



(*E*)-7-[(4-Nitrophenyl)diazenyl]-3a-(*p*-tolyl)-2,3,3a,4-tetrahydro-1*H*-benzo[*d*]pyrrolo[1,2-*a*]-imidazol-1-one 0.58-dimethyl sulfoxide 0.42-acetonitrile solvate: crystal structure, Hirshfeld analysis and DFT estimation of the energy of intermolecular interactions

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Received 14 August 2017

Accepted 27 September 2017

Edited by A. V. Yatsenko, Moscow State University, Russia

Keywords: crystal structure; X-ray diffraction analysis; π - π interactions; DFT calculations; energy of π - π interactions; Hirshfeld analysis; azo dye.

CCDC reference: 843227

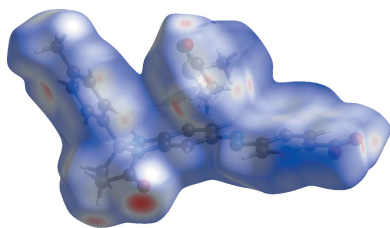
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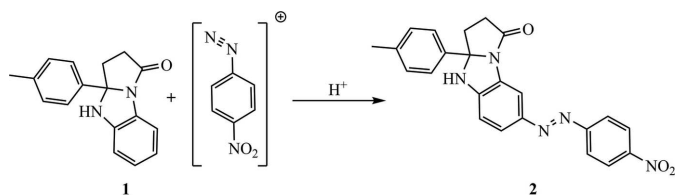
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In the crystal structure of the title compound, $C_{23}H_{19}N_5O_3 \cdot 0.58C_2H_6OS \cdot 0.42C_2H_3N$, prepared by the azo coupling of the 4-nitrophenyldiazonium salt with 3a-(*p*-tolyl)-2,3,3a,4-tetrahydro-1*H*-benzo[*d*]pyrrolo[1,2-*a*]imidazol-1-one, the azo molecules are linked by N—H···O hydrogen bonds into chains along the *a*-axis direction, and by the π - π interaction into [101] chains. The dimethyl sulfoxide and acetonitrile solvent molecules occupy the same positions, with populations of 0.585 (3) and 0.415 (3), respectively. These molecules take part in C—H···O(N) and C—H··· π contacts. The energy of the π - π interactions was estimated using DFT calculations. The Hirshfeld molecular surface analysis revealed the positions of the most important intermolecular contacts, such as hydrogen bonds and π - π interactions.

1. Chemical context

Compounds prepared by azo coupling of aryldiazonium salts with 3a-aryl-2,3,3a,4-tetrahydro-1*H*-benzo[*d*]pyrrolo[1,2-*a*]imidazol-1-one (**1**) are crystalline substances with deep color varying from yellow to red, depending on the structure of the initial diazonium cation. Since several nucleophilic centers in **1** can be attacked by the electrophilic diazonium cation, it was of interest to study the effect of heteroatoms, as well as other molecular fragments, on the molecular reactivity. The presence of the secondary amino group allows the formation of triazine derivatives. However, the most likely site of electrophilic attack is a fused aromatic ring activated by N heteroatoms. The azo dye molecules constructed in this way can exist in two forms, *E* and *Z*, depending on the presence or absence of certain stabilizing factors: bulky substituents, intramolecular hydrogen bonds, non-covalent interactions, etc. One of the representatives of the synthesized series is 7-[(4-nitrophenyl)diazenyl]-3a-(*p*-tolyl)-2,3,3a,4-tetrahydro-1*H*-benzo[*d*]pyrrolo[1,2-*a*]imidazol-1-one (**2**), which was prepared from 4-nitrophenyldiazonium chloride and **1**. For the final determination of the structure of the azo product, an X-ray diffraction study of a crystal grown from DMSO–acetonitrile solution as a mixed DMSO/acetonitrile solvate of **2** was performed.





2. Structural commentary

The asymmetric unit of the title compound is shown in Fig. 1. The molecules of **2** have the *E*-configuration that was expected because of the *para* position of the nitro group in the aryl-diazenyl fragment. Part of the molecule of **2**, including the 4-nitrophenyl and benzimidazole fragments linked by the azo group, is close to planar, with the dihedral angle formed by two aromatic rings being $2.73(7)^\circ$. The largest deviation from the mean plane of the benzimidazole ring system is $0.1300(9)$ Å for C4. The 1*H*-imidazole ring adopts an envelope conformation with C4 atom as the flap, thus introducing some non-planarity into the conjugated part of the molecule. The pyrrolidone ring is twisted with respect to the C2–C3 bond, thus the environment of the N2 amide atom becomes non-planar and this atom deviates by $0.267(1)$ Å from the plane formed by the three neighboring C atoms. As a result, the C1–N2 distance [$1.3737(17)$ Å] is larger than average for γ -lactams [$1.347(14)$ Å; Allen *et al.*, 1987]. The relatively long N2–C10 distance [$1.4091(17)$ Å] indicates weak π -conjugation and gives an insight into why substitution takes place at the 8 position.

3. Supramolecular features

In the crystal, molecules of **2** are linked by N–H \cdots O hydrogen bonds into chains along the *a*-axis direction (Table 1, Fig. 2). These molecules are also linked by π – π interactions between the aromatic rings of the benzimidazole fragments and 4-nitrophenyl substituents as well as between *p*-tolyl substituents (Table 2, Fig. 3), thus forming chains along the

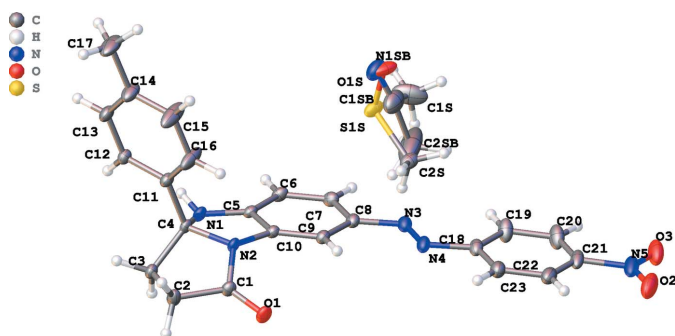


Figure 1

The asymmetric unit of the title compound with overlapping solvent molecules of DMSO and acetonitrile. Displacement ellipsoids are drawn at the 50% probability level.

Table 1

Hydrogen-bond geometry (Å, °).

