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Adsorption and sensing mechanisms of Ni-doped PtTe₂ monolayer upon NO₂ and O₃ in air-insulated switchgears†

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Under partial discharge, air would be converted into O₃ and NO₂ in air-insulated switchgears, therefore, the detection of such two gases can be used to evaluate the operation status of such electrical equipment. In this study, first-principles simulations are implemented to investigate the Ni-doping behavior on the pristine PtTe₂ monolayer, and the adsorption and sensing performances of the Ni-doped PtTe₂ (Ni–PtTe₂) monolayer upon O₃ and NO₂ in air-insulated switchgears. The formation energy (E_{form}) of Ni-doping on the PtTe₂ surface was calculated to be -0.55 eV, which indicates the exothermicity and spontaneity of the Ni-doping process. Strong interactions occurred in the O₃ and NO₂ systems given the significant adsorption energy (E_{ad}) of -2.44 and -1.93 eV, respectively. Using the band structure and frontier molecular orbital analysis, the sensing response of the Ni–PtTe₂ monolayer upon such two gas species is quite close and large enough for gas detections. Combined with the extremely long recovery time for gas desorption, it is presumed that the Ni–PtTe₂ monolayer is a promising one-shot gas sensor for O₃ and NO₂ detection with a strong sensing response. This study aims at proposing a novel and promising gas sensing material for the detection of the typical fault gases in air-insulated switchgears, so as to ensure their good operation in the whole power system.

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1. Introduction

Air-insulated switchgear is an important type of high-voltage electrical equipment, mainly applied to protect and control the equipment in the power system, such as transformers, generators, capacitors, and cables, to ensure the safety and normal operation of the whole power system.¹ In air-insulated switchgears, the air is used as the insulation medium to prevent partial discharge in the equipment, therefore, they are tremendously applied in medium and low-voltage power equipment owing to their unique merits of low cost and good performance.² However, after a long time running, some inevitable insulation defects, including partial discharge and electric corona, would occur in such kind of equipment, decomposing the air into two toxic gas species, including NO₂ and O₃,^{2,3} wherein the partial overheat can facilitate such process dramatically. With the increasing percentage of unserviceable gases, the insulation performance of air, in which the N₂ as the dominant species undertakes the insulation mission, would be weakened.⁴ This would pose a great threat to the safe operation of air-insulated switchgears and the whole power

grid.⁵ From this regard, the detection of the above-mentioned typical gases has been proposed and proved as a workable and effective approach to estimate the severity degree of the inner insulation defects in the air-insulated switchgears and their working conditions,^{6–8} which has become the hotspot in the gas sensing field.^{9,10}

Regarding gas sensing techniques, the nano-sensing method in recent years has been promoted largely owing to the dramatic developments of materials science and the remarkable advantages of such a method with good electric response and easy preparation.^{11,12} In comparison to traditional sensing materials, such as metal oxides and carbon nanotubes,^{13,14} two-dimensional (2D) nanomaterials can perform better adsorption performance upon gas species owing to their larger surface areas and stronger electron mobility.¹⁵ Therefore, the electrical response of the 2D nanomaterial-based gas sensor is promoted largely, giving rise to their strong potential for gas sensing explorations.^{16–19} Very recently, the transition metal dichalcogenides, namely TMDs, are paid remarkable attention to gas sensing and nano-electronic devices.^{20–23} Especially, noble TMDs, such as PtS₂ and PtSe₂, are successfully synthesized and theoretically verified with different electronic properties in comparison to MoS₂ and MoSe₂.^{24,25} For example, PtS₂ and PtSe₂ monolayers, especially through the transition (TM) doping process, are thoroughly proven with great sensing properties and electrical response upon the gas molecules, with desirable potential for exploration as the nano-electronic devices,^{26–29} in

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which the TM dopant plays an effective role to enhance the charge-transfer in the gas interactions, therefore, promoting the sensing response in the gas detections.

On the other hand, the PtTe₂ monolayer, to the best of our knowledge, is purposed for NO₂ sensing with ultra-high sensitivity and selectivity.³⁰ Besides, the Ru-doped PtTe₂ monolayer has been investigated for the detection of exhaled breath gases so as to explore its potential for the diagnosis of lung cancer.³¹ Also, there are several up-to-date reports about its structural and electronic properties,^{32–34} as well as the TM-doping performance.³⁵ In this research, the unique electronic properties and the desirable chemical reactivity are proved, which is beneficial to explore its potential as a gas sensor, especially with the TM atom doped. Considering the remarkable catalytic behavior of the Ni dopant in the gas interactions to enhance the sensing response of the 2D nano-systems,^{36–38} we propose Ni-doped PtTe₂ (Ni–PtTe₂) monolayer in this paper, based on the first-principles calculations, as a novel candidate for NO₂ and O₃ detection in the air insulated switchgear. A Ni atom is substituted by a Te atom to model the Ni–PtTe₂ configuration and to analyze the Ni-doping performance on the physico-chemical properties of the PtTe₂ monolayer, with a similar method from ref. 39. The findings in this theoretical work can provide detailed guidance to explore PtTe₂-based gas sensors for toxic gas detection in electrical engineering, which is also helpful for their further explorations in many other fields.

