

## 4-Benzyloxy-3-methoxybenzonitrile

Muhammad Hanif,<sup>a</sup> Muhammad Rafiq,<sup>a</sup> Muhammad Saleem,<sup>b</sup> Ghulam Qadeer<sup>a\*</sup> and Wai-Yeung Wong<sup>c‡</sup>

<sup>a</sup>Department of Chemistry, Quaid-i-Azam University, Islamabad 45320, Pakistan,

<sup>b</sup>Department of Chemistry, University of Sargodah, Sargodah, Pakistan, and

<sup>c</sup>Department of Chemistry, Hong Kong Baptist University, Waterloo Road, Kowloon Tong, Hong Kong, People's Republic of China

Correspondence e-mail: qadeerqau@yahoo.com

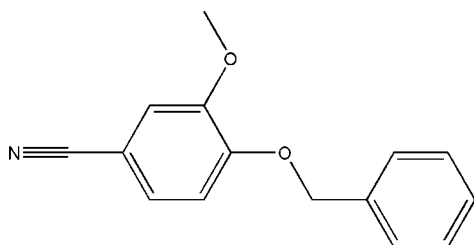
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Key indicators: single-crystal X-ray study;  $T = 173$  K; mean  $\sigma(\text{C}-\text{C}) = 0.002$  Å;  $R$  factor = 0.038;  $wR$  factor = 0.114; data-to-parameter ratio = 18.3.

In the molecule of the title compound,  $\text{C}_{15}\text{H}_{13}\text{NO}_2$ , the aromatic rings are oriented at a dihedral angle of  $81.65(3)^\circ$ . In the crystal structure, weak intermolecular  $\text{C}-\text{H}\cdots\text{N}$  hydrogen bonds link the molecules into chains along the  $b$  axis.

### Related literature

For the potential application of highly conjugated molecules in nanoelectronics, see: Tour (2003) and in optoelectronics, see: Lind *et al.* (2004); Ornelas *et al.* (2005, 2008). Terminal cyano groups provide the ability to coordinate to transition metal centres such as  $\text{RuCp}$ , see: Garcia *et al.* (2001); Ornelas *et al.* (2005). For bond-length data, see: Allen *et al.* (1987).



### Experimental

#### Crystal data

$\text{C}_{15}\text{H}_{13}\text{NO}_2$   
 $M_r = 239.26$   
 Monoclinic,  $P2_1/c$   
 $a = 14.9434(12)$  Å

$b = 9.5469(8)$  Å  
 $c = 8.8522(7)$  Å  
 $\beta = 102.663(2)^\circ$   
 $V = 1232.16(17)$  Å<sup>3</sup>

$Z = 4$   
 Mo  $K\alpha$  radiation  
 $\mu = 0.09$  mm<sup>-1</sup>

$T = 173$  K  
 $0.32 \times 0.25 \times 0.23$  mm

#### Data collection

Bruker SMART CCD area-detector diffractometer  
 Absorption correction: multi-scan (SADABS; Bruker, 2001)  
 $T_{\min} = 0.864$ ,  $T_{\max} = 0.980$   
 7286 measured reflections  
 2983 independent reflections  
 2499 reflections with  $I > 2\sigma(I)$   
 $R_{\text{int}} = 0.018$

#### Refinement

$R[F^2 > 2\sigma(F^2)] = 0.038$   
 $wR(F^2) = 0.114$   
 $S = 1.02$   
 2983 reflections  
 163 parameters  
 H-atom parameters constrained  
 $\Delta\rho_{\max} = 0.23$  e Å<sup>-3</sup>  
 $\Delta\rho_{\min} = -0.17$  e Å<sup>-3</sup>

**Table 1**

Hydrogen-bond geometry (Å, °).

$D-H\cdots A$	$D-H$	$H\cdots A$	$D\cdots A$	$D-H\cdots A$
$\text{C14}-\text{H14B}\cdots\text{N1}^i$	0.98	2.58	3.5170 (17)	160

Symmetry code: (i)  $-x, y - \frac{1}{2}, -z + \frac{1}{2}$ .

