

Bis(μ -di-2-pyridyl disulfide- $\kappa^3N,S:N'$)di- μ_3 -iodido-di- μ_2 -iodido-tetracopper(I)

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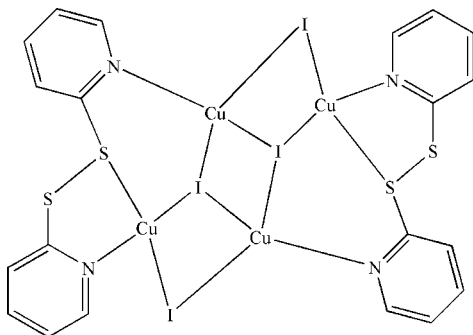
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Key indicators: single-crystal X-ray study; $T = 223$ K; mean $\sigma(C-C) = 0.010$ Å; R factor = 0.048; wR factor = 0.087; data-to-parameter ratio = 20.9.

In the centrosymmetric tetranuclear title compound, $[Cu_4I_4(C_{10}H_8N_2S_2)_2]$, there are two different Cu^I atoms with tetrahedral coordination geometries. One is chelated by a pyridine N atom and an S-donor from one di-2-pyridyl disulfide ligand and coordinated by two I atoms, while the second Cu^I atom is coordinated by a pyridine-N and three I atoms. Iodine bridges between the Cu^I atoms form a tetranuclear structure.

Related literature

For the structures and luminescence properties of Cu(I) complexes, see: Caradoc-Davies & Hanton (2003); Ford *et al.* (1999); Rath *et al.* (1986); Song *et al.* (2003); Song, Sun & Yang (2011); Song, Sun, Yang & Yang (2011); Su *et al.* (1997). For metal complexes with di-2-pyridyl disulfide, see: Bell *et al.* (2000); Delgado *et al.* (2007); Kadooka *et al.* (1976); Niu *et al.* (2007); Wu *et al.* (2011).



Experimental

Crystal data

$[Cu_4I_4(C_{10}H_8N_2S_2)_2]$	$V = 3006$ (3) Å ³
$M_r = 1202.37$	$Z = 4$
Orthorhombic, $Pbca$	Mo $K\alpha$ radiation
$a = 10.460$ (6) Å	$\mu = 7.20$ mm ⁻¹
$b = 14.434$ (8) Å	$T = 223$ K
$c = 19.908$ (12) Å	$0.38 \times 0.11 \times 0.10$ mm

Data collection

Rigaku Saturn diffractometer	10455 measured reflections
Absorption correction: multi-scan (<i>REQAB</i> ; Jacobson, 1998)	3410 independent reflections
$T_{min} = 0.171$, $T_{max} = 0.533$	2944 reflections with $I > 2\sigma(I)$
	$R_{int} = 0.041$

Refinement

$R[F^2 > 2\sigma(F^2)] = 0.048$	163 parameters
$wR(F^2) = 0.087$	H-atom parameters constrained
$S = 1.16$	$\Delta\rho_{max} = 0.73$ e Å ⁻³
3410 reflections	$\Delta\rho_{min} = -0.68$ e Å ⁻³

Data collection: *CrystalClear* (Rigaku, 2001); cell refinement: *CrystalClear*; data reduction: *CrystalStructure* (Rigaku, 2001); program(s) used to solve structure: *SHELXS97* (Sheldrick, 2008); program(s) used to refine structure: *SHELXL97* (Sheldrick, 2008); molecular graphics: *SHELXTL* (Sheldrick, 2008); software used to prepare material for publication: *SHELXTL*.

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Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: MW2029).

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Acta Cryst. (2012). E68, m18 [doi:10.1107/S160053681105152X]

Bis(μ -di-2-pyridyl disulfide- $\kappa^3N,S:N'$)di- μ_3 -iodido-di- μ_2 -iodido-tetracopper(I)

Y.-H. Wang and X.-H. Zhu

Comment

Polynuclear Cu^I compounds with luminescent properties have been designed and synthesized by reaction of Cu^I cations and suitable bridging ligands (Ford *et al.*, 1999). Flexible heterocyclic ligands containing nitrogen donors can form metal complexes with considerable structural diversity and reports of Cu^I complexes with such ligands have increased in recent years (Caradoc-Davies & Hanton, 2003; Rath *et al.*, 1986; Song *et al.*, 2003; Song, Sun & Yang, 2011; Song, Sun, Yang & Yang, 2011; Song, Sun, & Yang, 2011; Su *et al.*, 1997). di-2-pyridyl disulfide is an interesting heterocyclic ligand. A number of metal complexes of it and related ligands have been reported (Bell *et al.*, 2000; Delgado *et al.*, 2007; Kadooka *et al.*, 1976; Niu *et al.*, 2007; Wu *et al.*, 2011). The present study details the structure of a CuI adduct of di-2-pyridyl disulfide.