2. Computational details

The spin-unrestricted calculations were used in this work to perform all the first-principles simulations within the DMol³ package,⁴⁰ by which the Perdew–Burke–Ernzerhof (PBE) function within the generalized gradient approximation (GGA) was determined to deal with the exchange–correlation interaction.⁴¹ The crystal as a whole is geometrically optimized using the BFGS method in the DMol³ package. The DFT-D2 corrected dispersion was applied to consider the van der Waals force and long-range interactions.⁴² A *k*-point mesh of 10 × 10 × 1 in the Monkhorst–Pack–Brillouin zone was defined for the whole geometric and electronic calculations.⁴³ Besides, we defined the energy tolerance accuracy, self-consistent loop energy, and the global orbital cut-off radius to be 10^{−5} Ha, 10^{−6} Ha, and 5.0 Å,⁴⁴ respectively.

In the current work, we established a 3 × 3 × 1 supercell for the pristine PtTe₂ monolayer to conduct the following calculations. Besides, the vacuum region was determined to be 18 Å along the *z* direction to prevent probable interactions between the neighboring units.⁴⁵ The formation energy (*E*_{form}) is defined to reflect the required energy to substitute a Te atom with a Ni dopant on the pristine PtTe₂ surface, and the gas adsorption energy (*E*_{ad}) is defined to reflect the gas interaction strength on the Ni–PtTe₂ monolayer, calculated by:⁴⁶

$$E_{\text{form}} = E_{\text{Ni–PtTe}_2} - E_{\text{PtTe}_2} - \mu_{\text{Ni}} + \mu_{\text{X}} \quad (1)$$

$$E_{\text{ad}} = E_{\text{Ni–PtTe}_2/\text{gas}} - E_{\text{Ni–PtTe}_2} - E_{\text{gas}} \quad (2)$$

where *E*_{Ni–PtTe₂} and *E*_{PtTe₂} represent the total energy of the Ni-doped and the pristine PtTe₂ monolayer, respectively, while the μ_{Ni} and μ_{X} represent the chemical potential of per Ni and replaced Te or Pt atom in their bulk structures, respectively. *E*_{Ni–PtTe₂/gas} and *E*_{gas} represent the total energies of the gas-adsorbed system and isolated gas molecule, respectively. Moreover, the charged value of the Ni dopant (*Q*_{Ni}) in its substituting process and that of the gas species (*Q*_T) in the gas adsorption systems were considered by the Mulliken analysis.⁴⁷

3. Results and discussion

3.1 Ni-doping behavior on the PtTe₂ monolayer

The Ni-doping process on the PtTe₂ monolayer is plotted in Fig. 1, in which Fig. 1(a) shows the pristine PtTe₂ monolayer, whereas Fig. 1(b) and (c) show the configurations of Ni–PtTe₂ monolayer and related charge density difference (CDD). For the pristine PtTe₂ monolayer, the lattice constant and the Pt–Te bond with geometric optimization were obtained as 4.03 and 2.74 Å, respectively, which are consistent with the previously reported values of 4.02 and 2.76 Å,³² respectively, verifying that the calculations in this work were reliable. When the Ni–PtTe₂ monolayer is formed by substituting a Te atom with the Ni atom, one can see that the Ni atom is shrunk within the Te atomic layer and the Ni–Pt bond length, measured to be 2.51 Å, is much shorter than that of the Pt–Te bond (2.74 Å). This to a large extent is attributed to the much shorter atomic radii of the Ni dopant compared to the Te atom.⁴⁸ Besides, the *E*_{form} for Ni-doping on the PtTe₂ surface by replacing the Te atom was calculated to be −0.55 eV, and by replacing the Pt atom was calculated to be 1.08 eV. That is, the substitution of a Te atom by a Ni dopant on the PtTe₂ monolayer was exothermic while the substitution of a Pt atom by a Ni dopant on the PtTe₂ monolayer was not energy-favorable. In other words, the Ni-doping process would be more likely to occur by replacing the Te atom on the pristine PtTe₂ monolayer, and by such means, the synthesis of Ni–PtTe₂ monolayer was fully procurable and spontaneous at room temperature,^{49,50} which is beneficial to reduce the difficulty for its synthesis from the large-scale point of view. Moreover, we performed the DFT + *U* method (the *U* value was set as 3 eV) to calculate the *E*_{form} for Ni-doping on the pristine PtTe₂ monolayer by replacing the Te atom, and the *E*_{form} was calculated as −0.58 eV. Such a finding indicates that the DFT + *U* method has no big impact on the Ni-doping process.

