

PRELIMINARY COMMUNICATION

Open Access

# Reaction of pharmacological active tris-(2-hydroxyethyl)ammonium 4-chlorophenylsulfanylacetate with $ZnCl_2$ or $NiCl_2$ : first conversion of a protic ionic liquid into metallated ionic liquid

Anna N Mirskova<sup>1</sup>, Sergey N Adamovich<sup>1</sup>, Rudolf G Mirskov<sup>1</sup> and Uwe Schilde<sup>2\*</sup>

## Abstract

The reaction of pharmacological active protic ionic liquid tris-(2-hydroxyethyl)ammonium 4-chlorophenylsulfanylacetate  $H^+N(CH_2CH_2OH)_3 \cdot (^-OOCCH_2SC_6H_4Cl-4)$  (1) with zinc or nickel chloride in a ratio of 2:1 affords stable at room temperature powder-like adducts  $[H^+N(CH_2CH_2OH)_3]_2 \cdot [M(OOCCH_2SC_6H_4Cl-4)_2Cl_2]^{2-}$ ,  $M = Zn$  (2),  $Ni$  (3). By recrystallization from aqueous alcohol compound 2 unexpectedly gives  $Zn(OOCCH_2SC_6H_4Cl-4)_2 \cdot 2H_2O$  (4). Unlike 2, compound 3 gives crystals  $[N(CH_2CH_2OH)_3]_2Ni^{2+} \cdot [OOCCH_2SC_6H_4Cl-4]_2$  (5), which have a structure of metallated ionic liquid. The structure of 5 has been proved by X-ray diffraction analysis. It is the first example of the conversion of a protic ionic liquid into potentially biological active metallated ionic liquid (1  $\rightarrow$  3  $\rightarrow$  5).

## Findings

Alkanolammonium salts of inorganic and carbonic acids, also known as protic ionic liquids (PILs), have been the subject of many studies [1,2]. Depending on cation and anion structure, PILs can be liquid (room temperature ionic liquids) [2] or solid compounds with m.p. up to 100°C and even higher (176°C [3]). For example, 2-hydroxyethylammonium nitrate,  $H_3^+NCH_2CH_2OH \cdot NO_3^-$ , synthesized in 1888, has m.p. 52°C [4]. At the same time, 2-hydroxy-ethylammonium formiate,  $H_3^+NCH_2CH_2OH \cdot ^-OOCH$ , represents a typical room temperature PIL with extremely low freezing point (-82°C) [5]. Alkanolammonium PILs are used as catalysts in chemical reactions, as electrolytes in full cells, gas (such as  $CO_2$  and  $SO_2$ ) solvents and crystalline cellulose solvents, for desulfurization of fuel, enzymes stabilizers and promoters of their activity and for protein purification [6-10]. Also, they are employed for the design of nano-structured compounds [11]. Their toxicity and biological degradation have been studied [12,13].

Among the objects of our previous investigations were PILs containing cations of biologically active 2-hydroxyethylamines and anions of aroxy- and aryl(heteryl) sulfanyl(sulfonyl)acetic acids  $R_1R_2N^+H(CH_2CH_2OH)_{3-n} \cdot (^-OOCCH_2XR)$ ,  $R = Ar, Het$ ;  $R_1, R_2 = H, Alk$ ;  $X = O, S, SO_2$ ;  $n = 0-2$ .

These PILs are air-stable solids (m.p. 37-95°C) or viscous liquids, well soluble in water and polar solvents, representing a new class of pharmacologically active substances. Showing low toxicity ( $LD_{50} = 1500-6000$  mg/kg), they possess antiaggregation, antithrombotic, membrane-stabilizing, antioxidant, antisclerotic, adaptogenic, analgesic, cardiotropic, hypocholesterolemic, hemo- and immunotropic activities. These PILs protect the mammals and humans from shock, toxic stress, alcohol and heavy metal intoxication, and radiation. Their antitumor activity considerably exceeds or differs from the effect of the initial biologically active acids and alkanolamine [14-19].

They also exert pronounced growth-stimulating activity at very low concentrations ( $10^{-4} - 10^{-10}$  wt %) toward beneficial bacteria, yeasts, and fungi used in large-scale biotechnology processes (white biotechnology [20]) for

\* Correspondence: us@chem.uni-potsdam.de

<sup>2</sup>University of Potsdam, Institute of Chemistry, Karl-Liebknecht-Str. 24-25, D-14476 Potsdam, Germany

Full list of author information is available at the end of the article

manufacture of fodder, baker's yeasts and citric acid, barley sprouting for the preparation of brewer's malt, and breeding of silkworms [21].