The title compound, (I), is a tetranuclear complex, [Cu₄I₄(L)₂] having crystallographically-imposed centrosymmetry (Fig. 1). In the unique ligand, S(2) and N(1) chelate to Cu(1) while N(2) coordinates to Cu(2A). Doubly-bridging I2 and triply-bridging I1 serve to complete the formation of the tetranuclear core in which each copper adopts approximately tetrahedral geometry. Cu(1) is coordinated by a pyridine N atom and an S donor from one L ligand and two iodine atoms while Cu(2) is coordinated by a pyridine N atom from the second L ligand and three iodine atoms. The bond angles around Cu(2) range from 104.38 (4) to 114.42 (5)° and are normal for a slightly distorted tetrahedral geometry while Cu(1) exists in a more distorted tetrahedral environment with the angles around Cu(1) in the range 88.58 (17) to 122.55 (7)°. All the Cu—N and Cu—S bond distances are comparable with those of other tetrameric clusters (Rath *et al.*, 1986; Song *et al.*, 2003; Su *et al.*, 1997). The Cu—I bond lengths show significant differences depending on the Cu(I) environment. Cu(2) bonds to three iodine atoms with Cu—I bond distances ranging from 2.6864 (15) to 2.7623 (17) Å, while Cu(1) bonds to two iodine atoms with equivalent shorter Cu—I bond distances averaging 2.5764 (14) Å. The Cu...Cu distance of 2.6648 (18) Å is shorter than the sum of their van der Waals radii (2.83 Å), indicating the existence of Cu...Cu interactions. Such short Cu...Cu separations have been observed in some complexes such as [Cu₄I₄(quin)₄] [2.582 (10)Å](Rath *et al.*, 1986), [Cu₄I₄(MPTQ)₂] [2.607 (1)Å] (Su *et al.*, 1997), [Cu₄I₄(C₁₉H₃₆N₂S₂)₂]_n [2.734 (2)Å] (Song, Sun & Yang, 2011), [Cu₄I₄(C₁₅H₁₂N₂S)₂] [2.743 (1)Å] (Song *et al.*, 2003) and [CuI(bbbm)]_n [2.7683 (9)Å] (Niu *et al.*, 2007).

The tetrameric Cu₄I₄ core adopts a distorted chair-like structure which is defined by three Cu—I edges from different Cu₂I₂ moieties: the strictly planar Cu(2)—I(1A)—Cu(2A)—I(1) unit and the non-planar Cu(1)—I(2)—Cu(2)—I(1) and symmetry-related Cu(1A)—I(A)—Cu(2A)—I(1A) units. Similar results were also found in related tetranuclear complexes (Rath *et al.*, 1986; Song *et al.*, 2003; Su *et al.*, 1997).

Experimental

The ligand di-2-pyridyl disulfide (22.0 mg, 0.1 mmol) in CHCl₃ (3 ml) was added to CuI (17.8 mg, 0.1 mmol) dissolved in MeCN (5 ml). The reaction mixture was kept in the dark and allowed to evaporate slowly. Yellow block-shaped single

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crystals suitable for X-ray analysis were obtained in 58% yield. Analysis found: C 20.39, H 1.40, N 4.84%; calculated for $C_{20}H_{16}Cu_4I_4N_4S_4$: C 19.98, H 1.34, N 4.66%.

Refinement

H atoms were placed in calculated positions and included as riding contributions with C—H distances of 0.94 Å (aromatic H) and with $U_{iso} = 1.2U_{eq}(C)$.

Figures

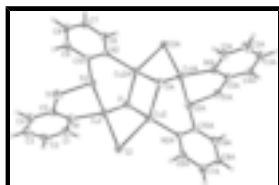


Fig. 1. View of (I) shown with 30% probability displacement ellipsoids and small spheres for the H atoms.

