

Effect of Monomer Solubility on the Evolution of Copolymer Morphology during Polymerization-Induced Self-Assembly in Aqueous Solution

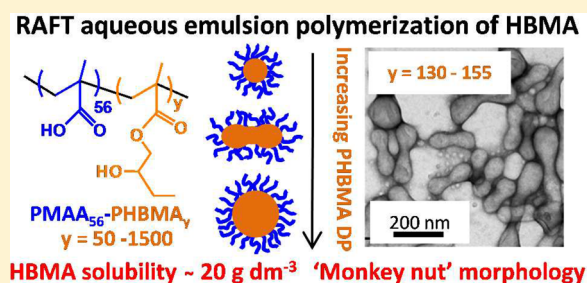
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Supporting Information

ABSTRACT: Polymerization-induced self-assembly (PISA) has become a widely used technique for the rational design of diblock copolymer nano-objects in concentrated aqueous solution. Depending on the specific PISA formulation, reversible addition–fragmentation chain transfer (RAFT) aqueous dispersion polymerization typically provides straightforward access to either spheres, worms, or vesicles. In contrast, RAFT aqueous emulsion polymerization formulations often lead to just kinetically-trapped spheres. This limitation is currently not understood, and only a few empirical exceptions have been reported in the literature. In the present work, the effect of monomer solubility on copolymer morphology is explored for an aqueous PISA formulation. Using 2-hydroxybutyl methacrylate (aqueous solubility = 20 g dm⁻³ at 70 °C) instead of benzyl methacrylate (0.40 g dm⁻³ at 70 °C) for the core-forming block allows access to an unusual “monkey nut” copolymer morphology over a relatively narrow range of target degrees of polymerization when using a poly(methacrylic acid) RAFT agent at pH 5. These new anisotropic nanoparticles have been characterized by transmission electron microscopy, dynamic light scattering, aqueous electrophoresis, shear-induced polarized light imaging (SIPLI), and small-angle X-ray scattering.



INTRODUCTION

In recent years, polymerization-induced self-assembly (PISA) has become a widely recognized route for the synthesis of many types of diblock copolymer nano-objects.^{1–5} Compared to post-polymerization processing techniques (solvent exchange, film rehydration, or pH switch), PISA is much more efficient and can be performed at relatively high solids (10–50% w/w).^{3,6–8} This approach involves growth of an insoluble block from a soluble homopolymer in a suitable solvent to give well-defined sterically stabilized diblock copolymer nanoparticles. For example, reversible addition-fragmentation chain transfer (RAFT) aqueous dispersion polymerization involves polymerization of a water-miscible monomer such as 2-hydroxypropyl methacrylate (HPMA) from a water-soluble stabilizer, e.g. poly(glycerol monomethacrylate).^{9,10} Such formulations enable the production of various copolymer morphologies such as spheres, worms or vesicles.^{11–19} RAFT aqueous emulsion polymerization has similarly received significant attention.^{2,6,7,20–24} In this case, a water-immiscible monomer is used to produce the hydrophobic core-forming block, but according to many literature reports only kinetically-trapped spheres can be obtained.^{6,7,24–31} Exceptionally, Charleux and co-workers reported the synthesis of diblock copolymer worms (described as “nanofibers”) and vesicles, as well as spheres.^{23,32–35} Recent empirical experiments have undoubtedly

provided some useful insights,³⁶ but the critical synthesis parameters that determine whether only kinetically-trapped spheres are obtained or the full range of morphologies are observed have not yet been established. In this context, Truong et al. recently synthesized novel “filomicelle nanomaterials” directly in water by employing RAFT aqueous emulsion polymerization followed by temperature-induced morphological transition. Morphological transitions from spherical micelles to filomicelles (worms) and/or vesicles were observed on cooling in the presence of additional monomer, which apparently acts as a plasticizer for the frustrated core-forming block.³⁷ However, this approach does not seem to be particularly attractive from a commercial perspective, unless the additional monomer can be polymerized.

The present work explores the effect of monomer solubility on copolymer morphology. As noted above, water-miscible monomers such as HPMA (aqueous solubility ~100 g dm⁻³ at 70 °C) are required for RAFT aqueous dispersion polymerization, whereas water-immiscible monomers such as benzyl methacrylate (BzMA; aqueous solubility ~0.40 g dm⁻³ at 70 °C) are required for RAFT aqueous emulsion polymerization.

