


 Cite this: *RSC Adv.*, 2020, 10, 38205

# Fully inkjet-printed multilayered graphene-based flexible electrodes for repeatable electrochemical response†

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Graphene has proven to be useful in biosensing applications. However, one of the main hurdles with printed graphene-based electrodes is achieving repeatable electrochemical performance from one printed electrode to another. We have developed a consistent fabrication process to control the sheet resistance of inkjet-printed graphene electrodes, thereby accomplishing repeatable electrochemical performance. Herein, we investigated the electrochemical properties of multilayered graphene (MLG) electrodes fully inkjet-printed (IJP) on flexible Kapton substrates. The electrodes were fabricated by inkjet printing three materials – (1) a conductive silver ink for electrical contact, (2) an insulating dielectric ink, and (3) MLG ink as the sensing material. The selected materials and fabrication methods provided great control over the ink rheology and material deposition, which enabled stable and repeatable electrochemical response: bending tests revealed the electrochemical behavior of these sensors remained consistent over 1000 bend cycles. Due to the abundance of structural defects (e.g., edge defects) present in the exfoliated graphene platelets, cyclic voltammetry (CV) of the graphene electrodes showed good electron transfer ( $k = 1.125 \times 10^{-2} \text{ cm s}^{-1}$ ) with a detection limit (0.01 mM) for the ferric/ferrocyanide redox couple,  $[\text{Fe}(\text{CN})_6]^{-3/-4}$ , which is comparable or superior to modified graphene or graphene oxide-based sensors. Additionally, the potentiometric response of the electrodes displayed good sensitivity over the pH range of 4–10. Moreover, a fully IJP three-electrode device (MLG, platinum, and Ag/AgCl) also showed quasi-reversibility compared to a single IJP MLG electrode device. These findings demonstrate significant promise for scalable fabrication of a flexible, low cost, and fully-IJP wearable sensor system needed for space, military, and commercial biosensing applications.

 Received 30th May 2020  
 Accepted 9th October 2020

DOI: 10.1039/d0ra04786d

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## Introduction

Graphene has been used for many electrochemical applications such as in fuel cells, electric double-layer capacitors, and even in lithium-ion batteries.<sup>1–5</sup> So far research has been conducted for graphene oxide electrodes, screen printed graphene electrodes, and IJP graphene electrodes modified with PEDOT-PSS or polyaniline, but fully IJP printed bare graphene-based electrodes with high stability, sensitivity, and repeatability have not been developed.<sup>2,6–18</sup> Conventional fabrication processes for sensor

development, such as vacuum deposition, photolithography, and epitaxial growth of electronic materials, tend to be complicated and expensive, often requiring lithographic patterning and high-temperature processing.<sup>19</sup> As a result, additive electronics manufacturing techniques, such as inkjet printing (IJP), aerosol jet printing (AJP), and micro-dispense printing (MDP), are being explored as potential low-cost scalable fabrication methods for flexible sensor systems.<sup>20–24</sup> Previous studies have demonstrated that inkjet printing, a drop-on-demand process, eliminates the need for the prefabricated masks or stencils required for lithographic and contact-printing processes.<sup>25–28</sup> An inkjet-printed ion-selective single layer reduced graphene oxide-based sensor by Claussen *et al.* demonstrated a wide sensing range and low detection limits.<sup>29</sup> However, such studies involved rigid substrates and high annealing temperatures that are not compatible with flexible substrates or included lithographic processes in the overall device design.

Over the last two decades, carbon-based materials such as graphene or functionalized/doped graphene, glassy carbon,

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† Electronic supplementary information (ESI) available. See DOI: 10.1039/d0ra04786d



carbon nanotubes (CNTs), fullerenes, and boron-doped diamond have been investigated for their use as electrodes in electrochemical sensing due to their advantageous structural and electrical properties.<sup>13,18,30–36</sup> As each carbon allotrope possesses a unique structure and surface chemistry, the electrochemical behavior of each is also unique. For electrochemical applications, highly ordered pyrolytic graphite, pseudographite, graphene, and orientated CNTs are excellent candidates due to their high conductivity, large surface area, and unique heterogeneous electron transfer rates.<sup>32,35,37</sup>

Graphene, a monolayer of  $sp^2$  hybridized carbon in a 2-dimensional hexagonal lattice structure, has received much attention in the research community due to its unique physical and chemical properties. The  $sp^2$  bonding between the carbon atoms in graphene creates three  $\sigma$ -bonds, which are responsible for its high mechanical strength and high in-plane thermal conductivity.<sup>38–42</sup> Graphene's remarkable conductivity is associated with overlapping  $p_z$  orbitals above and below the molecular plane, which creates a delocalized  $\pi$  - electron system to allow for free movement of electrons. These unique bonding characteristics give rise to a linear band structure with a zero-band gap near the  $K$  and  $K'$  points, leading to graphene's high electrical conductivity.<sup>39,43</sup> Moreover, graphene makes for an excellent candidate for electrochemical applications due to its high conductivity, large surface area, unique heterogeneous electron transfer rate, and low production cost. The edge plane and basal plane-defect sites of the highly ordered pyrolytic graphite greatly favors electrochemical activity.<sup>2,18</sup> Three common techniques used to obtain graphene are exfoliation, chemical vapor deposition (CVD), and epitaxial growth. While these are widely used techniques, they are known to introduce defects to the graphene structure that are detrimental to electrical and thermal transport properties, while conversely improving chemical and electrochemical sensitivity.<sup>1,5,42,44–48</sup>

The method of fabricating electrochemical graphene sensors is vital in creating edge and basal plane defects to improve chemical sensitivity. Work such as Banerjee *et al.* reported ultrahigh electrochemical current densities for graphene edges embedded in dielectric nanopores.<sup>1</sup> Yuan *et al.* further demonstrated that the electrochemical activity on the edge states of single-layer CVD grown graphene is higher than on the basal plane.<sup>49</sup> Shang *et al.* showed that increased graphitic edge and basal plane defects in CVD grown multilayer graphene resulted in superior electron transfer kinetics compared to the edge state of highly ordered pyrolytic graphite.<sup>50</sup> Fisher *et al.* used microwave plasma-enhanced chemical vapor deposition grown multilayered graphene petal nanosheets to develop a versatile glucose sensor on a silicon wafer with high sensitivity, selectivity, and stability.<sup>51</sup> Furthermore, Tang *et al.* showed excellent electrocatalytic activity for reduced graphene oxide sheets synthesized by chemical exfoliation and cast onto a glassy carbon electrode.<sup>52</sup> We chose a high yield, solvent assisted exfoliation method to synthesize multilayer graphene to retain desirable edge and basal-plane defects that promote electrochemical activity. Moreover, it has been previously shown that the annealing conditions, ink properties, and number of print passes impact the electrical and structural

porosity of printed graphene microstructures.<sup>53</sup> Such porosity in functionalized graphene electrodes has a significant impact on enhancing the electrochemical performance as well.<sup>54</sup>

In this report, the electrochemical performance of IJP MLG electrodes and fully IJP three-electrode sensors is investigated by measuring the cyclic voltammetry response of a ferric/ferrocyanide redox couple and by performing pH sensitivity studies. Additionally, the effect of electrode porosity is examined with a comparison between the electrochemical performance of MLG electrodes having different porosities as a result of the printing process. It is observed that the structure–property–processing correlations of fully additively manufactured graphene-based electrochemical electrodes are essential factors in improving consistency, repeatability, and uniformity of such fully printed sensor systems. Finally, the IJP MLG electrodes are shown to exhibit robust electrochemical performance over 1000 bend cycles, highlighting the attractive properties and behavior of IJP MLG electrodes for use in wearable electroanalysis. Advances such as this will further enable additive electronics manufacturing of flexible sensors for human performance monitoring in space, military, and commercial applications.

## Results and discussion

### Characterization of multilayer graphene (MLG) ink

The multilayer graphene ink (MLG) was formulated *via* solvent assisted exfoliation of bulk graphite powder using a process reported previously.<sup>53</sup> In Fig. S1,† we see the characteristic of the individual graphene flakes with TEM (Fig. S1c†) and Raman spectroscopy (inset Fig. S1b†). Raman spectroscopy reveals the characteristic of D, G, and 2D peaks for graphene at  $1350\text{ cm}^{-1}$ ,  $1580\text{ cm}^{-1}$ , and  $2700\text{ cm}^{-1}$ , respectively. The ratio of the D/G peak intensities ( $I_D/I_G$ ) determines the graphene flakes' quality (defect/disorder). The  $I_D/I_G$  peak ratio of 0.24 is lower than previously reported values (0.33–0.7), suggesting the exfoliated flakes are of higher quality with fewer defects.<sup>55</sup> TEM images show the graphene flakes vary in lateral size from  $\sim 50$ – $300\text{ nm}$ . To correlate the TEM and Raman data, the  $I_D/I_G$  peak ratio and 532 nm excitation wavelength was used in Cancado's general equation:<sup>56</sup>

$$L_a\text{ (nm)} = (2.4 \times 10^{-10}) \lambda_1^4 \left( \frac{I_D}{I_G} \right)^{-1}$$
 to extract the crystal size ( $L_a \approx 102\text{ nm}$ ) of the graphene flakes. AFM characterization of the flakes shows the thickness ( $t_g$ ) ranges from monolayer to flakes with an average thickness of  $t_g = 16\text{ nm} \pm 15\text{ nm}$  (Fig. S1d†).

