

Self-assembled nanoparticles based on amphiphilic chitosan derivative and arginine for oral curcumin delivery

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Abstract: Curcumin (Cur) is a striking anticancer agent, but its low aqueous solubility, poor absorption, hasty metabolism, and elimination limit its oral bioavailability and consequently hinder its development as a drug. To redress these limitations, amphiphilic chitosan (CS) conjugate with improved mucoadhesion and solubility over a wider pH range was developed by modification with hydrophobic acrylonitrile (AN) and hydrophilic arginine (Arg); the synthesized conjugate (AN-CS-Arg), which was well characterized by Fourier transform infrared and ¹H nuclear magnetic resonance spectroscopy. Results of critical aggregation concentration revealed that the AN-CS-Arg conjugate had low critical aggregation concentration and was prone to form self-assembled nanoparticles (NPs) in aqueous medium. Cur-encapsulated AN-CS-Arg NPs (AN-CS-Arg/Cur NPs) were developed by a simple sonication method and characterized for the physicochemical parameters such as zeta potential, particle size, and drug encapsulation. The results showed that zeta potential of the prepared NPs was 40.1±2.81 mV and the average size was ~218 nm. A considerable improvement in the aqueous solubility of Cur was observed after encapsulation into AN-CS-Arg/Cur NPs. With the increase in Cur concentration, loading efficiency increased but encapsulation efficiency decreased. The in vitro release profile exhibited sustained release pattern from the AN-CS-Arg/Cur NPs in typical biological buffers. The ex vivo mucoadhesion study revealed that AN-CS-Arg/Cur NPs had greater mucoadhesion than the control CS NPs. Compared with free Cur solution, AN-CS-Arg/Cur NPs showed stronger dose-dependent cytotoxicity against HT-29 cells. In addition, it was observed that cell uptake of AN-CS-Arg/Cur NPs was much higher compared with free Cur. Furthermore, the in vivo pharmacokinetic results in rats demonstrated that the AN-CS-Arg/Cur NPs could remarkably improve the oral bioavailability of Cur. Therefore, the developed AN-CS-Arg/Cur NPs might be a promising nano-candidate for oral delivery of Cur.

Keywords: curcumin, self-assembled, nanoparticles, cytotoxicity, cell uptake studies, oral bioavailability

Introduction

Cancer is the multifaceted and most distressing disease which represents one of the leading causes of death worldwide. Statistics from the World Health Organization reports 8.2 million cancer deaths in the year 2012 and the number of patients is increasing every year.¹ So far, the most common approach for the treatment of cancer is the removal of tumor followed by extensive chemotherapy, radiotherapy, and immunotherapy. Chemotherapy remains the mainstay of treatment, and most of the chemotherapeutic drugs are administered parenterally, causing severe adverse effects due to impulse concentrations in blood stream leading to poor patients' compliance and

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ultimately to failure of therapy.²⁻⁶ Oral delivery of anticancer agents could be a feasible alternative to parenteral administration, since it can avoid fluctuation of drug concentration, minimize drug-related adverse effects, and thus improve the overall efficacy of drugs; moreover, patients are more receptive to it.⁷⁻⁹ However, the oral delivery of anticancer agents is often associated with low bioavailability due to the first-pass effect and physical absorption barriers in gastrointestinal tract (GIT).^{10,11} Therefore, it has increasingly become an urgent medical need to develop a safer, targeted, and better oral anticancer drug delivery system.

Curcumin (Cur), a naturally active component extracted from rhizomes of turmeric (*Curcuma longa*), has a well-established history of utilization in the People's Republic of China and India not only as additive in foods but also as a natural remedy in medicines. Moreover, it possesses diverse biological and pharmacological activities, including antioxidant, anti-inflammatory, anticancer, and antidepressant properties.¹²⁻¹⁴ It is worth notable that no severe adverse drug reaction associated with Cur has been reported so far in conducted clinical trials and research signifying Cur as safest anticancer drug.¹⁵ Currently, the US National Cancer Institute has considered Cur as a third-generation chemotherapeutic drug.¹⁶ Unfortunately, the clinical progression of Cur has been hindered due to several challenges. For example, the low aqueous solubility of Cur leads to the limited bioavailability and poor in vivo pharmacokinetics, which subsequently restricts the therapeutic effect.¹⁷ Another drawback is the insufficient systemic absorption through the digestive tract after oral administration. The improvement in Cur's oral bioavailability has become a medical and pharmaceutical need as major portion of the drug never reaches the plasma to exert its pharmacological effects. Plasma concentrations as well as therapeutic and toxic effects are directly influenced by the extent of bioavailability, resulting after oral Cur administration. Hence, it is therapeutically important to improve the oral bioavailability of Cur.¹⁸ Therefore; efforts are being made for the exploration of novel formulations to circumvent these pitfalls and to utilize the full potential of Cur as anticancer agent.^{19,20}

Till date, various Cur delivery systems have been investigated, including polymeric micelles,²¹ solid lipid nanoparticle (NP),²² polymeric NPs,²³ biodegradable microspheres,²⁴ phospholipids,²⁵ cyclodextrin,²⁶ hydrogel,²⁷ and liposome.^{28,29} Among these delivery systems, polymeric NPs, due to their ultra-small size, possess the added advantage that they can pass through the smallest capillary vessels and avoid hasty clearance by phagocytes, and thus their duration is greatly

extended in blood stream.³⁰ Furthermore, these NPs can be utilized for targeted delivery of drugs to improve bioavailability, to sustain drug in target tissues, and to improve the stability of therapeutic agents.^{31,32} Properties of the NPs are largely dependent on the polymer employed.³³

Among these polymeric NPs, chitosan (CS)-based nano-carriers have received much attention as a drug delivery system.³⁴ CS has been well reported to be the most widespread biopolymer having nontoxic, biocompatible, and biodegradable characteristics.³⁵ In addition, the most common approach for oral Cur delivery is the encapsulation of drug using mucoadhesive polymers such as CS by which the drug is physically protected from enzymatic degradation and intestinal juices. Besides, mucoadhesive in nature CS can enhance the penetration of drugs across mucosal barrier by its interactions with epithelial tight junctions, thereby facilitating transport of drugs across mucosal membrane and improving pharmacological effects.^{36,37} Although the CS possess impressive array of biomedical benefits, the effectiveness of CS application has been limited due to its drawback of poor aqueous solubility at physiological pH, which can be overcome by chemical modification.^{38,39} After structural manipulation and functionalization of CS, drug molecule can be encapsulated and the advantages of encapsulating Cur within the CS polymer allow for a greater control of the active drug pharmacokinetic behavior.¹⁸

Amino acids, the building blocks of proteins, are considered as chiral molecules having low molecular weight and different properties based on their side chains. The inclusion of amino acid moieties to the backbone of CS has been shown to improve the physicochemical properties of CS such as solubility and mucoadhesiveness and gives rise to some interesting synergistic characteristics for use in drug delivery and tissue engineering, as well as other potentially useful biomedical properties such as anticancer, anticoagulant, antimicrobial, and cholesterol-lowering activities.⁴⁰ Gao et al⁴¹ reported the synthesis of conjugate based on arginine (Arg)-grafted CS and observed a great improvement in the water solubility of CS. As an amino acid with cationic surface charge, the Arg has been reported to promote the internalization ability of Arg-grafted CS NPs through the electrostatic interaction between positively surface-charged NPs and negatively charged cell membrane.⁴² Park et al⁴³ demonstrated that the conjugation of CS with Arg enhanced small interfering RNA delivery to cancer cells. In another published report, Zhu et al⁴⁴ showed that Arg-modified CS improved gene transfection in HeLa cells by 100-fold compared with the unmodified CS.

Self-assembling block copolymers or hydrophobically modified polymers have been extensively investigated in the field of biotechnology and pharmaceuticals.⁴⁵ Once hydrophobic segments are attached to the backbone of CS, the received amphiphilic conjugate could form self-assembled NPs by undergoing intermolecular association to encapsulate hydrophobic drugs and deliver drugs to the targeted sites of tumor via passive targeting.⁴⁶ Amphiphilic polysaccharide derivatives are potential drug carriers because of their good cytocompatibility and biodegradability.⁴⁷

Previous literature has proved that NPs based on acrylic acid derivatives such as poly(butyl cyanoacrylate) (PBCA) and monoacrylate are ideal carriers for passive drug targeting and controlled drug delivery due to their properties such as low toxicity, biodegradability, ease of synthesis, and purification.^{48–52} These NPs had been utilized efficiently as controlled drug delivery system for anticancer drugs, peptides, antibiotics, and analgesics.^{53–56} Sun et al⁵⁷ demonstrated that Cur-loaded PBCA NPs result in enhanced accumulation in mice brain. In another study, Bisht et al⁵² reported monoacrylate-based Cur NPs for superior tumor regression compared with free Cur. Similarly, Duan et al⁵⁸ reported that PBCA coated with Cur-loaded CS improves the solubility and anticancer potential of Cur against human hepatocellular carcinoma cells. Therefore, it is important and worth investigation to further explore acrylic acid-based derivatives as novel Cur carriers.

