

Acta Crystallographica Section E

## Structure Reports

Online

ISSN 1600-5368

2-[(*E*)-(4-Bromophenyl)iminomethyl]-4-chlorophenolXiao-Li Gao,<sup>a</sup> Si-Si Feng,<sup>b</sup> Cai-Xia Yuan<sup>b</sup> and Miao-Li Zhu<sup>b\*</sup><sup>a</sup>Department of Chemistry, Taiyuan Normal College, Taiyuan, Shanxi 030031, People's Republic of China, and <sup>b</sup>Institute of Molecular Science, Key Laboratory of Chemical Biology and Molecular Engineering of the Education Ministry, Shanxi University, Taiyuan, Shanxi 030006, People's Republic of China

Correspondence e-mail: miaoli@sxu.edu.cn

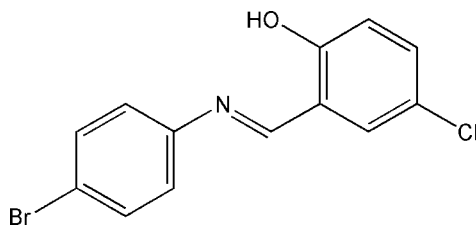
Received 12 November 2013; accepted 15 January 2014

Key indicators: single-crystal X-ray study;  $T = 298$  K; mean  $\sigma(\text{C}-\text{C}) = 0.006$  Å;  $R$  factor = 0.054;  $wR$  factor = 0.128; data-to-parameter ratio = 19.4.

In the title compound,  $\text{C}_{13}\text{H}_9\text{BrClNO}$ , the dihedral angle between the substituted benzene rings is  $44.25$  ( $11$ )°. There are strong intramolecular  $\text{O}-\text{H}\cdots\text{N}$  hydrogen bonds, which generate  $S(6)$  rings, and also intermolecular  $\text{Cl}\cdots\text{Cl}$  [ $3.431$  (3) Å] and  $\text{Br}\cdots\text{Br}$  [ $3.846$  (1) Å] contacts. The crystal packing a  $\text{C}-\text{H}\cdots\text{O}$  and  $\text{C}-\text{H}\cdots\pi$  interactions.

## Related literature

For background to the biological activity of Schiff bases, see: Akmal *et al.* (2007); Li *et al.* (2007, 2011); Lu *et al.* (2011); Ma *et al.* (2011); Rehmana *et al.* (2008); Ritter *et al.* (2009); Vanco *et al.* (2008); Yuan *et al.* (2009, 2010). For related structures, see: Ardakani *et al.* (2011). For hydrogen-bond motifs, see: Bernstein *et al.* (1995).



## Experimental

## Crystal data

 $\text{C}_{13}\text{H}_9\text{BrClNO}$  $M_r = 310.57$ Orthorhombic,  $Pccn$  $a = 6.9964$  (15) Å $b = 55.786$  (12) Å $c = 6.1443$  (14) Å $V = 2398.1$  (9) Å<sup>3</sup> $Z = 8$ Mo  $K\alpha$  radiation $\mu = 3.63$  mm<sup>-1</sup> $T = 298$  K $0.30 \times 0.25 \times 0.20$  mm

## Data collection

Bruker SMART 1K CCD area-detector diffractometer  
Absorption correction: multi-scan (*SADABS*; Sheldrick, 2000)  
 $T_{\min} = 0.409$ ,  $T_{\max} = 0.530$

30095 measured reflections  
3013 independent reflections  
2056 reflections with  $I > 2\sigma(I)$   
 $R_{\text{int}} = 0.061$

## Refinement

 $R[F^2 > 2\sigma(F^2)] = 0.054$  $wR(F^2) = 0.128$  $S = 1.14$ 

3013 reflections

155 parameters

H-atom parameters constrained

 $\Delta\rho_{\max} = 0.42$  e Å<sup>-3</sup> $\Delta\rho_{\min} = -0.95$  e Å<sup>-3</sup>

Table 1

Hydrogen-bond geometry (Å, °).

