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# DFT calculations bring insight to internal alkyne-to-vinylidene transformations at rhodium PNP- and PONOP-pincer complexes†

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Density Functional Theory (DFT) has been used to investigate the alkyne-to-vinylidene isomerisation reaction mediated by  $[\text{Rh}(\text{P}(\text{X})\text{NXP})]^+$  complexes ( $\text{X} = \text{CH}_2$ : 2,6-bis(di-*tert*-butylphosphinomethyl)pyridine (PNP) and  $\text{X} = \text{O}$ : 2,6-bis(di-*tert*-butylphosphinito)pyridine (PONOP)) for terminal alkynes  $\text{HC}\equiv\text{CR}$ , where  $\text{R} = \text{tBu}$  and  $\text{Ar}'$  (3,5- $\text{tBu}_2\text{C}_6\text{H}_3$ ). Calculations suggest the reaction mechanism proceeds *via* the slippage of  $\pi$ -bound alkyne at the Rh centre into a Rh-alkyne  $\sigma_{\text{C-H}}$  complex followed by an indirect 1,2-H shift to give the Rh-vinylidene species. NBO (Natural Bond Orbital) analysis of the transition states corresponding to the latter indirect 1,2-H shift step indicates that the migrating hydrogen atom exhibits protic character and hence, the basicity of the H-accepting centre ( $\text{C}_\beta$ ) is controlled by the substituents at that same atom and can tune the 1,2-H shift transition state. QTAIM (Quantum Theory of Atoms in Molecule) and NBO analyses of the Rh-vinylidene complexes indicate that these species exhibit a Rh  $\leftarrow$  C dative bond as well as  $\pi$ -back bonding from the Rh centre into the empty  $p_z$  orbital of the carbene centre ( $\text{C}_\alpha$ ), showing the Rh-vinylidene complexes are Fischer type carbenes. Analysis of the alkyne and vinylidene complex HOMOs show that the equilibrium between the isomers can be tuned by the P-Rh-P bite angle of the  $[\text{Rh}(\text{pincer})]^+$  fragment. Dictated by the nature of the pincer backbone, wider bite angles shift the equilibrium toward the formation of the Rh-vinylidene isomer (e.g.,  $\text{X} = \text{CH}_2$  and  $\text{R} = \text{Ar}'$ ), while tighter bite angles shift the equilibrium more to the formation of the Rh-alkyne isomer (e.g.,  $\text{X} = \text{O}$  and  $\text{R} = \text{Ar}'$ ).

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## Introduction

Transition metal (TM) vinylidene complexes are key species in organometallic chemistry.<sup>1–9</sup> These compounds can be subjected to a variety of reactions such as hydrogenation, dehydrogenation and C–C bond formation processes.<sup>10</sup> The most general strategy to form TM–vinylidene complexes under facile and simple conditions is the reaction of TM complexes with alkynes. This process can lead to the coordination of alkynes to a metal centre to give  $\eta^2$ -alkyne complexes, which are stable enough to be characterised experimentally<sup>11–17</sup> or can act as transient intermediates that undergo C–R ( $\text{R} = \text{C}_6\text{H}_5$ ,  $\text{Ar}'$  and H) bond cleavage and migration of the R group to yield TM–vinylidene complexes.<sup>18–31</sup> Such reactivity can be feasibly modulated by factors such as changing the nature of alkynes or the transition metal.

Mutoh, Ishii and co-workers described the reaction of  $[\text{CpRu}(\text{dppe})]^+$  (**1**) (dppe = 1,2-diphenylphosphinoethane) with internal alkynes ( $\text{PhC}\equiv\text{CAR}'$ ) and showed that both aryl

substituents undergo the 1,2-migration process to generate the disubstituted Ru–vinylidene species **2** (Fig. 1).<sup>16,32</sup> They demonstrated that groups featuring electron-donating

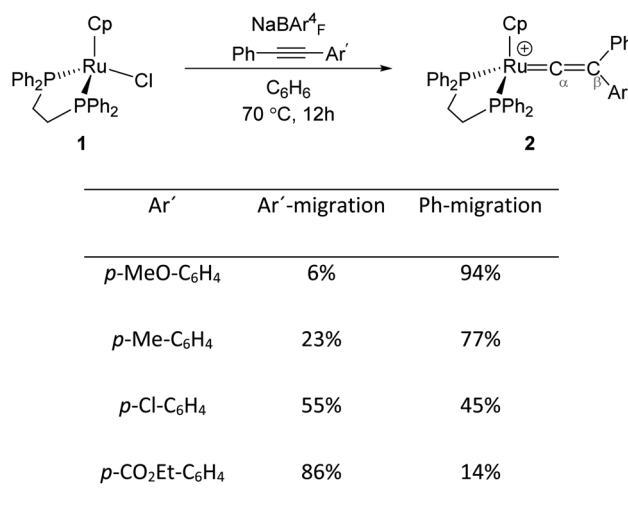
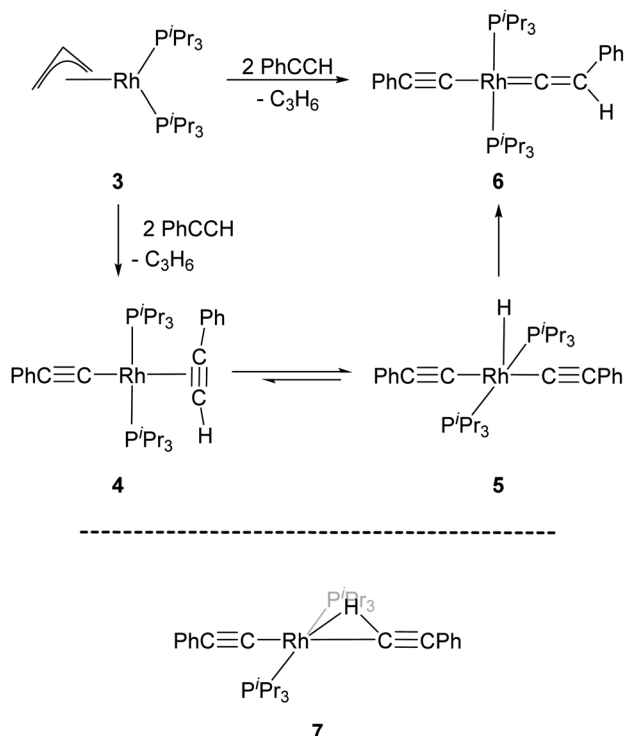


Fig. 1 Reaction of the Ru-precursor **1** with internal alkyne substrates  $\text{PhC}\equiv\text{CAR}'$  to give the Ru–vinylidene species **2**.<sup>16</sup>

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**Scheme 1** Reaction scheme for the formation of the Rh-vinylidene species **6**,<sup>34</sup> and the DFT computed Rh-H alkenyl intermediate **7** featuring a bridging hydrogen between the Rh and C centres.<sup>35</sup>

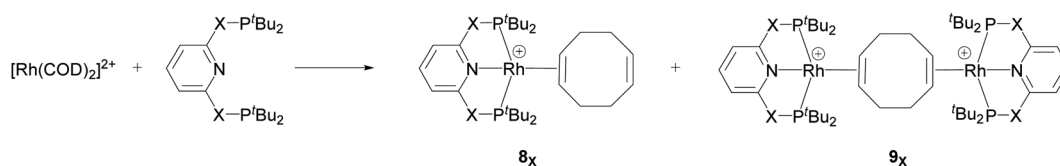
substituents decrease the relative migratory aptitude of the Ar' group over Ph. However, the migratory aptitude becomes opposite when Ar' groups have electron-withdrawing substituents. Based on these results, the authors suggested that this is due to the stabilisation of the negative charge on the Ar' group in the 1,2-Ar' shift transition state, suggesting the 1,2-Ar' migration proceeds *via* an electrophilic mechanism (Fig. 1).<sup>16</sup> DFT (Density Functional Theory) studies by Tsuchida, Takano and co-workers on internal alkyne transformations at  $[\text{CpRu}((\text{C}_6\text{H}_5)_3\text{C}\equiv\text{CC}_6\text{H}_4\text{R-p})(\text{dppf})]^+$  (R = OMe, CO<sub>2</sub>Et) into the corresponding vinylidene complexes shows that this reaction proceeds with a direct 1,2-migration mechanism.<sup>33</sup> In accordance with the experimental findings of Ishii *et al.*, Tsuchida, Takano and co-workers showed that alkyne substrates with an electron-donating substituent (R = OMe) favours the Ph-migration over the Ar'-migration reaction. However, using NBO (Natural Bond Orbital) donor-acceptor interaction analysis, they stated that this reaction proceeds *via* a nucleophilic

mechanism; as in the 1,2-Ar'/Ph migration process, the electron-donating substituted Ar' group stabilises the positive charge of the accepting carbon centre (C<sub>β</sub>), which facilitates the 1,2-Ph migration process.

