metal-organic compounds

4-Aminopyridinium *trans*-diaqua-dioxalatochromate(III) monohydrate

Ichraf Chérif, Jawher Abdelhak, Mohamed Faouzi Zid* and Ahmed Driss

Laboratoire de Matériaux et Cristallochimie, Faculté des Sciences de Tunis, Université de Tunis ElManar, 2092 Manar II Tunis, Tunisia
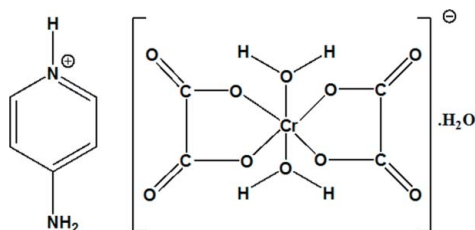Correspondence e-mail: faouzi.zid@fst.rnu.tn

In the non-centrosymmetric structure of the title compound, $(C_5H_7N_2)[Cr(C_2O_4)_2(H_2O)_2] \cdot H_2O$, the Cr^{III} ion has a slightly distorted octahedral coordination environment defined by two chelating oxalato ligands in equatorial positions and two water molecules in axial positions. An extensive three-dimensional network of hydrogen bonds involving all the water molecules, the 4-aminopyridinium cation and some of the oxalate O atoms contributes to the stabilization of the structure. π–π interactions between adjacent pyridine rings provide additional stability of the crystal packing, with a closest distance between pyridine mean planes of 3.613 (1) Å.

Related literature

For the structural characterization of salts containing the $[Cr(C_2O_4)_2(H_2O)_2]^-$ anion with various counter-cations, see: Bélombé *et al.* (2009); Nenwa *et al.* (2010). For oxalate coordination modes, see: Tang *et al.* (2002); Martak *et al.* (2009); Hernández-Molina *et al.* (2001); Zhao *et al.* (2004). For C—O distances in oxalate anions, see: Marinescu *et al.* (2000). For geometric parameters of the 4-aminopyridinium cation, see: Fun *et al.* (2008, 2009, 2010); Jebas *et al.* (2009); Quah *et al.* (2008); Ramesh *et al.* (2010); Rotondo *et al.* (2009); Pan *et al.* (2008). For discussion of hydrogen bonding, see: Blessing (1986); Brown (1976).

Experimental

Crystal data

$(C_5H_7N_2)[Cr(C_2O_4)_2(H_2O)_2] \cdot H_2O$	V = 1446.7 (7) Å³
M_r = 377.21	Z = 4
Orthorhombic, $Pna2_1$	Mo $K\alpha$ radiation
a = 21.102 (4) Å	μ = 0.85 mm⁻¹
b = 9.487 (3) Å	T = 298 K
c = 7.226 (2) Å	$0.54 \times 0.18 \times 0.12$ mm

Data collection

Enraf–Nonius CAD-4 diffractometer	3150 independent reflections
Absorption correction: ψ scan (North *et al.*, 1968)	2857 reflections with $I > 2\sigma(I)$
T_{min} = 0.832, T_{max} = 0.903	R_{int} = 0.017
3729 measured reflections	2 standard reflections every 120 min intensity decay: 2.9%

Refinement

$R[F^2 > 2\sigma(F^2)]$ = 0.025	All H-atom parameters refined
$wR(F^2)$ = 0.066	$\Delta\rho_{max}$ = 0.30 e Å⁻³
S = 1.04	$\Delta\rho_{min}$ = −0.31 e Å⁻³
3150 reflections	Absolute structure: Flack (1983), 1447 Friedel pairs
261 parameters	Flack parameter: 0.000 (16)
1 restraint	

Table 1
Selected bond lengths (Å).

Cr1—O1	1.9636 (15)	Cr1—O3	1.9810 (16)
Cr1—O4	1.9658 (15)	Cr1—OW1	1.9847 (17)
Cr1—O2	1.9715 (16)	Cr1—OW2	2.0079 (18)

Table 2
Hydrogen-bond geometry (Å, °).