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Crystal data

$[Cu_4I_4(C_{10}H_8N_2S_2)_2]$

$M_r = 1202.37$

Orthorhombic, *Pbca*

Hall symbol: -P 2ac 2ab

$a = 10.460$ (6) Å

$b = 14.434$ (8) Å

$c = 19.908$ (12) Å

$V = 3006$ (3) Å³

$Z = 4$

$F(000) = 2224$

$D_x = 2.657$ Mg m⁻³

Mo $K\alpha$ radiation, $\lambda = 0.71075$ Å

Cell parameters from 8826 reflections

$\theta = 3.0$ – 27.5°

$\mu = 7.20$ mm⁻¹

$T = 223$ K

Block, yellow

$0.38 \times 0.11 \times 0.10$ mm

Data collection

Rigaku Saturn
diffractometer

Radiation source: fine-focus sealed tube
graphite

Detector resolution: 14.63 pixels mm⁻¹

ω scans

Absorption correction: multi-scan
(*REQAB*; Jacobson, 1998)

$T_{min} = 0.171$, $T_{max} = 0.533$

10455 measured reflections

3410 independent reflections

2944 reflections with $I > 2\sigma(I)$

$R_{int} = 0.041$

$\theta_{max} = 27.5^\circ$, $\theta_{min} = 3.0^\circ$

$h = -9 \rightarrow 13$

$k = -11 \rightarrow 18$

$l = -23 \rightarrow 25$

Refinement

Refinement on F^2	Primary atom site location: structure-invariant direct methods
Least-squares matrix: full	Secondary atom site location: difference Fourier map
$R[F^2 > 2\sigma(F^2)] = 0.048$	Hydrogen site location: inferred from neighbouring sites
$wR(F^2) = 0.087$	H-atom parameters constrained
$S = 1.16$	$w = 1/[\sigma^2(F_o^2) + (0.0286P)^2]$
3410 reflections	where $P = (F_o^2 + 2F_c^2)/3$
163 parameters	$(\Delta/\sigma)_{\max} = 0.001$
0 restraints	$\Delta\rho_{\max} = 0.73 \text{ e } \text{\AA}^{-3}$
	$\Delta\rho_{\min} = -0.68 \text{ e } \text{\AA}^{-3}$

Special details

Geometry. All e.s.d.'s (except the e.s.d. in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell e.s.d.'s are taken into account individually in the estimation of e.s.d.'s in distances, angles and torsion angles; correlations between e.s.d.'s in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell e.s.d.'s is used for estimating e.s.d.'s involving l.s. planes.

Refinement. Refinement of F^2 against ALL reflections. The weighted R -factor wR and goodness of fit S are based on F^2 , conventional R -factors R are based on F , with F set to zero for negative F^2 . The threshold expression of $F^2 > \sigma(F^2)$ is used only for calculating R -factors(gt) *etc.* and is not relevant to the choice of reflections for refinement. R -factors based on F^2 are statistically about twice as large as those based on F , and R -factors based on ALL data will be even larger.

Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters (\AA^2)

	x	y	z	$U_{\text{iso}}^*/U_{\text{eq}}$
I1	0.28885 (4)	0.00696 (3)	0.52344 (2)	0.03779 (13)
I2	0.56758 (4)	0.03336 (3)	0.68154 (2)	0.04152 (14)
Cu1	0.43879 (8)	0.11605 (5)	0.58888 (4)	0.0440 (2)
Cu2	0.52112 (8)	-0.05428 (6)	0.56063 (5)	0.0430 (2)
S1	0.55004 (18)	0.33204 (12)	0.60075 (10)	0.0472 (5)
S2	0.57511 (16)	0.22524 (11)	0.53450 (9)	0.0381 (4)
N1	0.3452 (5)	0.2256 (4)	0.6339 (3)	0.0409 (13)
N2	0.4533 (4)	0.1941 (3)	0.4210 (3)	0.0300 (11)
C1	0.2348 (6)	0.2116 (5)	0.6675 (4)	0.0504 (19)
H1	0.1960	0.1530	0.6644	0.061*
C2	0.1761 (7)	0.2779 (6)	0.7059 (4)	0.058 (2)
H2	0.1002	0.2647	0.7293	0.069*
C3	0.2316 (7)	0.3650 (6)	0.7093 (4)	0.057 (2)
H3	0.1933	0.4124	0.7347	0.069*
C4	0.3452 (7)	0.3811 (5)	0.6745 (4)	0.0514 (19)
H4	0.3844	0.4397	0.6756	0.062*
C5	0.3986 (6)	0.3093 (4)	0.6386 (3)	0.0400 (16)
C6	0.3900 (6)	0.2177 (4)	0.3649 (3)	0.0374 (15)