Data collection: SMART (Bruker, 2001); cell refinement: SAINT (Bruker, 2002); data reduction: SAINT; program(s) used to solve structure: SHELXS97 (Sheldrick, 2009); program(s) used to refine structure: SHELXL97 (Sheldrick, 2008); molecular graphics: ORTEP-3 for Windows (Farrugia, 1997); software used to prepare material for publication: SHELXTL (Sheldrick, 2009).

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Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: HK2626).

### References

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‡ Additional contact author, e-mail: rwywong@net3.hkbu.edu.hk.

**supplementary materials**

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## 4-Benzyloxy-3-methoxybenzonitrile

M. Hanif, M. Rafiq, M. Saleem, G. Qadeer and W.-Y. wong Wong

### Comment

Schiff base compounds have attracted great attention for many years. They play an important role in the development of coordination chemistry related to catalysis and enzymatic reactions, magnetism, photochromism and thermochromism. We report herein the crystal structure of the title compound.

In the title compound (Fig. 1), the bond lengths (Allen *et al.*, 1987) and angles are within normal ranges. Rings A (C1-C6) and B (C8-C13) are, of course, planar, and they are oriented at a dihedral angle of 81.65 (3)°.

In the crystal structure, weak intermolecular C-H...N hydrogen bonds (Table 1) link the molecules into chains along the b axis (Fig. 2), in which they may be effective in the stabilization of the structure.

The preparation of highly conjugated molecules has been of great interest for their potential applications in fields such as nanoelectronics (Tour, 2003) or optoelectronics (Ornelas *et al.*, 2005, 2008; Lind *et al.*, 2004). Terminal cyano groups provide the ability to coordinate to transition metal centres such as RuCp (Cp = cyclopentadienyl); (Garcia *et al.*, 2001; Ornelas *et al.*, 2005) which should result in an increase of the physical properties such as the first molecular hyperpolarizability  $\beta$ , which is reported to rise with the coordination to cyclopentadienylruthenium type centres (Ornelas *et al.*, 2005, 2008). As such the preparation of the  $\pi$ -conjugated title compound was intended for the preparation of dinuclear ruthenium complexes for nanoelectronic application.

### Experimental

For the preparation of the title compound, 4-(benzyloxy)-3-methoxy benzenamine (2.29 g, 10 mmol) was treated with sodium nitrite (0.7 g, 10 mmol) in the presence of concentrated hydrochloric acid (10 ml) at 273-278 K. Aqueous cupreous cyanate solution (48%, 1.05 g, 10 mmol) was added into the resulting diazonium salt (1.95 g, 8 mmol). The obtained title compound was separated and recrystallized in ethanol/THF mixture (yield; 65%, m.p. 411-412 K).

### Refinement

H atoms were positioned geometrically, with C-H = 0.95, 0.99 and 0.98 Å for aromatic, methylene and methyl H, respectively, and constrained to ride on their parent atoms, with  $U_{\text{iso}}(\text{H}) = xU_{\text{eq}}(\text{C})$ , where  $x = 1.5$  for methyl H and  $x = 1.2$  for all other H atoms.

### Figures

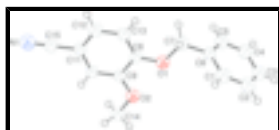


Fig. 1. The molecular structure of the title molecule, with the atom-numbering scheme.

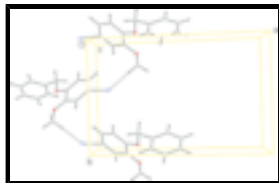


Fig. 2. A partial packing diagram of the title compound. Hydrogen bonds are shown as dashed lines.

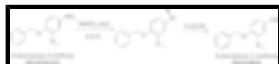


Fig. 3. The formation of the title compound.