D—H$\cdots A$	D—H	H$\cdots A$	$D\cdots A$	D—H$\cdots A$
OW3—HW31$\cdots$O8^i	0.78 (4)	2.06 (4)	2.823 (3)	167 (4)
OW3—HW32$\cdots$O8^{ii}	0.93 (4)	1.93 (4)	2.843 (3)	169 (4)
OW2—HW21$\cdots$O7^{iii}	0.89 (4)	1.78 (4)	2.640 (3)	161 (3)
OW2—HW22$\cdots$O9^{iv}	0.74 (4)	2.07 (4)	2.784 (3)	162 (4)
OW1—HW11$\cdots$OW3^v	0.90 (4)	1.67 (4)	2.572 (3)	172 (4)
OW1—HW12$\cdots$O10^{vi}	0.76 (5)	2.01 (5)	2.745 (3)	162 (5)
N2—H2$\cdots$O10^{vii}	0.93 (4)	2.01 (4)	2.940 (4)	179 (4)
N1—H1$\cdots$O3	1.00 (4)	2.27 (4)	3.061 (3)	135 (4)
N1—H1$\cdots$O9^{vi}	1.00 (4)	2.22 (4)	2.974 (3)	131 (4)

Symmetry codes: (i) $x, y + 1, z + 1$; (ii) $-x, -y + 1, z + \frac{1}{2}$; (iii) $-x - \frac{1}{2}, y + \frac{1}{2}, z + \frac{1}{2}$; (iv) $x, y, z - 1$; (v) $x, y - 1, z$; (vi) $-x - \frac{1}{2}, y - \frac{1}{2}, z - \frac{1}{2}$; (vii) $-x - 1, -y, z - \frac{1}{2}$.

Data collection: *CAD-4 EXPRESS* (Duisenberg, 1992; Macíček & Yordanov, 1992); cell refinement: *CAD-4 EXPRESS*; data reduction: *XCAD4* (Harms & Wocadlo, 1995); program(s) used to solve structure: *SHELXS97* (Sheldrick, 2008); program(s) used to refine structure: *SHELXL97* (Sheldrick, 2008); molecular graphics: *DIAMOND* (Brandenburg, 1998); software used to prepare material for publication: *WinGX* (Farrugia, 1999).

Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: WM2545).

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supplementary materials

4-Aminopyridinium *trans*-diaquadioxalatochromate(III) monohydrate

I. Chérif, J. Abdelhak, M. F. Zid and A. Driss

Comment

It is well established that the strategy of utilizing transition metals with versatile multidentate ligands like oxalate anions affords various structural topologies (Tang *et al.*, 2002; Martak *et al.*, 2009; Hernández-Molina *et al.*, 2001; Zhao *et al.*, 2004). In this article, we describe the crystal structure of a new chromium(III) salt, $(C_5H_7N_2)[Cr(C_2O_4)_2(H_2O)_2]\cdot H_2O$, containing oxalate (ox) anions which act as chelating ligands.

The structure of the title compound consists of discrete $[Cr(C_2O_4)_2(H_2O)_2]^-$ mononuclear anions, 4-aminopyridinium cations and uncoordinated water molecules (Fig. 1). In the $[Cr(C_2O_4)_2(H_2O)_2]^-$ anion, the Cr^{III} atom is six-coordinated in a slightly distorted octahedral coordination environment defined by two chelating bidentate oxalate anions and two water molecules in *trans* position. The expected 90° bond angles vary from 82.13 (6)° to 101.02 (6)°, and the ideal 180° bond angles vary from 175.74 (7)° to 178.68 (8)°. The four Cr—O$_{(ox)}$ bond lengths are different and slightly shorter than the Cr—O$_{(water)}$ bond lengths. This situation was previously observed in homologous salts involving quinolinium $(C_9H_8N)^+$ and 4-dimethylaminopyridinium $(C_7H_{11}N_2)^+$ counter cations (Bélombé *et al.*, 2009; Nenwa *et al.*, 2010). The C—C bond lengths in the oxalate ligand are as expected for single C—C bonds [1.553 (3) and 1.558 (3) Å for C1—C2 and C3—C4, respectively]. The bond length values of the peripheral and inner C—O bonds compare well with those reported for other oxalate complexes (Marinescu *et al.*, 2000), the shorter values being due to the greater double bond character of the free C—O bonds. The uncoordinated 4-aminopyridinium cations are located between the anions (Fig. 2), their geometric parameters do not show unusual features, they are in accordance with those previously reported (Fun *et al.*, 2008, 2009, 2010; Jebas *et al.*, 2009; Quah *et al.*, 2008; Ramesh *et al.*, 2010; Rotondo *et al.*, 2009; Pan *et al.*, 2008).