According to the Mulliken analysis, the Te dopant is positively charged by 0.229*e* in the pristine PtTe₂ monolayer, while the Ni dopant at the Te atomic site is positively charged by 0.211*e* in the Ni–PtTe₂ system. That is, the Te atom shows stronger electron-donating properties than the Ni dopant when bonding with the Pt atom, which is consistent with their electronegativities, in the order of Pt (2.28) > Te (2.10) > Ni (1.91).⁵¹ From the CDD, the Ni dopant is surrounded by electron depletion while the Ni–Pt bonds are embraced by the electron accumulation, which indicates the electron-releasing property of the Ni dopant when doping on the PtTe₂ surface as well as the strong binding force of the Ni–Pt bonds. Also, these electron

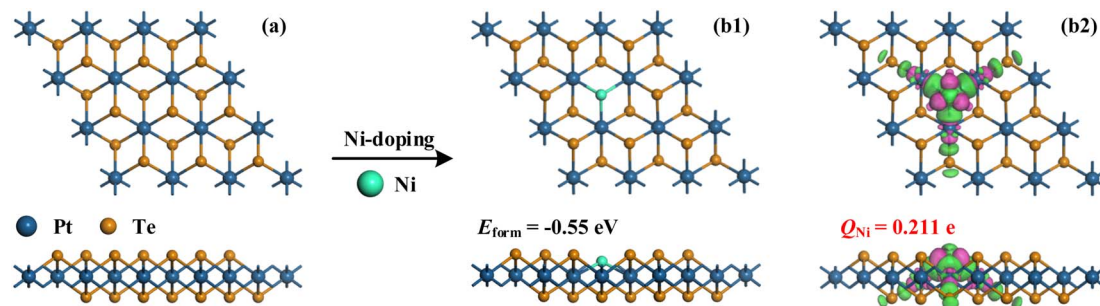


Fig. 1 Ni-doping configuration on the PtTe_2 monolayer. (a) Pristine PtTe_2 monolayer, (b1) and (b2) structure of Ni- PtTe_2 monolayer and related CDD. In CDD, the green area is electron accumulation, while the rosy area is electron depletion, and the isosurface was 0.01 eV \AA^{-3} .

distributions are in good agreement with the Mulliken population analysis.

Fig. 2 shows the band structure (BS) of the pristine and Ni-doped PtTe_2 monolayer as well as the orbital density of state (DOS) of the atoms in the Ni-Pt bond, in order to deeply comprehend the electronic property of Ni-doping on the PtTe_2 surface. From the BS of the pristine PtTe_2 monolayer, the top valence band and the bottom conduction band are r localized on the Γ and K point, respectively, and the bandgap was obtained as 0.612 eV. These findings are in good accordance with the previous report,³⁰ which revealed the indirect semiconducting property for the pristine PtTe_2 monolayer with a bandgap of 0.69 eV based on the GGA function, and it is worth noting that such a small difference in bandgap could be ascribed to the different basic set in the diverse calculations. For the Ni- PtTe_2 monolayer, one can see that the top valence band and the bottom conduction band are both localized on the K point with a bandgap of 0.506 eV. These findings indicate that Ni-doping not only narrows the bandgap, but also modifies the indirect semiconducting property of the pristine PtTe_2 monolayer making it become a direct semiconductor, which might be beneficial for its further exploration in the optical field. Besides, the spin-up is symmetric with the spin-down, which revealed the non-magnetic property of the Ni- PtTe_2 monolayer, namely

Ni-doping has no impact on the magnetic property of the PtTe_2 monolayer. From the atomic DOS, the Ni 3d orbital is highly hybrid with the Pt 5d orbital at -3.3 , -2.8 , -1.6 – 0 , and 0.4 – 0.9 eV, which revealed the admirable orbital interaction between the Ni dopant and the Pt atom, verifying the strong binding force of the Ni-Pt bonds.⁵²

3.2 Gas adsorptions on Ni- PtTe_2 monolayer

O_3 and NO_2 adsorptions on the Ni- PtTe_2 monolayer were performed around the Ni dopant using various configurations with the distance between the gas species and the Ni dopant appropriate 2.5 \AA , which is a proper empirical value to initiate surface interactions in the previous report.⁵³ For O_3 adsorption above the Ni atom, three configurations are considered, namely the molecule-parallel position, two non-bonded O atom-oriented positions, and the middle O atom-oriented position. Also, For NO_2 adsorption above the Ni atom, three configurations can be traced, namely the molecule-parallel position, two O atom-oriented positions, and the N atom-oriented position. With full relaxation, the most stable configuration (MSC) for gas adsorptions in line with related CDD is depicted in Fig. 3.