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Herein we utilize 2-hydroxybutyl methacrylate (HBMA) as a monomer of *intermediate* aqueous solubility ($\sim 20 \text{ g dm}^{-3}$ at 70°C) that has been previously reported to undergo RAFT aqueous emulsion polymerization.³⁸ The key question to be addressed is whether such formulations allow access to any copolymer morphologies other than kinetically-trapped spheres.

EXPERIMENTAL SECTION

Materials. Methacrylic acid (MAA), 2-hydroxybutyl methacrylate (HBMA; actually a 1:1 molar ratio of 2- and 4-isomers as judged by ^1H NMR spectroscopy¹⁵), and 4,4'-azobis(4-cyanovaleric acid) (ACVA) were purchased from Sigma-Aldrich U.K. and used as received unless otherwise specified. Deionized water was used in all experiments. 4-cyano-4-(2-phenylethanesulfanylthiocarbonyl)sulfanylpentanoic acid (PETTC) was prepared as described previously.³⁹ The trimethylsilyldiazomethane solution (2.0 M in diethyl ether), THF (HPLC, $\geq 99.9\%$), and glacial acetic acid ($\geq 99.85\%$) used for the preparation and analysis of the methylated diblock copolymers were also purchased from Sigma-Aldrich U.K. Methanol- d_4 , dimethyl sulfoxide- d_6 , and dimethylformamide- d_7 used for ^1H NMR spectroscopy were purchased from Goss Scientific Instruments Ltd. (Cheshire, U.K.). All other solvents were purchased from Sigma-Aldrich U.K. or VWR Chemicals.

Preparation of Poly(methacrylic acid) (PMAA) Macro-CTA Agent. PETTC RAFT agent (3.169 g, 0.0093 mol), MAA (45.00 g, 0.5227 mol), ACVA (0.523 g, 0.001 mol; CTA/initiator molar ratio = 5.0), and ethanol (73.04 g, 40% w/w) were weighed into a 500 mL round-bottom flask and degassed with nitrogen for 30 min in an ice bath. The reaction solution was then heated for 3 h at 70°C in a preheated oil bath. The resulting macro-CTA was then purified by precipitation into diethyl ether (1.5 L). The polymer was collected by filtration and redissolved in the minimum amount of ethanol, before a second precipitation step. The polymer was then collected and redissolved in the minimum amount of water for isolation by lyophilization. The mean degree of polymerization was calculated to be 56 for this macro-CTA by ^1H NMR. This synthesis was also performed using (4-cyano-4-(phenylcarbonothioylthio)pentanoic acid) (CPCP) as the RAFT agent.

RAFT Polymerization of HBMA in Water. A typical protocol for the synthesis of PMAA₅₆-PHBMA₅₀₀ was as follows: PMAA₅₆ macro-CTA (0.0489 g, 0.0094 mmol), ACVA (0.6 mg, 0.0019 mmol, macro-CTA/initiator molar ratio = 5.0), and water (3.20 g, 20% w/w) were weighed into a 14 mL vial. The pH was adjusted to pH 5 using 1 M NaOH. HBMA monomer (0.7500 g, 4.70 mmol) was then added, and the reaction vial was sealed and purged for 30 min before being placed in a preheated oil bath at 70°C for 18 h.

Purification of HBMA Monomer. As-received HBMA (3.0 g) was dissolved in water (300 g). This aqueous monomer solution was extracted using *n*-hexane to remove the dimethacrylate impurity. The aqueous monomer solution was then salted with NaCl (250 g/L), and HBMA was removed from the aqueous phase by extraction with diethyl ether. MgSO_4 was added to remove traces of water from the ether layer. Hydroquinone (0.1%) was added to prevent thermal polymerization prior to removal of the solvent by distillation under reduced pressure to afford purified HBMA monomer.

Copolymer Characterization. ^1H NMR Spectroscopy. All ^1H NMR spectra were recorded using a 400 MHz Bruker Advance-400 spectrometer using either methanol- d_4 , dimethyl sulfoxide- d_6 , or dimethylformamide- d_7 .

