

Poly[diaqua[μ_6 -4,4'-(1,4-phenylene)-bis(2,6-dimethylpyridine-3,5-dicarboxylato)]dilead(II)]

Yi Zhu, Ming-Xing Zhang, Shan-Shan Yang, Feng Xiao, Xiao-Ping Zhang, Yuan-Yuan Gao, Bing-Jie Li and Kun-Lin Huang*

College of Life Science, and College of Chemistry, Chongqing Normal University, Chongqing 400047, People's Republic of China

Correspondence e-mail: kunlin@jlu.edu.cn

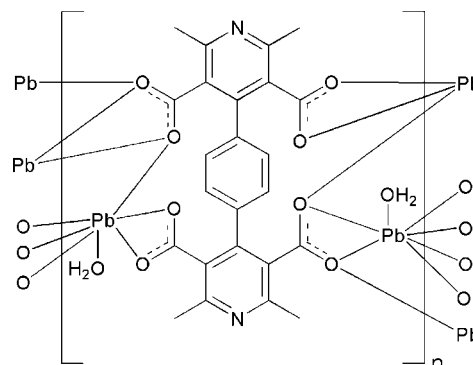
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Key indicators: single-crystal X-ray study; $T = 298$ K; mean $\sigma(\text{C}-\text{C}) = 0.007$ Å; R factor = 0.022; wR factor = 0.055; data-to-parameter ratio = 12.2.

The asymmetric unit of the title Pb-based coordination polymer, $[\text{Pb}_2(\text{C}_{24}\text{H}_{16}\text{N}_2\text{O}_8)(\text{H}_2\text{O})_2]_n$, consists of one Pb^{II} cation, half of a 4,4'-(1,4-phenylene)bis(2,6-dimethylpyridine-3,5-dicarboxylate (L^{4-}) ligand and one coordinating water molecule. The centers of the benzene ring of the ligand and the four-membered Pb/O/Pb/O ring are located on centers of inversion. The Pb^{II} ion is coordinated in form of a distorted polyhedron by seven O atoms from four separate L^{4-} ligands and by one water O atom. The PbO_7 polyhedra share O atoms, forming infinite zigzag $[\text{PbO}_4(\text{H}_2\text{O})]_n$ chains along [100] that are bridged by L^{4-} ligands, forming a two-dimensional coordination network parallel to (001). O—H...O hydrogen bonds involving the water molecule are observed.

Related literature

For background to metal-organic frameworks, see: Long & Yaghi (2009); Zhao *et al.* (2003). For related structures, see: Liu *et al.* (2002); O'Keeffe *et al.* (2008); Zhang *et al.* (2011). For lead complexes, see: Harrowfield *et al.* (2004); Yang *et al.* (2007). For typical Pb—O distances, see: Chen *et al.* (2012); Wei *et al.* (2005). For the photoluminescent mechanism of ligand–metal charge transfer, see: Hu *et al.* (2010); Zhang *et al.* (2012).



Experimental

Crystal data

$[\text{Pb}_2(\text{C}_{24}\text{H}_{16}\text{N}_2\text{O}_8)(\text{H}_2\text{O})_2]$
 $M_r = 910.80$
 Triclinic, $P\bar{1}$
 $a = 7.2182$ (12) Å
 $b = 9.0635$ (14) Å
 $c = 9.9589$ (15) Å
 $\alpha = 79.202$ (2)°
 $\beta = 71.683$ (2)°

$\gamma = 85.494$ (3)°
 $V = 607.43$ (17) Å³
 $Z = 1$
 Mo $K\alpha$ radiation
 $\mu = 13.90$ mm⁻¹
 $T = 298$ K
 0.25 × 0.23 × 0.23 mm

Data collection

Bruker SMART APEXII CCD diffractometer
 Absorption correction: multi-scan (SADABS; Sheldrick, 2003)
 $T_{\text{min}} = 0.129$, $T_{\text{max}} = 0.142$

3168 measured reflections
 2119 independent reflections
 1932 reflections with $I > 2\sigma(I)$
 $R_{\text{int}} = 0.015$

Refinement

$R[F^2 > 2\sigma(F^2)] = 0.022$
 $wR(F^2) = 0.055$
 $S = 1.01$
 2119 reflections

174 parameters
 H-atom parameters constrained
 $\Delta\rho_{\text{max}} = 1.15$ e Å⁻³
 $\Delta\rho_{\text{min}} = -1.16$ e Å⁻³

Table 1

Hydrogen-bond geometry (Å, °).

