



Cite this: *RSC Adv.*, 2018, 8, 24428

Insights into the spontaneous formation of hybrid PdO_x/PEDOT films: electroless deposition and oxygen reduction activity†

Julian A. Vigil,^a Michael T. Brumbach,^b Jonathon Duay^a and Timothy N. Lambert^{ID}*^a

Hybrid palladium oxide/poly(3,4-ethylenedioxythiophene) (PdO_x/PEDOT) films were prepared through a spontaneous reaction between aqueous PdCl₄²⁻ ions and a nanostructured film of electropolymerized PEDOT. Spectroscopic and electrochemical characterization indicate the presence of mixed-valence Pd species as-deposited (19 ± 7 at% Pd⁰, 64 ± 3 at% Pd²⁺, and 18 ± 4 at% Pd⁴⁺ by X-ray photoelectron spectroscopy) and the formation of stable, electrochemically reversible Pd^{0/α}-PdO_x active species in alkaline electrolyte and furthermore in the presence of oxygen. The elucidation of the Pd speciation as-deposited and in solution provides insight into the mechanism of electroless deposition in neutral aqueous conditions and the electrocatalytically active species during oxygen reduction in alkaline electrolyte. The PdO_x/PEDOT film catalyses 4e⁻ oxygen reduction (*n* = 3.97) in alkaline electrolyte at low overpotential (0.98 V vs. RHE, onset potential), with mass- and surface area-based specific activities competitive with, or superior to, commercial 20% Pt/C and state-of-the-art Pd- and PEDOT-based nanostructured catalysts. The high activity of the nanostructured hybrid PdO_x/PEDOT film is attributed to effective dispersion of accessible, stable Pd active sites in the PEDOT matrix.

Received 23rd April 2018
Accepted 18th June 2018

DOI: 10.1039/c8ra03505a
rsc.li/rsc-advances

Introduction

Overcoming the kinetic challenges of the electrocatalytic oxygen reduction reaction (ORR) is central to addressing some of the limitations of next-generation energy conversion devices, including the cathodic overpotential in fuel cells.¹ Pt-based materials have been the benchmark for electrocatalytic ORR activity, but concerns regarding the abundance, cost, and stability of Pt have motivated the identification of Pt-free electrocatalyst materials.² Furthermore, the enhanced stability of Pt-free materials and more facile kinetics in alkaline conditions, along with recent advances in anion exchange membranes, have revitalized interest in Anion Exchange Membrane Fuel Cells (AEMFCs).² Thus, materials based on alternative transition metals, heteroatom-doped carbon, and transition metal oxides that exhibit promising ORR activity and stability in alkaline electrolytes have been developed.²⁻⁶

As an emerging class of organic materials, conducting polymers (CPs) offer high electrical conductivity and, being polymeric in nature, may serve as a single replacement for both carbon and binders in composite electrode assemblies.⁷ In addition, catalytic activity discovered from CP-only electrodes such as polyaniline and poly(3,4-ethylenedioxythiophene) (PEDOT) has generated interest in developing CP-based electrocatalysts for the ORR.⁷⁻⁹ Winther-Jensen *et al.* first reported ORR activity competitive with Pt from a PEDOT air electrode prepared by vapor phase polymerization.⁹ Since then, studies have notably demonstrated that the ORR pathway on PEDOT-only electrodes is determined by the polymerization method¹⁰ and that an oxygen-induced doping mechanism may be responsible for the sustained conductivity of PEDOT at the highly reducing potentials of the ORR.¹¹ PEDOT has also served as a conductive support in composite ORR electrocatalysts with semiconducting metal oxides, including CoMn₂O₄,¹² Fe₃O₄,¹³ FeO,¹⁴ and MnO_x.^{15,16} Our group has reported two approaches to prepare highly active MnO_x/PEDOT composite electrocatalysts for the ORR: anodic co-electrodeposition¹⁵ and spontaneous reduction of MnO₄⁻ ions to form MnO_x on PEDOT.¹⁶ The latter approach provided a composite film with improved specific activity and active site utilization for catalysing the ORR.¹⁶

In addition to first-row transition metal oxides, Pd-based electrocatalysts are promising alternatives to Pt-based AEMFC cathodes owing to similar electronic properties to Pt (*i.e.* d-band center) and superior electrochemical activity and stability in

^aDepartment of Materials, Devices & Energy Technologies, Sandia National Laboratories, Albuquerque, New Mexico 87185, USA. E-mail: tnlambe@sandia.gov; Tel: +1 505 284 6967

^bMaterials Characterization & Performance, Sandia National Laboratories, Albuquerque, New Mexico 87185, USA

† Electronic supplementary information (ESI) available: Characterization methods and calculations; additional characterization and electrocatalytic data, literature comparison for alkaline Pd/C ORR electrocatalysts. See DOI: 10.1039/c8ra03505a