## 4-Benzoyloxy-3-methoxybenzonitrile

### Crystal data

$C_{15}H_{13}NO_2$

$M_r = 239.26$

Monoclinic,  $P2_1/c$

Hall symbol: -P 2ybc

$a = 14.9434$  (12) Å

$b = 9.5469$  (8) Å

$c = 8.8522$  (7) Å

$\beta = 102.663$  (2)°

$V = 1232.16$  (17) Å<sup>3</sup>

$Z = 4$

$F_{000} = 504$

$D_x = 1.290$  Mg m<sup>-3</sup>

Mo  $K\alpha$  radiation

$\lambda = 0.71073$  Å

Cell parameters from 7856 reflections

$\theta = 2.6$ – $28.3$ °

$\mu = 0.09$  mm<sup>-1</sup>

$T = 173$  K

Block, colorless

$0.32 \times 0.25 \times 0.23$  mm

### Data collection

Bruker SMART CCD area-detector diffractometer

Radiation source: fine-focus sealed tube

Monochromator: graphite

$T = 173$  K

$\omega$  and  $\varphi$  scans

Absorption correction: multi-scan (SADABS; Bruker, 2001)

$T_{\min} = 0.864$ ,  $T_{\max} = 0.980$

7286 measured reflections

2983 independent reflections

2499 reflections with  $I > 2\sigma(I)$

$R_{\text{int}} = 0.018$

$\theta_{\text{max}} = 28.3$ °

$\theta_{\text{min}} = 2.6$ °

$h = -17 \rightarrow 19$

$k = -11 \rightarrow 12$

$l = -11 \rightarrow 11$

### Refinement

Refinement on  $F^2$

Least-squares matrix: full

$R[F^2 > 2\sigma(F^2)] = 0.038$

$wR(F^2) = 0.114$

$S = 1.02$

2983 reflections

Secondary atom site location: difference Fourier map

Hydrogen site location: inferred from neighbouring sites

H-atom parameters constrained

$w = 1/[\sigma^2(F_o^2) + (0.0606P)^2 + 0.2089P]$

where  $P = (F_o^2 + 2F_c^2)/3$

$(\Delta/\sigma)_{\text{max}} < 0.001$

$\Delta\rho_{\text{max}} = 0.23$  e Å<sup>-3</sup>

163 parameters

$$\Delta\rho_{\min} = -0.17 \text{ e } \text{\AA}^{-3}$$

Primary atom site location: structure-invariant direct methods

Extinction correction: none

### Special details

**Geometry.** All e.s.d.'s (except the e.s.d. in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell e.s.d.'s are taken into account individually in the estimation of e.s.d.'s in distances, angles and torsion angles; correlations between e.s.d.'s in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell e.s.d.'s is used for estimating e.s.d.'s involving l.s. planes.

**Refinement.** Refinement of  $F^2$  against ALL reflections. The weighted  $R$ -factor  $wR$  and goodness of fit  $S$  are based on  $F^2$ , conventional  $R$ -factors  $R$  are based on  $F$ , with  $F$  set to zero for negative  $F^2$ . The threshold expression of  $F^2 > \sigma(F^2)$  is used only for calculating  $R$ -factors(gt) *etc.* and is not relevant to the choice of reflections for refinement.  $R$ -factors based on  $F^2$  are statistically about twice as large as those based on  $F$ , and  $R$ -factors based on ALL data will be even larger.

### Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters ( $\text{\AA}^2$ )