$D-H\cdots A$	$D-H$	$H\cdots A$	$D\cdots A$	$D-H\cdots A$
O5—H2...O4 ⁱ	0.85	2.04	2.834 (6)	155
O5—H1...O3 ⁱⁱ	0.85	2.05	2.879 (5)	165

Symmetry codes: (i) $-x, -y, -z + 1$; (ii) $-x + 1, -y, -z + 1$.

Data collection: APEX2 (Bruker, 2010); cell refinement: SAINT-Plus (Bruker, 2008); data reduction: SAINT-Plus; program(s) used to solve structure: SHELXS97 (Sheldrick, 2008); program(s) used to refine structure: SHELXL97 (Sheldrick, 2008); molecular graphics: SHELXTL (Sheldrick, 2008); software used to prepare material for publication: PLATON (Spek, 2009).

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Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: IM2422).

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supplementary materials

Acta Cryst. (2013). E69, m232–m233 [doi:10.1107/S1600536813007733]

Poly[*diaqua*[μ_6 -4,4'-(1,4-phenylene)bis(2,6-dimethylpyridine-3,5-dicarboxylato)]dilead(II)]

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Comment

In recent years, the chemistry of novel metal-organic hybrid coordination polymers has been the subject of intensive research, due to their interesting molecular structures and their potential as a new class of solid-state materials applied in catalysis, molecular recognition, gas storage, drug delivery, and so on (Liu *et al.*, 2002; O'Keeffe *et al.*, 2008). Generally speaking, the diversity of potential applications in the framework structures of such materials greatly depends on the selection of the metal centers and organic spacers. Recently, carboxylate groups are frequently exploited in the design, syntheses, and crystallization of coordination frameworks, because they exhibit diverse coordination modes, which can enhance the robustness of the architectures. Furthermore, the flexibility of carboxylate groups is always efficient to form fascinating structures. In this paper, we choose a new flexible and multidentate carboxylate ligand, 4,4'-(1,4-phenylene)bis(2,6-dimethylpyridine-3,5-dicarboxylic acid) (H_4L).

Up to date, research on coordination polymers has focused on transition metal ions as coordination centers, while less concentration has been given to heavy *p*-block metal ion, e.g. lead(II). In contrast to transition metal ions, lead(II), with its large radius, flexible coordination environment, and variable stereochemical activity, provides unique opportunities for the formation of unusual structures with interesting properties (Harrowfield *et al.*, 2004; Yang *et al.*, 2007). In addition, the intrinsic features of lead(II), the presence of a 6 s^2 outer electron configuration, inspire chemists extensive interest in coordination chemistry, photophysics, and photochemistry. Herein, we report a new photoluminescent complex $[Pb(L)(H_2O)]_n$ (**1**) from the flexible 4,4'-(1,4-phenylene)bis(2,6-dimethylpyridine-3,5-dicarboxylic acid) (H_4L) and lead salt.

X-ray diffraction analyses reveal that each asymmetric unit of **1** contains half deprotonated L^{4-} ligand, one H_2O molecule and one crystallographically independent Pb^{II} center (Fig. 1). $Pb1$ center is coordinated with seven O atoms: six ($O1^{#1}$, $O1^{#2}$, $O2^{#2}$, $O2^{#3}$, $O3$, $O4$) from four H_4L ligands and one ($O5$) from the H_2O molecule. Of particular interest is the weak coordinative bond that exists between $Pb1$ and $O2^{#3}$. $Pb1, O1^{#1}$, $O1^{#2}$, $O2^{#2}$, $O2^{#3}$, $O3$, $O4$, $O5$ furnish a polyhydral coordination environment (PbO_7) with the $Pb-O$ bond lengths are in agreement with those reported in other $Pb(II)$ complexes of O-chelating ligands (Wei *et al.*, 2005; Chen *et al.*, 2012).

