

RESEARCH ARTICLE

A Practical Anodic and Cathodic Curve Intersection Model to Understand Multiple Corrosion Potentials of Fe-Based Glassy Alloys in OH⁻ Contained Solutions

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Abstract

A practical anodic and cathodic curve intersection model, which consisted of an apparent anodic curve and an imaginary cathodic line, was proposed to explain multiple corrosion potentials occurred in potentiodynamic polarization curves of Fe-based glassy alloys in alkaline solution. The apparent anodic curve was selected from the measured anodic curves. The imaginary cathodic line was obtained by linearly fitting the differences of anodic curves and can be moved evenly or rotated to predict the number and value of corrosion potentials.

Introduction

Fe-based glassy alloys have been widely used in industry because of their high glass-forming ability, soft magnetic properties, good corrosion and wear resistance and ultrahigh strength properties [1–5]. Recently, the Fe₇₈Si₉B₁₃ metallic glass is the leading product in Chinese Fe-based amorphous ribbons market [6]. In 1988 year, Yoshizawa and co-workers [1] invent the alloy labeled as Finemet with the composition of Fe_{73.5}Si_{13.5}B₉Cu₁Nb₃, which has a super soft magnetic property and is used widely in electronic devices. For Fe-based glassy alloys, there exhibits a clear passive zone in the polarization curve in OH⁻ contained solutions, while the passive zone is very short in the solution without OH⁻ [5]. Meanwhile, the corrosion resistance becomes an important factor to consider in using Fe-based glasses [7,8]. Hence, it is valuable to study the electrochemical behavior of Fe₇₈Si₉B₁₃ and Fe_{73.5}Si_{13.5}B₉Cu₁Nb₃ glasses in different environments.

Measuring potentiodynamic polarization curve is one of the traditional methods to evaluate electrochemical properties [9,10]. And many works on multiple corrosion potentials in polarization curve for stainless steels in acidic solution have been reported in the literature [11–15]. Regarding the influencing factors on the corrosion potentials, Escrivà-Cerdán et al. [16] have

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found that high-alloyed austenitic stainless steel UNS N08031 (Alloy 31) has three corrosion potentials in phosphoric acid at 60 and 80°C since the temperature can favor the cathodic reaction. Qiao et al. [17] have reported that a certain sulfuric acid and oxygen concentration can induce three corrosion potentials during the polarization of a nitrogen bearing stainless steel. It has been found that the corrosion behavior of pure titanium shows three corrosion potentials when corroded in the H₂SO₄ solution containing 0.001 M and 0.002 M fluoride ions [18]. In addition, the speed jet of phosphoric acid solution to the stainless steel can also induce three corrosion potentials [19]. However, to the authors' knowledge, three corrosion potentials occurring in the alkaline solution for Fe-based glassy alloys have not been reported so far.

The occurrence of multiple corrosion potentials was reported to be attributed to the instability of passive film. Kelly et al. [20] have reported that the formation mechanism of three corrosion potentials is that the anodic and cathodic Evans lines intersect at three points and there exist one anodic loop and one cathodic loop between the three corrosion potentials. It was also pointed that the origin of the cathodic loop is related to the greater rate of the cathodic reaction than that of the passive current density and the former can conceal the latter at these potentials near the active-passive transition. Qiao et al. [17] have constructed ideal polarization curve models based on the aforementioned mechanism to understand the appearance of three corrosion potentials under certain sulfuric acid concentration and oxygen content. However, the authors have not given the direct relation between the ideal polarization curve and experimental data, which is valuable to discover.

In addition, many works have pointed out that the scanning rate and solution concentration have a significant impact on electrochemical behavior. Zhang et al. [21] have found that the scanning rate can directly affect the Tafel slope. It have been confirmed by Manning et al. that the pitting potential and scanning rate have a functional relationship [22]. In addition, Nakagawa et al. [23] have pointed that the fluoride concentration and pH can affect the polarization test of titanium in NaF solution with various concentrations and pH values.

In this study, our major aim is to build a practical accurate model to explain the occurrence of multiple corrosion potentials of Fe-based glassy ribbons when scanning rate (ν) or NaOH concentration (c_{NaOH}) reaches a specific value. In our model, we process a new method to construct the imaginary cathodic line from the measured anodic polarization curves.

Materials and Methods

The commercial Fe₇₈Si₉B₁₃ glassy ribbons and master alloys were supplied by Qingdao Yunlu Energy Technology Company Ltd. The original size of the commercial ribbons was 40 mm in width, 35 μm in thickness and about 20000 mm in length, from which the ribbon samples with size of 3×40 mm were cut for electrochemical tests. Fe₇₈Si₉B₁₃ master alloy was prepared by melting iron, pure silicon and Fe-B alloy in medium induction furnace, and Fe_{73.5}Si_{13.5}B₉Cu₁Nb₃ master alloy was obtained by melting of pure iron, silicon, copper, niobium and ferroboron alloy. The Fe_{73.5}Si_{13.5}B₉Cu₁Nb₃ glassy ribbons were obtained using a single copper roller in the lab. X-ray diffraction analysis (as shown in Refs [24] and [25]) proves that the as-spun ribbon samples are in fully amorphous state. Since the most of metallic materials in the industrial application are in crystalline state. In this paper, we focus on the metallic glasses, which are sparser than the metallic crystalline materials; by the way, we are also interested in the electrochemical behavior of their crystalline counterpart, i.e. master alloys. Hence, we choose the master alloys as the comparison.

All electrochemical measurements were carried out using a standard three-electrode system: working electrode, platinum counter electrode and Hg/HgO reference electrode. All the measurements were only conducted on the wheel sides of the ribbons and their free sides were

covered by silicone rubber. CHI 660E electrochemical workstation was used for measuring the polarization curves. All experiments were performed at room temperature (298 K). In the measurement of potentiodynamic polarization curves with a series of scanning rates (ν), the electrolyte was 0.6 M NaCl + 0.12 M NaOH. In the measurement of polarization curves with different NaOH concentrations (c_{NaOH}), ν was fixed at 1 mV/s and the electrolyte was 0.6 M NaCl + x M NaOH ($0.04 \leq x \leq 0.7$). All measurements were repeated at least three times to ensure good reproducibility.

To understand the anodic process of $\text{Fe}_{78}\text{Si}_9\text{B}_{13}$ glassy ribbon in 0.6 M NaCl + 0.12 M NaOH, the ribbons were polarized to three different potentials which were located after the first anodic peak, 2nd anodic peak and three corrosion potentials, respectively. The sample surface morphology was examined using scanning electron microscopy (SEM, SU-70). The surface was also analyzed by X-ray photoelectron spectroscopy (XPS, ESCALAB 250Xi) performed on a photoelectron spectrometer with Al-K α excitation.

Results

Polarization with different scanning rates

Fig 1A shows the electrochemical polarization curves of $\text{Fe}_{78}\text{Si}_9\text{B}_{13}$ glassy ribbons in 0.6 M NaCl + 0.12 M NaOH solution with the scanning rates $\nu = 0.5\text{--}5$ mV/s. As ν is reduced from 5 to 1.5 mV/s, the polarization curves have the similar shapes with two current density peaks P1 and P2 in A-B-C and C-D-E zones, respectively. As ν was further reduced to 1 mV/s, three corrosion potentials appear with the sacrifice of P1, and the polarization curve is divided into I, II, I' and II' zones. The current density in both I and I' zones is negative, and that in II and II' is

