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**(Dimethylphosphoryl)methanaminium
hydrogen oxalate–oxalic acid (2/1)**

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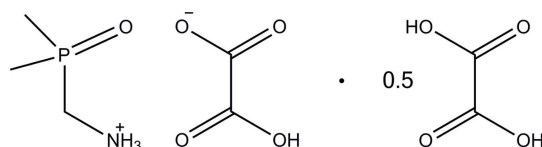
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Key indicators: single-crystal X-ray study; $T = 296$ K; mean $\sigma(\text{C}-\text{C}) = 0.002$ Å; R factor = 0.019; wR factor = 0.052; data-to-parameter ratio = 15.6.

The reaction of (dimethylphosphoryl)methanamine (dpma) with oxalic acid in ethanol yielded the title solvated salt, $\text{C}_3\text{H}_{11}\text{NOP}^+\cdot\text{C}_2\text{HO}_4^-\cdot 0.5\text{C}_2\text{H}_2\text{O}_4$. Its asymmetric unit consists of one dpmaH^+ cation, one hydrogen oxalate anion and a half-molecule of oxalic acid located around a twofold rotation axis. The H atom of the hydrogen oxalate anion is statistically disordered over two positions that are *trans* to each other. The hydrogen oxalate monoanion is not planar (bend angle $\sim 16^\circ$) whereas the oxalic acid molecule shows a significantly smaller bend angle ($\sim 7^\circ$). In the crystal, the components are connected by strong $\text{O}-\text{H}\cdots\text{O}$ and much weaker $\text{N}-\text{H}\cdots\text{O}$ hydrogen bonds, leading to the formation of layers extending parallel to (001). The structure was refined from a racemically twinned crystal with twin components in an approximate 1:1 ratio.

Related literature

For transition metal complexes of the cationic dpmaH^+ ligand, see: Reiss (2013*a,c*). For simple dpmaH^+ salts, see: Reiss & Jörgens (2012); Bianga *et al.* (2013); Buhl *et al.* (2013); Lambertz *et al.* (2013); Reiss (2013*b,d*).



Experimental

Crystal data

$\text{C}_3\text{H}_{11}\text{NOP}^+\cdot\text{C}_2\text{HO}_4^-\cdot 0.5\text{C}_2\text{H}_2\text{O}_4$
 $M_r = 242.14$
 Orthorhombic, $P2_12_12$

$a = 11.1482$ (6) Å
 $b = 13.0903$ (7) Å
 $c = 7.0432$ (4) Å

$V = 1027.84$ (10) Å³
 $Z = 4$
 Mo $K\alpha$ radiation

$\mu = 0.29$ mm⁻¹
 $T = 296$ K
 $0.52 \times 0.24 \times 0.14$ mm

Data collection

Bruker APEXII CCD
 diffractometer
 92924 measured reflections

3002 independent reflections
 2990 reflections with $I > 2\sigma(I)$
 $R_{\text{int}} = 0.020$

Refinement

$R[F^2 > 2\sigma(F^2)] = 0.019$
 $wR(F^2) = 0.052$
 $S = 1.05$
 3002 reflections
 192 parameters
 H atoms treated by a mixture of
 independent and constrained
 refinement

$\Delta\rho_{\text{max}} = 0.24$ e Å⁻³
 $\Delta\rho_{\text{min}} = -0.16$ e Å⁻³
 Absolute structure: refined as an
 inversion twin
 Absolute structure parameter:
 0.54 (7)

Table 1

Hydrogen-bond geometry (Å, °).