$\text{Cg1}$  and  $\text{Cg2}$  are the centroids of the  $\text{C1}-\text{C6}$  and  $\text{C8}-\text{C13}$  benzene rings, respectively.

$D-\text{H}\cdots A$	$D-\text{H}$	$\text{H}\cdots A$	$D\cdots A$	$D-\text{H}\cdots A$
$\text{O1}-\text{H1}\cdots\text{N1}$	0.82	1.87	2.593 (4)	147
$\text{C2}-\text{H2}\cdots\text{Cg1}^{\text{i}}$	0.93	2.82	3.489 (5)	129
$\text{C5}-\text{H5}\cdots\text{Cg1}^{\text{ii}}$	0.93	2.85	3.513 (5)	129
$\text{C10}-\text{H10}\cdots\text{Cg2}^{\text{iii}}$	0.93	2.75	3.460 (5)	133
$\text{C13}-\text{H13}\cdots\text{Cg2}^{\text{iv}}$	0.93	2.78	3.473 (5)	132

Symmetry codes: (i)  $-x + \frac{1}{2}, y, z - \frac{1}{2}$ ; (ii)  $-x + \frac{3}{2}, y, z - \frac{3}{2}$ ; (iii)  $-x + \frac{1}{2}, y, z - \frac{3}{2}$ ; (iv)  $-x + \frac{3}{2}, y, z - \frac{1}{2}$ .

Data collection: *SMART* (Bruker, 2000); cell refinement: *SAINT* (Bruker, 2000); data reduction: *SAINT*; program(s) used to solve structure: *SHELXS97* (Sheldrick, 2008); program(s) used to refine structure: *SHELXL97* (Sheldrick, 2008); molecular graphics: *ORTEP-3 for Windows* (Farrugia, 2012); software used to prepare material for publication: *publCIF* (Westrip, 2010).

This work was supported financially by the National Natural Science Foundation of China (grant Nos. 21001070 and 21171109), the Specialized Research Fund for the Doctoral Program of Higher Education (grant No. 20111401110002) and the Natural Science Foundation of Shanxi Province (grant Nos. 2010011011-2, 2011011009-1 and 2011021006-2).

Supporting information for this paper is available from the IUCr electronic archives (Reference: FJ2653).

## References

- Akmal, S. G., Mohsen, S. A., Atiat, S. B. & Said, M. T. (2007). *Spectrochim. Acta Part A*, **67**, 114–121.
- Ardakani, A. A., Kia, R., Kargar, H. & Tahir, M. N. (2011). *Acta Cryst. E* **67**, o597.
- Bernstein, J., Davis, R. E., Shimon, L. & Chang, N.-L. (1995). *Angew. Chem. Int. Ed. Engl.* **34**, 1555–1573.
- Bruker (2000). *SMART* and *SAINT*. Bruker AXS Inc., Madison, Wisconsin, USA.
- Farrugia, L. J. (2012). *J. Appl. Cryst.* **45**, 849–854.
- Li, Y., Lu, L.-P., Zhu, M.-L., Wang, Q.-M., Yuan, C.-X., Xing, S., Fu, X.-Q. & Mei, Y.-H. (2011). *BioMetals*, **24**, 993–1004.
- Li, Y.-G., Shi, D.-H., Zhu, H.-L., Yan, H. & Ng, S. W. (2007). *Inorg. Chim. Acta*, **360**, 2881–2889.
- Lu, L.-P., Yue, J.-J., Yuan, C.-X., Zhu, M.-L., Han, H., Liu, Z.-W. & Guo, M.-L. (2011). *J. Inorg. Biochem.* **105**, 1323–1328.
- Ma, L., Lu, L.-P., Zhu, M.-L., Wang, Q.-M., Li, Y., Xing, S., Fu, X.-Q., Gao, Z.-Q. & Dong, Y.-H. (2011). *Dalton Trans.* **40**, 6532–6540.
- Rehmana, W., Samana, F. & Ahmadb, I. (2008). *Russ. J. Coord. Chem.* **34**, 678–682.