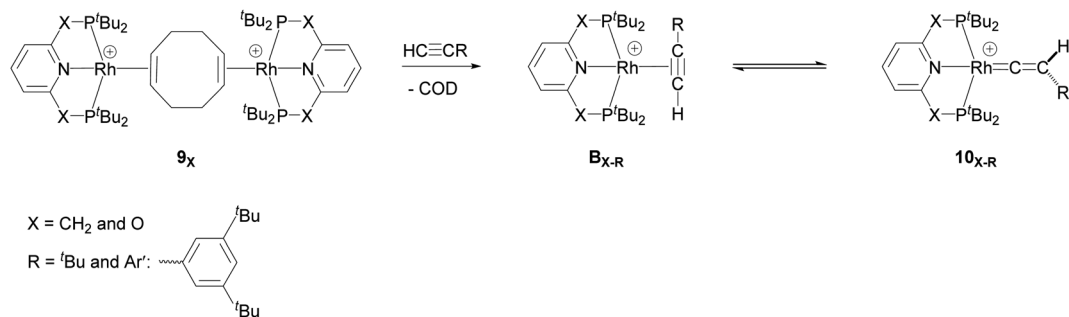
Werner and co-workers proposed the reaction of  $[(\eta^3\text{-C}_3\text{H}_5)_2\text{Rh}(\text{P}^i\text{Pr}_3)_2]$  (**3**) with terminal alkynes HC≡CR (R = H, Me, <sup>t</sup>Bu and Ph) eliminates C<sub>3</sub>H<sub>6</sub> and yields the Rh-alkyne species **4**, which undergoes a C-H bond oxidative cleavage at the Rh centre to form the Rh-H species **5** (Scheme 1).<sup>34</sup> Complex **5** then undergoes a further transformation and gives the Rh-vinylidene species, **6**. A DFT study by Hall and co-workers featured these reaction systems and shows that oxidative cleavage of the first HC≡CR substrate (R = Ph) followed by a C-H reductive elimination at the Rh centre releases propene and forms a Rh-alkenyl intermediate. Addition of a second HC≡CR substrate to the Rh-alkenyl intermediate forms the bridging hydrogen intermediate **7** (Scheme 1), through which the H-migration proceeds with a 1,2-H shift mechanism to give  $[(\text{PhC}\equiv\text{C})_2\text{Rh}(\text{P}^i\text{Pr}_3)_2]$  (**6**).

A DFT study by Angelis and co-workers on the isomerisation of  $[(\text{Cp})(\text{PMe}_3)_2\text{Ru}(\text{HC}\equiv\text{CR})]^+$  to the vinylidene isomer shows that this reaction proceeds *via* an indirect 1,2-H shift reaction.<sup>36</sup> The second possible pathway to from the vinylidene isomer is the C-H oxidative cleavage of the alkyne at the Ru centre followed by a 1,3-H migration reaction to give the Ru-vinylidene complex. However, this was reported to be kinetically inaccessible, as the energy barrier of the 1,3-H shift reaction is significantly higher than the 1,2-H shift pathway. Later on, the same authors demonstrated that upon increasing the electron-richness of the Ru centre, the oxidative cleavage product can be both kinetically and thermodynamically accessible.<sup>37</sup> Particularly notable is the joint experimental and computational study by Lynam, Fey and co-workers on factors such as the nature of substituents, the metal and the ligands, which can control the thermodynamic preference of metal-vinylidene isomers over the corresponding metal-alkyne isomers.<sup>38</sup> The authors proposed a protocol to design ideal conditions to stabilise metal-vinylidene isomers. For instance in  $[\text{RuCl}_2(\text{PR}'_3)_2](=\text{C}=\text{CHR})$ , with electron-withdrawing substituents (R) and phosphine ligands with electron-rich groups (R'), formation of the Ru-vinylidene isomer is thermodynamically favoured over the Ru-alkyne isomer.

Recently, Chaplin and co-workers described the substitution reaction in  $[\text{Rh}(\text{COD})_2][\text{BAR}^F_4]$  with the pincer ligands PNXNP (COD = 1,5-cyclooctadiene, Ar<sup>F</sup> = 3,5-(CF<sub>3</sub>)<sub>2</sub>C<sub>6</sub>H<sub>3</sub>) (X = CH<sub>2</sub>; 2,6-bis(di-*tert*-butylphosphinomethyl)pyridine (PNP) and X = O; 2,6-bis(di-*tert*-butylphosphinito)pyridine (PONOP)) which leads



**Scheme 2** Reaction of the pincer ligand with the dimeric Rh precursor,  $[\text{Rh}(\text{COD})_2]^{2+}$ , to form **8<sub>x</sub>** and **9<sub>x</sub>**. The counter ions ( $\text{BAR}^F_4$ ) are omitted for clarity.<sup>39</sup>



**Scheme 3** Reaction of the dimeric  $\{\text{Rh}(\text{PNP})\}^{2+}$  complex  $9_X$  with terminal alkynes  $\text{HC}\equiv\text{CR}$  ( $\text{R} = \text{tBu}$  and  $\text{Ar}'$  ( $3,5\text{-tBu}_2\text{C}_6\text{H}_3$ )) to form the Rh-alkyne ( $\text{B}_{X-R}$ ) and the Rh-vinylidene complex ( $10_{X-R}$ ). The counter ions ( $\text{BAR}^F_4$ ) are omitted for clarity.<sup>39</sup>

to the formation of the corresponding PNP and PONOP-pincer complexes  $8_X$  and  $9_X$ .<sup>39</sup> As shown in Scheme 2, this reaction gives the monomeric complex  $8_X$  ( $[\text{Rh}(\text{PXNXP})(\eta^2\text{-COD})][\text{BAR}^F_4]$ ) as the major product and the dimeric complex  $9_X$  ( $[\{\text{Rh}(\text{PXNXP})\}_2(\mu\text{-}\eta^2\text{-}\eta^2\text{-COD})][\text{BAR}^F_4]_2$ ) as the minor product.

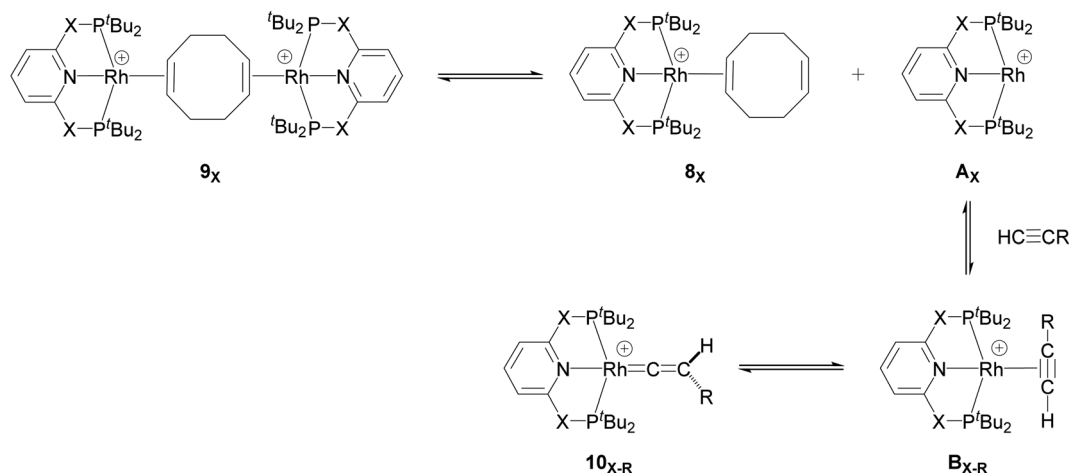
They demonstrated that reaction of the dimeric and monomeric Rh complexes  $8_X$  and  $9_X$  with L-type ligands such as CO results in the formation of the Rh-CO adduct, inferring the COD ligand dissociation from the Rh species forms the cationic  $14e^- \{\text{Rh}(\text{pincer})\}^+$  species, which can be trapped out by CO to form the adduct. The authors also explored the reactivity of the active species,  $\{\text{Rh}(\text{pincer})\}^+$ , by the reaction of the dimeric complex  $9_X$  with  $\text{HC}\equiv\text{CR}$  ( $\text{R} = \text{tBu}$  and  $\text{Ar}'$  ( $3,5\text{-tBu}_2\text{C}_6\text{H}_3$ )) in 1,2-difluorobenzene (DFB) solvent at room temperature which gives the Rh-vinylidene complex  $10_{X-R}$  (Scheme 3).<sup>39</sup> With the dimeric  $\{\text{Rh}(\text{PNP})\}^{2+}$ , this process was found to be very fast (*ca.* five minutes) whereas, with the dimeric  $\{\text{Rh}(\text{PONOP})\}^+$ , it took significantly longer (*ca.* 18 h). It should be noted that with the dimeric  $\{\text{Rh}(\text{PONOP})\}^{2+}$  and  $\text{R} = \text{Ar}'$ , this process forms the Rh-alkyne as the major product and the Rh-vinylidene complex as the minor product (70% *vs.* 30%). This suggests that the alkyne-to-vinylidene transformation process is reversible, a process

that is also seen at  $[\text{RuCl}(\eta^5\text{-C}_9\text{H}_7)(\text{PPh}_3)_2]^{17}$  and  $[\text{CpRu}(\text{dppe})]^+$  complexes.<sup>32</sup>

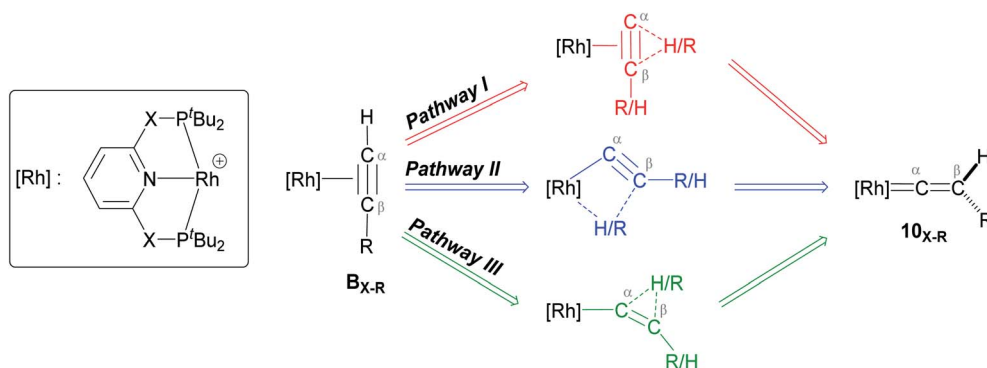
Herein, DFT calculations were carried out to rationalise the mechanism of the formation of these Rh-vinylidene complexes, to characterise the nature of the Rh-carbene bond in the Rh-vinylidene complexes and to understand factors that can affect the equilibrium between the Rh-alkyne and Rh-vinylidene isomers.