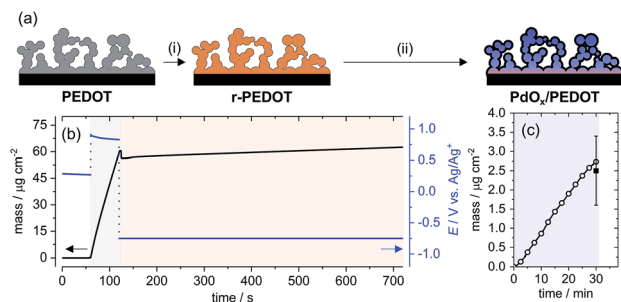


Fig. 1 (a) Schematic representation of the preparation of PdO_x/PEDOT: (i) reduction of electrodeposited PEDOT in CH₃CN, (ii) reaction of r-PEDOT with PdCl₄²⁻ in neutral aqueous solution; (b) working electrode potential and mass change (measured by QCM) during preparation of the r-PEDOT film: $t = 0-60$ s, open circuit potential; $t = 60-120$ s, electropolymerization; $t = 120-720$ s, PEDOT reduction; (c) mass change of a representative r-PEDOT film during electroless deposition of Pd species (circles) and average mass loading of Pd species after 30 min (square, 3 measurements).

alkaline electrolytes.^{17,18} This may be attributed to reduced dissolution of oxidized Pd (compared to Pt) and suitable reversibility of Pd redox in the ORR potential window.¹⁹ Pd is also approximately 50 times more abundant on Earth than Pt.¹⁷

Here we report on PdO_x/PEDOT films prepared by a spontaneous redox reaction between PdCl₄²⁻ ions and a reduced PEDOT film in neutral aqueous solution (Fig. 1a), and the electrocatalytic activity of the hybrid films toward the ORR. Previously, spontaneous electroless deposition using Pd²⁺ ions and PEDOT has been reported exclusively in acidic aqueous conditions,²⁰⁻²² and the chemical nature (*e.g.* Pd valence) of the as-deposited Pd species has not been studied. Such information regarding Pd speciation would provide insight across a wide range of conditions as Pd⁰, PdO, and PdO₂ are all thermodynamically stable at pH > 2. For the first time, we elucidate the Pd speciation as-deposited in neutral aqueous conditions and in basic electrolyte to identify the deposition mechanism and active species involved in catalysing the ORR, respectively. The specific ORR activity of the hybrid PdO_x/PEDOT film at low overpotential (0.8–0.9 V vs. RHE) is superior to commercial 20% Pt/C and competitive with state-of-the-art Pd/C and PEDOT-based nanostructured electrocatalysts at low Pd loading (~5 wt%).

Experimental

Materials

Potassium tetrachloropalladate(II) (K₂PdCl₄, 99.99%), lithium perchlorate (LiClO₄, 99.99%), 3,4-ethylenedioxythiophene (EDOT, 97%), potassium hydroxide (KOH, reagent grade >85%) and Nafion solution (5 wt% in lower alcohols) were obtained from Sigma-Aldrich. Acetonitrile (CH₃CN, Certified ACS) and isopropanol (IPA, Certified ACS) was obtained from Fisher Scientific. Ethanol (EtOH, Certified ACS) was obtained from Pharmco-Aaper. Commercial 20% Pt/C (Vulcan XC-72) catalyst powder was obtained from E-TekSM. All materials were used as-received.

Preparation of r-PEDOT and PdO_x/PEDOT films

Glassy carbon (GC) electrodes were polished with 0.05 μm alumina slurry and rinsed with DI H₂O and EtOH. A standard three-electrode cell was used for electropolymerization of PEDOT, including a GC working electrode (pyrolyzed GC electrodes on SiO₂ substrates or commercial rotating ring disk working electrodes), graphite rod counter electrode, and Ag/Ag⁺ reference electrode (polished Ag wire, 10 mM AgNO₃/0.1 M LiClO₄, CH₃CN). The cell was controlled by a Solartron Analytical 1287 Potentiostat Galvanostat (Ametek, Inc.). Electropolymerization of PEDOT proceeded in a bath of 0.05 M EDOT monomer and 0.1 M LiClO₄ in CH₃CN at constant anodic current density of 1 mA cm⁻² for 60 s. The pristine PEDOT film was electrochemically reduced at constant potential of -0.75 V vs. Ag/Ag⁺ for ten min. The reduced PEDOT film (r-PEDOT) was transferred to a water bath for five min to allow for wetting and removal of unreacted monomer and supporting electrolyte. For PdO_x/PEDOT, the r-PEDOT film was transferred directly from the water bath to a deposition solution of 1 mM K₂PdCl₄ for 30 min at room temperature. The electrode surface was inverted in the deposition solution to ensure the Pd species deposited on the film originated from reaction with PEDOT, and not instability in the bulk solution resulting in precipitation onto the electrode surface. Finally, the PdO_x/PEDOT film was likewise transferred to a water bath for five min.