The aim of this research was to study the fabrication of self-assembled NPs based on acrylonitrile (AN) and Arg-modified CS (AN–CS–Arg) conjugate. We hypothesized that AN–CS–Arg conjugate could self-assemble to form NPs when dissolved in water and the hydrophobic Cur could be entrapped within the NPs that might lead to stabilize Cur in aqueous media. Moreover, the prolonged GIT passage at colonic mucosa offered by mucoadhesive drug delivery system will aid in sustained release of Cur within the colon. Cur-loaded AN- and Arg-modified CS NPs (AN–CS–Arg/Cur NPs) were developed by simple sonication method and evaluated for their potential as delivery system for the water insoluble drug Cur. Meanwhile, NPs were characterized for their key features such as size and morphology, zeta potential and drug-loading and encapsulation efficiency (EE). In vitro mucoadhesions were investigated. In vitro release study was carried out to study the release behavior of Cur. Furthermore, the cytotoxicity and cellular uptake in tumor cells, pharmacokinetics, and in vivo oral bioavailability of AN–CS–Arg/Cur NPs in rats were also investigated.

Materials and methods

Materials

CS ($M_n=50$ kDa, deacetylation degree of 90%) was obtained from Laizhou Haili Biological Product Co. Ltd. (Shandong, People's Republic of China). Arg, ethyl-3-(3-dimethylaminopropyl) carbodiimide hydrochloride, *N*-hydroxysuccinimide, and sodium tripolyphosphate were procured from Sigma-Aldrich Co. (St Louis, MO, USA). AN (purity >99%) was purchased from Wako Pure Chemical Industries, Ltd. (Osaka, Japan). Cur was purchased from Aladdin Chemistry Co., Ltd. (Shanghai, People's Republic of China). Pyrene was purchased from Sigma-Aldrich Co. and purified by double recrystallization from absolute ethanol. Human colorectal adenocarcinoma (HT-29) cell line was received as gift sample from Qingdao University Medical College (Qingdao, People's Republic of China). RPMI-1640 and calf serum were purchased from Hyclone Company (Utah, USA). 3-(4, 5-dimethylthiazol-2-yl)-2, 5-diphenyltetrazolium bromide (MTT) was obtained from Biomol Company (Hamburg, Germany). Dimethyl sulfoxide was received from EMD Millipore (Billerica, MA, USA). Injections of streptomycin sulfate (1 million units) and injection of sodium penicillin (800,000 units) were purchased from Shandong Lukang Pharmaceutical Co., Ltd. (Shandong, People's Republic of China). All other reagents used were of analytical grade or purer, and water was double distilled.

Synthesis of AN- and Arg-modified CS conjugate

First, CS was modified with AN (AN–CS) according to the previously published report.⁵⁹ Briefly, CS was dissolved in (1 wt%) acetic acid to form homogenous aqueous solution, afterward filtered by filter paper. The solution was neutralized with NaOH (1 M) solution with continuous stirring and the fine powder obtained was collected. Then, CS powder (0.5 g) was taken in a 100 mL round bottom flask and thoroughly mixed with AN (30 mL), and 1 mL of aqueous NaOH solution (0.6%) was mixed dropwise under magnetic stirring. This reaction was conducted at 77°C for 24 hours. The obtained AN–CS was rinsed with diethyl ether and acetone.

Then, Arg was grafted on the AN–CS according to a modified method.⁶⁰ Briefly, BOC-Arg(Pbf)-OH (1.05 g) and 1-ethyl-3-(3-dimethylaminopropyl) carbodiimide hydrochloride (EDC.HCl) (0.38 g) were dissolved completely in *N,N*-dimethylformamide (DMF) and stirred magnetically for 2 hours. *N*-hydroxysuccinimide (0.24 g) dissolved in DMF was added dropwise with continuous stirring. After stirring for another 4 hours, the resultant mixture was added into

AN-CS (0.16 g) dissolved in DMF. The solution obtained was continuously stirred at room temperature for 48 hours. Trifluoroacetic acid (1:1 v/v) was mixed into the reaction system for deprotection. Finally, the obtained solution was concentrated and precipitated in diethyl ether and dichloromethane. The yellowish powder of AN-CS-Arg was obtained and the "degree of substitution" (DS) of AN and Arg on the backbone of CS was calculated from ^1H nuclear magnetic resonance ($^1\text{HNMR}$) spectrum.

Physicochemical characterizations of AN- and Arg-modified CS conjugate

To demonstrate the conjugation of AN and Arg groups to the CS, Fourier transform infrared (FTIR) and $^1\text{HNMR}$ spectroscopy methods were used. The FTIR spectra of CS and AN-CS-Arg were recorded on an FTIR Spectrophotometer-430 (Jasco Company, Tokyo, Japan) following the method of Shigemasa et al.⁶¹ Approximately 2 mg of the samples was mixed with 100 mg of KBr and made into pellets for spectroscopic analysis at 20°C. All spectra were measured over 400–4,000 cm^{-1} .

$^1\text{HNMR}$ spectra of CS and AN-CS-Arg were recorded on a $^1\text{HNMR}$ Spectrometer (Bruker AV 600 MHz; Bruker Optik GmbH, Ettlingen, Germany) using 5 mm $^1\text{HNMR}$ tube at 25°C. Then, 5 mg samples were dissolved in 1 mL D_2O solution and ultrasonically mixed to dissolve completely, yielding a final concentration of 5 mg/mL. $^1\text{HNMR}$ spectra were measured at 298 K with 80 scans, a spectral width of 4,800 Hz, and a relaxation delay of 1 second between scans and acquisition time of 3.75 seconds.⁶²

Solubility test of CS and modified conjugate

Solubility of CS and modified conjugate was estimated from turbidity.⁶³ The pH-reliant water solubility of CS and conjugate was assessed at room temperature over a pH range of 2–11. Briefly, 10 mg samples of CS and its conjugate were separately dissolved in HCl (0.1 M) solution. Then, pH of solution was allowed to adjust by dropwise addition of NaOH (0.1 M). Finally, transmittance of solutions at 600 nm over selected pH range was observed using ultraviolet (UV)–visible spectrophotometer.

Measurement of fluorescence spectroscopy

Fluorescence measurements were studied to determine the critical aggregation concentration (CAC) as described by Liu et al.⁶⁴ Pyrene is one of the fluorescent probes that are widely used in the determination of CAC of surfactants and

polymers since the relative intensity of the first and third emission peaks (which are denoted as I_1 and I_3 at 373 nm and 393 nm, respectively) is a good marker of environment polarity.⁶⁵ At a concentration of 0.4 mg/mL, purified pyrene was dissolved in ethanol. Approximately 20 μL of the resulting solution was added into a 20 mL test tube and ethanol was driven off under the stream of nitrogen gas. Then, 10 mL of AN-CS-Arg NPs solution was subsequently added to the test tube, bringing the final concentration of pyrene to 2 μM . The mixture was incubated for 3 hours in a water bath at 65°C and shaken in a SHA-B shaking water bath (Guo Hua Company, Jintan, Jiangsu, People's Republic of China) overnight at 20°C. Using a RF-5301PC Fluorescence Spectrophotometer (Shimadzu, Kyoto, Japan), pyrene emission spectra were obtained. Both emission and excitation slits were set to 10 nm and 2.5 nm, respectively.⁶⁶ Plot of I_1/I_3 vs log concentration was drawn, and the concentration at which sudden increase in I_1/I_3 had been occurred was considered as CAC value.

Preparation of Cur-loaded AN- and Arg-modified CS NPs

Cur-loaded AN- and Arg-modified CS NP (AN-CS-Arg/Cur NPs) solution was synthesized using sonication method, similar to that employed by Lee et al.⁶⁷ First, 10 mg/mL of Cur ethanol solution was prepared. Then, AN-CS-Arg conjugate (20 mg) was dissolved into 10 mL distilled water and ultrasonically mixed. To this solution, Cur ethanol solution was added, placed in ice bath, and sonicated at 90 W for 2 minutes. This was repeated three times to make AN-CS-Arg/Cur NPs solution. In the process of this fabrication, Cur occupied the hydrophobic cores of NPs to form AN-CS-Arg/Cur NPs. The obtained solution was then transferred into a pre-swollen membrane with a molecular weight cutoff of 1,200 kDa and dialyzed against 40 mL of 5% ethanol for 48 hours during which the dialysate was changed three times to remove the unloaded Cur, ethanol, and other impurities, and thereafter, lyophilized and stored at 4°C for further use. Similarly, blank NPs without Cur were prepared, keeping the rest of the method the same. Figure 1 shows schematic drawing of self-assembled AN-CS-Arg/Cur NPs. The hydrophobic Cur occupies the hydrophobic core while Arg forms the outer hydrophilic shell.