## Computational methodology

Gas phase DFT calculations were run with Gaussian 09 (Revision D.01).<sup>40</sup> The Rh and P centres were described with the Stuttgart RECPs and associated basis sets,<sup>41</sup> and the 6-31G\*\* basis set was used for all other atoms (BS1).<sup>42</sup> A polarization function was also added to P ( $\zeta_d = 0.387$ ) and S ( $\zeta_d = 0.503$ ) atoms. Initial BP86 (ref. 43 and 44) optimizations were performed using the 'grid = ultrafine' option, with all stationary points being fully characterized *via* analytical frequency calculations as either minima (all positive eigenvalues) or transition states (one negative eigenvalue). IRC (Intrinsic Reaction Coordinate) calculations and subsequent geometry optimizations were used to confirm the minima are linked by the isolated



**Scheme 4** Formation of the  $14e^-$  Rh species  $\text{A}_X$  and its onward reaction with the alkyne substrates to form the Rh-alkyne ( $\text{B}_{X-R}$ ) and Rh-vinylidene species ( $10_{X-R}$ ).



Scheme 5 Transition states for three potential mechanisms (Pathways I, II and III) to form the Rh-vinylidene complex  $10_{X-R}$ .

transition state. All energies were recomputed with a larger basis set featuring cc-pVTZ-PP for Rh, cc-pVTZ-dk for S and the 6-311++G\*\* basis set for all other atoms (BS2). 1,2-Difluorobenzene is not a parameter defined solvent in the Gaussian program. To overcome this its dielectric constant ( $\epsilon = 13.4$ ) was specified when applying a polarizable continuum model solvent correction with BS1.<sup>45</sup> It should be noted that similar trends to the 1,2-difluorobenzene results were obtained when using fluorobenzene ( $\epsilon = 5.42$ , see ESI† for more details) as the solvation medium. Additional single-point dispersion corrections to the BP86 results employed Grimme's D3 parameter set with Becke-Johnson damping as implemented in Gaussian 09.<sup>46</sup> QTAIM (Quantum Theory of Atoms in Molecules)<sup>47</sup> and NBO (Natural Bonding Orbital)<sup>48</sup> were performed on the BP86-optimised geometries of the Rh-vinylidene complexes  $10_{X-R}$  to characterise the nature of the Rh-carbene bond. The complete computational approach is described in shorthand as BP86-D3BJ (BS2,1,2-difluorobenzene)//BP86(BS1). All free energies are reported in kcal mol<sup>-1</sup> and are given relative to  $9_X$ .

## Results and discussion

Our computed mechanism starts with the splitting of  $9_X$  (as reported by Chaplin and co-workers<sup>39</sup>) to give the 14e<sup>-</sup> Rh species  $A_X$  (Scheme 4, X = CH<sub>2</sub> and O, R = <sup>t</sup>Bu and Ar'). This process is endergonic by  $\Delta G = +20.0$  kcal mol<sup>-1</sup> for X = CH<sub>2</sub> and  $\Delta G = +23.0$  kcal mol<sup>-1</sup> for X = O. The cationic Rh moiety can be stabilised by the coordination of HC≡CR to  $A_X$  in a  $\pi$ -binding mode to the Rh centre to give the Rh-alkyne species  $B_{X-R}$ . For X = CH<sub>2</sub> and O and R = <sup>t</sup>Bu,  $B_{X-tBu}$  lies at a free energy of -0.2 and -3.7 kcal mol<sup>-1</sup> respectively, showing formation of the PONOP complex ( $B_{O-tBu}$ ) is energetically favoured by 3.5 kcal mol<sup>-1</sup>. Coordination of HC≡CAR' to the Rh centre significantly stabilises the cationic Rh species to form  $B_{X-Ar'}$  at  $\Delta G = -12.6$  kcal mol<sup>-1</sup> for X = CH<sub>2</sub> and  $\Delta G = -11.0$  kcal mol<sup>-1</sup> for X = O.

As shown in Scheme 5, once the Rh-alkyne  $\pi$ -complex  $B_{X-R}$  has formed, it can undergo an intra-molecular transformation *via* three possible pathways (I, II and III) to afford the Rh-vinylidene complex  $10_{X-R}$ . In Pathway I, the Rh-alkyne species  $B_{X-R}$  undergoes a direct 1,2-R/H migration to form  $10_{X-R}$ . In Pathway II, transfer of the R/H groups onto the Rh centre occurs

*via* a four-member transition structure to yield  $10_{X-R}$ . Pathway III proceeds with the initial formation of the Rh-C bond that induces the 1,2-H/R shift reaction simultaneously to give  $10_{X-R}$ .

Both the <sup>t</sup>Bu and Ar' alkyne substituents could migrate from the C<sub>β</sub> centre to the C<sub>α</sub> centre from  $B_{X-R}$  to form  $10_{X-R}$ . However, the energy barriers for migration of Ar' or <sup>t</sup>Bu groups *via* the pathways shown in Scheme 3 are computed to be too high (between 37 to 58 kcal mol<sup>-1</sup>)<sup>49</sup> and hence, unsurmountable under the experimental reaction conditions of  $9_X$  and HC≡CR<sup>39</sup> (see ESI, Tables S2 and S3†).

The most obvious route to form the Rh-vinylidene species involves H migration. In Pathway I (Table 1) for the PNP-pincer ligand and R = <sup>t</sup>Bu, the direct 1,2-H shift reaction in  $B_{C-tBu}$  proceeds *via* TS(B-10)<sub>C-tBu</sub> at  $\Delta G^\ddagger = +40.1$  kcal mol<sup>-1</sup> to afford the Rh-vinylidene species  $10_{C-tBu}$  at  $\Delta G = -12.1$  kcal mol<sup>-1</sup>. The free energy barrier for this process ( $\Delta\Delta G_1^\ddagger$ ) is computed to

Table 1 DFT calculated free energies<sup>a</sup> (kcal mol<sup>-1</sup>) for the transformation of the Rh-alkyne species  $B_{X-R}$  to the Rh-vinylidene species  $10_{X-R}$  *via* Pathway I

X	R	$B_{X-R}$	TS(B-10) <sub>X-R</sub>	$\Delta\Delta G_1^\ddagger$	$10_{X-R}$
CH <sub>2</sub>	<sup>t</sup> Bu	-0.2	+40.1	40.3	-12.1
O	<sup>t</sup> Bu	-3.7	+39.4	43.1	-8.9
CH <sub>2</sub>	Ar'	-12.6	+30.3	42.9	-15.7
O	Ar'	-11.0	+31.0	42.0	-10.9

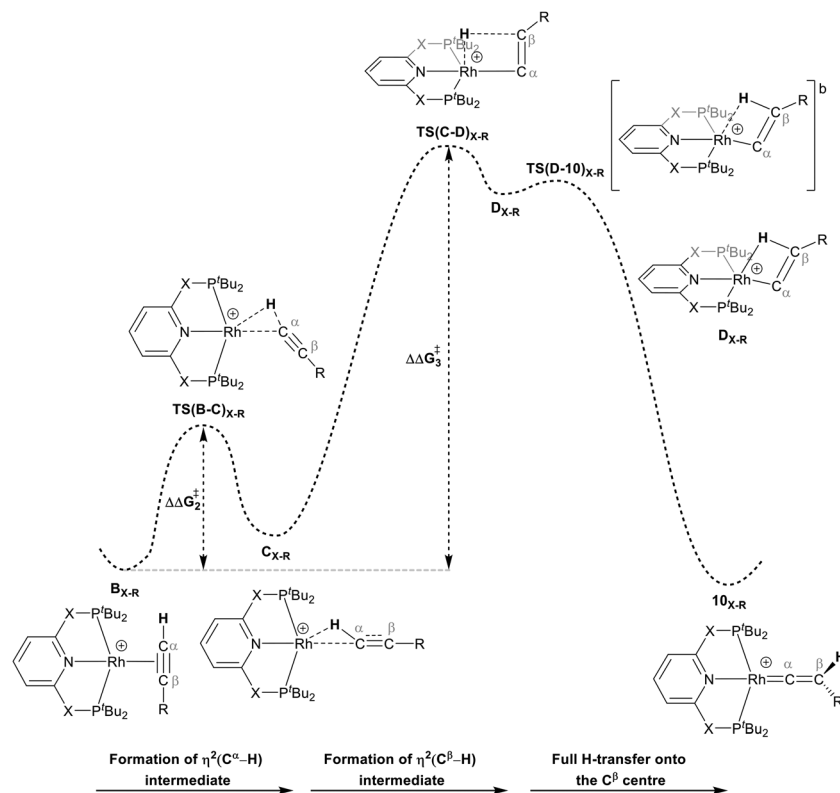
<sup>a</sup> DFT method = BP86-D3(BJ)-1,2-difluorobenzene/BS2//BP86/BS1. Computed energy barrier ( $\Delta\Delta G_1^\ddagger$ ) is relative to  $B_{X-R}$ . All free energies are quoted relative to  $9_X$  (X = CH<sub>2</sub> and O, R = <sup>t</sup>Bu and Ar').