Recently we have shown that metallated ionic liquid tris-(2-hydroxyethyl)amine-bis-(2-methylphenoxyacetate)zinc  $N(CH_2CH_2OH)_3Zn^{2+} \cdot 2(OOCCH_2OC_6H_4-Me-2)$  exhibits a pronounced anti-sclerotic effect [22].

We have assumed that the incorporation of essential metals (so-called "metals of life"), which are of vital importance for all living organisms: Ca, Mg, Zn, Mn, Cu, Fe, Co, Ni, etc. [23,24], can enhance or alter the biological activity of protic alkanolammonium ionic liquids.

To reach this goal, in this work we have studied the reaction of PIL tris-(2-hydroxyethyl)-ammonium 4-chlorophenylsulfanylacetate (**1**) (a non-toxic compound possessing antithrombotic, antioxidant and immunotropic activity) with Zn and Ni chlorides.

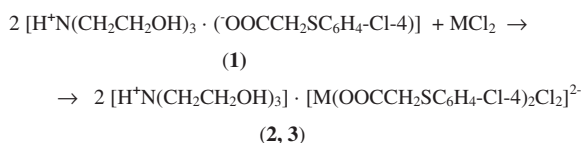
The interaction of **1** with metal salts furnishes powder compounds **2** and **3** (Scheme 1).

According to the data of IR spectroscopy, the compounds **2** and **3** contain coordination bonds  $HO \cdots M$  with the OH groups of two molecules of protonated triethanolamine and coordination bonds  $M-O$  with two carboxylate anions of 4-chlorophenylsulfanylacetic acid. So, the IR spectrum of **3** shows the absorption bands  $\nu(Ni-OC)$   $396\text{ cm}^{-1}$ ,  $\nu_{as}(COO^-)$  and  $\nu_s(COO^-)$  at  $1583$  and  $1401\text{ cm}^{-1}$ ,  $\Delta\nu = \nu_{as}(COO^-) - \nu_s(COO^-) = 182\text{ cm}^{-1}$ , characterizing bidentate coordination bonds of nickel atom with carboxylate anions; absorption bands typical for protonated triethanolamine  $HN^+(CH_2CH_2OH)_3$  at  $548$ ,  $549$  and  $397\text{ cm}^{-1}$ ,  $\nu(N^+H)$  is a broad band at  $2500-2700\text{ cm}^{-1}$ ; absorption band of the OH group at  $3312\text{ cm}^{-1}$ .

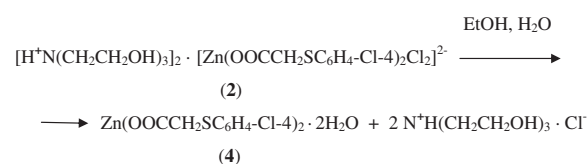
Powders **2** and **3** are stable at room temperature. However, on storage in solutions of organic solvents they change their composition and structure. So, for example, when recrystallized from aqueous alcohol ( $75^\circ\text{C}$ ), the powder adduct **2** is unexpectedly converted into zinc di-(4-chloro-phenylsulfonyl) acetate dihydrate **4** (Scheme 2).

Unlike compound **2**, compound **3** forms crystals **5** (Scheme 3).

The structure of compound **5** was established by X-ray crystal structure analysis. The molecular structure with the atom labeling scheme is given in Figure 1. The packing diagram is shown in Figure 2.