As shown in Fig. 1, each H_4L ligand connects six crystallographically equivalent Pb atoms. The carboxylate group with $O1$ and $O2$ coordinates three lead atoms producing two Pb_2O_2 rings that share one common lead atom. The other carboxylate moiety with donor atoms $O3$ and $O4$ coordinates one lead atom in a chelating mode. Notably, the resulting PbO_7 polyhedra share the $O1^{#4}$, $O1^{#5}$, $O2^{#2}$, $O2^{#3}$ atoms to form infinite zigzag chains composed of $[PbO_4(H_2O)]_n$ in which adjacent Pb atoms are coplanar and $Pb\cdots Pb$ distances are 4.077 Å and 4.161 Å respectively (Fig. 2). Another interesting structural feature of complex **1** is that the zigzag $[PbO_4(H_2O)]_n$ chains are bridged by H_4L ligands to form a two-dimensional (2-D) coordination network (Fig. 3).

The photoluminescence spectrum of compound **1** was measured in the solid state at room temperature, as shown in Fig. 4. At room temperature the photoluminescent emission maximum of free H₄L was observed at 426 nm (upon $\lambda_{\text{Ex, max}} = 208$ nm). For compound **1**, excitation at 380 nm leads to strong photoluminescence with an emission maximum at $\lambda = 465$ nm. The emission peak of complex **1** is red-shifted by about 40 nm compared to that of the pure H₄L ligand, which can be assigned to the ligand-metal charge transfer (LMCT) (Hu *et al.*, 2010; Zhang *et al.*, 2011; Zhang *et al.*, 2012).

Experimental

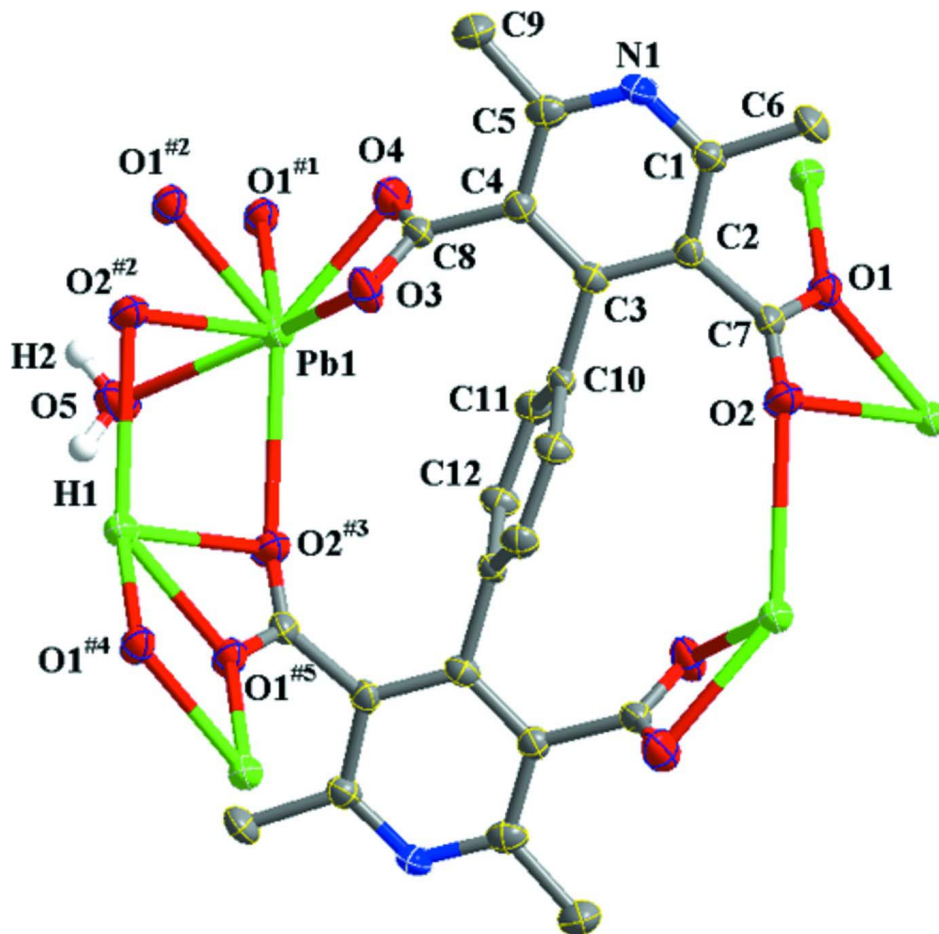
A mixture of Pb(NO₃)₂ × 6 H₂O (66 mg), H₄L (40 mg) and DMF (6 ml) was sealed in a 25 ml Teflon-lined stainless steel reactor. The mixture was heated to 373 K for 3 days and then cooled to room temperature. The crystal samples were washed with methanol to yield 18 mg of compound **1**.