Results and discussion

The EDOT monomer was electropolymerized galvanostatically from CH₃CN with LiClO₄ electrolyte onto a GC electrode (see Experimental section for details).¹⁶ The resulting oxidized PEDOT film (doped with perchlorate anions) was electrochemically reduced at constant potential in the CH₃CN electrolyte for 10 min, denoted r-PEDOT [Fig. 1a, (i)]. The r-PEDOT film was then used directly as a template for the electroless deposition of Pd species from aqueous K₂PdCl₄ for 30 min at room temperature [Fig. 1a, (ii)]; denoted PdO_x/PEDOT – see discussion below on the as-deposited Pd speciation.

The preparation of the hybrid PdO_x/PEDOT film was reproduced on Ti/Pt quartz crystal electrodes monitored by a Quartz Crystal Microbalance (QCM) to understand the mass change and doping/de-doping behaviour of the PEDOT film during the electrochemical reduction and Pd deposition steps. Anodic electropolymerization of PEDOT results in a linear increase in film mass at the electrode surface, with an average deposition rate of $0.98 \pm 0.03 \mu\text{g cm}^{-2} \text{ s}^{-1}$ (Fig. 1b). Subsequently, the reduction of PEDOT is accompanied by an initial decrease in film mass, explained by rapid de-doping of the PEDOT film,²³ followed by an overall increase in film mass over 10 min. The average mass of the r-PEDOT film was $63 \pm 5 \mu\text{g cm}^{-2}$.

Mass uptake of Pd species by the r-PEDOT film was approximately linear over the 30 min deposition time, resulting in a mass loading of Pd species of $2.5 \pm 0.9 \mu\text{g cm}^{-2}$ (Fig. 1c), or $3.9 \pm 1.4 \text{ wt}\%$ relative to the total film mass. Elemental analysis of PdO_x/PEDOT films digested in acid was also investigated by inductively coupled plasma-mass spectrometry, indicating 4.8

± 0.3 wt% Pd composition. Based on the following discussion, we propose that the spontaneous reaction of PdCl_4^{2-} with r-PEDOT produces a $\text{PdO}_x/\text{PEDOT}$ composite formed as follows: (i) initial e^- transfer between r-PEDOT and Pd^{2+} ions to form PEDOT-bound Pd^0 ; (ii) subsequent growth of a mixed-valence $\text{Pd}^{2+}/\text{Pd}^{4+}$ oxide by reactions between PEDOT-bound Pd^0 , Pd^{2+} ions, H_2O and dissolved or adsorbed O_2 .²⁴

Scanning electron microscope (SEM) images of the $\text{PdO}_x/\text{PEDOT}$ and r-PEDOT films (Fig. 2a and S1†) demonstrate a porous, interconnected nanostructure. The similar morphology of the r-PEDOT and $\text{PdO}_x/\text{PEDOT}$ films suggests the electroless deposition produces a homogenous coverage of Pd species on and within the polymeric film. This was confirmed by SEM with energy dispersive spectroscopy, where elemental mapping shows uniform distribution of Pd (Fig. S2†). Individual Pd species (particles) were not observed, suggesting they may be extremely small, on the order or smaller than surface features resulting from Pt sputter coating. X-ray diffraction studies on a $\text{PdO}_x/\text{PEDOT}$ film produced no diffraction pattern (not shown), which may also be indicative of small particle size. In contrast, electroless deposition of Pd species in 0.1 M H_2SO_4 results in the formation of spherical particles with >20 nm diameter within two minutes.²²

A representative high-resolution X-ray photoelectron spectroscopy (XPS) spectrum collected on a $\text{PdO}_x/\text{PEDOT}$ film (as-deposited) in the Pd 3d binding energy (E_b) region is shown in Fig. 2b. All $\text{PdO}_x/\text{PEDOT}$ Pd 3d spectra were fit with three symmetric components; average $E_b(\text{Pd } 3d_{5/2})/E_b(\text{Pd } 3d_{3/2})$ positions of $336.0 \pm 0.2/341.2 \pm 0.1$ eV, $337.8 \pm 0.1/343.0 \pm 0.1$ eV, and $338.8 \pm 0.1/344.5 \pm 0.2$ eV correspond with peak positions and spin-splitting values assigned to Pd^0 , Pd^{2+} , and Pd^{4+} , respectively.²⁵ In addition, the E_b shift of 1.8 eV between the Pd^0

and Pd^{2+} lines is indicative of the PdO stoichiometry, rather than Pd^0 with adsorbed oxygen (Pd-O_{ads}), which is characterized by a E_b shift of less than 1 eV.²⁵ Atomic concentrations of Pd^0 , Pd^{2+} , and Pd^{4+} calculated from the Pd 3d core levels are $19 \pm 7\%$, $64 \pm 3\%$, and $18 \pm 4\%$, respectively. The presence of Pd in the 0, 2+ and 4+ valence states reinforces the proposed mechanism for $\text{PdO}_x/\text{PEDOT}$ formation, and suggests an equilibrium favouring Pd^{2+} in what is most likely to be the PdO stoichiometry or a hydrated analogue (e.g. $\text{Pd}(\text{OH})_2$, $\text{PdO} \cdot n\text{H}_2\text{O}$). Full XPS quantitative analysis, E_b values, and survey spectra are provided in Fig. S3 and Tables S1 and S2.†