**Scheme 1** Synthesis of compounds **2** (M = Zn) and **3** (M = Ni).

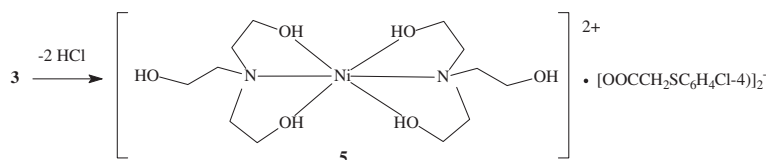


**Scheme 2** Conversion of compound **2** into compound **4**.

The Ni(II) cation is coordinated by four oxygen atoms of the hydroxyl groups and two nitrogen atoms forming an weakly distorted octahedral coordination environment. The asymmetric unit contains only the half of the cationic moiety. Because nickel is located on an inversion centre the second half is generated by inversion. The nitrogen atoms occupy the *trans* positions of the coordination polyhedron. Resulting from symmetry, the Ni-N bond lengths are equal being  $2.097(2)\text{ \AA}$ , and the N-Ni-N bond angle is  $180^\circ$ . The Ni-O distances are  $2.062(2)\text{ \AA}$  and  $2.070(2)\text{ \AA}$ . One hydroxyl group of each ethanol substituent is not involved in the coordination and directed away from the coordination centre. The N-Ni-O bond angles range from  $81.50(7)^\circ$  to  $98.50(7)^\circ$  and the O-Ni-O bond angles between  $85.12(7)^\circ$  and  $94.88(7)^\circ$ . Previously, structures containing the bis(triethanolamine)nickel (II) cation were described [25,26]. The 4-chloro-phenylsulfanyl unit is planar. The acetate substituent and the phenyl ring are almost in a tetrahedral arrangement with a C-S-C angle of  $101.4(1)^\circ$ . The carboxylic group is rotated around the S-C bond characterized by a C-S-C-C torsion angle of  $-50.7(3)^\circ$ . Cations and anions are linked by hydrogen bonds (see Figure 2). All oxygen atoms of the OH groups of tris-(2-hydroxy-ethyl)amine are involved in hydrogen bonds. Strong hydrogen bonds can be observed between those oxygen atoms, which are coordinated to the metal centre and both oxygen atoms of one carboxylic moiety. More weak hydrogen bonds are formed between the non-coordinated peripheric oxygen atoms of the OH groups and one oxygen atom of a carboxylic group. One oxygen atom (O4) of the carboxylic group forms bifurcated hydrogen bonds, one to the coordinated oxygen atom O3 and another one to the noncoordinated oxygen atom O3. That leads to a three-dimensional polymeric network.

## Experimental

IR spectra ( $\nu$ ,  $\text{cm}^{-1}$ ) were recorded on a Varian 3100 FT-IR75 spectrophotometer (KBr). NMR spectra ( $\delta$ , ppm) were measured on a DPX 400 instrument at  $25^\circ\text{C}$ . Reflections were collected using a STOE Imaging Plate Diffraction System (IPDS-II) at  $210\text{ K}$ . The data were corrected for Lorentz, polarisation and extinction effects. No absorption correction was applied. The structure was solved by direct methods as implemented in the program SHELXS-97 [27]. The refinement was carried out using SHELXL-97 [28]. All the non-hydrogen atoms



**Scheme 3** Conversion of compound 3 into compound 5.

were refined anisotropically. The hydrogen atoms of the phenyl groups were calculated in their expected positions. All the other hydrogen bonds were located from the difference Fourier map. The hydrogen atoms were refined isotropically. For the phenyl and methylene hydrogen atoms a riding model was used. The other hydrogen atoms were free refined. For the visualisation of the structure the program DIAMOND [29] was applied. CIF data: Additional file 1. CCDC reference number: 876072.

### Synthesis

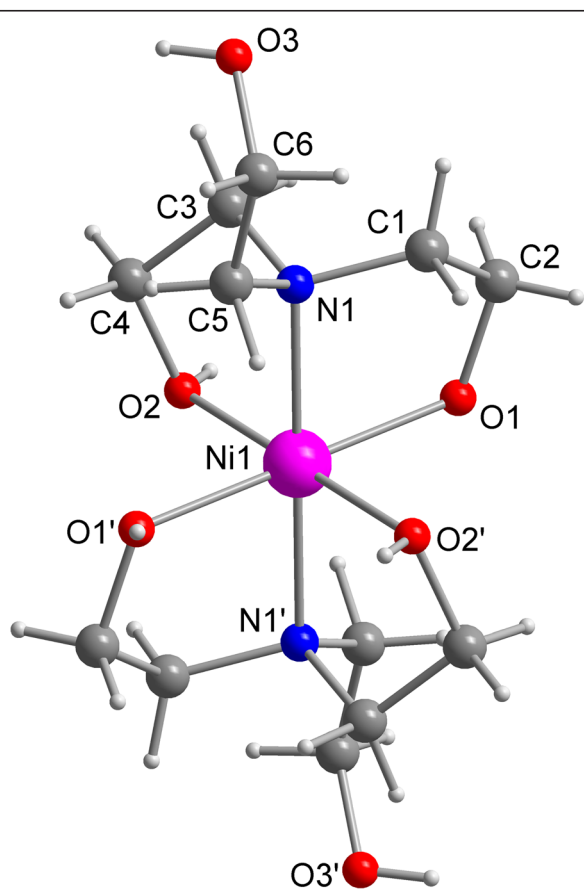
*Tris-(2-hydroxyethyl)ammonium 4-chlorophenylsulfanylacetate (1)* [30] was synthesized in the following manner. To a