- Ritter, E., Przybylski, P., Brzezinski, B. & Bartl, F. (2009). *Curr. Org. Chem.* **13**, 241–249.
- Sheldrick, G. M. (2000). *SADABS*. University of Göttingen, Germany.
- Sheldrick, G. M. (2008). *Acta Cryst.* **A64**, 112–122.
- Vanco, J., Marek, J., Travnicek, Z., Racanska, E., Muselik, J. & Svajlenova, O. (2008). *J. Inorg. Biochem.* **102**, 595–605.
- Westrip, S. P. (2010). *J. Appl. Cryst.* **43**, 920–925.
- Yuan, C., Lu, L., Gao, X., Wu, Y., Guo, M., Li, Y., Fu, X. & Zhu, M. (2009). *J. Biol. Inorg. Chem.* **14**, 841–851.
- Yuan, C., Lu, L., Wu, Y., Liu, Z., Guo, M., Xing, S., Fu, X. & Zhu, M. (2010). *J. Inorg. Biochem.* **104**, 978–986.

## supplementary materials

*Acta Cryst.* (2014). E70, o235–o236 [doi:10.1107/S1600536814000981]

## 2-[(*E*)-(4-Bromophenyl)iminomethyl]-4-chlorophenol

Xiao-Li Gao, Si-Si Feng, Cai-Xia Yuan and Miao-Li Zhu

### 1. Comment

Schiff bases are condensed by primary amines and carbonyl compounds, containing strong electronegative with atoms O and N, thus it is easy to coordinate with the metal ions to form stable complexes (Akmal *et al.*, 2007; Rehmana *et al.*, 2008). It is reported that metal complexes of Schiff base derivatives have a variety of important biological activities, such as anti-bacterial, anti-cancer, anti-tumor, hypoglycemic and so on (Vanco *et al.*, 2008; Li *et al.*, 2007; Ritter *et al.*, 2009). Our reports indicated that copper and vanadium complexes of Schiff bases are potential inhibitors over protein tyrosine phosphatases (Li *et al.*, 2011; Lu *et al.*, 2011; Ma *et al.*, 2011; Yuan *et al.*, 2009, 2010).

We report here the synthesis and characterization a potentially bidentate Schiff base derivative, (I), and prepared from the condensation reaction of an equimolar proportion of 5-chloro-salicylaldehyde and 4-bromo-aniline in absolute ethanol. The molecular structure is depicted in Fig. 1. X-ray structural analysis confirmed that in the title compound, (I), the dihedral angle between the substituted benzene rings is nearly  $44.25(11)^\circ$ , similar to the compound 4-bromo-2-[(*E*)-(4-chlorophenyl)iminomethyl]phenol (Ardakani *et al.*, 2011). In the crystal, there are strong intramolecular O—H $\cdots$ N hydrogen bonds with a distance of 2.593(5) Å between donor and acceptor, which generate S(6) ring and intermolecular C11 $\cdots$ C11<sup>i</sup> [3.430(2) Å, *i* -*x*, -*y*, -*z*] as well as Br1 $\cdots$ Br1<sup>ii</sup> [3.846(1) Å, *ii* 2 - *x*, -*y*, -*z*] contacts. The crystal packing is further stabilized by intermolecular C—H $\cdots$ O and C—H $\cdots$  $\pi$  interactions (Table 1).