In the O_3 adsorption system, the O_3 molecule is somewhat vertical to the Ni- PtTe_2 surface and two non-bonded O atoms

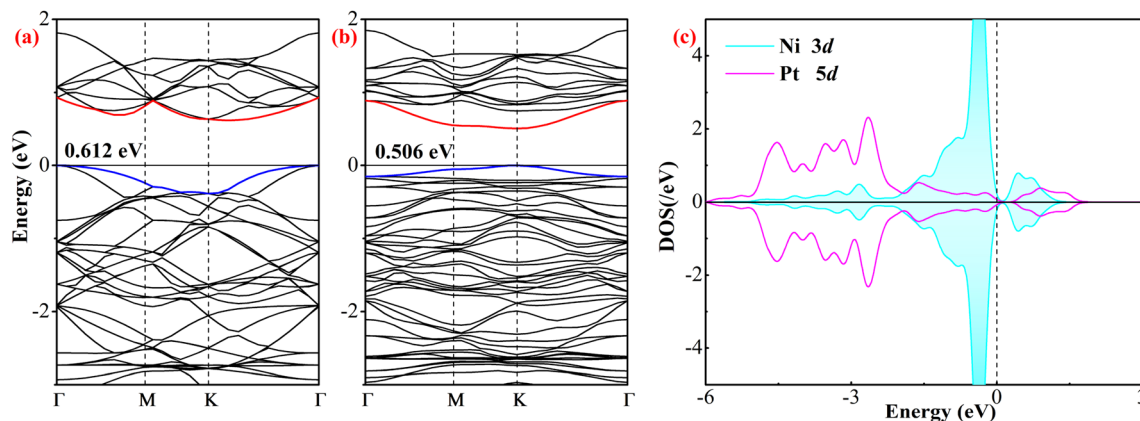


Fig. 2 Electronic property of pristine and Ni-doped PtTe_2 monolayer. (a) and (b) BS of pristine and Ni-doped PtTe_2 monolayer, and (c) orbital DOS of Ni dopant and Pt atom. In BS, the red and blue lines signify the bottom conduction band and top valence band, respectively, and the black values are the bandgaps. In DOS, the Fermi level was set as 0 eV.

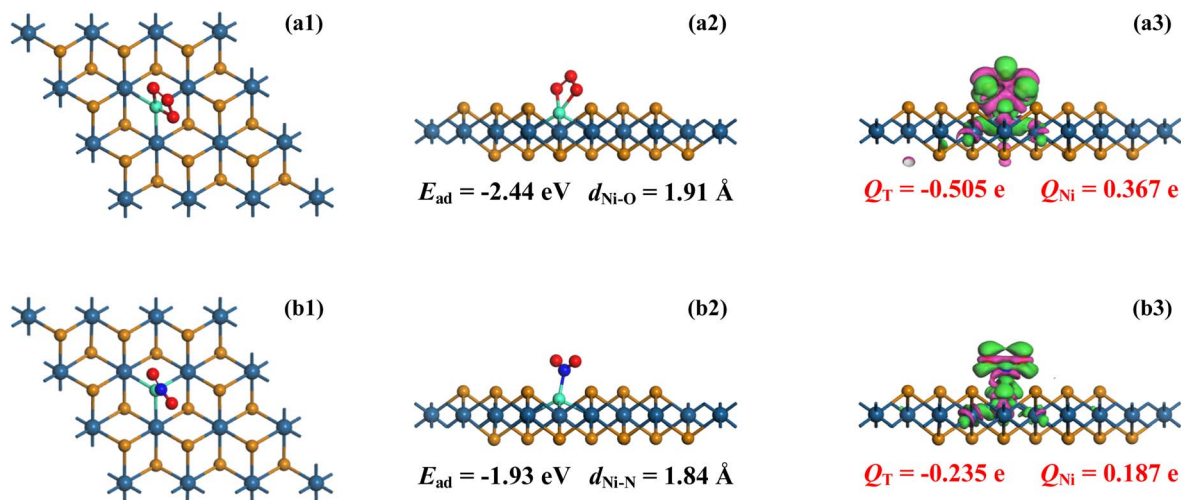


Fig. 3 MSC in the top and side view as well as the CDD of (a1)–(a3) O₃ system and (b1)–(b3) NO₂ system. In CDD, the basic sets are the same as that in Fig. 1.