Based on previous work on the redox behaviour of Pd-based films, Pd oxides have been classified by two types (α and β) with characteristic oxide-reduction profiles.¹⁹ Thus, the redox potentials and valence states of the $\text{PdO}_x/\text{PEDOT}$ film can also be revealed by examining oxide-reduction profiles by cyclic voltammetry (CV). α - PdO_x is characterized by Pd^{2+} valence in an anhydrous monolayer or dense multilayer structure and exhibits a reduction wave between 0.6 and 0.9 V vs. RHE, representative of the $\text{Pd}^{2+/0}$ couple.¹⁹ In contrast, β - PdO_x is typically amorphous, hydrated, and porous with valence assigned to Pd^{2+} , Pd^{4+} , or Pd^{6+} , exhibiting a reduction wave negative of 0.6 V vs. RHE.¹⁹ This wave has been interpreted as the reduction of both hydrated Pd^{2+} and Pd^{4+} oxides to Pd^0 .^{26,27}

An as-deposited $\text{PdO}_x/\text{PEDOT}$ film was transferred to a three-electrode cell and cycled in Ar-saturated 0.1 M KOH electrolyte between 0.03 and 1.23 V vs. RHE (Fig. 2c). Beginning from the open circuit potential (OCP) of approximately 1.1 V vs. RHE, the first cathodic scan shows a weak reduction wave with peak potential (E_p) of 0.65 V vs. RHE (C_1) and an intense reduction wave at 0.33 V vs. RHE (C_2), followed by a broad oxidation wave A_1 (E_p of 1.0 V vs. RHE) in the anodic scan. C_1 is indicative of α - PdO_x reduction and the dominant C_2 peak is associated with the reduction of the β - PdO_x species.^{19,26,27} A mixed $\text{Pd}^{2+}/\text{Pd}^{4+}$ oxide is also suggested by the OCP of the PEDOT/ PdO_x film, 1.1 V vs. RHE, a mixed-potential between the formal potentials of the $\text{Pd}^{2+/0}$ (0.9 V vs. RHE) and $\text{Pd}^{4+/2+}$ (1.2–1.28 V vs. RHE) couples.^{19,24} The OCP and reduction profile on the first cathodic scan are also consistent with the average quantitative analysis by XPS on $\text{PdO}_x/\text{PEDOT}$ films as-deposited (see above).

Further cycling demonstrates a decrease in current associated with C_2 and corresponding current increase and stabilization of C_1 (shifts to E_p of ~ 0.6 V vs. RHE), as shown in Fig. 2c. Hence, over time (and with cycling) the Pd oxide species on PEDOT becomes less oxidized and stabilizes in the α - PdO_x (Pd^{2+}) form with reversibility upon reduction to Pd^0 (C_1/A_1). It is worth noting that the β - PdO_x species can be regenerated and cycled if the anodic switching potential is extended to 1.5 V vs. RHE (Fig. S4†). The stability and electrochemical reversibility of the Pd^0/α - PdO_x species in the ORR potential window suggests it may act as an ORR-active species. However, the behaviour of the $\text{PdO}_x/\text{PEDOT}$ film in O_2 -saturated electrolyte using CV and Tafel analysis must also be considered to confirm the active species as Pd oxide films are known to exhibit different speciation and adsorption properties in the presence of O_2 .²⁸

The electrocatalytic activity of the r-PEDOT and $\text{PdO}_x/\text{PEDOT}$ films toward the ORR was investigated using CV and rotating