### 2. Experimental

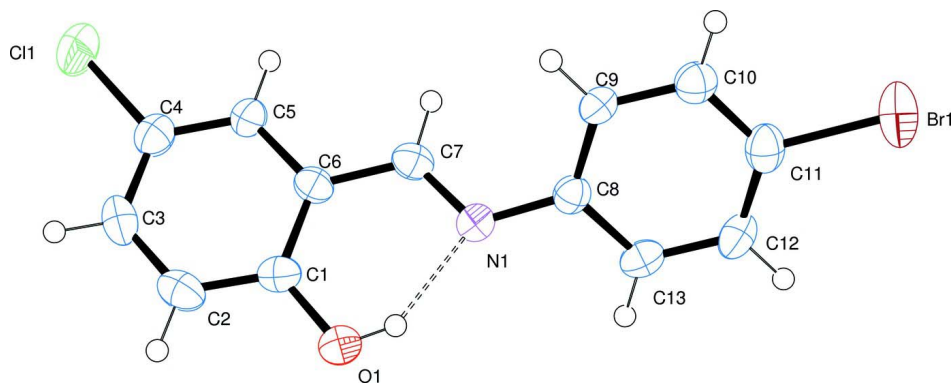
A 0.1566 g (1.0 mmol) 5-chloro-salicylaldehyde in 15 ml of absolute ethanol was heated until thoroughly dissolved and 0.1720 g (1.0 mmol) of 4-bromo-aniline in 5 ml of absolute ethanol was added dropwise with a constant stirring. The reaction mixture was heated under refluxing for 3 h. After cooling slowly, the orange powder was separated out. Orange-red crystal was obtained from filter after two weeks.

### 3. Refinement

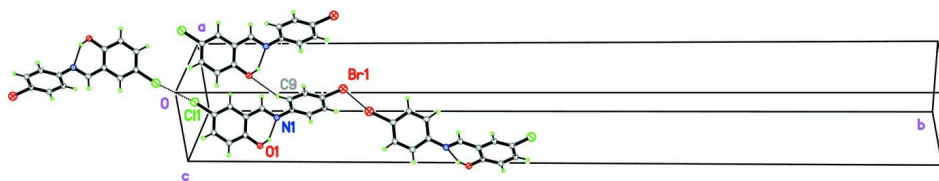
H atoms attached to C and O of (I) were placed in geometrically idealized positions with  $C_{sp^2}-H = 0.93$  Å and  $O-H = 0.84$  Å and constrained to refine with  $U_{iso}(H) = 1.2U_{eq}(C)$  and  $U_{iso}(H) = 1.5U_{eq}(O)$ .

### Computing details

Data collection: *SMART* (Bruker, 2000); cell refinement: *SAINTE* (Bruker, 2000); data reduction: *SAINTE* (Bruker, 2000); program(s) used to solve structure: *SHELXS97* (Sheldrick, 2008); program(s) used to refine structure: *SHELXL97* (Sheldrick, 2008); molecular graphics: *ORTEP-3 for Windows* (Farrugia, 2012); software used to prepare material for publication: *pubCIF* (Westrip, 2010).

**Figure 1**

A view of the structure of (I), with displacement ellipsoids drawn at the 30% probability level. Dashed line indicates hydrogen-bonding interaction.

**Figure 2**

The packing diagram of the title compound, showing the intermolecular O—H...N and intermolecular Cl...Cl, Br...Br and C—H...O interactions (dotted lines).

## 2-[(E)-(4-Bromophenyl)iminomethyl]-4-chlorophenol

### Crystal data

$C_{13}H_9BrClNO$

$M_r = 310.57$

Orthorhombic, *Pccn*

Hall symbol: -P 2ab 2ac

$a = 6.9964$  (15) Å

$b = 55.786$  (12) Å

$c = 6.1443$  (14) Å

$V = 2398.1$  (9) Å<sup>3</sup>

$Z = 8$

$F(000) = 1232$

$D_x = 1.720$  Mg m<sup>-3</sup>

Mo  $K\alpha$  radiation,  $\lambda = 0.71073$  Å

Cell parameters from 4170 reflections

$\theta = 2.2$ – $22.6^\circ$

$\mu = 3.63$  mm<sup>-1</sup>

$T = 298$  K

Block, colourless

$0.30 \times 0.25 \times 0.20$  mm

### Data collection

Bruker SMART 1K CCD area-detector  
diffractometer

Radiation source: fine-focus sealed tube

Graphite monochromator

$\omega$  scans

Absorption correction: multi-scan

(*SADABS*; Sheldrick, 2000)

