ligands. However, the arrangement of the 12 connection points enables double cross-linking by pairs of **Tripp** ligands (Fig. 1) meaning that the connectivity is effectively reduced to 6. Therefore, each pair of **Tripp** ligands can be simplified to a single node, the  $M_7F_{12}$  SBB can be treated as a 6-c node and the **Tripp** ligand as a 3-connected node. The outcome of this connectivity is structure that exhibits binodal 3,6-connected **pyr** topology (Fig. 2), of which there are relatively few examples<sup>22b,30</sup> when compared to other types of 3,6 nets such as those with **rtl**, **ant**, **sit** or **qom** topology.

The structure of **Tripp-1-M** is comprised of two types of double-walled cages; tetrahedral cages (TCs) and octahedral cages (OCs) with face-sharing configurations (Fig. 3). The  $M_7F_{12}^{2+}$  SBB is located at the vertex of every TC and OC, while the double-walls are constructed from **Tripp** ligands. The pore diameter of the TC and OC are *ca.* 7.7 and 7.6 Å, respectively. Every OC is surrounded by eight adjacent TCs and every TC is surrounded by four OCs. However, there are two faces and one

Fig. 3 The two types of double-walled cages (OC, left above; TC left below) found in **Tripp-1-M** (right).

face capped by double walls of **Tripp** ligands in the OCs and TCs, respectively. Therefore, each OC connects with only six TCs and each TC crosslinks with three OCs, which corresponds to the required 3,6-connectivity of a **pyr** net. The guest-accessible porosity (considering the presence of counter ions) of **Tripp-1-Co** is 53.5%, based on Platon software.<sup>31</sup> Two other structures with **pyr** topology using similar tripyridyl-based 3-c ligands, 2,4,6-tri(4-pyridyl)-1,3,5-triazine (TPT) and 1,3,5-benzene tricarboxylic acid tris[*N*-(4-pyridyl)amide (TPBTC), with 6-c metal ions  $Hg^{2+}$  and  $Cd^{2+}$ , were reported by Robson and Kitagawa, respectively.<sup>32</sup> These tripyridyl-based ligands can also be utilized as 3-c organic nodes to construct nets with high porosity.<sup>33</sup> Further, some discrete cages consisting of 4-c  $Pd^{2+}$  ions linked by TPT and cage-based three-dimensional nets were reported by Fujita and co-workers.<sup>34</sup> In contrast, the facile to prepare **Tripp** ligand has not been as widely studied as TPT and TPBTC.

The double-walled nature of **Tripp-1-M** is unusual and, to our knowledge, such a structure has not been prepared using a single ligand and a single SBB. However, Bu *et al.* recently reported two isostructural double-walled cage-based MOMs that were designed using a strategy based upon size-matching between two tritopic ligands (TPT and 2,4,6-tris[1-(3-carboxylphenoxy)-yl-methyl]mesitylene) connected by with same paddle-wheel unit.<sup>18</sup>

Analysis of crystal structure of **Tripp-1-Co** revealed that the Co–F distances in heptanuclear cluster lie the range from 2.036 (5) to 2.121 (7) Å, which are consistent with the values found in other fluoride-bridged Co(II) structures, *e.g.*  $[Co_5F_2(tetrazolate)_4(H_2O)_4]^{35}$  and  $[Co_{12}(RCOO)_6(PO_4)_4F_4(H_2O)_6](NO_3)_2$ .<sup>36</sup> The crystal structure of **Tripp-1-Co** also revealed that the charge of each heptanuclear cobalt cluster is balanced by one  $SiF_6^{2-}$  anion that exhibits three-fold  $F \cdots F$  (distance: 2.68 (2) Å) interactions with three bridging  $F^-$  anions (Fig. S1†). Energy-dispersive X-ray spectroscopy verified the presence of Si and F in crystals of **Tripp-1-Co** (Fig. 4). There are previous reports concerning the generation of  $F^-$  by decomposition of  $PF_6^-$  and  $BF_4^-$ ,<sup>26,37</sup> so we speculated that the source of the bridging fluoride anions in the  $Co_7F_{12}^{2+}$  SBBs was *in situ* decomposition of  $SiF_6^{2-}$ . Further, there are two drawbacks to the use of this synthetic method: the relatively low solubility of  $(NH_4)_2SiF_6$  in



