Acta Crystallographica Section E
Structure Reports Online
ISSN 1600-5368

Benzyl *N*-(1-{*N*′-[(*E*)-2,3-dihydroxybenzylidene]hydrazinecarbonyl}-2-hydroxyethyl)carbamate dihydrate

Solange M. S. V. Wardell,[a] Edward R. T. Tiekink,[b]* Marcus V. N. de Souza,[c] Alessandra C. Pinheiro[c] and James L. Wardell[d]‡

[a]CHEMSOL, 1 Harcourt Road, Aberdeen AB15 5NY, Scotland, [b]Department of Chemistry, University of Malaya, 50603 Kuala Lumpur, Malaysia, [c]Fundação Oswaldo Cruz, Instituto de Tecnologia, em Fármacos – Farmanguinhos, R. Sizenando Nabuco, 100, Manguinhos, 21041-250 Rio de Janeiro, RJ, Brazil, and [d]Centro de Desenvolvimento Tecnológico em Saúde (CDTS), Fundação Oswaldo Cruz (FIOCRUZ), Casa Amarela, Campus de Manguinhos, Av. Brasil 4365, 21040-900 Rio de Janeiro, RJ, Brazil
Correspondence e-mail: edward.tiekink@gmail.com

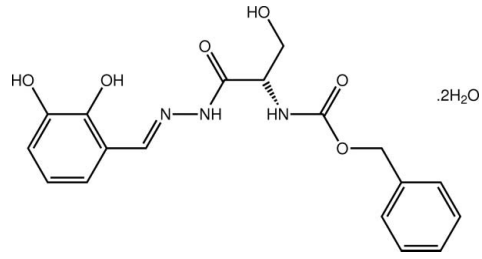
Received 20 July 2011; accepted 20 July 2011

Key indicators: single-crystal X-ray study; T = 120 K; mean σ(C–C) = 0.005 Å; R factor = 0.050; wR factor = 0.114; data-to-parameter ratio = 8.9.

The organic molecule in the title dihydrate, $C_{18}H_{19}N_3O_6\cdot 2H_2O$, adopts a twisted U-shape with the major twists evident about the chiral C atom [the C—N—C—C torsion angle is −88.2 (4) °] and about the oxygen–benzyl bond [C—O—C—C = 74.2 (4) °]. The conformation about the imine bond [1.290 (4) Å] is *E* and an intramolecular O—H···N hydrogen bond helps to establish the near coplanarity of the hydroxybenzene and hydrazine groups. The crystal packing features O—H···O and N—H···O hydrogen bonds, leading to two-dimensional supramolecular arrays in the *ab* plane with weak C—H···π connections between the arrays.

Related literature

For background to the use of L-serine derivatives in anti-tumour therapy, see: Jiao *et al.* (2009); Yakura *et al.* (2007). For background to *N*-acylhydrazone derivatives from L-serine for anti-tumour testing, see: de Souza *et al.* (2010, 2011); Pinheiro *et al.* (2010, 2011); Howie *et al.* (2011); Tiekink *et al.* (2011); Wardell *et al.* (2011).



Experimental

Crystal data

$C_{18}H_{19}N_3O_6\cdot 2H_2O$
M_r = 409.39
Orthorhombic, $P2_12_12_1$
a = 4.7570 (2) Å
b = 13.1011 (4) Å
c = 30.5511 (9) Å
V = 1904.00 (11) Å^3
Z = 4
Mo $K\alpha$ radiation
μ = 0.11 mm^{-1}
T = 120 K
0.18 × 0.12 × 0.10 mm

Data collection

Bruker–Nonius Roper CCD camera on κ-goniostat diffractometer
Absorption correction: multi-scan (*SADABS*; Sheldrick, 2007)
T_{min} = 0.887, T_{max} = 1.000
12958 measured reflections
2560 independent reflections
1984 reflections with $I > 2\sigma(I)$
R_{int} = 0.074

Refinement

$R[F^2 > 2\sigma(F^2)]$ = 0.050
$wR(F^2)$ = 0.114
S = 1.11
2560 reflections
289 parameters
9 restraints
H atoms treated by a mixture of independent and constrained refinement
$\Delta\rho_{max}$ = 0.23 e Å^{-3}
$\Delta\rho_{min}$ = −0.26 e Å^{-3}

Table 1
Hydrogen-bond geometry (Å, °).