$T_{\min} = 0.409$ ,  $T_{\max} = 0.530$

30095 measured reflections

3013 independent reflections

2056 reflections with  $I > 2\sigma(I)$

$R_{\text{int}} = 0.061$

$\theta_{\text{max}} = 28.4^\circ$ ,  $\theta_{\text{min}} = 2.2^\circ$

$h = -9 \rightarrow 9$

$k = -74 \rightarrow 74$

$l = -8 \rightarrow 8$

Refinement

Refinement on  $F^2$   
 Least-squares matrix: full  
 $R[F^2 > 2\sigma(F^2)] = 0.054$   
 $wR(F^2) = 0.128$   
 $S = 1.14$   
 3013 reflections  
 155 parameters  
 0 restraints  
 Primary atom site location: structure-invariant  
 direct methods

Secondary atom site location: difference Fourier  
 map  
 Hydrogen site location: inferred from  
 neighbouring sites  
 H-atom parameters constrained  
 $w = 1/[\sigma^2(F_o^2) + (0.027P)^2 + 7.4937P]$   
 where  $P = (F_o^2 + 2F_c^2)/3$   
 $(\Delta/\sigma)_{\max} = 0.001$   
 $\Delta\rho_{\max} = 0.42 \text{ e } \text{\AA}^{-3}$   
 $\Delta\rho_{\min} = -0.95 \text{ e } \text{\AA}^{-3}$

Special details

**Geometry.** All e.s.d.'s (except the e.s.d. in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell e.s.d.'s are taken into account individually in the estimation of e.s.d.'s in distances, angles and torsion angles; correlations between e.s.d.'s in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell e.s.d.'s is used for estimating e.s.d.'s involving l.s. planes.

**Refinement.** Refinement of  $F^2$  against ALL reflections. The weighted  $R$ -factor  $wR$  and goodness of fit  $S$  are based on  $F^2$ , conventional  $R$ -factors  $R$  are based on  $F$ , with  $F$  set to zero for negative  $F^2$ . The threshold expression of  $F^2 > \sigma(F^2)$  is used only for calculating  $R$ -factors(gt) *etc.* and is not relevant to the choice of reflections for refinement.  $R$ -factors based on  $F^2$  are statistically about twice as large as those based on  $F$ , and  $R$ -factors based on ALL data will be even larger.

Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters ( $\text{\AA}^2$ )

	$x$	$y$	$z$	$U_{\text{iso}}^*/U_{\text{eq}}$
Br1	-0.03618 (9)	0.228348 (8)	-0.07815 (11)	0.0634 (2)
C1	-0.0360 (6)	0.08714 (7)	0.5859 (7)	0.0339 (8)
C2	-0.0711 (6)	0.06455 (8)	0.6715 (7)	0.0404 (10)
H2	-0.1186	0.0631	0.8123	0.049*
C3	-0.0364 (6)	0.04435 (7)	0.5506 (8)	0.0426 (11)
H3	-0.0609	0.0293	0.6092	0.051*
C4	0.0346 (6)	0.04637 (7)	0.3425 (7)	0.0372 (10)
C5	0.0691 (6)	0.06843 (7)	0.2537 (7)	0.0333 (9)
H5	0.1169	0.0695	0.1129	0.040*
C6	0.0332 (6)	0.08932 (7)	0.3722 (6)	0.0306 (8)
C7	0.0538 (6)	0.11247 (7)	0.2675 (7)	0.0335 (9)
H7	0.0897	0.1132	0.1219	0.040*
C8	0.0157 (5)	0.15412 (7)	0.2599 (7)	0.0304 (8)
C9	-0.0626 (6)	0.15603 (7)	0.0518 (7)	0.0335 (9)
H9	-0.1044	0.1424	-0.0209	0.040*
C10	-0.0780 (6)	0.17821 (7)	-0.0459 (7)	0.0359 (9)
H10	-0.1315	0.1795	-0.1839	0.043*
C11	-0.0140 (6)	0.19838 (7)	0.0610 (8)	0.0392 (10)
C12	0.0610 (6)	0.19682 (7)	0.2681 (8)	0.0401 (10)
H12	0.1013	0.2106	0.3402	0.048*
C13	0.0759 (6)	0.17470 (7)	0.3676 (7)	0.0352 (9)
H13	0.1265	0.1736	0.5072	0.042*
Cl1	0.0736 (2)	0.02081 (2)	0.1879 (2)	0.0614 (4)
N1	0.0235 (5)	0.13186 (6)	0.3721 (5)	0.0322 (7)
O1	-0.0721 (5)	0.10666 (5)	0.7096 (5)	0.0474 (8)
H1	-0.0470	0.1188	0.6401	0.071*