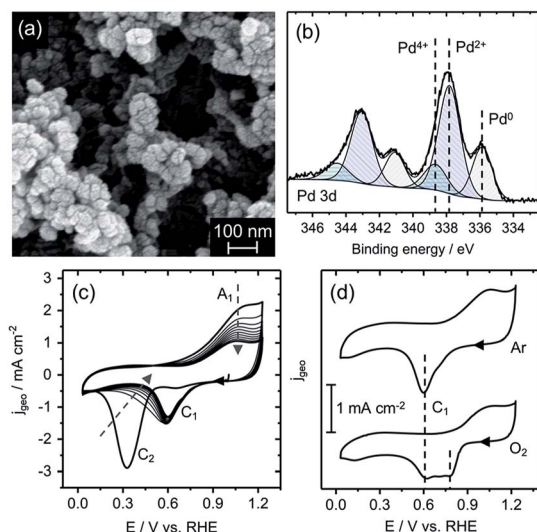


Fig. 2 (a) SEM image of a $\text{PdO}_x/\text{PEDOT}$ film as-deposited on glassy carbon; (b) representative XPS Pd 3d spectrum and symmetric core-level fits with valence assignments for a $\text{PdO}_x/\text{PEDOT}$ film as-deposited; (c) CV scans of a $\text{PdO}_x/\text{PEDOT}$ film in Ar-saturated 0.1 M KOH electrolyte, 50 mV s^{-1} ; (d) CV scans of a $\text{PdO}_x/\text{PEDOT}$ film sequentially cycled in Ar-saturated and O_2 -saturated 0.1 M KOH, 50 mV s^{-1} .

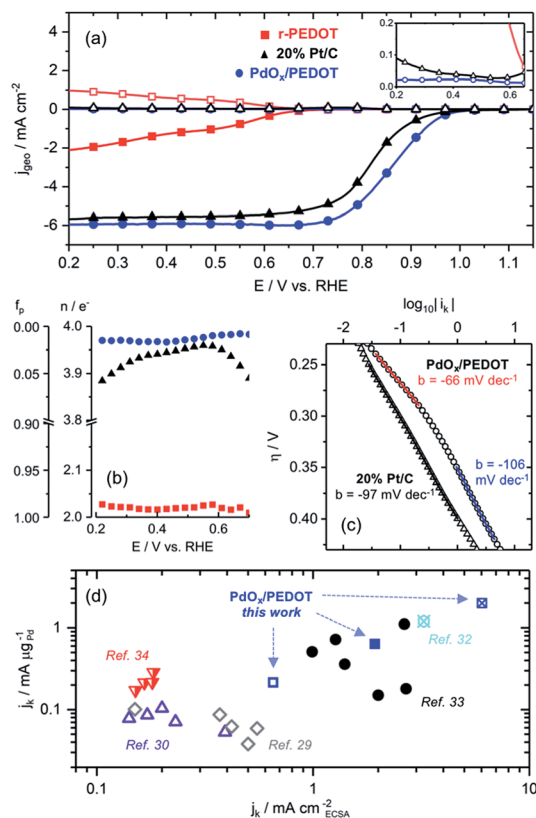


Fig. 3 (a) ORR LSVs measured at the disk (closed symbols) and ring (open symbols) electrodes in 0.1 M KOH electrolyte at 2500 RPM, inset: zoom view of ring currents; (b) n and f_p parameters calculated from disk and ring current in 0.1 M KOH electrolyte at 2500 RPM; (c) Tafel plots and slopes constructed from the kinetic ORR current in the range of ± 100 mV relative to E_{onset} ; (d) comparison of specific ORR activity on the basis of ECSA and Pd mass loading of the PdO_x/PEDOT film and Pd-based electrocatalysts in 0.1 M KOH: 0.8 V vs. RHE (strikethrough symbols), 0.85 V vs. RHE (closed symbols), 0.9 V vs. RHE (open symbols), unspecified potential (half-filled symbols).

ring disk electrode (RRDE) linear scanning voltammetry (LSV) methods. A cathodic scan of the PdO_x/PEDOT film in O₂-saturated 0.1 M KOH produced a catalytic wave with E_p of ~ 0.8 V vs. RHE, a 0.25 V reduction in ORR overpotential (η_{ORR}) relative to the r-PEDOT film, $E_p = 0.55$ V vs. RHE (Fig. 2d and S5†).[‡] Furthermore, the reduction of η_{ORR} is evident in RRDE LSVs, where the characteristic onset potential (E_{onset}) and half-wave potential ($E_{1/2}$) of PdO_x/PEDOT are 0.98 V and 0.86 V vs. RHE, respectively, compared to 0.7 V and 0.49 V vs. RHE for r-PEDOT (Fig. 3a). The PdO_x/PEDOT film also notably demonstrates a shift in E_{onset} and $E_{1/2}$ to lower η_{ORR} by 30–40 mV relative to the commercial 20% Pt/C catalyst film ($E_{\text{onset}} = 0.95$ V vs. RHE; $E_{1/2} = 0.82$ V vs. RHE).[§] Long-term stability was investigated by electrolysis for 3 h at 0.8 V vs. RHE; the PdO_x/PEDOT film

[‡] For a comparison to the ORR activity of an oxidized PEDOT film under the conditions reported here, refer to ref. 16.

[§] The ORR activity for the commercial 20% Pt/C electrocatalyst is the highest obtained in our hands. The activity is consistent with a range of reported values, where alkaline electrolytes are known to inhibit the ORR activity on Pt surfaces compared to acid (see ref. 2 and 18 for further discussion).

retained 87% of the initial ORR current, compared to 78% retention for 20% Pt/C (Fig. S6†).