Cg1 is the centroid of the C11–C16 ring.

<i>D</i> –H \cdots <i>A</i>	<i>D</i> –H	H \cdots <i>A</i>	<i>D</i> \cdots <i>A</i>	<i>D</i> –H \cdots <i>A</i>
N1–H1N \cdots O1 ⁱ	0.88 (2)	2.04 (2)	2.8550 (15)	154.9 (19)
C2–H2A \cdots N1S ⁱⁱ	0.99	2.52	3.43 (2)	153
C2–H2B \cdots O1S ⁱⁱⁱ	0.99	2.43	3.348 (14)	154
C2–H2B \cdots N1S ⁱⁱⁱ	0.99	2.43	3.37 (3)	158
C2S–H2SA \cdots Cg1 ⁱⁱ	0.96	2.93	3.766 (3)	146

Symmetry codes: (i) $x + 1, y, z$; (ii) $-x + 1, -y + 1, -z + 1$; (iii) $x - 1, y + 1, z$.

[101] direction. Comparing geometric parameters related to these π – π interactions (Table 2), one can conclude that those involving *p*-tolyl substituents are weaker. The dimethylsulfoxide and acetonitrile solvent molecules occupy the same positions with populations of 0.585 (3) and 0.415 (3), respectively. These molecules participate in intermolecular interactions as donors of H-atoms of the methyl groups of acetonitrile and DMSO, and as H-atom acceptors *via* the electronegative O and N atoms (Table 1).

4. Hirshfeld surface analysis

Hirshfeld surface analysis (Hirshfeld, 1977) of the title crystal structure allows us to visualize intermolecular interactions. The contribution of the H \cdots H intermolecular interactions amounts to 47.6%. The contributions of other important interactions are as follows: H \cdots O (21.2%), H \cdots C (11.2%) and H \cdots N (5.1%). Other contacts C \cdots O (3.9%), C \cdots C (3.8%), C \cdots N (3.6%), and H \cdots S (2.1%) are less than 5%.

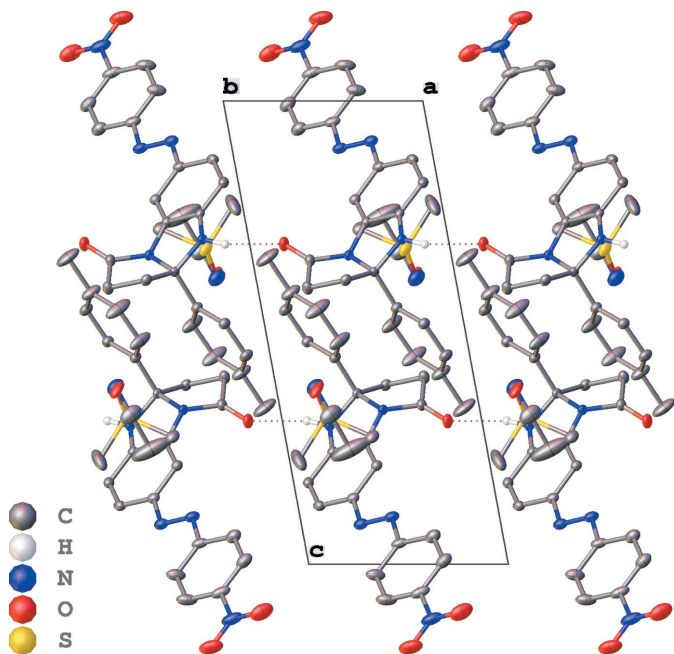


Figure 2

The packing diagram viewed along the *b* axis. N–H \cdots O hydrogen bonds are represented by dotted lines.

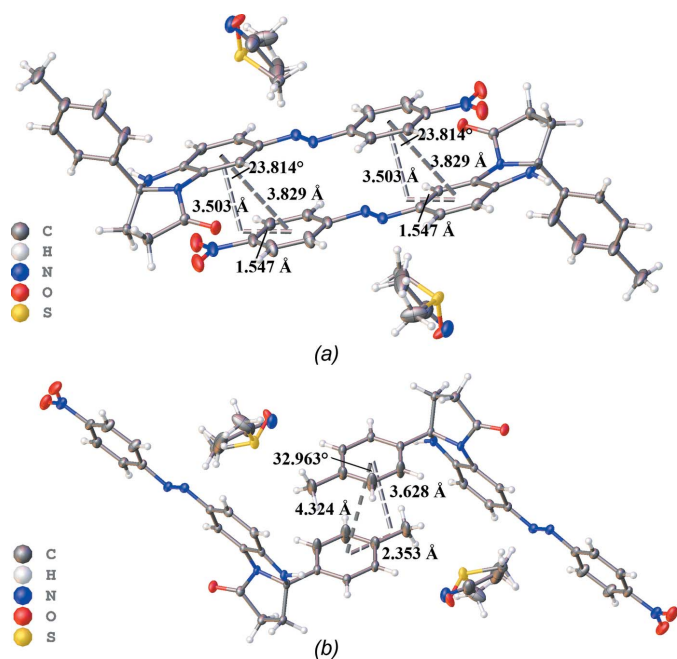


Figure 3
Diagram showing π - π interactions between molecules of **2** (a) between the aromatic rings of the benzimidazole group and the 4-nitrophenyl substituent, (b) between the aromatic rings of two *p*-tolyl substituents.

The Hirshfeld surface diagram, d_{norm} , with transparency (Fig. 4), indicates (in red) locations of the strongest intermolecular contacts with participation of atoms H6A, H2A and H2B (Fig. 4). The H \cdots H, H \cdots C, H \cdots S and H \cdots O contributions to the crystal packing are shown as two-dimensional fingerprint plots with blue dots (Fig. 5). The d_e (y axis) and d_i (x axis) values represent the closest external and internal distances (Å), respectively, from the given points on the Hirshfeld surface (Wolff *et al.*, 2012). The intermolecular hydrogen bonds are indicated by the H \cdots O contacts (21.2%) on the dotted diagram (Fig. 5c). Two sharp spikes with $d_e + d_i = \sim 2.0$ Å visualize the experimentally obtained value of 2.04 (2) Å for the H \cdots O distance corresponding to a hydrogen bond between azo molecules. The C \cdots C contacts (3.8%) reflect π - π interactions between the mentioned above