$D-\text{H}\cdots A$	$D-\text{H}$	$\text{H}\cdots A$	$D\cdots A$	$D-\text{H}\cdots A$
$\text{O2}-\text{H2}\cdots\text{O2}^{\text{i}}$	0.77 (3)	1.73 (4)	2.4959 (17)	177 (4)
$\text{O4}-\text{H4}\cdots\text{O4}^{\text{ii}}$	0.81 (4)	1.67 (4)	2.4694 (18)	171 (4)
$\text{N1}-\text{H1N}\cdots\text{O6}^{\text{iii}}$	0.84 (2)	2.56 (2)	3.2106 (16)	135.4 (19)
$\text{N1}-\text{H1N}\cdots\text{O7}$	0.84 (2)	2.27 (2)	2.9939 (17)	144 (2)
$\text{N1}-\text{H2N}\cdots\text{O5}$	0.86 (2)	2.03 (2)	2.8760 (14)	168 (2)
$\text{N1}-\text{H3N}\cdots\text{O3}^{\text{iv}}$	0.91 (2)	1.92 (2)	2.8030 (15)	166.2 (19)
$\text{N1}-\text{H3N}\cdots\text{O4}^{\text{iv}}$	0.91 (2)	2.49 (2)	3.0419 (16)	120.0 (16)
$\text{O6}-\text{H6}\cdots\text{O1}^{\text{v}}$	0.89 (3)	1.59 (3)	2.4701 (14)	168 (3)

Symmetry codes: (i) $-x + 1, -y, z$; (ii) $-x + 2, -y, z$; (iii) $-x + 1, -y + 1, z$; (iv) $-x + \frac{3}{2}, y + \frac{1}{2}, -z + 1$; (v) $x - \frac{1}{2}, -y + \frac{1}{2}, -z + 1$.

Data collection: APEX2 (Bruker, 2008); cell refinement: SAINT (Bruker, 2008); data reduction: SAINT; program(s) used to solve structure: SHELXS97 (Sheldrick, 2008); program(s) used to refine structure: SHELXL97 (Sheldrick, 2008); molecular graphics: DIAMOND (Brandenburg, 2012); software used to prepare material for publication: publCIF (Westrip, 2010).

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Supporting information for this paper is available from the IUCr electronic archives (Reference: WM5001).

References

- Bianga, C. M., Eggeling, J. & Reiss, G. J. (2013). *Acta Cryst.* **E69**, o1639–o1640.
 Brandenburg, K. (2012). *DIAMOND*. Crystal Impact GbR, Bonn, Germany.
 Bruker (2008). APEX2 and SAINT. Bruker AXS Inc., Madison, Wisconsin, USA.
 Buhl, D., Gün, H., Jablonka, A. & Reiss, G. J. (2013). *Crystals*, **3**, 350–362.
 Lambertz, C., Lupp, A. & Reiss, G. J. (2013). *Z. Kristallogr. New Cryst. Struct.* **228**, 227–228.
 Reiss, G. J. (2013*a*). *Acta Cryst.* **E69**, m248–m249.
 Reiss, G. J. (2013*b*). *Acta Cryst.* **E69**, o1253–o1254.
 Reiss, G. J. (2013*c*). *Acta Cryst.* **E69**, m250–m251.
 Reiss, G. J. (2013*d*). *Z. Kristallogr. New Cryst. Struct.* **228**, 431–433.
 Reiss, G. J. & Jörgens, S. (2012). *Acta Cryst.* **E68**, o2899–o2900.
 Sheldrick, G. M. (2008). *Acta Cryst.* **A64**, 112–122.
 Westrip, S. P. (2010). *J. Appl. Cryst.* **43**, 920–925.

supplementary materials

Acta Cryst. (2014). E70, o312 [doi:10.1107/S1600536814002931]

(Dimethylphosphoryl)methanaminium hydrogen oxalate–oxalic acid (2/1)**Sebastian Bialek, Rebecca Clemens and Guido J. Reiss****1. Comment**

The (dimethylphosphoryl)methanaminium (dpmaH⁺) cation is able to build various hydrogen-bonded, one dimensional structures (Bianga *et al.*, 2013; Buhl *et al.*, 2013; Lambertz *et al.*, 2013; Reiss, 2013*a,b*, Reiss & Jörgens, 2012).

However, we have shown that structures with a higher dimensional cross-linking by hydrogen bonds are also possible (Reiss, 2013*c,d*). Furthermore, it has been shown that the dpmaH⁺ cation is able to coordinate transition metal cations by the oxygen atom of its phosphoryl group (Reiss, 2013*a,c*). In this contribution we present a further example of a dpmaH⁺ salt, 2(C₃H₁₁NPO⁺ C₂HO₄⁻)C₂H₂O₄, owing a complex hydrogen bonding scheme.