are captured by the Ni dopant, forming two Ni–O bonds, with lengths of 1.91 and 1.93 Å, respectively. The E_{ad} for O₃ adsorption on the Ni–PtTe₂ monolayer was calculated to be -2.44 eV, whose absolute value is significantly larger than the critical value of 0.80 eV⁵⁴ to clarify such interaction as chemisorption. Besides, we should mention that the adsorption performance of the Ni–PtTe₂ monolayer upon O₃ is stronger than that of the Pd-doped GaN nanotube and Rh-doped ZnO monolayer given the more negative E_{ad} here compared with those of -1.548 eV² in the former system and -1.35 eV⁵⁵ in the latter system. At the same time, such a value indicates the strong binding force of the Ni–O bonds that gives rise to the remarkable E_{ad} in this system. From the Mulliken population analysis, the O₃ molecule is negatively charged by $0.505e$ while the Ni dopant is positively charged by $0.367e$. From these results, one can infer that in the O₃ adsorption process, the O₃ molecule accepts $0.505e$ in total from the Ni–PtTe₂ surface, among which $0.156e$ is donated by the Ni dopant while $0.349e$ is donated by the PtTe₂ monolayer. From the CDD, the O₃ molecule is embraced by the electron accumulation as well as the Ni–O bonds, while the Ni dopant is surrounded by the electron depletion. The electron distributions not only verified the Mulliken population analysis, but also proved the strong electron-accepting property of the O₃ molecule in the gas interactions.⁵⁶

For NO₂ adsorption on the Ni–PtTe₂ monolayer, it was seen that the NO₂ molecule, similar to that in the O₃ system, was also vertical to the PtTe₂ surface and the N atom was trapped by the Ni dopant forming a Ni–N bond as measured to be 1.84 Å. The E_{ad} herein was obtained as -1.93 eV, whose absolute value is smaller than that in the O₃ system. Even though, chemisorption could be identified as well given the judging terms of the critical value (absolute E_{ad} larger than 0.8 eV⁵⁷). Such E_{ad} is comparable with that in the PdN₃-doped CNT system in which the E_{ad} is obtained as -1.91 eV for NO₂ adsorption⁵⁸ and is more negative than that in the Rh-doped MoSe₂ monolayer in which the E_{ad} is obtained as -1.88 eV for NO₂ adsorption.⁵⁹ Besides, the bond

length of Ni–N in the NO₂ system is shorter than the Ni–O in the O₃ system by 0.07 Å. Considering the larger atomic radii of N compared to O by 0.08 Å (N is 71 Å and O is 63 Å⁴⁸), the stronger interaction in the O₃ system in comparison to the NO₂ system could also be verified, thereby narrowing the Ni–N bond much more than the Ni–O bond. According to the Mulliken population analysis, the NO₂ molecule is charged by $-0.235e$, while the Ni dopant is charged by $0.187e$. In other words, the NO₂ molecule and the Ni dopant accept 0.235 and $0.024e$, respectively, from the PtTe₂ monolayer which behaves as an electron donor, relieving $0.259e$ in total for NO₂ adsorption. From the CDD, there have strong electron accumulations dominantly on the NO₂ molecule, the Ni dopant, and the Ni–N bond, which agrees with the Mulliken population analysis well.

Moreover, we should note that considering the size effect of the Ni–PtTe₂ supercell, we performed the adsorption of O₃ and NO₂ on the $4 \times 4 \times 1$ Ni–PtTe₂ monolayer. To accomplish this, the pristine and Ni-doped $4 \times 4 \times 1$ PtTe₂ monolayer should be geometrically optimized as well. The obtained configurations are plotted in Fig. S1.† From Fig. S1† we can find that the E_{form} for Ni-doping on the pristine $4 \times 4 \times 1$ PtTe₂ monolayer (-0.60 eV) is somewhat impacted, becoming slightly negative in comparison with that of -0.55 eV in the $3 \times 3 \times 1$ PtTe₂ monolayer. On the other hand, the adsorption performance of the $4 \times 4 \times 1$ Ni–PtTe₂ monolayer has changed little compared with the $3 \times 3 \times 1$ Ni–PtTe₂ monolayer, indicating that the $3 \times 3 \times 1$ supercell would be large enough and suitable for performing adsorption of O₃ and NO₂.