This unique process results in a highly concentrated ink compatible with a variety of print modalities.<sup>53</sup> Specifically, the layered graphene/ethyl cellulose (EC) flakes were dispersed in a mixture of 85% cyclohexanone and 15% terpineol to yield a viscosity of  $\sim 8\text{ cP}$ , as measured with a Wells-Brookfield Cone/Plate viscometer at an ink concentration of  $3.5\text{ mg mL}^{-1}$ . To prepare stable and homogeneous MLG ink, one of the most important parameters is to optimize the ink viscosity. Viscosities and concentrations in this range have shown to be compatible with the ultrasonic atomizer of an aerosol jet printer

(AJP) and the Dimatix DMP 2850 materials inkjet printer (IJP).<sup>57</sup> Viscosity higher than 10 cP results in the printer's clogging and too low of value (<5 cP) doesn't allow to form a droplet. Furthermore, the ink's viscosity is greatly affected by the concentration and dispersing agents.<sup>58</sup> So, it is essential to carefully control the ink fabrication process as it can greatly affect the printability and quality of the printed patterns.

Microstructural and electrical characterization was performed on multilayered graphene (MLG) printed lines with silver contact pads in a transmission line measurement (TLM) structure with varying numbers of print passes (15–30, in increments of 5 passes) using a Dimatix inkjet printer (seen in Fig. 1a). To minimize interference due to excessive charging and fluorescence from the large surface roughness of Kapton HN substrates, the MLG and silver TLM structures for microstructural and electrical characterization were printed on glass instead. The tool platen temperature, nozzle diameter, and

cartridge temperature were optimized to ensure that the dimensions and material deposition were adequate to obtain uniform structures. Scanning electron microscopy (SEM) of a line with 25 print passes of IJP multilayered graphene (Fig. 1b – left), demonstrated good uniformity of the printed layer. Transmission electron microscopy (TEM) (Fig. 1b – right) showed multilayer graphene flakes ranging in thickness from 5–20 nm layers. The surfactant ethyl cellulose (EC) stabilized the graphene flakes in the solution but required decomposition after printing to achieve optimal electrical conductivity. Thermal gravimetric analysis (TGA) was employed to analyze the thermal stability of the EC coated graphene flakes. Fig. 1c shows weight percent (black) and differential scanning calorimetry (blue) as a function of temperature. The decomposition peak of EC occurs around 250 °C, while the decomposition temperature of the solvents is seen around 350 °C. Using the TGA, the graphene printed structures were then annealed in two

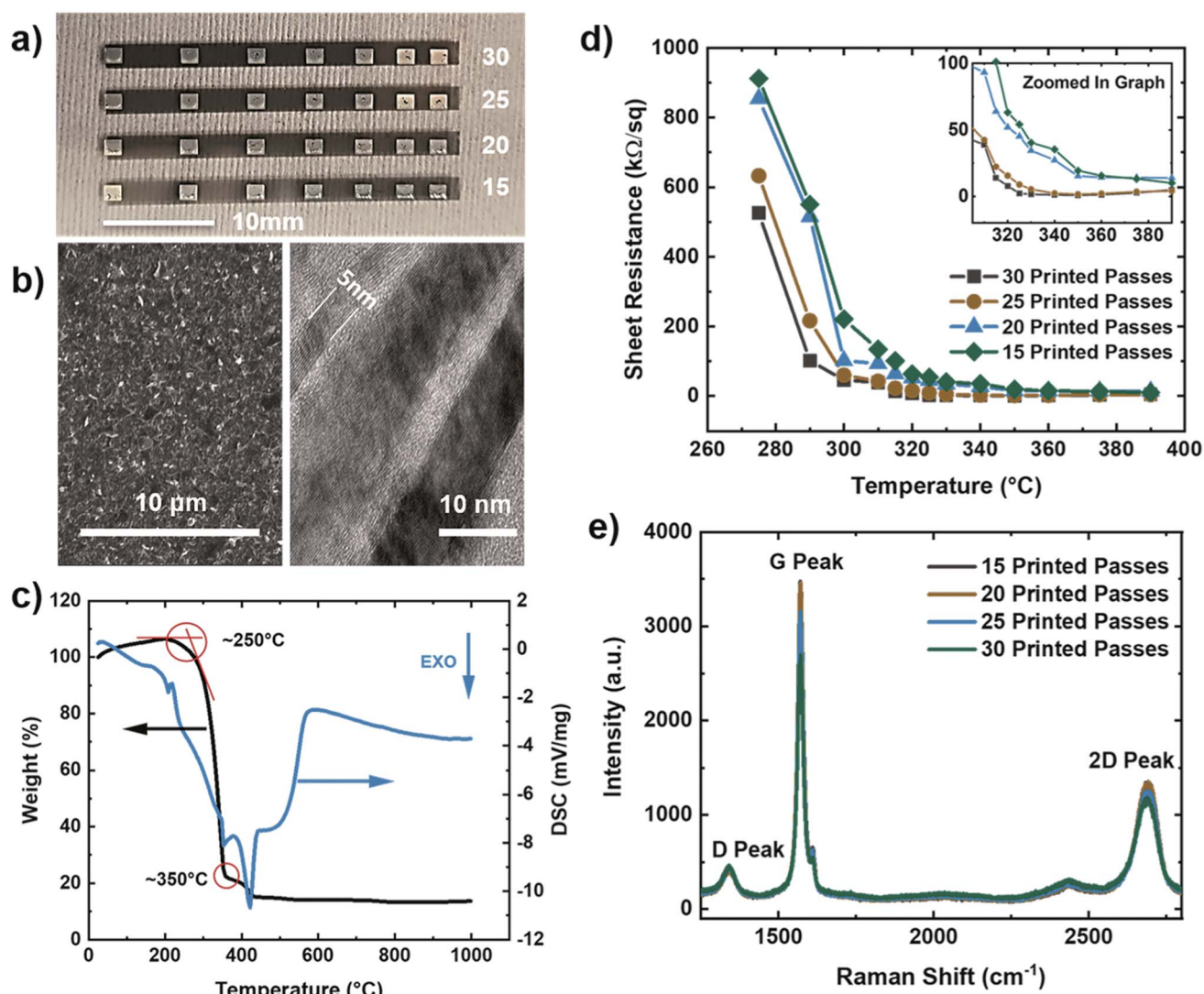


Fig. 1 Inkjet printed graphene layer characterization. (a) Optical image of inkjet printed graphene (15–30 printed passes) structure on glass. (b) SEM image of the 25-pass printed line (left) and TEM image of the multilayer graphene flakes (right) on Kapton HN. (c) TGA and DSC data of the graphene/ethyl cellulose ink. (d) Sheet resistance vs. temperature for 15 to 30 printed passes on glass. (e) Typical Raman spectra (532 nm excitation) for 15 to 30 printed passes MLG films.

stages: first at 250 °C for 30 minutes to evaporate the surfactant, and subsequently at 350 °C for another 30 minutes to remove the remaining solvents, leading to enhanced electrical conductivity. The silver contacts were deposited *via* IJP on top of the graphene in a TLM structure using commercially available 40 wt% silver ink.

To measure the electrical properties of the printed interconnects, a 2-point probe (Keithley 4200 SCS, Textronix, Beaverton, OR) measurements on the TLM structure were conducted. Fig. 1d shows the calculated sheet resistance as a function of annealing temperature for 15–30 print passes. Based on the TLM measurements, the lowest sheet resistance was calculated to be 0.89 kΩ sq<sup>-1</sup>, and 1.60 kΩ sq<sup>-1</sup> for 30 and 25 print passes, respectively, at an annealing temperature of 350 °C. Moreover, Raman spectroscopy results are shown in Fig. 1e reveal the graphitic nature of our electrodes through the characteristic D, G, and 2D peaks for the IJP graphene layers at 1350 cm<sup>-1</sup>, 1580 cm<sup>-1</sup>, and 2700 cm<sup>-1</sup>, and  $I_{2D}/I_G$  peak ratios in the range of 0.38–0.43.

### Electrochemical response of MLG electrodes

To investigate our IJP MLG electrode's electrochemical properties, the synthesized MLG was printed into 1 cm × 1 cm squares (15, 20, 25, and 30 printed passes) with silver contact pads (1 printed layer) and SU-8 (3 printed passes) as a passivation layer for the silver electrode. All layers were printed by IJP on 2 mil thick untreated Kapton HN film (Fig. 2a and b). Before printing MLG ink on Kapton, contact angle (CA) measurements of MLG inks on a Kapton HN substrate were performed to ensure wettability. In Fig. 2c, we present a low CA of 15.6°, suggesting good wettability of MLG ink on untreated Kapton HN substrates. An example of the flexible and fully printed MLG electrode with 25-layer print passes on Kapton HN is seen in Fig. 2d.