DMF; the requirement to decompose  $\text{SiF}_6^{2-}$  anions before the SBB can form. These drawbacks mean that unreacted  $(\text{NH}_4)_2\text{SiF}_6$  is isolated in a physical mixture with **Tripp-1-Co** crystals, mitigating against phase purity and also resulting in low product yield (*ca.* 10%). Therefore, we tested a different synthetic approach involving reflux of starting materials in DMF/MeOH with  $\text{NH}_4\text{F}$  instead of  $(\text{NH}_4)_2\text{SiF}_6$  as the  $\text{F}^-$  source. This method facilitated an increase in yield to 80%. The composition of **Tripp-1-Co** was changed since  $\text{SiF}_6^{2-}$  counterions are no longer present. Rather, two  $\text{NO}_3^-$  anions from the  $\text{Co}(\text{NO}_3)_2$  starting material balance charge as indicated by the presence of two diagnostic peaks measured using FT-IR at around 1320 and 1400  $\text{cm}^{-1}$  (Fig. S4†).<sup>38</sup> The relatively high solubility of  $\text{NH}_4\text{F}$  in MeOH enabled subsequent isolation of pure reaction product. The purity of bulk product was established by powder X-ray diffraction (PXRD) patterns of as-synthesized samples, which are good matches to those calculated from the crystal structure of **Tripp-1-Co** (Fig. S5†). To further verify the composition of the  $\text{M}_7\text{F}_{12}^{2+}$  cluster, X-ray photoelectron spectroscopy (XPS) analysis of a sample prepared from  $\text{NH}_4\text{F}$  was conducted, and a molar ratio for F : Co of 1.44 (expected 1.71) was observed (Fig. 4). A series of control experiments conducted without using  $\text{NH}_4\text{F}$  as a source of  $\text{F}^-$  revealed that different concentrations of metal and ligand and different reaction temperatures failed to afford the desired **Tripp-1-Co** product. Nevertheless, these experiments confirm the essential role that  $\text{F}^-$  plays in construction of the  $\text{M}_7\text{F}_{12}^{2+}$  SBB and subsequently the overall MOM framework. The isostructural nickel analogue of **Tripp-1-Co**, **Tripp-1-Ni**, was obtained *via* the same modified synthetic protocol with a yield of *ca.* 70% as verified by PXRD (Fig. S6†).

To address the thermal stability of **Tripp-1-M**, thermo-gravimetric analysis (TGA) was conducted for both as-synthesized and MeOH exchanged samples. The resulting TGA plots reveal that solvent guest molecules in the as-synthesized samples can be fully exchanged with MeOH, which in turn can be removed below 110 °C. No further weight loss until after 300 °C was observed, which we presume is a consequence of framework decomposition (Fig. S8†). Framework integrity was also verified by PXRD experiments conducted after desolvation of MeOH-exchanged samples at 120 °C. Furthermore, samples exposed to the air under ambient conditions for two months, were observed to exhibit PXRD patterns conforming to those calculated from single-crystal data (Fig. S5 and S6†). These results demonstrate that both **Tripp-1-Co** and **Tripp-1-Ni** possess good thermal and air/moisture stability, which we attribute to some extent to the high connectivity of the  $\text{M}_7\text{F}_{12}^{2+}$  SBB.