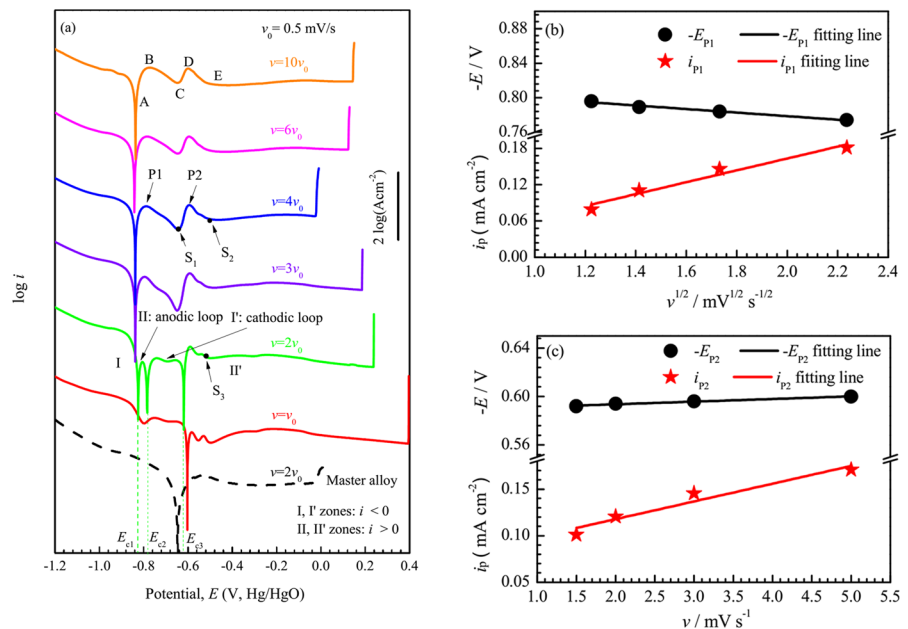


Fig 1. Electrochemical tests of samples. (a) Potentiodynamic polarization curves of $\text{Fe}_{78}\text{Si}_9\text{B}_{13}$ glassy ribbons and the corresponding master alloy with different scanning rates ν in 0.6 M NaCl + 0.12 M NaOH solution. For clarity, the curves of $\text{Fe}_{78}\text{Si}_9\text{B}_{13}$ glassy ribbons with $\nu = \nu_0, 2\nu_0, 3\nu_0, 4\nu_0, 6\nu_0$ and $10\nu_0$ are shifted upward by multiplying the raw data with $10^2, 10^4, 10^6, 10^8, 10^{10}$ and 10^{12} respectively. (b) Variation of the current densities i_{p1} and peak potentials E_{p1} with the square root of ν . (c) Variation of the current densities i_{p2} and peak potentials E_{p2} with ν . P1 and P2 indicate the two current density peaks such as A-B-C and C-D-E zones in the polarization curve with $\nu = 10\nu_0$ ($\nu_0 = 0.5$ mV/s), respectively.

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positive. Hence, II is called anodic loop and I' is called cathodic loop. The curve with $\nu = 0.5$ mV/s has only one corrosion potential, and there exist one and two small troughs before and after it respectively. It should be noted that, the measured polarization curve of Fe₇₈Si₉B₁₃ master alloy with $\nu = 1$ mV/s has one corrosion potential, and no distinct current density peak P1 or P2 (Fig 1A). This indicates that the occurrence of three corrosion potentials is not only related to ν , but also to microstructure.

Moreover, in Fig 1A, as the scanning rate ν is decreased from 5 to 2 mV/s, the pitting potential E_{pit} decreases, whereas E_{p1} increases with further decreasing ν from 1.5 to 0.05 mV/s. In the first changing range of ν , the positive relationship between E_{pit} and ν can be explained by the point defect model (PDM) [26,27]. However, in the second range, it is found that the passive current density i_{pass} is apparently lower than that in the first range, and according to the argument that cumulative anodic charge density, Q_c , is constant, which is presented as [28]:

$$Q_c \approx i_{pass} \times \Delta t_p = i_{pass} \frac{(E_{pt} - E_{ocp}) \times 10^{-3}}{\nu} \tag{1}$$

where E_{ocp} is the open circuit potential, Δt_p is the anodic polarization time duration until the stable pitting. Now we get:

$$E_{pit} = \frac{Q_c \cdot \nu}{i_{pass}} + E_{ocp} \tag{2}$$

When i_{pass} has a larger decreasing rate than the decreasing rate of ν , E_{pit} might increase. Hence, it is understood that E_{pit} increases when ν varies from 1.5–0.05 mV/s. This phenomenon is similar to the fact that Q_c and E_{pit} sometimes increase abnormally with decreasing ν in the experimental data of Zhang and co-workers' work [26].

As seen from Fig 1B, the values of current density i_p and the corresponding potential E_p of P1 vary linearly with the square root of the scanning rate ν . According to Müller and Calandra's model [29,30], P1 can be explained by ohmic resistance caused by the insoluble reaction products on the sample surface and the i_{p1} and E_{p1} can be expressed as

$$i_{p1} = \left(\frac{nF \rho k}{M} \right)^{1/2} (1 - \theta_p) \nu^{1/2} \tag{3}$$

$$E_{p1} = \left(\frac{nF \rho k}{M} \right)^{1/2} \left[\left(\frac{\delta}{k} \right) + R_0 A_0 (1 - \theta_p) \right] \nu^{1/2} \tag{4}$$

where k is the specific conductivity of the solution within the pores of the film, ρ the film density, δ the film thickness, M the molar mass of the film formed by the flow of Coulombs, R_0 the external resistance, θ_p the surface fractional coverage and A_0 the electrode area.

Fig 1C shows that the i_p of P2 decreases linearly but the corresponding E_p decreases slightly with the decrease of the scanning rate. According to Srinivasan and Gileadi's model [31], P2 is related to an adsorption process and the i_{p2} and E_{p2} can be defined by

$$i_{p2} = \left(\frac{\gamma F}{4RT} \right) \nu \tag{5}$$

$$E_{p2} = - \left(\frac{RT}{F} \right) \ln K \tag{6}$$

where K is the velocity constant ratio of adsorption to desorption, γ is the charge required to

form a monolayer of adsorbed species and other symbols have their usual meaning in electrochemistry.

In addition, the polarization curves of $\text{Fe}_{73.5}\text{Si}_{13.5}\text{B}_9\text{Cu}_1\text{Nb}_3$ glassy ribbons in 0.6 M NaCl + 0.12 M NaOH solution with $\nu = 6\text{--}10$ mV/s are shown in Fig 2. The polarization curves with $\nu = 8\text{--}10$ mV/s have a similar shape with one current density peak. With decreasing ν , the current density of the peak decreases, but its potential approximately remains constant. As ν decreases to 7 mV/s, three corrosion potentials present like $\text{Fe}_{78}\text{Si}_9\text{B}_{13}$ glassy ribbon with 1 mV/s and they divide the polarization curve into I, II, I' and II' zones with the current density symbol of negative, positive, negative and positive, respectively. The polarization curve with $\nu = 6$ mV/s has only one corrosion potential with a small trough before it. Similarly, the polarization curve of $\text{Fe}_{73.5}\text{Si}_{13.5}\text{B}_9\text{Cu}_1\text{Nb}_3$ master alloy with 7 mV/s does not have three corrosion potentials or a distinct current density peak in anodic part. In addition, the E_{pit} of samples creases with decreasing ν , which is similar to the phenomenon in the second changing range of ν in $\text{Fe}_{78}\text{Si}_9\text{B}_{13}$ (Fig 1A).

Polarization with different NaOH concentrations

Fig 3 displays the polarization curves of $\text{Fe}_{78}\text{Si}_9\text{B}_{13}$ glassy alloys in 0.6 M NaCl + x M NaOH solutions with the scanning rate $\nu = 1$ mV/s. The characteristic of the polarization curve in Fig 3 as well as those in Figs 1A and 2 are listed in Table 1. The curves with $c_{\text{NaOH}} = 0.2\text{--}0.7$ M have two current density peaks as the same as those obtained in 0.6 M NaCl + 0.12 M NaOH solution with $\nu = 5$ to 1.5 mV/s. With decreasing c_{NaOH} , the corrosion potential and both two peaks' potentials shift to right, and the passive range shortens. When c_{NaOH} decreases to 0.2 M, hydrolysis reaction near 0.6 V transforms into pitting corrosion reaction. Three corrosion potentials E_{c1} , E_{c2} and E_{c3} occur and divide the polarization curve into I, II, I' and II' zones looped with negative and positive symbols when c_{NaOH} decreases to 0.12 and 0.08 M. Apparently, the values of E_{c1} , E_{c2} or E_{c3} with $c_{\text{NaOH}} = 0.08$ M respectively are higher than the corresponding one with $c_{\text{NaOH}} = 0.12$ M. When c_{NaOH} decreases to 0.04 M, the number of potentials decreases to one and the pitting potential E_{pit} becomes further negative. The OH^- dependent E_{pit} obeys the following relation between the pitting potential and the PH of the electrolyte [32]:

$$E_{\text{pit}} = a + b\text{PH} \quad (7)$$

where the value of a is 690 mV and b is 21 mV for $\text{Sn}_{66}\text{Ni}_{34}$ alloy in NaCl solution with various pH values. We use Eq (7) to give only a qualitative explanation on the c_{NaOH} dependent E_{pit} is present paper.

For clarity, the characteristic parameters of the measured polarization curves of $\text{Fe}_{78}\text{Si}_9\text{B}_{13}$ and $\text{Fe}_{73.5}\text{Si}_{13.5}\text{B}_9\text{Cu}_1\text{Nb}_3$ glassy samples with various scanning rates and c_{NaOH} as well as their standard deviation are listed in Table 1.