The crystal structure is stabilized by strong intermolecular hydrogen bonds (Blessing, 1986; Brown, 1976). $(C_5H_7N_2)^+$ and $[Cr(C_2O_4)_2(H_2O)_2]^-$ are connected *via* N—H⋯O hydrogen bonds that link the amino N2—H2 group to the free O10 oxalato O atoms. The uncoordinated water molecule functions as both acceptor and donor while the coordinated water molecules function only as donors. Thus, all the constituents of $(C_5H_7N_2)[Cr(C_2O_4)_2(H_2O)_2]\cdot H_2O$ are connected to form a three-dimensional structure. Furthermore, π—π interactions between adjacent pyridine rings help to stabilize the crystal packing, the closest distance between two pyridine mean planes being 3.613 (1) Å (Fig. 3).

Experimental

Aqueous solutions of oxalic acid dihydrate, $H_2C_2O_4\cdot 2H_2O$ (2 mmol), and 4-aminopyridine (1 mmol) were added to $Cr(NO_3)_3\cdot 9H_2O$ (1 mmol) dissolved in 10 mL of water under continuous stirring at 323K. Slow evaporation of the resultant solution led to violet single crystals suitable for X-ray diffraction.

supplementary materials

Refinement

The hydrogen atoms were located in difference Fourier maps. Their displacement parameters were refined isotropically.
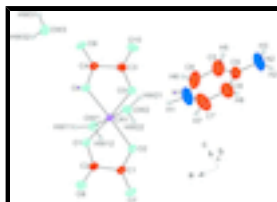
Figures

Fig. 1. : The asymmetric unit of $(C_5H_7N_2)[Cr(C_2O_4)_2(H_2O)_2]\cdot H_2O$. Thermal displacement parameters are drawn at the 50% probability level; H atoms are given as spheres of arbitrary radius.
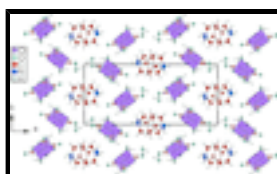

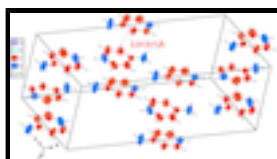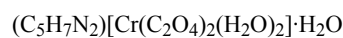Fig. 2. : Projection of the $(C_5H_7N_2)[Cr(C_2O_4)_2(H_2O)_2]\cdot H_2O$ structure along the *c* axis.

Fig. 3. : Perspective view showing the π—π interactions between adjacent 4-aminopyridinium groups. $[Cr(C_2O_4)_2(H_2O)_2]^-$ anions and uncoordinated water molecules have been omitted for clarity.