As illustrated in Fig. 1 the asymmetric unit of the title structure consists of one (dimethylphosphoryl)methanaminium (dpmaH⁺) cation, one hydrogen oxalate monoanion in general positions and one half of an oxalic acid molecule located around a twofold rotation axis. The geometric parameters of the dpmaH⁺ cation, the hydrogen oxalate monoanion and the oxalic acid molecule are generally in the expected ranges. The hydrogen oxalate monoanion is, as might be expected, not planar (bent angle ~ 16°). The oxalic acid molecule has an imposed twofold rotation symmetry. This neutral molecule shows significantly smaller bent angles (~7°) than the hydrogen oxalate monoanion. The neutral molecules, cations and anions are connected by strong O—H···O and much weaker N—H···O hydrogen bonds (Table 1). The O···O distances range from 2.4694 (18) Å to 2.4959 (17) Å whereas the N···O distances are 2.8030 (15) Å, 2.8760 (14) Å and 2.9939 (17) Å. The latter may be interpreted at least as a weak hydrogen bond.

For the further structural discussion we only consider the shorter hydrogen bonding connections (Fig. 1; O···O and N···O < 2.9 Å). The dpmaH⁺ cation appears, under this assumption, as a twofold hydrogen bond donor (connected to two anions) and a single hydrogen bond acceptor (connected to an oxalic acid molecule). As shown in Fig. 2, two dpmaH⁺ cations are connected to one oxalic acid molecule by two symmetry-related O—H···O hydrogen bonds of medium strength. The hydrogen oxalate monoanion acts as a single hydrogen bond donor and a twofold hydrogen bond acceptor. The hydrogen oxalate anions form head-to-tail connected polar chains running along [100] *via* medium strength O—H···O hydrogen bonds, as illustrated in Fig. 3. Caused by a 1:1 disorder of the hydrogen atom of this anion (attached to O2/O4) polar chains are present, which are oriented along and against [100], respectively. These chains are connected to the dpmaH⁺ cation through N—H···O hydrogen bonds into a two-dimensional-structure parallel to (001). Characteristic for this arrangement are gaps between the anion chains and the oxalic acid molecules and the hydrophobic areas where neighbouring layers are facing each other (Fig. 4).

2. Experimental

The title compound, 2(C₃H₁₁NPO⁺ C₂HO₄⁻)C₂H₂O₄, was prepared by dissolving 1.01 g dpma and 1.17 g oxalic acid in 5 ml ethanol. Within a few days under ambient conditions, colourless crystals were obtained by slow evaporation of the solvent.

3. Refinement

For the disordered hydrogen atom at the hydrogen oxalate anion the split positions were refined freely with a ratio of 1:1. One common U_{iso} value was refined for both sites. All other H atoms were identified in difference syntheses and refined freely with individual U_{iso} values.

Computing details

Data collection: *APEX2* (Bruker, 2008); cell refinement: *SAINT* (Bruker, 2008); data reduction: *SAINT* (Bruker, 2008); program(s) used to solve structure: *SHELXS97* (Sheldrick, 2008); program(s) used to refine structure: *SHELXL97* (Sheldrick, 2008); molecular graphics: *DIAMOND* (Brandenburg, 2012); software used to prepare material for publication: *pubCIF* (Westrip, 2010).