Exhaustive Methylation of Copolymers for GPC Analysis. Prior to gel permeation chromatography analysis, all copolymers were modified by exhaustive methylation of the carboxylic acid groups in the PMAA block. Excess trimethylsilyldiazomethane was added dropwise to a solution of copolymer (20 mg) in THF (2.0 mL), until the yellow color persisted. This reaction solution was then stirred overnight until all THF had evaporated. Degrees of methylation of the PMAA block were determined by ^1H NMR spectroscopy.

Gel Permeation Chromatography (GPC). THF GPC at 60°C was used to determine the molecular weights and dispersities of the modified copolymers. The GPC setup consisted of two 5 μM Mixed C columns connected to a WellChrom K-2301 refractive index detector. The mobile phase was HPLC-grade THF containing 1.0% glacial acetic acid and 0.05% w/v butylhydroxytoluene (BHT) at a flow rate of 1.0 mL min^{-1} . Molecular weights were calculated with respect to a series of near-monodisperse poly(methyl methacrylate) standards.

Dynamic Light Scattering (DLS). Aqueous copolymer dispersions (0.20% w/w) in disposable plastic cuvettes were analyzed using a Malvern Zetasizer NanoZS instrument. The mean hydrodynamic particle diameter was averaged over three consecutive runs.

Aqueous Electrophoresis. Measurements were performed using a Malvern Zetasizer instrument on dilute (0.20% w/w) copolymer dispersions containing background KCl (1 mM). The solution pH was adjusted by addition of either NaOH or HCl.

Transmission Electron Microscopy (TEM). One droplet (10 μL) of a dilute copolymer dispersion (0.20% w/w) was deposited onto a carbon-coated copper grid. The grid was then stained with 10 μL uranyl formate for 10 s and dried using a vacuum hose. TEM images were then obtained using a Philips CM100 instrument operating at 100 kV and equipped with a Gatan 1 k CCD camera.

Reverse-Phase High-Performance Liquid Chromatography (HPLC). The level of dimethacrylate impurity in the HBMA monomer was quantified by HPLC. The experimental setup consisted of an autosampler (Varian model 410), a solvent delivery module (Varian Module 230), a UV detector (Varian model 310), and an Zorbax Eclipse Plus C18, 3.5 μm , $4.6 \times 100 \text{ mm}$ HPLC column. HBMA (5.0 mg) was weighed into an autosampler vial and dissolved in acetonitrile (1.0 mL). The eluent was gradually varied from an initial composition of 5% acetonitrile in water to 95% acetonitrile in water after 15–20 min. The UV detector was set to a wavelength of 210 nm.

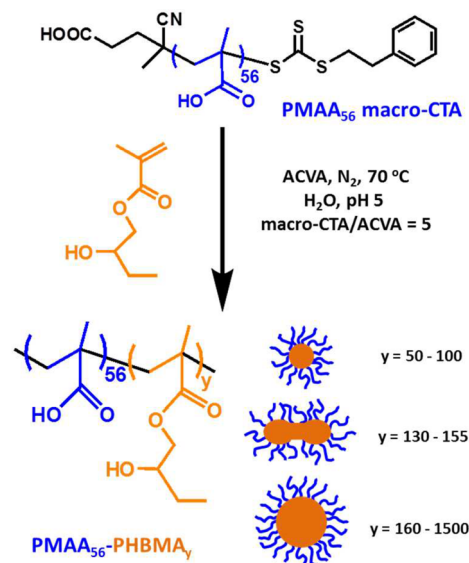
Shear-Induced Polarized Light Imaging (SIPLI). Shear alignment experiments were conducted using a mechano-optical rheometer (Anton Paar Physica MCR301 with SIPLI attachment). Measurements were performed using a plate–plate geometry composed of a 25 mm polished steel plate and a fused quartz plate connected to a variable temperature Peltier system. The gap between plates was set at 0.50 mm for all experiments. An additional Peltier hood was used to ensure good control of the sample temperature. Sample illumination was achieved using an Edmund Optics 150 W MI-150 high-intensity fiber-optic white light source. The polarizer and analyzer axes were crossed at 90° in order to obtain polarized light images (PLIs), which were recorded using a color CCD camera (Lumenera Lu165c).