4-Aminopyridinium *trans*-diaquadioxalatochromate(III) monohydrate

Crystal data

$(C_5H_7N_2)[Cr(C_2O_4)_2(H_2O)_2]\cdot H_2O$	$F(000) = 772$
$M_r = 377.21$	$D_x = 1.732$ Mg m^{-3}
Orthorhombic, $Pna2_1$	Mo $K\alpha$ radiation, $\lambda = 0.71073$ Å
Hall symbol: P 2c -2n	Cell parameters from 25 reflections
$a = 21.102$ (4) Å	$\theta = 10–15°$
$b = 9.487$ (3) Å	$\mu = 0.85$ mm^{-1}
$c = 7.226$ (2) Å	$T = 298$ K
$V = 1446.7$ (7) Å^3	Prism, violet
$Z = 4$	$0.54 \times 0.18 \times 0.12$ mm

Data collection

Enraf–Nonius CAD-4 diffractometer	2857 reflections with $I > 2\sigma(I)$
Radiation source: fine-focus sealed tube	$R_{int} = 0.017$
graphite	$\theta_{max} = 27.0°$, $\theta_{min} = 2.4°$
ω/2θ scans	$h = -26 \rightarrow 1$
Absorption correction: ψ scan (North *et al.*, 1968)	$k = -1 \rightarrow 12$

$T_{min} = 0.832$, $T_{max} = 0.903$

3729 measured reflections

3150 independent reflections

$l = -9 \rightarrow 9$

2 standard reflections every 120 min

intensity decay: 2.9%

Refinement

Refinement on F^2	Hydrogen site location: difference Fourier map
Least-squares matrix: full	All H-atom parameters refined
$R[F^2 > 2\sigma(F^2)] = 0.025$	$w = 1/[\sigma^2(F_o^2) + (0.0364P)^2 + 0.1566P]$ where $P = (F_o^2 + 2F_c^2)/3$
$wR(F^2) = 0.066$	$(\Delta/\sigma)_{max} < 0.001$
$S = 1.04$	$\Delta\rho_{max} = 0.30$ e Å^{-3}
3150 reflections	$\Delta\rho_{min} = -0.31$ e Å^{-3}
261 parameters	Extinction correction: *SHELXL97* (Sheldrick, 2008), $Fc^* = kFc[1 + 0.001xFc^2\lambda^3/\sin(2\theta)]^{-1/4}$
1 restraint	Extinction coefficient: 0.0022 (7)
Primary atom site location: structure-invariant direct methods	Absolute structure: Flack (1983), 1447 Friedel pairs
Secondary atom site location: difference Fourier map	Flack parameter: 0.000 (16)

Special details

Geometry. All esds (except the esd in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell esds are taken into account individually in the estimation of esds in distances, angles and torsion angles; correlations between esds in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell esds is used for estimating esds involving l.s. planes.

Refinement. Refinement of F^2 against ALL reflections. The weighted R-factor wR and goodness of fit S are based on F^2, conventional R-factors R are based on F, with F set to zero for negative F^2. The threshold expression of $F^2 > 2\text{sigma}(F^2)$ is used only for calculating R-factors(gt) etc. and is not relevant to the choice of reflections for refinement. R-factors based on F^2 are statistically about twice as large as those based on F, and R- factors based on ALL data will be even larger.

Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters (Å^2)

	x	y	z	$U_{iso}*/U_{eq}$
Cr1	−0.196274 (13)	0.09444 (3)	−0.00116 (5)	0.01985 (9)
O9	−0.19814 (7)	0.26620 (15)	0.5002 (3)	0.0295 (3)
O3	−0.27673 (7)	0.09078 (16)	0.1400 (2)	0.0262 (3)
O1	−0.11408 (7)	0.10787 (17)	−0.1273 (2)	0.0264 (3)
OW1	−0.17053 (9)	−0.08746 (18)	0.1118 (2)	0.0283 (3)
O4	−0.16685 (7)	0.19731 (16)	0.2188 (2)	0.0248 (3)
OW2	−0.22205 (8)	0.27650 (18)	−0.1212 (2)	0.0295 (3)
O7	−0.17687 (8)	−0.09782 (17)	−0.4830 (3)	0.0340 (4)
O2	−0.22078 (7)	−0.00965 (16)	−0.2262 (2)	0.0243 (3)
O10	−0.31987 (7)	0.17303 (18)	0.4006 (2)	0.0337 (4)
O8	−0.06261 (8)	0.0179 (2)	−0.3698 (3)	0.0430 (5)
C3	−0.27494 (10)	0.1550 (2)	0.2972 (3)	0.0228 (4)
C1	−0.17435 (10)	−0.0316 (2)	−0.3382 (3)	0.0222 (4)