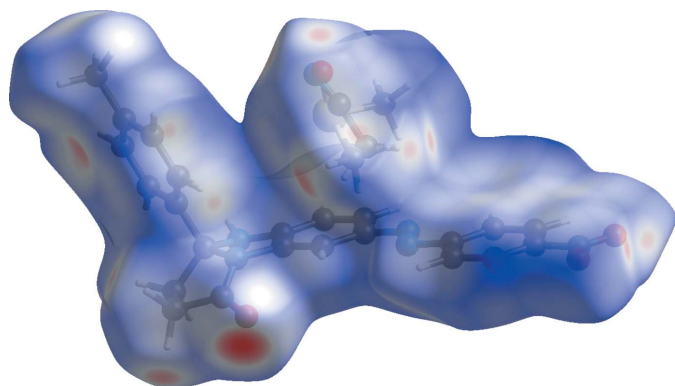


Figure 4
Hirshfeld surface diagram for the asymmetric unit of the title compound.

aromatic rings (Figs. 4, 5f). In addition, there are some H \cdots π contacts (H \cdots C), which are mostly located at hydrogen atoms of the CH₃ group of the *p*-tolyl substituent of one molecule and the π -system of the same substituent of the neighboring molecule (Fig. 5e).

5. Quantum chemical DFT calculations

To compare the energies of the two types of intermolecular π - π interactions found in the title crystal, we performed quantum chemical modeling of this system at the level of Density Functional Theory (DFT). All DFT calculations were made using GAUSSIAN09 package (Frisch *et al.*, 2010) and high-performance computing cluster of National Research Saratov State University. Crystallographic coordinates were used as a starting point, and full geometry optimization of monomer and dimers was performed using an mPW1B95 functional with a 6-31g(d) basis set. This hybrid meta density

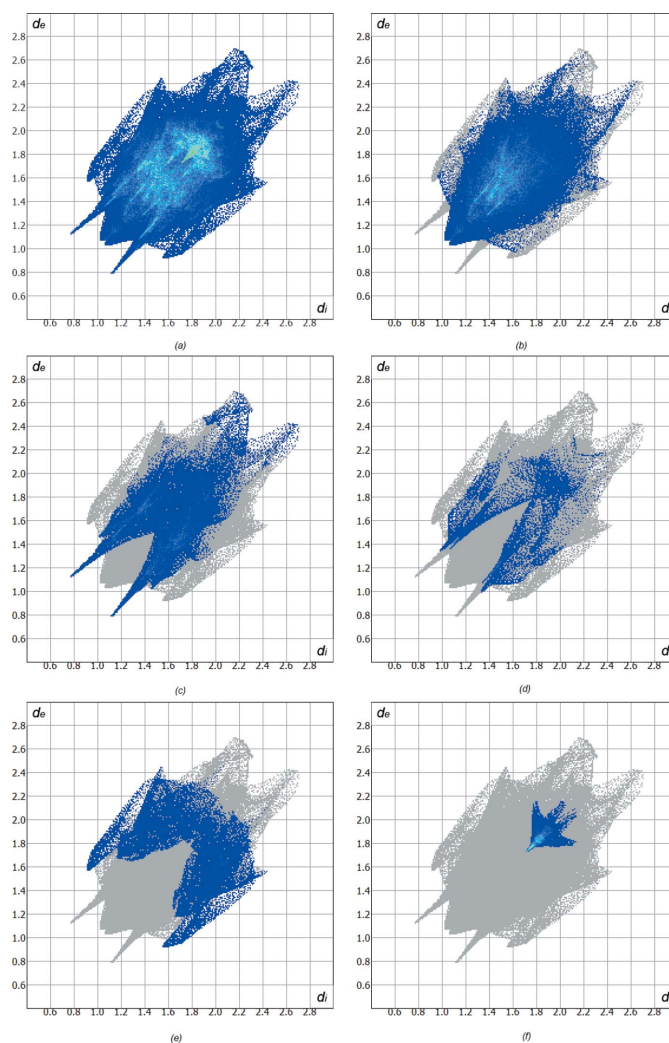


Figure 5
Diagrams showing (a) the full two-dimensional fingerprint plot, and those delineated into (b) H \cdots H, (c) O \cdots H/H \cdots O, (d) N \cdots H/H \cdots N, (e) C \cdots H/H \cdots C, and (f) C \cdots C contacts.

Table 2
Experimental and calculated parameters of π - π interactions in **2**.

Rings	Energy (kcal mol ⁻¹)	Intercentroid distance (Å)	Interplanar distance (Å)	Ring offset (Å)	Angle (°)
Benzimidazole/4-nitrophenyl	-16.48	exp 3.8290 (9)	calcd 3.876	exp 3.5025 (12)	calcd 3.485
<i>p</i> -Tolyl	-3.07	4.3241 (13)	4.807	3.628 (2)	3.740
				exp 1.547 (2)	calcd 1.698
				exp 2.353 (3)	calcd 3.018
				exp 23.814 (5)	calcd 25.977
				exp 32.963 (3)	calcd 38.920

functional theory (HMDFT) method based on the modified Perdew and Wang exchange functional (mPW) and Becke's 1995 correlation functional (B95) gives good results for the systems with non-covalent interactions, such as hydrogen bonding and weak van der Waals interactions (Zhao & Truhlar, 2004). The energy of the π - π interaction was estimated using the following simple equation:

$$E_{\text{interaction}} = E_{\text{dimer}} - 2 \times E_{\text{monomer}}$$

A comparison of some parameters of non-covalent interactions for the optimized geometry of **2** and for the crystallographic data is presented in Table 2. The chosen level of theory reproduces the geometrical parameters of the intermolecular interactions quite well. Thus, the energies of π - π interactions of both types, between the aromatic rings of the

benzimidazole fragment and of the 4-nitrophenyl substituent and between the two aromatic *p*-tolyl substituents at the 3*a* positions, can be estimated to be equal to -16.5 and -3.0 kcal mol⁻¹, respectively.