The permanent porosities of **Tripp-1-Co** and **Tripp-1-Ni** were established by measuring  $\text{CO}_2$  sorption isotherms at 195 K (Fig. 5). The apparent BET surface area was calculated to be 822 and 1149  $\text{m}^2 \text{g}^{-1}$  for **Tripp-1-Co** and **Tripp-1-Ni**, respectively. Pore volumes of 0.358 and 0.516  $\text{cm}^3 \text{g}^{-1}$  for **Tripp-1-Co** and **Tripp-1-Ni** were calculated by assuming liquid filling of  $\text{CO}_2$  at saturated state, which are close to the value of 0.587  $\text{cm}^3 \text{g}^{-1}$  estimated from the crystal data for **Tripp-1-Co**. The relatively lower uptake of **Tripp-1-Co** might be attributed to partial collapse of the framework during activation, which does not appear to occur for **Tripp-1-Ni**.  $\text{CO}_2$  and  $\text{N}_2$  sorption isotherms

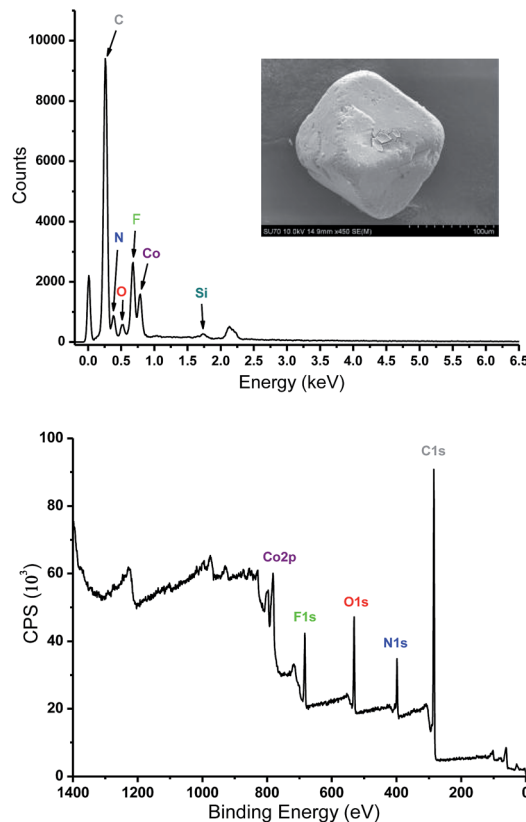


Fig. 4 Energy-dispersive X-ray spectroscopy (top) and X-ray photoelectron spectroscopy (bottom) of **Tripp-1-Co**.

of **Tripp-1-M** at 273, 283 and 293 K were also measured. As shown in Fig. 6,  $\text{CO}_2$  uptakes of 76.7, 61.4 and 49.2  $\text{cm}^3 \text{g}^{-1}$  at 273, 283 and 293 K, respectively, are much higher than the values of 3.7, 3.2 and 2.5  $\text{cm}^3 \text{g}^{-1}$  observed for  $\text{N}_2$  in **Tripp-1-Co**. **Tripp-1-Ni** exhibits higher  $\text{CO}_2$  uptakes of 99.3, 79.8 and 64.3  $\text{cm}^3 \text{g}^{-1}$  at 273, 283 and 293 K, respectively, than those observed for **Tripp-1-Co**. Meanwhile,  $\text{N}_2$  uptakes at 273, 283 and 293 K are only 5.2, 4.3 and 3.5  $\text{cm}^3 \text{g}^{-1}$  for **Tripp-1-Ni**. The high  $\text{CO}_2$  uptakes of **Tripp-1-M** could be attributed to high pore volume and polarized pore surface originating from the heptanuclear cluster and central pyridine ring of the **Tripp** ligand. These results indicate that both **Tripp-1-M** variants exhibit good selectivity for  $\text{CO}_2$  over  $\text{N}_2$ . Preliminary  $\text{CO}_2/\text{N}_2$  selectivity for **Tripp-1-Co** and **Tripp-1-Ni**, calculated from the uptakes of  $\text{CO}_2$  at 0.15 bar and  $\text{N}_2$  at 0.85 bar, are 41.4 and 40.8, 36.8 and 36.7, and 36.2 and 32.9 at 273, 283 and 293 K. These uptakes and selectivities are comparable to many well-known MOMs containing polar functional groups and/or open metal sites, both of which are absent in **Tripp-1-M** materials.<sup>39</sup>