3.3 Electronic property analysis

All the above analyses reveal the stronger adsorption performance of the Ni–PtTe₂ monolayer upon O₃ compared to NO₂, therefore, one can assume that the electronic response in the O₃ system might be much larger than that in the NO₂ system. Fig. 4 depicts the BS, frontier molecular orbital (FMO), and atomic DOS of two gas systems. It is worth noting that the FMO is also

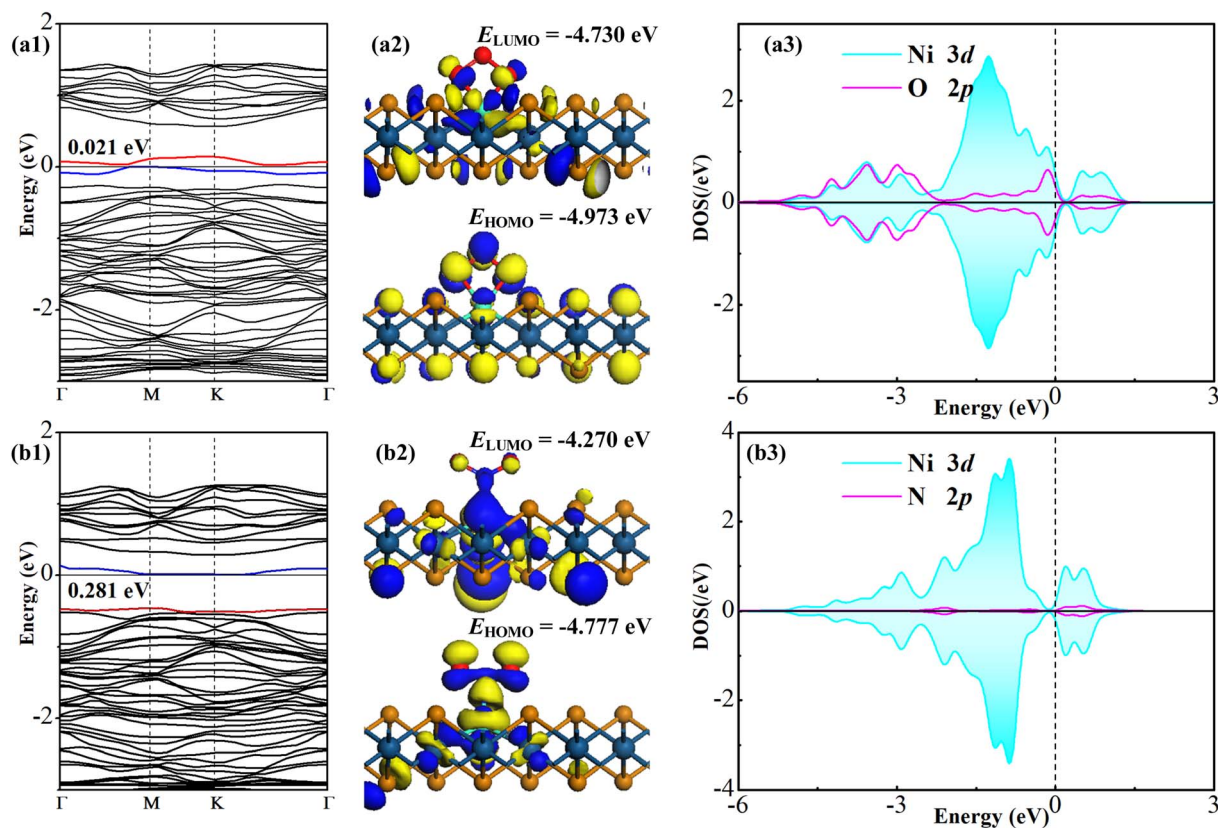


Fig. 4 BS, FMO and atomic DOS of (a1)–(a3) O₃ system and (b1)–(b3) NO₂ system.

an important parameter, which includes the highest occupied molecular orbital (HOMO) and lowest unoccupied molecular orbital (LUMO), to reflect the electrical conductivity (resistance) of the system according to the energy gap (E_g) between the HOMO and LUMO. The larger E_g indicates the smaller (larger) electrical conductivity (resistance).⁶⁰ Also, we should note that we only plotted the orbital DOS of the outermost electron shell of the bonded atoms in order to directly and effectively reflect the orbital interaction between them. Besides, we plotted the FMO of the pure Ni–PtTe₂ monolayer, as shown in Fig. 5, for comparison with the gas-adsorbed systems. We should note that in the BS distributions of the gas adsorbed systems, the spin-up is symmetric with the spin-down, revealing that gas adsorption has no impact on the magnetic property of the Ni–PtTe₂ monolayer.

For the BS of O₃ and NO₂ systems, one can see that the bandgaps are obtained as 0.021 and 0.281 eV, respectively, with different localizations for the top valence band and bottom

conduction band. That is, O₃ and NO₂ adsorptions largely deform the electronic property of the Ni–PtTe₂ monolayer, narrowing its bandgap significantly and altering its direct semiconducting property as well, which may be attributed to the induced novel states by the adsorption of O₃ and NO₂ that fall in the bandgap of the isolated Ni–PtTe₂ monolayer. Based on the dependence between the electrical resistance (σ) and the bandgap as depicted in eqn (5),⁶¹ wherein the λ is a constant, B_g is the bandgap, k is the Boltzmann constant and T is temperature, one can presume that the electrical resistance of the Ni–PtTe₂ monolayer could be largely reduced after adsorptions of such two gas species, especially for O₃. This provides the strong potential to explore the Ni–PtTe₂ monolayer as a gas sensor for the detection of O₃ and NO₂.