The experimental setup to study the electrochemical response of printed MLG electrodes is shown in Fig. 2e, like the setups used by Munoz and Richter.<sup>59–61</sup> Cyclic voltammetry (CV) measurements were carried out with a Bio-Logic VMP-300 potentiostat with scan rates from 10–100 mV s<sup>-1</sup> in 1 mM K<sub>3</sub>[Fe(CN)<sub>6</sub>] with 1 M KCl as the supporting electrolyte, at room temperature for MLG working electrodes made by 15 to 30 printed passes (ESI Fig. S2a, b† and 3a, b). Ferro/ferricyanide redox couple is often used in physiological experiments because of its sensitivity to relatively small surface changes. The iron is low spin and quickly reduce to the ferric/ferrocyanide Fe(CN)<sub>6</sub><sup>3-/4-</sup>.<sup>13</sup> For this experiment, the electrochemical cell is comprised of an IJP MLG working electrode, a standard platinum wire counter electrode and a conventional aqueous Ag/AgCl/KCl (sat.) reference electrode with saturated (sat.) KCl solution from SYC Technologies. A ~0.07 cm<sup>2</sup> circular surface area of IJP MLG is defined by the size of the O-ring in the liquid cell. Studies were performed with the 25 and 30 printed pass electrodes due to their superior electrochemical performance. The observed electrochemical behavior was evaluated, as can be seen from the CV curves in Fig. 3, distinct redox peaks can be observed. The 25 printed passes of MLG electrode exhibits

quasi-reversible CV characteristics with a peak to peak separation ~60 mV at a scan rate of 10 mV s<sup>-1</sup> in a 1 mM K<sub>3</sub>[Fe(CN)<sub>6</sub>] electrolyte. This peak to peak separation (*i.e.*, near-ideal  $\Delta E_p$  of 59 mV for Nernstian reactions) is evidence of the fast electrode kinetics, while shifts in peak to peak separation with an increase in scan rate point to the electrodes' quasi-reversible nature.

The kinetics (extracted dimensionless coefficient  $\phi$ ) of our IJP MLG electrode (25.14) is compared with that of other comparable graphene-based electrodes and inkjet-printed electrodes in Table 1. Electrodes shaded in pink shows functionality of a fully printed three electrode sensor system. Our electrode shows much better reversibility, stability, and repeatability on a flexible substrate than the other electrodes. Our data suggest that IJP MLG electrodes possess well-defined structures and electrochemical properties to support fast kinetics, comparable to results presented in the literature for MLG.<sup>51</sup> Although 30 printed passes of MLG shows a higher current and lower sheet resistance than 25 printed passes, a higher peak to peak separation of 80 mV (Fig. 3b) is observed. It is hypothesized that while the additional printed passes for the (30 printed passes sample) electrode increase uniformity and decrease resistance, they also create denser packing of the graphene flakes, thereby reducing the porosity and slowing the redox reaction on the electrode surface.

For further analysis, the cathodic and anodic peak currents ( $I_{pc}$  and  $I_{pa}$ , respectively) from the CV data were plotted *versus* the square root of the scan rate (V s<sup>(-1/2)</sup>) for both 25 and 30 printed passes (ESI Fig. S3a and b†). The cathodic ( $I_{pc}$ ) and anodic ( $I_{pa}$ ) peak currents for 25 and 30 printed passes graphene showed excellent linear regression, suggesting redox reaction controlled by diffusion. CV simulation was used to extract the diffusion coefficient ( $D$ ) and the rate constant ( $k$ ) from the experimental CV data (Fig. 3a and b). This simulation provided identical CV curves compared to the data for 25 and 30 printed passes, respectively, as seen in Fig. 3c and d. In order to calculate the active surface area, we accounted for surface roughness extracted from the AFM images of 25 printed passes and 30 printed passes (seen in ESI Fig. S4†). To compare our data, we also used Randles–Sevcik equation to calculate the electrochemically active surface area. All the steps for these calculations are presented in the supplementary active surface area calculations. The active surface area extracted from AFM images of ~0.086 cm<sup>2</sup> and ~0.084 cm<sup>2</sup> compared to the calculated electrochemically active surface area of ~0.091 cm<sup>2</sup>, and ~0.093 cm<sup>2</sup> for 25 printed passes 30 printed passes respectively, are in good agreement with 0.7% error. Using the calculated electrochemically active surface area, the diffusion coefficients are  $D = 4.17 \times 10^{-6}$  cm<sup>2</sup> s<sup>-1</sup> for 25 printed passes and  $D = 6.38 \times 10^{-6}$  cm<sup>2</sup> s<sup>-1</sup> for 30 printed passes, respectively. The calculated diffusion coefficient values are comparable with the ferric/ferrocyanide electrolyte reported in Konopka *et al.*<sup>62</sup> Furthermore, the average electron transfer rate constants, (25 printed passes)  $k = 1.125 \times 10^{-2}$  cm s<sup>-1</sup> with average double layer capacitance of 43.4 μF and (30 printed passes)  $k = 7.34 \times 10^{-3}$  cm s<sup>-1</sup> with average double layer capacitance of 45.5 μF, where  $\alpha = 0.5$  (shows symmetric free energy

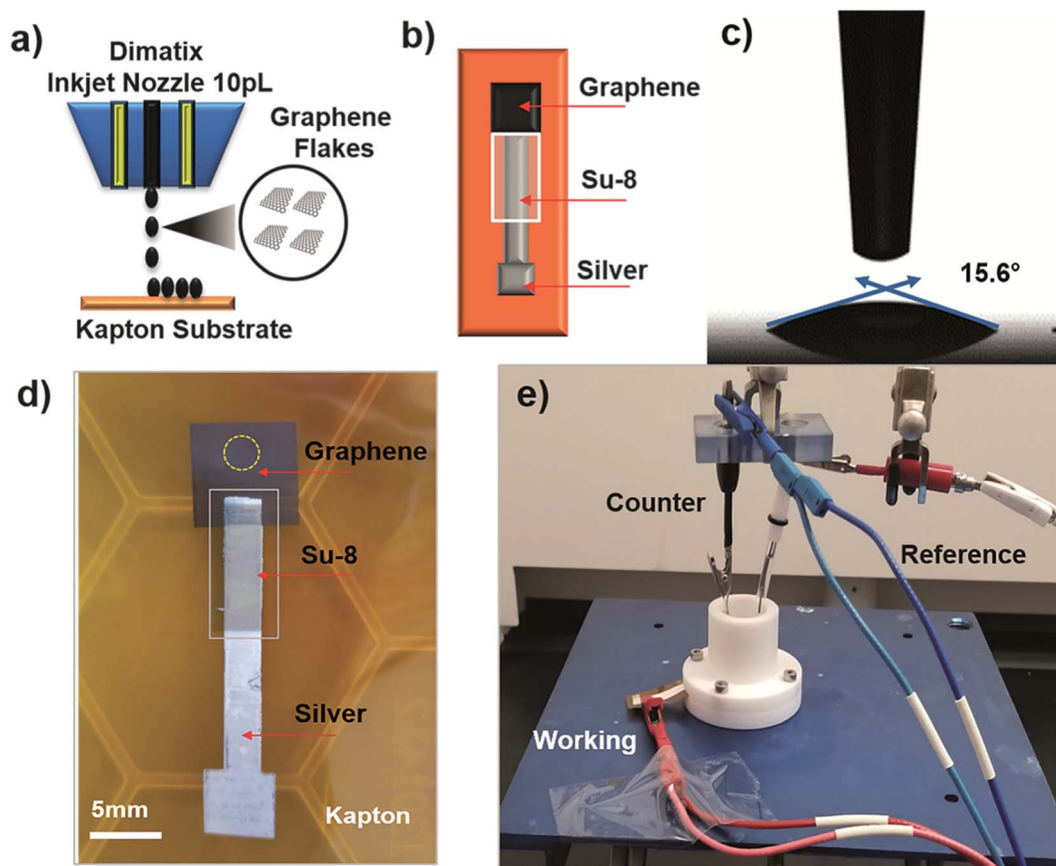


Fig. 2 (a) Sketch of the Dimatix inkjet printer printing graphene on Kapton. (b) Design and layers of the graphene electrode. (c) Contact angle measurements of graphene ink on Kapton. (d) Optical image of the printed graphene electrode, where the yellow ring indicates the surface area of the electrode exposed to the electrolyte. (e) Photograph of the electrochemical experimental setup of the graphene electrode in aqueous ferricyanide redox couple with Ag/AgCl as the reference electrode and platinum wire as the counter electrode.

curve, where the influence of applied voltage at transition state is about mid-way between reactants and products).<sup>63</sup>

To compare our extracted rate constant ( $k$ ), we used the dimensionless kinetic parameter equation to estimate the heterogenous standard rate constant ( $k_0$ ).<sup>64</sup>

$$\phi = (-0.6288 + 0.0021x)/(1 - 0.017x)$$

where the peak potential separation is ( $x$ ), multiplied by the number of electrons involved in the reaction ( $n$ ), with ferricyanide redox system is equal to one. The rate constant ( $k_0$ ) is then calculated using the equation

$$\phi = k_0 \left( \pi D \nu n \left( \frac{F}{RT} \right) \right)^{\frac{1}{2}}$$

where  $D$  is the diffusion coefficient of the redox mediator (potassium ferricyanide is about  $4.17 \times 10^{-6} \text{ cm}^2 \text{ s}^{-1}$ ),  $\nu$  is the scan rate ( $10 \text{ mV s}^{-1}$ ),  $F$  is the Faraday constant,  $R$  is the gas constant, and  $T$  is the temperature ( $25 \text{ }^\circ\text{C}$ ). The  $k_0$  of MLG was calculated as  $2.38 \times 10^{-2} \text{ cm s}^{-1}$  close to our extracted value of  $k = 1.125 \times 10^{-2} \text{ cm s}^{-1}$ .