Particle size and morphology

AN-CS-Arg/Cur NPs were characterized for size and morphology using an EM 400 transmission electron microscopy (TEM; Philips, Eindhoven, the Netherlands) operating at accelerating voltage of 80 kV. Five to ten drops of sample

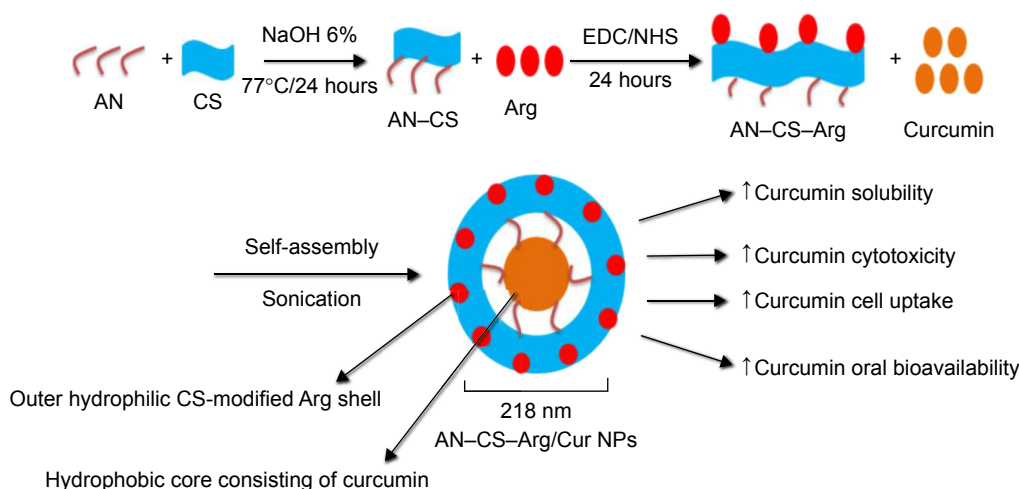


Figure 1 Schematic drawing of self-assembled Cur-encapsulated AN-CS-Arg NPs.

Abbreviations: AN, acrylonitrile; Arg, arginine; CS, chitosan; Cur, curcumin; EDC, ethyl-3-(3-dimethylaminopropyl) carbodiimide hydrochloride; NHS, N-hydroxysuccinimide; NPs, nanoparticles.

solution were carefully placed onto a carbon-coated copper grid, followed by the removal of excess solution by blotting the grid with filter paper. The samples were dried for 72 hours at room temperature. Afterward, the samples were stained negatively using phosphotungstic acid (2%) and allowed to dry further for 72 hours before final observation.

Particle size and zeta potential

The size distribution, mean particle size, and polydispersity index (PDI) of the AN-CS-Arg/Cur NPs were determined at room temperature by dynamic light scattering using a nano-ZS Zetasizer (Malvern Instruments, Malvern, UK). The concentration of the NPs solution was kept at 1 mg/mL and measurements were performed three times. All measurements were performed at the wavelength of 635 nm with a fixed scattering angle of 90°.

Zeta potential deals with the stability of formulations and is an indicator for charge present on the surface of NPs. A nano-ZS Zetasizer was used to measure the zeta potential of NPs. Each sample was analyzed in triplicate at 25°C, and the average of three readings was reported.

EE and LE of Cur-loaded AN- and Arg-modified CS NPs

The EE and loading efficiency (LE) of the synthesized NPs were determined by measuring the concentration of free Cur in the dispersion medium by centrifugation at 16,000×g for 20 minutes. The supernatant obtained was diluted with 90% ethanol, and the drug concentration in the resultant solution was measured at 420 nm by high-performance liquid chromatography (HPLC; Agilent Technologies, Santa Clara, CA, USA) with a standard curve. The HPLC analysis was carried

out on a reversed phase C18 column (4.6×150 mm, 5 μm, ZORBAX Eclipse XDB-C18) with a mobile phase composed of methanol and 0.3% acetic water solution (80:20 v/v) at a flow rate of 1 mL/min. The following formulas were used to calculate the EE and LE.

$$EE = \frac{\text{Total curcumin} - \text{Free curcumin}}{\text{Total curcumin}} \times 100\% \quad (1)$$

$$LE = \frac{\text{Total curcumin} - \text{Free curcumin}}{\text{Weight of freeze-dried nanoparticles}} \times 100\% \quad (2)$$

where the free Cur is the analyzed weight of drug in supernatant, total Cur is the feeding weight of drug in the preparation of NPs, and the weight of freeze-dried NPs is the total weight of carrier.

In vitro drug release study

The release property of Cur from AN-CS-Arg/Cur NPs was assessed by dialysis. Free Cur solution (5 mg/mL dissolved in 90% ethanol) with the same concentration was considered as control. Solutions of 2.5 mL AN-CS-Arg/Cur NPs or free Cur with Cur concentration of 0.2 mg/mL were transferred into dialysis bags (molecular weight cutoff, 6–8 kDa) and dialyzed against 10 mL release medium at different pH values (0.05 M glycine-hydrochloric acid buffer at pH 2.0, 0.05 M phosphate buffer at pH 5.0, and 0.05 M phosphate buffer at pH 6.8). To prevent the degradation of the released Cur, ascorbic acid at a concentration of 1% and butylated hydroxytoluene at a concentration of 0.1% were added into the buffer solutions.⁶⁸ Then, the release medium was continuously stirred at 37°C for 72 hours. At predetermined interval,

3.0 mL of sample was drawn out and replenished with 3.0 mL fresh buffer. The amount of released Cur was measured by the curve of concentration values obtained using HPLC method as described earlier and the following formula. All the release experiments were performed three times.

$$\text{Release (\%)} = \frac{\text{Released curcumin}}{\text{Total curcumin}} \times 100 \quad (3)$$

Ex vivo mucoadhesion measurements

All animal experiments were performed in accordance with the European community council directive of September 22, 2010 (2010/63/EU). The mucoadhesion experiment was conducted according to the previously published method.⁶⁹ Male Sprague Dawley rats (180–220 gm) were purchased from Qingdao Coastal Institute for Drug Control, Qingdao, People's Republic of China. After obtaining ethical approval from the Animal Care and Use Committee of the Ocean University of China, the rats were sacrificed in CO₂ chamber. The excised colon from rats was washed thoroughly with ice-cold phosphate-buffered saline to remove all its contents. Approximately 6 cm colon segments were cut open and everted with the help of tweezers and glass rod. One end was closed with closure, and the sac was filled through the other open end with 1 mL medium Dulbecco's Modified Eagle Medium: Nutrient Mixture F-12 (DMEM: F12). Afterward, the open end was also closed and kept at 4°C for further use. The sac was then suspended in a centrifuge tube (15 mL) containing 5 mg of the synthesized NP suspension dispersed in 5 mL of the same medium. Thereafter, the tube was gently agitated at 37°C for 60 minutes to assure the attachment equilibrium. Then, the tube was removed and centrifuged for the recovery of NPs that did not attach to colon. The amount obtained was subtracted from the initial amount of NPs added to the system to know the exact amount of NPs attached to the colon. To represent the loss of colonic tissue in experiment, the sac was incubated in NPs-free medium and used as a control. Similar procedure was carried out for CS NPs (prepared by dropwise addition of tripolyphosphate to acidic CS solution at 1:3 w/w), and the average of three results is reported in this article.

In vitro cytotoxicity study

To investigate the in vitro cytotoxicity, human colorectal adenocarcinoma (HT-29) cell lines were obtained from the cell bank of Qingdao University Medical College, Qingdao, People's Republic of China. HT-29 cells were then grown in RPMI-1640 medium supplemented with penicillin and streptomycin (1%) and fetal bovine serum (10%). The cells

were seeded at a density of 3–5×10⁴ cells/well into a 96-well plate and cultured in a humid atmosphere at 37°C having 5% CO₂ gas. After culturing for 24 hours, cytotoxicities of blank AN–CS–Arg NPs and AN–CS–Arg/Cur NPs were evaluated in HT-29 cells by the (MTT) dye assay method.^{70,71} Using microplate reader, the UV absorbance intensity was measured at 490 nm. All the experiments were performed in triplicate, and the cell viability is expressed by the following equation:

$$\text{Cell viability (\%)} = \frac{\text{Abs sample}}{\text{Abs control}} \times 100 \quad (4)$$

where Abs sample is the absorbance intensity of the cells treated with blank AN–CS–Arg NPs, Cur solution, or AN–CS–Arg/Cur NPs and Abs control is the absorbance intensity of untreated cells.

Intracellular uptake study

To examine the cellular uptake of unmodified Cur and AN–CS–Arg/Cur NPs, the fluorescence microscopic study was performed using HT-29 cells. Briefly, HT-29 cells at a density of 2.5×10⁴ cells/well were seeded into a six-well plate. After overnight incubation in a humid environment with 5% CO₂ at 37°C, cells were treated with a medium containing free Cur or AN–CS–Arg/Cur NPs (the concentration of Cur was 0.02 mg/mL). After incubation for 2 hours, the cells were rinsed three times with phosphate-buffered saline at 37°C to eliminate Cur or AN–CS–Arg/Cur NPs that were not taken up by the cells. Then, the cells were imaged under a Nikon ECLIPSE TE 2000-U Fluorescence Microscope (20× objective) equipped with a digital camera and a blue filter (420–490 nm excitation, 490–560 nm emission). Images were captured with NIS Elements F 3.0 software (Nikon Corporation, Tokyo, Japan).