solution of 4-chlorophenylsulfanylacetic acid 4-Cl-C<sub>6</sub>H<sub>4</sub>SCH<sub>2</sub>COOH (20.25 g, 0.1 mol) in MeOH (100 ml), was added dropwise a methanol (50 ml) solution of tris-(2-hydroxyethyl)amine (14.92 g, 0.1 mol). The mixture was stirred at 25°C for 30 min. The solvent was distilled in vacuum. The solid residue was repeatedly washed with ether and dried in vacuum to afford colorless powder (34.64 g, 98.5% yield), m.p. 90–92°C. For analytical characterization - see [30].

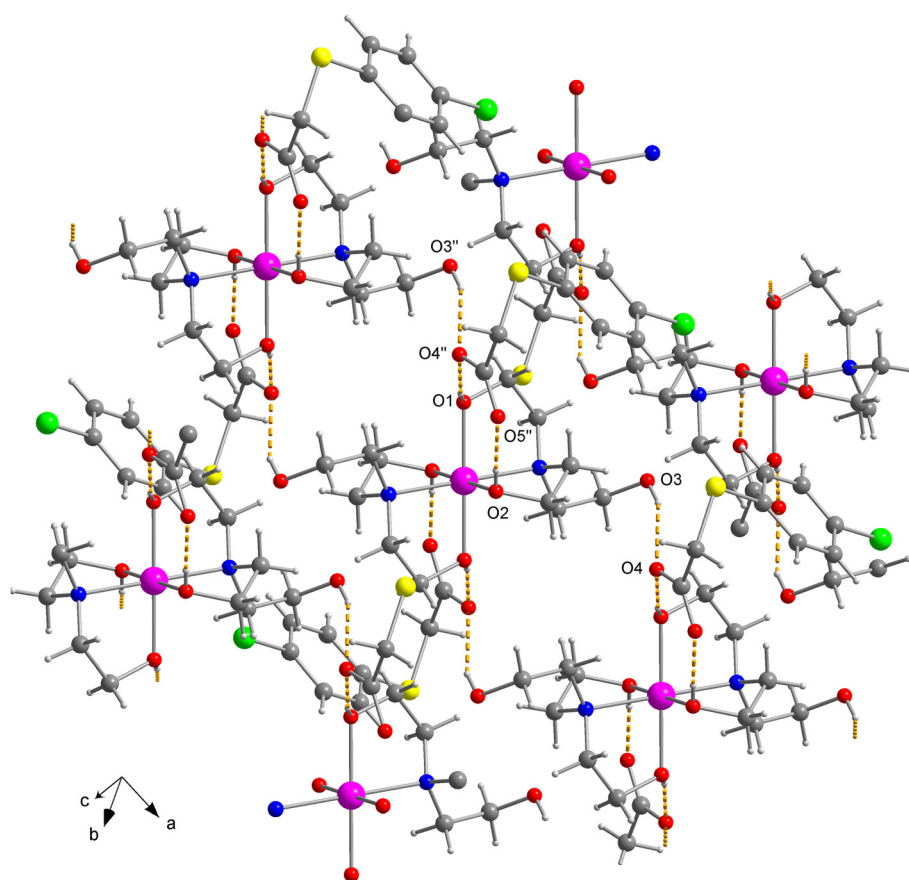
*Compound (2)*: To the solution of 7.03 g (0.02 mol) of **1** (m.p. 91°C) in 20 ml of MeOH the solution of 1.63 g (0.01 mol) ZnCl<sub>2</sub>·1.5H<sub>2</sub>O in 5 ml of MeOH was added dropwise. The reaction mixture was stirred at 25°C for 12 h, the solvent was removed in a vacuum. The solid residue was thoroughly washed with ether and dried over P<sub>2</sub>O<sub>5</sub> in a vacuum. **2** (3.71 g, 43%) of colorless powder was obtained, m.p. 152°C, readily soluble in alcohols and moderately in H<sub>2</sub>O. IR: 1439 ν<sub>s</sub>(COO), 1553 ν<sub>as</sub>(COO), 3305 (OH). <sup>1</sup>H NMR (100 MHz, *d*<sub>4</sub>-methanol): 7.30–7.23 (4H, m, C<sub>6</sub>H<sub>4</sub>), 3.84 (6H, t, OCH<sub>2</sub>), 3.63 (2H br s, SCH<sub>2</sub>), 3.28 (6H, t, NCH<sub>2</sub>). <sup>13</sup>C NMR (400 MHz, *d*<sub>4</sub>-methanol): 174.82 (C = O), 135.84–128.2 (C<sub>6</sub>H<sub>4</sub>), 55.52 (OCH<sub>2</sub>), 55.20 (NCH<sub>2</sub>), 37.37 (SCH<sub>2</sub>). Anal. Calc. for C<sub>28</sub>H<sub>44</sub>O<sub>10</sub>S<sub>2</sub>N<sub>2</sub>Cl<sub>4</sub>Zn: C 40.00; H 5.24; Cl 16.90; S 7.62; Zn 7.78. Found: C 39.67; H 5.78; Cl 16.97; S 8.26; Zn 7.82.