Furthermore, we conducted CV with ferrocene methanol ( $\text{C}_{11}\text{H}_{12}\text{FeO}$ ) an outer sphere redox species, which is not sensitive to surface oxides and only depended on the density of

states.<sup>65–67</sup> Unlike  $[\text{Fe}(\text{CN})_6]^{-3/-4}$  which is an inner sphere redox system that is sensitive to changes on the surface. From the CV scans in Fig. S7† we see that the peak separation remains close to 65 mV at a scan rate of  $10 \text{ mV s}^{-1}$  similar suggesting the quasi-reversible electrode kinetics. Our MLG demonstrates good electrochemical response due to many edge sites available on the surface of the electrodes.

This suggests that 25 printed passes of MLG has slightly faster electron transfer kinetics than 30 printed passes, likely due to a higher surface roughness of 25 printed passes extracted from the AFM images in the (ESI Fig. S4a and b†). This electron transfer rate ( $k = 1.125 \times 10^{-2} \text{ cm s}^{-1}$ ) is close to or better than that of graphite, graphene oxide, mechanically exfoliated graphene, and screen-printed carbon or graphene or CNT electrode which range from  $10^{-4} \text{ cm s}^{-1}$  to  $10^{-2} \text{ cm s}^{-1}$  for the ferric/ferrocyanide redox reaction.<sup>5,68–73</sup>

Cross-sectional TEM was used to image the porosity of the printed MLG electrodes (25 printed passes and 30 printed passes) on untreated Kapton substrates and better understand the structure of IJP deposited MLG. From the cross-sectional TEM images seen in Fig. 3e and f, it is evident that the 25 printed passes (Fig. 3e) IJP MLG sheets are less dense than the 30 printed passes of MLG (Fig. 3f). It is seen that 25 printed passes of graphene exhibit a higher disordered stacking than

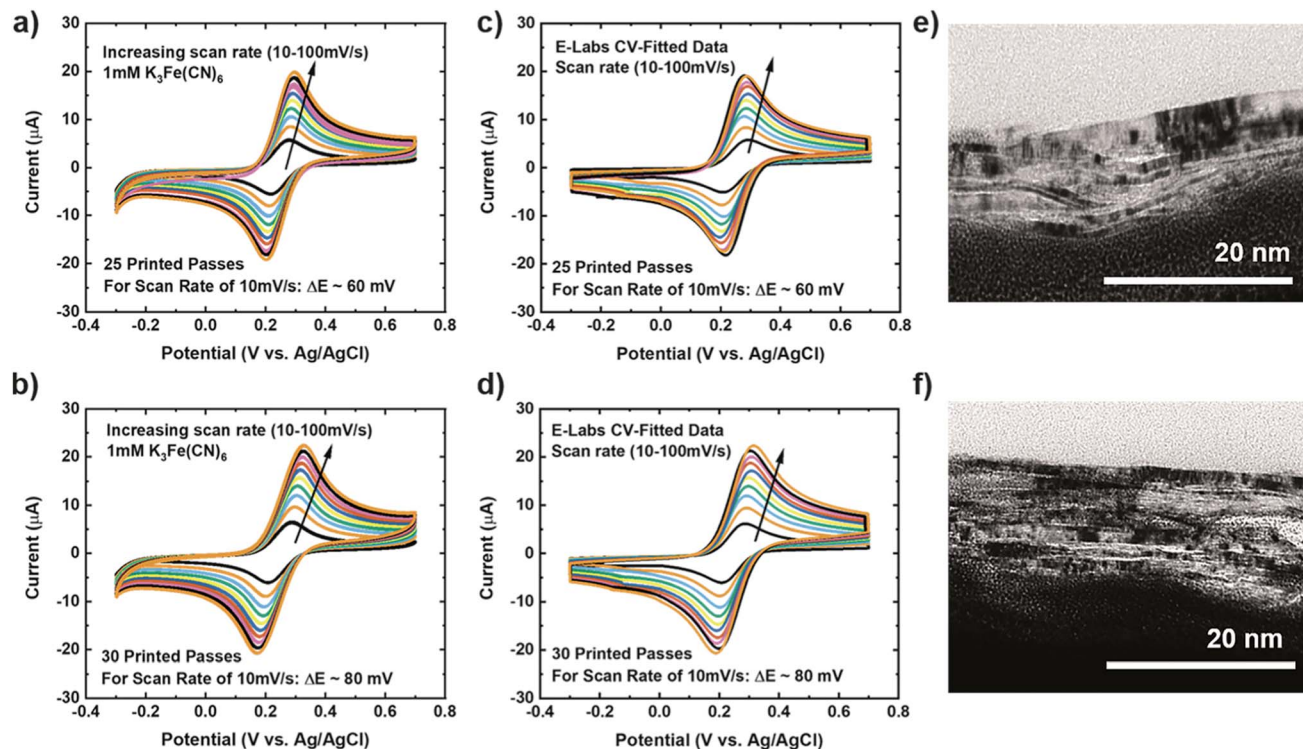


Fig. 3 Graphene electrode cyclic voltammetry (CV) scan rate data for 1 mM  $\text{K}_3[\text{Fe}(\text{CN})_6]$  in 1 M KCl as the supporting electrolyte with increasing scan rate  $10\text{--}100\text{ mV s}^{-1}$  for (a) 25 printed passes of graphene and (b) 30 printed passes of graphene. E-labs CV-Sim fitted data for (c) 25 printed passes of graphene and (d) 30 printed passes of graphene. Cross-section TEM images of (e) 25 printed passes graphene and (f) 30 printed passes graphene.

the 30 printed pass case. A higher porosity between the stacked multilayers of graphene is observed in 25 printed passes of graphene than in the 30 printed passes. Surface porosity and packing morphology play a significant role in electrochemical performance, as established by Punckt *et al.*<sup>54</sup> Moreover, using our CV data with varying scan rate ( $\nu$ ), we can obtain the values of max current at  $\nu = 10\text{ mV s}^{-1}$  and  $\nu = 100\text{ mV s}^{-1}$  and calculate the porosity ( $P$ ) according to the equation,

$$P = k \times \frac{I_{\max} \left( \nu = 100 \frac{\text{mV}}{\text{s}} \right)}{I_{\max} \left( \nu = 10 \frac{\text{mV}}{\text{s}} \right)} k = \frac{\sqrt{10}}{\sqrt{100}}. \text{ For an ideal planar}$$

electrode,  $P = 1$  since then  $I_{\max} \sim (\nu)^{1/2}$ . We find that for our IJP MLG electrodes,  $P = 1.17$  for 25 printed passes and  $P = 1.06$  for 30 printed passes. This suggests that 25 printed passes show more porosity than 30 printed passes, further supporting our hypothesis that enhanced electrocatalytic behavior is influenced by packing morphology in our printed graphene electrodes.<sup>74</sup>

### Stability of MLG electrodes

It is important that these electrodes are inherently stable in the electrolyte and can be reproduced *via* the printing methods, to enable biosensor design, optimization, and efficient collection of data in real-time.<sup>75</sup> To investigate the stability of our MLG electrodes, CV measurements were carried out for a series of redox mediator couple dilutions between 1 mM to 5 mM

$\text{K}_3[\text{Fe}(\text{CN})_6]$  with 1 M KCl as the supporting electrolyte. Scans were performed at room temperature for 25 (Fig. 4a–c) and 30 (Fig. 4d and e) printed passes with varying scan rate of  $10\text{ mV s}^{-1}$  (Fig. 4a and d),  $50\text{ mV s}^{-1}$  (Fig. 4b and e), and  $100\text{ mV s}^{-1}$  (Fig. 4c and f). We observed that 25 printed passes showed a lower peak to peak separation for all three scan rates in each dilution as compared to 30 printed passes of MLG. Additionally, time-dependent effects were investigated by recording the CV curves at  $100\text{ mV s}^{-1}$  every 5 min in the same electrolyte for up to 2 hours, similar to the study of Patel *et al.*<sup>76</sup> These studies were carried out with 25 printed passes of MLG electrode in 1 mM ferric/ferrocyanide solution. Fig. 5a shows great stability with negligible change in the peak to peak potential separation, even after 16 hours in the electrolyte.

The reproducibility of the MLG electrodes is demonstrated *via* the CV scans seen in Fig. 5b, showing a triplicate study with equivalent ink, printing, and other experimental and measurement conditions. The CV scans show consistent and reproducible results for all three electrodes, with a mean peak to peak potential separations of  $64 \pm 1\text{ mV}$ . To understand the sensing range of the MLG electrode, CV measurements were performed with varying concentrations of ferric/ferrocyanide solution from 10 mM to 1  $\mu\text{M}$ . The bare MLG flexible electrode exhibits a broad sensing range from 10 mM down to 0.01 mM (Fig. 5c), which is comparable to the sensing range for non-flexible, modified graphene, reduced graphene oxide, or CNT electrodes shown in Table 1 and reported in the literature.<sup>8,9,11,77–84</sup>

Table 1 Comparison of the  $\Psi$  (kinetic parameter) of different material- and fabrication-based electrodes printed three-electrode sensor system (shaded in pink)