6. Database survey

Molecule **2** may be considered as being composed of two fragments, a heterocyclic core and the 4-nitrophenyldiazenyl substituent. The latter is relatively abundant and a search in the Cambridge Structural Database (CSD, Version 5.37, update May 2016; Groom *et al.*, 2016) returned eight hits [CSD refcodes: EMAWUL (Yazıcı *et al.*, 2011), KEMFUE (Centore *et al.*, 2006), LEZXAO and LEZXUK (Šimůnek *et al.*, 2007), PIDVAA (Kasyan *et al.*, 2007), ROMNIR (Lu *et al.*, 2009), TIVBOQ (Rodriguez *et al.*, 2008), YEDYIQ (You *et al.*, 2006)], but no heterocyclic compounds were found among them. The closest to the heterocyclic core of **2** is the previously reported 3*a*-phenyl-2,3,3*a*,4-tetrahydro-1*H*-pyrrolo[1,2-*a*]-benzimidazol-1-one (CSD refcode CIGPEN01; Grinev and Egorova, 2013). Other examples of compounds containing the same heterocyclic core are disubstituted at the 2 position: 2-(4-isobutylphenyl)-2,3*a*-dimethyl-2,3,3*a*,4-tetrahydro-1*H*-pyrrolo[1,2-*a*]-benzimidazol-1-one (CSD refcode AKURII; Patil *et al.*, 2010) and 5*a*-*p*-tolyl-5*a*,5*b*,6,7,8,9,9*a*,10-octahydro-5*H*-isoindolo(2,1-*a*)-benzimidazol-10-one – a substituted benzimidazolone ring fused with cyclohexane (CSD refcode ZENVUJ; Sillanpää *et al.*, 1995). From comparison of the reported structure with literature data, one can notice that the N1–C5 bond length in the title structure is shorter than in the related heterocycles CIGPEN01 and AKURII. This is related to the π -acceptor properties of the nitrophenyldiazenyl group.

7. Synthesis and crystallization

The synthesis of **2** was carried out according to the procedure, proposed by Gavkus *et al.*, 2012, starting from 4-nitroaniline and **1**. The product was isolated with 87% yield and recrystallized from acetonitrile as ruby-red prisms. A suitable single crystal was obtained by slow cooling of the saturated solution of **2** in DMSO–acetonitrile mixture at 1:1 ratio.

8. Refinement

Crystal data, details of data collection and structure refinement details are summarized in Table 3. All non-H atoms, involving solvent molecules, were refined anisotropically. The N–H hydrogen atom was located from a difference map and

Table 3
Experimental details.

Crystal data	
Chemical formula	C ₂₃ H ₁₉ N ₅ O ₃ ·0.58C ₂ H ₆ OS·0.42C ₂ H ₃ N
<i>M_r</i>	476.17
Crystal system, space group	Triclinic, <i>P</i> $\bar{1}$
Temperature (K)	100
<i>a</i> , <i>b</i> , <i>c</i> (Å)	7.1755 (5), 10.7013 (8), 16.2586 (11)
α , β , γ (°)	86.072 (3), 78.868 (2), 73.222 (3)
<i>V</i> (Å ³)	1172.71 (14)
<i>Z</i>	2
Radiation type	Mo <i>K</i> α
μ (mm ⁻¹)	0.14
Crystal size (mm)	0.27 × 0.22 × 0.21
Data collection	
Diffraction	Bruker APEXII CCD area detector
Absorption correction	Multi-scan (SADABS; Bruker, 2008)
<i>T_{min}</i> , <i>T_{max}</i>	0.963, 0.971
No. of measured, independent and observed [<i>I</i> > 2 σ (<i>I</i>)] reflections	15024, 6820, 5126
<i>R_{int}</i>	0.024
(<i>sin</i> θ / λ) _{max} (Å ⁻¹)	0.704
Refinement	
<i>R</i> [<i>F</i> ² > 2 σ (<i>F</i> ²)], <i>wR</i> (<i>F</i> ²), <i>S</i>	0.050, 0.134, 1.00
No. of reflections	6820
No. of parameters	350
No. of restraints	25
H-atom treatment	H atoms treated by a mixture of independent and constrained refinement
$\Delta\rho_{\text{max}}$, $\Delta\rho_{\text{min}}$ (e Å ⁻³)	0.46, -0.50

Computer programs: APEX2 (Bruker, 2011), SAINT (Bruker, 2009), SHELXS97 and SHELXL97 (Sheldrick, 2008), OLEX2 (Dolomanov *et al.* and 2009), publCIF (Westrip, 2010).

refined isotropically. The C–H hydrogen atoms were positioned geometrically and refined using a riding model.

Funding information

This work was supported by the Russian Science Foundation (grant No. 15-13-10007 to Alevtina Yegorova).

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supporting information

Acta Cryst. (2017). E73, 1590-1594 [https://doi.org/10.1107/S2056989017013937]

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Computing details

Data collection: *APEX2* (Bruker, 2011); cell refinement: *S SAINT* (Bruker, 2009); data reduction: *S SAINT* (Bruker, 2009); program(s) used to solve structure: *SHELXS97* (Sheldrick, 2008); program(s) used to refine structure: *SHELXL97* (Sheldrick, 2008); molecular graphics: *OLEX2* (Dolomanov *et al.*, 2009); software used to prepare material for publication: *publCIF* (Westrip, 2010).

(E)-7-[(4-Nitrophenyl)diazenyl]-3a-(p-tolyl)-2,3,3a,4-tetrahydro-1H-benzo[d]pyrrolo[1,2-a]imidazol-1-one 0.58-dimethyl sulfoxide 0.42-acetonitrile solvate

Crystal data

$C_{23}H_{19}N_5O_3 \cdot 0.58C_2H_6OS \cdot 0.42C_2H_3N$

$M_r = 476.17$

Triclinic, $P\bar{1}$

Hall symbol: -P 1

$a = 7.1755$ (5) Å

$b = 10.7013$ (8) Å

$c = 16.2586$ (11) Å

$\alpha = 86.072$ (3)°

$\beta = 78.868$ (2)°

$\gamma = 73.222$ (3)°

$V = 1172.71$ (14) Å³

$Z = 2$

$F(000) = 499$

$D_x = 1.349$ Mg m⁻³

Mo $K\alpha$ radiation, $\lambda = 0.71073$ Å

Cell parameters from 679 reflections

$\theta = 3\text{--}30^\circ$

$\mu = 0.14$ mm⁻¹

$T = 100$ K

Prism, red

$0.27 \times 0.22 \times 0.21$ mm

Data collection

Bruker APEXII CCD area detector
diffractometer

Radiation source: fine-focus sealed tube

Graphite monochromator

ω scans

Absorption correction: multi-scan
(SADABS; Bruker, 2008)