Atomic displacement parameters ( $\text{\AA}^2$ )

	$U^{11}$	$U^{22}$	$U^{33}$	$U^{12}$	$U^{13}$	$U^{23}$
Br1	0.0722 (4)	0.0352 (2)	0.0829 (4)	0.0069 (2)	-0.0019 (3)	0.0149 (3)
C1	0.035 (2)	0.038 (2)	0.029 (2)	-0.0004 (18)	-0.0007 (18)	-0.0038 (17)
C2	0.036 (2)	0.053 (3)	0.032 (2)	-0.004 (2)	0.002 (2)	0.007 (2)
C3	0.046 (2)	0.034 (2)	0.048 (3)	-0.0034 (19)	0.000 (2)	0.0120 (19)
C4	0.039 (2)	0.0318 (19)	0.041 (2)	0.0016 (18)	-0.001 (2)	-0.0023 (17)
C5	0.037 (2)	0.0304 (19)	0.032 (2)	0.0017 (17)	0.0008 (19)	-0.0019 (16)
C6	0.030 (2)	0.0345 (19)	0.028 (2)	0.0024 (17)	-0.0034 (17)	0.0023 (15)
C7	0.034 (2)	0.038 (2)	0.029 (2)	0.0013 (17)	0.0027 (19)	0.0018 (17)
C8	0.029 (2)	0.0299 (18)	0.032 (2)	0.0004 (15)	0.0030 (17)	-0.0019 (16)
C9	0.039 (2)	0.0289 (18)	0.033 (2)	-0.0029 (17)	-0.0012 (19)	-0.0041 (16)
C10	0.033 (2)	0.038 (2)	0.037 (2)	0.0031 (17)	0.0000 (19)	-0.0019 (18)
C11	0.036 (2)	0.0315 (19)	0.050 (3)	0.0041 (17)	0.006 (2)	0.0035 (19)
C12	0.042 (2)	0.0299 (19)	0.048 (3)	-0.0017 (18)	-0.003 (2)	-0.0077 (18)
C13	0.033 (2)	0.037 (2)	0.036 (2)	-0.0005 (17)	-0.0029 (18)	-0.0080 (17)
Cl1	0.0858 (10)	0.0333 (5)	0.0652 (8)	-0.0007 (6)	0.0074 (8)	-0.0081 (6)
N1	0.0343 (18)	0.0305 (15)	0.0318 (18)	-0.0003 (14)	-0.0008 (15)	-0.0009 (14)
O1	0.069 (2)	0.0394 (16)	0.0342 (17)	0.0003 (16)	0.0096 (17)	-0.0042 (13)

Geometric parameters ( $\text{\AA}$ ,  $^\circ$ )

Br1—C11	1.884 (4)	C7—H7	0.9300
C1—O1	1.352 (5)	C8—C13	1.390 (5)
C1—C2	1.387 (6)	C8—C9	1.395 (6)
C1—C6	1.405 (6)	C8—N1	1.422 (5)
C2—C3	1.371 (6)	C9—C10	1.379 (5)
C2—H2	0.9300	C9—H9	0.9300
C3—C4	1.377 (6)	C10—C11	1.378 (6)
C3—H3	0.9300	C10—H10	0.9300
C4—C5	1.367 (6)	C11—C12	1.379 (7)
C4—C11	1.735 (4)	C12—C13	1.381 (6)
C5—C6	1.397 (5)	C12—H12	0.9300
C5—H5	0.9300	C13—H13	0.9300
C6—C7	1.450 (5)	O1—H1	0.8200
C7—N1	1.276 (5)		
O1—C1—C2	119.1 (4)	C13—C8—C9	119.5 (4)
O1—C1—C6	121.3 (4)	C13—C8—N1	118.6 (4)
C2—C1—C6	119.6 (4)	C9—C8—N1	121.7 (3)
C3—C2—C1	120.6 (4)	C10—C9—C8	119.8 (4)
C3—C2—H2	119.7	C10—C9—H9	120.1
C1—C2—H2	119.7	C8—C9—H9	120.1
C2—C3—C4	120.0 (4)	C11—C10—C9	120.0 (4)
C2—C3—H3	120.0	C11—C10—H10	120.0
C4—C3—H3	120.0	C9—C10—H10	120.0
C5—C4—C3	120.5 (4)	C10—C11—C12	120.8 (4)
C5—C4—C11	119.6 (3)	C10—C11—Br1	118.8 (3)
C3—C4—C11	119.9 (3)	C12—C11—Br1	120.4 (3)