C4	−0.20797 (10)	0.2123 (2)	0.3479 (3)	0.0218 (4)
C2	−0.11043 (10)	0.0354 (2)	−0.2779 (3)	0.0252 (4)
C6	−0.42595 (16)	0.1102 (4)	0.0242 (6)	0.0625 (10)
C8	−0.51027 (14)	−0.1061 (3)	0.0042 (7)	0.0524 (7)
N1	−0.40576 (12)	−0.0224 (3)	0.0161 (6)	0.0657 (8)
N2	−0.59497 (12)	0.0580 (3)	0.0089 (8)	0.0750 (10)
C9	−0.53327 (11)	0.0317 (3)	0.0117 (5)	0.0413 (5)
C7	−0.44661 (16)	−0.1296 (4)	0.0020 (8)	0.0726 (10)
OW3	−0.06036 (9)	0.9143 (3)	0.2640 (3)	0.0449 (5)
C5	−0.48794 (15)	0.1402 (3)	0.0211 (7)	0.0597 (8)
H8	−0.5357 (18)	−0.181 (4)	0.003 (6)	0.079 (11)*
H2	−0.6222 (18)	−0.014 (4)	−0.026 (6)	0.076 (12)*
H1	−0.359 (2)	−0.038 (5)	0.028 (7)	0.098 (14)*
HW31	−0.0613 (16)	0.955 (4)	0.358 (6)	0.047 (10)*
HW21	−0.2581 (18)	0.323 (4)	−0.100 (5)	0.068 (11)*
HW11	−0.1303 (19)	−0.085 (4)	0.155 (5)	0.058 (10)*
H6	−0.3942 (19)	0.183 (4)	0.031 (6)	0.079 (12)*
HW32	−0.022 (2)	0.948 (4)	0.220 (6)	0.077 (12)*
H5	−0.5032 (19)	0.233 (4)	0.043 (5)	0.077 (12)*
H3	−0.606 (2)	0.147 (5)	0.010 (7)	0.094 (13)*
H7	−0.428 (2)	−0.226 (5)	0.001 (8)	0.110 (15)*
HW12	−0.180 (2)	−0.146 (5)	0.045 (7)	0.094 (16)*
HW22	−0.2234 (18)	0.269 (4)	−0.223 (6)	0.070 (13)*

Atomic displacement parameters ($Å^2$)