The ORR electron transfer number (n) and fraction of peroxide product (f_p) were calculated using disk and ring currents measured by RRDE (see ESI† for calculations). These parameters provide insight regarding the ORR pathway, where the limits $n = 2$ and $n = 4$ correspond with the 2e⁻ reduction of O₂ to peroxide and the 4e⁻ reduction of O₂ to OH⁻, respectively, while n values of $2 < n < 4$ suggest a mixed pathway.² The increase from $n \approx 2$ for r-PEDOT to $n = 3.97$ for PdO_x/PEDOT (averaged over 0.2 V to 0.9 V vs. RHE) confirms the Pd species are the active site for catalysing efficient 4e⁻ reduction of O₂ to OH⁻ (Fig. 3b). The PdO_x/PEDOT composite was also more selective to the 4e⁻ pathway than commercial 20% Pt/C ($3.87 < n < 3.97$, Fig. 3b). Koutecky–Levich (K–L) analysis was employed to confirm the RRDE results shown in Fig. 3b (see ESI for calculations; Fig. S7†). Average n values calculated using the K–L equation and plots were 3.9 ± 0.1 for PdO_x/PEDOT and 3.86 ± 0.09 commercial 20% Pt/C. Furthermore, the calculated kinetic rate constant for the ORR is higher for PdO_x/PEDOT (0.12 ± 0.03 cm s⁻¹) than for commercial 20% Pt/C (0.07 ± 0.02 cm s⁻¹).

The relationship between kinetic ORR current (i_k) and overpotential is well established for Pd surfaces, where the rate-limiting step is proposed to be the first electron transfer to O₂.^{17,29–31} The Tafel slope (b) for PdO_x/PEDOT changes from -66 mV dec⁻¹ at low overpotential to -106 mV dec⁻¹ at higher current densities after E_{onset} (Fig. 3c). This behaviour is consistent with the literature, where b is expected to change from -60 mV dec⁻¹ to -120 mV dec⁻¹ in the kinetic-limited regime near E_{onset} , owing to the reduction or protonation of the surface oxides at higher ORR current densities.^{17,31} Considered together with CV profiles in Ar- and O₂-saturated 0.1 M KOH (Fig. 2c and d), we propose the following ORR-active species for PdO_x/PEDOT that correspond to distinct potential ranges. In the low current density region positive of E_{onset} (~ 1.23 – 0.98 V vs. RHE), the ORR is catalysed by the multi-layer α -PdO_x identified by CV (Fig. 2c and d) and with $b = -66$ mV dec⁻¹ (Fig. 3c). In the following high current density regime still under kinetic control (~ 0.98 – 0.85 V vs. RHE), the surface α -PdO_x species interacting with O₂ and adsorbed oxygen is reduced or protonated, as evidenced by the change in b to -106 mV dec⁻¹ (Fig. 3c). As the potential decreases below 0.85 V vs. RHE, bulk α -PdO_x is reduced to Pd⁰ and diffusion-limited ORR conditions are reached. E_p for the reduction of α -PdO_x to Pd⁰ is ~ 0.6 V vs. RHE in the absence and presence of O₂, and the film appears to be fully reduced to Pd⁰ by 0.4 V vs. RHE (Fig. 2d). The reduced bulk Pd⁰ likely remains covered with adsorbed oxide or hydroxide considering the high pH and constant n value between 0.7 and 0.2 V vs. RHE (Fig. 3b).

Finally, i_k was normalized on the basis of electrochemical surface area (ECSA) and Pd mass loading to provide measures of specific activity. The ECSA of the PdO_x/PEDOT film based on Pd oxide reduction as determined by CV was 1.05 cm² (see ESI† for calculations). The specific kinetic current density (j_k) of the hybrid PdO_x/PEDOT film based on ECSA (mA cm_{ECSA}⁻²) and Pd loading (mA μg_{Pd}⁻¹) between 0.8 and 0.9 V vs. RHE is shown in Fig. 3d. For comparison, Fig. 3d and Table S3† include reported

Table 1 ORR activity of hybrid PEDOT-based electrocatalysts in 0.1 M KOH

| | E_{onset} (V) | $E_{1/2}$ (V) | $j_k^{a,b}$ ($\text{mA } \mu\text{g}_M^{-1}/\text{mA cm}_{\text{ECSA}}^{-2}$) | Ref. |
|-------------------------|------------------------|-----------------|---|-----------|
| PdO _x /PEDOT | 0.98 | 0.86 | 1.99/6.03 | This work |
| P-MnO _x -20 | 0.87 | 0.83 | NR ^c | 16 |
| Pd/PEDOT/rGO | 0.98 | NR ^c | 1.2/3.25 | 32 |

^a at 0.8 V vs. RHE. ^b M = metal or metal oxide (Pd, MnO_x). ^c NR = not reported.