Small-Angle X-ray Scattering (SAXS). SAXS patterns for PMAA₅₆-PHBMA_y ($y = 150, 300, \text{ and } 1000$) were recorded at a synchrotron source (ESRF, station ID02, Grenoble, France) using monochromatic X-ray radiation (wavelength $\lambda = 0.0995 \text{ nm}$, with q ranging from 0.004 to 2.5 nm^{-1} , where $q = 4\pi \sin \theta / \lambda$ is the length of the scattering vector and θ is one-half of the scattering angle) and a Rayonix MX-170HS Kodak CCD detector. Measurements were conducted on 1.0% w/w aqueous dispersions at pH 5 using glass capillaries of 2.0 mm diameter. X-ray scattering data were reduced using standard routines from the beamline and were further analyzed using Irena SAS macros for Igor Pro.⁴⁰ The SAXS pattern for PMAA₅₆-PHBMA₅₀ was obtained using a Bruker AXS Nanostar laboratory instrument modified with a microfocus X-ray tube (GeniX3D, Xenocs) and motorized scatterless slits for the beam collimation (camera length = 1.46 m, Cu $K\alpha$ radiation, and HiSTAR multiwire gas detector). In this case the SAXS pattern was recorded for a 1.0% w/w aqueous dispersion at pH 5 over a q range of $0.08 \text{ nm}^{-1} < q < 1.6 \text{ nm}^{-1}$ using a glass capillary of 2.0 mm diameter and an exposure time of 1.0 h. Raw SAXS data were reduced using Nika macros for Igor Pro written by J. Ilavsky. All SAXS patterns were analyzed (background subtraction, data modeling and fitting) using Irena SAS macros for Igor Pro.⁴⁰

RESULTS AND DISCUSSION

A PMAA₅₆ macromolecular chain transfer agent (macro-CTA) was chain-extended with HBMA via RAFT polymerization at

Scheme 1. Synthesis of Poly(methacrylic acid)–Poly(2-hydroxybutyl methacrylate) (PMAA–PHBMA) Diblock Copolymer Nanoparticles via RAFT Polymerization in Aqueous Solution^a



^aIncreasing the target degree of polymerization of the PHBMA core-forming block alters both the particle diameter and particle morphology.

70 °C conducted in aqueous solution at pH 5 (see Scheme 1). The target degree of polymerization (DP) for the structure-directing PHBMA block was varied between 50 and 1500. All polymerizations proceeded to high conversion (>96%) as judged by ¹H NMR spectroscopy studies in DMF-*d*₇. Dynamic light scattering (DLS) and transmission electron microscopy (TEM) studies were conducted to determine the copolymer morphology (Figure 1). For target PHBMA DPs of 50–145, the PMAA₅₆–PHBMA_y diblock copolymer chains self-assembled to form well-defined spheres of 80–175 nm diameter. However, TEM studies indicated that a new “monkey nut” morphology could be obtained over a rather narrow range of *y* values (*y* = 150 or 155). These “monkey nuts” are approximately 100–800 nm in length, with widths varying from 25 to 125 nm; thus, the mean length/width ratio (or particle anisotropy) is approximately four. This unusual non-spherical morphology clearly demonstrates that using a monomer of intermediate aqueous solubility such as HBMA allows access to morphologies other than kinetically-trapped spheres. However, only relatively large spherical particles of 200–400 nm diameter were formed when targeting higher PHBMA DPs (up to *y* = 1500).

Aqueous electrophoresis was used to assess the mobility and zeta potential of these diblock copolymer nano-objects. The effect of varying pH on the apparent sphere-equivalent particle diameter and zeta potential of the PMAA₅₆–PHBMA₁₅₀ “monkey nut” nanoparticles was evaluated by DLS and aqueous electrophoresis, respectively (see Figure 2). Between pH 10 and pH 5.5, the PMAA stabilizer chains were highly ionized, leading to negative zeta potentials ranging from –50 to –45 mV. The PMAA stabilizer chains remained highly anionic over this pH range, with only a modest reduction in particle diameter from 200 to 165 nm being observed (DLS polydispersities ranged from 0.03 to 0.09). This is consistent