*Compound (3)*: To the solution of **1** (0.703 g, 0.002 mol) in methanol (10 ml) was added dropwise a methanol solution (10 ml) of NiCl<sub>2</sub>·6H<sub>2</sub>O (0.237 g, 0.001 mol). The reaction mixture was stirred at 25°C for 15 h. The solvent was distilled in vacuum to give light-green powder **3**, m.p. 170°C. Yield 0.55 g (59%). Well soluble in H<sub>2</sub>O, less soluble in alcohols. IR: 1401 ν<sub>s</sub>(COO), 1583 ν<sub>as</sub>(COO), 3312 (OH). <sup>1</sup>H NMR (100 MHz, *d*<sub>4</sub>-methanol): 7.33–7.01 (4H, m, C<sub>6</sub>H<sub>4</sub>S), 3.91 (6H, t, OCH<sub>2</sub>), 3.65 (2H, c, SCH<sub>2</sub>), 3.44 (6H, t, NCH<sub>2</sub>). <sup>13</sup>C NMR (400 MHz, *d*<sub>4</sub>-methanol): 177.09 (C = O), 131.76–128.25 (C<sub>6</sub>H<sub>4</sub>S), 55.33 (OCH<sub>2</sub>), 54.98 (NCH<sub>2</sub>), 37.69 (SCH<sub>2</sub>). Anal. Calc. for C<sub>28</sub>H<sub>44</sub>Cl<sub>4</sub>N<sub>2</sub>O<sub>10</sub>S<sub>2</sub>Ni: C 40.30; H 5.27; Cl 17.01; Ni 7.04. Found: C 41.13; H 4.99; Cl 17.29; Ni 6.88.

*Zinc di-(4-chlorophenylsulfanyl)acetate dihydrate (4)*: 0.5 g of **2** was dissolved in 10 ml of aqueous alcohol upon stirring (75°C), the solution was kept for one month at room temperature and filtered. The solid residue was washed with ether and dried over P<sub>2</sub>O<sub>5</sub> to obtain colorless plate crystals **4** with m.p. 202°C. IR: 1416 ν<sub>s</sub>(COO), 1540 ν<sub>as</sub>



**Figure 1** Molecular structure of **5**, showing the atom labelling. The bis-[tris-(2-hydroxyethyl)-amine]-nickel(II) cation is centrosymmetric.



**Figure 2** Crystal packing of **5**, illustrating the hydrogen bonds (dashed lines). Hydrogen bonds geometry: O1...O4'' 2.583(2) Å, H1...O4'' 2.00(3) Å, O1-H1...O4'' 171(5)°; O2...O5'' 2.563(2) Å, H2...O5'' 1.87(3) Å, O2-H2...O5'' 177(3)°; O3...O4 2.886(3) Å, H3...O4 2.13(3) Å, O3-H3...O4 158(3)°. Symmetry operator: " x-1, y, z.