SEM and XPS measurement after various polarizations

In order to investigate the nature of the anodic current density peaks of $\text{Fe}_{78}\text{Si}_9\text{B}_{13}$ glassy alloys in 0.6 M NaCl + 0.12 M NaOH, the scanning rate $\nu = 1$ mV/s and 2 mV/s were chosen to polarize the glassy ribbons until three different points S_1 , S_2 and S_3 , which are indicated in Fig 1A. Samples S_1 and S_2 were polarized with $\nu = 2$ mV/s until the end of the first and second anodic peaks, respectively; while sample S_3 was polarized to a similar potential to S_2 with $\nu = 1$ mV/s after experiencing three corrosion potentials. The surface morphologies of these samples are shown in Fig 4. A lot of white flake-like products in size of about 0.1–1 μm present on all the

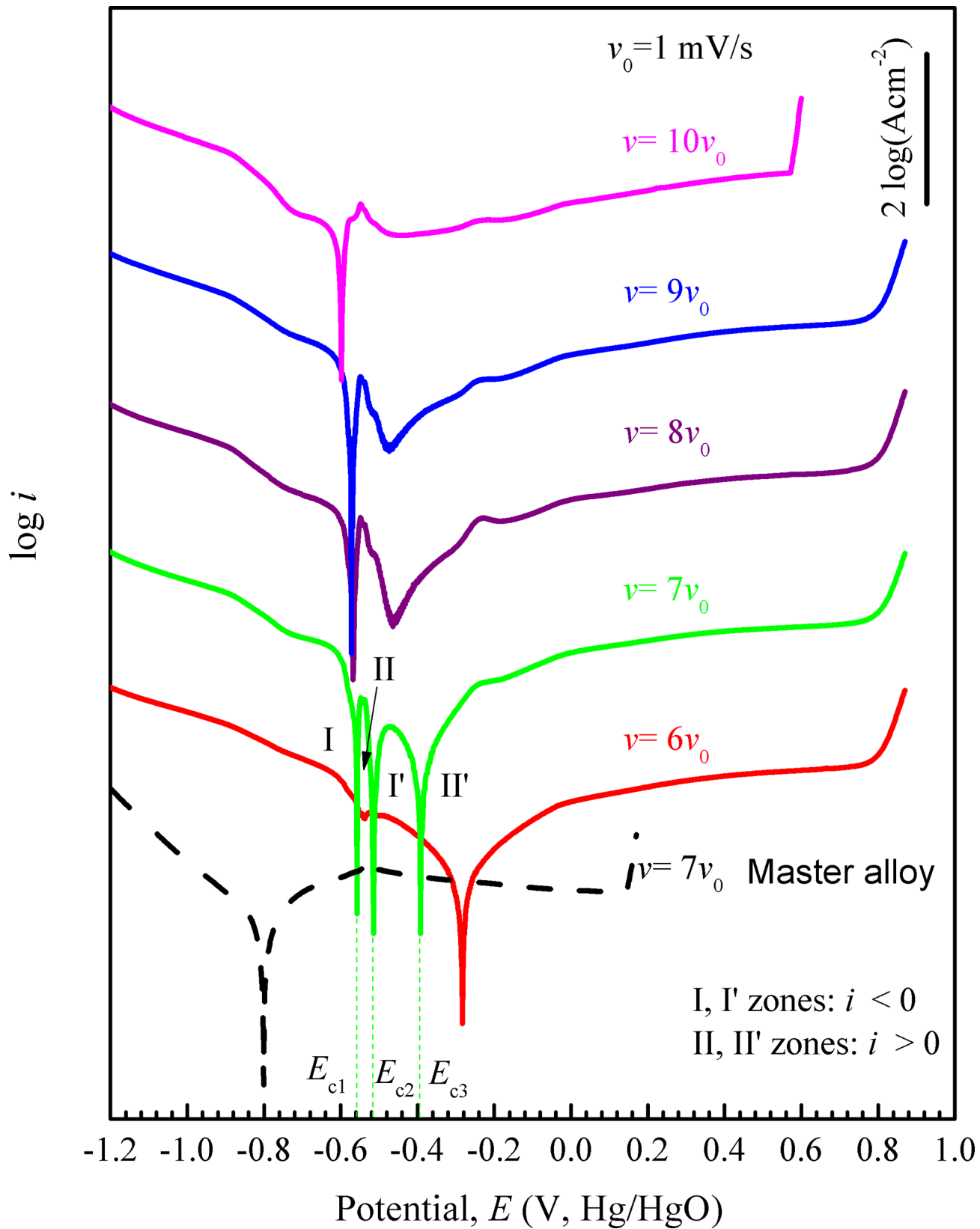


Fig 2. Potentiodynamic polarization curves of Fe_{73.5}Si_{13.5}B₉Cu₁Nb₃ glassy ribbons and the corresponding master alloy with different scanning rates ν in 0.6 M NaCl + 0.12 M NaOH solution. For clarity, the curves of Fe_{73.5}Si_{13.5}B₉Cu₁Nb₃ glassy ribbons with $\nu = 6 \nu_0, 7 \nu_0, 8 \nu_0, 9 \nu_0$ and $10 \nu_0$ are shifted upward by multiplying the raw data with $10^2, 10^4, 10^6, 10^8$ and 10^{10} respectively.

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samples and their size, i.e. the surface roughness of samples, increases in order of $S_1 < S_2 < S_3$. Meanwhile, the surface morphology of S_3 is similar to that of S_2 rather than S_1 .

Fig 5 shows the X-ray photoelectron spectroscopy (XPS) spectra of Fe 2p and Si 2p recorded from the surfaces of S_1, S_2 and S_3 . For the spectra of Fe 2p, the total peak intensity of S_1 is greater than that of S_2 or S_3 , possibly due to the higher surface roughness of S_2 and S_3 . The Fe 2p spectrum can be decomposed into two peaks belonging to FeO and Fe₂O₃ [33]. According to earlier work [34], the XPS spectrum of Si 2p of S_1 can be decomposed into three peaks: SiO₂, SiO_x/Si and Si. But for S_2 and S_3 , it can only be decomposed into two peaks corresponding to SiO₂ and SiO_x/Si. The decomposed peak fractions for Fe 2p and Si 2p spectra are listed in Table 2. Apparently the FeO and SiO₂ fractions of samples change in the order $S_1 > S_2 > S_3$, but the Fe₂O₃ and SiO_x/Si fractions change oppositely. Overall, the decomposed XPS result of S_3 is similar to that of S_2 , but different from S_1 .

Discussion

Practical model to explain the occurrence of multiple corrosion potentials of Fe-based glassy ribbons

From Fig 1A, the polarization curves of Fe₇₈Si₉B₁₃ glassy ribbons with the scanning rate $\nu \geq 1.5$ mV/s have two current density peaks P1 and P2 in the anodic part. As $\nu = 1$ mV/s, there are three corrosion potentials in the polarization curve. Kelly et al. [20] have pointed that three corrosion potentials E_{c1}, E_{c2} , and E_{c3} are the intersecting points of anodic and cathodic Evans lines based on the following equation

$$i = i_a - i_c = 0, \text{ when } E = E_{c1}, E_{c2} \text{ and } E_{c3} \quad (8)$$

where i_a is anodic current density, i_c is cathodic current density. Similarly, Qiao et al. [17] have supposed a schematic mechanism of corrosion with three E_c by intersecting an ideal anodic polarization curve with an ideal cathodic reaction line, there are different intersection numbers by changing the former's or latter's position.

It's known that the cathodic reaction in alkaline solution is the oxygen reduction reaction [35]



and /or the water reduction reaction



The cathodic part does not show the limiting value representing the oxygen reduction reaction, but shows a relatively higher current density (Fig 1A). In addition, a lot of gas bubbles can be observed on the sample corroding surface in the experiment. Both phenomena show that the cathodic reaction in alkaline solution is dominated by water reduction (reaction (10)) [35]. Hence, decreasing c_{NaOH} can enhance the cathodic reaction and increase the corrosion