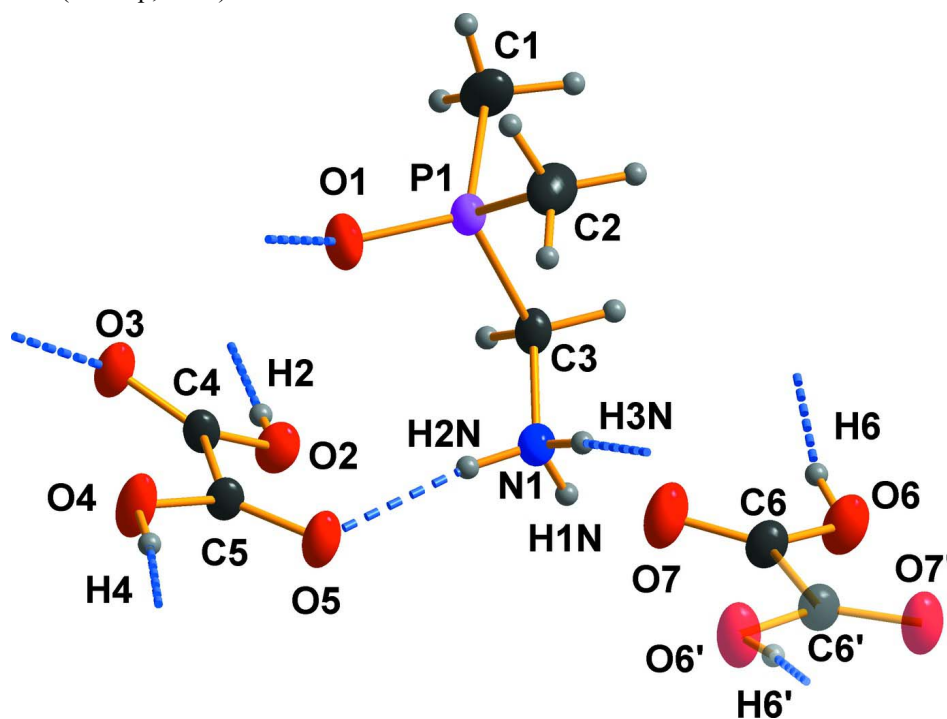


Figure 1

The molecular moieties of the title structure. Displacement ellipsoids are drawn at the 50% probability level. [Symmetry code: ' = 1 - x, 1 - y, z.]

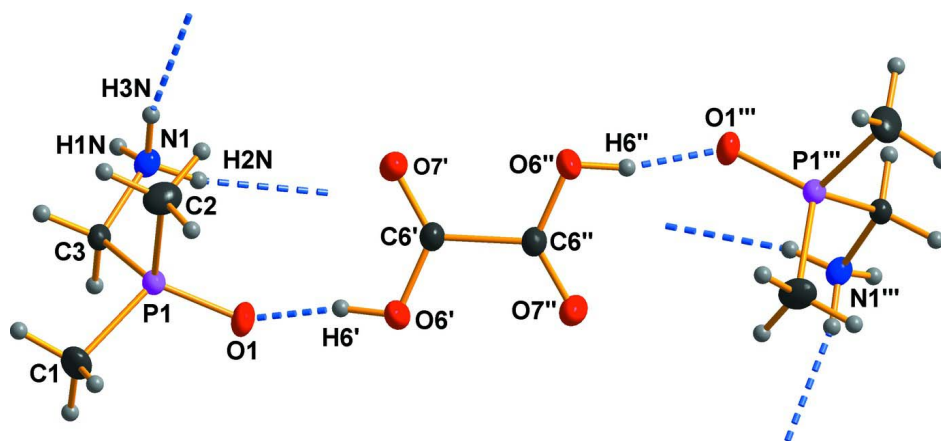


Figure 2

Two dpmaH^+ cations connected *via* an oxalic acid molecule. Displacement ellipsoids as in Fig. 1. [Symmetry codes: ' = $0.5 + x, 0.5 - y, 1 - z$; '' = $1.5 - x, -0.5 + y, 1 - z$; ''' = $2 - x, -y, z$.]