Refinement

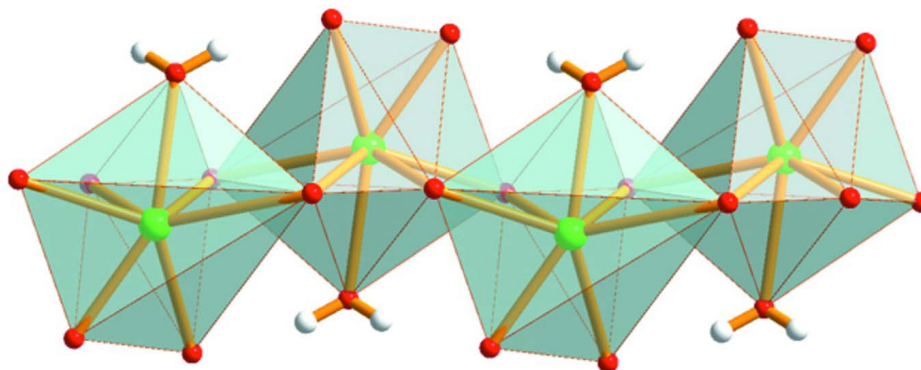
Methyl H atoms were constrained to an ideal geometry (C—H = 0.96 Å), with $U_{\text{iso}}(\text{H}) = 1.5U_{\text{eq}}(\text{C})$, but were allowed to rotate freely. Other H atoms attached to C atoms were refined using a riding model [C—H = 0.93 Å (CH) and $U_{\text{iso}}(\text{H}) = 1.2U_{\text{eq}}$ (parent atom)].

Computing details

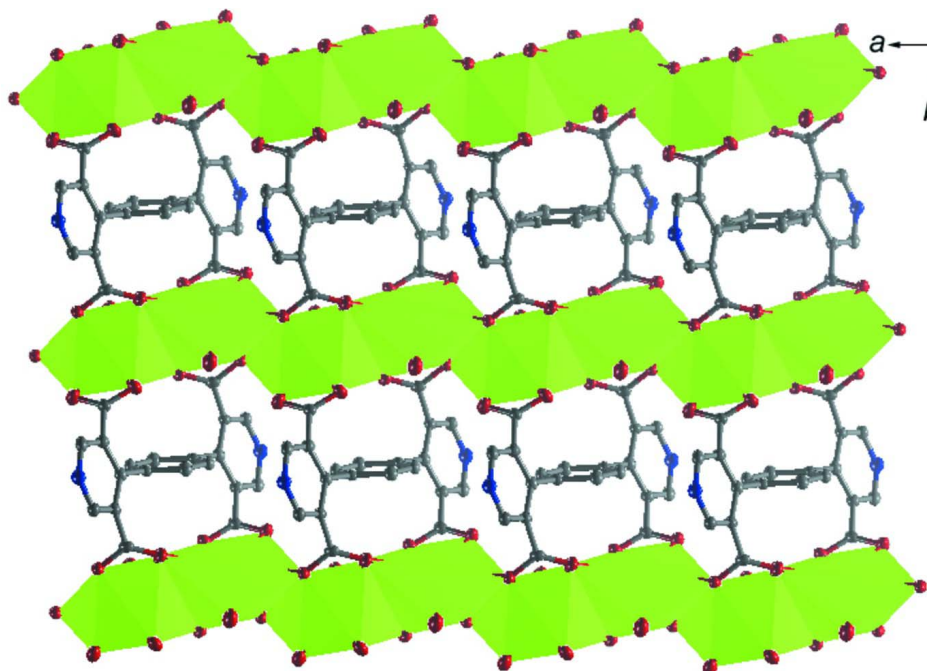
Data collection: *APEX2* (Bruker, 2010); cell refinement: *SAINTE-Plus* (Bruker, 2008); data reduction: *SAINTE-Plus* (Bruker, 2008); program(s) used to solve structure: *SHELXS97* (Sheldrick, 2008); program(s) used to refine structure: *SHELXL97* (Sheldrick, 2008); molecular graphics: *SHELXTL* (Sheldrick, 2008); software used to prepare material for publication: *PLATON* (Spek, 2009).