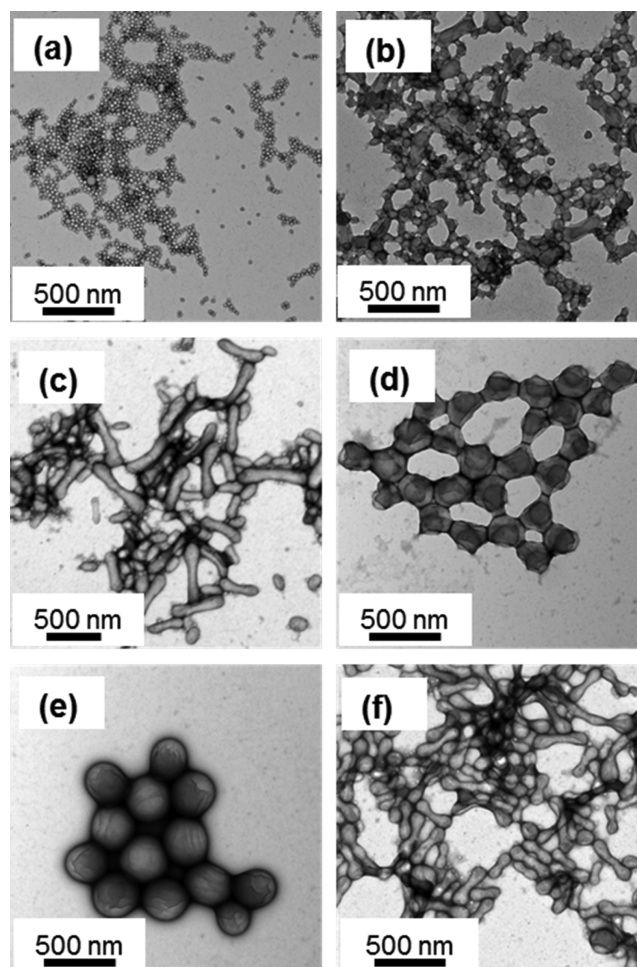


Figure 1. TEM images obtained for PMAA₅₆–PHBMA_y diblock copolymer nanoparticles prepared at 20% w/w solids via RAFT aqueous emulsion polymerization at 70 °C. Polymerization-induced self-assembly (PISA) leads to the formation of small spherical nanoparticles for (a) PMAA₅₆–PHBMA₃₀ and (b) PMAA₅₆–PHBMA₁₃₀. Synthesis of (c) PMAA₅₆–PHBMA₁₅₀ produces a distinctive “monkey nut” morphology. However, targeting a mean degree of polymerization for the PHBMA block of either (d) 300 or (e) 1000 only leads to the formation of relatively large spheres. (f) PISA syntheses conducted using a purified batch of HBMA monomer also produce a “monkey nut” morphology when targeting PMAA₅₆–PHBMA₁₅₀, clearly indicating that this unusual morphology is not simply the result of *in situ* cross-linking as a result of the dimethacrylate impurity in the HBMA monomer.

with ionized PMAA chains acting as a polyelectrolytic stabilizer block, conferring electrosteric stabilization during the PISA synthesis. Between pH 5.5 and pH 3.5, the zeta potential is gradually lowered to –20 mV as the PMAA chains become progressively more protonated. A concomitant reduction in apparent hydrodynamic particle diameter to 150 nm occurs as the PMAA chains start to collapse. However, an apparent particle diameter of 5 μm is observed by DLS at approximately pH 2.5. This is the result of flocculation of the PMAA₅₆–PHBMA₁₅₀ nanoparticles because the near-neutral PMAA stabilizer chains no longer confer effective electrosteric stabilization at this pH. Such aggregation proved to be reversible: on raising the solution pH from pH 3.5 to pH 10. The PMAA chains become ionized again, and approximately the original sphere-equivalent nanoparticle diameter was