be 40.3 kcal mol<sup>-1</sup> (relative to **B**<sub>C-tBu</sub>), which is too high to be accessible under the reaction conditions. Changing the pincer ligand from PNP to PONOP slightly lowers the free energy (by 0.7 kcal mol<sup>-1</sup>) of the transition state corresponding to the 1,2-H shift process (**TS**(**B**-**10**)<sub>O-tBu</sub>) to +39.4 kcal mol<sup>-1</sup> to give the Rh-vinylidene species **10**<sub>O-tBu</sub> at -8.9 kcal mol<sup>-1</sup>. The free energy barrier for this process is still too high ( $\Delta\Delta G_1^\ddagger = 43.1$  kcal mol<sup>-1</sup>), going from R = <sup>t</sup>Bu to R = Ar', with the PNP-pincer ligand the transition state **TS**(**B**-**10**)<sub>C-Ar'</sub> lies at  $\Delta G_1^\ddagger = +30.3$  kcal mol<sup>-1</sup>; significantly lower than when R = <sup>t</sup>Bu. With respect to **B**<sub>O-Ar'</sub>, the free energy barrier for this process is computed to be 42.9 kcal mol<sup>-1</sup> – again too high and not accessible at room temperature. A similar trend can be seen for the PONOP-pincer ligand when R = Ar' ( $\Delta\Delta G_1^\ddagger = 42.0$  kcal mol<sup>-1</sup>). Therefore, the high energy barriers associated with the direct 1,2-H shift reaction in Pathway I are not in agreement with experiment.

For Pathway II (Table 2), **B**<sub>C-tBu</sub> undergoes slippage of the  $\pi$ -coordination mode to a  $\sigma_{C-H}$  coordination mode *via* **TS**(**B**-**C**)<sub>C-tBu</sub> at  $\Delta G_1^\ddagger = +8.8$  kcal mol<sup>-1</sup> to give **C**<sub>C-tBu</sub>, featuring an elongated C <sub>$\alpha$</sub> -H bond (1.22 Å) to give a short Rh...H of 1.73 Å. NBO analysis of **C**<sub>C-tBu</sub> indicates a donor-acceptor interaction between the C <sub>$\alpha$</sub> -H bonding orbital and a vacant Rh orbital with stabilisation energy ( $\Delta E^{(2)}$ ) of 82.9 kcal mol<sup>-1</sup>, showing a C <sub>$\alpha$</sub> -H agostic interaction with the Rh centre (see ESI, Fig. S3† for details). With respect to **B**<sub>C-tBu</sub>, formation of **C**<sub>C-tBu</sub> proceeds with a free energy barrier ( $\Delta\Delta G_2^\ddagger$ ) of 9.0 kcal mol<sup>-1</sup> and is energetically downhill by 2.8 kcal mol<sup>-1</sup>. **C**<sub>C-tBu</sub> can undergo the H-transfer into the bridging position between the C <sub>$\beta$</sub>  and Rh centre *via* **TS**(**C**-**D**)<sub>C-tBu</sub> at +37.7 kcal mol<sup>-1</sup> to give the four-membered intermediate **D**<sub>C-tBu</sub> at  $\Delta G = +36.4$  kcal mol<sup>-1</sup>.

The H atom then fully transfers onto the C <sub>$\beta$</sub>  centre to give **10**<sub>C-tBu</sub>. Attempts to locate the transition state (**TS**(**D**-**10**)<sub>X-R</sub>) corresponding to this process were inconclusive as it was a very

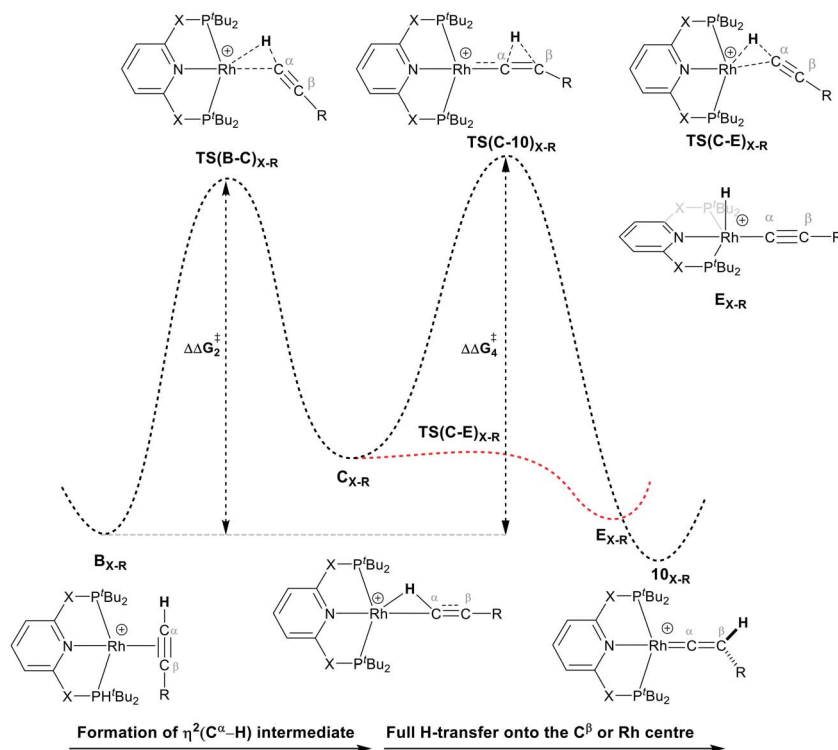
Table 2 DFT calculated free energies<sup>c</sup> (kcal mol<sup>-1</sup>) for the transformation of the Rh-alkyne species **B**<sub>X-R</sub> to the Rh-vinylidene species **10**<sub>X-R</sub> *via* Pathway II



X	R	<b>TS</b> ( <b>B</b> - <b>C</b> ) <sub>X-R</sub>	$\Delta\Delta G_2^\ddagger$	<b>C</b> <sub>X-R</sub>	<b>TS</b> ( <b>C</b> - <b>D</b> ) <sub>X-R</sub>	$\Delta\Delta G_3^\ddagger$	<b>D</b> <sub>X-R</sub>	<b>10</b> <sub>X-R</sub>
CH <sub>2</sub>	<sup>t</sup> Bu	+8.8	9.0	-3.0	+37.7	37.9	+36.4	-12.1
O	<sup>t</sup> Bu	+6.5	10.2	+1.3	+38.5	42.2	+37.0	-8.9
CH <sub>2</sub>	Ar'	+0.2	12.8	-4.5	+29.8 <sup>a</sup>	42.4 <sup>a</sup>	— <sup>b</sup>	-15.7
O	Ar'	+2.8	13.8	-1.0	+31.6 <sup>a</sup>	42.6 <sup>a</sup>	— <sup>b</sup>	-10.9

<sup>a</sup> Estimated free energies by freezing the key distances in **TS**(**C**-**D**)<sub>X-Ar'</sub>. <sup>b</sup> Transition state is not calculated. <sup>c</sup> DFT method = BP86-D3(BJ)-1,2-difluorobenzene/BS2//BP86/BS1. Computed energy barriers ( $\Delta\Delta G_2^\ddagger$  and  $\Delta\Delta G_3^\ddagger$ ) are relative to **B**<sub>X-R</sub>. All free energies are quoted relative to **8**<sub>X</sub> (X = CH<sub>2</sub> and O, R = <sup>t</sup>Bu and Ar').