	U^{11}	U^{22}	U^{33}	U^{12}	U^{13}	U^{23}
Cr1	0.01992 (14)	0.02436 (15)	0.01528 (14)	−0.00098 (11)	0.00098 (14)	−0.00316 (16)
O9	0.0352 (7)	0.0347 (7)	0.0187 (6)	−0.0031 (6)	0.0000 (7)	−0.0086 (10)
O3	0.0232 (7)	0.0314 (8)	0.0241 (8)	−0.0029 (6)	0.0011 (6)	−0.0063 (7)
O1	0.0233 (7)	0.0359 (9)	0.0200 (7)	−0.0051 (6)	0.0011 (6)	−0.0079 (6)
OW1	0.0317 (9)	0.0281 (8)	0.0253 (8)	0.0011 (7)	−0.0057 (7)	−0.0039 (6)
O4	0.0227 (7)	0.0314 (7)	0.0203 (7)	−0.0045 (6)	0.0026 (5)	−0.0064 (6)
OW2	0.0336 (8)	0.0330 (9)	0.0218 (8)	0.0075 (7)	0.0028 (7)	0.0014 (7)
O7	0.0371 (8)	0.0425 (8)	0.0225 (9)	−0.0094 (7)	0.0013 (9)	−0.0108 (8)
O2	0.0229 (7)	0.0299 (8)	0.0202 (6)	−0.0049 (6)	−0.0002 (6)	−0.0026 (6)
O10	0.0303 (8)	0.0364 (9)	0.0344 (9)	−0.0051 (7)	0.0134 (7)	−0.0095 (7)
O8	0.0265 (9)	0.0681 (13)	0.0345 (9)	−0.0075 (9)	0.0089 (7)	−0.0186 (9)
C3	0.0259 (10)	0.0191 (10)	0.0234 (10)	−0.0012 (8)	0.0037 (8)	−0.0008 (8)
C1	0.0266 (10)	0.0228 (9)	0.0172 (9)	−0.0029 (8)	−0.0010 (8)	−0.0013 (8)
C4	0.0267 (10)	0.0201 (9)	0.0184 (9)	0.0004 (8)	−0.0002 (8)	−0.0006 (8)
C2	0.0250 (10)	0.0308 (11)	0.0198 (10)	−0.0021 (9)	0.0015 (8)	−0.0043 (8)
C6	0.0454 (16)	0.075 (2)	0.067 (3)	−0.0191 (16)	−0.0159 (18)	0.013 (2)
C8	0.0456 (14)	0.0354 (13)	0.0761 (19)	0.0017 (11)	−0.001 (2)	0.0092 (18)
N1	0.0334 (12)	0.086 (2)	0.077 (2)	0.0050 (13)	−0.0007 (15)	0.025 (2)
N2	0.0337 (11)	0.0442 (14)	0.147 (3)	0.0059 (11)	−0.019 (2)	−0.022 (2)
C9	0.0334 (11)	0.0362 (12)	0.0542 (15)	0.0001 (9)	−0.0069 (14)	−0.0041 (15)
C7	0.0562 (18)	0.0569 (18)	0.105 (3)	0.0222 (15)	0.009 (3)	0.018 (3)

OW3	0.0304 (10)	0.0753 (15)	0.0290 (9)	0.0014 (10)	−0.0009 (8)	−0.0084 (10)
C5	0.0484 (15)	0.0403 (14)	0.090 (3)	−0.0076 (12)	−0.0156 (19)	−0.002 (2)

Geometric parameters (Å, °)