Electrode material	Fabrication technique	Substrate	Electrolyte	Redox couple	Scan rate ( $\text{mV s}^{-1}$ )	$\Psi$ , $E_p\Delta$ (59–200) mV	Ref.
Multi-layered graphene	Inkjet-printed	Kapton (polyimide)	1 M KCL	$\text{Fe}(\text{CN})_6^{3-/4-}$ (1 mM)	10	25.14	This work
Reduced graphene oxide	Laser-scribed	PET	1 M KCl	$\text{Fe}(\text{CN})_6^{3-/4-}$ (1 mM)	10	25.14	66
COOH-terminated graphene nanoflakes	Coated	Boron doped dimond	0.1 M $\text{KH}_2\text{PO}_4$	$\text{Fe}(\text{CN})_6^{3-/4-}$ (0.5 mM)	50	2.54	96
Reduced graphene oxide	Coated	Glassy carbon electrode	0.1 M KCl	$\text{Fe}(\text{CN})_6^{3-/4-}$ (5 mM)	30	2.54	97
Gold	Inkjet-printed	Kapton (polyimide)	100 mM KCl	$\text{Fe}(\text{CN})_6^{3-/4-}$ (1 mM)	100	1.71	98
Reduced graphene oxide	Screen-printed	Poly(vinyl chloride)	0.1 M KCl	$\text{Fe}(\text{CN})_6^{3-/4-}$ (5 mM)	50	1.61	15
Edge-oxidized graphene nanosheet	Inkjet-printed	Kapton (polyimide)	0.1 M $\text{KNO}_3$	$\text{Fe}(\text{CN})_6^{3-/4-}$ (2 mM)	50	1.61	99
Carbon nanotubes	Inkjet-printed	Kapton (polyimide)	0.1 M KCl	$\text{FeMeOH}$ (4 mM)	100	1.29	100
CNT	Inkjet-printed	PET	0.1 M $\text{KNO}_3$	$\text{FeMeOH}$ (2 mM)	25	1.28	101
Gold	Inkjet-printed	Paper	3 M KCl	$\text{Fe}(\text{CN})_6^{3-/4-}$ (3 mM)	20	0.75	102
Graphite	Screen printed	Ultra-flexible polyester materials	0.1 M KCl	$\text{Fe}(\text{CN})_6^{3-/4-}$ (1 mM)	10	0.60	103
Reduced graphene oxide	Inkjet-printed	FTO (fluorine-doped tin oxide) TEC-15	0.1 M TBAPF <sub>6</sub>	$\text{Co}(\text{bpy})_3^{2+}/\text{B}(\text{CN})_{4,5}$ (1 mM)	50	0.60	104
Graphene-PEDOT:PSS	Inkjet-printed	Screen-printed carbon electrode/ non-conformal	0.1 M KCl	$\text{Fe}(\text{CN})_6^{3-/4-}$ (0.36 $\mu\text{M}$ )	100	0.33	105
Functionalized graphene nanoribbons	Screen-printed	Polyethylene glycol terephthalate	0.2 M PBS	$\text{Fe}(\text{CN})_6^{3-/4-}$ (1 mM)	100	0.20	106
Graphene-polyaniline	Inkjet-printed	PET	0.1 M KCl	$\text{Fe}(\text{CN})_6^{3-/4-}$ (1 mM)	100	0.20	107
Graphite pencil	Drawn	Paper	0.1 M KCl	$\text{Fe}(\text{CN})_6^{3-/4-}$ (1 mM)	100	0.10	108
Graphene with platinum (laser annealed)	Inkjet mask-less lithography	Screen-printed carbon electrode/ non-conformal substrate	1 M PBS	$\text{Fe}(\text{CN})_6^{3-/4-}$ (5 mM)	10	$E_p\Delta > 300$ (mV vs. Ag/AgCl)	109
Reduced graphene/poly(lactic acid)	3D-printed	—	0.1 M KCl	$\text{Fe}(\text{CN})_6^{3-/4-}$ (1 mM)	100	$\sim 500$ (mV vs. Ag/AgCl)	110
Reduced graphene oxide	Inkjet-printed/ laser sintered	Cellulose-based paper	1 M KCl	$\text{Fe}(\text{CN})_6^{3-/4-}$	10	$\sim 0.7$ (V vs. Ag/AgCl)	111
Reduced graphene oxide	Inkjet-printed	Poly(ethylene 2,6-naphthalate) PEN	0.1 M KCl	$\text{Fe}(\text{CN})_6^{3-/4-}$ (1 mM)	10	$\sim 400$ to 5000 (mV vs. Ag/AgCl)	112
PEDOT:PSS	Inkjet-printed	Paper	0.1 M PBS	$\text{FeMeOH}$ (—)	20	$\sim 50$ (mV vs. Ag/AgCl)	113
Multi-walled nanotubes with silver	Inkjet-printed	Paper	0.5 M KCl	$\text{Fe}(\text{CN})_6^{3-/4-}$ (3 mM)	20	$\sim 55$ (mV vs. Ag/AgCl)	114
Graphite	Screen-printed	Chromatography paper	0.1 M $\text{H}_2\text{SO}_4$	$\text{Fe}(\text{CN})_6^{3-/4-}$ (1 mM)	100	$\sim 56$ (mV vs. Ag/AgCl) after 30 scans	115

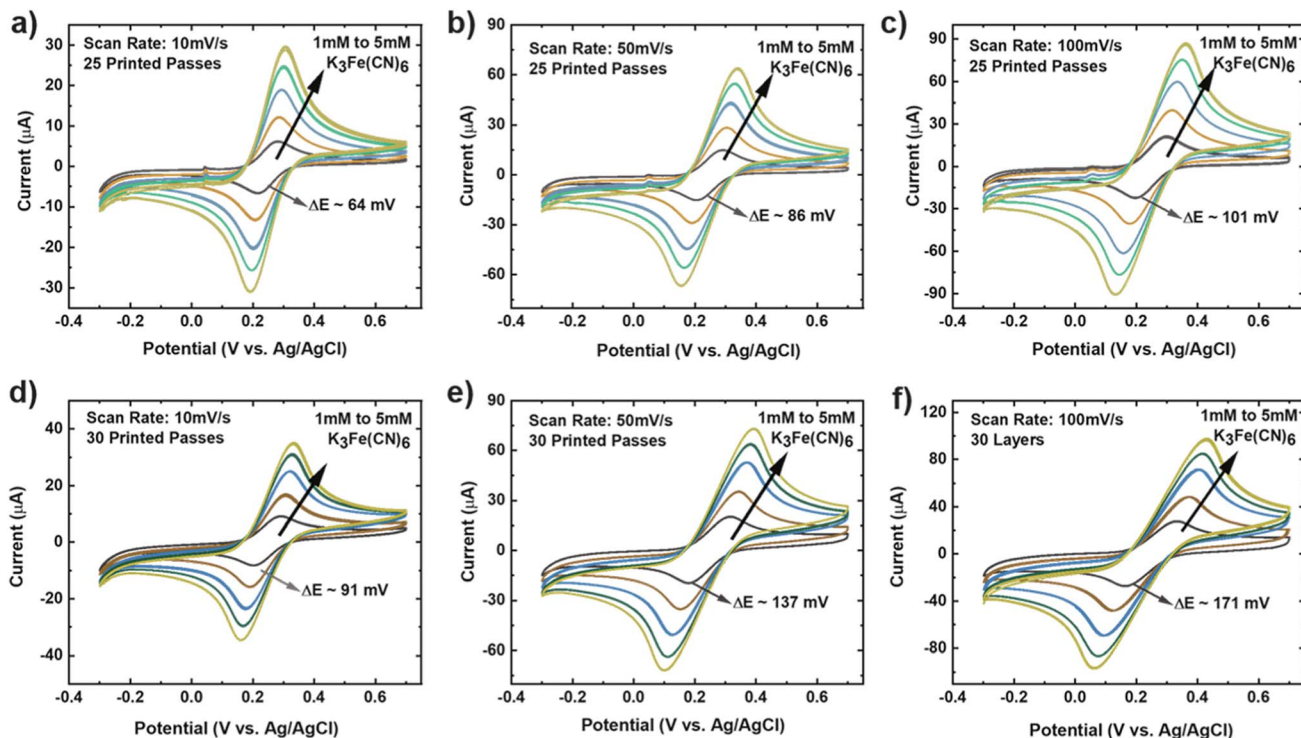


Fig. 4 Graphene electrode CV dilution data for 1–5 mM  $K_3[Fe(CN)_6]$  in 1 M KCl as the supporting electrolyte for 25 printed passes at scan rates of (a)  $10\text{ mV s}^{-1}$ , (b)  $50\text{ mV s}^{-1}$ , and (c)  $100\text{ mV s}^{-1}$  (d–f) are corresponding data for 30 printed passes, with scan rates of (d)  $10\text{ mV s}^{-1}$ , (e)  $50\text{ mV s}^{-1}$ , and (f)  $100\text{ mV s}^{-1}$ .