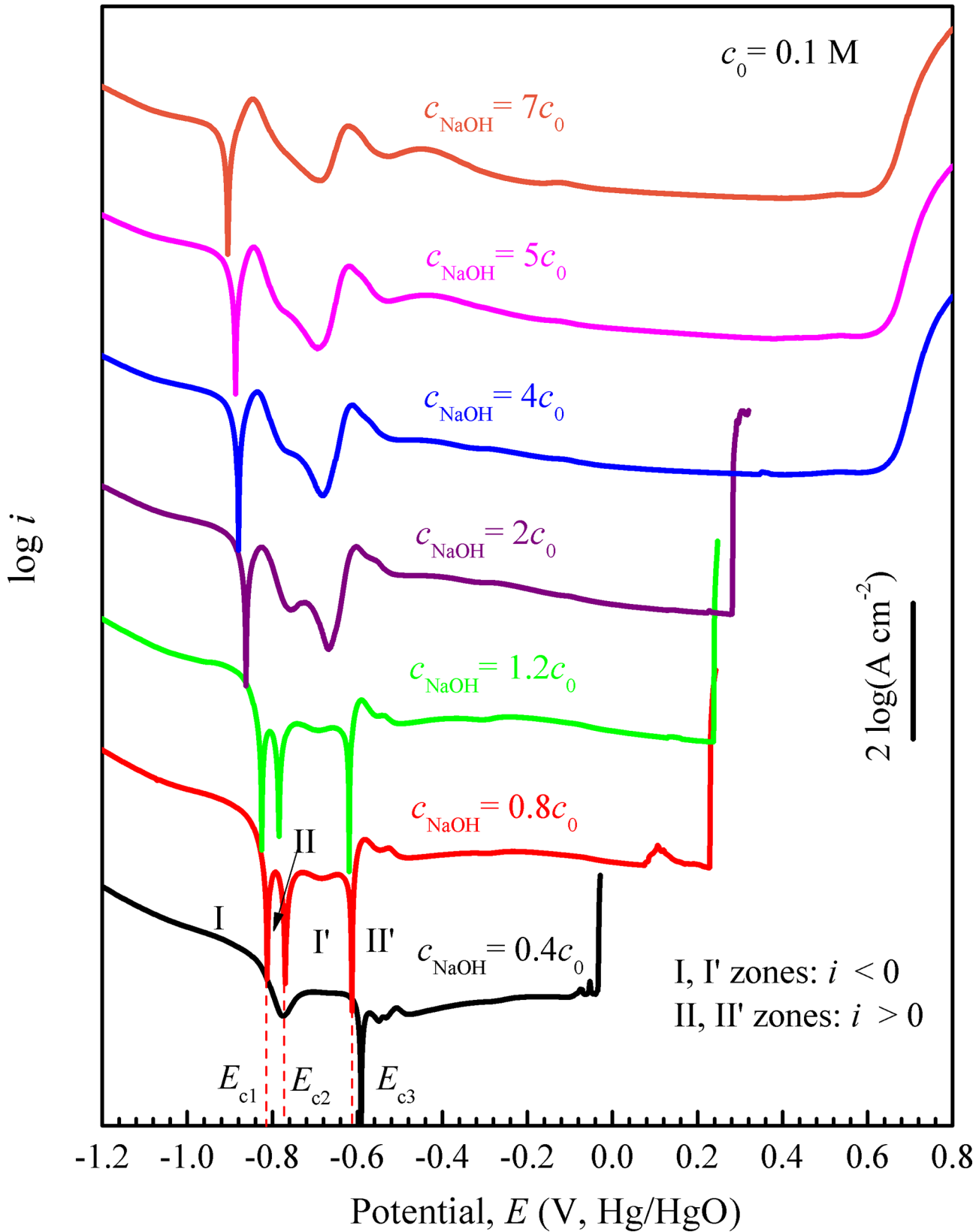


Fig 3. Potentiodynamic polarization curves of $\text{Fe}_{78}\text{Si}_9\text{B}_{13}$ glassy ribbons in 0.6 M NaCl + x M NaOH ($x = 0.04\text{--}0.7$) solution with $v = 1$ mV/s. For clarity, the curves with $c = 0.8 c_0, 1.2 c_0, 2 c_0, 4 c_0, 5 c_0$ and $7 c_0$ are shifted upward by multiplying the raw data with $10^2, 10^4, 10^6, 10^8, 10^{10}$ and 10^{12} respectively.

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Table 1. Variation of polarization parameters of Fe₇₈Si₉B₁₃ and Fe_{73.5}Si_{13.5}B₉Cu₁Nb₃ glassy ribbons.

Alloy	v or c_{NaOH}	E_{p1} (Vvs.Hg/HgO)	$\log j_{p1}$ (A cm ⁻²)	E_{p2} (Vvs.Hg/HgO)	$\log j_{p2}$ (A cm ⁻²)	$\log j_{\text{pass}}$ (A cm ⁻²)	E_{pit} (Vvs.Hg/HgO)
Fe ₇₈ Si ₉ B ₁₃	10 v_0	-0.77±0.01	-3.74±0.11	-0.60±0.01	-3.77±0.10	-4.13±0.10	0.14±0.03
	6 v_0	-0.78±0.01	-3.84±0.15	-0.59±0.02	-3.84±0.30	-4.17±0.05	0.13±0.04
	4 v_0	-0.78±0.02	-3.96±0.12	-0.59±0.01	-3.92±0.14	-4.21±0.12	-0.02±0.03
	3 v_0	-0.79±0.01	-4.20±0.06	-0.59±0.01	-4.00±0.01	-4.31±0.04	0.19±0.05
	2 v_0	-	-	-0.59±0.01	-4.02±0.2	-4.52±0.03	0.24±0.03
	v_0	-	-	-0.58±0.00	-4.83±0.3	-5.72±0.02	0.40±0.01
Fe _{73.5} Si _{13.5} B ₉ Cu ₁ Nb ₃	10 v'_0	-0.55±0.01	-4.67±0.01	-	-	-4.29±0.01	0.57±0.01
	9 v'_0	-0.55±0.01	-4.99±0.01	-	-	-4.38±0.01	0.79±0.01
	8 v'_0	-0.55±0.02	-4.80±0.01	-	-	-4.38±0.01	0.78±0.02
	7 v'_0	-	-	-	-	-4.35±0.05	0.79±0.01
	6 v'_0	-	-	-	-	-4.36±0.01	0.80±0.01
Fe ₇₈ Si ₉ B ₁₃	7 c_0	-0.84±0.01	-3.38±0.01	-0.62±0.02	-3.78±0.06	-4.68±0.01	0.60±0.01
	5 c_0	-0.84±0.01	-3.58±0.01	-0.62±0.02	-3.88±0.01	-4.75±0.01	0.62±0.01
	4 c_0	-0.83±0.01	-3.71±0.01	-0.61±0.02	-3.89±0.05	-4.79±0.01	0.64±0.01
	2 c_0	-0.83±0.01	-4.12±0.06	-0.60±0.01	-4.05±0.05	-4.82±0.01	0.28±0.01
	1.2 c_0	-	-	-0.58±0.01	-4.60±0.07	-4.69±0.01	0.24±0.03
	0.8 c_0	-	-	-0.59±0.01	-4.22±0.04	-4.59±0.01	0.23±0.01
	0.4 c_0	-	-	-0.57±0.01	-4.47±0.01	-4.55±0.01	-0.03±0.01

$v_0 = 0.5$ mV/s

$v'_0 = 1$ mV/s

$c_0 = 0.1$ M NaOH.

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potential (Fig 3). According to Lee [36], there are three possible determining steps in reaction (10).



The reaction rate U can be written as the following formula based on reaction (11a) by neglecting the back reaction

$$U = K_f \exp(-n\alpha F\psi/RT) \quad (12)$$

where n is the total number of electrons involved in the reaction, α the transfer coefficient, ψ the potential between the sample and the solution and K_f the rate constant for the forward reaction. Then the current density i can be written as

$$i = nFa_{\text{H}_2\text{O}}U \quad (13)$$

where $a_{\text{H}_2\text{O}}$ denotes the activity of H_2O . Combining Eqs (12) with (13), we will get

$$\log i = \log nFa_{\text{H}_2\text{O}}K_f + \frac{1}{2.303}RT - \frac{n\alpha F}{2.303}\psi \quad (14)$$

So the cathodic reaction can be represented by a line in the half logarithmic coordinates. Such linear relation in $\log i - \psi$ can be obtained in a straightforward way according to Ref. [36].