$$\sigma = \lambda \times e^{(-B_g/2kT)} \quad (3)$$

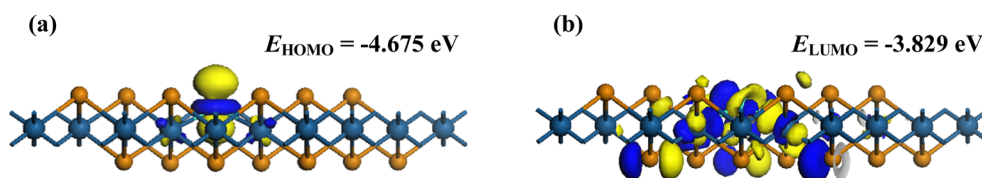


Fig. 5 HOMO (a) and LUMO (b) distributions of Ni–PtTe₂ monolayer.

Similar phenomena can also be found in the FMO distributions of the gas adsorbed systems, in comparison with the pure Ni–PtTe₂ counterpart. It is seen that the HOMO and LUMO are dominantly on the Ni dopant and their energies are –4.675 and –3.829 eV, respectively. At the same time, there are strong HOMO and LUMO distributions around Ni–O bonds in the O₃ system and around the Ni–N bond in the NO₂ system. These distributions of electron cloud manifest the strong interactions in the adsorptions and the strong binding force of Ni–O and Ni–N bonds.⁶² For the O₃ system, the energies of HOMO and LUMO are obtained as –4.973 and –4.730 eV, respectively, while for the NO₂ system, those are obtained as –4.777 and –4.270 eV, respectively. Accordingly, the E_g is calculated to be 0.846 eV for the pure Ni–PtTe₂ monolayer, and 0.243 and 0.507 eV for O₃ and NO₂ systems, respectively. The results from the FMO analysis are in good accordance with those of the BS analysis in that the adsorptions of O₃ and NO₂ can remarkably narrow the bandgap (energy gap) of the Ni–PtTe₂ monolayer, therefore, reducing its electrical resistance significantly.

Moreover, from the orbital DOS of the bonding atoms, one can see that the Ni 3d orbital is highly overlapped with the O 2p orbital at –4.8 to –2.3, –1.6 to –0.1 and 0.5–1.0 eV, which reveals the strong orbital interaction in the Ni–O bond for O₃ adsorption, while the Ni 3d orbital is slightly hybrid with the N 2p orbital at –2.1, –0.5, and 0.2–0.5 eV, which revealed the much weaker binding force of the Ni–N bond in NO₂ adsorption compared with the Ni–O bond. These findings are consistent with the above analysis on the adsorption performance of Ni–PtTe₂ monolayer upon two gas species.

3.4 Gas sensor explorations

The above section indicated the change of bandgap based on BS analysis or of E_g based on FMO analysis for the Ni–PtTe₂ monolayer after adsorptions of two gases, therefore, the exploration of Ni–PtTe₂ monolayer as a resistance-type gas sensor is full of potential. As reported, the sensing response (S) of a gas sensor can be evaluated by eqn (4):⁶³

$$S = (\sigma_{\text{gas}}^{-1} - \sigma_{\text{pure}}^{-1}) / \sigma_{\text{pure}}^{-1} \quad (4)$$

wherein the σ is electrical resistance that can be calculated by eqn (3).

Combined with eqn (3) and (4), the sensing responses of the Ni–PtTe₂ monolayer upon O₃ and NO₂ were obtained as –99.99% and –98.75%, according to the bandgap, and are to be –99.99% and –99.86%, respectively, according to the E_g . That is, the Ni–PtTe₂ monolayer has an admirable sensing response for O₃ and NO₂ detections, fully meeting the requirement for measurement by an electrochemical workstation⁶⁴ no matter what analyzing method is applied. Besides, the recovery property is further considered to analyze the possible reusability of the Ni–PtTe₂ monolayer for gas detection. Based on the van't-Hoff-Arrhenius expression,⁶⁵ the recovery time (τ) for gas desorption from the Ni–PtTe₂ surface can be calculated by eqn (5):

$$\tau = A^{-1} e^{(-E_{\text{ad}}/K_{\text{B}}T)} \quad (5)$$

in which A is the attempt frequency (10^{12} s^{-1}),⁶⁶ T is temperature and K_{B} is Boltzmann constant ($8.318 \times 10^{-3} \text{ kJ (mol K)}^{-1}$). Therefore, the recovery time for the O₃ and NO₂ systems at room temperature (298 K) was calculated to be 1.54×10^{29} and 4.42×10^{20} s, and those even at 498 K were 4.43×10^{12} and 3.43×10^7 s, respectively. In other words, the recycle detection for the Ni–PtTe₂ monolayer is almost impossible due to the extremely long recovery time for gas desorption from its surface. This also reflects the quite strong interactions in the O₃ and NO₂ system that gives rise to the large absolute value of E_{ad} . In this regard, ultraviolet light should be applied to illuminate the sensing material and facilitate the gas desorption from the nano-surface, which has been proven with desirable performance to significantly shut down the recovery time for the sensing material.⁶⁷