$T_{\min} = 0.963$, $T_{\max} = 0.971$

15024 measured reflections

6820 independent reflections

5126 reflections with $I > 2\sigma(I)$

$R_{\text{int}} = 0.024$

$\theta_{\max} = 30.0^\circ$, $\theta_{\min} = 2.0^\circ$

$h = -10 \rightarrow 10$

$k = -15 \rightarrow 14$

$l = -22 \rightarrow 22$

Refinement

Refinement on F^2

Least-squares matrix: full

$R[F^2 > 2\sigma(F^2)] = 0.050$

$wR(F^2) = 0.134$

$S = 1.00$

6820 reflections

350 parameters

25 restraints

Primary atom site location: structure-invariant
direct methods

Secondary atom site location: difference Fourier
map

Hydrogen site location: inferred from
neighbouring sites

H atoms treated by a mixture of independent
and constrained refinement

$w = 1/[\sigma^2(F_o^2) + (0.0632P)^2 + 0.525P]$

where $P = (F_o^2 + 2F_c^2)/3$

$(\Delta/\sigma)_{\max} < 0.001$

$\Delta\rho_{\max} = 0.46 \text{ e } \text{\AA}^{-3}$

$\Delta\rho_{\min} = -0.50 \text{ e } \text{\AA}^{-3}$

Special details

Geometry. All esds (except the esd in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell esds are taken into account individually in the estimation of esds in distances, angles and torsion angles; correlations between esds in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell esds is used for estimating esds involving l.s. planes.

Refinement. Refinement of F^2 against ALL reflections. The weighted R-factor wR and goodness of fit S are based on F^2 , conventional R-factors R are based on F , with F set to zero for negative F^2 . The threshold expression of $F^2 > 2\sigma(F^2)$ is used only for calculating R-factors(gt) etc. and is not relevant to the choice of reflections for refinement. R-factors based on F^2 are statistically about twice as large as those based on F , and R-factors based on ALL data will be even larger.

Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters (\AA^2)

	<i>x</i>	<i>y</i>	<i>z</i>	$U_{\text{iso}}^*/U_{\text{eq}}$	Occ. (<1)
O1	0.17180 (14)	0.76220 (11)	0.30872 (7)	0.0229 (2)	
O2	0.2794 (2)	0.02811 (13)	-0.10319 (9)	0.0436 (4)	
O3	0.5660 (2)	0.01862 (13)	-0.17906 (8)	0.0428 (3)	
N1	0.76692 (17)	0.79781 (11)	0.29851 (7)	0.0152 (2)	
H1N	0.876 (3)	0.810 (2)	0.3093 (13)	0.035 (5)*	
N2	0.48636 (16)	0.73034 (11)	0.33286 (7)	0.0141 (2)	
N3	0.70223 (19)	0.43649 (12)	0.08643 (7)	0.0200 (2)	
N4	0.54027 (19)	0.40565 (12)	0.10212 (7)	0.0203 (2)	
N5	0.4347 (3)	0.05826 (14)	-0.11836 (9)	0.0330 (3)	
C1	0.28738 (19)	0.79333 (14)	0.34383 (8)	0.0168 (3)	
C2	0.2492 (2)	0.90348 (15)	0.40379 (9)	0.0205 (3)	
H2A	0.2166	0.8749	0.4627	0.025*	
H2B	0.1403	0.9792	0.3914	0.025*	
C3	0.44855 (19)	0.93659 (13)	0.38670 (9)	0.0168 (3)	
H3A	0.4570	0.9966	0.3379	0.020*	
H3B	0.4686	0.9767	0.4362	0.020*	
C4	0.59959 (18)	0.80249 (13)	0.36856 (8)	0.0135 (2)	
C5	0.77798 (19)	0.70403 (13)	0.24332 (8)	0.0145 (2)	
C6	0.9259 (2)	0.65182 (14)	0.17582 (8)	0.0181 (3)	
H6A	1.0456	0.6763	0.1641	0.022*	
C7	0.8930 (2)	0.56275 (14)	0.12611 (8)	0.0189 (3)	
H7A	0.9922	0.5261	0.0797	0.023*	
C8	0.7174 (2)	0.52528 (13)	0.14252 (8)	0.0172 (3)	
C9	0.5690 (2)	0.57522 (13)	0.21261 (8)	0.0157 (3)	

H9A	0.4504	0.5494	0.2253	0.019*	
C10	0.60410 (18)	0.66223 (13)	0.26114 (8)	0.0138 (2)	
C11	0.6773 (2)	0.73681 (13)	0.44615 (8)	0.0166 (3)	
C12	0.8005 (2)	0.78988 (14)	0.48057 (8)	0.0169 (3)	
H12A	0.8305	0.8668	0.4564	0.020*	
C13	0.8798 (2)	0.73152 (15)	0.54974 (9)	0.0215 (3)	
H13A	0.9633	0.7692	0.5724	0.026*	
C14	0.8394 (3)	0.61915 (17)	0.58628 (11)	0.0332 (4)	
C15	0.7196 (4)	0.5662 (2)	0.55122 (14)	0.0515 (6)	
H15A	0.6917	0.4885	0.5749	0.062*	
C16	0.6385 (3)	0.62415 (19)	0.48177 (12)	0.0401 (5)	
H16A	0.5560	0.5858	0.4589	0.048*	
C17	0.9234 (4)	0.5562 (2)	0.66234 (13)	0.0463 (5)	
H17A	0.9225	0.4646	0.6666	0.069*	
H17B	0.8426	0.6022	0.7129	0.069*	
H17C	1.0596	0.5610	0.6568	0.069*	
C18	0.5271 (2)	0.31506 (14)	0.04503 (9)	0.0214 (3)	
C19	0.6755 (3)	0.26357 (16)	-0.02325 (9)	0.0289 (3)	
H19A	0.7968	0.2859	-0.0327	0.035*	
C20	0.6442 (3)	0.17971 (17)	-0.07702 (10)	0.0320 (4)	
H20A	0.7424	0.1453	-0.1243	0.038*	
C21	0.4679 (3)	0.14689 (15)	-0.06083 (10)	0.0275 (3)	
C22	0.3205 (3)	0.19436 (15)	0.00692 (10)	0.0273 (3)	
H22A	0.2010	0.1696	0.0168	0.033*	
C23	0.3516 (2)	0.27924 (15)	0.06024 (10)	0.0249 (3)	
H23A	0.2527	0.3131	0.1074	0.030*	
S1S	0.77680 (10)	0.20694 (10)	0.31863 (5)	0.0258 (3)	0.585 (3)
O1S	0.801 (2)	0.0884 (12)	0.3766 (7)	0.0334 (17)	0.585 (3)
C1S	0.9536 (5)	0.1596 (4)	0.2247 (2)	0.0525 (11)	0.585 (3)
H1SA	1.0805	0.1651	0.2321	0.079*	0.585 (3)
H1SB	0.9113	0.2167	0.1795	0.079*	0.585 (3)
H1SC	0.9634	0.0715	0.2118	0.079*	0.585 (3)
C2S	0.5644 (4)	0.2154 (3)	0.2752 (2)	0.0233 (6)	0.585 (3)
H2SA	0.4507	0.2204	0.3185	0.035*	0.585 (3)
H2SB	0.5891	0.1397	0.2419	0.035*	0.585 (3)
H2SC	0.5403	0.2923	0.2402	0.035*	0.585 (3)
N1SB	0.804 (5)	0.107 (2)	0.3821 (14)	0.036 (3)	0.415 (3)
C1SB	0.7529 (9)	0.1442 (7)	0.3226 (4)	0.0437 (11)	0.415 (3)
C2SB	0.685 (2)	0.2100 (8)	0.2476 (6)	0.101 (4)	0.415 (3)
H2S1	0.7907	0.2424	0.2135	0.151*	0.415 (3)
H2S2	0.5678	0.2835	0.2639	0.151*	0.415 (3)
H2S3	0.6526	0.1483	0.2149	0.151*	0.415 (3)