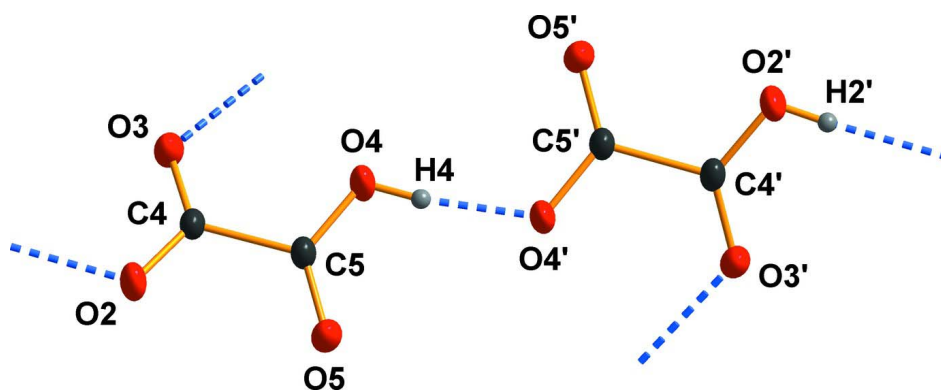


Figure 3

Part of the polar chain oriented along $[100]$, built by head-to-tail connected hydrogen oxalate anions. For clarity, just one of the disordered hydrogen atoms is shown. Displacement ellipsoids as in Fig. 1. [Symmetry code: ' = $2 - x, -y, z$.]

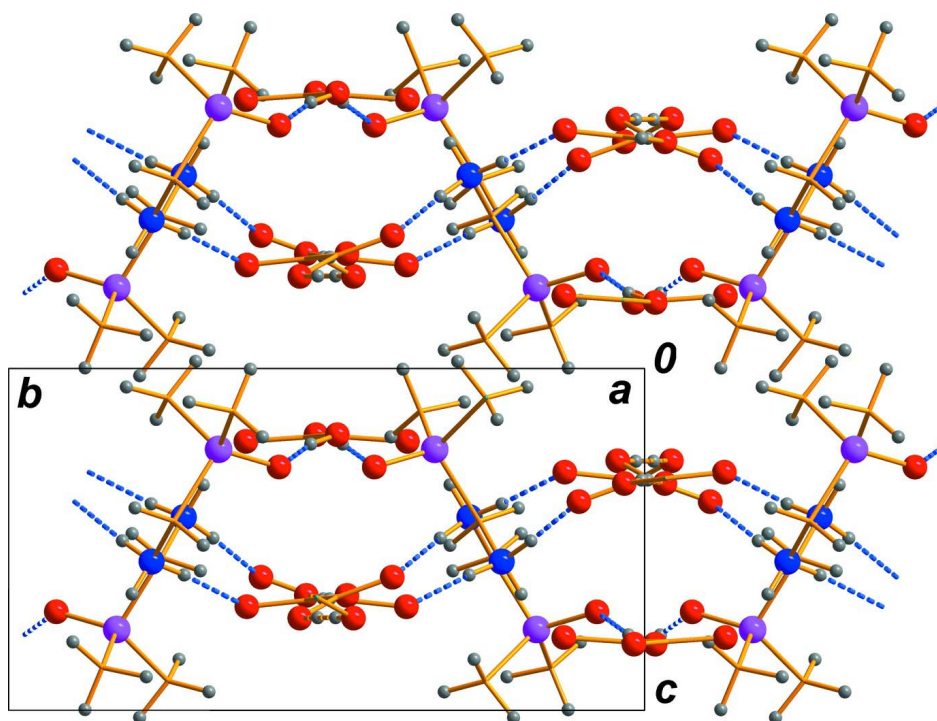


Figure 4

The two-dimensional framework parallel to (001) constructed by hydrogen bonding interactions (dashed lines).