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H6	0.3637	0.1702	0.3357	0.045*
C7	0.3617 (6)	0.3078 (4)	0.3479 (4)	0.0432 (17)
H7	0.3157	0.3215	0.3085	0.052*
C8	0.4027 (6)	0.3779 (4)	0.3903 (4)	0.0424 (17)
H8	0.3860	0.4401	0.3796	0.051*
C9	0.4671 (6)	0.3565 (4)	0.4474 (4)	0.0399 (16)
H9	0.4960	0.4031	0.4767	0.048*
C10	0.4893 (5)	0.2629 (4)	0.4615 (3)	0.0307 (14)

Atomic displacement parameters (\AA^2)

	U^{11}	U^{22}	U^{33}	U^{12}	U^{13}	U^{23}
I1	0.0346 (2)	0.0420 (3)	0.0368 (3)	-0.00164 (19)	-0.00135 (19)	-0.0036 (2)
I2	0.0539 (3)	0.0378 (2)	0.0329 (3)	-0.0041 (2)	-0.0081 (2)	0.00119 (19)
Cu1	0.0558 (5)	0.0339 (4)	0.0423 (6)	0.0019 (4)	-0.0071 (4)	-0.0047 (4)
Cu2	0.0515 (5)	0.0329 (4)	0.0447 (6)	0.0009 (4)	-0.0026 (4)	0.0002 (4)
S1	0.0593 (11)	0.0411 (10)	0.0411 (11)	-0.0087 (9)	-0.0065 (9)	-0.0111 (8)
S2	0.0420 (9)	0.0368 (9)	0.0354 (10)	0.0018 (8)	-0.0076 (8)	-0.0033 (7)
N1	0.045 (3)	0.041 (3)	0.037 (4)	0.008 (3)	-0.008 (3)	-0.001 (3)
N2	0.034 (3)	0.027 (3)	0.029 (3)	0.001 (2)	0.003 (2)	0.001 (2)
C1	0.047 (4)	0.059 (5)	0.045 (5)	-0.002 (4)	-0.005 (3)	-0.004 (4)
C2	0.049 (4)	0.089 (6)	0.036 (5)	0.016 (4)	-0.009 (4)	-0.011 (4)
C3	0.063 (5)	0.060 (5)	0.050 (5)	0.028 (4)	-0.016 (4)	-0.015 (4)
C4	0.071 (5)	0.044 (4)	0.039 (5)	0.014 (4)	-0.011 (4)	-0.008 (3)
C5	0.054 (4)	0.037 (4)	0.028 (4)	0.006 (3)	-0.010 (3)	-0.003 (3)
C6	0.037 (3)	0.041 (4)	0.033 (4)	0.005 (3)	-0.009 (3)	-0.005 (3)
C7	0.047 (4)	0.039 (4)	0.044 (4)	0.017 (3)	-0.008 (3)	0.004 (3)
C8	0.049 (4)	0.034 (3)	0.044 (4)	0.014 (3)	-0.007 (3)	0.002 (3)
C9	0.038 (3)	0.039 (4)	0.043 (4)	-0.004 (3)	0.000 (3)	-0.007 (3)
C10	0.029 (3)	0.029 (3)	0.034 (4)	0.001 (3)	0.000 (3)	-0.005 (3)

Geometric parameters (\AA , $^\circ$)

I1—Cu1	2.5760 (13)	N2—Cu2 ⁱ	2.068 (5)
I1—Cu2 ⁱ	2.6868 (15)	C1—C2	1.370 (10)
I1—Cu2	2.6892 (16)	C1—H1	0.9400
I2—Cu1	2.5772 (13)	C2—C3	1.385 (11)
I2—Cu2	2.7623 (17)	C2—H2	0.9400
Cu1—N1	2.064 (5)	C3—C4	1.395 (10)
Cu1—S2	2.385 (2)	C3—H3	0.9400
Cu1—Cu2	2.6650 (18)	C4—C5	1.377 (9)
Cu2—N2 ⁱ	2.068 (5)	C4—H4	0.9400
Cu2—I1 ⁱ	2.6868 (15)	C6—C7	1.377 (8)
Cu2—Cu2 ⁱ	2.912 (2)	C6—H6	0.9400
S1—C5	1.785 (7)	C7—C8	1.386 (9)
S1—S2	2.046 (3)	C7—H7	0.9400
S2—C10	1.792 (6)	C8—C9	1.357 (9)
N1—C5	1.334 (8)	C8—H8	0.9400