In vivo oral administration and pharmacokinetics

Twelve male Sprague Dawley rats (220–250 g) were purchased from Qingdao Coastal Institute for Drug Control, Qingdao, People's Republic of China and cared in air-conditioned quarters under a photoperiod schedule of 12-hour light/dark cycle. They were allowed to receive laboratory chow and tap water available ad libitum for 7 days. All the animal experiments were performed in accordance with the European Community Council Directive of September 22, 2010 (2010/63/EU). Experimental rats were first anesthetized with 0.5 mL chloral hydrate (10%). Then, divided into two groups; the first group also called the control group (n=6) was given aqueous

solution of free Cur with 10% ethanol mixture; the second group ($n=6$) received AN-CS-Arg/Cur NPs solution. The Cur dosage for both the experimental and control group was 100 mg/kg body weight and administered once by oral gavage, after blood samples (0.3 mL) from the jugular vein were collected into microcentrifuge tubes containing 200 μ L of ethylenediaminetetraacetic acid-K2 solution and replenished with 400 μ L heparin saline at fixed time intervals. The plasma was collected by centrifuging the blood sample at 4,000 rpm for 20 minutes. For protein precipitation, 200 μ L of 99% ethanol was mixed to 200 μ L of plasma. Subsequently, 10 μ L (1 mg/mL) β -17-estradiol acetate in ethyl acetate was added as the internal standard. Then, extraction with ethyl acetate was carried out for three times and redissolved in H₂O:acetonitrile (1:2 v/v) mixture. The samples were vortexed and centrifuged at 1,000 \times *g* for 30 minutes. Then, 200 μ L of supernatants was separated and processed for pharmacokinetic analysis using reversed-phase (RP)-HPLC as described earlier with the help of statistics calculated on DAS 2.1.1 software designed by Chinese Pharmacological Society.

Statistical analysis

All the results were expressed using Microsoft Excel 2013, Graph pad Prism 6.01, and DAS 2.1.1 softwares, and the chemical structure was processed by ChemDraw 7.0. The assays were formed in triplicate on separate occasions.

Results were expressed as mean value \pm standard deviation. Differences were considered to be statistically significant when the *P*-values were <0.05 .

Results and discussion

Synthesis and structural characterization of AN- and Arg-modified CS conjugate

The synthetic route of AN-CS-Arg was shown in figure 2. Both the groups were conjugated on the backbone of CS. The hydroxyl group present in CS was reacted with carbon-carbon double bond (HC=CH) in AN. The protected Arg was grafted on CS via amide bond formation, and the deprotection was carried out to get AN-CS-Arg. This modification had many advantages. The first advantage was the improvement in the solubility of CS after AN modification, the AN-modified CS was soluble in organic solvents, which was feasible for the following step of Arg grafting. Moreover, the modified CS was amphiphilic in behavior, which could undergo through self-assembly to form NPs. The second advantage was the improvement in the transmembrane ability of the modified CS NPs, as Arg has been reported to be the main component in cell-penetrating peptides.

FTIR spectra of CS and modified CS conjugate are shown in Figure 3. It was clear that a characteristic strong peak of C \equiv N emerged at 2,257.39 cm^{-1} . It showed that AN was successfully grafted on the backbone of CS. The second new

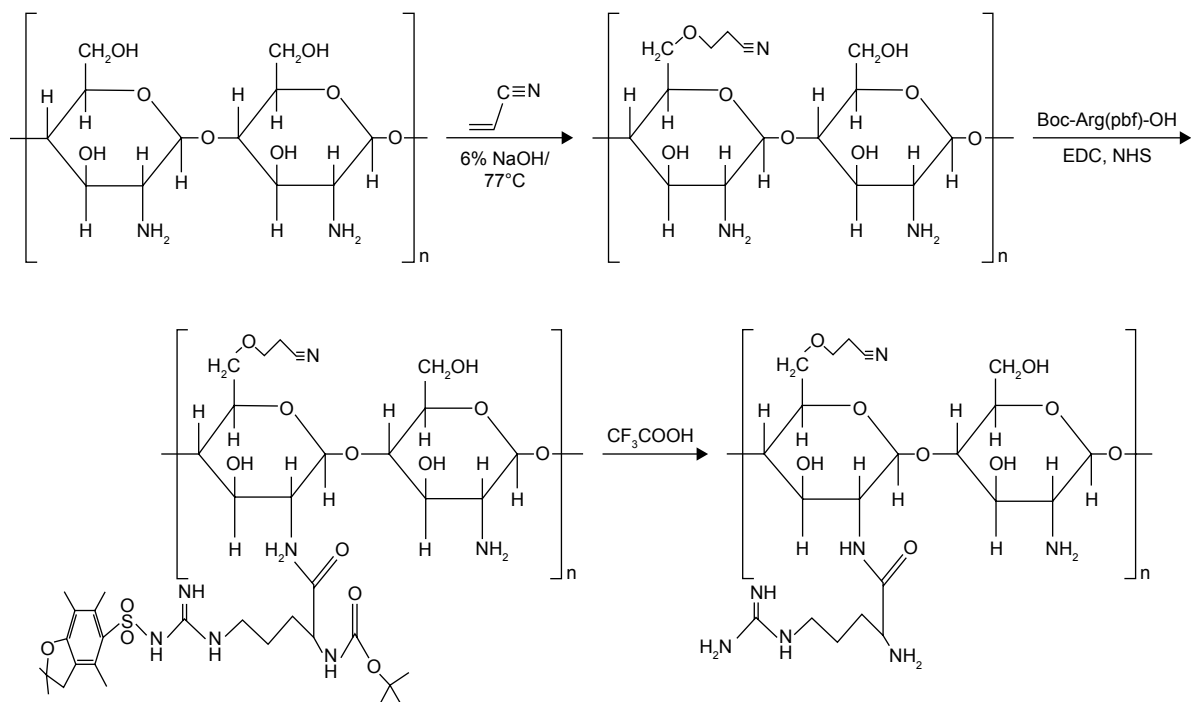


Figure 2 Process and structure of AN- and Arg-modified CS (AN-CS-Arg) synthesis.

Abbreviations: AN, acrylonitrile; Arg, arginine; CS, chitosan; EDC, ethyl-3-(3-dimethylaminopropyl) carbodiimide hydrochloride; NHS, N-hydroxysuccinimide.

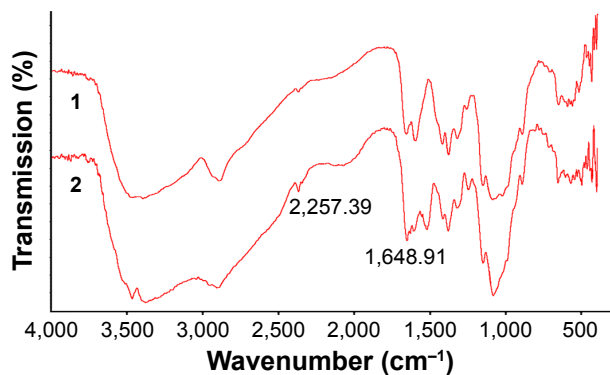


Figure 3 FTIR spectra of (1) CS (2) AN- and Arg-modified CS (AN-CS-Arg) conjugate.

Abbreviations: AN, acrylonitrile; Arg, arginine; CS, chitosan; FTIR, Fourier transform infrared.

and strong peak at $1,648.91\text{ cm}^{-1}$ could be attributed to the amide bond formation between CS and Arg.

The successful incorporation of the AN and Arg groups was further ascertained by $^1\text{H NMR}$ assay of CS and AN-CS-Arg conjugate (Figure 4). The $^1\text{H NMR}$ spectrum of CS is illustrated in the inset picture of Figure 4, in which the concentrated chemical shift from 2.2 ppm to 4.7 ppm was due to the chemical environment of the protons in CS chains. Figure 4 depicts the $^1\text{H NMR}$ spectrum of AN-CS-Arg

conjugate in which the additional peaks from $\sim 1.1\text{ ppm}$ to 2.0 ppm clearly indicated the conjugation of AN and Arg groups to the backbone of CS. The characteristic protons of e and f (Figure 4) in AN moiety and a and d in Arg appeared at $\sim 1.6\text{ ppm}$ and characteristic protons in CH_2CH_2 (b and c) unit were observed at $\sim 1.1\text{ ppm}$ and 1.2 ppm , respectively. The DS of AN and Arg on CS chains as calculated via the calculation from $^1\text{H NMR}$ spectra was 2.5% and 2.8%, respectively. The DS was calculated by using the following formula:

DS =

$$\frac{\text{Integrity of protons in } \text{CH}_2\text{CH}_2 \text{ (b and c) unit/4}}{\text{Integrity of protons in CS/7}} \times 100\%$$

(5)

On the basis of FTIR and $^1\text{H NMR}$ spectra of CS and AN-CS-Arg samples, both AN and Arg have been successfully grafted on CS main chains.