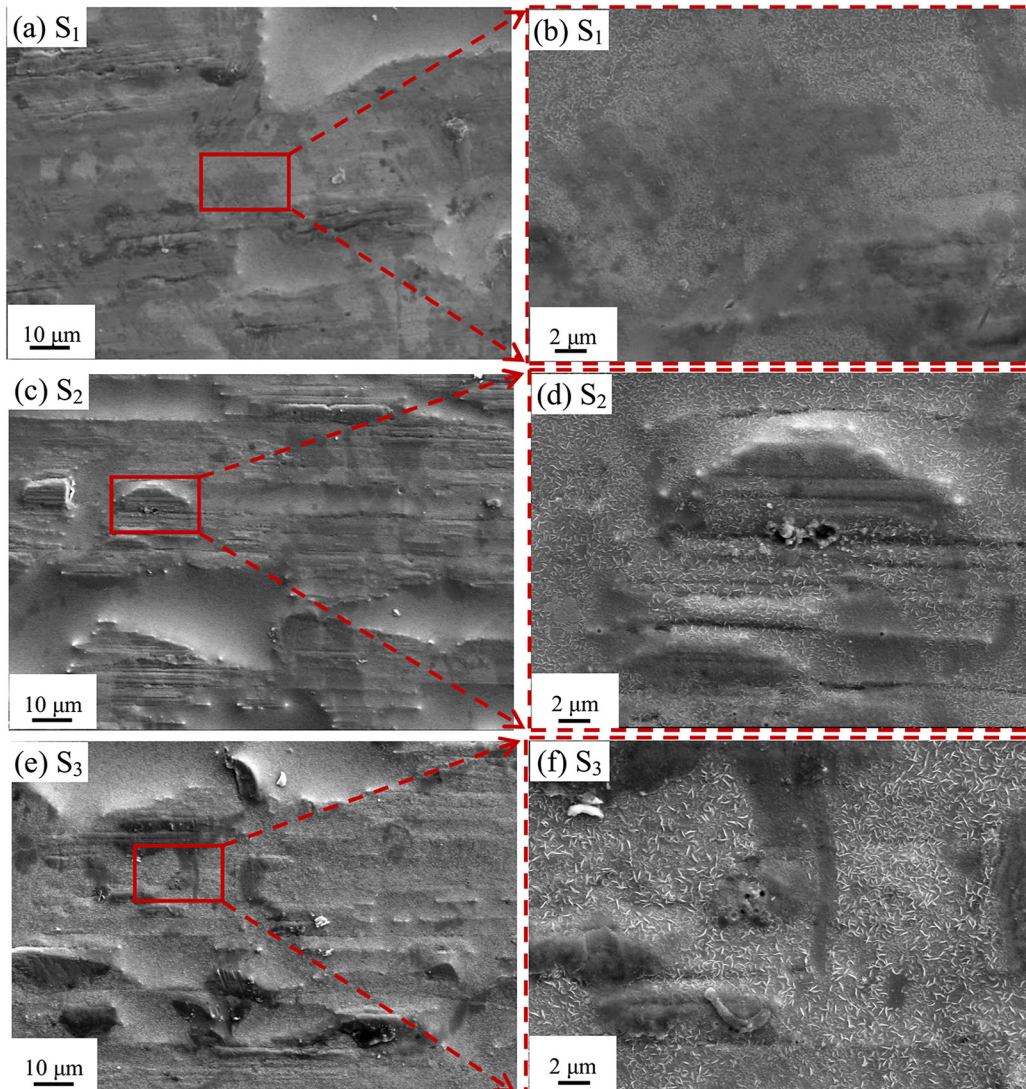


Fig 4. SEM micrographs for the surfaces of (a) and (b) S_1 , (c) and (d) S_2 , and (e) and (f) S_3 samples. S_1 and S_2 were polarized with $v = 2$ mV/s until the end of the first and second anodic peaks, respectively, while S_3 was polarized to a similar potential to S_2 with $v = 1$ mV/s after experiencing three corrosion potentials. S_1 , S_2 and S_3 are indicated in Fig 1A.

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Hence, it is understood that Kelly and Qiao used the cathodic line to explain the variation of polarization curves [17,20].

However, there are several unknown parameters like K_f , α and a_{H_2O} in present polarization process. In the following sections, we propose a simple method to deduce an imaginary cathodic line, evenly move it or rotate it, and try to discuss the evolution of the polarization curves, especially the formation of three corrosion potentials (Figs 1A, 2 and 3).

Evenly moving the imaginary cathodic line. As shown in Fig 1A, the active anodic peaks P1 and P2 with the scanning rate $v = 5$ mV/s can be divided into four potential ranges, i.e., AB, BC, CD and DE. According to Lu et al. [37], the measured current density $i(E)$ of anodic

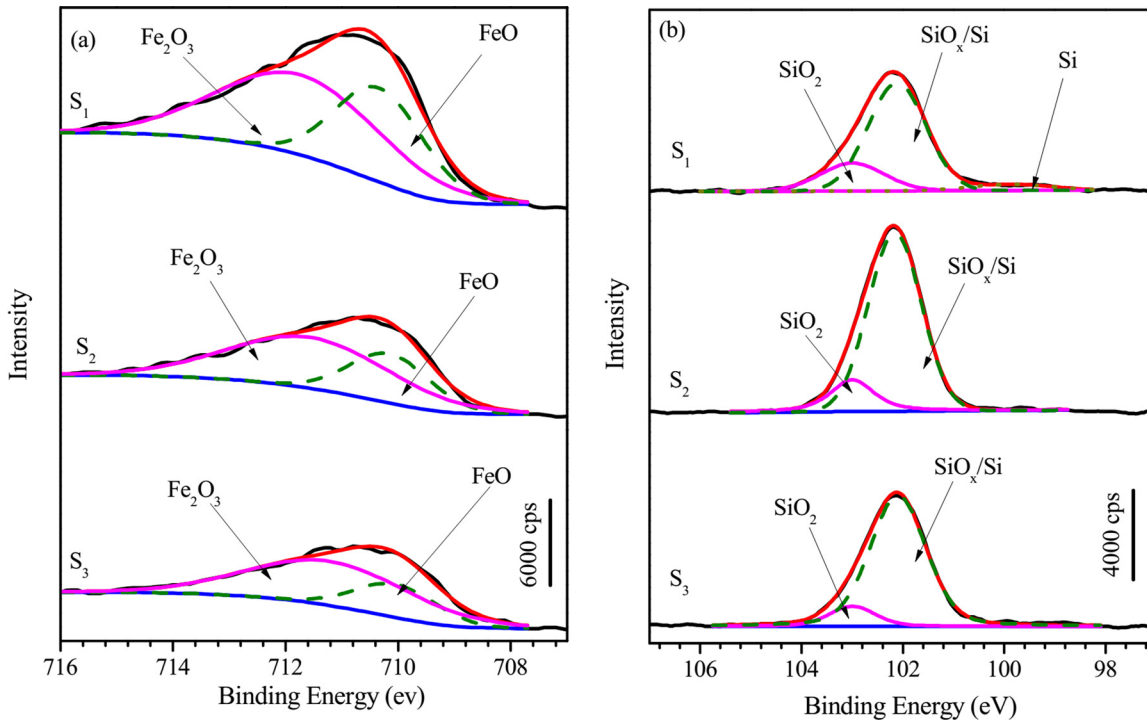


Fig 5. XPS spectra of (a) Fe 2p and (b) Si 2p recorded from the surfaces of S₁, S₂ and S₃. S₁ and S₂ were polarized with $v = 2$ mV/s until the end of the first and second anodic peaks, respectively, while S₃ was polarized to a similar potential to S₂ with $v = 1$ mV/s after experiencing three corrosion potentials. S₁, S₂ and S₃ are indicated in Fig 1A.

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polarization curve except the steady passive ranges can be expressed by

$$i(E) = \sum_m i_{0-m}(E) = \sum_m i_{0-m} \exp\left(\frac{E - E_0}{B}\right) \quad (15)$$

where i_{0-m} , E_0 and B are exchange current density, equilibrium potential and symmetry like coefficient, respectively. Generally, B is also denoted by

$$B = \frac{RT}{\beta F} \quad (15a)$$

here β is symmetry coefficient, R is gas constant, T is temperature, F is Faraday's constant [37]. In this paper, R , T and F are constants, which can be move before the exponent function.

From Figs 1A and 3 and Table 1, the second current density peak varies much smaller than the first peak with the variation of v and c_{NaOH} . So P2 could be omitted in following analysis for simplification like omitting the higher terms in series expansion. Hence, the anodic

Table 2. The fraction of decomposed peaks from XPS spectra for Fe 2p and Si 2p of the surfaces of S₁, S₂ and S₃.