C4—C5—C6	120.8 (4)	C11—C12—C13	119.6 (4)
C4—C5—H5	119.6	C11—C12—H12	120.2
C6—C5—H5	119.6	C13—C12—H12	120.2
C5—C6—C1	118.5 (4)	C12—C13—C8	120.3 (4)
C5—C6—C7	119.6 (4)	C12—C13—H13	119.9
C1—C6—C7	121.7 (4)	C8—C13—H13	119.9
N1—C7—C6	121.0 (4)	C7—N1—C8	120.2 (3)
N1—C7—H7	119.5	C1—O1—H1	109.5
C6—C7—H7	119.5		
O1—C1—C2—C3	179.9 (4)	C1—C6—C7—N1	5.4 (6)
C6—C1—C2—C3	0.9 (6)	C13—C8—C9—C10	0.6 (6)
C1—C2—C3—C4	0.3 (7)	N1—C8—C9—C10	176.0 (4)
C2—C3—C4—C5	-0.8 (7)	C8—C9—C10—C11	0.7 (6)
C2—C3—C4—C11	-178.6 (4)	C9—C10—C11—C12	-1.8 (6)
C3—C4—C5—C6	0.1 (6)	C9—C10—C11—Br1	179.7 (3)
C11—C4—C5—C6	178.0 (3)	C10—C11—C12—C13	1.4 (7)
C4—C5—C6—C1	1.1 (6)	Br1—C11—C12—C13	179.9 (3)
C4—C5—C6—C7	-174.2 (4)	C11—C12—C13—C8	0.0 (6)
O1—C1—C6—C5	179.5 (4)	C9—C8—C13—C12	-1.0 (6)
C2—C1—C6—C5	-1.6 (6)	N1—C8—C13—C12	-176.6 (4)
O1—C1—C6—C7	-5.3 (6)	C6—C7—N1—C8	-170.4 (3)
C2—C1—C6—C7	173.6 (4)	C13—C8—N1—C7	-148.7 (4)
C5—C6—C7—N1	-179.4 (4)	C9—C8—N1—C7	35.8 (6)

Hydrogen-bond geometry ( $\text{\AA}$ ,  $^\circ$ )

Cg1 and Cg2 are the centroids of the C1—C6 and C8—C13 benzene rings, respectively.

$D-H\cdots A$	$D-H$	$H\cdots A$	$D\cdots A$	$D-H\cdots A$
O1—H1 $\cdots$ N1	0.82	1.87	2.593 (4)	147
C2—H2 $\cdots$ Cg1 <sup>i</sup>	0.93	2.82	3.489 (5)	129
C5—H5 $\cdots$ Cg1 <sup>ii</sup>	0.93	2.85	3.513 (5)	129
C10—H10 $\cdots$ Cg2 <sup>iii</sup>	0.93	2.75	3.460 (5)	133
C13—H13 $\cdots$ Cg2 <sup>iv</sup>	0.93	2.78	3.473 (5)	132

Symmetry codes: (i)  $-x+1/2, y, z-1/2$ ; (ii)  $-x+3/2, y, z-3/2$ ; (iii)  $-x+1/2, y, z-3/2$ ; (iv)  $-x+3/2, y, z-1/2$ .