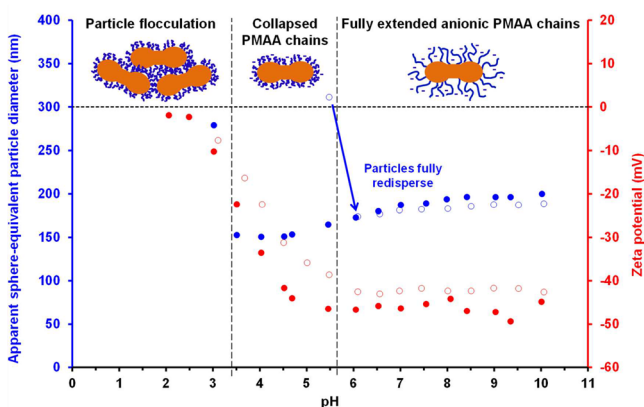


Figure 2. Variation in apparent sphere-equivalent particle diameter (as judged by DLS) and zeta potential with pH curves for PMAA₅₆–PHBMA₁₅₀ “monkey nut” nanoparticles prepared using purified HBMA monomer. Filled spheres (●) indicate titration from pH 10 to pH 2. Open spheres (○) indicate titration from pH 2 to pH 10. Particle flocculation is observed below pH 3.5 due to the loss of electrosteric stabilization as the PMAA chains become less anionic. This flocculation is reversible on addition of NaOH.

obtained, with a corresponding zeta potential of around -40 mV.

In principle, the molecular weight distributions of such PMAA₅₆–PHBMA_y diblock copolymers can be assessed by gel permeation chromatography (GPC). However, in practice the MAA residues require exhaustive methylation to prevent adsorption onto the GPC column. Unfortunately, the original methylated PMAA₅₆–PHBMA_y diblock copolymers proved to be insoluble in both THF and DMF, making GPC analysis impossible. This is believed to be the result of extensive cross-linking caused by the ~ 4.4 mol % dimethacrylate impurity in the HBMA monomer. Similar problems have been reported for PISA syntheses involving HPMA.^{10,13} In order to address this technical problem, the HBMA monomer was purified prior to the preparation of a second series of PMAA₅₆–PHBMA_y diblock copolymers. Moreover, analysis of such diblock copolymer nano-objects should establish whether the unusual “monkey nut” morphology is merely an artifact caused by *in situ* cross-linking. In this context, it is worth noting that Sugihara and co-workers reported a “lumpy rod” morphology for the synthesis of cross-linked nanoparticles prepared via RAFT aqueous dispersion copolymerization of HPMA with EGDMA when targeting more than six EGDMA units per copolymer chain.⁴¹ Thus an aqueous solution of the as-supplied HBMA monomer was extracted using *n*-hexane to remove the dimethacrylate impurity.⁴² The purified HBMA monomer was analyzed by reverse-phase HPLC, which indicated approximately 87% removal of the original dimethacrylate impurity, leaving around 0.57 mol % dimethacrylate still present. A series of PMAA₅₆–PHBMA_y diblock copolymers (targeting $y = 130$ – 300) were then prepared using this purified HBMA monomer. The MAA residues of the diblock copolymer chains were exhaustively methylated using excess trimethylsilyldiazomethane and proved to be fully soluble in a THF eluent containing 1.0% glacial acetic acid,⁴³ which indicates a substantial reduction in the degree of cross-linking. The molecular weight of the diblock copolymer chains increased as the target PHBMA DP was varied from 130 to 300, but dispersities ranged from 1.18 to 6.13, which suggests substantial branching (see Figure S1, Supporting Information).^{44,45} TEM

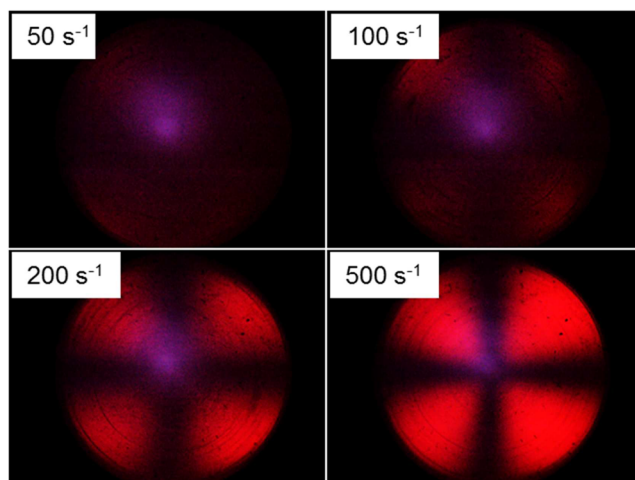


Figure 3. Shear-Induced Polarized Light Images (SIPLIs) obtained for a 20% w/w aqueous dispersion of PMAA₅₆–PHBMA₁₅₀ “monkey nut” nanoparticles at maximum shear rates of 50, 100, 200, and 500 s^{-1} . A Maltese cross is observed above a critical shear rate of 40 s^{-1} , indicating shear-induced alignment. Thus the mean relaxation time of the “monkey nut” nanoparticles corresponds to approximately 25 ms.