j_k values for state-of-the-art Pd-based electrocatalysts in 0.1 M KOH electrolyte.^{29,30,32–34} The combined mass- and ECSA-based specific activity of the hybrid PdO_x/PEDOT film is among the highest reported to date; for example, at 0.9 V vs. RHE (0.65 mA cm_{ECSA}⁻², 0.22 mA $\mu\text{g}_{\text{Pd}}^{-1}$), PdO_x/PEDOT surpasses the activity of size-controlled Pd nanocubes²⁹ and nanoparticles³⁰ supported on conductive carbon. In addition, the mass-specific activity of PdO_x/PEDOT is superior to commercial 20% Pt/C at 0.9 V vs. RHE (0.14 mA $\mu\text{g}_{\text{Pt}}^{-1}$).

Table 1 provides a comparison of the most active PEDOT-based ORR electrocatalysts reported in 0.1 M KOH.^{16,32} The E_{onset} values for the Pd-based composites are markedly positive of that for the MnO_x composite (by ~100 mV), suggesting significantly higher intrinsic activities at low η_{ORR} on Pd. Despite this, $E_{1/2}$ values for both PdO_x/PEDOT and the P-MnO_x-20 film¹⁶ are >0.8 V vs. RHE and within 30 mV, indicating fast kinetics for both systems in the kinetic-limited regime. The specific activities of PdO_x/PEDOT represent a >150% improvement in those reported for the Pd/PEDOT/rGO catalyst³² on the basis of both ECSA and Pd mass (also shown in Fig. 3d). This is notable considering the present work eliminates the need for additional conductive carbon (*e.g.* rGO), demonstrating the promise of PEDOT to act as both a conductive support and polymeric binder. Analogous to the highly active P-MnO_x-20 catalyst films,¹⁶ we attribute the high ORR activity of the PdO_x/PEDOT films to the electroless deposition method, which affords the following: (i) a porous nanostructure that improves surface area and mass transfer for catalysis; (ii) deposition of a stable, electrochemically reversible ORR-active Pd⁰/ α -PdO_x species; (iii) high dispersion and low-mass loading of exposed Pd species maximizes active site utilization and specific activity.

Conclusions

Nanostructured hybrid PdO_x/PEDOT films were prepared by a spontaneous reaction between aqueous PdCl₄²⁻ ions and reduced PEDOT films in neutral aqueous solution. The Pd species as-deposited on PEDOT was characterized for the first time by XPS and CV, revealing Pd in the 0, 2+ and 4+ valence states. Thus, we propose the spontaneous redox reactions proceed to form PEDOT-bound Pd⁰ with a mixed-valence Pd²⁺/Pd⁴⁺ oxide. Cycling the hybrid PdO_x/PEDOT films in the absence (and presence) of O₂ in alkaline electrolyte reveals a stable Pd⁰/ α -PdO_x species, which acts as the active site in electrocatalysis of the ORR. The PdO_x/PEDOT films display high activity and selectivity toward efficient 4e⁻ ORR, competitive with or superior to the activity of state-of-the-art Pd/C and PEDOT-based

composites and commercial 20% Pt/C, demonstrating the promise of the dual utility of PEDOT as a conductive substrate and polymeric binder.

Conflicts of interest

There are no conflicts to declare.

Acknowledgements

Sandia National Laboratories is a multi-mission laboratory managed and operated by National Technology and Engineering Solutions of Sandia, LLC, a wholly owned subsidiary of Honeywell International, Inc., for the U.S. Department of Energy's National Nuclear Security Administration under contract DE-NA0003525. The views expressed in the article do not necessarily represent the views of the U.S. Department of Energy or the United States Government.

Notes and references

- V. R. Stamenkovic, D. Strmcnik, P. P. Lopes and N. M. Markovic, *Nat. Mater.*, 2016, **16**, 57–69.
- Q. He and E. J. Cairns, *J. Electrochem. Soc.*, 2015, **162**, F1504–F1539.
- T. N. Lambert, J. A. Vigil, S. E. White, D. J. Davis, S. J. Limmer, P. D. Burton, E. N. Coker, T. E. Beechem and M. T. Brumbach, *Chem. Commun.*, 2015, **51**, 9511–9514.
- T. N. Lambert, J. A. Vigil, S. E. White, C. J. Delker, D. J. Davis, M. Kelly, M. T. Brumbach, M. A. Rodriguez and B. S. Swartzentruber, *J. Phys. Chem. C*, 2017, **121**, 2789–2797.
- J. A. Vigil, T. N. Lambert, J. Duay, C. J. Delker, T. E. Beechem and B. S. Swartzentruber, *ACS Appl. Mater. Interfaces*, 2018, **10**, 2040–2050.
- D. J. Davis, A.-R. O. Raji, T. N. Lambert, J. A. Vigil, L. Li, K. Nan and J. M. Tour, *Electroanalysis*, 2014, **26**, 164–170.
- E. Antolini and E. R. Gonzalez, *Appl. Catal., A*, 2009, **365**, 1–19.
- M. C. Lefebvre, Z. Qi and P. G. Pickup, *J. Electrochem. Soc.*, 1999, **146**, 2054–2058.
- B. Winther-Jensen, O. Winther-Jensen, M. Forsyth and D. R. MacFarlane, *Science*, 2008, **321**, 671.
- R. Kerr, C. Pozo-Gonzalo, M. Forsyth and B. Winther-Jensen, *ECS Electrochem. Lett.*, 2013, **2**, F29–F31.
- E. Mittra, M. J. Jafari, M. Vagin, X. Liu, M. Fahlman, T. Ederth, M. Berggren, M. P. Jonsson and X. Crispin, *J. Mater. Chem. A*, 2017, **5**, 4404–4412.