**Table 3** DFT calculated free energies<sup>a</sup> (kcal mol<sup>-1</sup>) for the transformation of the Rh-alkyne species **B**<sub>X-R</sub> to the Rh-vinylidene species **10**<sub>X-R</sub> (via Pathway III) and the Rh-H species **E**<sub>X-R</sub>



X	R	TS(C-10) <sub>X-R</sub>	C <sub>X-R</sub>	ΔΔG <sub>4</sub> <sup>‡</sup>	10 <sub>X-R</sub>	TS(C-E) <sub>X-R</sub>	E <sub>X-R</sub>
CH <sub>2</sub>	<sup>t</sup> Bu	+8.7	-3.0	8.9	-12.1	-2.8	-8.1
O	<sup>t</sup> Bu	+10.2	+1.3	13.9	-8.9	+0.8	-0.6
CH <sub>2</sub>	Ar'	+7.6	-4.5	20.2	-15.7	-5.5	-11.4
O	Ar'	+9.1	-1.0	20.1	-10.9	-1.2	-4.4

<sup>a</sup> DFT method = BP86-D3(BJ)-1,2-difluorobenzene/BS2//BP86/BS1. Computed energy barrier (ΔΔG<sub>4</sub><sup>‡</sup>) is relative to **B**<sub>X-R</sub>. All free energies are quoted relative to **9**<sub>X</sub> (X = CH<sub>2</sub> and O, R = <sup>t</sup>Bu and Ar').

flat potential free energy surface. A systematic increase of the Rh-H distance revealed that the energy barrier for this process is approximately 1.0 kcal mol<sup>-1</sup> (relative to **D**<sub>C-<sup>t</sup>Bu</sub>). Thus, formation of **10**<sub>C-<sup>t</sup>Bu</sub> via Pathway II is a high energy-process that requires a free energy barrier (ΔΔG<sub>3</sub><sup>‡</sup>) of 37.9 kcal mol<sup>-1</sup> (relative to **B**<sub>C-<sup>t</sup>Bu</sub>). Likewise, for the PONOP ligand and R = <sup>t</sup>Bu formation of the σ<sub>C-H</sub> intermediate **C**<sub>O-<sup>t</sup>Bu</sub> (C<sub>α</sub>-H: 1.21 Å and Rh...H: 1.75 Å, ΔE<sup>(2)</sup> = 72.2 kcal mol<sup>-1</sup>) proceeds with a low free energy barrier (ΔΔG<sub>2</sub><sup>‡</sup> = 10.2 kcal mol<sup>-1</sup>). However, this process is energetically uphill by 5.0 kcal mol<sup>-1</sup>, suggesting it can be reversed to reform the Rh-alkyne species **B**<sub>O-<sup>t</sup>Bu</sub>. The H transfer into the bridging position between the Rh and C<sub>β</sub> centres occurs via **TS**(C-D)<sub>O-<sup>t</sup>Bu</sub> at ΔG<sup>‡</sup> = +38.5 kcal mol<sup>-1</sup> to give **D**<sub>O-<sup>t</sup>Bu</sub> at ΔG = +37.0 kcal mol<sup>-1</sup>. A very flat free energy surface sees the full H-transfer onto the C<sub>β</sub> centre to give **10**<sub>O-<sup>t</sup>Bu</sub> at ΔG = -8.9 kcal mol<sup>-1</sup>. Thus, for Pathway II, formation of the Rh-vinylidene species **10**<sub>O-<sup>t</sup>Bu</sub> requires a high free energy barrier of 42.2 kcal mol<sup>-1</sup>. For R = Ar' and the PNP-pincer ligand, the

transition state (**TS**(C-D)<sub>C-Ar</sub>) corresponding to the formation of the intermediate **C**<sub>C-Ar</sub> (C<sub>α</sub>-H: 1.19 Å and Rh...H: 1.78 Å, ΔE<sup>(2)</sup> = 73.9 kcal mol<sup>-1</sup>) has a free energy of 12.8 kcal mol<sup>-1</sup> - higher than when R = <sup>t</sup>Bu and computed to be energetically uphill by 8.1 kcal mol<sup>-1</sup>. Unfortunately, all attempts to locate the four-member transition state (**TS**(C-D)<sub>X-Ar</sub>) corresponding to the H transfer onto the C<sub>β</sub> centre and the subsequent intermediate **D**<sub>X-Ar</sub> remained allusive. Thus, the transition state free energy value was estimated by freezing the corresponding bond distances and **TS**(C-D)<sub>C-Ar</sub> was calculated to be approximately ΔG<sup>‡</sup> = +29.8 kcal mol<sup>-1</sup>,<sup>50</sup> with a free energy barrier of 42.4 kcal mol<sup>-1</sup> that again is too high to be accessible under the reaction conditions. In comparison to the PNP ligand, the PONOP intermediate **C**<sub>O-Ar</sub> (C<sub>α</sub>-H: 1.18 Å and Rh...H: 1.82 Å, ΔE<sup>(2)</sup> = 67.1 kcal mol<sup>-1</sup>) formed via a slightly higher energy barrier (ΔΔG<sub>2</sub><sup>‡</sup> = 13.8 kcal mol<sup>-1</sup>). This process is computed to be more energetically uphill (by 1.9 kcal mol<sup>-1</sup>). The H transfer process onto the C<sub>β</sub> occurs with a high free energy barrier of

+42.6 kcal mol<sup>-1</sup> to form the Rh-vinylidene species. Thus, as outlined above, although formation of the  $\sigma_{C-H}$  intermediate  $C_{X-R}$  occur with low free energy barriers, the hydrogen transfer onto the  $C_\beta$  involves a large free energy barrier, which is not accessible under reaction conditions.

The indirect 1,2-H shift reaction *via* Pathway III involves the Rh sigma C-H complex  $C_{X-R}$ . As shown in Table 3, in  $C_{X-R}$ , the H centre can either transfer to the Rh centre to give  $E_{X-R}$  or it can transfer onto the  $C_\beta$  centre to form  $10_{X-R}$ . The H-transfer onto the Rh centre occurs with a C-H oxidative cleavage at the Rh centre *via*  $TS(C-E)_{X-R}$  to form the terminal Rh(III)-H species  $E_{X-R}$ . As listed in Table 3, in all cases, the H-transfer process to the Rh centre proceeds with a very flat free energy surface, an essentially barrierless process to give  $E_{X-R}$ , suggesting a facile reaction and importantly consistent with the findings reported by Hall and co-workers.<sup>35</sup> With respect to  $C_{X-R}$ , formation of the terminal Rh-H species  $E_{X-R}$  is slightly exergonic. However,  $E_{X-R}$  is considerably less thermodynamically favoured than the experimentally observed Rh-vinylidene species  $10_{X-R}$  (by  $x$  kcal mol<sup>-1</sup>).

For the H transfer process onto the  $C_\beta$  centre *via*  $C_{X-R}$ , with  $R = 'Bu$  and the PNP-pincer ligand, the bridging hydrogen indirectly can transfer onto the  $C_\beta$  centre *via*  $TS(C-10)_{C-tBu}$ , which lies at  $\Delta G^\ddagger = +8.7$  kcal mol<sup>-1</sup> to generate  $10_{C-tBu}$ . Interestingly, with respect to  $B_{C-tBu}$ , the free energy barrier for this process ( $\Delta\Delta G_4^\ddagger$ ) is computed to be only 8.9 kcal mol<sup>-1</sup>, significantly lower than those in Pathways I and II. With the PONOP-pincer ligand, the indirect 1,2-H shift process proceeds with a free energy barrier of 13.9 kcal mol<sup>-1</sup>, which is 4.0 kcal mol<sup>-1</sup> higher than with the PNP-pincer ligand. The free energy barriers of the indirect 1,2-H migration process for the aryl-substituted alkyne with the PNP or the PONOP-pincer ligands increase to 20.2 kcal mol<sup>-1</sup> and 20.1 kcal mol<sup>-1</sup> respectively. However, they are significantly lower than those in Pathways I and II and values would be accessible at room temperature. Thus, Pathway III exhibits the energetically viable mechanism, consistent with the facile formation of Rh vinylidene species  $10_{X-R}$  seen experimentally.

It should be noted that the Rh-H species  $E_{X-R}$  may undergo a reverse process *via* low free energy barriers (1.4–5.9 kcal mol<sup>-1</sup>) to reform  $C_{X-R}$  and then follow the H-transfer reaction onto the  $C_\beta$  centre to generate  $10_{X-R}$  which is more stable than the Rh-H species  $E_{X-R}$ . For  $R = 'Bu$  and  $X = CH_2$ , the rate-determining step (RDS) is the indirect 1,2-H shift ( $TS(C-10)_{X-R}$ ) while for other systems, RDS is the formation of the  $14e^-$  Rh intermediate  $A_{X-tBu}$ .

### Hydrogen character in the indirect 1,2-H shift reaction: protic or hydridic?

We have shown that formation of the Rh-vinylidene species proceeds *via* Pathway III, in which 1,2-H migration is a key step. However, it is not clear whether in the H-migration process, the H centre migrates as a proton or hydride (*i.e.*, *via* an electrophilic or a nucleophilic mechanism respectively). In this regard, as listed in Table 4, both  $TS(C-10)_{C-tBu}$  and  $TS(C-10)_{O-tBu}$  exhibit a shorter  $C_\alpha \cdots H$  distance than  $C_\beta \cdots H$  by 0.36 Å and 0.35 Å

Table 4 Key atomic distances (Å) and the NBO charges (a.u.) of selected atoms in the transition state  $TS(C-10)_{X-R}$  – corresponding to indirect 1,2-H shift reaction