Cr1—O1	1.9636 (15)	C3—C4	1.558 (3)
Cr1—O4	1.9658 (15)	C1—C2	1.553 (3)
Cr1—O2	1.9715 (16)	C6—N1	1.330 (5)
Cr1—O3	1.9810 (16)	C6—C5	1.339 (5)
Cr1—OW1	1.9847 (17)	C6—H6	0.96 (4)
Cr1—OW2	2.0079 (18)	C8—C7	1.362 (4)
O9—C4	1.231 (3)	C8—C9	1.396 (4)
O3—C3	1.289 (3)	C8—H8	0.89 (4)
O1—C2	1.290 (3)	N1—C7	1.337 (5)
OW1—HW11	0.91 (4)	N1—H1	1.00 (5)
OW1—HW12	0.77 (5)	N2—C9	1.326 (3)
O4—C4	1.282 (2)	N2—H2	0.92 (4)
OW2—HW21	0.89 (4)	N2—H3	0.88 (4)
OW2—HW22	0.74 (4)	C9—C5	1.407 (4)
O7—C1	1.222 (3)	C7—H7	1.00 (5)
O2—C1	1.287 (3)	OW3—HW31	0.78 (4)
O10—C3	1.219 (3)	OW3—HW32	0.93 (4)
O8—C2	1.219 (3)	C5—H5	0.95 (4)
O1—Cr1—O4	93.67 (6)	O2—C1—C2	114.75 (17)
O1—Cr1—O2	83.18 (6)	O9—C4—O4	125.64 (19)
O4—Cr1—O2	176.78 (6)	O9—C4—C3	120.54 (18)
O1—Cr1—O3	175.74 (7)	O4—C4—C3	113.82 (16)
O4—Cr1—O3	82.13 (6)	O8—C2—O1	125.6 (2)
O2—Cr1—O3	101.02 (6)	O8—C2—C1	120.68 (19)
O1—Cr1—OW1	90.32 (7)	O1—C2—C1	113.75 (18)
O4—Cr1—OW1	90.73 (7)	N1—C6—C5	120.9 (3)
O2—Cr1—OW1	88.59 (7)	N1—C6—H6	117 (2)
O3—Cr1—OW1	90.43 (7)	C5—C6—H6	122 (2)
O1—Cr1—OW2	89.02 (7)	C7—C8—C9	119.8 (3)
O4—Cr1—OW2	90.45 (7)	C7—C8—H8	118 (2)
O2—Cr1—OW2	90.20 (8)	C9—C8—H8	123 (2)
O3—Cr1—OW2	90.31 (7)	C6—N1—C7	121.1 (3)
OW1—Cr1—OW2	178.68 (8)	C6—N1—H1	117 (3)
C3—O3—Cr1	114.84 (14)	C7—N1—H1	122 (3)
C2—O1—Cr1	114.20 (13)	C9—N2—H2	118 (2)
Cr1—OW1—HW11	112 (2)	C9—N2—H3	117 (3)
Cr1—OW1—HW12	107 (4)	H2—N2—H3	123 (4)
HW11—OW1—HW12	119 (4)	N2—C9—C8	121.2 (2)
C4—O4—Cr1	115.46 (13)	N2—C9—C5	122.0 (3)
Cr1—OW2—HW21	126 (2)	C8—C9—C5	116.8 (3)
Cr1—OW2—HW22	111 (3)	N1—C7—C8	120.7 (3)
HW21—OW2—HW22	100 (4)	N1—C7—H7	116 (3)
C1—O2—Cr1	113.56 (13)	C8—C7—H7	123 (3)
O10—C3—O3	125.7 (2)	HW31—OW3—HW32	98 (4)

O10—C3—C4	120.81 (18)	C6—C5—C9		120.6 (3)
O3—C3—C4	113.51 (17)	C6—C5—H5		122 (2)
O7—C1—O2	126.0 (2)	C9—C5—H5		117 (2)
O7—C1—C2	119.22 (19)			

Hydrogen-bond geometry (Å, °)

D—H⋯A	D—H	H⋯A	D⋯A	D—H⋯A
OW3—HW31⋯O8[i]	0.78 (4)	2.06 (4)	2.823 (3)	167 (4)
OW3—HW32⋯O8[ii]	0.93 (4)	1.93 (4)	2.843 (3)	169 (4)
OW2—HW21⋯O7[iii]	0.89 (4)	1.78 (4)	2.640 (3)	161 (3)
OW2—HW22⋯O9[iv]	0.74 (4)	2.07 (4)	2.784 (3)	162 (4)
OW1—HW11⋯OW3[v]	0.90 (4)	1.67 (4)	2.572 (3)	172 (4)
OW1—HW12⋯O10[vi]	0.76 (5)	2.01 (5)	2.745 (3)	162 (5)
N2—H2⋯O10[vii]	0.93 (4)	2.01 (4)	2.940 (4)	179 (4)
N1—H1⋯O3	1.00 (4)	2.27 (4)	3.061 (3)	135 (4)
N1—H1⋯O9[vi]	1.00 (4)	2.22 (4)	2.974 (3)	131 (4)