N1—C1	1.350 (9)	C9—C10	1.399 (8)
N2—C10	1.334 (7)	C9—H9	0.9400
N2—C6	1.341 (7)		
Cu1—I1—Cu2 ⁱ	73.10 (5)	C1—N1—Cu1	120.4 (5)
Cu1—I1—Cu2	60.77 (4)	C10—N2—C6	117.0 (5)
Cu2 ⁱ —I1—Cu2	65.59 (5)	C10—N2—Cu2 ⁱ	125.6 (4)
Cu1—I2—Cu2	59.76 (4)	C6—N2—Cu2 ⁱ	117.4 (4)
N1—Cu1—S2	88.52 (17)	N1—C1—C2	123.7 (7)
N1—Cu1—I1	113.53 (16)	N1—C1—H1	118.1
S2—Cu1—I1	122.57 (7)	C2—C1—H1	118.1
N1—Cu1—I2	106.96 (16)	C1—C2—C3	118.3 (8)
S2—Cu1—I2	108.56 (6)	C1—C2—H2	120.9
I1—Cu1—I2	113.40 (5)	C3—C2—H2	120.9
N1—Cu1—Cu2	162.08 (17)	C2—C3—C4	118.9 (7)
S2—Cu1—Cu2	108.70 (7)	C2—C3—H3	120.5
I1—Cu1—Cu2	61.71 (4)	C4—C3—H3	120.5
I2—Cu1—Cu2	63.57 (4)	C5—C4—C3	118.5 (7)
N2 ⁱ —Cu2—Cu1	154.75 (15)	C5—C4—H4	120.7
N2 ⁱ —Cu2—I1 ⁱ	105.23 (13)	C3—C4—H4	120.7
Cu1—Cu2—I1 ⁱ	97.82 (4)	N1—C5—C4	123.2 (7)
N2 ⁱ —Cu2—I1	119.11 (13)	N1—C5—S1	120.5 (5)
Cu1—Cu2—I1	57.52 (3)	C4—C5—S1	116.2 (6)
I1 ⁱ —Cu2—I1	114.41 (5)	N2—C6—C7	123.5 (6)
N2 ⁱ —Cu2—I2	105.64 (15)	N2—C6—H6	118.3
Cu1—Cu2—I2	56.67 (3)	C7—C6—H6	118.3
I1 ⁱ —Cu2—I2	107.23 (5)	C6—C7—C8	118.2 (6)
I1—Cu2—I2	104.38 (4)	C6—C7—H7	120.9
N2 ⁱ —Cu2—Cu2 ⁱ	133.76 (15)	C8—C7—H7	120.9
Cu1—Cu2—Cu2 ⁱ	68.25 (5)	C9—C8—C7	119.9 (6)
I1 ⁱ —Cu2—Cu2 ⁱ	57.25 (4)	C9—C8—H8	120.0
I1—Cu2—Cu2 ⁱ	57.17 (3)	C7—C8—H8	120.0
I2—Cu2—Cu2 ⁱ	120.18 (6)	C8—C9—C10	118.0 (6)
C5—S1—S2	104.3 (2)	C8—C9—H9	121.0
C10—S2—S1	103.3 (2)	C10—C9—H9	121.0
C10—S2—Cu1	105.6 (2)	N2—C10—C9	123.4 (6)
S1—S2—Cu1	97.40 (10)	N2—C10—S2	114.0 (4)
C5—N1—C1	117.3 (6)	C9—C10—S2	122.6 (5)
C5—N1—Cu1	121.7 (5)		
Cu2 ⁱ —I1—Cu1—N1	128.20 (18)	I1—Cu1—S2—C10	33.8 (2)
Cu2—I1—Cu1—N1	-160.76 (18)	I2—Cu1—S2—C10	169.2 (2)
Cu2 ⁱ —I1—Cu1—S2	24.03 (6)	Cu2—Cu1—S2—C10	101.6 (2)
Cu2—I1—Cu1—S2	95.07 (7)	N1—Cu1—S2—S1	22.68 (17)
Cu2 ⁱ —I1—Cu1—I2	-109.45 (5)	I1—Cu1—S2—S1	139.91 (8)
Cu2—I1—Cu1—I2	-38.41 (4)	I2—Cu1—S2—S1	-84.72 (9)
Cu2 ⁱ —I1—Cu1—Cu2	-71.04 (4)	Cu2—Cu1—S2—S1	-152.27 (7)