**TS(C-10)<sub>X-R</sub>**

Structure	$C_\alpha \cdots H$	$C_\beta \cdots H$	Rh	H	$C_\alpha$	$C_\beta$
$TS(C-10)_{C-tBu}$	1.18	1.54	-0.54	+0.39	-0.28	+0.06
$TS(C-10)_{O-tBu}$	1.18	1.55	-0.64	+0.39	-0.29	+0.09
$TS(C-10)_{C-Ar}$	1.22	1.41	-0.55	+0.40	-0.20	-0.04
$TS(C-10)_{O-Ar}$	1.24	1.36	-0.66	+0.40	-0.17	-0.06

respectively. This shows that the H centre exhibits a greater tendency to interact with the  $C_\alpha$  centre than with the  $C_\beta$  centre. A similar trend can also be seen in  $TS(C-10)_{C-Ar}$  and  $TS(C-10)_{O-Ar}$  (0.19 Å and 0.12 Å respectively). However, for  $R = Ar'$  the H centre has a considerably shorter distance to  $C_\beta$ , showing a stronger interaction between the H and  $C_\beta$  centres when  $R = Ar'$ . NBO charge analysis of the transition state  $TS(C-10)_{X-R}$  shows that for  $R = 'Bu$ , the Rh centre exhibits a very similar negative charge value ( $q_{Rh} = -0.54$  to  $-0.55$ ), implying the nature of the X moiety in the pincer ligands has a negligible effect on the charge of the Rh centre. For  $R = Ar'$ , the Rh centre features a slightly more negative charge ( $q_{Rh} = -0.64$  to  $-0.66$ ) than when  $R = 'Bu$ , which can be attributed to the  $\pi$ -electrons of the aryl group of the alkyne. In all cases, the H centre features similar positive charges ( $q_H = +0.39$  to  $+0.40$ ), implying the H centre exhibits protic character, showing it can transfer *via* an electrophilic mechanism. For  $R = 'Bu$ , the  $C_\beta$  centre displays small positive charges, while for  $R = Ar'$ , the  $C_\beta$  centre possesses negative charges, indicating the electronic character of the  $C_\beta$  centre is sensitive to the nature of the R substituent. Interestingly, as the charge of the  $C_\beta$  centre becomes slightly negative (*e.g.*,  $R = Ar'$ ), the transition state  $TS(C-10)_{X-R}$  lies at a lower free energy (by 1.1 kcal mol<sup>-1</sup> for both PNP and PONOP relative to  $R = 'Bu$ ), whereas, when the  $C_\beta$  has a positive charge (*e.g.*,  $R = 'Bu$ ), the transition state  $TS(C-10)_{X-R}$  lies at a higher free energy. This suggests that  $\pi$ -donor R substituents donate electron density to the empty  $C_\beta$   $p_z$  orbital and hence increases its basicity, consistent with the electrophilic mechanism in the 1,2-H migration process. In order to examine this, the indirect 1,2-H shift process was computed for  $X = CH_2$  and  $R = NMe_2$ . Our calculations show that the transition state  $TS(C-10)_{C-NMe_2}$  exhibits a short  $C_\beta-H$  distance of 1.26 Å, showing a stronger interaction between the H and the  $C_\beta$  centres. This is due to  $NMe_2$  being a strong  $\pi$ -donor substituent that significantly populates the  $C_\beta$   $p_z$  orbital to give a more stable transition state

(at  $\Delta G^\ddagger = +4.7 \text{ kcal mol}^{-1}$ ), supporting the electrophilic character of the H centre in the indirect 1,2-H shift process.

### Characterisation of the Rh–vinylidene complexes $10_{X-R}$ : Fischer or Schrock-type complex?

TM–carbene complexes can be described as a Fischer or Schrock carbene complex, depending on the nature of the metal–carbene bond. In Fischer type carbene complexes, the carbene centre donates two electrons to the metal centre *via* a dative interaction whereas the metal centre donates electron density to the  $p_z$  orbital of the carbene centre through  $\pi$ -back bonding (*i.e.*, singlet carbene). In Schrock type carbene complexes, the carbene centre shares two single electron with the metal centre to form a  $\sigma$  and  $\pi$  metal–carbene bonds (*i.e.*, triplet carbene). Thus, the nature of the metal–carbene interactions in  $10_{X-R}$  can provide useful information whether these species are Fischer or Schrock-type carbene complexes. In this context, in order to elucidate the nature of the metal–carbene bond in the Rh–vinylidene species  $10_{X-R}$ , QTAIM and NBO analyses were performed on the BP86-optimised geometries of  $10_{X-R}$ . As shown in Fig. 2, the QTAIM plot of  $10_{C-tBu}$  exhibits a BCP (Bond Critical Point) along the Rh $\cdots$ C $_{\alpha}$  bond path, which is characterised by a positive Laplacian of electron density and small negative total energy density. This is indicative of a dative interaction between the Rh and carbene centres.<sup>51</sup> A similar trend can also be observed for the other Rh–vinylidene species (see ESI, Fig. S4†). NBO analysis of  $10_{C-tBu}$  identifies a Rh  $\sigma$ -bonding orbital that is dominated by carbene character (Rh (34.5%) and C $_{\alpha}$  (66.5%)), showing a Rh  $\leftarrow$  C $_{\alpha}$  dative interaction. Additionally, the NBO analysis reveals the presence of  $\pi$ -back donation from the Rh centre into the empty  $p_z$  orbital of the carbene, which is dominated by the Rh character (Rh (77.6%) and C $_{\alpha}$  (22.4%)), see ESI, Fig. S4† for more details). Thus, the QTAIM and NBO data are fully consistent with a Rh(I) centre ligated by vinylidene (C $_{\alpha}$ =C $_{\beta}$ HR) ligand and hence the Rh–vinylidene complexes  $10_{X-R}$  can be described as a Fischer TM–carbene complex, seen previously in d<sup>8</sup> TM–carbene complexes.<sup>52</sup> It should be noted that in all the Rh–vinylidene complexes, the Rh–C $_{\alpha}$   $\sigma$ -bond is polarised toward the C $_{\alpha}$  centre,

Table 5 The difference between the free energies of  $B_{X-Ar}$  and  $10_{X-Ar}$  represented by  $\Delta G_r$  (in kcal mol<sup>-1</sup>) and HOMOs of  $B_{X-Ar}$  and  $10_{X-Ar}$  and their difference represented by  $\Delta_{HOMO}$  (in eV; X = NH, CH<sub>2</sub>, O and S)

X	R	$\Delta G_r$	HOMO ( $B_{X-R}$ )	HOMO ( $10_{X-R}$ )	$\Delta_{HOMO}$
CH <sub>2</sub>	<sup>t</sup> Bu	-11.9	-7.05	-7.31	+0.26
O	<sup>t</sup> Bu	-5.2	-7.54	-7.65	+0.11
S	Ar'	-3.2	-7.25	-7.31	-0.06
CH <sub>2</sub>	Ar'	-3.1	-6.87	-6.96	-0.09
NH	Ar'	-2.1	-7.01	-7.04	-0.03
O	Ar'	+0.1	-7.29	-7.23	+0.06

while the Rh–C $_{\alpha}$   $\pi$ -interaction is polarised toward the Rh centre (see ESI, Fig. S4†).

### Stability of Rh–alkyne vs. Rh–vinylidene complexes

As shown above, for R = Ar' and the PNP pincer ligand, the Rh–vinylidene isomer is more stable than the Rh–alkyne isomer, while with the PONOP pincer ligand, the Rh–alkyne isomer is more stable. In order to explore the equilibrium between the Rh–alkyne and Rh–vinylidene isomers, the difference between the free energies of  $B_{X-R}$  and  $10_{X-R}$  species were calculated and is represented by  $\Delta G_r$  ( $\Delta G_r = \Delta G(10_{X-R}) - \Delta G(B_{X-R})$ , Table 5). A negative  $\Delta G_r$  value indicates that the equilibrium is toward the Rh–vinylidene isomer while a positive value indicates the equilibrium is toward the Rh–alkyne isomer. As listed in Table 5, our computational findings show that for R = <sup>t</sup>Bu and both PNP or PONOP-pincer ligands,  $\Delta G_r$  is negative, with the PNP-pincer complex being significantly more negative than the PONOP-pincer complex (-11.9 and -5.2 kcal mol<sup>-1</sup>, respectively). For R = Ar' and the PNP-pincer ligand,  $\Delta G_r$  is also negative (-3.1 kcal mol<sup>-1</sup>) yet for the PONOP-pincer ligand  $\Delta G_r$  becomes slightly positive (+0.1 kcal mol<sup>-1</sup>). This trend implies that with the aryl substituent, the Rh–vinylidene complex is less stable than the alkyl substituent. This becomes critically important when switching from the PNP pincer ligand to the PONOP pincer ligand, as in this case, the Rh–alkyne isomer becomes more stable than the Rh–vinylidene isomer. This

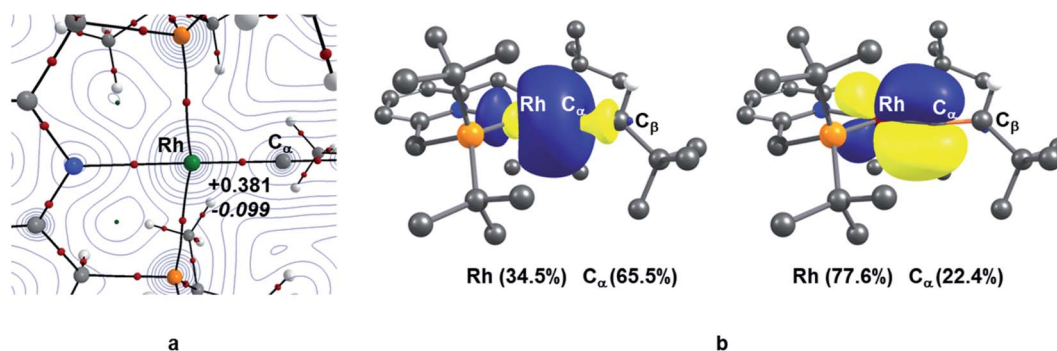


Fig. 2 (a) QTAIM molecular graph of  $10_{C-tBu}$ : the electron density contours are computed in the {Rh/P/C $_{\alpha}$ } plane with bond critical points (BCPs) shown as small red spheres; for the Rh–C $_{\alpha}$  BCP, the Laplacian of the electron density ( $\nabla^2\rho(r)$  in eÅ<sup>-5</sup>, in plain) and total energy density ( $H(r)$  in a.u., in italic) (b) NBO  $\sigma$ - and  $\pi$ -bonding orbitals of the Rh–C $_{\alpha}$  interaction and the contribution of the Rh and C $_{\alpha}$  centres in the corresponding interactions in  $10_{C-tBu}$ .