Atomic displacement parameters (\AA^2)

	U^{11}	U^{22}	U^{33}	U^{12}	U^{13}	U^{23}
O1	0.0135 (5)	0.0311 (6)	0.0268 (5)	-0.0084 (4)	-0.0049 (4)	-0.0074 (4)
O2	0.0700 (10)	0.0360 (7)	0.0395 (7)	-0.0281 (7)	-0.0255 (7)	0.0013 (6)

O3	0.0742 (10)	0.0321 (7)	0.0254 (6)	-0.0148 (7)	-0.0157 (6)	-0.0063 (5)
N1	0.0126 (5)	0.0217 (6)	0.0142 (5)	-0.0085 (4)	-0.0030 (4)	-0.0009 (4)
N2	0.0124 (5)	0.0180 (5)	0.0143 (5)	-0.0068 (4)	-0.0037 (4)	-0.0023 (4)
N3	0.0273 (6)	0.0175 (6)	0.0159 (5)	-0.0057 (5)	-0.0064 (5)	0.0002 (4)
N4	0.0296 (6)	0.0182 (6)	0.0152 (5)	-0.0078 (5)	-0.0074 (5)	-0.0006 (4)
N5	0.0608 (10)	0.0216 (7)	0.0228 (7)	-0.0140 (7)	-0.0196 (7)	0.0026 (5)
C1	0.0131 (6)	0.0214 (7)	0.0169 (6)	-0.0066 (5)	-0.0021 (5)	-0.0014 (5)
C2	0.0149 (6)	0.0245 (7)	0.0225 (7)	-0.0052 (5)	-0.0022 (5)	-0.0077 (6)
C3	0.0175 (6)	0.0181 (6)	0.0171 (6)	-0.0062 (5)	-0.0059 (5)	-0.0025 (5)
C4	0.0124 (5)	0.0178 (6)	0.0135 (6)	-0.0080 (5)	-0.0038 (4)	-0.0009 (5)
C5	0.0137 (6)	0.0177 (6)	0.0133 (6)	-0.0048 (5)	-0.0052 (4)	0.0019 (5)
C6	0.0145 (6)	0.0250 (7)	0.0156 (6)	-0.0073 (5)	-0.0031 (5)	0.0016 (5)
C7	0.0182 (6)	0.0228 (7)	0.0133 (6)	-0.0029 (5)	-0.0015 (5)	-0.0007 (5)
C8	0.0211 (6)	0.0170 (6)	0.0140 (6)	-0.0047 (5)	-0.0054 (5)	0.0003 (5)
C9	0.0168 (6)	0.0171 (6)	0.0146 (6)	-0.0060 (5)	-0.0046 (5)	0.0005 (5)
C10	0.0127 (5)	0.0163 (6)	0.0127 (6)	-0.0041 (5)	-0.0031 (4)	0.0007 (5)
C11	0.0176 (6)	0.0192 (6)	0.0153 (6)	-0.0069 (5)	-0.0060 (5)	0.0004 (5)
C12	0.0178 (6)	0.0194 (6)	0.0162 (6)	-0.0082 (5)	-0.0049 (5)	-0.0001 (5)
C13	0.0246 (7)	0.0247 (7)	0.0197 (7)	-0.0098 (6)	-0.0106 (5)	0.0001 (5)
C14	0.0520 (11)	0.0288 (8)	0.0300 (8)	-0.0189 (8)	-0.0262 (8)	0.0108 (7)
C15	0.0889 (17)	0.0467 (11)	0.0498 (12)	-0.0506 (12)	-0.0491 (12)	0.0314 (10)
C16	0.0626 (12)	0.0416 (10)	0.0396 (10)	-0.0394 (10)	-0.0355 (9)	0.0208 (8)
C17	0.0751 (15)	0.0387 (10)	0.0420 (11)	-0.0259 (10)	-0.0427 (11)	0.0187 (8)
C18	0.0346 (8)	0.0171 (6)	0.0144 (6)	-0.0080 (6)	-0.0088 (6)	0.0016 (5)
C19	0.0455 (10)	0.0283 (8)	0.0166 (7)	-0.0192 (7)	-0.0002 (6)	-0.0026 (6)
C20	0.0537 (11)	0.0285 (8)	0.0158 (7)	-0.0175 (8)	-0.0014 (7)	-0.0033 (6)
C21	0.0514 (10)	0.0172 (7)	0.0191 (7)	-0.0114 (7)	-0.0175 (7)	0.0026 (5)
C22	0.0363 (9)	0.0213 (7)	0.0295 (8)	-0.0092 (6)	-0.0172 (7)	0.0017 (6)
C23	0.0309 (8)	0.0211 (7)	0.0239 (7)	-0.0053 (6)	-0.0104 (6)	-0.0023 (6)
S1S	0.0239 (4)	0.0302 (5)	0.0298 (4)	-0.0135 (3)	-0.0133 (3)	0.0054 (3)
O1S	0.0293 (17)	0.039 (4)	0.038 (3)	-0.015 (2)	-0.021 (2)	0.020 (3)
C1S	0.0248 (15)	0.059 (2)	0.056 (2)	-0.0019 (15)	0.0106 (14)	0.0246 (19)
C2S	0.0258 (13)	0.0254 (14)	0.0225 (14)	-0.0121 (11)	-0.0089 (10)	0.0073 (10)
N1SB	0.040 (4)	0.040 (5)	0.035 (3)	-0.018 (3)	-0.017 (3)	0.001 (3)
C1SB	0.062 (3)	0.036 (3)	0.049 (3)	-0.027 (2)	-0.027 (2)	0.009 (2)
C2SB	0.212 (13)	0.060 (5)	0.073 (6)	-0.066 (7)	-0.095 (7)	0.028 (4)