### MLG electrode pH sensitivity

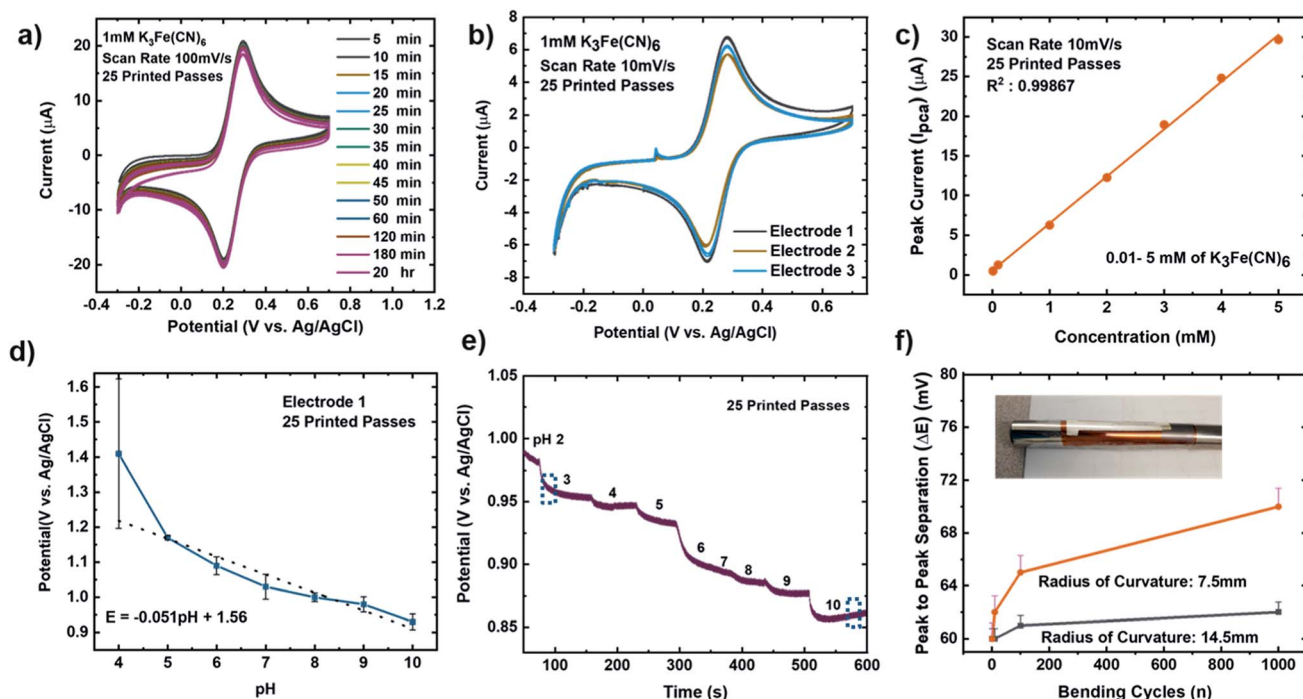
The pH of a system is critical to chemical/biological/biochemical processes.<sup>78,80,85,86</sup> It is also an essential factor for accurately determining the stability and sensitivity of a biosensor as biochemical reactions that take place on the sensor are highly dependent on pH. Potentiometric pH sensors can extract information about pH values by measuring the open circuit potentials.<sup>11</sup> For this study, pH sensitivity experiments were conducted on bare MLG electrodes to observe the potentiometric response of the electrode as the pH was varied in the range of 1–10. First, static pH data were acquired using commercially available potassium phosphate monobasic with sodium hydroxide pH buffer solutions (pH 1, 4, 7, and 10) on bare 25 printed passes of MLG printed electrode. Solutions with pH values of 2, 3, 5, 6, 8, and 9 were then formulated by mixing the high pH standard solution with low pH solutions. The pH of the buffer solutions was confirmed using a glass-electrode benchtop pH meter (Mettler-Toledo, Columbus, OH) in a stirred solution. Chronopotentiometry measurements were performed with the printed MLG electrodes while varying the pH buffer solution.

The open circuit potential values were captured for different pH solutions for a 120 second duration. Since the MLG electrode electrochemical process is reversible, the Nernst equation for pH calculation from open circuit potentials is as follows,  $E = E_0 - \frac{RT}{nF} \ln(Q)$ , where for an ideal electrode,  $E = E' + 0.0591\text{pH}$ , with  $E$  the measured open circuit potential and  $E'$  the standard potential,  $R$  the gas constant ( $8.314\text{ J K}^{-1}\text{ mol}^{-1}$ ),  $T$  the

absolute temperature (K),  $n$  the signed ionic charge and  $F$  is the Faraday constant ( $96487.3415\text{ C mol}^{-1}$ ).<sup>87</sup> The equation of the fitted line (pH values 4–10) is as follows:  $E = 1.56 - 0.051\text{pH}$ , where the slope of  $51\text{ mV pH}^{-1}$  is close to that of an ideal electrode (*i.e.*,  $59\text{ mV pH}^{-1}$ ) seen in Fig. 5d and the raw data is seen in the ESI Fig. S5.† To examine the reproducibility of the MLG electrodes, the pH experiment was conducted on three different 25 print pass electrodes fabricated with identical print conditions as described above. ESI Fig. S6† displays the potential (MLG vs. Ag/AgCl) versus pH data for all three electrodes. Additionally, from ESI Fig. S6,† it can be determined that the MLG electrodes are capable of providing a consistent response with potential (V vs. Ag/AgCl). Furthermore, we employed a response time experiment of our electrode with changing pH values. We started by measuring pH 2 buffer solution and added aliquots ( $100\text{ }\mu\text{L}$  to  $1\text{ mL}$ ) of pH 10 buffer solution to change the pH of the solution tested from 2 to 10 and recorded the change in the open circuit potential. The solution tested was stirred between measurements with a magnetic stirrer placed under the cell. Fig. 5e displays the change in potential with the pH of the solution. This suggests that the IJP of bare MLG on a Kapton substrate showed a significant response to the change in pH in the solution.

Using the separate solutions methods (SSM) with different interfering ions of  $K^+$ ,  $Na^+$ , and  $NH_4^+$  (pH 6) to estimate the potentiometric selectivity coefficients  $K_{ij}^{\text{pot}}$  at different ion concentration ( $10^{-2}\text{ M}$ ), where  $K$  is the selectivity coefficient,  $I$  is the primary ion, and  $J$  is the interfering ion.<sup>88,89</sup> Equation is as follows:





**Fig. 5** (a) Time-dependent CV scans for 25 IJP printed passes, 1 mM  $K_3[Fe(CN)_6]$  in 1 M KCl as the supporting electrolyte at a scan rate of  $100\text{ mV s}^{-1}$ . (b) CV repeatability data for 25 IJP printed passes (3 graphene electrodes), 1 mM dilution  $K_3[Fe(CN)_6]$  in 1 M KCl as the supporting electrolyte at a scan rate of  $10\text{ mV s}^{-1}$ . (c) CV dilution data showing peak current ( $I_{pca}$ ) vs. concentration from 5 mM to 0.01 mM of  $K_3[Fe(CN)_6]$  in 1 M KCl as the supporting electrolyte for 25 IJP printed passes, one graphene electrode at scan rate from  $10\text{ mV s}^{-1}$ . (d) Static linear pH data vs. potential (potassium phosphate monobasic with sodium hydroxide commercial pH buffer solutions: 4–10 pH) using the 25 passes graphene printed electrode. Error bars represents the interelectrode standard variation in slope compared to the theoretical values based on the Nernst equation.<sup>93–95</sup> (e) Time vs. potential data with changing pH from 2 to 10 for a single 25 printed passes graphene electrode. The error bar represents the standard deviation of potential across three independent samples. (f) Bending cycles (1, 10, 100, and 1000) conducted on the electrodes with radius of curvature either 7.5 mm (orange) or 14.5 mm (black) vs. peak to peak separation potential.

**Table 2** SSM for calculating selectivity coefficients of MLG sensors ( $H^+$  ions against interfering ions)

Ions ( $J$ )	$\log K_{IJ}^{pot}$	$K_{IJ}^{pot}$
$Na^+$	-5.34	$4.52 \times 10^{-6}$
$K^+$	-4.48	$3.32 \times 10^{-5}$
$NH_4^+$	-6.87	$1.34 \times 10^{-7}$

$$K_{IJ}^{pot} = a_I / (a_J)^{z_I/z_J}$$

where  $a$  is the activity of the ion and  $z$  is the sign or the magnitude for the corresponding charge of the ion. For good selectivity of  $H^+$  ions over the cations, the  $K$  value is  $<1$ . Table 2 shows the selectivity coefficient for MLG sensor. The result does show that MLG has good ion selectivity compared to literature.<sup>90</sup>

To investigate the flexibility of the IJP MLG electrodes, bending cycle testing ( $n$  = number of bending cycles) was performed on five IJP printed MLG electrodes fabricated with identical printing conditions (25 printed passes), and having similar resistance measurements. Bending cycles were performed with 7.5 mm and 14.5 mm radius of curvature tubes, and tests were performed from  $n = 1$  to 1000, with CV

measurements performed at  $n = 10$  intervals, as shown in the images in Fig. 5f. Bending is expected to increase the resistance of the electrodes, which should increase the peak to peak separation. Fig. 5f, shows that the IJP MLG electrodes show a robust performance over this range of bend cycle testing. Additionally, bending cycles resulted in a  $\sim 2\%$  increase peak-to-peak separation with a 14.5 mm radius of curvature and an  $\sim 10\%$  increase in peak-to-peak separation with a 7.5 mm radius of curvature at  $n = 1000$ . A summary of flexible, graphene-based sensors is listed in Table 1. Electrodes shaded in pink shows functionality of a fully printed three electrode sensor system, however only our study show functionality or bare graphene electrodes with inkjet printed Pt counter electrodes. To the best of our knowledge, this is the first study amongst all other flexible, graphene-based sensors, to report on and demonstrate the stability of these types of sensors subjected to cyclic bending tests.