analysis of this second series of PMAA₅₆–PHBMA_y nano-objects prepared using purified HBMA monomer confirmed that a “monkey nut” copolymer morphology could still be obtained. Thus, such nano-objects do not appear to be an artifact caused by cross-linking. Moreover, the “monkey nut” morphology is observed for PHBMA DPs of 130–155, which is somewhat a somewhat broader range than that obtained when using the as-received HBMA monomer.

PMAA₅₆–PHBMA₁₅₀ “monkey nut” nanoparticles prepared using purified HBMA monomer were analyzed using the shear-induced polarized light imaging (SIPLI) technique.^{46–49} It is well-known that anisotropic nanoparticles can be aligned when subjected to an applied shear.^{50,51} Above a certain critical shear rate, alignment in the direction of flow leads to shear-thinning behavior and the observation of birefringence. In a SIPLI experiment, linearly polarized white light is directed through a transparent quartz plate on which an aqueous dispersion of PMAA₅₆–PHBMA₁₅₀ “monkey nuts” at 20% w/w solids is placed. After transmission through the dispersion, the light is reflected by a polished steel plate and then analyzed at 90° to the plane of polarization using a CCD camera. Because the reflected light is analyzed at 90° to the incident light, only rotated light is detected. Particle alignment leads to the observation of a characteristic Maltese cross pattern. The PMAA₅₆–PHBMA₁₅₀ “monkey nut” nanoparticles were subjected to maximum shear rates ranging from 50 to 500 s^{-1} (Figure 3). There is a shear rate gradient across the polished steel plate from its center to the periphery, with the maximum shear rate being obtained at the plate edge. A characteristic Maltese cross pattern was observed at maximum shear rates of either 200 or 500 s^{-1} , indicating alignment of anisotropic nanoparticles. The critical shear rate for nanoparticle alignment can be calculated from the image at a maximum shear rate of 100 s^{-1} , where a partial Maltese cross pattern is obtained with a dark circle in the center. The critical shear rate under these conditions is 40 s^{-1} , which corresponds to a mean relaxation time of approximately 25 ms. This represents the time scale required to produce an isotropic dispersion after cessation of the applied shear.

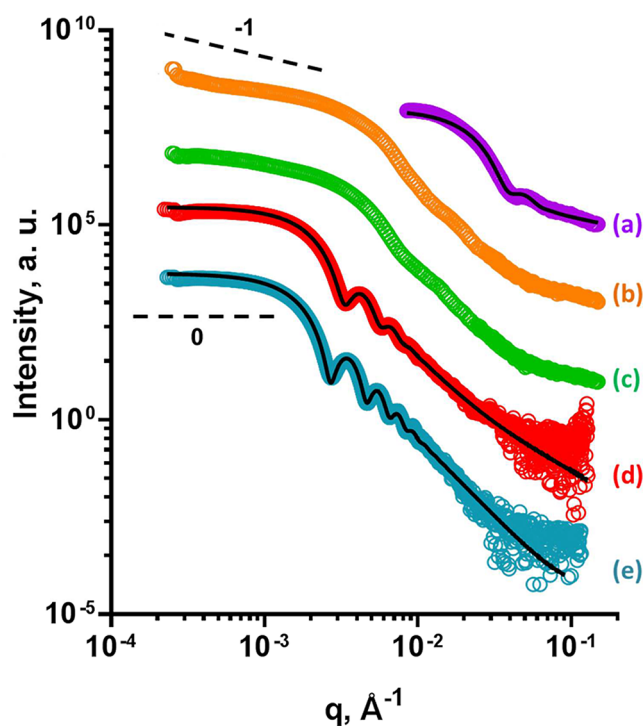


Figure 4. SAXS patterns obtained from 1.0% w/w aqueous dispersions of PMAA₅₆-PHBMA_y diblock copolymer nano-objects at pH 5: (a) PMAA₅₆-PHBMA₅₀ spheres, (b) PMAA₅₆-PHBMA₁₅₀ “monkey nut” nanoparticles (prepared using purified HBMA monomer), (c) PMAA₅₆-PHBMA₁₅₀ “monkey nut” nanoparticles (prepared using as-received HBMA monomer), (d) PMAA₅₆-PHBMA₃₀₀ spheres, and (e) PMAA₅₆-PHBMA₁₀₀₀ spheres.