To assess the strength of interaction between  $\text{CO}_2$  and framework, the  $\text{CO}_2$  isotherms measured at 273, 283 and 293 K were fitted using the virial equation (Fig. S9†), and the isosteric heat of adsorption ( $Q_{\text{st}}$ ) was calculated using the Clausius–Clapeyron equation. The enthalpies at zero loading for **Tripp-1-Co** and **Tripp-1-Ni** are 25.6 and 26.3  $\text{kJ mol}^{-1}$ , respectively (Fig. S10†). These values are also consistent with those observed



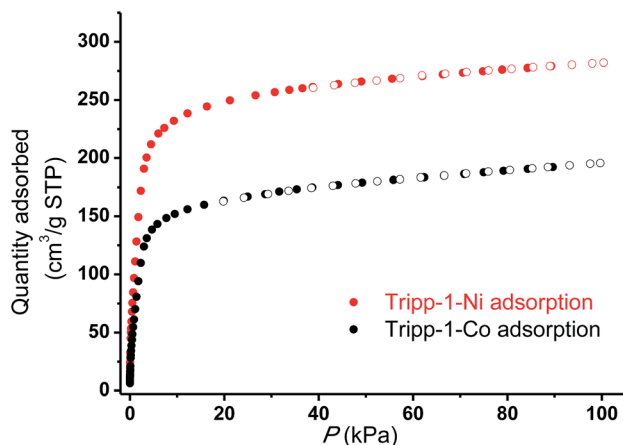


Fig. 5 CO<sub>2</sub> adsorption isotherms (filled symbols) and desorption (empty symbols) for Tripp-1-Co (black) and Tripp-1-Ni (red) conducted at 195 K.

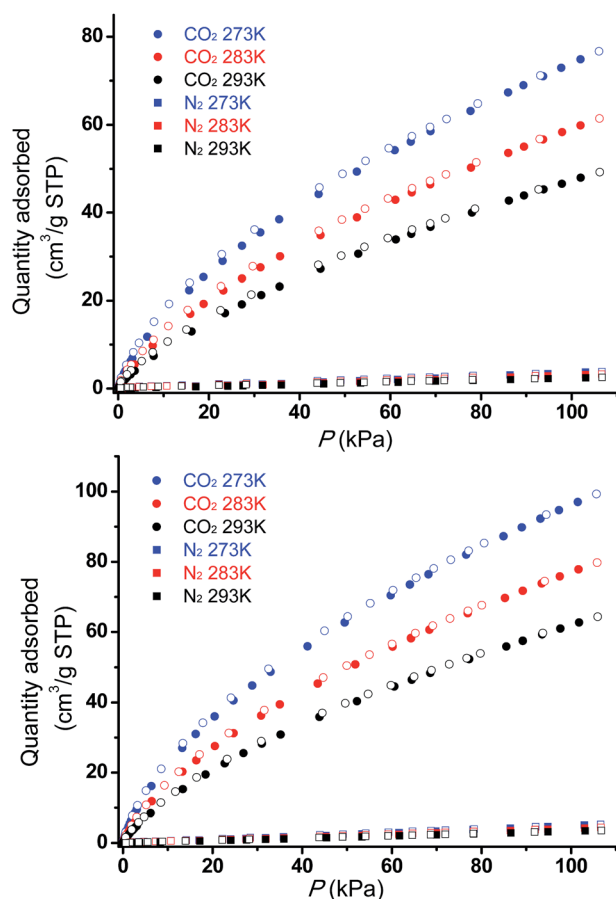


Fig. 6 CO<sub>2</sub> and N<sub>2</sub> adsorption (filled symbols) and desorption (empty symbols) isotherms for Tripp-1-Co (top) and Tripp-1-Ni (bottom) at three temperatures, 273 K, 283 K and 293 K.