Solubility of CS and modified conjugate at pH 2–11

The solubility of CS and modified conjugate measured by transmittance as function of pH values in aqueous solution is

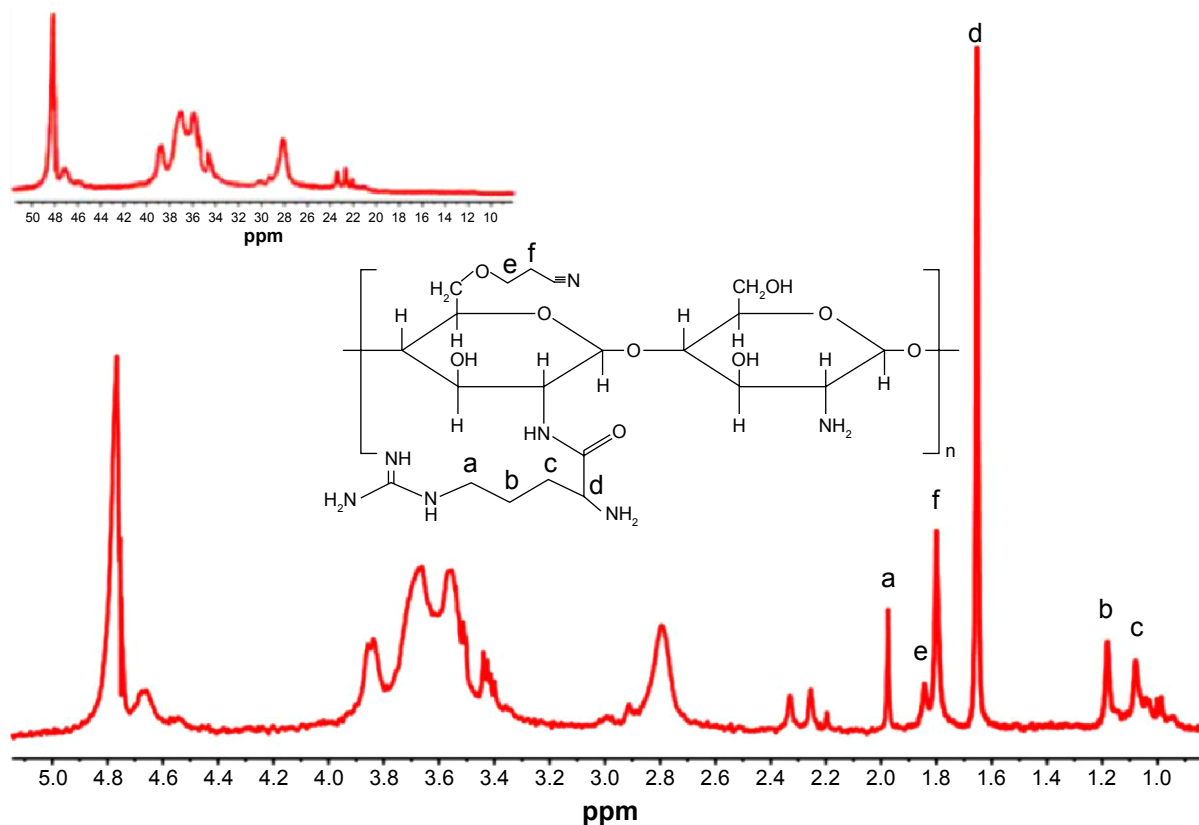


Figure 4 $^1\text{H NMR}$ spectrum of AN-CS-Arg conjugate (inset: $^1\text{H NMR}$ spectrum of CS).

Notes: a, b, c and d represent characteristic protons in Arg moiety; e and f represent characteristic protons in AN moiety.

Abbreviations: AN, acrylonitrile; Arg, arginine; CS, chitosan; $^1\text{H NMR}$, ^1H nuclear magnetic resonance; ppm, parts per million.

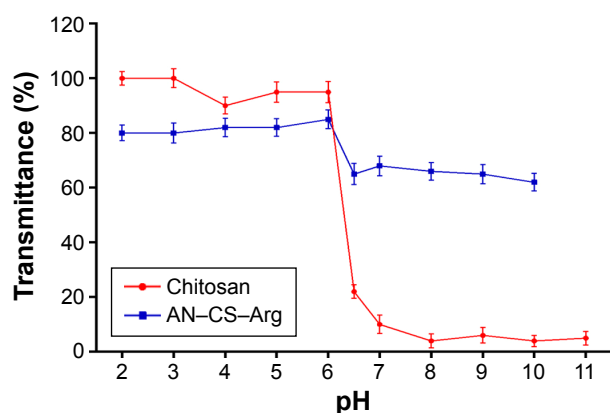


Figure 5 Solubility of CS and AN-CS-Arg conjugate.

Note: Data are expressed as mean \pm SD (n=3).

Abbreviations: AN, acrylonitrile; Arg, arginine; CS, chitosan; SD, standard deviation.

shown in Figure 5. The turbidity of CS solution as measured by relative transmittance was remarkably decreased from its original value, 100% to 90% at pH 2–6 downward to 22% to 2.9% at pH 6.5–11. On the contrary, the modified CS showed an improved solubility over the entire tested pH range. This result is consistent with the previous published reports.^{72,73}

Self-aggregates of AN- and Arg-modified CS in solution

Fluorescence spectroscopy was used to determine the rheology of AN-CS-Arg in aqueous media. The variation of fluorescence intensity ratio for peak I/III against the logarithm of AN-CS-Arg concentration was shown in Figure 6. At low concentrations (0.0001–0.01 mg/mL), the intensity ratios of peak I/III remain fairly constant. However, as AN-CS-Arg concentration increased from 0.075 mg/mL to 1 mg/mL, the intensity ratio decreased significantly. The concentration corresponding to the breakpoint shown in Figure 6 was the

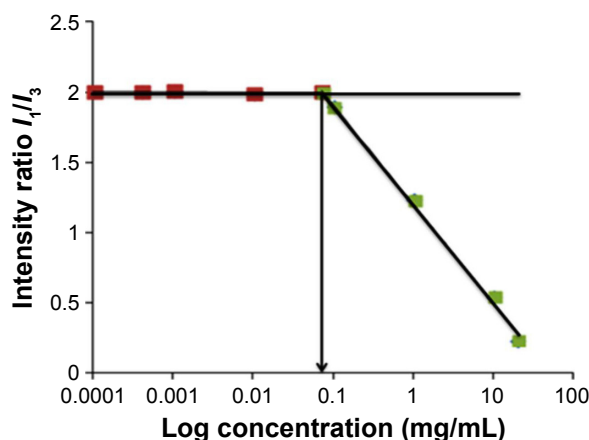


Figure 6 The variation of fluorescence intensity ratio I_1/I_3 against the logarithm of AN-CS-Arg conjugate concentration in distilled water.

Abbreviations: AN, acrylonitrile; Arg, arginine; CS, chitosan.

CAC value of AN-CS-Arg, which was \sim 0.081 mg/mL. The low CAC value indicated that AN-CS-Arg NPs could be formed at a low copolymer concentration.

Characterization of self-assembled Cur-loaded AN- and Arg-modified CS NPs

Morphology, particle size, PDI, and zeta potential

The size and morphology of the prepared NPs were confirmed by TEM. Figure 7 shows the TEM image of AN-CS-Arg/Cur NPs, indicating spherical particles within an average size range of 218 nm.

Particle size is a very important parameter for drug delivery carriers. It has been reported that particles smaller than 1 μ m can undergo capillary distribution and uniform perfusion at the desired target site.⁷⁴ Most solid tumors have high levels of vascular permeability. Particles <400 nm can cross vascular endothelia and accumulate at the tumor site via the enhanced permeability and retention (EPR) effect.^{75,76} Figure 8A reveals the size distribution of AN-CS-Arg/Cur NPs solution prepared under sonication. The particle size of AN-CS-Arg/Cur NPs obtained under ultrasound condition was found to be 218 nm and PDI was 0.325 ± 0.027 which indicated a narrow particle size distribution with better effect.