(Dimethylphosphoryl)methanaminium hydrogen oxalate; oxalic acid

Crystal data

$C_3H_{11}NOP^+ \cdot C_2HO_4^- \cdot 0.5C_2H_2O_4$

$M_r = 242.14$

Orthorhombic, $P2_12_12$

$a = 11.1482$ (6) Å

$b = 13.0903$ (7) Å

$c = 7.0432$ (4) Å

$V = 1027.84$ (10) Å³

$Z = 4$

$F(000) = 508$

$D_x = 1.565$ Mg m⁻³

Mo $K\alpha$ radiation, $\lambda = 0.71073$ Å

Cell parameters from 9422 reflections

$\theta = 2.9$ – 35.9°

$\mu = 0.29$ mm⁻¹

$T = 296$ K

Lath, colourless

$0.52 \times 0.24 \times 0.14$ mm

Data collection

Bruker APEXII CCD
diffractometer

Radiation source: sealed tube

Graphite monochromator

φ and ω scans

92924 measured reflections

3002 independent reflections

2990 reflections with $I > 2\sigma(I)$

$R_{int} = 0.020$

$\theta_{max} = 30.0^\circ$, $\theta_{min} = 2.4^\circ$

$h = -15 \rightarrow 15$

$k = -18 \rightarrow 18$

$l = -9 \rightarrow 9$

Refinement

Refinement on F^2

Least-squares matrix: full

$R[F^2 > 2\sigma(F^2)] = 0.019$

$wR(F^2) = 0.052$

$S = 1.05$

3002 reflections

192 parameters

0 restraints

Primary atom site location: structure-invariant
direct methods

Secondary atom site location: difference Fourier map
 Hydrogen site location: difference Fourier map
 H atoms treated by a mixture of independent and constrained refinement

$w = 1/[\sigma^2(F_o^2) + (0.030P)^2 + 0.150P]$
 where $P = (F_o^2 + 2F_c^2)/3$
 $(\Delta/\sigma)_{\max} = 0.001$
 $\Delta\rho_{\max} = 0.24 \text{ e } \text{Å}^{-3}$
 $\Delta\rho_{\min} = -0.16 \text{ e } \text{Å}^{-3}$
 Absolute structure: refined as an inversion twin
 Absolute structure parameter: 0.54 (7)

Special details

Geometry. All e.s.d.'s (except the e.s.d. in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell e.s.d.'s are taken into account individually in the estimation of e.s.d.'s in distances, angles and torsion angles; correlations between e.s.d.'s in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell e.s.d.'s is used for estimating e.s.d.'s involving l.s. planes.

Refinement. Refined as a 2-component inversion twin.

Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters (Å²)

	x	y	z	$U_{\text{iso}}^*/U_{\text{eq}}$	Occ. (<1)
P1	0.59355 (2)	0.16995 (2)	0.76357 (4)	0.02120 (7)	
O1	0.66126 (9)	0.07493 (7)	0.71183 (16)	0.0320 (2)	
C1	0.45282 (14)	0.14728 (13)	0.8751 (3)	0.0401 (3)	
H1A	0.402 (3)	0.101 (2)	0.800 (4)	0.077 (8)*	
H1B	0.415 (2)	0.216 (2)	0.902 (4)	0.066 (7)*	
H1C	0.467 (3)	0.120 (2)	0.997 (4)	0.069 (8)*	
C2	0.67941 (16)	0.25082 (12)	0.9144 (2)	0.0362 (3)	
H2A	0.696 (3)	0.208 (2)	1.021 (4)	0.066 (8)*	
H2B	0.637 (2)	0.3147 (18)	0.947 (3)	0.048 (6)*	
H2C	0.754 (2)	0.268 (2)	0.851 (4)	0.069 (8)*	
C3	0.55056 (10)	0.24000 (9)	0.55333 (18)	0.0226 (2)	
H3A	0.5027 (19)	0.1946 (16)	0.482 (3)	0.039 (5)*	
H3B	0.5075 (19)	0.3008 (15)	0.593 (3)	0.038 (5)*	
N1	0.65260 (10)	0.27530 (8)	0.43508 (17)	0.0253 (2)	
H1N	0.6245 (19)	0.3065 (17)	0.341 (3)	0.043 (5)*	
H2N	0.694 (2)	0.2243 (17)	0.397 (3)	0.040 (5)*	
H3N	0.6989 (19)	0.3213 (16)	0.497 (3)	0.037 (5)*	
O2	0.60038 (7)	0.04219 (7)	0.26969 (15)	0.02816 (18)	
O3	0.67262 (8)	-0.10318 (7)	0.38932 (17)	0.0314 (2)	
O4	0.89578 (8)	-0.03192 (7)	0.33626 (19)	0.0354 (2)	
O5	0.82187 (8)	0.12549 (7)	0.30304 (17)	0.0312 (2)	
H2	0.539 (3)	0.014 (3)	0.272 (5)	0.035 (7)*	0.5
H4	0.962 (3)	-0.007 (3)	0.328 (5)	0.035 (7)*	0.5
C4	0.68433 (9)	-0.01553 (8)	0.33008 (17)	0.0212 (2)	
C5	0.81095 (9)	0.03311 (9)	0.32141 (17)	0.0212 (2)	
H6	0.283 (2)	0.476 (2)	0.220 (4)	0.070 (8)*	
O6	0.34552 (9)	0.51669 (8)	0.19359 (19)	0.0380 (3)	
O7	0.45436 (10)	0.37300 (8)	0.2131 (2)	0.0404 (3)	
C6	0.44455 (11)	0.46465 (10)	0.20114 (19)	0.0271 (2)	