D—H···A	D—H	H···A	D···A	D—H···A
O1—H1o···N1	0.84 (2)	1.88 (2)	2.604 (3)	143 (3)
O2—H2o···O1w^{i}	0.84 (2)	1.79 (2)	2.625 (3)	170 (3)
O4—H4o···O1^{ii}	0.82 (3)	1.97 (3)	2.791 (3)	176 (3)
O1w—H1w···O2^{iii}	0.85 (2)	2.05 (2)	2.894 (3)	169 (3)
O1w—H2w···O2w^{ii}	0.84 (2)	2.04 (2)	2.879 (3)	176 (2)
O2w—H3w···O3	0.85 (2)	2.38 (2)	3.188 (3)	160 (3)
O2w—H4w···O3^{iv}	0.86 (3)	1.97 (2)	2.818 (3)	168 (3)
N2—H2n···O5^{iv}	0.87 (2)	2.08 (2)	2.892 (3)	154 (2)
N2—H2n···N3	0.87 (2)	2.34 (2)	2.705 (3)	106 (2)
N3—H3n···O2w^{ii}	0.85 (3)	2.28 (3)	3.078 (3)	157 (3)
C18—H18···$Cg1^{v}$	0.95	2.94	3.700 (3)	138

Symmetry codes: (i) $x, y-1, z$; (ii) $-x+1, y+\frac{1}{2}, -z+\frac{1}{2}$; (iii) $x-1, y+1, z$; (iv) $x+1, y, z$; (v) $x-\frac{1}{2}, -y+\frac{3}{2}, -z+1$.

Data collection: *COLLECT* (Hooft, 1998); cell refinement: *DENZO* (Otwinowski & Minor, 1997) and *COLLECT*; data reduction: *DENZO* and *COLLECT*; program(s) used to solve structure: *SHELXS97* (Sheldrick, 2008); program(s) used to refine structure: *SHELXL97* (Sheldrick, 2008); molecular graphics: *ORTEP-3* (Farrugia, 1997) and *DIAMOND* (Brandenburg, 2006); software used to prepare material for publication: *publCIF* (Westrip, 2010).

The use of the EPSRC X-ray crystallographic service at the University of Southampton, England, and the valuable assistance of the staff there is gratefully acknowledged. JLW acknowledges support from CAPES (Brazil).

‡ Additional correspondence author, e-mail: j.wardell@abdn.ac.uk.

Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: HB6326).

References

Brandenburg, K. (2006). *DIAMOND*. Crystal Impact GbR, Bonn, Germany.

Farrugia, L. J. (1997). *J. Appl. Cryst.* **30**, 565.

Hooft, R. W. W. (1998). *COLLECT*. Nonius BV, Delft, The Netherlands.

Howie, R. A., de Souza, M. V. N., Pinheiro, A. C., Kaiser, C. R., Wardell, J. L. & Wardell, S. M. S. V. (2011). *Z. Kristallogr.* **226**, 483–491.

Jiao, X., Wang, L., Xiao, Q., Xie, P. & Liang, X. (2009). *J. Asian Nat. Prod. Res.* **11**, 274–280.

Otwinowski, Z. & Minor, W. (1997). *Methods in Enzymology*, Vol. 276, *Macromolecular Crystallography*, Part A, edited by C. W. Carter Jr & R. M. Sweet, pp. 307–326. New York: Academic Press.

Pinheiro, A. C., de Souza, M. V. N., Tiekink, E. R. T., Wardell, J. L. & Wardell, S. M. S. V. (2010). *Acta Cryst.* E**66**, o1004–o1005.

Pinheiro, A. C., de Souza, M. V. N., Tiekink, E. R. T., Wardell, S. M. S. V. & Wardell, J. L. (2011). *Acta Cryst.* E**67**, o581–o582.

Sheldrick, G. M. (2007). *SADABS*. Bruker AXS Inc., Madison, Wisconsin, USA.

Sheldrick, G. M. (2008). *Acta Cryst.* A**64**, 112–122.

Souza, M. V. N. de, Pinheiro, A. C., Tiekink, E. R. T., Wardell, S. M. S. V. & Wardell, J. L. (2010). *Acta Cryst.* E**66**, o3253–o3254.

Souza, M. V. N. de, Pinheiro, A. C., Tiekink, E. R. T., Wardell, S. M. S. V. & Wardell, J. L. (2011). *Acta Cryst.* E**67**, o1868–o1869.

Tiekink, E. R. T., de Souza, M. V. N., Pinheiro, A. C., Wardell, S. M. S. V. & Wardell, J. L. (2011). *Acta Cryst.* E**67**, o1866–o1867.

Wardell, J. L., de Souza, M. V. N., Pinheiro, A. C., Tiekink, E. R. T. & Wardell, S. M. S. V. (2011). *Acta Cryst.* E**67**, o1888–o1889.

Westrip, S. P. (2010). *J. Appl. Cryst.* **43**, 920–925.