**Figure 1**

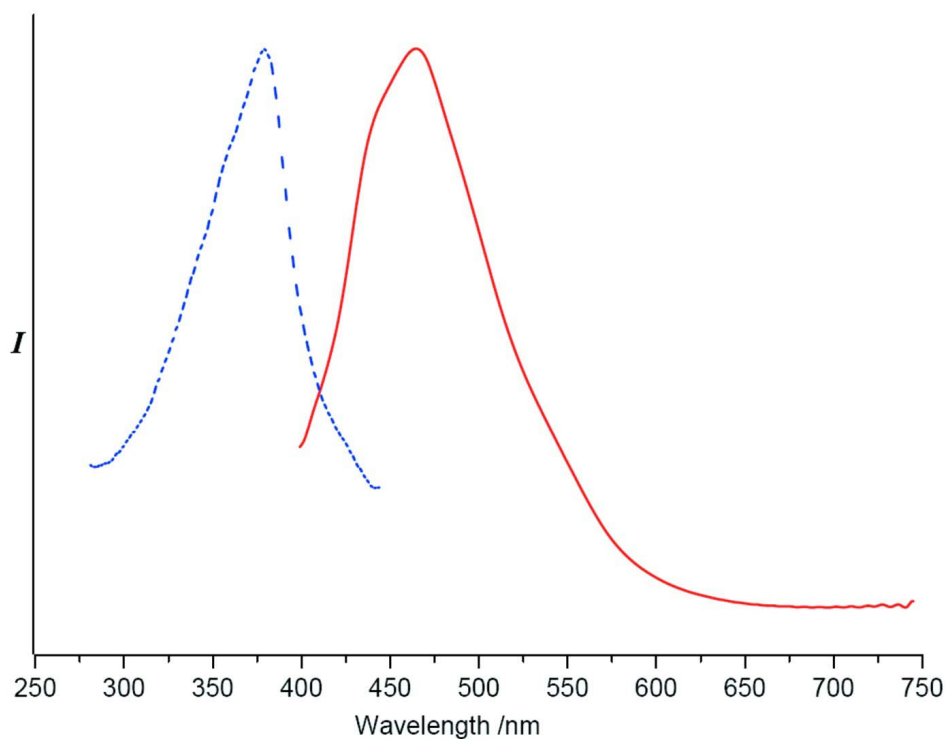
Part of crystal structure of the L^+ ligand and Pb^{II} centres in **1**. All H atoms have been omitted for clarity. [symmetry code: (#1) $-x, y+1, -z+1$; (#2) $x, y-1, z$; (#3) $-x+1, -y+1, -z+1$; (#4) $x+1, y-1, z$; (#5) $-x+1, -y+1, -z+1$.]

**Figure 2**

Sectional crystal structure of zigzag chain $[PbO_4(H_2O)]_n$. Pb, green; O, red; H, white

**Figure 3**

View along c axis of two-dimensional coordination polymer from Pb ions and L^+ ligands. PbO_7 , polyhedron: green; O: red; N: blue; C: grey.

**Figure 4**

Photoluminescent spectra of **1** (λ_{em} at 465 nm, upon λ_{ex} at 380 nm). I = relative intensity, em = emission, and ex = excitation.