The zeta potential of the NPs as measured and graphed in Figure 8B was 40.1 ± 2.81 mV, which should be ascribed to the positive charge, arisen from CS units. The size and zeta potential of NPs are critical parameters for cellular uptake. It has been reported that positively charged surfaces facilitate adhesion of the NPs to negatively charged cellular membranes, which may improve their potential as drug delivery

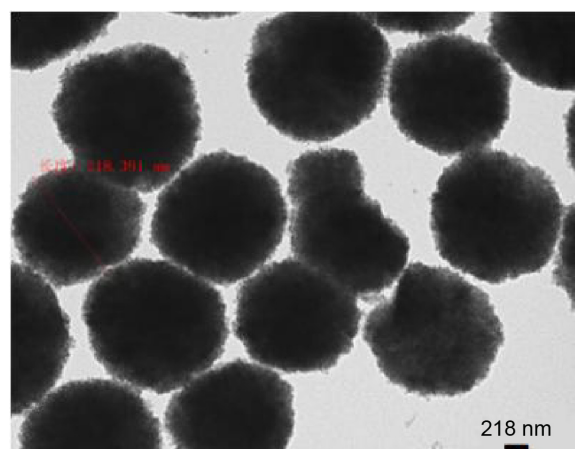


Figure 7 TEM image of AN-CS-Arg/Cur NPs solution.

Note: TEM image of AN-CS-Arg/Cur NPs solution at magnification (100,000 \times), scale bar represent 218 nm.

Abbreviations: AN, acrylonitrile; Arg, arginine; CS, chitosan; Cur, curcumin; NPs, nanoparticles; TEM, transmission electron microscopy.

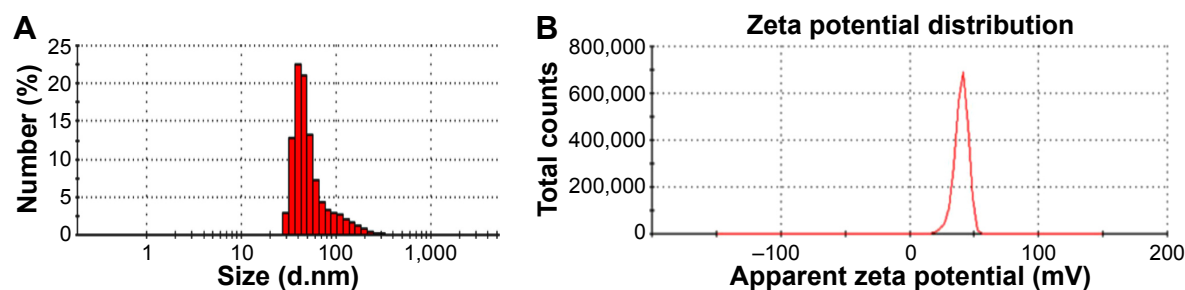


Figure 8 (A) Size distribution of 1 mg/mL AN-CS-Arg/Cur NPs solution. (B) Zeta potential of 1 mg/mL AN-CS-Arg/Cur NPs solution. **Abbreviations:** AN, acrylonitrile; Arg, arginine; CS, chitosan; Cur, curcumin; NPs, nanoparticles.

carriers.⁷⁷ The prepared NPs possess positive surface charge and are suitable to be used as drug delivery carrier.

A comparative effect of Cur concentration on particle size, PDI, and zeta potential is presented in Table 1. With increasing amounts of Cur in the formulation, the particle size was reduced from 218.3 ± 7.2 nm to 185.3 ± 5.9 nm and the values of PDI decreased from 0.325 ± 0.027 to 0.290 ± 0.012 , while no significant effect was observed on zeta potential values. The reduction in size with increasing concentration of Cur might be due to the formation of denser and stronger hydrophobic core. Earlier studies with other types of Cur-encapsulated nanoparticulates demonstrated similar results.⁷⁸

Drug EE and drug loading

AN- and Arg-modified CS NPs (AN-CS-Arg/Cur NPs) were successfully used to encapsulate the hydrophobic Cur inside the hydrophobic core. As the concentration of Cur increased, the LE increased from $6.48\% \pm 0.27\%$ to $27.82\% \pm 2.55\%$, whereas the EE decreased from $76.53\% \pm 3.58\%$ to $66.34\% \pm 3.12\%$ as shown in Table 2.

Furthermore, it is obvious from Table 2 that with increasing concentration of Cur, EE and LE increase but when the concentration becomes >8 mg EE declines. This might be due to the reason that the carrying capacity of AN-CS-Arg conjugate is limited; when the concentration reaches to some

point, then more Cur will be buried into the hydrophobic microdomains of AN-CS-Arg/Cur NPs but the drug-loading amount will increase with the increase in feeding amount of Cur.^{79,80} Formulation 2 was selected in the remainder of experimental studies due to optimized properties.

Entrapment efficiency and LE of Cur in AN-CS-Arg/Cur NPs were found to be $76.53\% \pm 3.58\%$ and $27.82\% \pm 2.55\%$, respectively, as mentioned earlier. These results are consistent with the previous findings in reports^{74-76,81} and also suggest that the concentration of Cur and AN-CS-Arg can be adjusted to give desired entrapment efficiency and LE.

Aqueous solubility study

Cur immediately precipitates in water as practically having an extremely low aqueous solubility, but a significant enhancement in its aqueous solubility was observed after encapsulation into AN-CS-Arg NPs. While maintaining clear and transparent appearance, the solubility of Cur in water was largely enhanced from the original 11 ng/mL⁸² to 0.5 mg/mL after encapsulation into AN-CS-Arg NPs, $\sim 5 \times 10^4$ -fold increase. These results indicated that the obtained NPs system could significantly improve the solubility of Cur in aqueous solution, and the images shown in Figure 9 support the fact. The main reason for the low oral bioavailability of Cur is its insolubility in aqueous solutions, and many of researchers are

Table 1 Characterization of AN-CS-Arg/Cur NPs by varying drug concentration

Cur formulation	Cur (mg)	Size (nm)	PDI	Zeta potential (mV)
AN-CS-Arg/Cur NPs 1	5	218.3 ± 7.2	0.325 ± 0.027	40.1 ± 2.81
AN-CS-Arg/Cur NPs 2	8	199.4 ± 6.3	0.310 ± 0.015	36.3 ± 2.65
AN-CS-Arg/Cur NPs 3	10	185.3 ± 5.9	0.290 ± 0.012	39.4 ± 1.52

Note: Data are expressed as mean \pm SD.

Abbreviations: Arg, arginine; AN, acrylonitrile; CS, chitosan; Cur, curcumin; NPs, nanoparticles; PDI, polydispersity index; SD, standard deviation.

Table 2 EE and LE of AN-CS-Arg/Cur NPs with increasing amount of Cur added

Cur formulation	Cur (mg)	EE (%)	LE (%)
AN-CS-Arg/Cur NPs 1	5	57.24 ± 2.73	6.48 ± 0.27
AN-CS-Arg/Cur NPs 2	8	76.53 ± 3.58	14.23 ± 1.65
AN-CS-Arg/Cur NPs 3	10	66.34 ± 3.12	27.82 ± 2.55

Note: Data are expressed as mean \pm SD.

Abbreviations: Arg, arginine; AN, acrylonitrile; CS, chitosan; Cur, curcumin; EE, encapsulation efficiency; LE, loading efficiency; NPs, nanoparticles; SD, standard deviation.

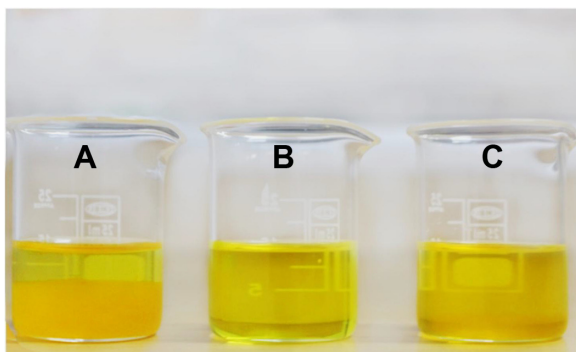


Figure 9 Images of free Cur and AN-CS-Arg/Cur NPs solution.

Notes: (A) Images of free Cur solution; (B) AN-CS-Arg/Cur formulated as lyophilized powder of 0.2 mg/ml Nps solution; and (C) AN-CS-Arg/Cur formulated as lyophilized powder of 0.5 mg/ml Nps solution; in water at pH 7.4, clearly showing the enhanced solubility of Cur, while maintaining clear and transparent appearance but free Cur forms flakes indicating poor solubility.