Yakura, T., Yoshimoto, Y., Ishida, C. & Mabuchi, S. (2007). *Tetrahedron*, **63**, 4429–4438.

supplementary materials

Acta Cryst. (2011). E**67**, o2155-o2156 [**doi:10.1107/S1600536811029370**]

Benzyl *N*-(1-{*N*'-[(*E*)-2,3-dihydroxybenzylidene]hydrazinecarbonyl}-2-hydroxyethyl)carbamate dihydrate

S. M. S. V. Wardell, E. R. T. Tiekink, M. V. N. de Souza, A. C. Pinheiro and J. L. Wardell

Comment

The motivation for investigations of molecules related to the title compound, (I), arise from the anti-tumour activity of *L*-serine derivatives (Jiao *et al.*, 2009; Yakura *et al.*, 2007) and, in particular, the development of *N*-acylhydrazone derivatives from *L*-serine for use in anti-tumour testing (Pinheiro *et al.*, 2010; de Souza *et al.*, 2010; Pinheiro *et al.*, 2011: Howie *et al.*, 2011; de Souza *et al.*, (2011); Howie *et al.* (2011); Tiekink *et al.* (2011); Wardell *et al.* (2011).

The crystallographic asymmetric unit of (I) comprises an organic molecule and two water molecules of crystallization. While the absolute structure could not be determined experimentally, the assignment of the *S*-configuration at the C9 atom is based on a starting reagent, *L*-serine. Overall, the organic molecule has a twisted U-shape with the two benzene rings lying to the same side of the molecule. Twists are evident about the chiral centre [the C11—N3—C9—C8 torsion angle is -88.2 (4) °] and about the benzyl group [C11—O6—C12—C13 is 74.2 (4) °]. The co-planarity at the hydroxylbenzene/hydrazine residue arises as a result of an intramolecular O1—H···N1 hydrogen bond (Table 1). The conformation about the N1═C7 imine bond [1.290 (4) Å] is *E*.

In the crystal, each of the acidic hydrogen atoms forms a significant hydrogen bond (Table 1). Thus, the O2- and O4-hydroxy groups form hydrogen bonds to the water-O1w and hydroxy-O1 atoms, respectively. The O1w-water molecule forms hydrogen bonds to the hydroxy-O2 and water-O2w atoms, while the O2w-water molecule forms connections to symmetry related carbonyl-O3 atoms, implying the latter is bifurcated. Finally, amine-N2—H is connected to carbonyl-O5, and amine-N3—H is connected to water-O2w. The result of the hydrogen bonding scheme is the formation of layers in the *ab* plane, Fig. 2. Connections between layers that stack along the *c* axis are of the form C—H···π, Table 1.

Experimental

To a stirred solution of methyl (2*S*)-2-[(benzyloxycarbonyl)amino]-3-hydroxypropanoate (0.3 g, 1.17 mmol), prepared from (2*S*)-2-amino-3-hydroxypropanoate hydrochloride and benzyl chloroformate (21 ml, 0.15 mol), in ethanol (10 ml) was added N_2H_4.H2O (80%, 5.5 mmol). The reaction mixture was stirred for 24 h at room temperature, rotary evaporated and the residue washed with cold ethanol (3 *x* 10 ml) to give benzyl (1*S*)-2-hydrazino-1-(hydroxymethyl)-2-oxoethylcarbamate in 78% yield, which was used as such for the next stage. To a stirred solution of (*S*)-$PhCH_2OCONHCH(CH_2OH)CONHNH_2$ (1.0 mmol) in ethanol (10 ml) at room temperature was added 2, 3-dihydroxybenzaldehyde (1.05 mmol). The reaction mixture was refluxed for 4 h, rotary evaporated and the residue purified by washing with cold ethanol (3 *x* 10 ml), affording the title compound, *M*.pt. 423 K, yield 81%. The sample for the structure determination was recrystallized from EtOH as pale-brown blocks of the dihydrate. The water molecules were presumably absorbed from the atmosphere. ^{1}H NMR (500 MHz, DMSO-d6) δ (p.p.m.): 11.76 (1*H*, s, NHN), 10.95 (1*H*, s, C1—OH or C2—OH), 9.21 (1*H*, s, C1—OH or C2—OH), 8.41 (1*H*, s, N=CH),7.44 (1*H*, d, J= 7.4, NHCH), 7.40–7.20 (5*H*, m, Ph), 6.94 (1*H*, d, J= 7.8, H6), 6.85–6.80 (1*H*, m, H4), 6.73 (1*H*, t, J= 7.8, H5), 5.04 (3*H*, m, CH_2Ph and OH), 4.15 (1*H*, m, CH), 3.75–3.55 (2*H*, m, CH_2OH). ^{13}C NMR (125 MHz,

DMSO-d6) δ (p.p.m.): 171.1, 156.0, 145.6, 145.2, 141.6, 136.9, 128.4, 127.9, 127.8, 127.7, 120.0, 119.2, 117.4, 116.5, 65.4, 61.4, 56.3. IR (cm^{-1}, KBr): 3270 (O—H), 1676 (COCH and COO). MS/ESI: [M—H]: 372.3.

Refinement

The C-bound H atoms were geometrically placed (C–H = 0.95–1.00 Å) and refined as riding with U_{iso}(H) = 1.2–1.5U_{eq}(C). The O– and N-bound H atoms were located from a difference map and refined with the distance restraints O–H = 0.84 ± 0.01 and N–H = 0.86±0.01 Å, and with U_{iso}(H) = zU_{eq}(carrier atom); z = 1.5 for O and z = 1.2 for N. In the absence of significant anomalous scattering effects, 1782 Friedel pairs were averaged in the final refinement. However, the absolute configuration was assigned on the basis of the chirality of the *L*-serine starting material.