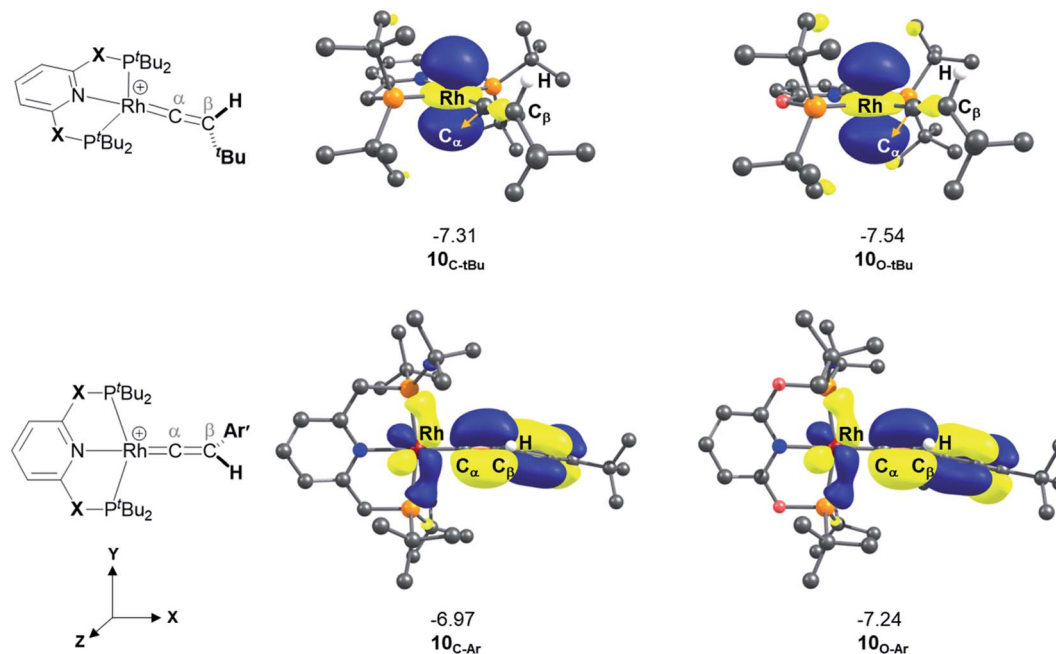


Fig. 3 HOMO (in eV) of Rh-vinylidene species  $10_{X-R}$  ( $X = \text{CH}_2$  and  $\text{O}$ ;  $R = \text{tBu}$  and  $\text{Ar}'$ ). Substituent hydrogen atoms removed for clarity.

highlights the influence of the nature of both substituent and pincer ligand on the stability of the Rh-vinylidene species.

A detailed analysis of the electronic structure of the Rh-alkyne and Rh-vinylidene complexes can help to understand the origin of the equilibrium between the two isomers. As shown in Fig. 3, the molecular orbital analysis of the Rh-vinylidene species  $10_{X-R}$  shows that when  $R = \text{tBu}$ , the HOMO of  $10_{X-R}$  predominantly consists of the Rh  $d_{z^2}$  character, whereas for  $R = \text{Ar}'$ , the HOMO consists of the Rh  $d_{xy}$  orbital and the  $\pi$ -system of the aryl moiety. With the PNP-pincer ligand, when  $R = \text{Ar}'$ , the Rh-vinylidene complex exhibits a higher HOMO than when  $R = \text{tBu}$  ( $-7.31$  eV and  $-6.96$  eV, respectively). This can be attributed to the presence of the  $\pi$ -system of the aryl moiety, which in conjugation with the  $\text{C}_\alpha\text{-C}_\beta$   $\pi$ -bond and the  $\{\text{Rh}(\text{pincer})\}^+$  moiety can destabilise the HOMO of the system and hence affects the stability of the Rh-vinylidene complex. Thus, it suggests that the HOMO energy is a reasonable descriptor that can capture the stability of the Rh-vinylidene isomer *versus* the Rh-alkyne isomer.

Table 5 lists the energy difference between the HOMOs of  $\text{B}_{X-R}$  and  $10_{X-R}$  species, represented by  $\Delta_{\text{HOMO}}$  ( $\Delta_{\text{HOMO}} = \Delta_{\text{HOMO}}(10_{X-R}) - \Delta_{\text{HOMO}}(\text{B}_{X-R})$ ). For  $R = \text{tBu}$  and the PNP pincer ligand,  $\Delta_{\text{HOMO}}$  is  $+0.26$  eV, showing the Rh-vinylidene isomer possess a more stable HOMO than the Rh-alkyne isomer. Similarly, for  $R = \text{tBu}$  and the PONOP pincer ligand,  $\Delta_{\text{HOMO}}$  is also a positive value ( $+0.11$  eV). However, the PONOP analogue has a smaller  $\Delta_{\text{HOMO}}$  than the PNP complex. A similar trend can also be seen for  $R = \text{Ar}'$  ( $+0.09$  eV). Interestingly, for  $R = \text{Ar}'$ , when the pincer ligand is PONOP, the Rh-alkyne complex features a more stable HOMO than the Rh-vinylidene isomer ( $\Delta_{\text{HOMO}} = -0.06$  eV). This suggests that for  $R = \text{Ar}'$ , the nature of the pincer ligand plays an important role on the stability of the

HOMO of the Rh-vinylidene isomer *versus* the Rh-alkyne isomer. In order to investigate this further, the relative free energies of the vinylidene complex against the alkyne isomer was also computed for  $X = \text{NH}$  and  $\text{S}$ , and  $R = \text{Ar}'$ . Both  $X = \text{NH}$  and  $\text{S}$  showed a similar trend to  $X = \text{CH}_2$ , with the Rh-vinylidene complex being more stable than the Rh-alkyne complex ( $\Delta G_{\text{r}} = -2.1$  and  $-3.2$  kcal mol $^{-1}$  respectively) in which the Rh-vinylidene isomer features a more stable HOMO than the Rh-alkyne isomer ( $\Delta_{\text{HOMO}} = -0.03$  and  $-0.06$  eV respectively). Interestingly, a reasonable correlation ( $R^2 = 0.95$ , Fig. S1 $^\dagger$ ) was found between  $\Delta G_{\text{r}}$  and  $\Delta_{\text{HOMO}}$  values, showing when  $\Delta_{\text{HOMO}}$  is negative, the Rh-vinylidene species is more stable than the Rh-alkyne species and when  $\Delta_{\text{HOMO}}$  is positive, the Rh-alkyne species is more stable than the Rh-vinylidene species. This therefore indicates that for  $R = \text{Ar}'$ ,  $\Delta_{\text{HOMO}}$  can capture the stability of the Rh-vinylidene *versus* the Rh-alkyne species, however, raises the question; why is  $\Delta_{\text{HOMO}}$  negative for  $X = \text{CH}_2$ ,  $\text{NH}$  or  $\text{S}$ , while for  $X = \text{O}$ ,  $\Delta_{\text{HOMO}}$  is positive?

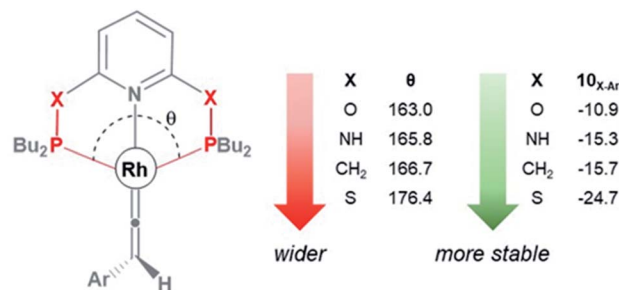


Fig. 4 The P-Rh-P bite angle ( $\theta$ , in degree) and the free energy (kcal mol $^{-1}$ ) of the Rh-vinylidene species  $10_{X-Ar}$ .

As shown in Fig. 3, the HOMO of the Rh–vinylidene complex  $\mathbf{10}_{X-Ar}$  consists of the Rh  $d_{xy}$  orbital that is affected by the P–Rh–P bite angle ( $\theta$ ) of the supporting pincer ligand. For X = O, the  $\theta$  value of  $\mathbf{10}_{O-Ar}$  is  $163.0^\circ$  and its HOMO energy is  $-7.23$  eV (Fig. 4). In comparison to X = O, for X = S, the  $\theta$  value of  $\mathbf{10}_{S-Ar}$  is considerably larger ( $176.4^\circ$ ) and its HOMO is more stable than  $\mathbf{10}_{O-Ar}$  ( $-7.31$  eV). Interestingly, an excellent correlation ( $R^2 = 0.9999$ , Fig. S2†) was found between the free energy of the Rh–vinylidene species  $\mathbf{10}_{X-R}$  and its  $\theta$  value, showing tighter P–Rh–P bite angles result in more reactive Rh–vinylidene species (*i.e.*, higher HOMOs), while wider bite angles give more stable Rh–vinylidene species  $\mathbf{10}_{X-R}$  (*i.e.*, lower HOMO), consistent with recent work from Mansell.<sup>53</sup> Thus, the bite angle is dictated by the nature of the X moiety and can therefore tune the reactivity of the Rh–vinylidene species *versus* the Rh–alkyne species  $\mathbf{B}_{X-R}$ .