(COO), 3240 (OH).  $^1\text{H}$  NMR (100 MHz,  $d_4$ -methanol): 7.31-7.21 (4H, m,  $\text{C}_6\text{H}_4$ ), 3.63 (2H, br.s,  $\text{SCH}_2$ ).  $^{13}\text{C}$  NMR (400 MHz,  $d_4$ -methanol): 175.02 (C = O), 131.80-128.65 ( $\text{C}_6\text{H}_4$ ), 37.37 ( $\text{SCH}_2$ ). Anal. Calc. for  $\text{C}_{16}\text{H}_{16}\text{O}_6\text{S}_2\text{Cl}_2\text{Zn}$ : C 38.01; H 3.17; Zn 12.95. Found: C 38.17; H 3.18; Zn 12.82.

*Bis-[(tris-2-hydroxyethyl)ammonium]nickel(II) di-(4-chlorophenylsulfanyl)acetate (5)*: From a solution **3** (aqueous alcohol, 75°C), the blue crystals **5** were obtained (20°C, for one month), m.p. 176°C.  $^1\text{H}$  NMR (100 MHz,  $d_4$ -methanol): 7.13-6.79 (4H, m,  $\text{C}_6\text{H}_4\text{S}$ ), 3.49 (6H, t,  $\text{OCH}_2$ ), 3.29 (2H, c,  $\text{SCH}_2$ ), 2.67 (6H, t,  $\text{NCH}_2$ ).  $^{13}\text{C}$  NMR (400 MHz,  $d_4$ -methanol): 178.90 (C = O), 136.16 -128.40 ( $\text{C}_6\text{H}_4\text{S}$ ), 56.50 ( $\text{OCH}_2$ ), 54.43 ( $\text{NCH}_2$ ), 37.62 ( $\text{SCH}_2$ ). Anal. Calc. for  $\text{C}_{28}\text{H}_{42}\text{Cl}_2\text{N}_2\text{O}_{10}\text{S}_2\text{Ni}$ : C 44.18; H 5.79; Cl 9.33; Ni 7.72. Found: C 43.88; H 5.63; Cl 9.54; Ni, 8.01.

#### Crystal data of **5**

$\text{C}_{28}\text{H}_{42}\text{Cl}_2\text{N}_2\text{NiO}_{10}\text{S}_2$ ,  $M = 760.37$ , triclinic,  $a = 8.0020$  (8),  $b = 9.4980$ (10),  $c = 11.6401$ (12) Å,  $V = 823.29$ (15) Å<sup>3</sup>,  $T = 210$ (2) K, space group  $P-1$  (no.2),  $Z = 1$ ,  $\mu(\text{MoK}\alpha) = 0.936$  mm<sup>-1</sup>; 5252 reflections measured, 2714 unique

( $R_{\text{int}} = 0.031$ ) which were used in all calculations. Final  $R$  values:  $wR_2(F^2) = 0.0603$ ,  $R_1 = 0.0445$  (all data);  $wR_2(F^2) = 0.0575$ ,  $R_1 = 0.0300$  [ $I > 2\sigma$ ].

#### Conclusion

The reaction of pharmacological active ionic liquid  $\text{H}^+\text{N}(\text{CH}_2\text{CH}_2\text{OH})_3 \cdot (\text{OOCCH}_2\text{SC}_6\text{H}_4\text{Cl-4})$  (**1**) with zinc or nickel chloride affords stable at room temperature powder-like adducts  $[\text{H}^+\text{N}(\text{CH}_2\text{CH}_2\text{OH})_3]_2 \cdot [\text{M}(\text{OOCCH}_2\text{SC}_6\text{H}_4\text{Cl-4})_2\text{Cl}_2]^{2-}$ ,  $\text{M} = \text{Zn}$  (**2**),  $\text{Ni}$  (**3**). By recrystallization compound **2** unexpectedly gives  $\text{Zn}(\text{OOCCH}_2\text{SC}_6\text{H}_4\text{Cl-4})_2 \cdot 2\text{H}_2\text{O}$  (**4**). Unlike **2**, compound **3** gives crystals  $[\text{N}(\text{CH}_2\text{CH}_2\text{OH})_3]_2\text{Ni}^{2+} \cdot [\text{OOCCH}_2\text{SC}_6\text{H}_4\text{Cl-4}]_2$  (**5**), which have a structure of metallated ionic liquid. It is the first example of the conversion of a protic ionic liquid into metallated ionic liquid (compound **1**  $\rightarrow$  **3**  $\rightarrow$  **5**). The structure of **5** has been proved by X-ray diffraction analysis. The investigation of physiological activity of metallated ionic liquids will be conducted in a new future.



## Additional file

**Additional file 1: Crystallographic information. Contains all relevant CIF information.**

### Competing interests

The authors declare that they have no competing interests.