Abbreviations: AN, acrylonitrile; Arg, arginine; CS, chitosan; Cur, curcumin; NPs, nanoparticles.

interested in how to improve the solubility of Cur. We can say that an increase in Cur solubility from these NPs may be due to the combination of hydrophobic microenvironment of Cur and the solubilizing effect of the carrier.^{58,83,84}

In vitro release study

Cur has been well reported for the hasty elimination in the metabolism process.^{82,85} Nanocarrier-based delivery may provide an alternative approach to shield Cur from metabolic degradation before Cur reaches to its targeted site of action. To achieve this goal, sustained release of Cur is required.⁸⁶

The release profiles of AN-CS-Arg/Cur NPs were obtained in three media at pH 2.0, pH 5.0, and pH 6.8, simulating various physiological environments. The different three pH values were chosen to simulate different

physiological environments, including stomach (pH 1.0–2.5), intestine (pH 6.4–7.5), intracellular components (pH 4.5–8.0), and solid tumor extracellular microenvironment (pH ~6.5). As shown in Figure 10, Cur was released much slower after being encapsulated by AN-CS-Arg NPs compared with free Cur, which reveals the sustained release of AN-CS-Arg/Cur NPs. For example, after 72 hours, 57%, 50%, and 63% of the Cur were released from the AN-CS-Arg NPs at pH 2.0, pH 5.0, and pH 6.8, respectively; by contrast, native Cur had completed its release/diffusion at the three pHs studied. It suggests that the AN-CS-Arg/Cur NPs can be used as sustained formulations applied either in oral administration or in intravenous administration. The release mechanism from the NPs might be related to the diffusion of drug and disintegration of polymer.^{87,88} The drug release process was mainly as follows: first, the media gradually diffused into the interior NPs (hydrophobic core) to dissolve Cur and the dissolved drug spread to the media slowly. Then, the carrier was degraded and the Cur was released with a slow rate.⁸⁹ To investigate the release mechanism of Cur from AN-CS-Arg/Cur NPs, we analyzed regression coefficients with three kinetic models (first-order, zero-order, and Higuchi model). As a result, the Higuchi's square-root plot was best fitted with release kinetic with regression coefficients (R^2) of 0.941. This result further suggests that AN-CS-Arg/Cur NPs exhibit a continuous and slow-release behavior for Cur.

Mucoadhesion

The percentage adhesion of CS NPs and AN-CS-Arg/Cur NPs to the colon of excised rat tissue is presented in Figure 11. As shown in Figure 11, 74.21%±4.33% of the experimental

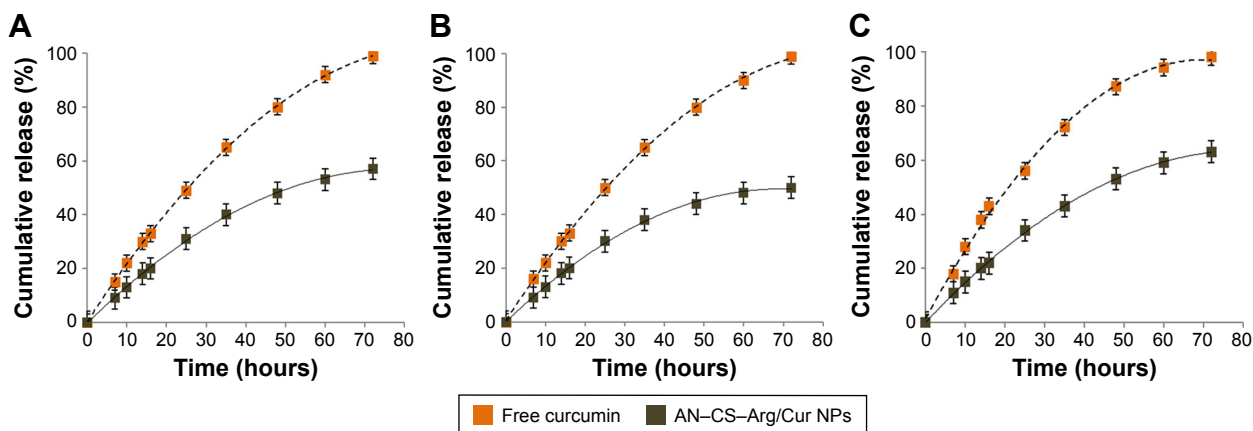


Figure 10 (A) Sustained release pattern of Cur from AN-CS-Arg/Cur NPs at pH 2.0. (B) Sustained release pattern of Cur from AN-CS-Arg/Cur NPs at pH 5.0. (C) Sustained release pattern of Cur from AN-CS-Arg/Cur NPs at pH 6.8.

Abbreviations: AN, acrylonitrile; Arg, arginine; CS, chitosan; Cur, curcumin; NPs, nanoparticles.

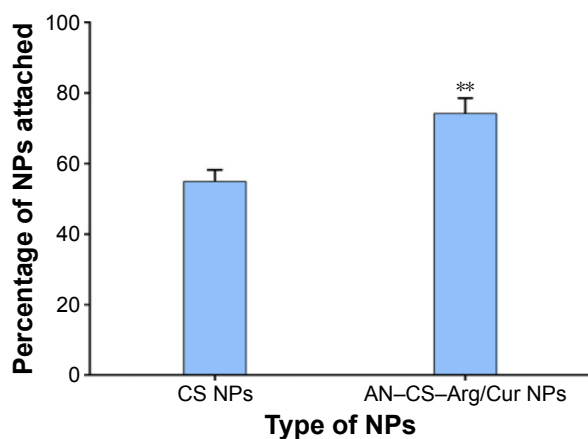


Figure 11 Percentage of attachment of NPs to excised colon of rat after 1-hour incubation at 37°C.

Notes: Data are expressed as mean \pm SD ($n=3$). ** $P<0.01$ compared with CS NPs.
Abbreviations: AN, acrylonitrile; Arg, arginine; CS, chitosan; Cur, curcumin; NPs, nanoparticles; SD, standard deviation.

NPs were bound to the colon, while $55.12\% \pm 3.20\%$ of control CS NPs were found attached with colon showing a difference of 19% ($P<0.01$). This result demonstrated that NPs based on amphiphilic CS derivative showed greater mucoadhesion compared with unmodified CS. This increased mucoadhesion might be due to combinational effects of CS, AN, Arg, and Cur as fabricated in the synthesized NPs. CS has been well reported to have mucoadhesive nature,^{36,37,90} AN may promote mucoadhesion by hydrophobic interactions as earlier reports demonstrate that acrylic acid-based derivatives such as poly acrylic acid and cyanoacrylate possess mucoadhesive properties.^{91–93} This mucoadhesive nature of two polymers was further augmented by Arg to increase the overall surface charge of NPs while Cur partly play its role via π - π interaction with mucus glycoproteins. The result is being in agreement with previously published report.⁹⁴

In vitro cytotoxicity study

Figure 12A shows the percentage of cell viability as a function of the concentration of the blank AN-CS-Arg NPs. Results showed that blank AN-CS-Arg NPs in all the studied concentrations have no effects on cell viability of HT-29 cells, suggesting blank AN-CS-Arg NPs might be good biocompatible carriers.

Figure 12B shows the percentage of cell viability as a function of the concentration of the Cur. In HT-29 cell lines, the IC_{50} (half maximal inhibitory concentration) of AN-CS-Arg/Cur NPs and Cur solution was $15 \mu\text{g/mL}$ and $28.50 \mu\text{g/mL}$, respectively. Results showed that for both Cur solution and AN-CS-Arg/Cur NPs, the antitumor activity increased the concentration dependently. For the concentration of Cur below $2.7 \mu\text{g/mL}$, no significant differences in antitumor activity were observed for two formulations. However, compared with Cur solution, the cytotoxicity of AN-CS-Arg/Cur NPs increased significantly ($P<0.05$) when the concentration of Cur was above $5 \mu\text{g/mL}$. The increased cytotoxic effect of the drug-loaded AN- and Arg-modified CS NPs could be attributed to the better penetration of the drug in the tumor cells due to the carrier (AN- and Arg-modified CS). Arg has been reported as a potential valuable tool for drug delivery due to its ability to improve the cellular uptake and enhance the permeation with minimal toxicity.^{95–97}

Intracellular uptake of Cur-loaded AN- and Arg-modified CS NPs

To further confirm the relationship between the anticancer activities and cell uptake, we observed the fluorescence of HT-29 cells after incubation for 3 hours at 37°C with free Cur and AN-CS-Arg/Cur NPs, respectively, at an equal concentration of $7.5 \mu\text{g/mL}$. As shown in Figure 13, it clearly

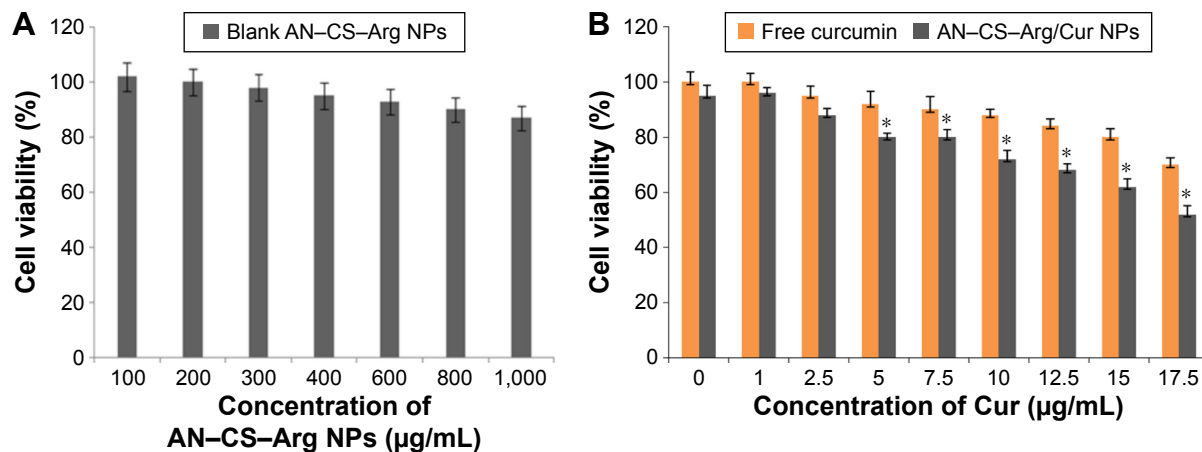


Figure 12 (A) Cytotoxicity of different concentrations of blank AN-CS-Arg NPs in HT-29 cells. **(B)** Cytotoxicity effect of different concentrations of AN-CS-Arg/Cur NPs in HT-29 cells. Data are expressed as mean \pm SD ($n=3$). * $P<0.05$ compared with free Cur solution.