Sample	Fe 2p		Si 2P		
	Fe ₂ O ₃ (%)	FeO (%)	SiO ₂ (%)	SiO _x /Si (%)	Si (%)
S ₁	61.5	38.5	21.5	72.5	6
S ₂	68.2	31.8	16.2	83.8	-
S ₃	73.3	26.7	11.0	89.0	-

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polarization curves with $\nu = 5$ and 3 mV/s could be described as following:

$$i = i_{0-1}e^{\beta_1\psi} - i_{0-2}e^{\beta_2\psi} \tag{16}$$

$$i' = i'_{0-1}e^{\beta'_1\psi} - i'_{0-2}e^{\beta'_2\psi} \tag{17}$$

Since the curve shape with $\nu = 5$ mV/s is similar to that with $\nu = 3$ mV/s, a further simplification and assumption could be made as

$$\Delta i_0 = i_{0-1} - i'_{0-1} \approx i_{0-2} - i'_{0-2} < 0 \tag{18}$$

and

$$\beta_1 = \beta'_1, \beta_2 = \beta'_2, \beta_2 = \beta_1 + \delta/\psi \tag{19}$$

hence, the current density difference Δi can be expressed as

$$\Delta i = i - i' = \Delta i_0 e^{\beta_1\psi} (1 - e^\delta) \tag{20}$$

Herein, we set

$$\delta = e^{-a^0\psi} \tag{21}$$

where a^0 is constant and $a^0 > \beta_1$. Now, by expanding the right item in Eq (21) with Talor`s series and omitting the higher order terms, the difference Δi can be expressed as

$$\Delta i = \Delta i_0 e^{\beta_1\psi} (1 - e^{-a^0\psi}) = -\Delta i_0 e^{\beta_1\psi} e^{-a^0\psi} \tag{22}$$

$$\log \Delta i = \log(-\Delta i_0) + 0.434(\beta_1 - a^0)\psi \tag{23}$$

Eq (23) shows a liner relation between $\log \Delta i$ and ψ . Its intercept and slope can be obtained by linearly fitting the difference $i-i'$ of the measured curves which is shown in Fig 6A. When ν changes, the exchange current density i'_{0-1} should change according to Eq (17), and then $\log(-\Delta i_0)$, i.e. the intercept of the Eq (23), changes; while the slope of Eq (23) does not change. In other words, the line of $\log \Delta i-\psi$ moves evenly with changing ν .

Apparently, Eq (23) is similar to Eq (14). If Eq (20) is approximately regarded as Butler-Volmer formula, the current density difference i and current density i with $\nu = 5$ mV/s can be regarded as imaginary cathodic and anodic current densities, respectively, i.e. i' can be deduced if Δi and i are known. In other words, we can predict the characteristic of a new current density i' by subtracting an imaginary cathodic line i from the apparent cathodic curve i with $\nu = 5$ mV/s. In present work, the construction of an imaginary cathodic line is explained in the following and illustrated in Fig 6 and Table 3.

Firstly, we subtract the curves with $\nu = 3$ mV/s from that with $\nu = 5$ mV/s, and fitting the difference into a line in $\log \Delta i-E$ graph with a intercept and a slope as shown in Fig 6A and denoted in Table 3 with the ν change of 5 mV/s $\rightarrow 3$ mV/s; we also subtract the curves with $\nu = 2$ mV/s and 1.5 mV/s from that with $\nu = 5$ mV/s and obtain the groups of slopes and intercepts, which are listed in Table 3 and denoted by the ν changes of 5 mV/s $\rightarrow 2$ mV/s and 5 mV/s $\rightarrow 1.5$ mV/s, respectively. Here, the slopes fluctuate with an average value; while the intercepts decrease with decreasing ν , which is consistent with Eq (23). Secondly, we choose the average value of above slopes and the linearly extrapolated value from the known intercepts in Table 3 to make the imaginary cathodic line for $\nu = 1$ mV/s i.e. moving the known imaginary cathodic line upward evenly, which is denoted by the ν change of 5 mV/s $\rightarrow 1$ mV/s.

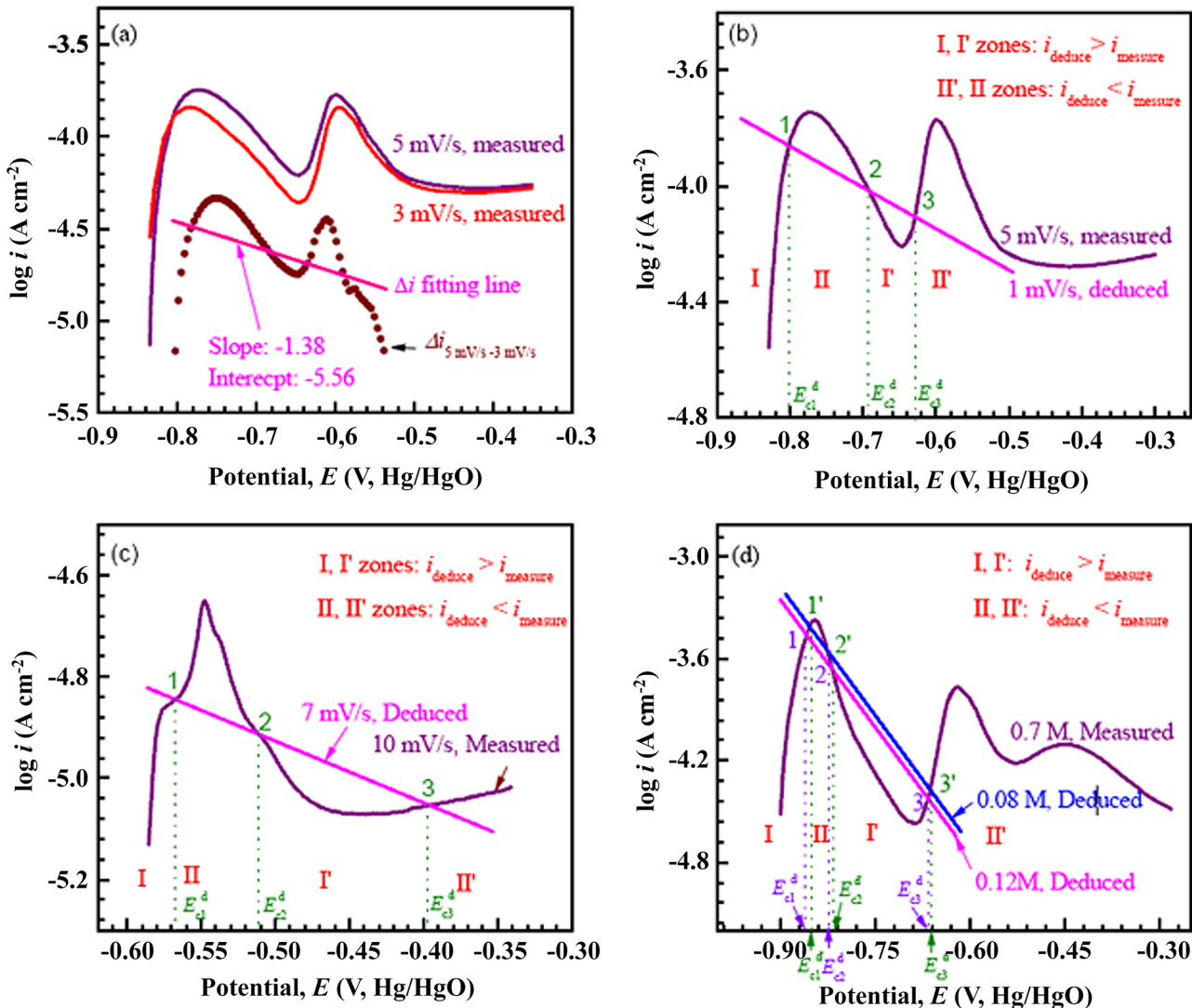


Fig 6. The schematic of practical cathodic and anodic curve intersection model. (a) Deducing of imaginary cathodic line from the difference of measured anodic curves; (b) deducing the corrosion potentials of Fe₇₈Si₉B₁₃ glassy ribbon with 1 mV/s from that with $v = 5$ mV/s in 0.6 M NaCl + 0.12 M NaOH solution by imposing an extrapolated imaginary cathodic line, i.e. moving evenly the imaginary cathodic line; (c) deducing the polarization curve of Fe_{73.5}Si_{13.5}B₉Cu₁Nb₃ glassy ribbon with $v = 7$ mV/s from that with $v = 10$ mV/s in 0.6 M NaCl + 0.12 M NaOH solution by rotating the imaginary cathodic line; and (d) deducing the polarization curve of Fe₇₈Si₉B₁₃ glassy ribbons with $c_{NaOH} = 0.12$ and 0.08 M from that with 0.7 M by rotating the imaginary cathodic line.

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After constructing the imaginary cathodic line for $v = 1$ mV/s, we impose it on the apparent anodic curve with $v = 5$ mV/s, and intersecting the apparent anodic curve at three points, and indicating that the deduced polarization curve with 1 mV/s has three corrosion potentials (Fig 6B). In addition, according to Eq (20), the current density symbol of I and I' zones, in which the anodic curve is lower than the imaginary cathodic line, is negative, being consistent with the measured polarization curve (Fig 1A). As $v = 0.5$ mV/s, the imaginary cathodic line, whose slope and intercept are determined with the v change of 5 mV/s \rightarrow 0.5 mV/s in Table 3, intersects the apparent anodic curve at about -0.6 V_{Hg/HgO} and is nearly tangent to P1 of the apparent anodic curve at about -0.8 V_{Hg/HgO}. Thus, a trough rather than E_{c1} or E_{c2} appears in the measured polarization curves with $v = 0.5$ mV/s. In short words, the imaginary cathodic line

Table 3. The values of slope and intercept for imaginary cathodic line (ICL) for the anodic and cathodic curve intersection model of the two samples; solution: 0.6 M NaCl+0.12 M NaOH solution.