in other classes of MOMs such as **MOF-5** (34 kJ mol<sup>-1</sup>),<sup>40</sup> **HKUST-1** (35 kJ mol<sup>-1</sup>),<sup>41</sup> **MAF-25** (26 kJ mol<sup>-1</sup>),<sup>42</sup> **InOF-1** (29 kJ mol<sup>-1</sup>),<sup>43</sup> **NOTT-140** (25 kJ mol<sup>-1</sup>).<sup>44</sup>

## Conclusions

In summary, we report a novel fluoride-bridged heptanuclear metal cluster-based SBB which has not been previously observed as a discrete structure. This cluster has 12-connection points, but 3-*c* **Tripp** ligands doubly cross-link to adjacent SBBs in order to form **Tripp-1-M**, two isostructural MOMs with binodal 3,6-connected **pyr** network topology. Good thermal and air/moisture stabilities were observed and gas sorption experiments demonstrate that both **Tripp-1-Co** and **Tripp-1-Ni** exhibit permanent porosity. The novel M<sub>7</sub>F<sub>12</sub><sup>2+</sup> SBB reported herein has the potential to serve as an SBB for a wider range of MOMs with tailored pore sizes and surface chemistries. Follow-on studies on this platform will address properties related to gas sorption, catalysis and magnetism and are currently underway.

## Acknowledgements

The authors thank the Science Foundation Ireland for funding of this research (SFI Award 13/RP/B2549). We also thank Dr Yina Guo and Dr Fathima Iaffir for help with Energy-dispersive X-ray spectroscopy and X-ray photoelectron spectroscopy measurements, respectively.

## Notes and references

- J. J. Perry IV, J. A. Perman and M. J. Zaworotko, *Chem. Soc. Rev.*, 2009, **38**, 1400–1417.
- (a) L. R. MacGillivray, *Metal-Organic Frameworks: Design and Application*, Wiley&Sons, 2010; (b) S. R. Batten, S. M. Neville and D. R. Turner, *Coordination Polymers: Design, Analysis and Application*, RSC publishing, 2009; (c) A. J. Blake, N. R. Champness, P. Hubberstey, W.-S. Li, M. A. Withersby and M. Schröder, *Coord. Chem. Rev.*, 1999, **183**, 117–138; (d) M. W. Hosseini, *Acc. Chem. Res.*, 2005, **38**, 313–323.
- S. Kitagawa, R. Kitaura and S.-I. Noro, *Angew. Chem., Int. Ed.*, 2004, **43**, 2334–2375.
- S. S.-Y. Chui, S. M.-F. Lo, J. P. H. Charmant, A. G. Orpen and I. D. Williams, *Science*, 1999, **283**, 1148–1150.
- S. Ma, D. Sun, M. Ambrogio, J. A. Fillinger, S. Parkin and H.-C. Zhou, *J. Am. Chem. Soc.*, 2007, **129**, 1858–1859.
- H. Li, M. Eddaoudi, T. L. Groy and O. M. Yaghi, *J. Am. Chem. Soc.*, 1998, **120**, 8571–8572.
- D. N. Dybtsev, H. Chun and K. Kim, *Angew. Chem., Int. Ed.*, 2004, **43**, 5033–5036.
- O. K. Farha, A. Özgür Yazaydın, I. Eryazici, C. D. Malliakas, B. G. Hauser, M. G. Kanatzidis, S. T. Nguyen, R. Q. Snurr and J. T. Hupp, *Nat. Chem.*, 2010, **2**, 944–948.
- C. Serre, C. Mellot-Draznieks, S. Surblé, N. Audebrand, Y. Filinchuk and G. Férey, *Science*, 2007, **315**, 1828–1831.
- G. Férey, C. Mellot-Draznieks, C. Serre, F. Millange, J. Dutour, S. Surblé and I. Margiolaki, *Science*, 2005, **309**, 2040–2042.
- K. Wang, D. Feng, T.-F. Liu, J. Su, S. Yuan, Y.-P. Chen, M. Bosch, X. Zou and H.-C. Zhou, *J. Am. Chem. Soc.*, 2014, **136**, 13983–13986.