### Fully printed three-electrode devices

The studies mentioned above provide insights into the electrochemical performance of individual working IJP MLG electrodes using conventional external reference and counter electrodes, as seen in Fig. 2e. Here we compare the electrochemical performance of fully flexible IJP three-electrode (working, counter, and reference) sensor systems to the

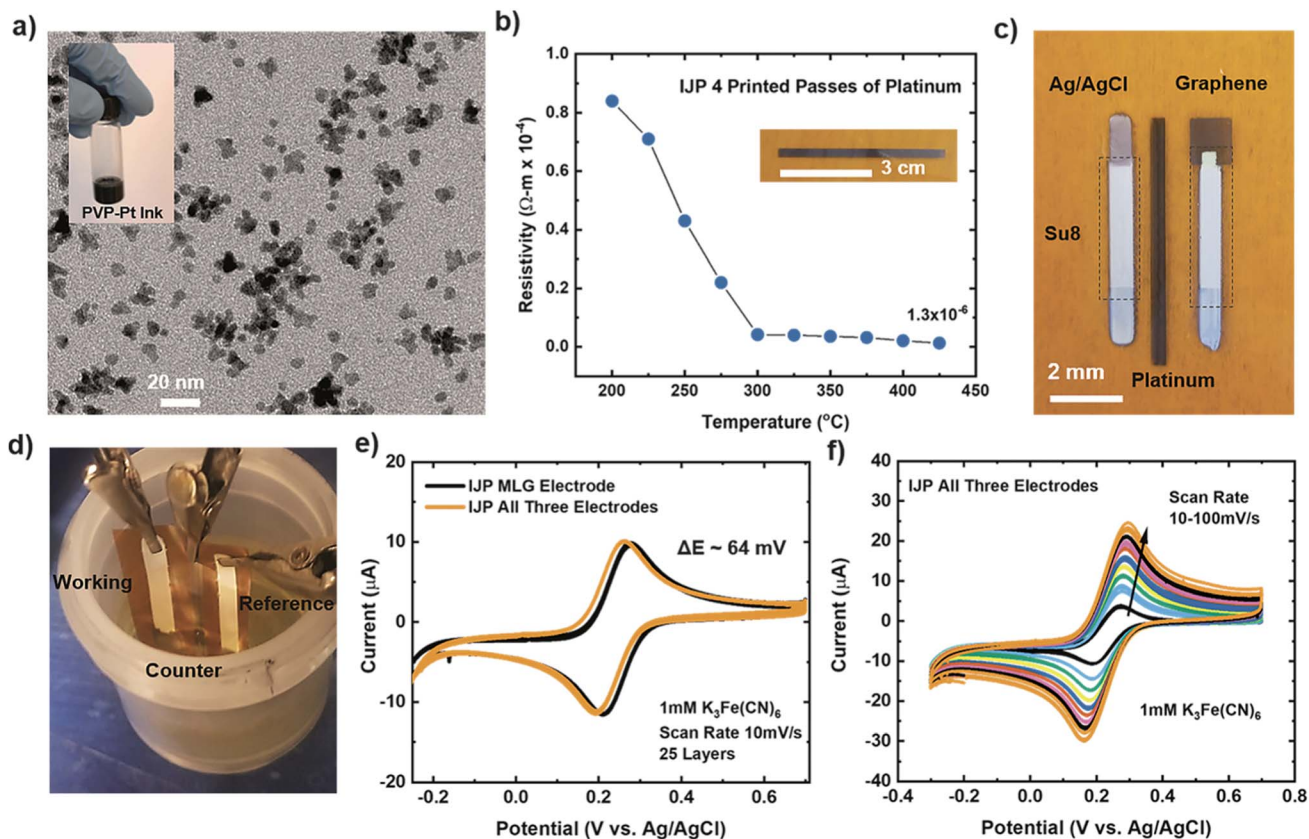


Fig. 6 (a) TEM image of the platinum nanoparticles (PtNP) with an optical image of the platinum ink (inset) (b) resistivity vs. temperature for 4 printed passes of platinum lines on Kapton (inset). (c) Optical image of IJP all three electrodes (Ag/AgCl as the reference electrode, platinum as the counter electrode, and MLG as the working electrode). (d) Optical picture of the electrochemical setup for the fully printed three electrode sensors. (e) Cyclic voltammetry (CV) data of IJP MLG (black) compared to all three electrodes (orange) in 1 mM  $\text{K}_3\text{Fe}(\text{CN})_6$  and 1 M KCl as the supporting electrolyte with scan rate  $10 \text{ mV s}^{-1}$  and (f) CV scans of all three printed electrodes with increasing scan rate from  $10$ – $100 \text{ mV s}^{-1}$ .

individual IJP MLG electrode. Fully IJP flexible electrodes could enable large scale, roll-to-roll level production of such sensors. To fabricate the IJP three-electrode sensor system, custom made polyvinyl pyrrolidone capped Pt nanoparticle (PVP–PtNP) ink (seen inset of Fig. 6a) was prepared to print a counter electrode. The TEM image seen in Fig. 6a shows the PtNP ranging from 5–8 nm. To measure the electrical properties of the printed platinum lines, a 4-point probe (Keithley 4200 SCS, Textronix) measurements were conducted. Fig. 6b shows the calculated resistivity as a function of annealing temperature for 4 print passes. Based on the graph, the lowest resistivity was calculated to be  $1.3 \times 10^{-6} \Omega\text{m}$  for 4 print passes at an annealing temperature of  $\sim 425 \text{ }^\circ\text{C}$  which is  $\sim 10\times$  greater than bulk platinum ( $1.06 \times 10^{-7} \Omega\text{m}$ ).<sup>91</sup> Silver (NovaCentrix), Su8 (PriElex Microchem), NaClO, and a cocktail mixture of PVP (Butvar B-98) ink solutions were used to fabricate the Ag/AgCl/( $\text{Na}^+$ ) reference electrode (see Methods section for further details). Fig. 6c displays the fully IJP three-electrode sensor system using MLG as the working electrode, Pt counter electrode and Ag/AgCl/( $\text{Na}^+$ ) reference electrode. Su8 ink was used as a passivation layer for the MLG electrode, and Ag/AgCl/( $\text{Na}^+$ ) electrode.

The experimental setup shown in Fig. 6d was used to study and compare the electrochemical response of a fully printed three-electrode device to our individual IJP MLG electrode. CV

measurements were carried out with a scan rate of  $10 \text{ mV s}^{-1}$  in  $1 \text{ mM K}_3[\text{Fe}(\text{CN})_6]$  with 1 M KCl as the supporting electrolyte, for the individual IJP MLG electrode (25 printed passes) and the fully printed three-electrode device. Fig. 6e demonstrates that the response of the fully printed three-electrode devices is comparable to the individual IJP MLG electrodes, and that the three-electrode device exhibits excellent reversibility with a peak to peak separation of  $\sim 64 \text{ mV}$ . Furthermore, CV measurements were carried out with an increasing scan rate from  $10$ – $100 \text{ mV s}^{-1}$  in  $1 \text{ mM K}_3[\text{Fe}(\text{CN})_6]$ , as seen in Fig. 6f. Again, CV simulation was used to extract the rate constant ( $k$ ) from the experimental CV data seen in Fig. 6f. The electron rate transfer constant extracted for the fully IJP three-electrode sensors was determined to be  $k = 1.22 \times 10^{-2} \text{ cm s}^{-1}$  for a scan rate of  $10 \text{ mV s}^{-1}$ . These CV measurements suggest that the fully printed three-electrode device shows fast electron transfer with this redox system, similar to the results presented for the individual IJP MLG electrodes.

## Conclusion

In this study, the electrochemical behavior of fully inkjet printed multilayer graphene electrodes on flexible Kapton substrates was investigated. Cyclic voltammetry was used to

analyze the electrochemical reversibility of a fully inkjet-printed MLG electrode and a fully inkjet-printed three-electrode device using the ferric/ferrocyanide  $[\text{Fe}(\text{CN})_6]^{-3/-4}$  redox couple as the analyte. It was confirmed that electrodes optimized at 25 printed passes (with adequate inert edge defects and surface porosity) showed quasi-reversibility with a low peak to peak potential separation of 60 mV and fast electron-transfer kinetics ( $k = 1.125 \times 10^{-2} \text{ cm s}^{-1}$ ). Moreover, it was verified that the printed MLG electrode was responsive to varying solution pH and displayed good electrochemical stability even after 1000 bending cycles (7.5 mm radius of curvature) with less than 10% change in peak to peak separation. Cross-sectional TEM images also revealed that the morphology of the printed graphene electrodes enhanced the electrochemical response and behavior of the printed electrodes. These studies indicate that fully IJP three-electrode sensors are a promising approach to fabricating flexible electrodes with an excellent electrochemical response comparable to those reported in the literature. These electrodes can be produced quickly, easily, and repeatedly, thus showing excellent potential for scalable manufacturing and flexible biosensing applications. The approach reported here enables a deeper understanding of how the combination of ink rheology and additive electronics manufacturing can enable the scalable manufacturing of flexible biosensors for space, military, and commercial applications.

## Methods section

### Preparation and characterization of multilayered graphene ink

Similar to previous work, graphene flakes were obtained by solvent assisted exfoliation of 50 mg mL<sup>-1</sup> graphite powder in a suspension of 2% ethyl cellulose (EC) in ethanol using a Qsonica (Q125) (Newtown, CT) probe tip sonicator for 90 minutes.<sup>20,53,57</sup> To remove the larger graphite flakes, the dispersion was centrifuged (Heraeus™ Megafuge™ 8 Centrifuge TX-150 rotor) at 3402 RCF for 60 min and the supernatant was collected immediately. In a 1 : 2 volume ratio, the collected supernatant and 0.04 g mL<sup>-1</sup> aqueous solution of NaCl (Sigma-Aldrich, >99.5%) were centrifuged for 30 min at 3402 RCF to facilitate the flocculation of graphene flakes. The resulting graphene/EC dispersion was dried overnight on a PTFE (Teflon) plate. To tailor the concentration and viscosity of ink to be compatible with the Dimatix IJP (Fujifilm, Santa Clara, CA), the dried graphene/EC paper was then dispersed by sonication for 30 min in a mixture of 85% cyclohexanone and 15% terpeneol solution, followed by centrifugation at 3402 RCF for 15 min. The resulting ink concentration was 3.5 mg mL<sup>-1</sup> with a viscosity of 8 cP (Wells-Brookfield Cone/Plate Middleboro, MA). The ink concentration was quantified by UV-vis absorption spectroscopy and Beer-Lamberts law at  $\lambda = 600 \text{ nm}$ .