Poly[*diaqua*[μ_6 -4,4'-(1,4-phenylene)bis(2,6-dimethylpyridine-3,5-dicarboxylato)]dilead(II)]

Crystal data

[Pb₂(C₂₄H₁₆N₂O₈)(H₂O)₂]

M_r = 910.80

Triclinic, *P* $\bar{1}$

Hall symbol: -P 1

a = 7.2182 (12) Å

b = 9.0635 (14) Å

c = 9.9589 (15) Å

α = 79.202 (2)°

β = 71.683 (2)°

γ = 85.494 (3)°

V = 607.43 (17) Å³

Z = 1

F(000) = 422

D_x = 2.490 Mg m⁻³

Mo *K*α radiation, λ = 0.71073 Å

Cell parameters from 1580 reflections

θ = 2.7–23.1°

μ = 13.90 mm⁻¹

T = 298 K

Block, colorless

0.25 × 0.23 × 0.23 mm

Data collection

Bruker SMART APEXII CCD

diffractometer

Radiation source: fine-focus sealed tube

Graphite monochromator

phi and ω scans

Absorption correction: multi-scan

(*SADABS*; Sheldrick, 2003)

T_{min} = 0.129, *T_{max}* = 0.142

3168 measured reflections

2119 independent reflections

1932 reflections with *I* > 2σ(*I*)

R_{int} = 0.015

θ_{\max} = 25.0°, θ_{\min} = 2.2°

h = -8→8

k = -10→10

l = -10→11

Refinement

Refinement on *F*²

Least-squares matrix: full

R[*F*² > 2σ(*F*²)] = 0.022

wR(*F*²) = 0.055

S = 1.01

2119 reflections

174 parameters

0 restraints

Primary atom site location: structure-invariant

direct methods

Secondary atom site location: difference Fourier map

Hydrogen site location: inferred from neighbouring sites

H-atom parameters constrained

w = 1/[σ²(*F_o*²) + (0.0347*P*)²]

where *P* = (*F_o*² + 2*F_c*²)/3

(Δ/σ)_{max} < 0.001

Δρ_{max} = 1.15 e Å⁻³

Δρ_{min} = -1.16 e Å⁻³

Special details

Geometry. All e.s.d.'s (except the e.s.d. in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell e.s.d.'s are taken into account individually in the estimation of e.s.d.'s in distances, angles and torsion angles; correlations between e.s.d.'s in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell e.s.d.'s is used for estimating e.s.d.'s involving l.s. planes.

Refinement. Refinement of *F*² against ALL reflections. The weighted *R*-factor *wR* and goodness of fit *S* are based on *F*², conventional *R*-factors *R* are based on *F*, with *F* set to zero for negative *F*². The threshold expression of *F*² > σ(*F*²) is used only for calculating *R*-factors(gt) etc. and is not relevant to the choice of reflections for refinement. *R*-factors based on *F*² are statistically about twice as large as those based on *F*, and *R*-factors based on ALL data will be even larger.

Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters (Å²)

	<i>x</i>	<i>y</i>	<i>z</i>	<i>U_{iso}</i> */ <i>U_{eq}</i>
Pb1	0.24557 (3)	0.11141 (2)	0.47717 (2)	0.02892 (9)
O1	0.0452 (5)	0.9117 (4)	0.6086 (4)	0.0281 (8)
O2	0.3224 (5)	0.8538 (4)	0.6568 (5)	0.0365 (9)
O3	0.2253 (5)	0.1995 (4)	0.7194 (4)	0.0343 (9)

O4	-0.0248 (6)	0.2448 (5)	0.6332 (5)	0.0382 (10)
O5	0.3782 (6)	-0.0957 (5)	0.3132 (5)	0.0441 (11)
H1	0.4881	-0.1412	0.3019	0.066*
H2	0.2924	-0.1629	0.3379	0.066*
N1	-0.2153 (6)	0.5547 (5)	0.9379 (5)	0.0275 (10)
C1	-0.1131 (7)	0.6798 (6)	0.8704 (6)	0.0245 (11)
C2	0.0591 (7)	0.6774 (5)	0.7553 (5)	0.0220 (11)
C3	0.1282 (7)	0.5404 (6)	0.7125 (5)	0.0213 (10)
C4	0.0148 (7)	0.4140 (6)	0.7758 (5)	0.0233 (11)
C5	-0.1579 (7)	0.4255 (6)	0.8888 (6)	0.0271 (12)
C6	-0.1922 (8)	0.8214 (6)	0.9291 (6)	0.0341 (13)
H6A	-0.2907	0.8662	0.8872	0.051*
H6B	-0.0882	0.8906	0.9060	0.051*
H6C	-0.2481	0.7975	1.0315	0.051*
C7	0.1542 (7)	0.8222 (6)	0.6705 (6)	0.0242 (11)
C8	0.0752 (7)	0.2736 (6)	0.7103 (6)	0.0243 (11)
C9	-0.2867 (8)	0.2946 (7)	0.9630 (7)	0.0379 (14)
H9A	-0.2869	0.2692	1.0611	0.057*
H9B	-0.2388	0.2102	0.9154	0.057*
H9C	-0.4172	0.3203	0.9603	0.057*
C10	0.3206 (7)	0.5253 (5)	0.5999 (5)	0.0192 (10)
C11	0.3313 (7)	0.5098 (6)	0.4608 (5)	0.0249 (11)
H11	0.2177	0.5160	0.4347	0.030*
C12	0.5090 (7)	0.4854 (6)	0.3615 (5)	0.0241 (11)
H12	0.5146	0.4761	0.2689	0.029*