Sample	Change of ν	Slope (A cm ⁻² /V)	Intercept(A cm ⁻²)	Method	Category
Fe ₇₈ Si ₉ B ₁₃	5 mV/s → 3 mV/s	-1.37±0.18	-5.56±0.11	d.m.	
	5 mV/s → 2 mV/s	-1.33±0.10	-5.20±0.07	d.m.	
	5 mV/s → 1.5 mV/s	-1.40±0.10	-5.10±0.07	d.m.	
	5 mV/s → 1 mV/s	-1.37	-4.95	e.d.m.	Evenly moving ICL
	5 mV/s → 0.5 mV/s	-1.37	-4.79	e.d.m.	Evenly moving ICL
Fe _{73.5} Si _{13.5} B ₉ Cu ₁ Nb ₃	10 mV/s → 9 mV/s	-0.87±0.18	-5.49±0.09	d.m.	
	10 mV/s → 8 mV/s	-1.05±0.03	-5.58±0.03	d.m.	
	10 mV/s → 7 mV/s	-1.23	-5.54	e.d.m.	Rotating ICL
	10 mV/s → 6 mV/s	-1.41	-5.54	e.d.m.	Rotating ICL

d.m.—from difference of measured polarization curves

e.d.m.—extrapolated from d.m.

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and apparent anodic curve intersection model can be used to deduce the multiple corrosion potentials in polarization curves.

Rotating the imaginary cathodic line. For the polarization curves of Fe_{73.5}Si_{13.5}B₉Cu₁Nb₃ glassy ribbons with various ν (Fig 2), the anodic and cathodic curve intersection model is also applied based on Eq (20). Here, because the curve with $\nu = 10$ mV/s is similar to that with $\nu = 9$ mV/s, Eqs (16) and (17) are used to describe the former and latter curves, respectively. Assuming

$$i_{0-1} = i'_{0-1} = \text{constant}, \quad i_{0-2} = i'_{0-2} = \text{constant} \tag{24a}$$

$$\beta_1 = \beta'_1, \quad \beta'_2 = \beta_2 + \frac{\delta'}{\psi} \tag{24b}$$

We will get

$$\Delta i = i - i' = i_{0-2} e^{\beta_2 \psi} (e^{\delta'} - 1) \tag{25}$$

Here, we set

$$\delta' = e^{-a' \psi} \tag{26}$$

where a' varies with ν and $a' > \beta_2$. Combining Eqs (25) with (26), expanding the right item of Eq (26) with Talor`s series and omitting the higher order terms, we will have

$$\log \Delta i = \log i_{0-2} + 0.434(\beta_2 - a')\psi \tag{27}$$

Eq (27) is also similar to Eq (14), so it can be considered as an imaginary cathodic line. With changing ν , its intercept is constant, but its slope changes, i.e. the line rotates around the intersection with vertical axis ($\psi = 0$), which is different from Eq (23). For predicating the corrosion potential with a given ν (Fig 2), we use the active anodic region of the curve with 10 mV/s as the apparent anodic curve. After subtracting the measured curves with 9 and 8 mV/s and linearly fitting the differences, the intercepts and slopes of the lines in $\log \Delta i - E$ can be obtained, which are listed in Table 3 and denoted by changing ν with 10 mV/s → 9 mV/s and 10 mV/s → 8 mV/s, respectively. Here, the variation of intercept is smaller than that of slope, which is consistent with Eq (27). So the average value of intercepts is chosen for constructing an imaginary cathodic line for $\nu = 7$ mV/s, and its slope is linearly extrapolated from the first two slopes

Table 4. The measured and deduced value E_{corr}^m and E_{corr}^d for the two samples; solution: 0.6 M NaCl + 0.12 M NaOH solution.

Sample	ν	E_{cl}^m (mV _{Hg/HgO})	E_{cl}^d (mV _{Hg/HgO})	E_{c2}^m (mV _{Hg/HgO})	E_{c2}^d (mV _{Hg/HgO})	E_{c3}^m (mV _{Hg/HgO})	E_{c3}^d (mV _{Hg/HgO})
Fe ₇₈ Si ₉ B ₁₃	1 mV/s	-824	-802	-784	-690	-618	-630
Fe _{73.5} Si _{13.5} B ₉ Cu ₁ Nb ₃	7 mV/s	-558	-568	-514	-511	-392	-398

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i.e. rotating the known imaginary cathodic line around its intersection with vertical axis (as shown by the line in Fig 6C and denoted by the ν change of 10 mV/s \rightarrow 7 mV/s in Table 3). After constructing the imaginary cathodic line with $\nu = 7$ mV/s, we impose it on the apparent anodic curve and get three interceptions at different potentials, which is consistent with the measured curve (Fig 2). In addition, according to Eq (25), the current density symbol of I and I' zones is negative, being consistent with the measured polarization curve (Fig 2). As $\nu = 6$ mV/s, the imaginary cathodic line, corresponding to the ν change of 10 mV/s \rightarrow 6 mV/s in Table 3, intersects the apparent anodic curve at about $-0.3 V_{Hg/HgO}$ and simultaneously tends to be tangent to the current density peak of the apparent anodic curve. Thus, the measured polarization curve with $\nu = 6$ mV/s contains a trough before the corrosion potential.

Meanwhile, the measured corrosion potentials E_{cl}^m , E_{c2}^m and E_{c3}^m and the deduced ones E_{cl}^d , E_{c2}^d and E_{c3}^d for Fe₇₈Si₉B₁₃ and Fe_{73.5}Si_{13.5}B₉Cu₁Nb₃ glassy ribbons in 0.6 M NaCl + 0.12 M NaOH solution with $\nu = 1$ and 7 mV/s are summarized in Table 4, respectively. For Fe₇₈Si₉B₁₃ glassy ribbons, the deduced E_{cl}^d or E_{c2}^d respectively is slightly higher than the corresponding measured E_{cl}^m or E_{c2}^m , but E_{c3}^d is lower than E_{c3}^m . It is explained by the fact that the first peak of measured curves moves to left side with decreasing ν and their second peak moves to right (Fig 1A). For Fe_{73.5}Si_{13.5}B₉Cu₁Nb₃ glassy ribbons, each measured corrosion potential is very close to the deduced counterpart, which is consistent with the fact that the current density peaks of the polarization curves approximately have no lateral displacement (Fig 2).

In addition, using the anodic and cathodic curve intersection model, the evolution of polarization curves with various hydroxyl concentrations can also be analyzed by the imaginary cathodic line with Eqs (23) or (27). The active anodic region of the curve obtained in the solution with $c_{NaOH} = 0.7$ M is chosen as the apparent anodic curve. After subtracting the curves obtained in solutions with $c_{NaOH} = 0.5, 0.4$ and 0.2 M and linearly fitting the difference, various intercepts and slopes could be obtained, which are summarized in Table 5 and denoted by the c_{NaOH} changes of 0.7 M \rightarrow 0.5 M, 0.7 M \rightarrow 0.4 M and 0.7 M \rightarrow 0.2 M, respectively. The intercept decreases slightly with the decrease of NaOH concentration, which is opposite to the desired varying tendency of an imaginary cathodic line in our intersection model. Hence, to deduce the corrosion potentials with $c_{NaOH} = 0.12, 0.08$ and 0.04 M, we select Eq (27) as the

Table 5. The values of slope and intercept for imaginary cathodic line (ICL) for the anodic and cathodic curve intersection model of Fe₇₈Si₉B₁₃ glassy ribbon; solution: 0.6 M NaCl + x M NaOH.

Change of c_{NaOH}	Slope(A cm ⁻² /V)	Intercept(A cm ⁻²)	Method	Category
0.7 M \rightarrow 0.5 M	-4.18 \pm 0.15	-7.57 \pm 0.11	d.m.	
0.7 M \rightarrow 0.4 M	-4.42 \pm 0.20	-7.71 \pm 0.15	d.m.	
0.7 M \rightarrow 0.2 M	-4.81 \pm 0.25	-7.92 \pm 0.19	d.m.	
0.7 M \rightarrow 0.12 M	-4.98	-7.73	e.d.m.	Rotating ICL
0.7 M \rightarrow 0.08 M	-5.06	-7.73	e.d.m.	Rotating ICL
0.7 M \rightarrow 0.04 M	-5.15	-7.73	e.d.m.	Rotating ICL

d.m.—from difference of measured polarization curves

e.d.m.—extrapolated from d.m.