Furthermore, we would like to mention that gas adsorptions on the Ni–PtTe₂ monolayer may largely change its work function (WF), which would bring about its sensing potential for O₃ and NO₂ detections based on the WF-measurement devices. For this purpose, we calculated the WF of the freestanding Ni–PtTe₂ monolayer and those after gas adsorptions, which are fully presented in Fig. 6. From this figure, one can see that the WF of the freestanding Ni–PtTe₂ monolayer is calculated to be 4.73 eV, which is somewhat larger than the WF of graphene (4.60 eV⁶⁸), indicating the weaker electron affinity of the Ni–PtTe₂ monolayer. In other words, the Ni–PtTe₂ monolayer behaves relatively weaker electron-liberating properties than the graphene.⁶⁹ In addition, the WF in the O₃ and NO₂ systems are increased to 4.84 and 5.01 eV, respectively. That is, the WF of the Ni–PtTe₂ monolayer is increased by 2.32% after the adsorption of the O₃ molecule, and by 5.92% after the adsorption of the NO₂ molecule. These findings reveal that the WF of the Ni–PtTe₂ monolayer after adsorptions of two species is too small to be explored as a WF-based gas sensor for the detection of O₃ and NO₂ using the Kelvin oscillator device.⁷⁰

In short, the Ni–PtTe₂ monolayer is a promising resistance-type gas sensing material to be explored for O₃ and NO₂ detections, by which the detection is based on the change of

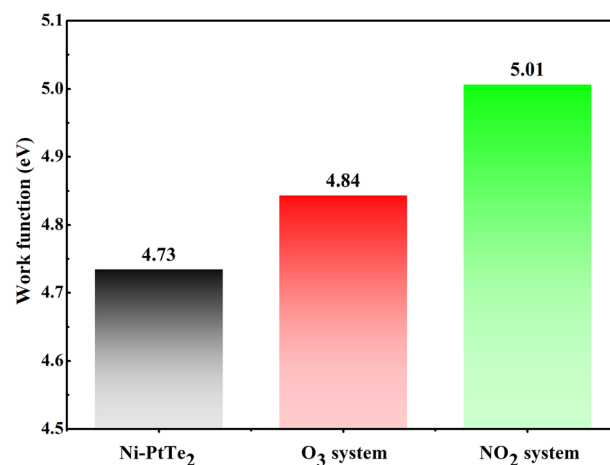


Fig. 6 WF of Ni–PtTe₂ system and Ni–PtTe₂/gas systems.

electrical resistance. Therefore, from the aspect of operation status evaluation for the air-insulated switchgear, the Ni–PtTe₂ monolayer is a promising candidate *via* the detections of O₃ and NO₂. However, such a sensor can only realize the one time operation for gas detection with a desirable sensing response, due to the long recovery time for recycling use.⁵⁹ Apart from that, the Ni–PtTe₂ monolayer is not suitable for the exploration of a WF-based O₃ and NO₂ gas sensor due to the limited WF change after their adsorptions.

4. Conclusions

In this work, we proposed the Ni–PtTe₂ monolayer as a novel and promising gas sensing material upon O₃ and NO₂ detection in the air-insulated switchgear, using the first-principles theory. The main conclusions are as follows:

(i) The E_{form} for Ni-doping on the PtTe₂ surface was calculated to be -0.55 eV, which indicated the exothermicity and spontaneity for Te-substitution by a Ni atom.

(ii) Strong interactions, namely chemisorption, occurred in the O₃ and NO₂ systems given the significant E_{ad} values of -2.44 and -1.93 eV, respectively.

(iii) The sensing response of the Ni–PtTe₂ monolayer upon such two gas species is large enough to realize their detections, on the basis of the band structure and frontier molecular orbital analysis.

(iv) Ni–PtTe₂ monolayer is a promising resistance-type and one-shot gas sensor for O₃ and NO₂ detection with a strong sensing response, given the extremely long recovery time for gas desorption.

Overall, this work can broaden the family members of the gas sensing materials for the detection of the typical fault gases in the high-voltage insulation devices, which is significant to guarantee the good operation of the power system.

Conflicts of interest

The authors declare no conflicts of interest.

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