Atomic displacement parameters (\AA^2)

	U^{11}	U^{22}	U^{33}	U^{12}	U^{13}	U^{23}
Pb1	0.02297 (13)	0.02526 (13)	0.03649 (14)	-0.00185 (8)	-0.00454 (9)	-0.00765 (9)
O1	0.0269 (19)	0.0225 (19)	0.037 (2)	-0.0014 (15)	-0.0143 (17)	-0.0010 (16)
O2	0.024 (2)	0.029 (2)	0.054 (3)	-0.0052 (16)	-0.0106 (19)	-0.0020 (18)
O3	0.034 (2)	0.029 (2)	0.044 (2)	0.0113 (17)	-0.0165 (19)	-0.0119 (18)
O4	0.030 (2)	0.036 (2)	0.056 (3)	0.0017 (17)	-0.017 (2)	-0.022 (2)
O5	0.029 (2)	0.048 (3)	0.059 (3)	0.0011 (19)	-0.013 (2)	-0.021 (2)
N1	0.021 (2)	0.031 (3)	0.026 (2)	0.0005 (19)	0.0000 (19)	-0.008 (2)
C1	0.020 (3)	0.027 (3)	0.026 (3)	0.002 (2)	-0.006 (2)	-0.008 (2)
C2	0.021 (2)	0.018 (3)	0.027 (3)	0.002 (2)	-0.009 (2)	-0.004 (2)
C3	0.020 (2)	0.026 (3)	0.018 (2)	0.002 (2)	-0.006 (2)	-0.005 (2)
C4	0.023 (3)	0.022 (3)	0.024 (3)	0.001 (2)	-0.007 (2)	-0.005 (2)
C5	0.021 (3)	0.033 (3)	0.027 (3)	-0.001 (2)	-0.009 (2)	-0.002 (2)
C6	0.031 (3)	0.029 (3)	0.037 (3)	0.006 (2)	0.000 (3)	-0.014 (2)
C7	0.024 (3)	0.019 (3)	0.027 (3)	0.001 (2)	-0.001 (2)	-0.007 (2)
C8	0.020 (3)	0.021 (3)	0.027 (3)	0.002 (2)	-0.002 (2)	-0.003 (2)
C9	0.030 (3)	0.034 (3)	0.041 (4)	-0.006 (3)	0.001 (3)	-0.003 (3)
C10	0.018 (2)	0.017 (2)	0.020 (2)	0.0026 (19)	-0.003 (2)	-0.0028 (19)
C11	0.020 (2)	0.028 (3)	0.029 (3)	0.001 (2)	-0.011 (2)	-0.005 (2)
C12	0.023 (3)	0.030 (3)	0.019 (3)	0.001 (2)	-0.005 (2)	-0.004 (2)

Geometric parameters (Å, °)