Small-angle X-ray scattering (SAXS; see Figure 4) was used to confirm the copolymer morphologies indicated by TEM studies. To a good first approximation, the copolymer morphology is indicated by the gradient in the low q regime. Spherical micelles are characterized by a gradient of zero and rigid rods possess a gradient of negative unity.⁵² Although they exhibit considerable flexibility, highly anisotropic diblock copolymer worms prepared via PISA behave more or less like rigid rods in terms of their SAXS patterns.^{53,54} Inspecting Figure 4, the $I(q)$ vs q scattering pattern recorded for a 1.0% w/w dispersion of PMAA₅₆-PHBMA₅₀ diblock copolymer nanoparticles can be satisfactorily fitted to a previously reported spherical micelle model, with a volume-average core diameter of 19 ± 3 nm.⁵⁵ The same spherical micelle model also provided good fits to the scattering patterns obtained for the PMAA₅₆-PHBMA₃₀₀ and PMAA₅₆-PHBMA₁₀₀₀ nanoparticles. In each case, the gradient at low q of approximately zero confirms the spherical morphology indicated by TEM studies, with SAXS volume-average diameters estimated to be 262 ± 26 and 330 ± 22 nm, respectively. These analyses enable us to reject our initial hypothesis that the latter nano-objects might be thick-walled vesicles, not least because there is no evidence for any membrane structure. Moreover, the presence of multiple fringes in these latter two scattering patterns suggests relatively narrow size distributions in each case. In contrast, the scattering patterns recorded for the PMAA₅₆-PHBMA₁₅₀ “monkey nut” nanoparticles (synthesized with either the as-received or the purified HBMA monomer) cannot be fitted using the spherical model. These patterns have low q gradients of -0.82 and -0.71 , respectively, confirming that these

nanoparticles possess significant anisotropic character (as suggested by TEM analysis). In addition, the lack of a well-defined local minimum at high q suggests that these “monkey nut” nanoparticles are relatively polydisperse in terms of their mean widths. Although not yet fully analyzed, these preliminary SAXS data are important because they are much more statistically robust than TEM analyses. They confirm a unique evolution in copolymer morphology for this PMAA₅₆-PHBMA_y RAFT aqueous emulsion formulation, from small spheres to “monkey nuts” to large spheres with increasing y values. We hypothesize that a higher aqueous monomer solubility facilitates more effective plasticization of the core-forming block on the time scale of the polymerization, which in turn facilitates the stochastic 1D fusion of the growing monomer-swollen spheres to form the “monkey nut” copolymer morphology, rather than kinetically-trapped spheres. However, a detailed mechanistic explanation for these morphological observations will clearly require further studies that are beyond the scope of the present work.

In summary, the RAFT aqueous emulsion polymerization of HBMA at pH 5 using a PMAA₅₆ macro-CTA leads to the formation of elongated nanoparticles with a highly unusual “monkey nut” morphology over a relatively narrow range of core-forming block DPs. This nanoparticle anisotropy is confirmed by SAXS analysis and is sufficient to enable alignment under shear, as indicated by shear-induced polarized light imaging studies. This suggests that the aqueous solubility of the monomer can play an important role in determining the copolymer morphology obtained during aqueous PISA syntheses. In future work, we plan to fit the SAXS scattering patterns obtained for these “monkey nut” nanoparticles using an appropriate new analytical model.

■ ASSOCIATED CONTENT

Supporting Information

The Supporting Information is available free of charge on the ACS Publications website at DOI: 10.1021/acs.macromol.6b02309.

THF GPC curves obtained when using the purified HBMA monomer, summary table for SAXS structural parameters, additional polymerization kinetics data, and TEM images of diblock copolymers made using a shorter PMAA₂₉ macro-CTA (PDF)

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Notes

The authors declare no competing financial interest.

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