Figures

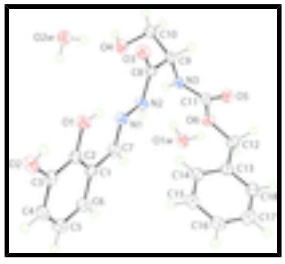


Fig. 1. The molecular structure of (I) showing displacement ellipsoids at the 50% probability level.

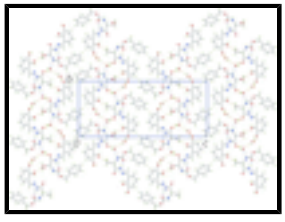


Fig. 2. A view in projection down the *a* axis of the stacking of 2-D supramolecular arrays in the *ab* plane in (I), and with the O—H···O and N—H···O hydrogen bonding shown as orange and blue dashed lines, respectively.

Benzyl *N*-(1-{*N*'-[(*E*)-2,3- dihydroxybenzylidene]hydrazinecarbonyl}-2-hydroxyethyl)carbamate dihydrate

Crystal data

$C_{18}H_{19}N_3O_6·2H_2O$	F(000) = 864
M_r = 409.39	D_x = 1.428 Mg m^{-3}
Orthorhombic, $P2_12_12_1$	Mo $K\alpha$ radiation, λ = 0.71073 Å
Hall symbol: P 2ac 2ab	Cell parameters from 9601 reflections
a = 4.7570 (2) Å	θ = 2.9–27.5°
b = 13.1011 (4) Å	μ = 0.11 mm^{-1}
c = 30.5511 (9) Å	T = 120 K
V = 1904.00 (11) $Å^3$	Block, pale-brown
Z = 4	0.18 × 0.12 × 0.10 mm

Data collection

Bruker–Nonius Roper CCD camera on κ-goniostat diffractometer	2560 independent reflections

Radiation source: Bruker–Nonius FR591 rotating anode	1984 reflections with $I > 2\sigma(I)$
graphite	$R_{int} = 0.074$
Detector resolution: 9.091 pixels mm^{-1}	$\theta_{max} = 27.5°$, $\theta_{min} = 3.1°$
φ and ω scans	$h = -6\rightarrow5$
Absorption correction: multi-scan (*SADABS*; Sheldrick, 2007)	$k = -17\rightarrow16$
$T_{min} = 0.887$, $T_{max} = 1.000$	$l = -37\rightarrow39$
12958 measured reflections	

Refinement

Refinement on F^2	Primary atom site location: structure-invariant direct methods
Least-squares matrix: full	Secondary atom site location: difference Fourier map
$R[F^2 > 2\sigma(F^2)] = 0.050$	Hydrogen site location: inferred from neighbouring sites
$wR(F^2) = 0.114$	H atoms treated by a mixture of independent and constrained refinement
$S = 1.11$	$w = 1/[\sigma^2(F_o^2) + (0.034P)^2 + 0.4778P]$ where $P = (F_o^2 + 2F_c^2)/3$
2560 reflections	$(\Delta/\sigma)_{max} < 0.001$
289 parameters	$\Delta\rho_{max} = 0.23$ e Å^{-3}
9 restraints	$\Delta\rho_{min} = -0.26$ e Å^{-3}

Special details

Geometry. All s.u.'s (except the s.u. in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell s.u.'s are taken into account individually in the estimation of s.u.'s in distances, angles and torsion angles; correlations between s.u.'s in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell s.u.'s is used for estimating s.u.'s involving l.s. planes.

Refinement. Refinement of F^2 against ALL reflections. The weighted *R*-factor *wR* and goodness of fit *S* are based on F^2, conventional *R*-factors *R* are based on *F*, with *F* set to zero for negative F^2. The threshold expression of $F^2 > 2\sigma(F^2)$ is used only for calculating *R*-factors(gt) *etc*. and is not relevant to the choice of reflections for refinement. *R*-factors based on F^2 are statistically about twice as large as those based on *F*, and *R*- factors based on ALL data will be even larger.

Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters (Å^2)

	x	*y*	*z*	U_{iso}*/U_{eq}
O1	0.5296 (5)	0.14889 (15)	0.34576 (7)	0.0256 (5)
H1O	0.435 (7)	0.2000 (19)	0.3378 (10)	0.038*
O2	0.8776 (5)	0.00115 (16)	0.38088 (8)	0.0289 (6)
H2O	0.719 (4)	−0.022 (3)	0.3735 (12)	0.043*
O3	0.0260 (5)	0.33623 (15)	0.27559 (7)	0.0288 (6)
O4	0.3510 (5)	0.54140 (18)	0.23084 (7)	0.0293 (6)
H4O	0.391 (10)	0.571 (3)	0.2077 (11)	0.044*
O5	−0.3437 (5)	0.58789 (16)	0.34957 (7)	0.0246 (5)

O6	−0.0248 (5)	0.71487 (15)	0.35850 (6)	0.0249 (6)
N1	0.4021 (7)	0.34210 (18)	0.34263 (8)	0.0222 (6)
N2	0.2671 (6)	0.42717 (19)	0.32640 (8)	0.0225 (6)
H2N	0.337 (7)	0.4863 (14)	0.3341 (10)	0.027*
N3	0.0278 (6)	0.60659 (19)	0.30326 (8)	0.0218 (6)
H3N	0.172 (9)	0.642 (2)	0.2984 (10)	0.026*
C1	0.7354 (7)	0.2756 (2)	0.39369 (9)	0.0208 (7)
C2	0.7133 (7)	0.1748 (2)	0.37819 (10)	0.0204 (7)
C3	0.8820 (8)	0.0997 (2)	0.39619 (10)	0.0230 (7)
C4	1.0684 (8)	0.1236 (2)	0.42946 (10)	0.0264 (8)
H4	1.1846	0.0716	0.4414	0.032*
C5	1.0887 (9)	0.2223 (2)	0.44573 (10)	0.0283 (8)
H5	1.2161	0.2376	0.4688	0.034*
C6	0.9210 (8)	0.2979 (2)	0.42797 (9)	0.0257 (8)
H6	0.9318	0.3655	0.4391	0.031*
C7	0.5738 (8)	0.3584 (2)	0.37454 (9)	0.0228 (7)
H7	0.5949	0.4257	0.3857	0.027*
C8	0.0844 (8)	0.4187 (2)	0.29279 (9)	0.0213 (7)
C9	−0.0437 (8)	0.5184 (2)	0.27644 (9)	0.0213 (7)
H9	−0.2529	0.5109	0.2763	0.026*
C10	0.0533 (8)	0.5356 (3)	0.22979 (10)	0.0253 (8)
H10A	−0.0080	0.4785	0.2109	0.030*
H10B	−0.0269	0.5998	0.2181	0.030*
C11	−0.1316 (8)	0.6322 (2)	0.33806 (10)	0.0213 (7)
C12	−0.1522 (9)	0.7377 (2)	0.40029 (9)	0.0267 (8)
H12A	−0.3592	0.7348	0.3973	0.032*
H12B	−0.1009	0.8081	0.4090	0.032*
C13	−0.0615 (8)	0.6646 (2)	0.43599 (10)	0.0240 (7)
C14	0.1605 (8)	0.5972 (2)	0.43100 (10)	0.0286 (8)
H14	0.2612	0.5948	0.4042	0.034*
C15	0.2360 (8)	0.5330 (3)	0.46517 (10)	0.0305 (8)
H15	0.3859	0.4859	0.4613	0.037*
C16	0.0959 (9)	0.5366 (3)	0.50491 (10)	0.0299 (8)
H16	0.1499	0.4931	0.5283	0.036*
C17	−0.1231 (9)	0.6044 (3)	0.50978 (11)	0.0342 (9)
H17	−0.2201	0.6079	0.5369	0.041*
C18	−0.2039 (8)	0.6676 (3)	0.47569 (10)	0.0295 (8)
H18	−0.3573	0.7132	0.4795	0.035*
O1W	0.4103 (6)	0.91059 (18)	0.35680 (8)	0.0343 (6)
H1W	0.248 (4)	0.936 (3)	0.3598 (11)	0.051*
H2W	0.416 (8)	0.876 (2)	0.3336 (7)	0.051*
O2W	0.5751 (6)	0.28389 (19)	0.22018 (7)	0.0310 (6)
H3W	0.420 (4)	0.282 (3)	0.2339 (10)	0.047*
H4W	0.695 (6)	0.303 (3)	0.2392 (9)	0.047*

Atomic displacement parameters (Å^2)