Atomic displacement parameters (\AA^2)

	U^{11}	U^{22}	U^{33}	U^{12}	U^{13}	U^{23}
P1	0.01823 (12)	0.01867 (12)	0.02670 (13)	0.00132 (9)	0.00031 (11)	-0.00042 (10)
O1	0.0279 (4)	0.0260 (4)	0.0421 (5)	0.0105 (3)	-0.0041 (4)	-0.0037 (4)
C1	0.0283 (6)	0.0398 (8)	0.0521 (9)	0.0016 (6)	0.0134 (7)	0.0109 (7)
C2	0.0451 (8)	0.0336 (7)	0.0299 (6)	-0.0056 (6)	-0.0105 (6)	-0.0022 (5)
C3	0.0179 (5)	0.0198 (5)	0.0300 (5)	0.0019 (4)	-0.0031 (4)	-0.0012 (4)
N1	0.0271 (5)	0.0211 (4)	0.0275 (5)	0.0033 (4)	0.0014 (4)	0.0007 (4)
O2	0.0134 (3)	0.0272 (4)	0.0439 (5)	-0.0003 (3)	-0.0022 (4)	0.0045 (4)
O3	0.0184 (4)	0.0201 (4)	0.0557 (6)	-0.0023 (3)	0.0030 (4)	0.0049 (4)
O4	0.0126 (3)	0.0220 (4)	0.0715 (7)	0.0001 (3)	-0.0009 (4)	0.0047 (4)
O5	0.0194 (4)	0.0187 (4)	0.0556 (6)	-0.0024 (3)	0.0005 (4)	0.0019 (4)
C4	0.0131 (4)	0.0197 (5)	0.0307 (5)	-0.0016 (4)	0.0017 (4)	-0.0027 (4)
C5	0.0129 (4)	0.0203 (4)	0.0305 (5)	-0.0011 (4)	0.0010 (4)	0.0004 (4)
O6	0.0199 (4)	0.0279 (5)	0.0661 (8)	-0.0022 (3)	0.0020 (4)	0.0074 (5)
O7	0.0291 (5)	0.0237 (4)	0.0685 (8)	-0.0027 (4)	-0.0058 (5)	0.0032 (5)
C6	0.0224 (5)	0.0237 (5)	0.0351 (6)	-0.0033 (4)	-0.0017 (4)	0.0014 (4)

Geometric parameters (\AA , $^\circ$)