### Fabrication of the MLG electrodes

The MLG was printed using a Dimatix inkjet printer. The tool platen was heated to 60 °C to ensure rapid drying of the ink once deposited on the Kapton HN (Dupont, Wilmington, DE)

substrate. A 10 pL cartridge was used to print the MLG. The waveform, jetting voltage, and drop spacing were adjusted to achieve uniform droplets in volume and velocity of the MLG ink. Moreover, 4 nozzles were used to print, and the cartridge was at room temperature. After printing, any remaining ethyl cellulose and solvent was driven out of the lines with a 250 °C bake for 30 min followed by a ramp to a 350 °C bake for 45 min to maximize conductivity of the printed features. Next, NovaCentrix, Metalon (JS-B40G, Austin, TX) silver ink was used to print the contact pad connecting the printed MLG before sintering at 250 °C for 15 min. Lastly, PriElex Microchem (Westborough, MA) SU-8 ink was printed on top of the silver contact pad connecting the printed MLG. SU-8 acted as a passivation layer to isolate the silver while conducting electrochemical experiments.

### Fabrication of the platinum ink and platinum electrodes

98% sodium tetrahydroborate (Alfa Aesar), 99.999% hexachloroplatinic(IV) acid hydrate (40% platinum metals basis, BeanTown Chemical), 10 kDa polyvinylpyrrolidone (Alfa Aesar), 95–100.5% sodium hydroxide pellets (Macron), ethylene glycol (VWR), ≥99.5% carboxymethyl cellulose sodium salt, and 40 kDa MWCO Slide-A-Lyzer Dialysis flasks (Thermo Scientific) were purchased from commercial sources and used without further purification. A 0.58 M hexachloroplatinic acid ( $\text{H}_2\text{PtCl}_6$ ) stock solution was prepared with the addition of 5 g of  $\text{H}_2\text{PtCl}_6$  to 20 mL of nanopure (18 M $\Omega$ ) water. Additionally, a 2.2 M stock solution of sodium tetrahydroborate ( $\text{NaBH}_4$ ) was prepared with the addition of 0.500 g of  $\text{NaBH}_4$  to 6 mL of nanopure (18 M $\Omega$ ) water buffered to a pH of 12 with NaOH. Both solutions were used without further purification or dilution.

A platinum nanoparticle ink containing ~20 wt% of 5–8 nm of polyvinyl pyrrolidone capped Pt nanoparticles (PVP-PtNP) suspended in a water/ethylene glycol co-solvent mixture was prepared to be compatible with ink jet printing. The synthesis of PVP stabilized Pt nanoparticles was accomplished through wet chemical methods where 10 mL of stock  $\text{H}_2\text{PtCl}_6$  solution was added to 1.5 L of nano-pure water containing 6 g of dissolved PVP. The  $\text{H}_2\text{PtCl}_6/\text{H}_2\text{O}/\text{PVP}$  solution was allowed to stir for two hours and was followed by the drop-wise addition of 6 mL of stock  $\text{NaBH}_4$  to form PVP-PtNP. The resulting suspension was allowed to stir vigorously for 24 hours, and this process was repeated until a total of 20 g of  $\text{H}_2\text{PtCl}_6$  had been reduced or 6 L of PVP-PtNP suspension had been synthesized to form ~8 g of PVP capped platinum nanoparticles.

The removal of excess capping agent and reaction by-products was performed through dialysis while utilizing 40 kDa MWCO Slide-A-Lyzer Dialysis flasks. As the dialysis process can be extremely time consuming, the PVP capped Pt nanoparticle suspension was dialyzed against a very high concentration solution of carboxymethyl cellulose in order to accelerate this process. A total of 6 L of PVP-PtNP was concentrated to 50 mL, which was followed by rotary evaporation to further concentrate the suspension to 20 mL. The viscosity of the PVP-PtNP suspension was tuned through the addition of ethylene glycol solution to ensure the ink rheology was compatible with inkjet printing.

The PVP–Pt was printed using a Dimatix inkjet printer. The tool platen was heated to 30 °C to ensure rapid drying of the ink once deposited on the Kapton HN (Dupont, Wilmington, DE) substrate. A 10 pL cartridge was used to print the PVP–Pt. The waveform, jetting voltage, and drop spacing were adjusted to achieve uniform droplets in volume and velocity of the PVP–Pt ink. Moreover, 2 nozzles were used to print, and the cartridge was at room temperature. After printing, any remaining surfactant and solvent was driven out of the lines with a 150 °C bake for 15 min followed by a ramp to a 400 °C bake for 45 min to maximize conductivity of the printed features.

### Fabrication of Ag/AgCl electrodes

A good method to fabricate fully inkjet printed Ag/AgCl reference electrodes has been described by Moya *et al.*<sup>92</sup> NovaCentrix, Metalon (JS-B40G, Austin, TX) silver ink was used to print the silver layers on Kapton substrate and then sintered at 250 °C for 30 min. Then, PriElex Microchem (Westborough, MA) SU-8 ink was printed on top of the silver layers for passivation and then sintered 250 °C for 50 min. For chlorination, diluted NaClO (5 v/v% purchased from Sigma-Aldrich) was printed (2 passes) on the exposed silver and then washed with deionized water. For the formulation of a protecting membrane, a cocktail mixture of PVB (Butvar B-98) (10 w%) in methanol (40%), xylene (30%), diacetone alcohol (15%), and 1-butanol (15%), all purchased from Sigma-Aldrich, was printed (5 passes) on top of the chlorinated area. Lastly, the electrode was left in a fume hood to dry overnight.

### Thermogravimetric analysis

To determine the annealing temperature of printed graphene features, thermogravimetric analysis (TGA) was employed. A Netzsch STA 449 F1 Jupiter (Burlington, MA) TGA instrument was used to measure the weight percent loss as a function of temperature (25 °C to 1000 °C) at a heating rate of 5 °C min<sup>-1</sup> in air. TGA analysis revealed the decomposition peak of ethyl cellulose is around 250 °C, while the other volatile solvent components (cyclohexanone and terpineol) are driven off at 390 °C (seen in Fig. 1e).

### SEM, TEM and Raman imaging

A FEI Teneo (Hillsboro, OR), field emission SEM was used to image the printed films. Using the Dimatix inkjet printer, MLG and silver ink were used to print a transmission line measurement (TLM) structure with varying number of print passes from 15–30, with increments of 5 passes, on glass substrate (seen in Fig. 1a). The SEM image for 25-layer pass line of IJP multilayered graphene is shown in Fig. 1b to demonstrate its uniformity. TEM images were obtained using a JOEL JEM 2100 (Peabody, MA) system, with the particles characterized using ImageJ software. 25 and 30 printed passes of MLG were printed on untreated Kapton and the TEM samples were prepared by a FIB (focused ion beam) tool at the Center for Advanced Energy Studies. Lastly, Raman spectra were obtained using a Horiba LabRAM HR Evolution Raman microscope (Irvine, CA) with a 532 nm excitation wavelength. The spectra (1000–3000 cm<sup>-1</sup>)

were collected at a relative laser power of 25% with a 100× objective and 30 s exposure time.

### Electrochemical setup

The electrochemical experiments were conducted using a customized 3D printed cell (Fig. 2e), and potentiostat (Bio-Logic VMP-300 instrument, Knoxville, TN) with EC-Lab for the software. The 3D printed cell allowed for the printed MLG (working electrode) to be placed on the bottom of the cell, with only 0.07 cm<sup>2</sup> of the MLG exposed to the solution in the cell. Ag/AgCl was used as a reference electrode and with a platinum mesh as the working electrode. Room temperature cyclic voltammetry (CV) measurements were carried out in a fume hood with an increasing scan rate from 10–100 mV s<sup>-1</sup>, with dilutions ranging from 1 mM to 5 mM K<sub>3</sub>[Fe(CN)<sub>6</sub>] and 1 M KCl (both purchased from Sigma Aldrich) as the supporting electrolyte. The EC-lab software was used to extract the fitted CV data seen in Fig. 3c and d.

### Authors contributions

T. P. designed and printed the test structures, and together with C. C., P. B., and J. C. characterized the devices. J. K. and D. E. conceived of the experimental design, and together with H. S. and C. X. supervised the experiments. K. F. synthesized the platinum (PVP–Pt) inks. All authors contributed to the development of the manuscript.

### Conflicts of interest

The authors declare no competing interests.

### Acknowledgements

This work was supported by NASA EPSCoR under Award #ID-80NSSC17M0029. The authors also thank Dr Yaqiao Wu at the Center for Advanced Energy Studies for his support on FIB and TEM experiments. D. E. also acknowledges career development support by Institutional Development Awards (IDeA) from the National Institute of General Medical Sciences of the National Institutes of Health under Grants #P20GM103408 and #P20GM109095.

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