	U^{11}	U^{22}	U^{33}	U^{12}	U^{13}	U^{23}

O1	0.0249 (15)	0.0216 (11)	0.0302 (12)	0.0024 (10)	−0.0065 (11)	−0.0023 (10)
O2	0.0241 (14)	0.0196 (12)	0.0430 (13)	−0.0005 (12)	−0.0046 (12)	0.0010 (10)
O3	0.0296 (15)	0.0218 (11)	0.0350 (12)	−0.0035 (12)	−0.0039 (12)	−0.0053 (10)
O4	0.0240 (14)	0.0372 (14)	0.0268 (12)	−0.0026 (12)	0.0030 (12)	0.0055 (11)
O5	0.0228 (14)	0.0230 (11)	0.0280 (11)	−0.0033 (11)	0.0013 (11)	−0.0011 (10)
O6	0.0295 (15)	0.0199 (11)	0.0254 (11)	−0.0061 (11)	0.0002 (11)	−0.0010 (9)
N1	0.0241 (16)	0.0187 (13)	0.0237 (13)	0.0026 (13)	−0.0005 (13)	0.0024 (11)
N2	0.0248 (17)	0.0155 (13)	0.0273 (13)	0.0008 (13)	−0.0066 (13)	0.0003 (11)
N3	0.0163 (16)	0.0215 (14)	0.0276 (14)	−0.0010 (13)	−0.0002 (13)	0.0001 (11)
C1	0.0201 (18)	0.0213 (16)	0.0210 (14)	−0.0004 (15)	0.0020 (15)	0.0014 (13)
C2	0.0199 (18)	0.0199 (16)	0.0214 (15)	−0.0011 (15)	0.0019 (15)	−0.0001 (13)
C3	0.0212 (19)	0.0184 (15)	0.0293 (16)	0.0006 (15)	0.0038 (15)	0.0021 (13)
C4	0.0189 (19)	0.0291 (17)	0.0314 (17)	0.0024 (16)	−0.0020 (16)	0.0060 (14)
C5	0.027 (2)	0.0305 (17)	0.0274 (16)	−0.0039 (18)	−0.0052 (16)	0.0053 (14)
C6	0.026 (2)	0.0254 (16)	0.0255 (15)	−0.0048 (17)	0.0005 (16)	0.0014 (14)
C7	0.0238 (19)	0.0212 (16)	0.0235 (15)	0.0002 (15)	0.0015 (16)	0.0008 (13)
C8	0.0195 (18)	0.0248 (16)	0.0195 (14)	−0.0026 (16)	0.0002 (14)	−0.0012 (13)
C9	0.0191 (18)	0.0201 (14)	0.0248 (15)	0.0008 (15)	0.0003 (15)	−0.0009 (12)
C10	0.023 (2)	0.0295 (17)	0.0232 (15)	−0.0018 (16)	−0.0027 (16)	0.0013 (14)
C11	0.024 (2)	0.0158 (14)	0.0240 (16)	−0.0003 (15)	−0.0037 (15)	0.0018 (12)
C12	0.034 (2)	0.0211 (16)	0.0247 (15)	−0.0012 (16)	0.0032 (16)	−0.0017 (13)
C13	0.0264 (19)	0.0186 (15)	0.0269 (16)	−0.0055 (16)	−0.0033 (15)	−0.0019 (13)
C14	0.026 (2)	0.0315 (18)	0.0279 (17)	−0.0005 (17)	−0.0006 (16)	−0.0013 (15)
C15	0.026 (2)	0.0289 (18)	0.0366 (19)	0.0015 (17)	−0.0031 (17)	0.0013 (16)
C16	0.030 (2)	0.0300 (18)	0.0299 (17)	−0.0053 (18)	−0.0082 (18)	0.0059 (15)
C17	0.042 (2)	0.0332 (19)	0.0276 (17)	−0.003 (2)	0.0038 (18)	−0.0001 (15)
C18	0.029 (2)	0.0257 (17)	0.0337 (18)	0.0031 (17)	0.0026 (17)	−0.0021 (15)
O1W	0.0290 (15)	0.0288 (13)	0.0451 (14)	0.0034 (13)	−0.0043 (13)	−0.0095 (11)
O2W	0.0283 (15)	0.0302 (12)	0.0346 (13)	0.0016 (14)	−0.0016 (12)	−0.0039 (11)

Geometric parameters (Å, °)

O1—C2	1.364 (4)	C6—H6	0.9500
O1—H1O	0.842 (10)	C7—H7	0.9500
O2—C3	1.374 (4)	C8—C9	1.526 (4)
O2—H2O	0.843 (10)	C9—C10	1.515 (4)
O3—C8	1.233 (3)	C9—H9	1.0000
O4—C10	1.418 (4)	C10—H10A	0.9900
O4—H4O	0.83 (4)	C10—H10B	0.9900
O5—C11	1.216 (4)	C12—C13	1.515 (4)
O6—C11	1.350 (4)	C12—H12A	0.9900
O6—C12	1.445 (3)	C12—H12B	0.9900
N1—C7	1.290 (4)	C13—C14	1.385 (5)
N1—N2	1.379 (3)	C13—C18	1.390 (4)
N2—C8	1.350 (4)	C14—C15	1.388 (4)
N2—H2N	0.875 (10)	C14—H14	0.9500
N3—C11	1.349 (4)	C15—C16	1.386 (5)
N3—C9	1.456 (4)	C15—H15	0.9500
N3—H3N	0.84 (4)	C16—C17	1.378 (5)