Symmetry codes: (i) x, $y+1$, $z+1$; (ii) $-x$, $-y+1$, $z+1/2$; (iii) $-x-1/2$, $y+1/2$, $z+1/2$; (iv) x, y, $z-1$; (v) x, $y-1$, z; (vi) $-x-1/2$, $y-1/2$, $z-1/2$; (vii) $-x-1$, $-y$, $z-1/2$.
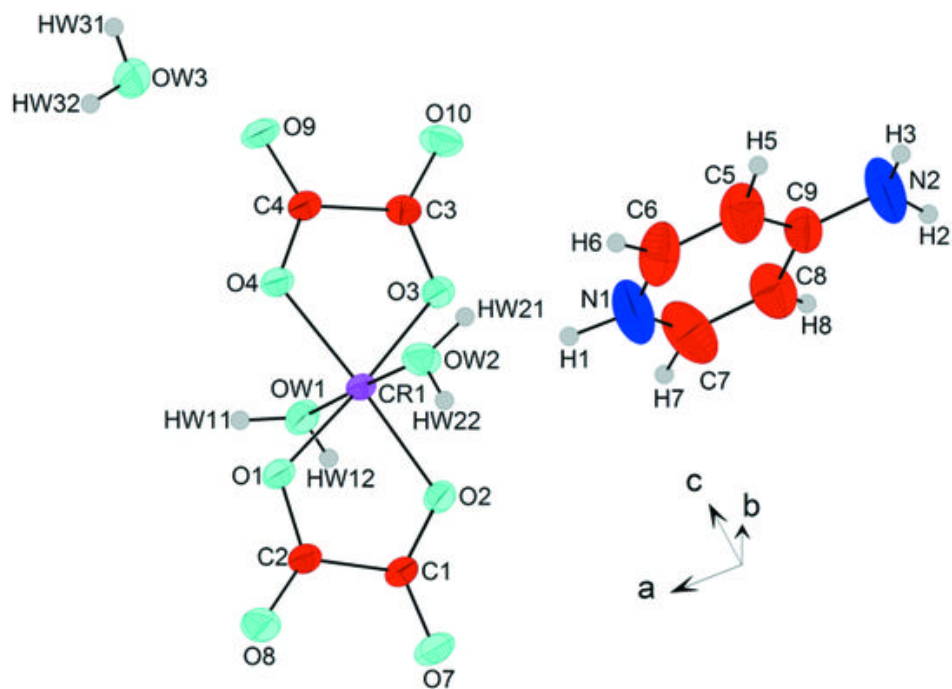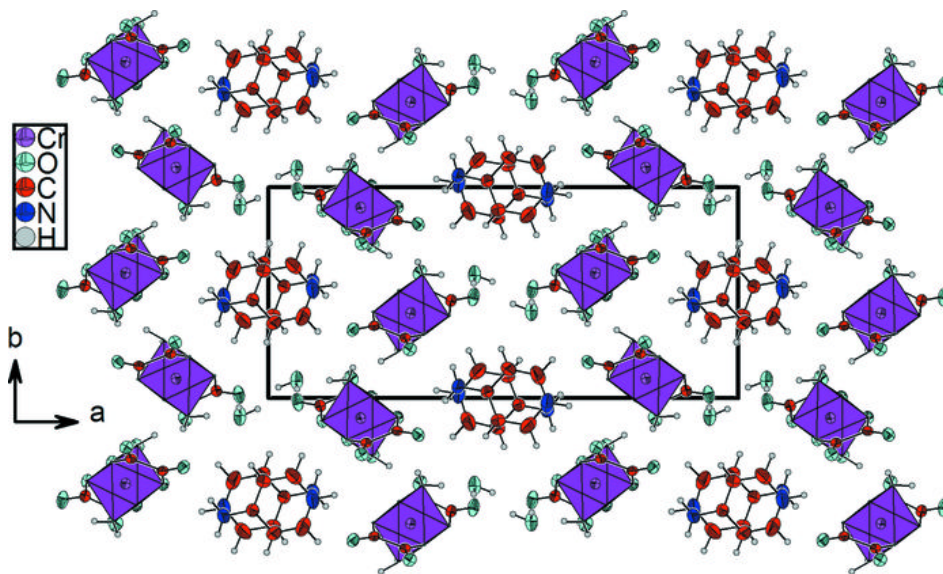
Fig. 1

Fig. 2

Fig. 3