Pb1—O1 ⁱ	2.327 (4)	C2—C3	1.393 (7)
Pb1—O4	2.472 (4)	C2—C7	1.507 (7)
Pb1—O1 ⁱⁱ	2.538 (3)	C3—C4	1.390 (7)
Pb1—O3	2.638 (4)	C3—C10	1.501 (7)
Pb1—O5	2.644 (4)	C4—C5	1.405 (7)
O3—C8	1.248 (6)	C5—C9	1.494 (8)
O4—C8	1.277 (7)	C6—H6A	0.9600
C8—C4	1.510 (7)	C6—H6B	0.9600
O1—C7	1.294 (6)	C6—H6C	0.9600
O1—Pb1 ⁱⁱⁱ	2.327 (4)	C9—H9A	0.9600
O1—Pb1 ⁱⁱ	2.538 (3)	C9—H9B	0.9600
N1—C5	1.338 (7)	C9—H9C	0.9600
N1—C1	1.348 (7)	C10—C12 ^{iv}	1.391 (7)
O2—C7	1.230 (6)	C10—C11	1.395 (7)
O5—H1	0.8500	C11—C12	1.383 (7)
O5—H2	0.8500	C11—H11	0.9300
C1—C2	1.403 (7)	C12—C10 ^{iv}	1.391 (7)
C1—C6	1.507 (7)	C12—H12	0.9300
O1 ⁱ —Pb1—O4	79.28 (13)	C3—C4—C5	119.1 (5)
O1 ⁱ —Pb1—O1 ⁱⁱ	66.21 (14)	C3—C4—C8	117.7 (4)
O4—Pb1—O1 ⁱⁱ	75.30 (12)	C5—C4—C8	122.9 (5)
O1 ⁱ —Pb1—O3	89.30 (13)	N1—C5—C4	121.7 (5)
O4—Pb1—O3	51.16 (12)	N1—C5—C9	116.0 (5)
O1 ⁱⁱ —Pb1—O3	124.84 (11)	C4—C5—C9	122.3 (5)
O1 ⁱ —Pb1—O5	78.92 (13)	C1—C6—H6A	109.5
O4—Pb1—O5	151.48 (13)	C1—C6—H6B	109.5
O1 ⁱⁱ —Pb1—O5	79.19 (12)	H6A—C6—H6B	109.5
O3—Pb1—O5	146.08 (13)	C1—C6—H6C	109.5
C8—O3—Pb1	87.8 (3)	H6A—C6—H6C	109.5
C8—O4—Pb1	94.7 (3)	H6B—C6—H6C	109.5
O3—C8—O4	122.3 (5)	O2—C7—O1	121.6 (5)
O3—C8—C4	121.6 (5)	O2—C7—C2	124.0 (5)
O4—C8—C4	115.8 (4)	O1—C7—C2	114.4 (4)
C7—O1—Pb1 ⁱⁱⁱ	104.5 (3)	C5—C9—H9A	109.5
C7—O1—Pb1 ⁱⁱ	136.9 (3)	C5—C9—H9B	109.5
Pb1 ⁱⁱⁱ —O1—Pb1 ⁱⁱ	113.79 (14)	H9A—C9—H9B	109.5
C5—N1—C1	119.3 (4)	C5—C9—H9C	109.5
Pb1—O5—H1	125.1	H9A—C9—H9C	109.5
Pb1—O5—H2	107.8	H9B—C9—H9C	109.5
H1—O5—H2	106.8	C12 ^{iv} —C10—C11	119.3 (4)
N1—C1—C2	122.0 (5)	C12 ^{iv} —C10—C3	118.9 (4)
N1—C1—C6	115.8 (5)	C11—C10—C3	121.6 (4)
C2—C1—C6	122.2 (5)	C12—C11—C10	120.7 (5)
C3—C2—C1	118.6 (5)	C12—C11—H11	119.7
C3—C2—C7	120.8 (5)	C10—C11—H11	119.7
C1—C2—C7	120.2 (4)	C11—C12—C10 ^{iv}	120.0 (5)
C4—C3—C2	118.8 (5)	C11—C12—H12	120.0

C4—C3—C10	119.1 (4)	C10 ^{iv} —C12—H12	120.0
C2—C3—C10	122.1 (5)		

Symmetry codes: (i) $x, y-1, z$; (ii) $-x, -y+1, -z+1$; (iii) $x, y+1, z$; (iv) $-x+1, -y+1, -z+1$.

Hydrogen-bond geometry (Å, °)

$D-H\cdots A$	$D-H$	$H\cdots A$	$D\cdots A$	$D-H\cdots A$
O5—H2 \cdots O4 ^v	0.85	2.04	2.834 (6)	155
O5—H1 \cdots O3 ^{vi}	0.85	2.05	2.879 (5)	165

Symmetry codes: (v) $-x, -y, -z+1$; (vi) $-x+1, -y, -z+1$.