	<i>x</i>	<i>y</i>	<i>z</i>	$U_{\text{iso}}^*/U_{\text{eq}}$
O1	0.26381 (5)	0.56473 (9)	0.00751 (9)	0.0409 (2)
O2	0.26362 (5)	0.38653 (9)	0.22449 (9)	0.0423 (2)
N1	-0.09205 (7)	0.55876 (14)	0.35630 (13)	0.0545 (3)
C1	0.44027 (9)	0.66914 (15)	-0.09352 (15)	0.0492 (3)
H1A	0.4473	0.7296	-0.0065	0.059*
C2	0.51719 (9)	0.62309 (16)	-0.14277 (16)	0.0520 (3)
H2A	0.5765	0.6523	-0.0893	0.062*
C3	0.50825 (9)	0.53599 (14)	-0.26781 (15)	0.0479 (3)
H3A	0.5612	0.5033	-0.3000	0.058*
C4	0.42195 (9)	0.49571 (15)	-0.34716 (16)	0.0510 (3)
H4A	0.4155	0.4364	-0.4351	0.061*
C5	0.34464 (8)	0.54158 (13)	-0.29886 (14)	0.0440 (3)
H5A	0.2855	0.5137	-0.3543	0.053*
C6	0.35295 (8)	0.62749 (12)	-0.17068 (13)	0.0375 (2)
C7	0.26977 (8)	0.66610 (13)	-0.11153 (14)	0.0425 (3)
H7A	0.2760	0.7620	-0.0680	0.051*
H7B	0.2141	0.6623	-0.1961	0.051*
C8	0.18984 (7)	0.57131 (11)	0.07225 (12)	0.0339 (2)
C9	0.18915 (7)	0.47256 (11)	0.19135 (12)	0.0328 (2)
C10	0.11690 (7)	0.46987 (12)	0.26518 (12)	0.0343 (2)
H10A	0.1166	0.4048	0.3462	0.041*
C11	0.04377 (7)	0.56410 (12)	0.21963 (12)	0.0358 (2)
C12	0.04428 (8)	0.66074 (13)	0.10355 (13)	0.0403 (3)
H12A	-0.0055	0.7241	0.0736	0.048*
C13	0.11769 (8)	0.66500 (13)	0.03071 (13)	0.0399 (3)
H13A	0.1185	0.7324	-0.0479	0.048*
C14	0.26552 (8)	0.28301 (13)	0.34102 (13)	0.0411 (3)
H14A	0.3220	0.2280	0.3534	0.062*
H14B	0.2124	0.2210	0.3105	0.062*
H14C	0.2635	0.3288	0.4393	0.062*

## supplementary materials

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C15                    −0.03198 (7)                    0.55965 (13)                    0.29606 (13)                    0.0411 (3)

### Atomic displacement parameters ( $\text{\AA}^2$ )

	$U^{11}$	$U^{22}$	$U^{33}$	$U^{12}$	$U^{13}$	$U^{23}$
O1	0.0369 (4)	0.0459 (5)	0.0439 (4)	0.0084 (3)	0.0176 (3)	0.0109 (3)
O2	0.0311 (4)	0.0469 (5)	0.0509 (5)	0.0104 (3)	0.0132 (3)	0.0138 (4)
N1	0.0373 (5)	0.0777 (8)	0.0499 (6)	0.0106 (5)	0.0128 (4)	0.0015 (5)
C1	0.0493 (7)	0.0525 (7)	0.0454 (6)	−0.0055 (5)	0.0094 (5)	−0.0083 (5)
C2	0.0378 (6)	0.0625 (8)	0.0544 (7)	−0.0090 (6)	0.0071 (5)	−0.0007 (6)
C3	0.0408 (6)	0.0529 (7)	0.0547 (7)	−0.0047 (5)	0.0204 (5)	0.0033 (6)
C4	0.0481 (7)	0.0573 (8)	0.0528 (7)	−0.0108 (6)	0.0220 (6)	−0.0127 (6)
C5	0.0383 (6)	0.0496 (7)	0.0463 (6)	−0.0109 (5)	0.0142 (5)	−0.0036 (5)
C6	0.0408 (6)	0.0350 (5)	0.0392 (5)	−0.0020 (4)	0.0139 (4)	0.0061 (4)
C7	0.0475 (6)	0.0398 (6)	0.0435 (6)	0.0055 (5)	0.0171 (5)	0.0084 (5)
C8	0.0317 (5)	0.0370 (5)	0.0337 (5)	0.0025 (4)	0.0085 (4)	−0.0009 (4)
C9	0.0263 (5)	0.0351 (5)	0.0362 (5)	0.0026 (4)	0.0054 (4)	−0.0002 (4)
C10	0.0295 (5)	0.0386 (5)	0.0348 (5)	0.0000 (4)	0.0068 (4)	−0.0003 (4)
C11	0.0293 (5)	0.0427 (6)	0.0354 (5)	0.0022 (4)	0.0072 (4)	−0.0078 (4)
C12	0.0367 (5)	0.0433 (6)	0.0401 (5)	0.0119 (4)	0.0069 (4)	−0.0025 (5)
C13	0.0421 (6)	0.0411 (6)	0.0373 (5)	0.0090 (5)	0.0106 (4)	0.0032 (4)
C14	0.0366 (5)	0.0411 (6)	0.0431 (6)	0.0037 (4)	0.0034 (4)	0.0065 (5)
C15	0.0320 (5)	0.0523 (7)	0.0385 (6)	0.0063 (5)	0.0067 (4)	−0.0043 (5)