Geometric parameters (Å, °)

O1—C1	1.2219 (16)	C13—H13A	0.9500
O2—N5	1.224 (2)	C14—C15	1.378 (2)
O3—N5	1.228 (2)	C14—C17	1.512 (2)
N1—C5	1.3670 (17)	C15—C16	1.396 (2)
N1—C4	1.4787 (17)	C15—H15A	0.9500
N1—H1N	0.88 (2)	C16—H16A	0.9500
N2—C1	1.3737 (17)	C17—H17A	0.9800
N2—C10	1.4091 (17)	C17—H17B	0.9800
N2—C4	1.4828 (16)	C17—H17C	0.9800

N3—N4	1.2728 (18)	C18—C23	1.393 (2)
N3—C8	1.4012 (18)	C18—C19	1.399 (2)
N4—C18	1.4214 (18)	C19—C20	1.386 (2)
N5—C21	1.4727 (19)	C19—H19A	0.9500
C1—C2	1.5109 (19)	C20—C21	1.381 (3)
C2—C3	1.5409 (19)	C20—H20A	0.9500
C2—H2A	0.9900	C21—C22	1.381 (2)
C2—H2B	0.9900	C22—C23	1.388 (2)
C3—C4	1.5354 (19)	C22—H22A	0.9500
C3—H3A	0.9900	C23—H23A	0.9500
C3—H3B	0.9900	S1S—O1S	1.517 (11)
C4—C11	1.5231 (18)	S1S—C2S	1.776 (3)
C5—C6	1.3912 (19)	S1S—C1S	1.785 (4)
C5—C10	1.4162 (18)	C1S—H1SA	0.9599
C6—C7	1.388 (2)	C1S—H1SB	0.9601
C6—H6A	0.9500	C1S—H1SC	0.9600
C7—C8	1.4021 (19)	C2S—H2SA	0.9598
C7—H7A	0.9500	C2S—H2SB	0.9600
C8—C9	1.4160 (19)	C2S—H2SC	0.9601
C9—C10	1.3659 (18)	N1SB—C1SB	1.108 (16)
C9—H9A	0.9500	C1SB—C2SB	1.460 (8)
C11—C16	1.379 (2)	C2SB—H2SB	1.1778
C11—C12	1.3936 (18)	C2SB—H2SC	1.1738
C12—C13	1.3874 (19)	C2SB—H2S1	0.9800
C12—H12A	0.9500	C2SB—H2S2	0.9800
C13—C14	1.386 (2)	C2SB—H2S3	0.9800
C5—N1—C4	109.49 (10)	C15—C14—C13	117.95 (14)
C5—N1—H1N	120.1 (14)	C15—C14—C17	120.77 (16)
C4—N1—H1N	118.6 (13)	C13—C14—C17	121.28 (15)
C1—N2—C10	126.55 (11)	C14—C15—C16	121.51 (16)
C1—N2—C4	112.88 (11)	C14—C15—H15A	119.2
C10—N2—C4	110.00 (10)	C16—C15—H15A	119.2
N4—N3—C8	114.52 (12)	C11—C16—C15	120.37 (15)
N3—N4—C18	113.95 (12)	C11—C16—H16A	119.8
O2—N5—O3	123.77 (14)	C15—C16—H16A	119.8
O2—N5—C21	118.60 (16)	C14—C17—H17A	109.5
O3—N5—C21	117.63 (16)	C14—C17—H17B	109.5
O1—C1—N2	123.73 (13)	H17A—C17—H17B	109.5
O1—C1—C2	129.41 (13)	C14—C17—H17C	109.5
N2—C1—C2	106.85 (11)	H17A—C17—H17C	109.5
C1—C2—C3	102.37 (11)	H17B—C17—H17C	109.5
C1—C2—H2A	111.3	C23—C18—C19	120.23 (14)
C3—C2—H2A	111.3	C23—C18—N4	115.52 (14)
C1—C2—H2B	111.3	C19—C18—N4	124.25 (14)
C3—C2—H2B	111.3	C20—C19—C18	119.51 (16)
H2A—C2—H2B	109.2	C20—C19—H19A	120.2
C4—C3—C2	102.76 (11)	C18—C19—H19A	120.2