C1—C6	1.401 (4)	C16—H16	0.9500
C1—C2	1.407 (4)	C17—C18	1.385 (5)
C1—C7	1.453 (4)	C17—H17	0.9500
C2—C3	1.383 (4)	C18—H18	0.9500
C3—C4	1.385 (5)	O1W—H1W	0.851 (10)
C4—C5	1.388 (4)	O1W—H2W	0.845 (10)
C4—H4	0.9500	O2W—H3W	0.848 (10)
C5—C6	1.383 (5)	O2W—H4W	0.852 (10)
C5—H5	0.9500		
C2—O1—H1O	111 (3)	N3—C9—H9	108.3
C3—O2—H2O	116 (3)	C10—C9—H9	108.3
C10—O4—H4O	104 (3)	C8—C9—H9	108.3
C11—O6—C12	114.7 (3)	O4—C10—C9	106.9 (3)
C7—N1—N2	115.6 (2)	O4—C10—H10A	110.3
C8—N2—N1	120.5 (2)	C9—C10—H10A	110.3
C8—N2—H2N	121 (2)	O4—C10—H10B	110.3
N1—N2—H2N	116 (2)	C9—C10—H10B	110.3
C11—N3—C9	120.6 (3)	H10A—C10—H10B	108.6
C11—N3—H3N	118 (2)	O5—C11—N3	125.2 (3)
C9—N3—H3N	122 (2)	O5—C11—O6	124.2 (3)
C6—C1—C2	119.6 (3)	N3—C11—O6	110.7 (3)
C6—C1—C7	118.6 (3)	O6—C12—C13	112.7 (3)
C2—C1—C7	121.8 (3)	O6—C12—H12A	109.1
O1—C2—C3	118.9 (3)	C13—C12—H12A	109.1
O1—C2—C1	121.7 (3)	O6—C12—H12B	109.1
C3—C2—C1	119.4 (3)	C13—C12—H12B	109.1
O2—C3—C4	118.2 (3)	H12A—C12—H12B	107.8
O2—C3—C2	121.6 (3)	C14—C13—C18	119.1 (3)
C4—C3—C2	120.2 (3)	C14—C13—C12	122.8 (3)
C3—C4—C5	121.2 (3)	C18—C13—C12	118.1 (3)
C3—C4—H4	119.4	C13—C14—C15	120.1 (3)
C5—C4—H4	119.4	C13—C14—H14	120.0
C6—C5—C4	119.1 (3)	C15—C14—H14	120.0
C6—C5—H5	120.4	C16—C15—C14	120.9 (3)
C4—C5—H5	120.4	C16—C15—H15	119.5
C5—C6—C1	120.5 (3)	C14—C15—H15	119.5
C5—C6—H6	119.7	C17—C16—C15	118.7 (3)
C1—C6—H6	119.7	C17—C16—H16	120.6
N1—C7—C1	121.0 (3)	C15—C16—H16	120.6
N1—C7—H7	119.5	C16—C17—C18	121.0 (3)
C1—C7—H7	119.5	C16—C17—H17	119.5
O3—C8—N2	122.8 (3)	C18—C17—H17	119.5
O3—C8—C9	121.4 (3)	C17—C18—C13	120.3 (3)
N2—C8—C9	115.8 (3)	C17—C18—H18	119.9
N3—C9—C10	109.9 (3)	C13—C18—H18	119.9
N3—C9—C8	113.7 (2)	H1W—O1W—H2W	109 (2)
C10—C9—C8	108.3 (3)	H3W—O2W—H4W	105 (2)
C7—N1—N2—C8	179.6 (3)	O3—C8—C9—N3	174.4 (3)

C6—C1—C2—O1	178.0 (3)	N2—C8—C9—N3	−6.7 (4)
C7—C1—C2—O1	−3.1 (5)	O3—C8—C9—C10	−63.2 (4)
C6—C1—C2—C3	−2.0 (5)	N2—C8—C9—C10	115.7 (3)
C7—C1—C2—C3	176.9 (3)	N3—C9—C10—O4	64.5 (3)
O1—C2—C3—O2	3.0 (5)	C8—C9—C10—O4	−60.2 (3)
C1—C2—C3—O2	−177.0 (3)	C9—N3—C11—O5	−1.4 (5)
O1—C2—C3—C4	−179.4 (3)	C9—N3—C11—O6	179.0 (3)
C1—C2—C3—C4	0.7 (5)	C12—O6—C11—O5	10.5 (4)
O2—C3—C4—C5	178.4 (3)	C12—O6—C11—N3	−169.9 (3)
C2—C3—C4—C5	0.7 (5)	C11—O6—C12—C13	74.2 (4)
C3—C4—C5—C6	−0.6 (6)	O6—C12—C13—C14	12.3 (5)
C4—C5—C6—C1	−0.7 (5)	O6—C12—C13—C18	−169.4 (3)
C2—C1—C6—C5	2.0 (5)	C18—C13—C14—C15	0.8 (5)
C7—C1—C6—C5	−176.9 (3)	C12—C13—C14—C15	179.1 (3)
N2—N1—C7—C1	−178.2 (3)	C13—C14—C15—C16	−1.4 (5)
C6—C1—C7—N1	179.3 (3)	C14—C15—C16—C17	0.8 (5)
C2—C1—C7—N1	0.4 (5)	C15—C16—C17—C18	0.4 (5)
N1—N2—C8—O3	1.2 (5)	C16—C17—C18—C13	−1.0 (5)
N1—N2—C8—C9	−177.7 (3)	C14—C13—C18—C17	0.4 (5)
C11—N3—C9—C10	150.3 (3)	C12—C13—C18—C17	−178.0 (3)
C11—N3—C9—C8	−88.2 (4)		