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Cu2—I2—Cu1—N1	163.59 (17)	S2—Cu1—N1—C5	-19.2 (5)
Cu2—I2—Cu1—S2	-102.17 (6)	I1—Cu1—N1—C5	-144.4 (4)
Cu2—I2—Cu1—I1	37.66 (4)	I2—Cu1—N1—C5	89.8 (5)
N1—Cu1—Cu2—N2 ⁱ	-11.1 (6)	Cu2—Cu1—N1—C5	145.1 (4)
S2—Cu1—Cu2—N2 ⁱ	152.3 (3)	S2—Cu1—N1—C1	169.7 (5)
I1—Cu1—Cu2—N2 ⁱ	-90.1 (3)	I1—Cu1—N1—C1	44.5 (6)
I2—Cu1—Cu2—N2 ⁱ	50.3 (3)	I2—Cu1—N1—C1	-81.3 (5)
N1—Cu1—Cu2—I1 ⁱ	-167.0 (5)	Cu2—Cu1—N1—C1	-26.0 (9)
S2—Cu1—Cu2—I1 ⁱ	-3.57 (6)	C5—N1—C1—C2	-0.4 (11)
I1—Cu1—Cu2—I1 ⁱ	114.03 (5)	Cu1—N1—C1—C2	171.2 (6)
I2—Cu1—Cu2—I1 ⁱ	-105.52 (5)	N1—C1—C2—C3	1.5 (12)
N1—Cu1—Cu2—I1	79.0 (5)	C1—C2—C3—C4	-0.9 (11)
S2—Cu1—Cu2—I1	-117.60 (7)	C2—C3—C4—C5	-0.7 (11)
I2—Cu1—Cu2—I1	140.45 (3)	C1—N1—C5—C4	-1.4 (10)
N1—Cu1—Cu2—I2	-61.4 (5)	Cu1—N1—C5—C4	-172.8 (5)
S2—Cu1—Cu2—I2	101.95 (7)	C1—N1—C5—S1	175.7 (5)
I1—Cu1—Cu2—I2	-140.45 (3)	Cu1—N1—C5—S1	4.3 (7)
N1—Cu1—Cu2—Cu2 ⁱ	143.0 (5)	C3—C4—C5—N1	1.9 (11)
S2—Cu1—Cu2—Cu2 ⁱ	-53.58 (6)	C3—C4—C5—S1	-175.3 (5)
I1—Cu1—Cu2—Cu2 ⁱ	64.02 (4)	S2—S1—C5—N1	16.5 (6)
I2—Cu1—Cu2—Cu2 ⁱ	-155.53 (5)	S2—S1—C5—C4	-166.2 (5)
Cu1—I1—Cu2—N2 ⁱ	150.78 (17)	C10—N2—C6—C7	0.1 (9)
Cu2 ⁱ —I1—Cu2—N2 ⁱ	-125.65 (17)	Cu2 ⁱ —N2—C6—C7	177.3 (5)
Cu2 ⁱ —I1—Cu2—Cu1	83.57 (5)	N2—C6—C7—C8	1.2 (11)
Cu1—I1—Cu2—I1 ⁱ	-83.57 (5)	C6—C7—C8—C9	-1.0 (10)
Cu2 ⁱ —I1—Cu2—I1 ⁱ	0.0	C7—C8—C9—C10	-0.3 (10)
Cu1—I1—Cu2—I2	33.31 (3)	C6—N2—C10—C9	-1.6 (9)
Cu2 ⁱ —I1—Cu2—I2	116.88 (6)	Cu2 ⁱ —N2—C10—C9	-178.5 (4)
Cu1—I1—Cu2—Cu2 ⁱ	-83.57 (5)	C6—N2—C10—S2	-179.2 (4)
Cu1—I2—Cu2—N2 ⁱ	-160.06 (14)	Cu2 ⁱ —N2—C10—S2	4.0 (6)
Cu1—I2—Cu2—I1 ⁱ	88.08 (5)	C8—C9—C10—N2	1.7 (10)
Cu1—I2—Cu2—I1	-33.68 (3)	C8—C9—C10—S2	179.1 (5)
Cu1—I2—Cu2—Cu2 ⁱ	26.43 (5)	S1—S2—C10—N2	-157.5 (4)
C5—S1—S2—C10	83.7 (3)	Cu1—S2—C10—N2	-55.7 (5)
C5—S1—S2—Cu1	-24.3 (2)	S1—S2—C10—C9	25.0 (6)
N1—Cu1—S2—C10	-83.4 (3)	Cu1—S2—C10—C9	126.7 (5)

Symmetry codes: (i) $-x+1, -y, -z+1$.

Fig. 1