## Conclusions

DFT calculations on the transformation of Rh–alkyne  $\mathbf{B}_{X-R}$  into the Rh–vinylidene complex  $\mathbf{10}_{X-R}$  reveal that this process involves the slippage of the Rh–alkyne  $\pi$ -complex into the Rh–alkyne  $\sigma_{C-H}$  complex – followed by an indirect 1,2-H shift reaction. The H atom in the indirect 1,2-H shift transition state has protic character. The presence of  $\pi$ -donor groups at the accepting carbon centre ( $C_\beta$ ) can populate the  $C_\beta$  empty  $p_z$  orbital and hence increase its basicity to give stronger  $C_\beta \cdots H$  interactions, which consequently lower the free energy of the 1,2-H shift transition state. The equilibrium between the alkyne and vinylidene species can be tuned by the P–Rh–P bite angle, which is dictated by the nature of the X moiety in the PXNXP pincer ligand coordinated to the Rh-centre.

## Conflicts of interest

There are no conflicts to declare.

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## Notes and references

- 1 F. Ye, J. Chen and T. Ritter, *J. Am. Chem. Soc.*, 2017, **139**, 7184–7187.
- 2 Y. Yamamoto, K. Matsui and M. Shibuya, *Chem.–Eur. J.*, 2015, **21**, 7245–7255.
- 3 S. Takano, T. Kochi and F. Kakiuchi, *Organometallics*, 2016, **35**, 4112–4125.
- 4 S. W. Roh, K. Choi and C. Lee, *Chem. Rev.*, 2019, **119**, 4293–4356.
- 5 L. Nunes dos Santos Comprido, J. E. M. N. Klein, G. Knizia, J. Kästner and A. S. K. Hashmi, *Chem.–Eur. J.*, 2016, **22**, 2892–2895.
- 6 F. Kakiuchi, S. Takano and T. Kochi, *ACS Catal.*, 2018, **8**, 6127–6137.
- 7 R.-S. Liu, *Synlett*, 2008, **2008**, 801–812.
- 8 B. M. Trost and J. T. Masters, *Chem. Soc. Rev.*, 2016, **45**, 2212–2238.
- 9 M. Otsuka, N. Tsuchida, Y. Ikeda, N. Lambert, R. Nakamura, Y. Mutoh, Y. Ishii and K. Takano, *Organometallics*, 2015, **34**, 3934–3943.
- 10 A. B. Antonova, *Coord. Chem. Rev.*, 2007, **251**, 1521–1560.
- 11 P. Barrio, M. A. Esteruelas and E. Oñate, *Organometallics*, 2003, **22**, 2472–2485.
- 12 M. J. Cowley, J. M. Lynam and J. M. Slattery, *Dalton Trans.*, 2008, 4552–4554, DOI: 10.1039/B806358C.
- 13 V. K. Singh, E. Bustelo, I. de los Ríos, I. Macías-Arce, M. C. Puerta, P. Valerga, M. Á. Ortuño, G. Ujaque and A. Lledós, *Organometallics*, 2011, **30**, 4014–4031.
- 14 M. Jiménez-Tenorio, M. C. Puerta, P. Valerga, M. A. Ortuño, G. Ujaque and A. Lledós, *Inorg. Chem.*, 2013, **52**, 8919–8932.
- 15 Y. Ikeda, Y. Mutoh, K. Imai, N. Tsuchida, K. Takano and Y. Ishii, *Organometallics*, 2013, **32**, 4353–4358.
- 16 Y. Mutoh, Y. Ikeda, Y. Kimura and Y. Ishii, *Chem. Lett.*, 2009, **38**, 534–535.
- 17 M. Bassetti, V. Cadierno, J. Gimeno and C. Pasquini, *Organometallics*, 2008, **27**, 5009–5016.
- 18 D. B. Grotjahn, X. Zeng and A. L. Cooksy, *J. Am. Chem. Soc.*, 2006, **128**, 2798–2799.
- 19 W. Debrouwer and A. Fürstner, *Chem.–Eur. J.*, 2017, **23**, 4271–4275.
- 20 I. de los Ríos, E. Bustelo, M. C. Puerta and P. Valerga, *Organometallics*, 2010, **29**, 1740–1749.
- 21 P.-Y. Chia, C.-C. Kuo, S.-L. Huang, Y.-H. Liu, L.-K. Liu and Y.-C. Lin, *Chem.–Asian J.*, 2018, **13**, 3885–3894.
- 22 J. M. Lynam, C. E. Welby and A. C. Whitwood, *Organometallics*, 2009, **28**, 1320–1328.
- 23 R. H. Lam, D. B. Walker, M. H. Tucker, M. R. D. Gatus, M. Bhadbhade and B. A. Messerle, *Organometallics*, 2015, **34**, 4312–4317.
- 24 F. Mo, H. N. Lim and G. Dong, *J. Am. Chem. Soc.*, 2015, **137**, 15518–15527.
- 25 V. Cadierno, M. P. Gamasa, J. Gimeno, C. González-Bernardo, E. Pérez-Carreño and S. García-Granda, *Organometallics*, 2001, **20**, 5177–5188.
- 26 M. Bassetti, P. Alvarez, J. Gimeno and E. Lastra, *Organometallics*, 2004, **23**, 5127–5134.
- 27 I. Jourdain, L. Vieille-Petit, S. Clément, M. Knorr, F. Villafañe and C. Strohmann, *Inorg. Chem. Commun.*, 2006, **9**, 127–131.
- 28 M. J. Cowley, J. M. Lynam and A. C. Whitwood, *J. Organomet. Chem.*, 2010, **695**, 18–25.
- 29 J. M. Lynam, *Chem.–Eur. J.*, 2010, **16**, 8238–8247.
- 30 D. G. Johnson, J. M. Lynam, J. M. Slattery and C. E. Welby, *Dalton Trans.*, 2010, **39**, 10432–10441.
- 31 F. De Angelis, A. Sgamellotti and N. Re, *Organometallics*, 2007, **26**, 5285–5288.

- 32 Y. Mutoh, K. Imai, Y. Kimura, Y. Ikeda and Y. Ishii, *Organometallics*, 2011, **30**, 204–207.
- 33 M. Otsuka, N. Tsuchida, Y. Ikeda, Y. Kimura, Y. Mutoh, Y. Ishii and K. Takano, *J. Am. Chem. Soc.*, 2012, **134**, 17746–17756.
- 34 M. Schäfer, J. Wolf and H. Werner, *Organometallics*, 2004, **23**, 5713–5728.
- 35 B. A. Vastine and M. B. Hall, *Organometallics*, 2008, **27**, 4325–4333.
- 36 F. De Angelis, A. Sgamellotti and N. Re, *Organometallics*, 2002, **21**, 5944–5950.
- 37 F. De Angelis, A. Sgamellotti and N. Re, *Dalton Trans.*, 2004, 3225, DOI: 10.1039/b408452g.
- 38 O. J. S. Pickup, I. Khazal, E. J. Smith, A. C. Whitwood, J. M. Lynam, K. Bolaky, T. C. King, B. W. Rawe and N. Fey, *Organometallics*, 2014, **33**, 1751–1761.
- 39 M. R. Gyton, T. M. Hood and A. B. Chaplin, *Dalton Trans.*, 2019, **48**, 2877–2880.
- 40 M. J. Frisch, *et al.*, *Gaussian 09*, Gaussian Inc., Wallingford, CT, 2009 (see the full reference in ESI†).
- 41 D. Andrae, U. Häussermann, M. Dolg, H. Stoll and H. Preuss, *Theor. Chim. Acta*, 1990, **77**, 123–141.
- 42 P. C. Hariharan and J. A. Pople, *Theor. Chim. Acta*, 1973, **28**, 213–222.
- 43 J. P. Perdew, *Phys. Rev. B: Condens. Matter Mater. Phys.*, 1986, **33**, 8822–8824.
- 44 A. D. Becke, *Phys. Rev. A: At., Mol., Opt. Phys.*, 1988, **38**, 3098–3100.
- 45 J. Tomasi, B. Mennucci and R. Cammi, *Chem. Rev.*, 2005, **105**, 2999–3094.
- 46 S. Grimme, S. Ehrlich and L. Goerigk, *J. Comput. Chem.*, 2011, **32**, 1456–1465.
- 47 *AIMAll (Version 13.02.26)*, TK Gristmill Software, Overland Park KS, USA, 2015.
- 48 A. E. Reed, J. E. Carpenter and F. Weinhold, *NBO Version 3.1*.
- 49 For Pathway **III**, no intermediate with bridging 'Bu or Ar' between the Rh and C<sub>β</sub> centres was characterised. For R = 'Bu, the indirect transfer of 'Bu led to the formation of the Rh-'Bu complex and for R = Ar', the indirect transfer of Ar' group results in the formation of the Rh-vinylidene complex (see ESI for more details).†
- 50 The computed geometry features only one imaginary frequency.
- 51 N. A. Rajabi, PhD thesis, Heriot-Watt University, 2018.
- 52 N. M. Kostic and R. F. Fenske, *Organometallics*, 1982, **1**, 974–982.
- 53 S. M. Mansell, *Dalton Trans.*, 2017, **46**, 15157–15174.