C4—C3—H3A	111.2	C21—C20—C19	118.97 (16)
C2—C3—H3A	111.2	C21—C20—H20A	120.5
C4—C3—H3B	111.2	C19—C20—H20A	120.5
C2—C3—H3B	111.2	C22—C21—C20	122.72 (14)
H3A—C3—H3B	109.1	C22—C21—N5	118.46 (15)
N1—C4—N2	101.19 (10)	C20—C21—N5	118.82 (16)
N1—C4—C11	109.81 (10)	C21—C22—C23	118.16 (15)
N2—C4—C11	113.51 (11)	C21—C22—H22A	120.9
N1—C4—C3	116.20 (11)	C23—C22—H22A	120.9
N2—C4—C3	102.46 (10)	C22—C23—C18	120.39 (15)
C11—C4—C3	112.96 (11)	C22—C23—H23A	119.8
N1—C5—C6	129.84 (12)	C18—C23—H23A	119.8
N1—C5—C10	110.23 (11)	O1S—S1S—C2S	105.3 (5)
C6—C5—C10	119.90 (12)	O1S—S1S—C1S	106.9 (6)
C7—C6—C5	117.82 (12)	C2S—S1S—C1S	96.38 (17)
C7—C6—H6A	121.1	S1S—C1S—H1SA	109.8
C5—C6—H6A	121.1	S1S—C1S—H1SB	109.3
C6—C7—C8	121.87 (13)	H1SA—C1S—H1SB	109.5
C6—C7—H7A	119.1	S1S—C1S—H1SC	109.3
C8—C7—H7A	119.1	H1SA—C1S—H1SC	109.5
N3—C8—C7	115.87 (12)	H1SB—C1S—H1SC	109.5
N3—C8—C9	123.65 (12)	S1S—C2S—H2SA	111.1
C7—C8—C9	120.48 (12)	S1S—C2S—H2SB	109.6
C10—C9—C8	116.92 (12)	H2SA—C2S—H2SB	109.5
C10—C9—H9A	121.5	S1S—C2S—H2SC	107.8
C8—C9—H9A	121.5	H2SA—C2S—H2SC	109.5
C9—C10—N2	130.64 (12)	H2SB—C2S—H2SC	109.5
C9—C10—C5	122.94 (12)	N1SB—C1SB—C2SB	172.7 (15)
N2—C10—C5	106.39 (11)	C1SB—C2SB—H2SB	93.7
C16—C11—C12	118.43 (13)	C1SB—C2SB—H2SC	130.4
C16—C11—C4	123.11 (12)	H2SB—C2SB—H2SC	83.6
C12—C11—C4	118.40 (12)	C1SB—C2SB—H2S1	109.5
C13—C12—C11	120.64 (13)	C1SB—C2SB—H2S2	109.5
C13—C12—H12A	119.7	H2S1—C2SB—H2S2	109.5
C11—C12—H12A	119.7	C1SB—C2SB—H2S3	109.5
C14—C13—C12	121.10 (13)	H2S1—C2SB—H2S3	109.5
C14—C13—H13A	119.5	H2S2—C2SB—H2S3	109.5
C12—C13—H13A	119.5		
C8—N3—N4—C18	-179.71 (11)	N1—C5—C10—C9	175.62 (12)
C10—N2—C1—O1	27.6 (2)	C6—C5—C10—C9	-2.8 (2)
C4—N2—C1—O1	168.43 (13)	N1—C5—C10—N2	-2.58 (14)
C10—N2—C1—C2	-151.40 (13)	C6—C5—C10—N2	179.04 (11)
C4—N2—C1—C2	-10.58 (15)	N1—C4—C11—C16	-115.12 (17)
O1—C1—C2—C3	-150.17 (15)	N2—C4—C11—C16	-2.7 (2)
N2—C1—C2—C3	28.76 (14)	C3—C4—C11—C16	113.40 (17)
C1—C2—C3—C4	-35.31 (13)	N1—C4—C11—C12	61.98 (16)
C5—N1—C4—N2	-16.28 (13)	N2—C4—C11—C12	174.42 (12)

C5—N1—C4—C11	103.95 (12)	C3—C4—C11—C12	-69.49 (15)
C5—N1—C4—C3	-126.30 (12)	C16—C11—C12—C13	-0.9 (2)
C1—N2—C4—N1	-132.55 (11)	C4—C11—C12—C13	-178.10 (13)
C10—N2—C4—N1	14.76 (13)	C11—C12—C13—C14	0.2 (2)
C1—N2—C4—C11	109.88 (13)	C12—C13—C14—C15	0.7 (3)
C10—N2—C4—C11	-102.81 (12)	C12—C13—C14—C17	-179.21 (18)
C1—N2—C4—C3	-12.25 (14)	C13—C14—C15—C16	-0.9 (4)
C10—N2—C4—C3	135.06 (11)	C17—C14—C15—C16	179.0 (2)
C2—C3—C4—N1	138.38 (11)	C12—C11—C16—C15	0.7 (3)
C2—C3—C4—N2	29.10 (12)	C4—C11—C16—C15	177.80 (19)
C2—C3—C4—C11	-93.40 (12)	C14—C15—C16—C11	0.2 (4)
C4—N1—C5—C6	-169.39 (13)	N3—N4—C18—C23	-179.91 (12)
C4—N1—C5—C10	12.44 (14)	N3—N4—C18—C19	-0.5 (2)
N1—C5—C6—C7	-175.73 (13)	C23—C18—C19—C20	1.9 (2)
C10—C5—C6—C7	2.29 (19)	N4—C18—C19—C20	-177.45 (14)
C5—C6—C7—C8	-0.1 (2)	C18—C19—C20—C21	-1.3 (3)
N4—N3—C8—C7	-179.04 (12)	C19—C20—C21—C22	0.1 (3)
N4—N3—C8—C9	1.62 (19)	C19—C20—C21—N5	179.70 (15)
C6—C7—C8—N3	178.86 (12)	O2—N5—C21—C22	-1.3 (2)
C6—C7—C8—C9	-1.8 (2)	O3—N5—C21—C22	178.84 (14)
N3—C8—C9—C10	-179.31 (12)	O2—N5—C21—C20	179.07 (15)
C7—C8—C9—C10	1.4 (2)	O3—N5—C21—C20	-0.8 (2)
C8—C9—C10—N2	178.58 (13)	C20—C21—C22—C23	0.5 (2)
C8—C9—C10—C5	0.86 (19)	N5—C21—C22—C23	-179.14 (13)
C1—N2—C10—C9	-44.5 (2)	C21—C22—C23—C18	0.2 (2)
C4—N2—C10—C9	173.73 (13)	C19—C18—C23—C22	-1.3 (2)
C1—N2—C10—C5	133.46 (13)	N4—C18—C23—C22	178.05 (13)
C4—N2—C10—C5	-8.26 (14)		

Hydrogen-bond geometry (Å, °)

Cg1 is the centroid of the C11–C16 ring.

<i>D</i> —H... <i>A</i>	<i>D</i> —H	H... <i>A</i>	<i>D</i> ... <i>A</i>	<i>D</i> —H... <i>A</i>
N1—H1N...O1 ⁱ	0.88 (2)	2.04 (2)	2.8550 (15)	154.9 (19)
C2—H2A...N1S ^B ⁱⁱ	0.99	2.52	3.43 (2)	153
C2—H2B...O1S ⁱⁱⁱ	0.99	2.43	3.348 (14)	154
C2—H2B...N1S ^B ⁱⁱⁱ	0.99	2.43	3.37 (3)	158
C2S—H2SA...Cg1 ⁱⁱ	0.96	2.93	3.766 (3)	146

Symmetry codes: (i) $x+1, y, z$; (ii) $-x+1, -y+1, -z+1$; (iii) $x-1, y+1, z$.