P1—O1	1.5000 (9)	N1—H2N	0.86 (2)
P1—C2	1.7791 (15)	N1—H3N	0.91 (2)
P1—C1	1.7796 (15)	O2—C4	1.2759 (14)
P1—C3	1.8064 (12)	O2—H2	0.77 (3)
C1—H1A	0.99 (3)	O3—C4	1.2278 (15)
C1—H1B	1.01 (3)	O4—C5	1.2767 (14)
C1—H1C	0.95 (3)	O4—H4	0.81 (4)
C2—H2A	0.96 (3)	O5—C5	1.2223 (14)
C2—H2B	0.99 (2)	C4—C5	1.5497 (15)
C2—H2C	0.97 (3)	O6—C6	1.2984 (15)
C3—N1	1.4836 (16)	O6—H6	0.89 (3)
C3—H3A	0.94 (2)	O7—C6	1.2076 (15)
C3—H3B	0.97 (2)	C6—C6 ⁱ	1.544 (2)
N1—H1N	0.84 (2)		
O1—P1—C2	111.58 (7)	N1—C3—H3B	106.5 (12)
O1—P1—C1	114.37 (7)	P1—C3—H3B	108.3 (12)
C2—P1—C1	108.05 (9)	H3A—C3—H3B	112.9 (17)
O1—P1—C3	110.81 (6)	C3—N1—H1N	107.9 (15)
C2—P1—C3	109.28 (7)	C3—N1—H2N	110.4 (15)
C1—P1—C3	102.28 (7)	H1N—N1—H2N	109 (2)
P1—C1—H1A	112.1 (17)	C3—N1—H3N	112.0 (13)
P1—C1—H1B	107.9 (15)	H1N—N1—H3N	105.7 (19)
H1A—C1—H1B	114 (2)	H2N—N1—H3N	111.2 (19)
P1—C1—H1C	108.4 (18)	C4—O2—H2	111 (3)
H1A—C1—H1C	110 (2)	C5—O4—H4	113 (3)
H1B—C1—H1C	104 (2)	O3—C4—O2	126.07 (10)
P1—C2—H2A	103.0 (16)	O3—C4—C5	119.62 (10)
P1—C2—H2B	112.4 (13)	O2—C4—C5	114.31 (10)

H2A—C2—H2B	114 (2)	O5—C5—O4	126.48 (10)
P1—C2—H2C	109.1 (17)	O5—C5—C4	120.09 (10)
H2A—C2—H2C	110 (2)	O4—C5—C4	113.42 (9)
H2B—C2—H2C	109 (2)	C6—O6—H6	110.0 (18)
N1—C3—P1	114.51 (8)	O7—C6—O6	126.95 (12)
N1—C3—H3A	109.3 (12)	O7—C6—C6 ⁱ	121.53 (15)
P1—C3—H3A	105.6 (12)	O6—C6—C6 ⁱ	111.48 (13)
O1—P1—C3—N1	-61.08 (10)	O2—C4—C5—O5	16.86 (17)
C2—P1—C3—N1	62.27 (11)	O3—C4—C5—O4	15.53 (18)
C1—P1—C3—N1	176.61 (10)	O2—C4—C5—O4	-163.53 (12)
O3—C4—C5—O5	-164.08 (13)		

Symmetry code: (i) $-x+1, -y+1, z$.

Hydrogen-bond geometry (Å, °)

<i>D—H...A</i>	<i>D—H</i>	<i>H...A</i>	<i>D...A</i>	<i>D—H...A</i>
O2—H2...O2 ⁱⁱ	0.77 (3)	1.73 (4)	2.4959 (17)	177 (4)
O4—H4...O4 ⁱⁱⁱ	0.81 (4)	1.67 (4)	2.4694 (18)	171 (4)
N1—H1N...O6 ⁱ	0.84 (2)	2.56 (2)	3.2106 (16)	135.4 (19)
N1—H1N...O7	0.84 (2)	2.27 (2)	2.9939 (17)	144 (2)
N1—H2N...O5	0.86 (2)	2.03 (2)	2.8760 (14)	168 (2)
N1—H3N...O3 ^{iv}	0.91 (2)	1.92 (2)	2.8030 (15)	166.2 (19)
N1—H3N...O4 ^{iv}	0.91 (2)	2.49 (2)	3.0419 (16)	120.0 (16)
O6—H6...O1 ^v	0.89 (3)	1.59 (3)	2.4701 (14)	168 (3)

Symmetry codes: (i) $-x+1, -y+1, z$; (ii) $-x+1, -y, z$; (iii) $-x+2, -y, z$; (iv) $-x+3/2, y+1/2, -z+1$; (v) $x-1/2, -y+1/2, -z+1$.