### Authors' contributions

SNA carried out the synthetic experiments and drafted the manuscript. ANM has formulated the research idea and prepared the manuscript draft version. RGM prepared the manuscript for submission and coordinated final formulation. US collected the X-ray data and performed the structure solution. All authors read and approved the final manuscript.

### Author details

<sup>1</sup>A. E. Favorsky Irkutsk Institute of Chemistry, Siberian Branch, Russian Academy of Sciences, 1 Favorsky Str, 664033 Irkutsk, Russia. <sup>2</sup>University of Potsdam, Institute of Chemistry, Karl-Liebknecht-Str. 24-25, D-14476 Potsdam, Germany.

Received: 5 December 2012 Accepted: 11 February 2013

Published: 19 February 2013

### References

- Wasserscheid P, Welton T: *Ionic liquids in synthesis wiley-VCH GmbH & KGaA*. Germany: Weinheim; 2008.
- Hallett JP, Welton T: Room-temperature ionic liquids: solvents for synthesis and catalysis. *2. Chem Rev* 2011, **111**:3508–3576.
- Bica K, Rijkse C, Nieuwenhuyzen M, Rogers RD: In search of pure liquid salt forms of aspirin: ionic liquid approaches with acetylsalicylic acid and salicylic acid. *Phys Chem Chem Phys* 2010, **12**:2011–2017.
- Gabriel S: Ueber vinylamin und bromäthylamin. *2. Chemische Berichte* 1888, **21**(2):2664–2669.
- Bicak N: A new ionic liquid: 2-hydroxy ethylammonium formate. *J Mol Liquids* 2005, **116**:15–18.
- Alizadeh A, Khodaei MM, Eshghi A: Amphiphilic dual activation role of a task-specific ionic liquid: 2-hydroxyethylammonium formate as a recyclable promoter and medium for the green synthesis of  $\beta$ -nitrostyrenes. *J Org Chem* 2010, **75**:8295–8298.
- Zhang Q, Zhang S, Deng Y: Recent advances in ionic liquid catalysis. *Green Chem* 2011, **13**:2619–2637.
- Pinkert A, Marsh KN, Pang S: Alkanolamine ionic liquids and their inability to dissolve crystalline cellulose. *Ind Eng Chem Res* 2010, **49**:11809–11813.
- Mann JP, McCluskey A, Atkin R: Activity and thermal stability of lysozyme in alkylammonium formate ionic liquids - influence of cation modification. *Green Chem* 2009, **11**:785–792.
- Hekmat D, Hebel D, Joswig S, Schmidt M, Weuster-Botz D: Advanced protein crystallization using water-soluble ionic liquids as crystallization additives. *Biotechnol Lett* 2007, **29**:1703–1711.
- Greaves TL, Kennedy DF, Mudie ST, Drummond SJ: Diversity observed in the nanostructure of protic ionic liquids. *J Phys Chem B* 2010, **114**:10022–10031.
- Pavlovica S, Zicmanis A, Gzibovska E, Klavins M, Peteris M: (2-Hydroxyethyl) ammonium lactates - highly biodegradable and essentially Non-toxic ionic liquids. *Green Sustain Chem* 2011, **1**:103–110.
- Wood N, Ferguson JL, Nimal Gunaratne HQ, Seddon KR, Goodacre R, Stephens RG: Screening ionic liquids for use in biotransformations with whole microbial cells. *Green Chem* 2011, **13**:1843–1851.
- Kolesnikova OP, Mirskova AN, Adamovich SN, Mirskov RG, Kudaeva OT, Voronkov MG: Alkanolammonium salts of o-cresoxy- and p-chlorocresoxyacetic acids as immunopoiesis modulators and cytostatics. *Dokl Biol Sci* 2009, **425**:107–111.
- Voronkov MG, Sofronov GA, Starchenko DA, Adamovich SN, Mirskova AN: Protective properties of chlorocresacine against adverse impact of electromagnetic radiation. *Dokl Biol Sci* 2009, **428**:398–402.
- Mirskova AN, Adamovich SN, Mirskov RG, Voronkov MG: Tris-(2-hydroxyethyl)ammonium 2-methyl- and 2-methyl-4-chlorophenoxyacetate serve as effective inhibitors of thrombocyte aggregation and antioxidants. *Dokl Biol Sci* 2010, **433**:244–246.