Hydrogen-bond geometry (Å, °)

D—H···*A*	*D*—H	H···*A*	*D*···*A*	*D*—H···*A*
O1—H1o···N1	0.84 (2)	1.88 (2)	2.604 (3)	143 (3)
O2—H2o···O1w^{i}	0.84 (2)	1.79 (2)	2.625 (3)	170 (3)
O4—H4o···O1^{ii}	0.82 (3)	1.97 (3)	2.791 (3)	176 (3)
O1w—H1w···O2^{iii}	0.85 (2)	2.05 (2)	2.894 (3)	169 (3)
O1w—H2w···O2w^{ii}	0.84 (2)	2.04 (2)	2.879 (3)	176 (2)
O2w—H3w···O3	0.85 (2)	2.38 (2)	3.188 (3)	160 (3)
O2w—H4w···O3^{iv}	0.86 (3)	1.97 (2)	2.818 (3)	168 (3)
N2—H2n···O5^{iv}	0.874 (19)	2.08 (2)	2.892 (3)	154 (2)
N2—H2n···N3	0.874 (19)	2.34 (2)	2.705 (3)	106 (2)
N3—H3n···O2w^{ii}	0.85 (3)	2.28 (3)	3.078 (3)	157 (3)
C18—H18···Cg1^{v}	0.95	2.94	3.700 (3)	138

Symmetry codes: (i) *x*, *y*−1, *z*; (ii) −*x*+1, *y*+1/2, −*z*+1/2; (iii) *x*−1, *y*+1, *z*; (iv) *x*+1, *y*, *z*; (v) *x*−1/2, −*y*+3/2, −*z*+1.

Fig. 1

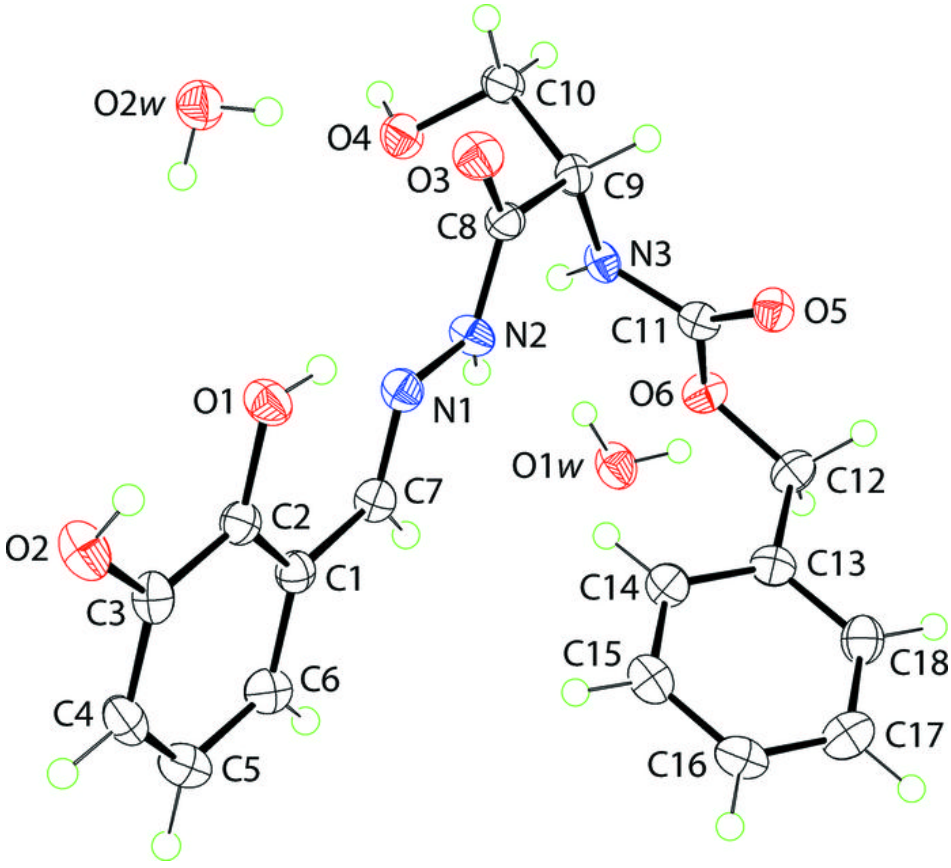

O2w
O4
C10
O3
C9
C8
N3
O5
C11
N2
O1
N1
O6
O1w
C12
C7
C2
O2
C1
C13
C3
C14
C15
C18
C6
C4
C17
C5
C16

Fig. 2