- 12 H. Li, M. Eddaoudi, M. O'Keeffe and O. M. Yaghi, *Nature*, 1999, **402**, 276–279.
- 13 H. K. Chae, D. Y. Siberio-Perez, J. Kim, Y. Go, M. Eddaoudi, A. J. Matzger, M. O'Keeffe and O. M. Yaghi, *Nature*, 2004, **427**, 523–527.
- 14 J. H. Cavka, S. Jakobsen, U. Olsbye, N. Guillou, C. Lamberti, S. Bordiga and K. P. Lillerud, *J. Am. Chem. Soc.*, 2008, **130**, 13850–13851.
- 15 G. J. McManus, Z. Wang and M. J. Zaworotko, *Cryst. Growth Des.*, 2004, **4**, 11–13.
- 16 Y.-S. Wei, K.-J. Chen, P.-Q. Liao, B.-Y. Zhu, R.-B. Lin, H.-L. Zhou, B.-Y. Wang, W. Xue, J.-P. Zhang and X.-M. Chen, *Chem. Sci.*, 2013, **4**, 1539–1546.
- 17 F. Nouar, J. F. Eubank, T. Bousquet, L. Wojtas, M. J. Zaworotko and M. Eddaoudi, *J. Am. Chem. Soc.*, 2008, **130**, 1833–1835.
- 18 D. Tian, Q. Chen, Y. Li, Y.-H. Zhang, Z. Chang and X.-H. Bu, *Angew. Chem., Int. Ed.*, 2014, **53**, 837–841.
- 19 (a) Z.-Z. Lu, R. Zhang, Y.-Z. Li, Z.-J. Guo and H.-G. Zheng, *J. Am. Chem. Soc.*, 2011, **133**, 4172–4174; (b) M.-H. Zeng, Q.-X. Wang, Y.-X. Tan, S. Hu, H.-X. Zhao, L.-S. Long and M. Kurmoo, *J. Am. Chem. Soc.*, 2010, **132**, 2561–2563.
- 20 J. J. Perry IV, V. C. Kravtsov, G. J. McManus and M. J. Zaworotko, *J. Am. Chem. Soc.*, 2007, **129**, 10076–10077.
- 21 S. Y. Yang, L. S. Long, Y. B. Jiang, R. B. Huang and L. S. Zheng, *Chem. Mater.*, 2002, **14**, 3229–3231.
- 22 (a) A. Schoedel, L. Wojtas, S. P. Kelley, R. D. Rogers, M. Eddaoudi and M. J. Zaworotko, *Angew. Chem., Int. Ed.*, 2011, **50**, 11421–11424; (b) E. Y. Lee, S. Y. Jang and M. P. Suh, *J. Am. Chem. Soc.*, 2005, **127**, 6374–6381.
- 23 A. Schaate, P. Roy, T. Preuße, S. J. Lohmeier, A. Godt and P. Behrens, *Chem.–Eur. J.*, 2011, **17**, 9320–9325.
- 24 (a) M. Zhang, Y.-P. Chen, M. Bosch, T. Gentle, K. Wang, D. Feng, Z. U. Wang and H.-C. Zhou, *Angew. Chem., Int. Ed.*, 2014, **53**, 815–818; (b) W. Morris, B. Voloskiy, S. Demir, F. Gándara, P. L. McGrier, H. Furukawa, D. Cascio, J. F. Stoddart and O. M. Yaghi, *Inorg. Chem.*, 2012, **51**, 6443–6445; (c) J. E. Mondloch, W. Bury, D. Fairen-Jimenez, S. Kwon, E. J. DeMarco, M. H. Weston, A. A. Sarjeant, S. T. Nguyen, P. C. Stair, R. Q. Snurr, O. K. Farha and J. T. Hupp, *J. Am. Chem. Soc.*, 2013, **135**, 10294–10297.
- 25 K. M. Ok and D. O'Hare, *Dalton Trans.*, 2008, 5560–5562.
- 26 J. C. Jansen, H. Van Koningsveld and J. Reedijk, *Nature*, 1977, **269**, 318–319.
- 27 M. Romanelli, G. A. Kumar, T. J. Emge, R. E. Riman and J. G. Brennan, *Angew. Chem., Int. Ed.*, 2008, **47**, 6049–6051.
- 28 (a) T. Birk, K. S. Pedersen, C. A. Thuesen, T. Weyhermüller, M. Schau-Magnussen, S. Piligkos, H. Weihe, S. Mossin, M. Evangelisti and J. Bendix, *Inorg. Chem.*, 2012, **51**, 5435–5443; (b) G. A. Timco, S. Carretta, F. Troiani, F. Tuna, R. J. Pritchard, C. A. Muryn, E. J. L. McInnes, A. Ghirri, A. Candini, P. Santini, G. Amoretti, M. Affronte and R. E. P. Winpenny, *Nat. Nanotechnol.*, 2009, **4**, 173–178.
- 29 C. B. Smith, C. L. Raston and A. N. Sobolev, *Green Chem.*, 2005, **7**, 650–654.