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Table 6. The measured and deduced value E_{corr}^m and E_{corr}^d for $Fe_{78}Si_9B_{13}$ glassy ribbons; solution: 0.6 M NaCl + 0.12 and 0.08 M NaOH.

c_{NaOH}	E_{c1}^m (mV _{Hg/HgO})	E_{c1}^d (mV _{Hg/HgO})	E_{c2}^m (mV _{Hg/HgO})	E_{c2}^d (mV _{Hg/HgO})	E_{c3}^m (mV _{Hg/HgO})	E_{c3}^d (mV _{Hg/HgO})
0.12 M	-824	-858	-784	-819	-618	-667
0.08 M	-812	-855	-770	-825	-614	-663

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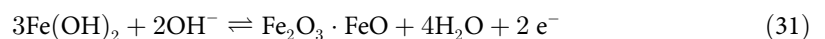
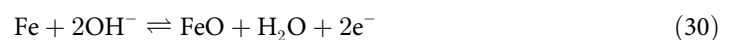
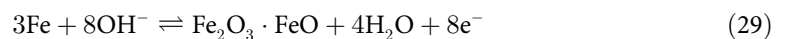
starting point, set the intercept as a constant equal to the average value of the first three intercepts, and linearly extrapolate the slope by the first three slopes i.e. rotating the known imaginary cathodic line around its intersection with vertical axis ($\psi = 0$ (as shown by the lines in Fig 6D) and denoted by the c_{NaOH} changes of 0.7 M \rightarrow 0.12 M and 0.7 M \rightarrow 0.08 M in Table 5)).

After getting the imaginary cathodic line, the anodic and cathodic curve intersection model can be used to deduce the corrosion potentials with 0.12 and 0.08 M NaOH solutions (Fig 6D). As $c_{NaOH} = 0.12$ and 0.08 M, both the imaginary cathodic lines whose slope and intercept are determined with the c_{NaOH} changes of 0.7 M \rightarrow 0.12 M and 0.7 M \rightarrow 0.08 M, respectively in Table 5, intersect the apparent anodic curve at three corrosion potentials (Fig 6D). According to Eq (25), the deduced current density symbol of I and I' zones is negative and consistent with the measured polarization curve (Fig 3). As $c_{NaOH} = 0.04$ M, the imaginary cathodic line, whose slope and intercept are determined with the c_{NaOH} change of 0.7 M \rightarrow 0.04 M in Table 5, intersects the apparent anodic curve at about -0.6 V_{Hg/HgO} and is nearly to be tangent to P1 of the apparent anodic curve, being consistent with the emergence of the trough in the measured polarization curve with $c_{NaOH} = 0.04$ M. The measured and deduced corrosion potentials of 0.12 and 0.08 M are summarized in Table 6. Apparently, three corrosion potentials E_{c1} , E_{c2} and E_{c3} in the polarization curves are successfully predicted. The deduced corrosion potential E_{c1}^d , E_{c2}^d or E_{c3}^d is slightly lower than the corresponding measured E_{c1}^m , E_{c2}^m or E_{c3}^m respectively, which is consistent with the fact that two current density peaks move to right with decreasing c_{NaOH} (Fig 3).

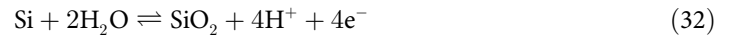
It should be noted that choosing the starting polarization curve is very important in our anodic and cathodic curve intersection model. If the ν or c_{NaOH} of the starting polarization curves is far beyond the ν or c_{NaOH} for multiple potentials phenomenon, our model will become invalid. Moreover, if the ν or c_{NaOH} of the starting polarization curves is far below the ν or c_{NaOH} for the multiple potentials phenomenon, our model will fail to predict it. It should be noted that the applicable ν range of our intersection model for $Fe_{78}Si_9B_{13}$ ribbon is different from that for $Fe_{73.5}Si_{13.5}B_9Cu_1Nb_3$ ribbon, due to the occurrence of their multiple corrosion potentials at different ν .

Nature of the first and second peaks in the anodic polarization curves of Fe-Si-B glassy ribbons

According to the previous explanation about the anodic dissolution of Fe-based alloys in alkaline solution [38], the first current density peak (P1) can be expressed

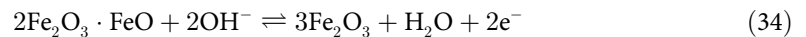
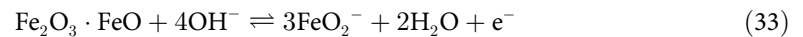


Kelsall et al. [39] have reported that the oxidation of silicon in steel is in preference to iron at low potentials. In present case, Eq (32) should be included in the first current peak P1.



Since SiO_2 bears a high chemical stability and a dielectric character [40,41], it is understood that the first peak is an ohmic resistance controlled reaction (Fig 1B).

For the second peak (P2), it is a process of the formation of stable Fe_2O_3 [42], and can be expressed



Kelsall et al. [39,43] have pointed out that Fe_2SiO_4 can be formed from iron oxide and SiO_2 /Si (IV) species at higher potentials. So the second current peak P2 should also contain the following reaction



According to the previous paper [44], both Eqs (33) and (34) are controlled by the adsorption step of OH^- , and it is expected that the second peak is characterized as the adsorption controlled reaction (Fig 1C).

As shown in Fig 1A, sample S_1 has only experienced the first peak (P1) with the scanning rate $\nu = 2$ mV/s in 0.6 M NaCl + 0.12 M NaOH solution, S_2 has experienced the first and second peaks (P1 and P2) with $\nu = 2$ mV/s. As indicated by reactions (28)–(32), the first current peak P1 is the formation process of FeO, Fe_2O_3 and SiO_2 . And the second current peak P2 is the conversion process from FeO to Fe_2O_3 (Eqs (33) and (34)). In addition, the formation of P2 consumes some SiO_2 according to Eq (35). These reactions are confirmed by the higher SiO_2 and FeO fractions of S_1 in XPS results compared with S_2 (Fig 5 and Table 2).

Sample S_3 has experienced three corrosion potentials (E_{c1} , E_{c2} and E_{c3}) and P2, ending the polarization at a similar potential to S_2 with $\nu = 1$ mV/s. The XPS decomposition similarity between S_2 and S_3 (Fig 5 and Table 2) indicates that, although the polarization curves of S_3 contains three corrosion potentials E_{c1} , E_{c2} and E_{c3} at the expense of the formation of P1, the anodic reaction of P1 still happened, which is hidden by the enhanced cathodic reaction. In other words, it is reasonable to obtain the measured anodic curves with $\nu = 5$ mV/s as the apparent anodic curve in our practical model.

According to the above analysis, the distinct peak P1 with formation of FeO, Fe_2O_3 and SiO_2 is necessary condition for the appearance of three corrosion potentials according to the practical model; while the second peak is not the necessary condition for their appearance in polarization curves. It should be noted that the formation of the anodic peak P1 or P2 is closely associated with the composition and microstructure of Fe-based alloys.

Conclusions

Two anodic peaks (P1 and P2) present in the polarization curve of $\text{Fe}_{78}\text{Si}_9\text{B}_{13}$ glassy ribbon in NaCl + NaOH solution when the scanning rate ν or NaOH concentration c_{NaOH} is high enough. The former peak formation is ohmic resistance controlled, which is associated with the formation of SiO_2 , and the latter is adsorption controlled, accompanying the damage of the SiO_2 film. The formation of P1 or P2 is closely associated with the composition and microstructure of Fe-based alloys. Three corrosion potentials (E_{c1} , E_{c2} and E_{c3}) are observed in the

polarization curve for Fe-based glassy ribbons in OH^- contained solutions when v or c_{NaOH} reaches a specific value.

The occurrence of multiple corrosion potentials is explained by a practical anodic and cathodic curve intersection model. In this model, we choose the measured anodic curve as the apparent anodic curve, from it subtract the measured anodic curves with various v or c_{NaOH} and linearly fit the difference as the imaginary cathodic line. By moving the cathodic line evenly or rotating it and imposing it with the apparent anodic curve, the number of intersections can be obtained and the number of corrosion potentials can be predicted. The practical model shows that the distinct P1 is the necessary condition for the occurrence of three corrosion potentials.

Acknowledgments

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Author Contributions

Conceived and designed the experiments: WMW. Performed the experiments: YJL YGW BA YL. Analyzed the data: WMW. Contributed reagents/materials/analysis tools: WMW. Wrote the paper: WMW LCZ HYM. Modified the revised manuscript: HX.

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