- 12 A. D. Chowdhury, N. Agnihotri, P. Sen and A. De, *Electrochim. Acta*, 2014, **118**, 81–87.
- 13 G. Gnana kumar, C. Joseph Kirubaharan, D. J. Yoo and A. R. Kim, *Int. J. Hydrogen Energy*, 2016, **41**, 13208–13219.
- 14 R. Kerr, C. Pozo-Gonzalo, M. Forsyth and B. Winther-Jensen, *Electrochim. Acta*, 2015, **154**, 142–148.
- 15 J. A. Vigil, T. N. Lambert and K. Eldred, *ACS Appl. Mater. Interfaces*, 2015, **7**, 22745–22750.
- 16 J. A. Vigil, T. N. Lambert, M. Kelly and R. Aidun, *Mater. Chem. Front.*, 2017, **1**, 1668–1675.
- 17 D. Diabaté, T. W. Napporn, K. Servat, A. Habrioux, S. Arriclaens, A. Trokourey and K. B. Kokoh, *J. Electrochem. Soc.*, 2013, **160**, H302–H308.
- 18 F. H. B. Lima, J. Zhang, M. H. Shao, K. Sasaki, M. B. Vukmirovic, E. A. Ticianelli and R. R. Adzic, *J. Phys. Chem. C*, 2007, **111**, 404–410.
- 19 M. Grdeń, M. Łukaszewski, G. Jerkiewicz and A. Czerwiński, *Electrochim. Acta*, 2008, **53**, 7583–7598.
- 20 S. N. Eliseeva, V. V. Malev and V. V. Kondratiev, *Russ. J. Electrochem.*, 2009, **45**, 1045.
- 21 S. N. Eliseeva, E. V. Ubyivovk, A. S. Bondarenko, O. F. Vyvenko and V. V. Kondratiev, *Russ. J. Gen. Chem.*, 2010, **80**, 1143–1148.
- 22 V. V. Kondratiev, T. A. Babkova and S. N. Eliseeva, *Russ. J. Electrochem.*, 2012, **48**, 205–211.
- 23 A. Bund and S. Neudeck, *J. Phys. Chem. B*, 2004, **108**, 17845–17850.
- 24 J. P. Hoare, *J. Electrochem. Soc.*, 1964, **111**, 610–615.
- 25 K. S. Kim, A. F. Gossmann and N. Winograd, *Anal. Chem.*, 1974, **46**, 197–200.
- 26 V. I. Birss, M. Chan, T. Phan, P. Vanysek and A. Zhang, *J. Chem. Soc. Faraday Trans.*, 1996, **92**, 4041–4047.
- 27 A. E. Bolzán and A. J. Arvia, *J. Electroanal. Chem.*, 1992, **322**, 247–265.
- 28 L. M. Vracar, D. B. Sepa and A. Damjanovic, *J. Electrochem. Soc.*, 1989, **136**, 1973–1977.
- 29 M. Lüsi, H. Erikson, A. Sarapuu, K. Tammeveski, J. Solla-Gullón and J. M. Feliu, *Electrochem. Commun.*, 2016, **64**, 9–13.
- 30 L. Jiang, A. Hsu, D. Chu and R. Chen, *J. Electrochem. Soc.*, 2009, **156**, B643–B649.
- 31 L. M. Vracar, D. B. Sepa and A. Damjanovic, *J. Electrochem. Soc.*, 1986, **133**, 1835–1839.
- 32 J. E. Choe, M. S. Ahmed and S. Jeon, *J. Power Sources*, 2015, **281**, 211–218.
- 33 Q. Wu, Z. Rao, L. Yuan, L. Jiang, G. Sun, J. Ruan, Z. Zhou and S. Sang, *Electrochim. Acta*, 2014, **150**, 157–166.
- 34 W. Yan, Z. Tang, L. Li, L. Wang, H. Yang, Q. Wang, W. Wu and S. Chen, *ChemElectroChem*, 2017, **4**, 1349–1355.