- 30 (a) X.-J. Li, F.-L. Jiang, M.-Y. Wu, L. Chen, J.-J. Qian, K. Zhou, D.-Q. Yuan and M.-C. Hong, *Inorg. Chem.*, 2014, **53**, 1032–1038; (b) J. Liefbrig, O. Jeannin and M. Fourmigué, *J. Am. Chem. Soc.*, 2013, **135**, 6200–6210; (c) J. J. Henkelis, S. A. Barnett, L. P. Harding and M. J. Hardie, *Inorg. Chem.*, 2012, **51**, 10657–10674; (d) H. K. Chae, J. Kim, O. D. Friedrichs, M. O'Keeffe and O. M. Yaghi, *Angew. Chem., Int. Ed.*, 2003, **42**, 3907–3909; (e) L. Xu, Y.-U. Kwon, B. de Castro and L. Cunha-Silva, *Cryst. Growth Des.*, 2013, **13**, 1260–1266; (f) M. J. Manos, M. S. Markoulides, C. D. Malliakas, G. S. Papaefstathiou, N. Chronakis, M. G. Kanatzidis, P. N. Trikalitis and A. J. Tasiopoulos, *Inorg. Chem.*, 2011, **50**, 11297–11299; (g) D. B. Cordes and L. R. Hanton, *Inorg. Chem.*, 2007, **46**, 1634–1644; (h) B. Gao, S.-X. Liu, C.-D. Zhang, L.-H. Xie, C.-Y. Sun and M. Yu, *J. Coord. Chem.*, 2007, **60**, 911–918; (i) M. Du, Z.-H. Zhang, L.-F. Tang, X.-G. Wang, X.-J. Zhao and S. R. Batten, *Chem.–Eur. J.*, 2007, **13**, 2578–2586; (j) E. Lefebvre, F. Conan, N. Cosquer, J.-M. Kerbaol, M. Marchivie, J. Sala-Pala, M. M. Kubicki, E. Vigier and C. J. Gomez Garcia, *New J. Chem.*, 2006, **30**, 1197–1206; (k) B. Gao, S.-X. Liu, L.-H. Xie, M. Yu, C.-D. Zhang, C.-Y. Sun and H.-Y. Cheng, *J. Solid State Chem.*, 2006, **179**, 1681–1689; (l) A. Mondal, G. Mostafa, A. Ghosh, I. Rahaman Laskar and N. Ray Chaudhuri, *J. Chem. Soc., Dalton Trans.*, 1999, 9–10.
- 31 A. L. Spek, *J. Appl. Crystallogr.*, 2003, **36**, 7–13.
- 32 (a) S. Hasegawa, S. Horike, R. Matsuda, S. Furukawa, K. Mochizuki, Y. Kinoshita and S. Kitagawa, *J. Am. Chem. Soc.*, 2007, **129**, 2607–2614; (b) S. R. Batten, B. F. Hoskins and R. Robson, *Angew. Chem., Int. Ed.*, 1995, **34**, 820–822.
- 33 (a) X. Zhao, X. Bu, Q.-G. Zhai, H. Tran and P. Feng, *J. Am. Chem. Soc.*, 2015, **137**, 1396–1399; (b) F.-Y. Yi, J. Zhang, H.-X. Zhang and Z.-M. Sun, *Chem. Commun.*, 2012, **48**, 10419–10421.
- 34 (a) T. Osuga, T. Murase and M. Fujita, *Angew. Chem., Int. Ed.*, 2012, **51**, 12199–12201; (b) Y. Inokuma, S. Yoshioka, J. Ariyoshi, T. Arai, Y. Hitora, K. Takada, S. Matsunaga, K. Rissanen and M. Fujita, *Nature*, 2013, **495**, 461–466; (c) S. Horiuchi, T. Murase and M. Fujita, *Chem.–Asian J.*, 2011, **6**, 1839–1847.
- 35 W. Ouellette, K. Darling, A. Prosvirin, K. Whitenack, K. R. Dunbar and J. Zubietta, *Dalton Trans.*, 2011, 12288–12300.
- 36 S. A. Kumalah Robinson, M.-V. L. Mempo, A. J. Cairns and K. T. Holman, *J. Am. Chem. Soc.*, 2011, **133**, 1634–1637.
- 37 S. C. Lee and R. H. Holm, *Inorg. Chem.*, 1993, **32**, 4745–4753.
- 38 (a) K. B. Yatsimirskii, *Pure Appl. Chem.*, 1977, **49**, 115; (b) Y.-Q. Chen, G.-R. Li, Z. Chang, Y.-K. Qu, Y.-H. Zhang and X.-H. Bu, *Chem. Sci.*, 2013, **4**, 3678–3682.
- 39 (a) M. H. Mohamed, S. K. Elsaidi, L. Wojtas, T. Pham, K. A. Forrest, B. Tudor, B. Space and M. J. Zaworotko, *J. Am. Chem. Soc.*, 2012, **134**, 19556–19559; (b) P. S. Nugent, V. L. Rhodus, T. Pham, K. Forrest, L. Wojtas, B. Space and M. J. Zaworotko, *J. Am. Chem. Soc.*, 2013, **135**, 10950–10953; (c) B. Arstad, H. Fjellvåg, K. Kongshaug, O. Swang and R. Blom, *Adsorption*, 2008, **14**, 755–762; (d) J. Kim,