Abbreviations: AN, acrylonitrile; Arg, arginine; CS, chitosan; Cur, curcumin; NPs, nanoparticles; SD, standard deviation.

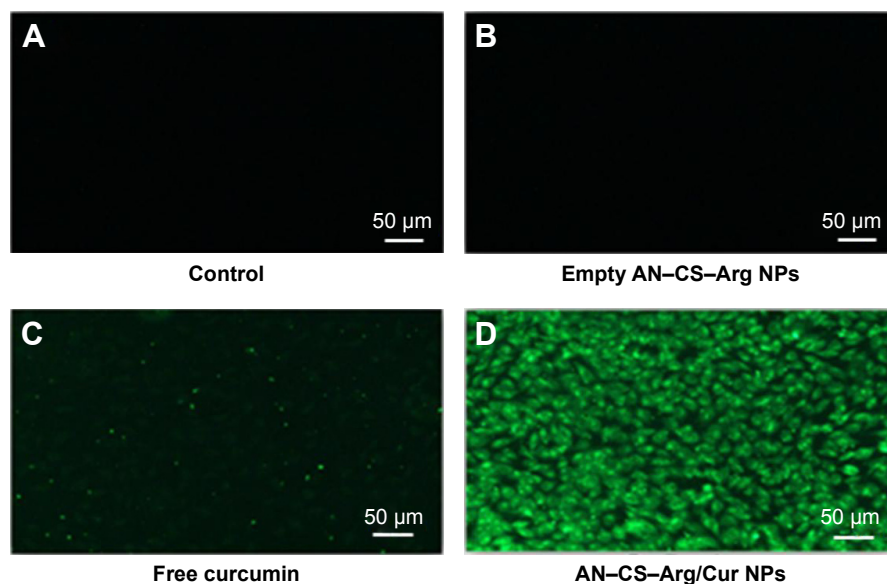


Figure 13 Fluorescence microscopic images of the uptake of HT-29 cells of free Cur and AN-CS-Arg/Cur NPs after incubation with 7.5 µg/mL at 37°C for 3 hours. **Notes:** (A) Fluorescence images of control cells (without any treatment); (B) Cells treated with blank AN-CS-Arg NPs; (C) Cells treated with free Cur solution; and (D) Cells treated with AN-CS-Arg/Cur Nps. Original magnification 100×. **Abbreviations:** AN, acrylonitrile; Arg, arginine; CS, chitosan; Cur, curcumin; NPs, nanoparticles.

observed that the cells treated with AN-CS-Arg/Cur NPs showed profound fluorescence intensity compared with cells treated with native Cur. The fluorescence of cells treated with native Cur is relatively faint. However, the fluorescence of HT-29 cells is greatly enhanced, and it can be clearly seen in the of HT-29 cells after treated with AN-CS-Arg/Cur NPs. This result showed that AN-CS-Arg/Cur NPs were more effective in the enhancement of the uptake of Cur. Our results are being in agreement with other Cur nanoparticulate formulations that were more effective than free Cur against the growth of cancer cell lines due to enhanced cell uptake.^{70,98}

Pharmacokinetics and in vivo oral bioavailability

The main pharmacokinetic parameters of free Cur and Cur-loaded AN- and Arg-modified CS NPs (AN-CS-Arg/Cur NPs) following the oral administration of 100 mg/kg in rats are listed in Table 3, and the mean plasma concentration-time curves are shown in Figure 14. As listed in Table 3, the peak plasma concentration (C_{max}) was significantly higher for NPs-treated group compared with the group treated with free

Cur ($P < 0.01$). Similarly, area under the concentration-time curve (area under the concentration-time curve [AUC]_{0–24 hours}) of the NPs-administered group was greater than the control group ($P < 0.01$). Overall, these results show that currently synthesized NPs have higher plasma concentration, lower clearance, and longer half-life compared with free Cur. The relative bioavailability of NPs as calculated by the following equation: $F_{rel} (\%) = (100 \times AUC_A \text{ dose}_B / AUC_B \text{ dose}_A)$ was 454%, indicating that the absorption of Cur increased when it was encapsulated in a nanoformulation. Similar results were also found in other Cur nanoformulations.^{99,100} This remarkable increase in absorption might be due to higher affinity between NPs and mucous membrane. In GIT, the membranes on epithelial cells are lipidic in nature causing better dispersibility of Cur, thereby prolonging residence time of drug in GIT. The mucoadhesive nature of CS resulted from interaction between positively charged CS and negatively charged mucin contributes to further prolong the contact time between CS and absorptive surface. This property of CS was further augmented in our NPs due to fabrication with Arg, which not only contribute to increase the surface charge of NPs

Table 3 Pharmacokinetic parameters of Cur and AN-CS-Arg/Cur NPs in mice plasma after oral administration

Formulations	Dose (mg/kg)	AUC _{0–24 hours} (mg/L hour)	MRT _{0–24} (hours)	CL/F (l/h per kg)	$t_{1/2}$ (hours)	T_{max} (hours)	C_{max} (mg/L)	F_{rel} (%)
Cur solution	100	65.12±7.42	4.33±0.67	0.34±0.29	0.90±0.89	1.00	11.00±1.17	–
AN-CS-Arg/Cur NPs	100	295.47±82.44**	3.40±0.75	0.72±0.80*	1.15±0.30*	2.00	99.72±30.47**	453.73

Notes: * $P < 0.05$; ** $P < 0.01$ vs Cur solution. Data are expressed as mean ± SD.

Abbreviations: Arg, arginine; AN, acrylonitrile; AUC, area under the curve; CS, chitosan; Cur, curcumin; NPs, nanoparticles; MRT, mean residence time; CL/F, apparent total clearance of the drug from plasma after oral administration.

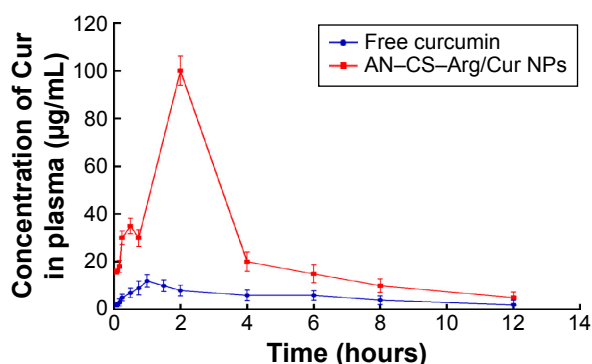


Figure 14 Plasma concentration–time profile of Cur in rats following the single oral administration with free Cur and AN–CS–Arg/Cur NPs at a dose of 100 mg/kg.

Note: Data are expressed as mean \pm SD ($n=6$).

Abbreviations: AN, acrylonitrile; Arg, arginine; CS, chitosan; Cur, curcumin; NPs, nanoparticles; SD, standard deviation.

but also enhance the permeation across the cell membrane. Thus, promoting absorption results in higher bioavailability. In addition, the carrier could shield Cur against metabolic degradation in GIT and in blood circulation before Cur is released from NPs, which could be a second reason attributed to the improvement of bioavailability.^{101,102} However, this is worth noticing that the bioavailability was still not high enough with regard to the administered dose that might be due to the production of metabolites during the process of absorption, metabolism, and biotransformation.⁸⁵

Conclusion

Our results demonstrated that AN- and Arg-modified CS conjugate (AN–CS–Arg) was successfully synthesized. Self-assembled NPs based on AN–CS–Arg conjugate were developed by sonication method. The aqueous solubility of Cur was significantly enhanced after encapsulation into AN- and Arg-modified CS NPs. The Cur-loaded AN- and Arg-modified CS NPs (AN–CS–Arg/Cur) showed the advantages of narrow size distribution, high drug encapsulation and LE, sustained release, and more importantly improved mucoadhesion. Furthermore, they exhibited enhanced cell uptake, superior growth inhibition against tumor cells, and improved oral bioavailability over native Cur. These preliminary studies indicated the potential of the currently prepared NPs in its possible applications to improve the solubility and anticancer activity of Cur. However, further studies in experimentally induced colitis animal models are needed to evaluate the efficacy of these NPs in the treatment of colorectal cancers and inflammatory bowel disease.

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Disclosure

The authors report no conflicts of interest in this work.

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