### Geometric parameters ( $\text{\AA}$ , $^\circ$ )

C1—C2	1.3872 (18)	C8—C13	1.3870 (15)
C1—C6	1.3916 (17)	C8—C9	1.4161 (15)
C1—H1A	0.9500	C9—O2	1.3624 (12)
C2—C3	1.3673 (19)	C9—C10	1.3794 (14)
C2—H2A	0.9500	C10—C11	1.4046 (14)
C3—C4	1.3810 (18)	C10—H10A	0.9500
C3—H3A	0.9500	C11—C12	1.3822 (16)
C4—C5	1.3878 (17)	C11—C15	1.4411 (14)
C4—H4A	0.9500	C12—C13	1.3897 (15)
C5—C6	1.3834 (16)	C12—H12A	0.9500
C5—H5A	0.9500	C13—H13A	0.9500
C6—C7	1.4967 (15)	C14—O2	1.4244 (13)
C7—O1	1.4478 (13)	C14—H14A	0.9800
C7—H7A	0.9900	C14—H14B	0.9800
C7—H7B	0.9900	C14—H14C	0.9800
C8—O1	1.3535 (12)	C15—N1	1.1402 (15)
C8—O1—C7	117.61 (8)	O1—C8—C13	125.25 (10)
C9—O2—C14	117.38 (8)	O1—C8—C9	115.10 (9)
C2—C1—C6	120.51 (12)	C13—C8—C9	119.65 (9)
C2—C1—H1A	119.7	O2—C9—C10	125.03 (9)
C6—C1—H1A	119.7	O2—C9—C8	115.01 (9)
C3—C2—C1	120.43 (12)	C10—C9—C8	119.96 (9)

C3—C2—H2A	119.8	C9—C10—C11	119.51 (10)
C1—C2—H2A	119.8	C9—C10—H10A	120.2
C2—C3—C4	119.72 (12)	C11—C10—H10A	120.2
C2—C3—H3A	120.1	C12—C11—C10	120.68 (10)
C4—C3—H3A	120.1	C12—C11—C15	120.10 (10)
C3—C4—C5	120.20 (12)	C10—C11—C15	119.22 (10)
C3—C4—H4A	119.9	C11—C12—C13	119.87 (10)
C5—C4—H4A	119.9	C11—C12—H12A	120.1
C6—C5—C4	120.58 (11)	C13—C12—H12A	120.1
C6—C5—H5A	119.7	C8—C13—C12	120.33 (10)
C4—C5—H5A	119.7	C8—C13—H13A	119.8
C5—C6—C1	118.55 (11)	C12—C13—H13A	119.8
C5—C6—C7	120.06 (10)	O2—C14—H14A	109.5
C1—C6—C7	121.28 (11)	O2—C14—H14B	109.5
O1—C7—C6	106.16 (9)	H14A—C14—H14B	109.5
O1—C7—H7A	110.5	O2—C14—H14C	109.5
C6—C7—H7A	110.5	H14A—C14—H14C	109.5
O1—C7—H7B	110.5	H14B—C14—H14C	109.5
C6—C7—H7B	110.5	N1—C15—C11	178.73 (13)
H7A—C7—H7B	108.7		

*Hydrogen-bond geometry* ( $\text{\AA}$ ,  $^\circ$ )

$D-H\cdots A$	$D-H$	$H\cdots A$	$D\cdots A$	$D-H\cdots A$
C14—H14B $\cdots$ N1 <sup>i</sup>	0.98	2.58	3.5170 (17)	160

Symmetry codes: (i)  $-x, y-1/2, -z+1/2$ .