- S.-T. Yang, S. B. Choi, J. Sim, J. Kim and W.-S. Ahn, *J. Mater. Chem.*, 2011, **21**, 3070–3076; (e) R. Vaidhyanathan, S. S. Iremonger, K. W. Dawson and G. K. H. Shimizu, *Chem. Commun.*, 2009, 5230–5232; (f) K.-J. Chen, R.-B. Lin, P.-Q. Liao, C.-T. He, J.-B. Lin, W. Xue, Y.-B. Zhang, J.-P. Zhang and X.-M. Chen, *Cryst. Growth Des.*, 2013, **13**, 2118–2123.
- 40 Z. Zhao, Z. Li and Y. S. Lin, *Ind. Eng. Chem. Res.*, 2009, **48**, 10015–10020.
- 41 Q.-M. Wang, D. Shen, M. Bülow, M. Ling Lau, S. Deng, F. R. Fitch, N. O. Lemcoff and J. Semanscin, *Microporous Mesoporous Mater.*, 2002, **55**, 217–230.
- 42 J.-B. Lin, J.-P. Zhang and X.-M. Chen, *J. Am. Chem. Soc.*, 2010, **132**, 6654–6656.
- 43 J. Qian, F. Jiang, D. Yuan, M. Wu, S. Zhang, L. Zhang and M. Hong, *Chem. Commun.*, 2012, **48**, 9696–9698.
- 44 C. Tan, S. Yang, N. R. Champness, X. Lin, A. J. Blake, W. Lewis and M. Schröder, *Chem. Commun.*, 2011, **47**, 4487–4489.