- Mirskova AN, Mirskov RG, Adamovich SN, Voronkov MG: Indole-3-ylsulfanyl acetate tris-(2-hydroxy-ethyl)ammonium is an effective antioxidant and cell membranes stabilizer. *Dokl Biol Sci* 2010, **435**:390–392.
- Mirskova AN, Levkovskaya GG, Kolesnikova OP, Perminova OM, Rudyakova EV, Adamovich SN: Directed synthesis and immunoactive properties of (2-hydroxyethyl)ammonium salts of 1-R-indol-3-ylsulfanyl-(sulfonyl) alkanecarboxylic acids. *Russ Chem Bull (Engl Transl)* 2010, **59**:2236–2246.
- Mirskova AN, Mirskov RG, Adamovich SN, Voronkov MG: 2-Hydroxyethylammonium salts organylsulfanyl(sulfonyl)acetic acids - new pharmacological compounds. *Russ J Chem Sustainable Develop (Engl Transl)* 2011, **19**:467–478.
- Gupta MN, Raghava S: Relevance of chemistry to white biotechnology. *Chem Central J* 2007, **1**:17.
- Mirskova AN, Levkovskaya GG, Mirskov RG, Voronkov MG: Hydroxyalkylammonium salts of organylsulfanyl(sulfonyl)acetic acids - New stimulators of biological processes. *Russ J Org Chem* 2008, **44**:1478–1485.
- Rasulov MM, Voronkov MG, Nurbekov MK, Zvereva MV, Mirskova AN, Adamovich SN, Mirskov RG: Complex bis-(2-methylphenoxyacetate)zinc with tris-(2-hydroxyethyl)amine - as a activator of synthesis of total triptophanyl - tRNA-synthetase. *Dokl Biol Sci* 2012, **444**:219–222.
- Gielen M, Tiekink ERT: *Metallotherapeutic drugs and metal-based diagnostic agent. The use of metals in medicine*. Chichester, UK: Wiley & Sons Ltd; 2005.
- Oberleas D, Harland BF, Skalny AV: *The biological role of macro and trace elements in humans and animals*. St. Petersburg: Nauka; 2008.
- Guo H-X, Du ZX, Li XZ: Bis(triethanolamine)nickel(II) sulfate. *Acta Crystallogr Sect E Struct Rep* 2009, **65**:m810–m811. Online available. <http://journals.iucr.org/e/issues/2009/07/00/pk2167/pk2167.pdf>.
- Haukka M, Kirillov AM, Kopylovich MN, Pombeiro AJL: Bis(triethanolamine- $\kappa^3N,O,O'$ )nickel(II) benzene-1,4-dicarboxylate. *Acta Crystallogr Sect E Struct Rep* 2005, **61**:m2746–m2748. Online available. <http://journals.iucr.org/e/issues/2005/12/00/rn6069/index.html>.
- Sheldrick GM: *SHELXS-97: Program for the solution of crystal structures*. Germany: University of Göttingen; 1997.
- Sheldrick GM: *SHELXL-97, Program for the refinement of crystal structures*. Germany: University of Göttingen; 1997.
- Brandenburg K: *Diamond*. Version 31st edition. Bonn: Crystal Impact; 2005.
- Levkovskaya GG, Guseva SA, Kazimirovskaya VB, Kholdeeva LN, Voronkov MG, Mirskova AN: 2-Hydroxyalkylammonium salts aryltioacetic asides and there influence on functional activity of trombosides. II. *Russ Chem Pharm J* 1986, **22**:295–300.

doi:10.1186/1752-153X-7-34

**Cite this article as:** Mirskova et al.: Reaction of pharmacological active tris-(2-hydroxyethyl)ammonium 4-chlorophenylsulfanylacetate with ZnCl<sub>2</sub> or NiCl<sub>2</sub>: first conversion of a protic ionic liquid into metallated ionic liquid. *Chemistry Central Journal* 2013 **7**:34.

Publish with **ChemistryCentral** and every scientist can read your work free of charge

“Open access provides opportunities to our colleagues in other parts of the globe, by allowing anyone to view the content free of charge.”

W. Jeffery Hurst, The Hershey Company.

- available free of charge to the entire scientific community
- peer reviewed and published immediately upon acceptance
- cited in PubMed and archived on PubMed Central
- yours — you keep the copyright

Submit your manuscript here:  
<http://www.chemistrycentral.com/manuscript/>

  
**ChemistryCentral**