Fig. 1

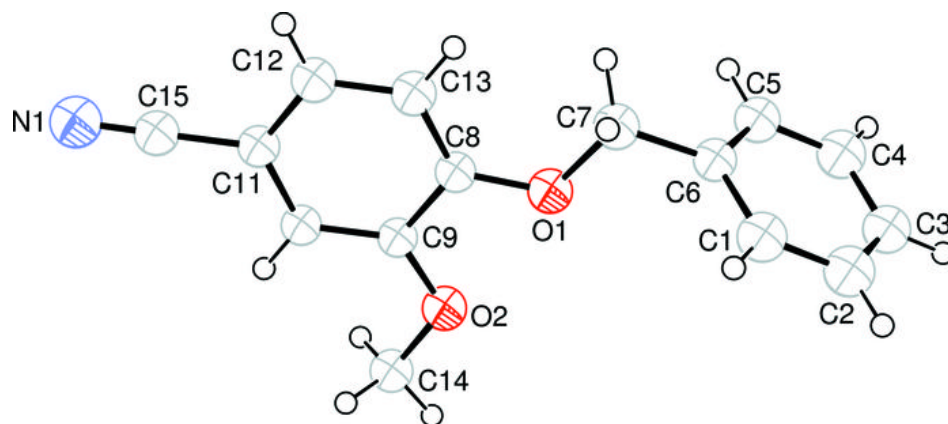




Fig. 2

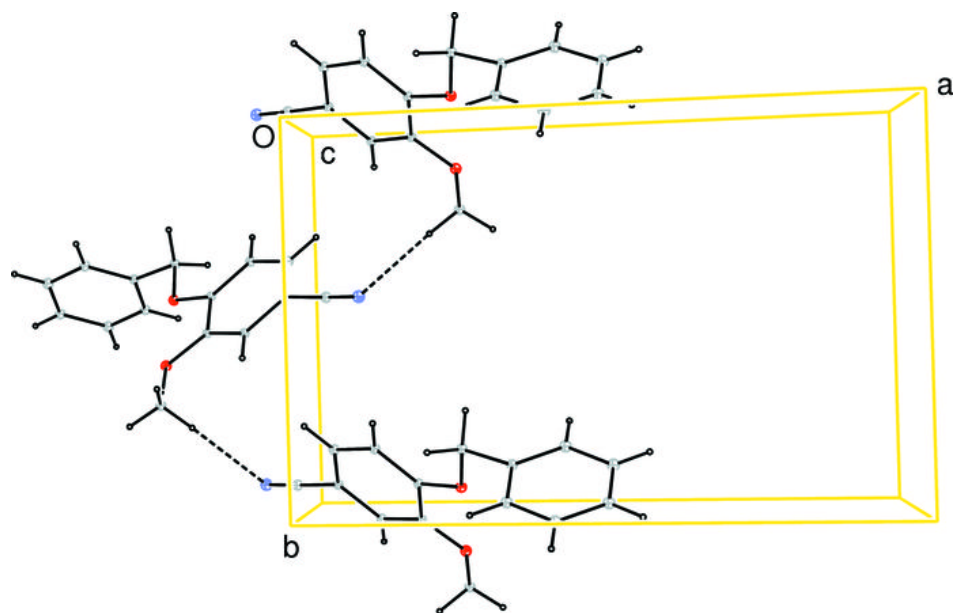


Fig. 3

